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Isopropanol-assisted synthesis of highly stable MAPbBr₃/p-g-C₃N₄ intergrowth composite photocatalysts and their interfacial charge carrier dynamics†

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Two phase photocatalysts can be intergrown with each other, resulting in superior photocatalytic properties. Herein, methylamine lead bromide (MAPbBr₃) wrapped/entrapped protonated graphitic carbon nitride (p-g-C₃N₄) intergrowth microcrystals were fabricated by mixing a perovskite precursor with p-g-C₃N₄ colloidal sol. A highly stable isopropanol (IPA) solvent based photocatalytic system for dye degradation was demonstrated. The composite with an optimal p-g-C₃N₄ mass percentage of 3.3 wt% (denoted as MAPbBr₃/p-g-C₃N₄-1.0 mg) exhibited the highest photocatalytic degradation of malachite green (99.8%) within 10 min under visible light, which was 5.3-fold and 16-fold greater than that exhibited by its constituents separately. The strong chemical interaction and fundamental photophysical processes in MAPbBr₃/p-g-C₃N₄ were systematically evaluated by spectroscopic and electrochemical techniques, confirming the effective separation of photogenerated electron-hole pairs and faster interfacial charge transfer behavior. Furthermore, active superoxide radicals (O₂^{•-}) played a vital role in the catalytic reaction, because of the significant photoinduced electron transfer rate (*k_{et}*) in the inverted type-I core/shell MAPbBr₃/p-g-C₃N₄ band configuration structure. In addition, MAPbBr₃/p-g-C₃N₄ has good cycling stability for 10 cycles and versatility for other cationic (RhB) and anionic (MO) dye pollutants, indicating the great potential for solar energy conversion into chemical energy.

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1. Introduction

In 2016, Park and coworkers first introduced the dynamic equilibrium concept utilizing organic-inorganic hybrid halide perovskite (CH₃NH₃PbX₃) catalysts.¹ Subsequently, both all inorganic and hybrid perovskites with the general formula APbX₃ (A = methylammonium (MA), formamidinium (FA) or cesium (Cs); X = Br, I) were explored in photocatalytic water splitting,^{2,3} CO₂ reduction in ethyl acetate solvent^{4,5} and selective organic synthesis fields.^{6,7} Additionally, halide perovskites for the photocatalytic and photoelectrochemical degradation of organic dyes were reported.^{8–12} For example, CsPbX₃ can decompose methyl orange (MO) within 100 min under visible light irradiation.¹⁰ Cardenas-Morcoso *et al.* confirmed that the photocatalytic degradation of a 2-mercaptobenzothiazole (MBT) compound resulted from hole injection from CsPbBr₃ nanodots to MBT.¹¹ Moreover, a novel alcohol-based photocatalytic

system using lead-free Cs₂AgBiBr₆ perovskite was proposed for degradation of dyes including rhodamine B (RhB), methyl red (MR), *etc.*¹² However, to the best of our knowledge, literature reports on halide perovskite MAPbBr₃ photocatalysts for degradation of dyes in isopropanol (IPA) solvent are still rare.

Because of the ionic migration nature, the inherent instability issues of organic halide perovskites make them generally incompatible with typical photocatalytic reactions in aqueous media. Therefore, special organic alcohol-catalyzed substitutional growth processes^{13,14} and polymer modification¹⁵ were used to prepare highly air (up to ≈ 1.5 months), thermally and chemically stable halide perovskites with unique surface chemical states. In particular, large MAPbBr₃ microcrystals were fabricated through an isopropanol-assisted solvothermal method in our previous report.¹⁴ Furthermore, surface modification and defect passivation strategies using carbon-based materials had been investigated to increase the chemical stability, for example, *via* integrating graphene oxide (GO) with MAPbI₃ (ref. 3) and CsPbX₃ (ref. 4), and polymeric carbon nitride (C₃N₄) with CsPbBr₃ quantum dots⁵ and MAPbBr₃ nanocrystals.¹⁶ The stability enhancement was attributed to the interaction and intimate contact at the interface between the nanocrystals and the encapsulation layer, facilitating strong interfacial electronic hybridization and coupling.¹⁵ Although

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grafting rGO onto MAPbI₃ microparticles was reported,³ it is clear that hybrid perovskite microcrystals with high crystallinity will reduce defects, but at the same time the long diffusion distance and reduced active surface area are not beneficial to the photocatalytic activity relative to nanocrystal composites. Recently, two phases of cocatalysts and host photocatalysts were intergrown with each other, which resulted in superior photocatalytic properties.¹⁷ Befitting from this two phase intergrowth feature, good contact and strong coupling between cocatalysts and host photocatalysts give rise to efficient charge transfer and surface redox reaction. On the other hand, in perovskite solar cells, new concepts such as incorporating g-C₃N₄ into perovskite precursor solution were developed to enhance the binding strength in perovskite films and finally improve crystallization and reduce the intrinsic defect density of the perovskite layer.^{18,19} In addition, as another rising star, graphite-like carbon nitride (g-C₃N₄) is an important class of metal-free organic photocatalysts that has attracted particular research interest since its first demonstration of photocatalytic H₂ production in 2009.²⁰ Further research had been motivated by their tunable band gaps, earth-abundance, low cost of synthesis and ease of surface functionalization. In particular, g-C₃N₄ is an ideal choice for coupling with perovskite-type metal oxide compounds such as AgNbO₃, KTa_{0.75}Nb_{0.25}O₃ and bismuth oxyhalide-based semiconductors, and black phosphorus^{21–24} to form hybrid photocatalysts that greatly hinder the annihilation of charge carriers and extend the life of electrons. However, there are few reports on the preparation of highly stable MAPbBr₃/C₃N₄ hybrid intergrowth microcrystal photocatalysts as well as investigation of interfacial charge carrier dynamics.

Herein, we designed highly crystalline MAPbBr₃ microcrystal wrapped/entrapped protonated graphitic carbon nitride (p-g-C₃N₄) intergrowth structures by a one-step injection of C₃N₄ sol into the perovskite precursor without any surfactants. The unique chemical covalent C–N–Br interaction at the interface could eliminate the defects and enhance the stability of MAPbBr₃/p-g-C₃N₄, which facilitated the charge separation in the photocatalytic process. In addition, we systematically studied the effects of the composition and carrier dynamics behavior of the composites on photocatalytic activity for molecular oxygen activation.

2. Experimental section

2.1. Materials and preparation procedures

All chemicals and reagents were analytically pure and used without further purification. The dark-yellow g-C₃N₄ powder was synthesized according to a modified approach reported previously.²⁰ g-C₃N₄ was treated with HBr at room temperature for 10 hours, collected, and centrifuged to obtain a light yellow powder (protonated p-g-C₃N₄). A series of characterization experiments with bulk b-g-C₃N₄ and protonated p-g-C₃N₄ nanosheets showed that the p-g-C₃N₄ nanosheets have a thickness of 4–5 nm (Fig. S1–S3†). Next, 30 mg (0.08 mmol) lead acetate (Pb(CH₃COO)₂·3H₂O; 99.99% metal basis; Aladdin Industrial Corporation), 5 mL HBr solution (*Z* ≥ 47.0%; AR; Aladdin Industrial Corporation), and 450 μL CH₃NH₂ alcohol

solution (30–33%; AR; Aladdin Industrial Corporation) were added into a beaker with stirring. After stirring constantly for 15 min to get a homogeneous mixture, different amounts of p-g-C₃N₄ (0, 0.25/0.08 wt%, 0.5/1.6 wt%, 1/3.3 wt%, 1.5 mg/5 wt%) dissolved in isopropanol (IPA) were added. The mixture was transferred into a 50 mL stainless-steel Teflon-lined autoclave and the container was closed and maintained at 120 °C for 4 h. Then the autoclave was allowed to cool down to room temperature under ambient conditions. The resulting solid product was collected, rinsed with distilled isopropanol several times and then dried in a vacuum oven at 55 °C for 12 h.

