## **RSC Advances**



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Cite this: RSC Adv., 2019, 9, 10414

# A highly sensitive and selective fluorescent probe for quantitative detection of Al<sup>3+</sup> in food, water, and living cells†

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Three novel  $\beta$ -pinene-based fluorescent probes 2a-2c were designed and synthesized for the selective detection of  $Al^{3+}$ . Probe 2a showed higher fluorescence intensity toward  $Al^{3+}$  than the other two compounds. Probe 2a determined the concentration of  $Al^{3+}$  with a rapid response time (45 s), wide pH range (pH = 1-9), excellent sensitivity (LOD =  $8.1 \times 10^{-8}$  M) and good selectivity. The recognition mechanism of probe 2a toward  $Al^{3+}$  was confirmed by  $^{1}H$  NMR, HRMS and DFT analysis. Probe 2a was successfully used as a signal tool to quantitatively detect  $Al^{3+}$  in food samples and environmental water samples. Furthermore, probe 2a was successfully utilized to label intracellular  $Al^{3+}$ , indicating its promising applications in living cells.

Received 18th January 2019 Accepted 20th March 2019

DOI: 10.1039/c9ra00447e

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#### 1. Introduction

As the third most abundant element and the most widely used metal ion on Earth, aluminum is extensively used in a variety of fields, but is harmful to the environment and living systems. <sup>1-4</sup> The increasing concentration of Al<sup>3+</sup> in pollutants can deeply influence the growth of plants, and lead to soil acidification<sup>5,6</sup> and underground water contamination. <sup>7</sup> Moreover, the accumulation of Al<sup>3+</sup> in the body can greatly affect the absorption of calcium in bone tissue and induce several diseases such as Alzheimer's, Parkinson's epilepsy, seizures, and renal and liver damage. <sup>8-11</sup> In 1989, the World Health Organization (WHO) and the United Nations Food and Agriculture Organization (UNFAO) identified Al<sup>3+</sup> as a food pollutant to be controlled. In 2011, WHO/UNFAO revised the weekly allowable intake of Al<sup>3+</sup> from 7 mg kg<sup>-1</sup> to 2 mg kg<sup>-1</sup>. <sup>12</sup> Therefore, it is significant to detect and control the Al<sup>3+</sup> concentration in water and food samples.

Compared to traditional detection methods, a fluorescent probe has become one of the most widely used tools for detecting metal ions. <sup>13–15</sup> In the past few years, many fluorescent probes for Al<sup>3+</sup> detection have been reported. <sup>16–25</sup> However, there are still some shortcomings in reported Al<sup>3+</sup> fluorescent probes, including a complex synthesis process, poor selectivity and sensitivity, easy interference by other metal ions, Zn<sup>2+</sup>, Cu<sup>2+</sup>, Cr<sup>3+</sup>, Hg<sup>2+</sup>, and even F<sup>-</sup>, <sup>26–30</sup> and lack of applicability to water samples

Nopinone is obtained by the oxidation of  $\beta$ -pinene, which is a primary ingredient in natural turpentine. It is often used in the production of medicine and perfume. In addition, the rigid structure of the nopinone molecule can reduce the energy loss of non-radiative transitions in the fluorescence emission process. Furthermore, molecules with a nopinone skeleton structure have good biological compatibility and low cytotoxicity. Thus, the development of novel fluorescence probes from nopinone is very promising.

In this paper, we synthesized three new indazole derivatives for the specific detection of  $Al^{3+}$ , using natural nopinone as the starting material. Probes **2a–2c** could be synthesized by a simple two-step reaction and displayed a rapid ratiometric fluorescence toward  $Al^{3+}$  in aqueous solution (pH = 7.4). In addition, probe **2a** had a high selectivity for  $Al^{3+}$  over other metal ions and a rapid response time. The detection limit of probe **2a** was found to be 8.1  $\times$  10<sup>-8</sup> M, which is lower than that of many reported  $Al^{3+}$  fluorescence probes. The <sup>1</sup>H NMR, HRMS and theoretical calculations showed the detection mechanism. Furthermore, probe **2a** was successfully proved for the quantitative detection of  $Al^{3+}$  in food and environmental water samples. More importantly, cell experiments also demonstrated that probe **2a** could be used as a signal tool to detect the concentration of  $Al^{3+}$  in living cells.

### 2. Experimental

#### 2.1. General information

All reagents and solvents were of analytical grade and bought from commercial sources. UV-Vis absorption spectra were

and food samples.<sup>31-35</sup> Therefore, developing a fluorescence probe with high sensitivity, good selectivity, rapid response and low toxicity for detecting Al<sup>3+</sup> is particularly meaningful.

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<sup>†</sup> Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra00447e

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recorded by a PerkinElmer Lambda 950. PL emission spectra were tested using PerkinElmer LS55. The <sup>1</sup>H NMR and <sup>13</sup>C NMR spectra were recorded in CDCl<sub>3</sub> solutions on a Bruker AV 400 spectrometer. The purity of the synthesized compounds was recorded by America Agilent 1260 Infinity liquid chromatography. All pH measurements were recorded using a Sartorius Basic pH-Meter PB-20. High-resolution mass spectra (HRMS) were tested by an America Agilent 5975c mass spectrometer. Melting points were recorded using an X-6 microscopic melting point apparatus. All measurements were performed at room temperature. The  $5.0 \times 10^{-5}$  M solutions of various metal ions (Ag<sup>+</sup>, Hg<sup>2+</sup>, Na<sup>+</sup>, Mg<sup>2+</sup>, K<sup>+</sup>, Ca<sup>2+</sup>, Fe<sup>2+</sup>, Fe<sup>3+</sup>, Co<sup>2+</sup>, Ni<sup>2+</sup>, Cu<sup>2+</sup>, Zn<sup>2+</sup>, Ba<sup>2+</sup>, Sn<sup>2+</sup>, Pb<sup>2+</sup>, Cr<sup>2+</sup>, Pb<sup>2+</sup>, Bi<sup>3+</sup> and Al<sup>3+</sup>) were prepared in deionized water from their nitrate, chloride or sulfate, and stored at room temperature.

