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Silica gel-induced aryne generation from *o*-triazenylarylboronic acids as stable solid precursors†

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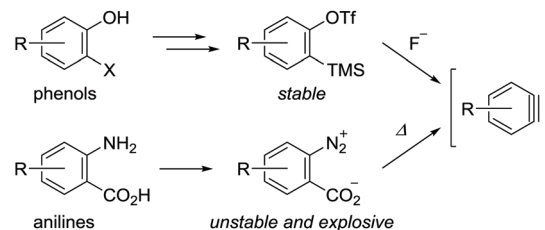
We report the development of *o*-triazenylarylboronic acids as new aniline-based aryne precursors. The readily available and shelf-stable solid precursors generate (hetero)arynes under remarkably mild conditions using silica gel as the sole reagent, which subsequently undergo reactions with a range of aryne-philic. Furthermore, solid-state aryne reactions under solvent-free conditions were accomplished. Aryne generation proceeded *via* a dual activation mechanism, as rationalized using Jaffé's plot analysis based on Hammett constants.

Introduction

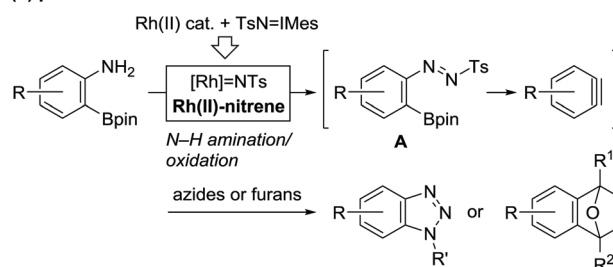
Arynes and heteroarynes are highly reactive synthetically useful reaction intermediates that enable the simultaneous creation of two bonds, including C–C, C–H, and C–X (X = heteroatom) bonds, on adjacent aromatic carbons *via* reactions with a range of aryne-philic.^{1–3} Because unstable arynes are typically generated *in situ*, a judicious choice of precursors that generate aryne under conditions compatible with the selected aryne-philic is crucial for achieving the desired transformations. Over the last few decades, the use of 2-trimethylsilylphenyl triflates⁴ in combination with fluoride ions has significantly contributed to the advancement of aryne chemistry, including the expansion of the aryne-philic scope, elegant reaction design, and syntheses of functional and biologically active compounds (Scheme 1a). These achievements are attributed to the stability and accessibility of the precursors, obtained from ubiquitous phenols, as well as the use of mild reaction conditions. In addition to phenol derivatives, aniline derivatives, represented by benzenediazonium 2-carboxylates,⁵ have been used as aryne precursors since the 1960s.⁵ However, despite their latent synthetic utility associated with the ubiquity of anilines, being comparable to that of phenols, their use is presently limited owing to their explosive character.⁶ In this context, we envisioned that the development of a methodology for aryne generation from readily available and stable aniline derivatives under mild conditions would con-

tribute to further advancements of aryne chemistry, which would be distinct from that achieved hitherto with phenol-based precursors.⁷

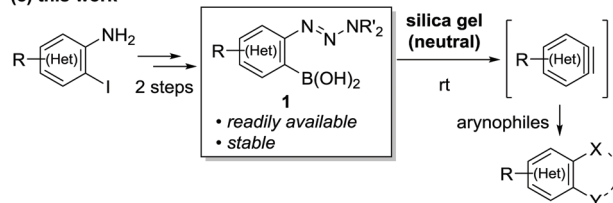
(a) representative aryne precursors



(b) previous work



(c) this work



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Scheme 1 Aryne generation from phenol- and aniline-based precursors.