2.2. Characterization of samples

The morphologies and microstructures were investigated using a scanning electron microscope (FESEM) (Hitachi S-4800). The Fourier transform infrared (FTIR) spectra were acquired using a Nexus 870 FTIR instrument (USA). The photoluminescence (PL) and time resolved PL decay spectra were measured using a home-built PL scanned imaging microscope coupled with a time-correlated single photon counting (TCSPC) module at 500 nm laser excitation at room temperature. Electrochemical experiments were performed using an electrochemical workstation (CHI-660E) with a three-electrode system, employing a Pt coil as the counter electrode (CE) and a Ag/AgCl electrode as the reference electrode (RE). Dichloromethane (DCM) and tetrabutyl ammonium hexafluorophosphate (Bu₄NPF₆) were used as the solvent and conducting electrolyte, respectively. Hall coefficients (*R*_H) were measured in a Quantum Design Physical Property Measurement System (PPMS). The charge carrier concentration (*n*_H) was calculated by using $n_H = 1/R_H e$, where *e* is the elementary charge. The Hall carrier mobility (*μ*_H) was calculated according to the relation $\mu_H = \sigma R_H$.

2.3. Photocatalytic experiments

The photocatalytic activity of the samples was evaluated by the degradation of malachite green (MG) under visible light. Typically, 5.0 mg catalyst was well dispersed in 40 mL MG isopropanol solution at an initial concentration of 10 mg mL^{−1} (*C*₀). Subsequently, the dispersion was stirred for 60 min in the dark to reach the adsorption equilibrium. Then, the container was irradiated under visible light from a 300 W Xe lamp with a vertical distance of 20 cm. Finally, the concentration of MG (*C*) was analyzed using an UV-vis spectrophotometer at given time intervals. ESR-trapping measurements were carried out in 2 mL acetonitrile mixed with 20 μL of 5,5-dimethyl-1-pyrroline-*N*-oxide (DMPO) and 4 mg samples. After being irradiated, the mixture was characterized using a Bruker EMX plus model spectrometer operating at room temperature.

3. Results and discussion

3.1. Structure and morphology of MAPbBr₃/p-g-C₃N₄

As we all know, the growth of perovskite single crystals, nanocrystals and polycrystalline films is a complicated process which strongly depends on precursor composition, solvent choice, deposition temperature, *etc.* Recently, the substitutional growth



mechanism catalyzed by polar protic alcohols has been used for the preparation of highly chemically stable MAPbI₃ powders up to ≈ 1.5 months. The reaction between alcohols and the precursors is initiated by the polar nature of the alcohols that can readily donate hydrogen to the reaction medium.¹³ On the other hand, the stable p-g-C₃N₄ colloidal suspension was prepared by HBr treatment methods, without using other acids such as HCl or HNO₃ in order to avoid introducing impurities. Last but not least, it is reported that suitable incorporation of g-C₃N₄ could reduce the intrinsic defect density by passivating the charge recombination centers, improving the crystal quality of the perovskite film.¹⁹ Therefore, isopropanol and HBr are very suitable precursors for our present fabricated system. Fig. 1

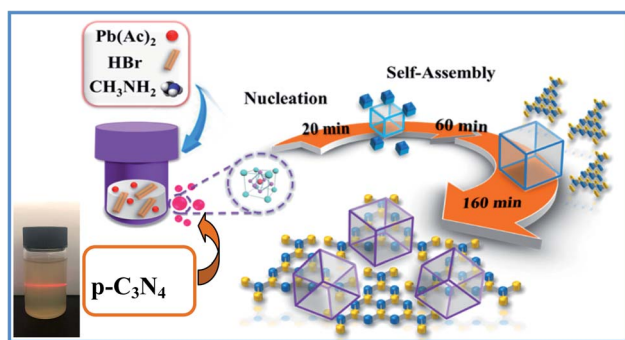


Fig. 1 Schematic illustration of the formation and morphological evolution process of MAPbBr₃/p-g-C₃N₄ composites using a p-g-C₃N₄ sol solution through a solvothermal method.

shows a schematic illustration of the formation and morphological evolution process of the MAPbBr₃/p-g-C₃N₄ samples using the stable p-g-C₃N₄ colloidal suspension solution. A typical Tyndall effect of p-g-C₃N₄ colloidal suspension is clearly evident, as indicated by the beam passing through the transparent sol solution. According to our previous research work,¹⁴ the MAPbBr₃ microparticles (with an average size of about 5 μ m) become rougher due to the wrapped/entrapped p-g-C₃N₄ nanosheets in Fig. 2a and b. With increasing the amount of 1.5 mg p-g-C₃N₄, the sample appeared as irregular shaped particles in Fig. 2c, indicating that the high amount of C₃N₄ nanosheet sol could affect the crystal nucleation and growth *via* modulation of the local dielectric environment and precursor concentrations during growth kinetics processes.^{15,19} It is important to understand the nature of the interaction between MAPbBr₃ and p-g-C₃N₄, which influences the photo-electrochemical properties of the composite. The composition and chemical structure of the MAPbBr₃-p-g-C₃N₄ samples with different contents of p-g-C₃N₄ (denoted as MAPbBr₃/p-g-C₃N₄-x, where x = 0, 0.25, 0.5, 1.0, and 1.5 mg) were studied by XRD and FTIR spectroscopy. The typical peaks of Fig. 2d at 2θ of 14.9°, 21.1°, 30.2°, 33.8°, 37.0°, 43.1°, and 45.6° correspond to (100), (110), (200), (210), (211), (220) and (300) planes, respectively, indicating that the product contains pure cubic phase CH₃NH₃PbBr₃ (MAPbBr₃) with $a = 5.9394$ Å (space group = $Pm\bar{3}m$). Besides, the strong one at 27.4° represents the stacking of the π -conjugated p-g-C₃N₄ (p-C₃N₄) layers. No other impurity phases are discovered, reflecting the presence of a high crystal quality two-phase composite. Particularly, the intensity of XRD peaks

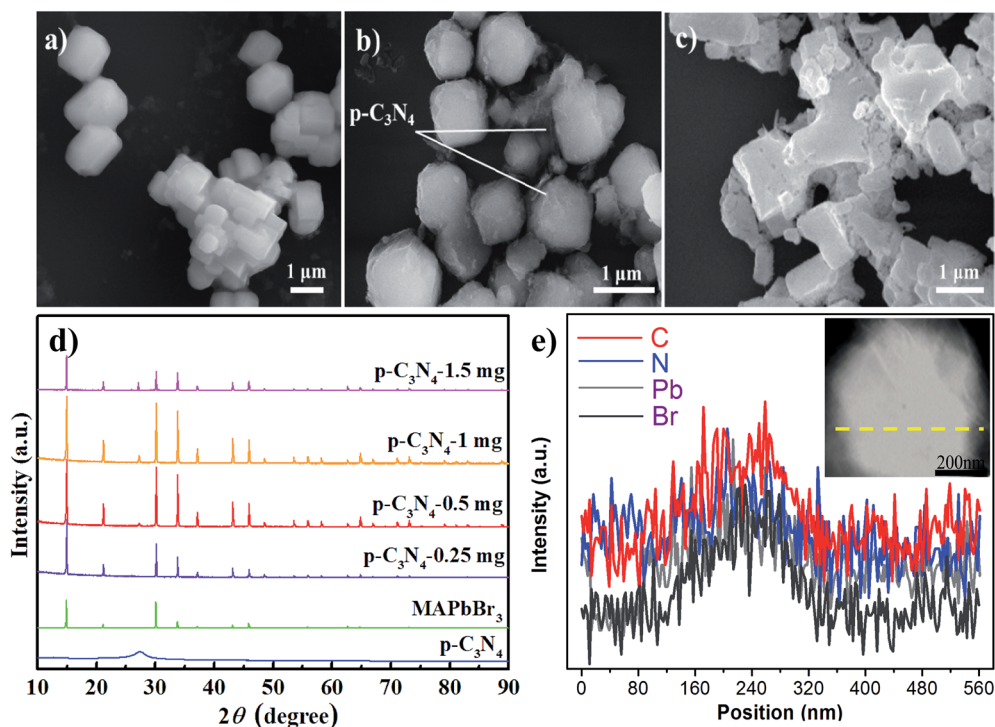


Fig. 2 (a–c) SEM images of MAPbBr₃/p-C₃N₄-0, -1.0, and -1.5 mg samples, respectively. (d and e) The corresponding XRD patterns and EDS line scan image of an individual MAPbBr₃ microcrystal encapsulated nanosheet.



for the MAPbBr₃/p-g-C₃N₄-0.25 mg is stronger than that of pure MAPbBr₃, suggesting that the crystallinity of the composites was improved. This is attributed to the relatively controlled nucleation and growth behavior due to p-g-C₃N₄ providing heterogeneous nucleation sites in the crystallization process through interaction between the hydrogen atoms of the MA and basic sites on p-g-C₃N₄.¹⁹ Fig. 2e shows the EDS line profiles recorded from an individual MAPbBr₃/p-g-C₃N₄ intergrowth microcrystal, indicating that the content of C and N is slight higher than that of Pb and Br due to wrapped/entrapped p-g-C₃N₄ nanosheets. Such microstructures could not be prepared when bulk g-C₃N₄ powder was used. With increasing the content of p-g-C₃N₄, the diffraction peak at 26.4° gradually increases, while the diffraction peak intensity of MAPbBr₃ increases to the highest value for MAPbBr₃-p-g-C₃N₄-1 mg, and then decreases for MAPbBr₃-p-g-C₃N₄-1.5 mg. The reduced crystallinity of MAPbBr₃-p-g-C₃N₄-1.5 mg may be attributed to the occurrence of micro-strain in the MAPbBr₃ crystals, induced by crystal imperfections/structural defects including dislocations, vacancies, stacking faults, *etc.* It is worth noting that the peaks at 26.4° of p-g-C₃N₄ still exist for MAPbBr₃-p-g-C₃N₄ after ultrasonic washing treatment, showing that the MAPbBr₃/p-g-C₃N₄ intergrowth hybrids are very robust due to chemical interaction, which is beneficial to electron transportation.

The interaction will be further discussed in the FTIR spectra. For MAPbBr₃ microcrystals, the N–H stretch vibrations around 3100–3200 cm⁻¹, and C–H bending and CH₃ rocking modes at 1471 cm⁻¹ and 916 cm⁻¹ are in good agreement with previous results.¹⁴ Evidently, all of the bands assigned to p-g-C₃N₄ and MAPbBr₃ are visible in MAPbBr₃/p-g-C₃N₄ and the absorption intensity at 810 cm⁻¹ is substantially increased and shifted toward a higher wavenumber with the introduction of greater amounts of p-g-C₃N₄ in Fig. 3a. It is reported that “nitrogen pots” of g-C₃N₄ with abundant amino sites (NH_x, x = 1, 2) on the edges of heptazine units can take up the PbBr₂ precursors *via* surface N–Br bonding interactions, forming an effective nucleation point for crystal growth.⁵ In contrast, it is worthy of note that when 1.5 mg p-g-C₃N₄ was introduced, the peak intensities of MAPbBr₃/p-g-C₃N₄ decrease, which is possible due to greater

particle agglomeration around the edges of the nanosheets. Furthermore, the typical stretching modes of aromatic C–N and C=N heterocycles in MAPbBr₃/p-g-C₃N₄ at 1241 cm⁻¹ are systematically shifted to lower values with the increasing content of p-g-C₃N₄ (Fig. 3b), which is ascribed to the high electron density of aromatic heterocycles due to C–N–Br bond chemical interaction. An analogous absorption change trend has also been observed, showing high interfacial coupling quality.⁵ In addition, cation–π interaction, an example of non-covalent bonding between an electron-rich π system and an adjacent cation, is proposed for perovskite films.^{18,19} In brief, chemical interaction at the interface could eliminate the defects and enhance the stability of MAPbBr₃/p-g-C₃N₄, which is also confirmed by the following photocatalytic tests.

The optical properties of the samples were investigated by UV-vis diffuse reflectance spectroscopy (UV-vis DRS). As is shown in Fig. 4a, MAPbBr₃-p-g-C₃N₄-x has an absorption edge at about 595 nm, which is the characteristic peak of MAPbBr₃.² Additionally, the optical band gaps of MAPbBr₃-p-g-C₃N₄-x estimated from the Tauc plots using the equation $ah\nu \propto (h - E_g)^2$ are 2.132, 2.143, 2.175, 2.206, and 2.101 eV, respectively (Fig. 4b). The band gap of our sample is smaller than the previously reported optical band gap of MAPbBr₃ nanoparticles ($E_g = 2.39$ eV),² because the latter are nanoscale particles that can present the quantum effect. Simultaneously, the band gap values of our samples are close to the values reported for MAPbBr₃ single crystals ($E_g = 2.18, 2.21$ eV) grown through an inverse temperature crystallization and antisolvent vapor-assisted approach.^{2,25–27} It is well known that the Br (4s²3d¹⁰4p⁵) 4p orbital and Pb (6s²5d¹⁰4f¹⁴6p²) 6p orbital contribute the most to the valence band maximum (VBM) and conduction band minimum (CBM) of MAPbBr₃, respectively.³ So the band gap variation with MAPbBr₃/p-g-C₃N₄ composition is probably mainly due to the shrinkage of PbBr₆ octahedra due to the chemical interaction, which is consistent with FTIR results.

3.2. Photocatalytic degradation performance

After the structural analysis, photocatalytic degradation of malachite green (MG) in IPA solvent was employed as a model reaction to probe the structure–activity relations. MG, a cationic

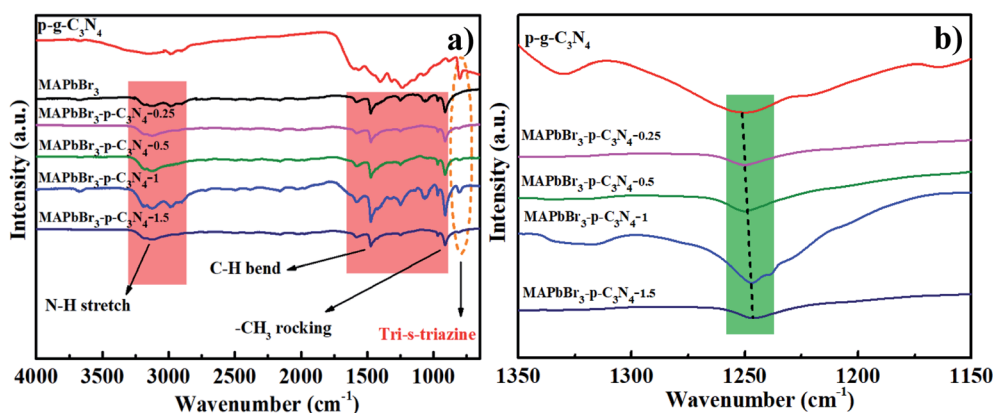


Fig. 3 (a and b) Full and partial regional FTIR spectra of products prepared with different contents of p-g-C₃N₄ confirming that the C–N–Br chemical bonding interaction is present in MAPbBr₃/p-g-C₃N₄.