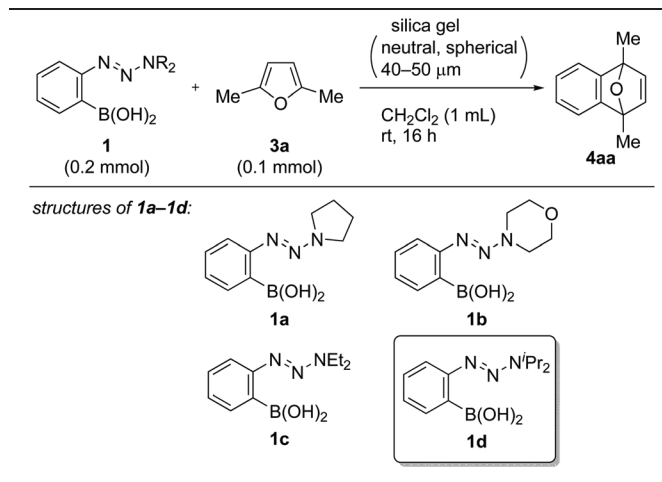
We recently reported the development of a new methodology for aryne generation from *o*-aminophenylboronates *via* the *in situ* preparation of a new aryne precursor, namely, *N*-tosyldiazene **A** (Scheme 1b).⁸ However, under the conditions implemented for the *in situ* preparation of unstable **A**, including the generation of reactive Rh(II)-nitrene species from Rh(II) catalyst and iminoiodinane (TsN = IMes), the only applicable aryneophiles were found to be azides and furans. These results prompted us to design a more stable precursor based on **A**, not requiring *in situ* preparation. After some consideration, we newly designed *o*-triazenylarylboronic acids **1** by replacing the diazene moiety of **A** with a triazenyl group, which is well-known as a masked diazonio group.^{6c,9} Herein, we describe the unexpected discovery of (hetero)aryne generation from *o*-triazenylarylboronic acids **1** under remarkably mild reaction conditions using neutral silica gel as the sole reagent (Scheme 1c).

Results and discussion

Following a literature procedure,¹⁰ *o*-triazenylarylboronic acids **1** were synthesized from *o*-iodoarylamines **2** in 32–72% yields over two steps, including triazene formation *via* diazotization followed by borylation *via* halogen–lithium exchange (Scheme 2). Notably, synthesized **1** were obtained as solids after purification by filtration and were shelf-stable at ambient temperature in air for over three months.

Initially, we examined the reaction of **1a** with 2,5-dimethylfuran (**3a**) in CH₂Cl₂ without the use of additives (Table 1, entry 1). Interestingly, TLC analysis indicated the formation of cycloadduct **4aa**, whereas the ¹H NMR spectrum of the crude product suggested that the reaction had not occurred. Indeed, **4aa** was obtained in 36% isolated yield after column chromatography on silica gel (neutral, spherical, 40–50 μm). Inspired by these results, we examined the reaction of **1d** (0.2 mmol) in the presence of silica gel.¹¹ The yield of **4aa** increased with increasing amounts of silica gel, and a maximum yield of 88% was observed with 200 mg of silica gel (entries 2–4).¹² Regarding alkyl substituents on the triazenyl group, isopropyl groups were proven to be optimal, and quantitative ¹H NMR yield was obtained with the use of **1d** (entries 4–7). Notably, when **1d** was used on 5 mmol scale, **4aa** was isolated in 92% yield (entry 7). Silica gel displayed virtually no loss of activity when it was used without dryness (entry 8). Although the

Table 1 Silica gel-induced reaction of *o*-triazenylphenylboronic acids **1** and 2,5-dimethylfuran (**3a**)^a