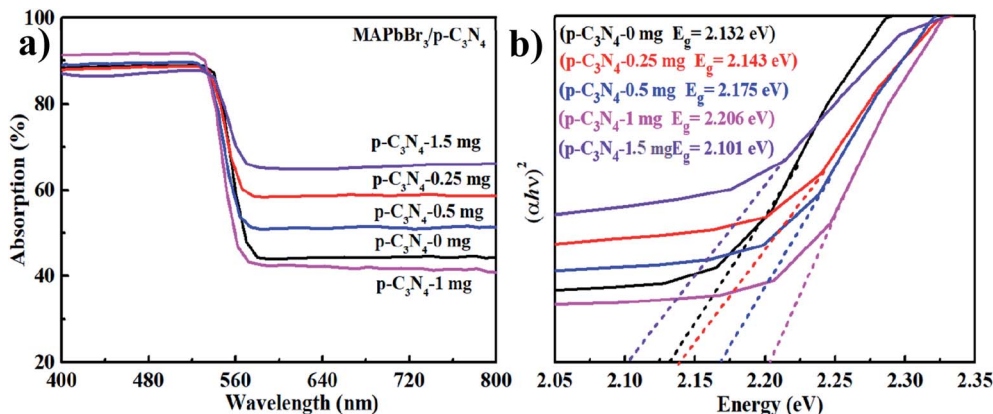


Fig. 4 (a and b) UV-vis diffuse reflectance spectra (DRS) and Tauc plots.

triphenylmethane dye (Fig. S4[†]), has been extensively used in the textile, leather, and pharmaceutical industries due to its low cost and high efficiency for disease prevention.^{8,28} Different control experiments were conducted to compare the ability of MAPbBr₃, p-g-C₃N₄, and MAPbBr₃/p-g-C₃N₄ composites under visible-light illumination. Fig. 5a shows the concentration–time profiles of the MG degradation over various catalysts. After irradiation for 10 min, 30%, 35%, 52%, 46%, 98% and 60% MG were reduced by MAPbBr₃, p-g-C₃N₄, and MAPbBr₃/p-g-C₃N₄-0.25, -0.5, -1.0 and -1.5, respectively. Remarkably, MAPbBr₃/

C₃N₄-1.0 showed significantly faster degradation rates. To the best of our knowledge, there has been no study for the photocatalytic degradation of MG within 10 min under visible light irradiation. A comparison of previously reported values for photocatalytic degradation of MG is shown in Table S3.[†] The time-dependent 3D absorption spectra of a malachite green solution in the presence of MAPbBr₃/p-g-C₃N₄ are shown in Fig. 5b. During the 10 min irradiation, the characteristic absorption at about 618 nm in the MG spectra displayed a gradually declining trend, which indicated the process of *N*-demethylation, *i.e.*, the non-selective attack of reactive oxygen species on the C–N bond as reported.²⁹ The absorbance peaks at 425 nm reduce in intensity, which evidently indicates that the whole conjugated chromophore structure of MG has been destroyed.²⁹ The surface properties of a photocatalyst is usually closely correlated with the photocatalytic activity. It is well known that a photocatalyst with a large surface area can provide more active sites and good adsorption for reactants, both of which can speed up the heterogeneous photocatalytic process. In the current case, N₂-adsorption analysis indicated that the addition of a small amount of p-g-C₃N₄ slightly increased the surface area of samples. The BET values of MAPbBr₃/p-g-C₃N₄, MAPbBr₃, and p-g-C₃N₄ samples are determined to be 10.2, 7.2, and 12.2 m² g⁻¹, respectively. Clearly, the consistency between the BET surface area and catalytic performance is not established due to similar specific surface areas, indicating that the surface properties of the MAPbBr₃/p-g-C₃N₄ are not the key factor.

According to the kinetic curves in Fig. 6a, the rate constant *k* of pseudo-first-order kinetics follows the order MAPbBr₃/p-g-C₃N₄ (0.195 min⁻¹) > MAPbBr₃ (0.037 min⁻¹) > p-g-C₃N₄ (0.018 min⁻¹) > b-C₃N₄ (0.012 min⁻¹). Clearly, the MAPbBr₃/p-g-C₃N₄-1.0 composite has the highest rate constant, which is almost 5.3 and 16 times higher than that of pure MAPbBr₃ and C₃N₄, respectively (Fig. 6b). These results allow us to conclude that controlling the addition ratios of p-g-C₃N₄ is important for the achievement of an optimal synergistic interaction between MAPbBr₃ and p-g-C₃N₄. Greater amounts of p-g-C₃N₄ nano-sheets (1.5 mg) are not good for the photocatalytic performance. This can be attributed to two reasons: (1) the irregular micro-crystals with low crystallinity resulted in trap states and thus

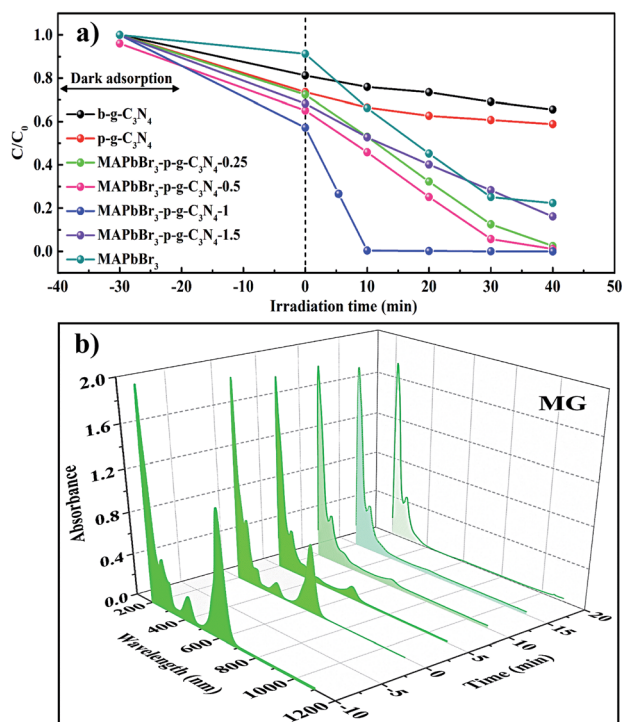


Fig. 5 (a) Photocatalytic activities of concentration–time profiles of the degradation of MG over various catalysts including b-g-C₃N₄, p-g-C₃N₄, MAPbBr₃, and MAPbBr₃/p-g-C₃N₄ samples under visible light. (b) 3D UV-vis spectra of malachite green (MG) at different irradiation times in the presence of the MAPbBr₃-p-g-C₃N₄-1.0 mg sample.



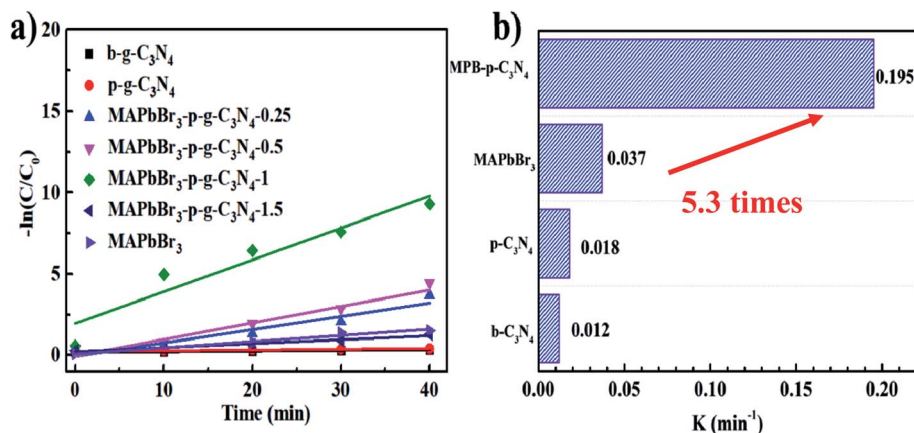


Fig. 6 (a) Photocatalytic kinetic behavior for the corresponding dependence of $\ln(C/C_0)$ on irradiation time. (b) Comparison of slope values of different samples.

reduced the efficiency of charge separation. (2) The increased p-g-C₃N₄ surface coverage on perovskite materials also blocked the absorption of light of MAPbBr₃ and decreased the exposed surface area of MAPbBr₃ to the reagent solution.^{4,5}