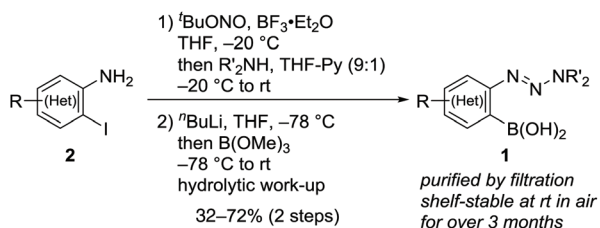


| Entry | Triazene | Silica gel ^b (mg) | Variation from standard conditions | Yield ^c (%) |
|-------|-----------|------------------------------|---|----------------------------|
| 1 | 1a | 0 | None | ND (36) ^d |
| 2 | 1a | 40 | None | 33 |
| 3 | 1a | 120 | None | 75 |
| 4 | 1a | 200 | None | 88 |
| 5 | 1b | 200 | None | 71 |
| 6 | 1c | 200 | None | 93 |
| 7 | 1d | 200 | None | Quant. (92) ^{d,e} |
| 8 | 1d | 200 | Silica gel (undried) | 95 |
| 9 | 1d | 200 | 1d : 3a = 1.5 : 1 | 98 |
| 10 | 1d | 200 | 1d : 3a = 1 : 2 | 59 |
| 11 | 1d | 200 | MeCN instead of CH ₂ Cl ₂ | 54 |
| 12 | 1d | 200 | THF ^f instead of CH ₂ Cl ₂ | NR |
| 13 | 1d | 200 | Toluene instead of CH ₂ Cl ₂ | 98 |
| 14 | 1d | 200 | Hexane instead of CH ₂ Cl ₂ | 97 |

^a Reaction conditions: **1** (0.200 mmol), **3a** (0.100 mmol), silica gel in CH₂Cl₂ (1.0 mL). ^b Silica gel was used after heating under vacuum to dryness. ^c Determined by ¹H NMR spectroscopy using 1,1,2,2-tetrachloroethane as an internal standard. ^d Yields in parentheses refer to the yields of the isolated products. ^e In 5 mmol scale (**1d**: 10.0 mmol, **3a**: 5.00 mmol). ^f Stabilizer-free.

amount of **1d** could be reduced to 1.5 equiv. without significant loss of yield (entry 9), using **1d** as the limiting reagent (**1d** : **3a** = 1 : 2) decreased the yield to 59% (entry 10). Solvent screening revealed that polar solvents, such as MeCN and THF, were markedly less effective than CH₂Cl₂ (entries 11 and 12). In particular, virtually no conversion of **1d** was observed in THF. In contrast, the use of toluene provided a yield comparable to that obtained in CH₂Cl₂ (entry 13). Interestingly, a high product yield was also obtained using hexane, even though **1d** was hardly soluble (entry 14).

With the optimized conditions in hand, we next examined the performance of functionalized aryne precursors **1e–j** and heteroaryne precursors **1k–m** (Table 2). Precursors **1e–i** bearing electron-donating or electron-withdrawing groups at the 4-, 5-, and 6-position provided **4ea–ia** in 68–98% yields (entries 1–7). The results obtained with chlorine-substituted (**1f**, **1f'**) and methoxy-substituted precursors (**1g**, **1g'**) demonstrated that the position of the substituent had little impact on



Scheme 2 Syntheses of *o*-triazenylarylboronic acids **1**.



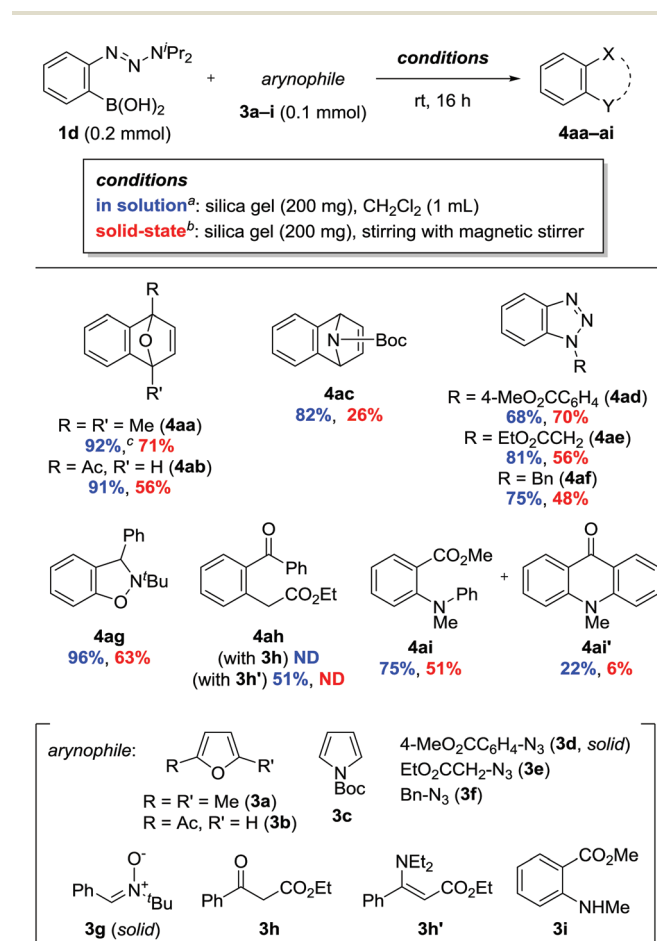
Table 2 Scope of aryne precursors 1

| Entry | Triazene | Product; yield ^a |
|-------|--|---|
| | | |
| 1 | 1e | 4ea ; 84% |
| | 1f-i | 4fa-ia (R = R ¹ or R ²) |
| 2 | R ¹ = H, R ² = Cl (1f) | R = Cl (4fa); 98% |
| 3 | R ¹ = Cl, R ² = H (1f') | R = Cl (4fa); 95% |
| 4 | R ¹ = H, R ² = OMe (1g) | R = OMe (4ga); 98% |
| 5 | R ¹ = OMe, R ² = H (1g') | R = OMe (4ga); 86% |
| 6 | R ¹ = H, R ² = CF ₃ (1h) | R = CF ₃ (4ha); 68% |
| 7 | R ¹ = H, R ² = CN (1i) | R = CN (4ia); 96% |
| 8 | 1j | 4ja ; 93% |
| 9 | 1k | 4ka 16% (1k:3a = 2:1) 13% (1k:3a = 1:2) 62% (1k:3a = 1:5) |
| 10 | 1l | 4la 28% (1l:3a = 2:1) 63% (1l:3a = 1:2) |
| 11 | 1m | 4ma 96% (1m:3a = 2:1) 83% (1m:3a = 1:2) |

^a Isolated yields.

the product yield (entries 2–5). In addition to benzynes, the present protocol was applicable to the reaction of 2,3-naphthalene (entry 8) and heteroarynes (entries 9–11).³ Under the standard conditions, 3,4-pyridine precursor **1k** resulted in the formation of **4ka** in only 16% yield (entry 9). However, we found that the use of an excess amount of aryneophile **3a** significantly improved the yield to 62%. Similar results were obtained with 5,6-quinolyne precursor **1l**, whereby only 2 equiv. of **3a** was sufficient to give **4la** in 63% yield (entry 10). In contrast to **1k** and **1l**, 4,5-indolyne precursor **1m** afforded **4ma** in excellent yields in both cases using **1m** and **3a** as the limiting reagent (entry 11).

Next, we investigated the reaction of **1d** with a range of aryneophiles (Scheme 3). The precursor was applicable to [4 + 2] and [3 + 2] cycloadditions with various aryneophiles, including furan **3b**, pyrrole **3c**, azides **3d–f**, and nitron **3g**. Unfortunately, β -ketoester **3h** failed to produce the desired product **4ah**. Instead, using enamine **3h'** provided **4ah** in 51% yield.¹³ The reaction with methyl *N*-methylantranilate (**3i**) gave *N*-phenylated product **4ai** in 75% yield along with 22% of *N*-methylacridone (**4ai'**).¹⁴ Surprisingly, the reactions also proceeded in the absence of solvent by mixing **1d**, **3**, and silica gel