3.3. Discussion of underlying photocatalytic mechanisms

An in-depth understanding of plausible reaction mechanisms for the photodegradation of MG over semiconductors is desirable. Under visible light irradiation, the photogenerated superoxide radical ($\cdot\text{O}_2^-$), holes (h^+) and hydroxyl radical ($\cdot\text{OH}$) play an important role in the degradation of organic dyes.⁸⁻¹² For unambiguous identification of reactive oxygen species (ROS) involved in the dye degradation process, the reactive-species-trapping experiments were performed to clarify the contribution of different active species to the reaction. In general, *p*-benzoquinone (*p*-BQ), ethylenediaminetetraacetic acid disodium salt (EDTA-2Na), *tert*-butanol (*t*-BuOH) and AgNO₃ are known as effective scavengers for $\cdot\text{O}_2^-$, h^+ , $\cdot\text{OH}$ and photogenerated electrons (e^-), respectively. As illustrated in Fig. 7a, the addition of *t*-BuOH had little effect on the photodegradation of MG. It is worth noting that isopropanol (IPA) and *t*-BuOH have commonly been employed to impair the hydroxyl radical ($\cdot\text{OH}$) in photocatalytic reaction systems; therefore, based on our present IPA solvent system, it is ruled out that $\cdot\text{OH}$ plays a role. However, the presence of *p*-BQ, AgNO₃ and EDTA-2Na brought about obvious deactivation of catalysts and the photodegradation performance was decreased. Interestingly, among the scavengers employed, BQ can significantly suppress the degradation rate from 98% to 46% after 15 min, indicating that $\cdot\text{O}_2^-$ is a major reactive species in the system. In addition, spin-trapping electron paramagnetic resonance (EPR) experiments were also performed to give direct proof for the generation of $\cdot\text{O}_2^-$ over the MAPbBr₃/g-C₃N₄-1.0 composites. In Fig. 7b, no EPR signal was observed in the dark, indicating that there is no detectable $\cdot\text{O}_2^-$ species formed. Upon visible-light irradiation for 5 min, a four-line EPR signal with an intensity ratio of 1 : 1 : 1 : 1 was clearly observed, which is characteristic of the (5,5-dimethyl-1-pyrroline *N*-oxide) DMPO- $\cdot\text{O}_2^-$ adduct, demonstrating the

formation of the superoxide radical $\cdot\text{O}_2^-$ in the photocatalytic reaction.⁷ The observations strongly suggest that oxygen plays a critical role as an electron acceptor. Namely, O₂ accepts the electrons from the semiconductor conduction band to form $\cdot\text{O}_2^-$ due to a strong thermodynamic driving force to reduce O₂ (E° , (O₂/ $\cdot\text{O}_2^-$) = -0.33 V vs. Ag/AgCl at pH = 7) to superoxide $\cdot\text{O}_2^-$. Very interestingly, $\cdot\text{O}_2^-$ was also observed in other perovskite photocatalytic systems including FAPbBr₃/TiO₂ and Cs₂AgBiBr₆.^{7,12} Very interestingly, $\cdot\text{OH}$ active species for CsPbX-based catalysts in an aqueous solution containing water molecules are often observed, which is attributed to the

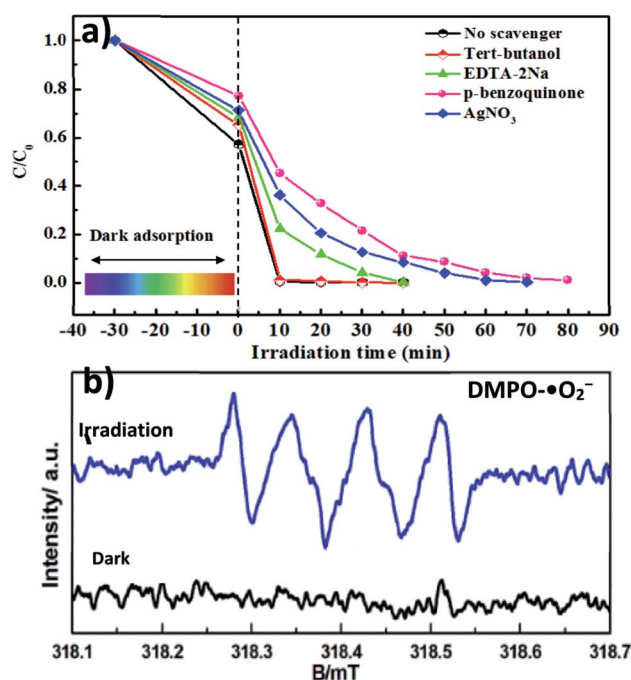
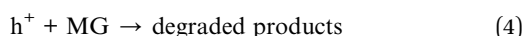
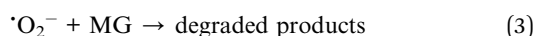
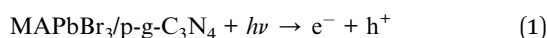


Fig. 7 (a) Effect of radical scavenger/quencher including *p*-benzoquinone (*p*-BQ), ethylenediaminetetraacetic acid disodium salt (EDTA-2Na), *tert*-butanol (*t*-BuOH) and AgNO₃ on the degradation process of MG. (b) ESR spectra of the MAPbBr₃/p-g-C₃N₄-1.0 mg sample without and with visible-light irradiation for 5 min for the detection of $\cdot\text{O}_2^-$ radicals.



oxidation of water molecules with its negative side (δ^- , O) by the photo-formed positive charges (holes) on the semiconductor surface.^{10,30} However, $\cdot\text{O}_2^-$ is the main active species in the photocatalytic dye degradation in organic alcohol solvent using organic-inorganic hybrid halide perovskites,^{7,12} demonstrating that the incorporation of organic solvent inhibits the generation of other ROS, especially $\cdot\text{OH}$. Hence it can be concluded that tuning the solvent system may control the active species for organic synthesis. The phenomena may be attributed to two facts that photocatalytic reactions were performed in organic solvent media without H_2O and the conjugated carbon nitride with amino nitrogen components and vacancy defects could provide more sites to adsorb O_2 by Lewis acid-base interaction. In conclusion, according to the analysis above, the reactions listed below may be involved:



3.4. Interfacial charge transfer

Subsequently, to further investigate the interfacial charge transfer in $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ intergrowth microcrystals, the PL spectra and time-resolved transient PL spectroscopy (TRPL) spectra of these samples with different loading amounts of p-g- C_3N_4 were recorded and are shown in Fig. 8. The PL spectrum of the MAPbBr_3 crystal in Fig. 8a displayed a symmetric broad emission band with a peak at 530 nm, which is attributed to charge carrier recombination from the CB to the VB *via* a radiative pathway.² With an increase in the amount of p-g- C_3N_4 , a slight shift (521 nm) was observed, which may arise from the formation of MAPbBr_3 with different structural defect densities. On the other hand, a gradual decrease in the PL intensity was

found, suggesting that an additional energy-transfer pathway exists in addition to the intrinsic radiative channel for excited-state electron transfer. High crystalline quality MAPbBr_3 -p-g- C_3N_4 -1 mg presented the weakest PL emission and the PL intensity is strongly quenched, indicating that the excited electrons at the interface of the $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ composite will rapidly be separated and recombination will be suppressed.

The decay traces of MAPbBr_3 (MPB_3) and $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -*x* were fitted using triexponential decay kinetics and are shown in Fig. 8b. The derivation of the three components and intensity-average lifetime (τ_{av}) was carried out using the following equation, and the results are listed in Table S1:†

$$\tau_{\text{av}} = (A_1\tau_1^2 + A_2\tau_2^2 + A_3\tau_3^2)/(A_1\tau_1 + A_2\tau_2 + A_3\tau_3) \quad (5)$$

where *A* and τ denote the amplitudes and emission lifetimes of each component. The τ_{av} values were determined to be 43.3 ns for MAPbBr_3 , 28.2 ns for $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -0.25, 23.1 ns for $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -0.5, 20.8 ns for $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -1.0 and 48.6 ns for $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -1.5, exhibiting a PL decay time on the nanosecond scale. Significantly, the τ_{av} value decreased to a minimum value for the $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -1.0 sample, showing that the charge-separated state lifetime of $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -1.0 is ~ 2.08 -fold longer than that of MAPbBr_3 , in agreement with previous observations for $\text{MAPbI}_3/\text{rGO}$ and $\text{CsPbBr}_3/\text{rGO}$ composites with shorter lifetime compared to pure perovskite.^{3,4} More importantly, the contribution of the fast decay component (τ_1) underwent a significant increase from 9.70% for MAPbBr_3 to 91.1% for $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ -1.0, illustrating that more pronounced interfacial electron transfer occurred due to the existence of non-radiative quenching routes.⁴ Therefore, based on the PL and TRPL results, it can be concluded that the N-Br chemical bond of the MAPbBr_3 wrapped/entrapped p-g- C_3N_4 microcrystals could allow a more efficient electron transfer, and reduce the proportion of accumulated unreactive electrons, which is presumably key to the increased photoactivity. The electron and hole transfer rate constants (k_{et} and k_{ht}) can be estimated using the following expression:¹⁶

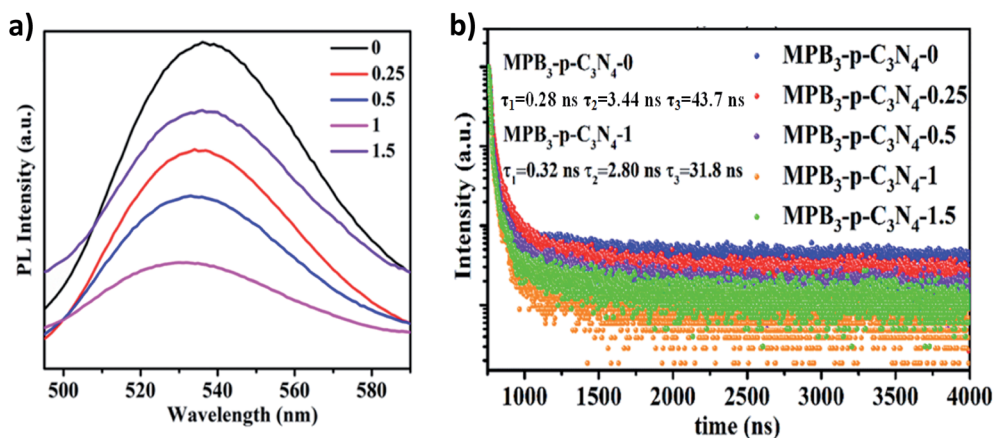


Fig. 8 (a) Steady-state PL spectra and (b) time-resolved photoluminescence decay spectra (TRPL) of $\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4$ (loading amount: 0–1.5 mg) samples.



$$k_{\text{et}}(\text{p-C}_3\text{N}_4 \rightarrow \text{MAPbBr}_3) = \frac{1}{\tau_{\text{I}}(\text{MAPbBr}_3/\text{p-C}_3\text{N}_4)} - \frac{1}{\tau_{\text{I}}(\text{p-C}_3\text{N}_4)} \quad (6)$$

$$k_{\text{ht}}(\text{p-C}_3\text{N}_4 \rightarrow \text{MAPbBr}_3) = \frac{1}{\tau_{\text{av}}(\text{MAPbBr}_3/\text{p-C}_3\text{N}_4)} - \frac{1}{\tau_{\text{av}}(\text{p-C}_3\text{N}_4)} \quad (7)$$

The calculated k_{et} and k_{ht} values for samples of MAPbBr₃/p-g-C₃N₄-1.0 were approximately $2.86 \times 10^9 \text{ s}^{-1}$ and $4.90 \times 10^6 \text{ s}^{-1}$, respectively. This k_{et} value is 4.65 fold higher than the reported k_{et} value ($6.15 \times 10^8 \text{ s}^{-1}$) by Pu *et al.*,¹⁶ indicating that highly crystalline MAPbBr₃ wrapped/entrapped p-g-C₃N₄ intergrowth microcrystals have great potential as new photocatalysts to convert solar energy into chemical energy.

3.5. Electrochemical analysis of MAPbBr₃/p-g-C₃N₄

Recently, there have been some literature studies on the rich solid state chemistry aspects of perovskites through electrochemical experiments^{9,31–34} where defects and grain boundaries as well as chemical changes occurring at the various interfaces could be monitored. To gain deeper insight into the electron transport behaviors, electrochemical impedance spectroscopy (EIS) was conducted in dichloromethane solution to determine the charge-transfer resistance of the electrodes. The spectra depicted the interfacial charge transfer efficiency of the samples with the radius of the arc correlating with the reaction rate at the surface of the electrode.³ Hence, as shown in Fig. 9a, a smaller diameter of the Nyquist semicircle at high frequencies is realized with MAPbBr₃/p-g-C₃N₄-1.0, implying that the charge transfer resistance is smaller than that of MAPbBr₃ and the other MAPbBr₃/p-g-C₃N₄ composites. Evidently, the pronounced decrease of the charge-transfer resistance by incorporating a suitable amount of p-g-C₃N₄ validates better charge transport ability and a more facilitated charge separation process of the MAPbBr₃/p-g-C₃N₄-1.0 sample.

On the other hand, the band edge positions of the perovskite materials were obtained from electrochemical experiments. Electrochemical Mott–Schottky (MS) analysis (Fig. 9b) was also

performed to estimate the flat band potential (V_{FB}) of different optically active perovskites, which can further illustrate their band structure. The Mott–Schottky equation can be written as follows (eqn (8))

$$\frac{1}{C^2} = \frac{2}{N_{\text{D}}q\epsilon\epsilon_0} \left(E - E_{\text{fb}} - \frac{k_{\text{B}} T}{q} \right) \quad (8)$$

where C denotes the space charge capacitance, N_{D} is the density of charge carriers, q is the electronic charge ($1.602 \times 10^{-19} \text{ C}$), ϵ and ϵ_0 refer to the dielectric constant of the semiconductor and permittivity in a vacuum ($8.85 \times 10^{-12} \text{ F cm}^{-2}$), respectively, E and E_{fb} denote the applied potential and flat-band potential, respectively, k_{B} is the Boltzmann constant, and T is the temperature in Kelvin. The value of E_{fb} is determined from the x -intercept of the linear section of the Mott–Schottky curve. For instance, Gimenez's research group demonstrated that a Mott–Schottky plot on a CsPbBr₃ quantum dots/c-TiO₂/FTO film showed n-type behavior and a flat band potential of -0.109 V vs. NHE .⁹ Furthermore, Li *et al.* found that the position of the valence band maximum (VBM) could be determined from the MS plots. The V_{FB} of MAPbBr₃ is $1.07\text{--}1.23 \text{ V vs. NHE}$, which is 0.4 V more positive than V_{FB} of MAPbI₃.³³ Similarly, Samu *et al.* reported that the VB edge of MAPbI_xBr_{1-x} becomes significantly more positive upon increasing Br incorporation into the lattice ($0.77 \rightarrow 1.15 \text{ V vs. Ag/AgCl}$).³⁴ It is generally accepted that the CB and VB potentials in n or p-type semiconductors are approximately equal to the flat band potential and the potentials vs. Ag/AgCl were converted to normal hydrogen electrode (NHE) potentials using the Nernst equation^{35,36}

$$V_{\text{NHE}} = V_{\text{Ag/AgCl}} + V_{\text{Ag/AgCl vs. NHE}}^{\circ} \quad (9)$$

where $V_{\text{Ag/AgCl vs. NHE}}^{\circ}$ is 0.199 V at $25 \text{ }^{\circ}\text{C}$. Fig. 6d shows n-type semiconducting features and the flat band potentials of MAPbBr₃ and p-g-C₃N₄ are -1.03 V and $-1.34 \text{ V vs. Ag/AgCl}$ derived from the intersection. Gelderman *et al.* have described the nature of the semiconductor–electrolyte interface, together with the Mott–Schottky equation for determining the flat-band potential which is approximately equal to that of the