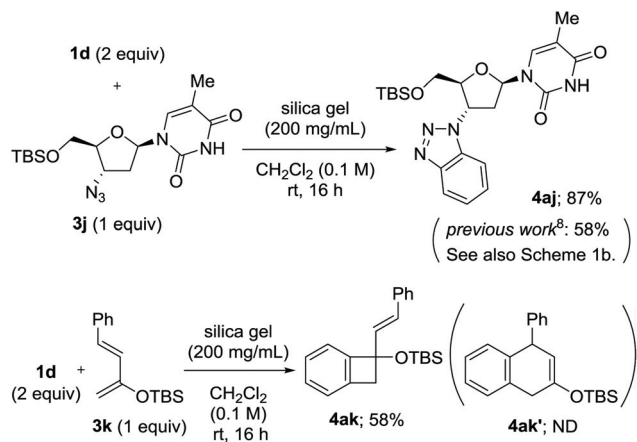
Scheme 3 Reactions of **1d** with aryneophiles **3a–i**. ^a Isolated yields. ^b ¹H NMR yields using 1,1,2,2-tetrachloroethane as an internal standard. ^c 5 mmol scale.



using a magnetic stirrer.¹⁵ The formation of **4** under these solvent-free conditions indicated that the aryne had formed *via* the contact of **1d** and silica gel at the solid–solid interface. Furthermore, the generated aryne reacted effectively even with solid arynophiles **3d** and **3g**. Notably, compared with the outcome of reactions conducted in solution, no significant loss of yield was observed, except in the cases of **3c** and **3h'**, despite a lower diffusion rate in the solid-state than in solution. The results suggested that the lifetime of aryne was sufficiently long to encounter the arynophile through diffusion on the silica gel surface. Although a growing number of studies on solid-state reactions have emerged in recent years,^{11b,d,16,17} our results represent a pioneering example of solid-state intermolecular reactions occurring *via* short-lived reactive species.^{16d}

Along with the feasibility of solid-state operation, excellent functional group tolerance is a salient feature of the present method. In particular, the fluoride-free conditions allowed for the use of arynophiles bearing silyl functionalities (Scheme 4). The reaction of **1d** and *O*-*tert*-butyldimethylsilyl (TBS)-protected zidovudine **3j** provided **4aj** in a significantly higher yield than that obtained in a previous study, with the added advantage of inexpensive and environmentally benign conditions (Scheme 4a).⁸ 2-Siloxy-1,3-diene **3k**, which is sensitive to various conditions, was also applicable as an arynophile to provide [2 + 2] cycloadduct **4ak** in 58% yield, while no evidence of [4 + 2] cycloaddition was observed.

To gain insight into the reaction mechanism, we performed a time-course study (Fig. 1). To a series of reaction vessels containing **3a** (0.1 mmol), silica gel, and CH₂Cl₂ was added **1d** (0.2 mmol), and each mixture was filtered after stirring for the indicated time to remove silica gel. ¹H NMR analysis of the crude mixture after stirring for 5 min indicated that only ~0.05 mmol of **1d** (~25% of initial amount) was contained in the mixture. The amount of **1d** slowly decreased over 4 h. Meanwhile, the formation of **4aa** proceeded in >90% yield within 4 h according to the rate of N₂ gas evolution.¹⁸ This



Scheme 4 Utilization of arynophiles bearing silyl functionalities.