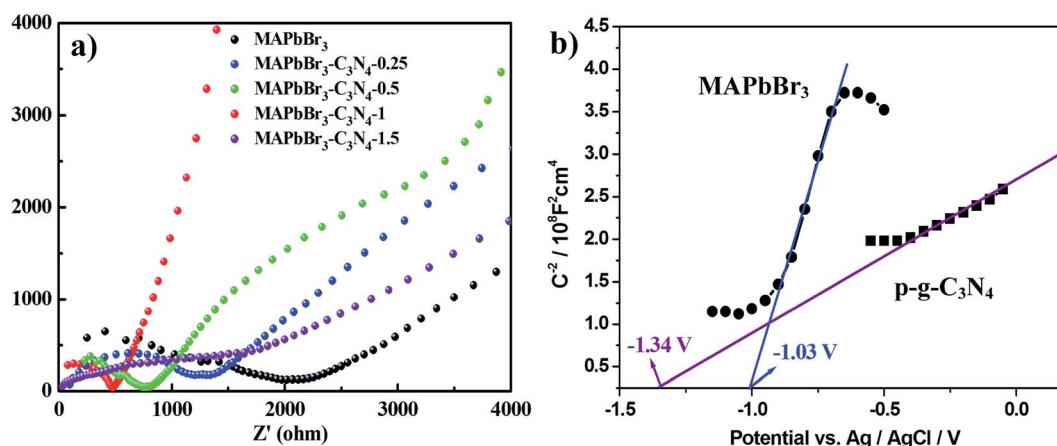


Fig. 9 (a) Electrochemical impedance spectroscopy (EIS) spectra of the MAPbBr₃/p-g-C₃N₄ composites and (b) Mott–Schottky plot on MAPbBr₃ and p-g-C₃N₄, showing n-type semiconducting features and flat band potentials of -1.03 V and $-1.34 \text{ V vs. Ag/AgCl}$, respectively.



Fermi level.³⁷ The Fermi level is defined as the energy level at which the probability of occupation by an electron is 1/2, and it lies at the mid-point of the band gap in an intrinsic semiconductor. But doping could change the Fermi level due to the redistribution of electrons within the solid. The Fermi level of an n-type semiconductor lies just below the conduction band, while for a p-type semiconductor it lies just above the valence band.^{37,38} Consequently, the Fermi level (flat-band potential) of a p-type catalyst is close to the VB while that of an n-type sample is located near the CB, which was also demonstrated for porous few-layer carbon nitride by Xiao *et al.*³⁶ and a NiO/g-C₃N₄ heterojunction by Tang *et al.*³⁹ Therefore, the potentials of the VB of MAPbBr₃ are calculated to be ~1.1 eV by subtracting the optical band gap (2.13 eV) from the CBM, which is in reasonably good agreement with those reported by Li *et al.*³³ and Samu *et al.*³⁴ Additionally, the flat-band potential of p-g-C₃N₄-Sol was measured to be -1.34 V *versus* Ag/AgCl, which is similar to the value of ultrathin g-C₃N₄ (-1.32 V *vs.* Ag/AgCl)³⁶ and slightly smaller than the value (-1.43 V) reported by Wang *et al.* So, the potentials of the VB of p-g-C₃N₄ are calculated to be ~1.4 eV by subtracting the optical band gap (2.7 eV)^{16,20} from the CB, which is slightly bigger than that of MAPbBr₃ (1.1 eV).

Generally, based on the bandgap and electronic energy level of the semiconductors, there are five different cases for semiconductor heterojunctions: straddling alignment (type-I), staggered alignment (type-II), Z-scheme system, p-n heterojunctions, and homojunctions.⁴⁰ In type-I band alignment composed of two kinds of semiconductors, both VB and CB edges of one semiconductor are localized within the energy gap of another semiconductor. Recently, Zhao *et al.* reported that the Ag-CsPbBr₃/CN composite has two electron transfer pathways. In fact, the first pathway is analogous to the type-I model, where the CB and VB edge potentials of pure CN are about -1.33 and 1.40 eV, and the CB and VB edge potentials of CsPbBr₃ are about -1.13 and 1.1 eV, respectively.⁴⁰ In our present work, the MAPbBr₃/p-g-C₃N₄ intergrowth composites have a structure similar to this so-called inverted type-I core/shell structure, where a material with a narrower band gap is grown epitaxially around the core material with a higher band gap.⁴⁰ Both electrons and holes would be rationally driven to the shell by the built-in energy potential, and the charge carriers (electrons and holes) are distributed largely in the shell region, which promotes separation of the photoexcited charge carriers and efficiently facilitates charge transfer to the shell layer surface to enhance the redox reaction. In contrast, with contact of p-type and n-type semiconductors with each other, the bands of the semiconductors will bend and the Fermi levels will equilibrate because of the formation of a space charge region after the diffusion of electrons and holes. Thus, the built-in electrical potential in the space charge region from the n-type side to the p-type side can direct the electrons and holes to quickly travel in the opposite directions, and allow more effective separation and longer lifetime of electron-hole pairs.⁴⁰ These advantages endow the p-n type heterostructures with an enhanced photocatalytic performance. For example, when p-type NiO and n-type g-C₃N₄ are connected closely, the electrons in n-type g-C₃N₄ transfer to p-type NiO, leading to a positive charge; meanwhile, the holes in NiO transfer to g-C₃N₄, generating a negative charge.

Consequently, the space charge region called the internal electric field in p-n heterojunctions is established when the electrons and holes accumulate until reaching an equilibrium, which could act as a potential barrier to decrease the electron-hole pair recombination.³⁹ Based on the Mott-Schottky (MS) analysis, the corresponding inverted type-I band configuration structure alignments of MAPbBr₃ and p-g-C₃N₄ are schematically depicted in Fig. 10. In other words, in MAPbBr₃ microcrystal wrapped/entrapped p-g-C₃N₄ microstructures, MAPbBr₃ with a smaller band gap is the shell and p-g-C₃N₄ with a bigger band gap is the core, which is evidently different from the work of Pu *et al.*¹⁶ Both electrons and holes would rationally move from the p-g-C₃N₄ core to the MAPbBr₃ shell driven by the built-in energy potential, and then diffuse to the MAPbBr₃ shell layer surface to enhance the redox reaction, which promotes separation of the photoexcited charge carriers (Fig. 10b).

Meanwhile, based on the slopes of the Mott-Schottky plots, the density of the charge carrier, N_D , can also be calculated from eqn (10).

$$N_D = \frac{2}{q\epsilon\epsilon_0} \frac{dE}{d\frac{1}{C^2}} = \frac{2}{q\epsilon\epsilon_0} \frac{1}{\text{slope}} \quad (10)$$

The relative ϵ dielectric constant of a MAPbBr₃ single crystal is about 25.5.²⁷ Therefore, the N_D of MAPbBr₃ ($5.43 \times 10^{16} \text{ cm}^{-3}$) was found to be comparable to previously reported values of MAPbI₃ (on the order of $\sim 10^{16} \text{ cm}^{-3}$)³³ and was also in line with our Hall effect measurement results and smaller than the donor density of n-type CsPbBr₃ films ($6.77 \times 10^{19} \text{ cm}^{-3}$).⁹ The deviation is most probably due to the action of surface states in the polycrystalline electrode capturing and immobilizing the carriers.