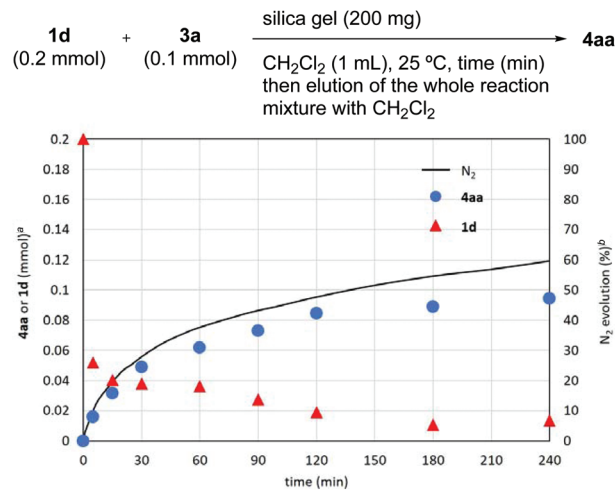


Fig. 1 Time course studies of the reaction of **1d** and **3a**. ^a Determined by ¹H NMR spectroscopy using 1,1,2,2-tetrachloroethane as an internal standard. ^b Based on **1d** (0.8 mmol).

result indicated that **1d** was strongly adsorbed on the silica gel surface preceding aryne generation.

Next, we performed competition experiments between parent precursor **1d** and substituted precursors **1f–h** or **1f'**, **g'**, and analyzed the relative rate of the reaction (k_R/k_H) using Hammett constants based on the triazenyl group (σ_N) and those based on the borono group (σ_B) (Fig. 2a).¹⁹ As a result, the plots according to Jaffé's eqn (2) and (3),²⁰ derived from the Hammett eqn (1), exhibited linearity with good R^2 values, as shown in Fig. 2b and c, respectively (see ESI for details[†]).

$$\log(k_R/k_H) = \rho_N \sigma_N + \rho_B \sigma_B \quad (1)$$

$$\log(k_R/k_H)/\sigma_N = \rho_N + \rho_B (\sigma_B/\sigma_N) \quad (2)$$

$$\log(k_R/k_H)/\sigma_B = \rho_N (\sigma_N/\sigma_B) + \rho_B \quad (3)$$

The obtained negative ρ_N and positive ρ_B values suggest simultaneous build-up of positive and negative charges on the nitrogen atom and the boron atom, respectively, in the rate-determining step (see ESI for details[†]). Combining the results in Fig. 1 and 2, we propose a plausible reaction mechanism, as illustrated in Scheme 5. Precursor **1** is adsorbed onto the silica gel surface *via* the boronate moiety, followed by the formation of zwitterionic intermediate **B**.²¹ In other words, triazenyl and borono groups were activated as diazonio and boronate groups, respectively. Upon this dual activation, highly stable precursor **1** is capable of generating an aryne under remarkably mild conditions without heating, photo-irradiation, or the use of acids or bases.⁹ The role of silica gel remains unclear.²² While the use of excess acetic acid instead of silica gel induced aryne generation from **1d**, the yield of **4aa** was only 47% (Scheme 6a). Thus, weak acidity of silica gel seems to be insufficient to activate **1**. On the other hand, (\pm)-camphorsulfonic acid [(\pm)-CSA] led to a formation of **4aa** in 97% yield (Scheme 6b). Thus, both heterogeneous conditions using silica



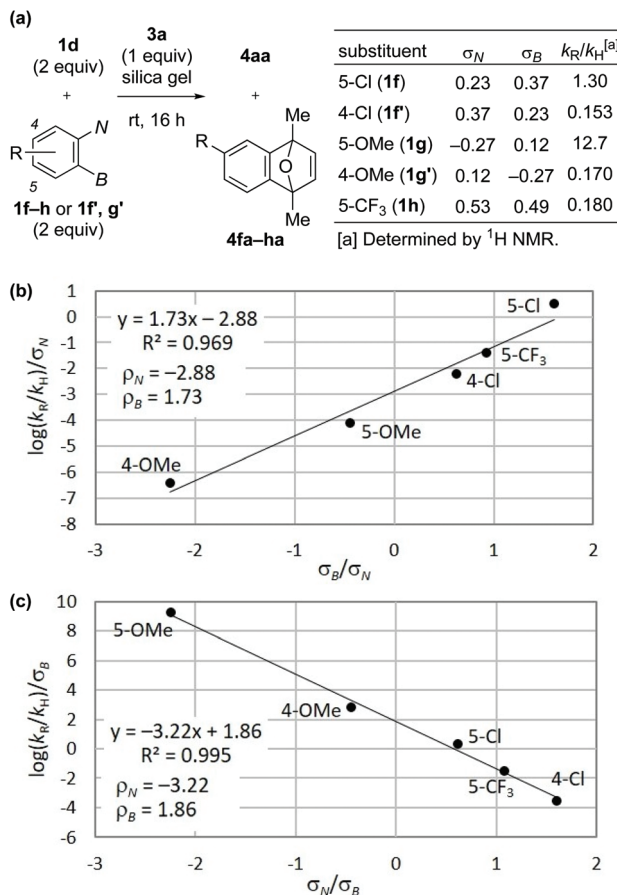
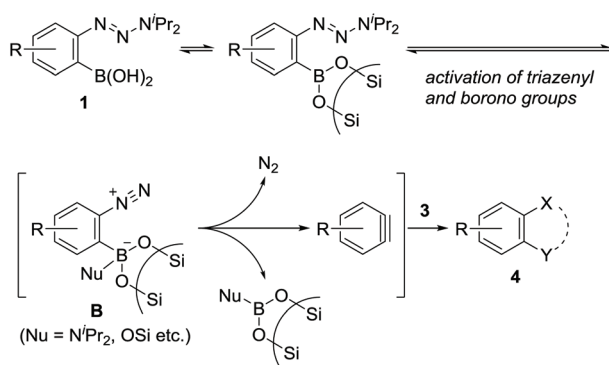
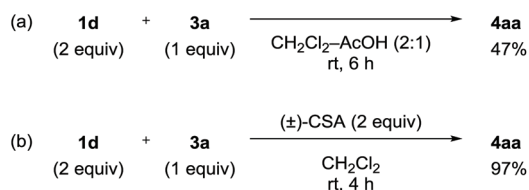


Fig. 2 (a) Competition experiments between **1d** and **1f–h**, or **1f'**, **g'**. (b) Jaffé's plot based on eqn (2). (c) Jaffé's plot based on eqn (3).



Scheme 5 Plausible dual activation mechanism for aryne generation from **1**.



Scheme 6 Reactions of **1d** with **3a** using Brønsted acid.

gel and homogeneous conditions using Brønsted acid were applicable to aryne generation from **1**.

Conclusions

We have developed new aniline-based aryne precursors **1**, which generate arynes under remarkably mild conditions through the use of silica gel as the activating agent. The protocol was applicable to a wide range of (hetero)arynes and various arynophiles reacting in solution and in the solid-state. The reaction proceeded *via* a dual activation mechanism to generate arynes, as rationalized through Jaffé's plot analysis based on Hammett constants. Investigation of further synthetic applications of the protocol as well as mechanistic studies on the role of silica gel are currently in progress.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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- 21 No reaction was observed with 2-(trimethylsilyl)phenyltriazene under the conditions using silica gel.
- The reaction scheme shows 2-(trimethylsilyl)phenyltriazene (a benzene ring with a TMS group at the 2-position and an N=N-NiPr₂ group at the 1-position) reacting with 2,5-dimethylfuran (a five-membered furan ring with methyl groups at the 2 and 5 positions). The reaction conditions are silica gel, CH₂Cl₂, rt, 16 h. The product is labeled as NR (No Reaction).
- 22 No reaction was observed with the use of other adsorbents, such as alumina (neutral or basic) and MS 4A. Compared to that of spherical silica gel (40–50 μm, specific surface area: 630–730 m² g⁻¹), crushed-shape silica gel (40–63 μm, specific surface area: 480–540 m² g⁻¹) displayed lower activity. In addition, recycling of spherical silica gel slightly diminished the activity. Thus, the activity of silica gel depends on the properties of their surface. Additionally, we consider that the silica gel did not play the role of a photocatalyst, because quantitative conversion of **1d** was observed in the dark. See ESI for details.†