Moreover, Hall effect measurements were also carried out at room temperature, showing the n-type semiconductor behavior for all MAPbBr₃/p-g-C₃N₄ samples. Both the samples have a carrier concentration of 10^{16} cm^{-3} , as shown in Table S2,† which is in agreement with the Mott-Schottky results. This mobility values are of the same order of magnitude compared to the previous result of a MAPbBr₃ single crystal (up to $60 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$).²⁷ Our results are in accord with the fact that g-C₃N₄ doping increases the conductivity and carrier mobility in the perovskite film.¹⁹ With regard to MAPbBr₃, another important parameter is the carrier diffusion length, L_D (the distance over which the limiting photocarrier diffuses before it recombines), so we can calculate the carrier diffusion length by combining carrier lifetime with carrier mobility given by the following equation:^{27,28}

$$L_D = (\mu\tau k_B T e)^{1/2} \quad (11)$$

where Boltzmann's constant $k_B = 1.38 \times 10^{-23} \text{ J K}^{-1}$ and T is the sample's temperature (298 K). By using the average carrier lifetime (τ_{av}) obtained from the present work and carrier mobility $\mu = 37.21\text{--}165.02 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, the carrier diffusion lengths of $L_D(\text{MAPbBr}_3)$ and $L_D(\text{MAPbBr}_3/\text{p-g-C}_3\text{N}_4\text{-1.0})$ are found to be 6.44 and 9.39 μm , respectively, which is similar to a previously reported value of 4.3 μm for MAPbBr₃,²⁷ and an estimated 3–17 μm in MAPbBr₃ using a shorter lifetime and lower mobility.²⁸ Such long carrier diffusion lengths will be



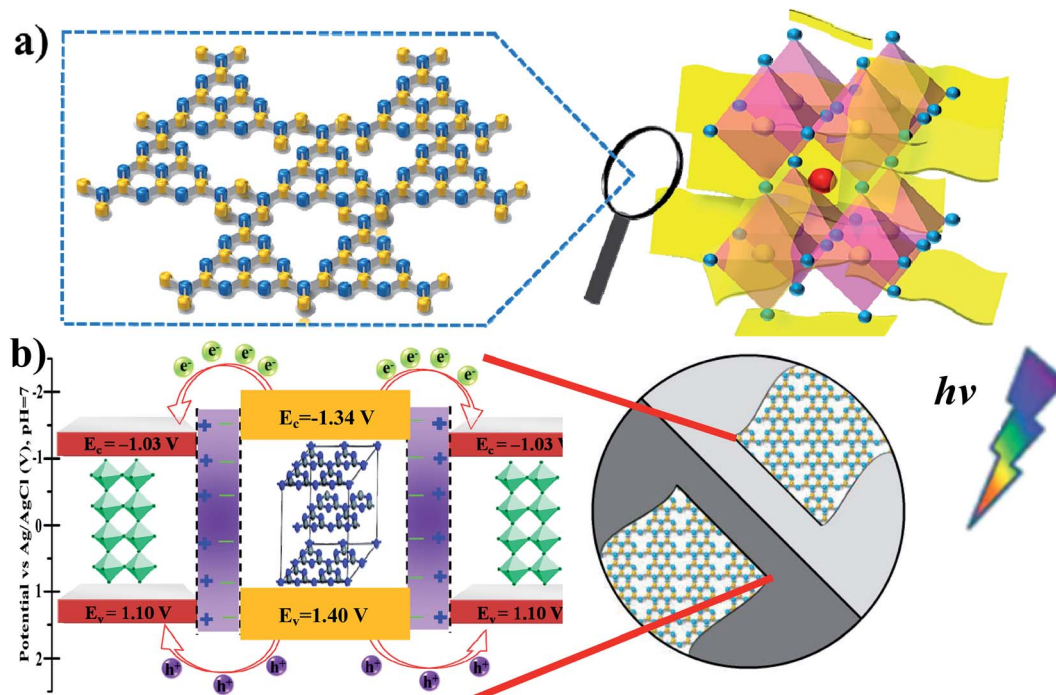


Fig. 10 (a) Schematic illustrations of the MAPbBr₃/p-g-C₃N₄ intergrowth microcrystal composite wrapped/entrapped nanosheets on the molecular level. (b) The proposed corresponding inverted type-I band configuration structure alignments of MAPbBr₃ and p-g-C₃N₄, which are similar to core/shell heterostructure band structures.

favourable for photocarrier diffusion and separation for an efficient halide perovskite microcrystal photocatalyst.

3.6. Stability of the as-prepared samples

As long-term stability is also an important metric for photocatalytic systems, the recycling experiment runs for MG degradation under visible light irradiation were further tested in Fig. 11a and b. Compared with pure MAPbBr₃ (Fig. 11a), no evident reduction in activity is observed in the MAPbBr₃/p-g-C₃N₄-1.0 composite sample during at least ten reaction cycles (Fig. 11b), indicating the excellent chemical stability of the photocatalyst under testing conditions. The degradation rates were slightly decreased because of the inevitable loss of the photocatalyst mass during the recycling process. Since it is widely accepted that microcrystals are more stable than nanocrystals due to low surface Gibbs energy, it is reasonable that our samples have desirable high stability. In order to further evaluate the stability of the MAPbBr₃/p-g-C₃N₄ samples, FTIR, XRD and SEM measurements after the recycling experiment were employed. The FTIR spectra of IPA and MAPbBr₃/p-g-C₃N₄-1.0 were identical before and after the stability test (Fig. S5[†]), showing no obvious changes in chemical functional groups for the solvent and catalyst. Simultaneously, no appreciable change in the crystal structure and morphology of MAPbBr₃/p-g-C₃N₄-1.0 can be observed, suggesting that the chemical stability of the MAPbBr₃/p-g-C₃N₄ sample is high, and is comparable to the stability reported for double-perovskite Cs₂AgBiBr₆.¹⁰ Clearly, the as-synthesized MAPbBr₃/p-g-C₃N₄ composite could thus be an effective and stable catalyst for photocatalytic reaction. On



Fig. 11 (a and b) Long-term cycling stability tests of the MAPbBr₃ and MAPbBr₃/p-g-C₃N₄-1.0 under visible-light irradiation for photoreduction of malachite green.



the other hand, for electrocatalyst experiments, the $I-t$ polarization curve after many CV cycles in alkaline solution is usually used to evaluate the long-term stability. The related electrocatalyst work on the MAPbBr₃/p-g-C₃N₄ composite is under investigation. Besides the degradation of MG, our sample exhibited a good photocatalytic performance for other cationic (RhB) and anionic (MO) dye pollutants (Fig. S6†), and the solution turned completely colorless within 20 min. To sum up, it is believed that the N-Br chemical bonding interaction between MAPbBr₃ and p-g-C₃N₄ not only enhances charge separation and transportation efficiency, but also provides an alternative way for surface passivation of uncoordinated halides to improve their chemical stability, exhibiting huge potential for thermocatalysis, electrocatalysis and photocatalysis.

4. Conclusion

In summary, we demonstrated that the construction of a MAPbBr₃/p-g-C₃N₄ intergrowth composite photocatalyst was realized by the abundant NH_x sites along the edges of p-g-C₃N₄ bonded with MAPbBr₃. The spectroscopic and electrochemical results showed that the unique chemical covalent C-N-Br interaction at the interface could eliminate the defects and enhance the stability of MAPbBr₃/p-g-C₃N₄, which facilitated the charge separation in the photocatalytic process. The photocatalytic activity (k_{MG}) and interfacial charge transfer efficiency (k_{et}) in the inverted type-I core/shell MAPbBr₃/p-g-C₃N₄ band configuration structure were strongly dependent on the composition of p-g-C₃N₄ nanosheets. The MAPbBr₃/g-C₃N₄ composites exhibited outstanding photocatalytic degradation of malachite green, with 15 and 3 times higher activity compared to their pure constituents, respectively. The photoexcited electron induced radical anions ($\cdot O_2^-$) were the major species for MG degradation. Furthermore, the MAPbBr₃/g-C₃N₄ composites were found to be rather stable for at least ten cycles, revealing their promise for utilization in long-term photocatalysis. The excellent activity and good stability could enable this hybrid system to be a highly efficient visible light-driven photocatalyst for practical harvesting of energy from sunlight.

Conflicts of interest

There are no conflicts to declare.

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