#### 2.2. Synthesis

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2.2.1. Synthesis of compounds 1a-1c. Under nitrogen atmosphere, nopinone (2 mol) was aroylated with a-c (3 mol) catalyzed by NaH (6 mol) in 1,2-dimethoxyethane. After cooling to room temperature, 20 ml of distilled water was added to the reaction mixture, and it was extracted with ethyl acetate. The ethyl acetate phase was washed with distilled water to neutrality. After recovering the solvent, the obtained crude 1a was purified by column chromatography (100-200 mesh silica gel, eluent: PE : EA = 10 : 1, v/v) to obtain compounds 1a-1c.

Compound 1a was a pale yellow grease, yield: 71.3%; <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>)  $\delta$ : 15.58 (s, 1H), 8.68 (dd, J = 4.9, 1.7 Hz, 1H), 8.00 (d, J = 8.0 Hz, 1H), 7.88-7.72 (m, 1H), 7.35-7.30 (m, 1H), 3.28-2.95 (m, 2H), 2.59 (dt, J = 22.5, 5.4 Hz, 2H), 2.32 (qt, J $= 5.9, 2.6 \text{ Hz}, 1\text{H}), 1.49 (d, J = 9.9 \text{ Hz}, 1\text{H}), 1.35 (s, 3\text{H}), 0.96 (s, 3\text{$ 3H); <sup>13</sup>C NMR (100 MHz, DMSO) δ: 153.35, 148.51, 137.18, 127.72, 124.90, 123.41, 121.44, 105.43, 57.40, 54.31, 48.29, 28.34, 27.26, 25.41, 21.18; HRMS (m/z):  $[M + H]^+$  calculated for  $C_{15}H_{17}NO_2 + H^+$ , 244.1341; found, 244.1335.

Compound **1b** was a pale yellow grease, yield: 73.8%; <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>)  $\delta$ : 15.41 (s, 1H), 8.95 (dd, J = 2.2, 0.9 Hz, 1H), 8.67 (dd, J = 4.9, 1.7 Hz, 1H), 8.03 (dt, J = 8.0, 2.0 Hz, 1H),  $7.40 \, (ddd, J = 7.9, 4.8, 0.9 \, Hz, 1H), 2.71 \, (dd, J = 4.0, 3.0 \, Hz, 2H),$ 2.65-2.53 (m, 2H), 2.32 (tt, J = 5.9, 3.1 Hz, 1H), 1.47 (d, J =9.5 Hz, 1H), 1.36 (s, 3H), 0.97 (s, 3H); <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>) δ: 209.55, 169.51, 150.69, 149.01, 135.67, 131.29, 123.24, 104.89, 54.84, 39.77, 39.61, 28.07, 27.66, 25.79, 21.54; HRMS (*m/z*): [M +  $H^+$  calculated for  $C_{15}H_{17}NO_2 + H^+$ , 244.1338; found, 244.1331.

Compound 1c was a pale yellow grease, yield: 76.5%; <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>)  $\delta$ : 15.26 (s, 1H), 8.78–8.67 (m, 2H), 7.63– 7.52 (m, 2H), 2.77–2.64 (m, 2H), 2.68–2.54 (m, 2H), 2.32 (tt, J =6.0, 3.1 Hz, 1H), 1.47 (d, J = 9.8 Hz, 1H), 1.36 (s, 3H), 0.97 (s, 3H); <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>)  $\delta$ : 209.98, 168.74, 150.06, 142.47, 122.14, 105.30, 54.92, 39.65, 39.62, 27.95, 27.60, 25.78, 21.55; HRMS (m/z):  $[M + H]^+$  calculated for  $C_{15}H_{17}NO_2 + H^+$ , 244.1339; found, 244.1332.

**2.2.2.** Synthesis of 2a–2c. 1a–1c (0.5 mmol),  $60\% N_2H_4 \cdot H_2O$ (1 mmol), and 10 ml of absolute ethanol were added into a round-bottom flask equipped with a thermometer and reflux condenser. The cyclization reaction was carried out at reflux for

4 h. After recovering the ethanol, the obtained crude 2a-2c was purified by column chromatography (100-200 mesh silica gel, eluent: PE : EA = 15 : 1, v/v) to obtain compounds 2a-2c.

Compound 2a was a white solid powder, yield: 85.6%, mp: 146.2–147.2 °C; <sup>1</sup>H NMR (400 MHz, DMSO)  $\delta$ : 12.59 (s, 1H), 8.61 (d, J = 4.8 Hz, 1H), 7.96-7.58 (m, 2H), 7.29 (t, J = 6.3 Hz, 1H),3.00 (dd, J = 16.3, 3.0 Hz, 1H), 2.89 (dd, J = 16.3, 2.6 Hz, 1H),2.80 (t, I = 5.2 Hz, 1H), 2.71 (dt, I = 9.3, 5.8 Hz, 1H), 2.29 (tt, I = 9.3) 5.6, 2.8 Hz, 1H), 1.39 (s, 3H), 1.27 (d, J = 9.3 Hz, 1H), 0.65 (s, 3H); <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>) δ: 160.79, 149.58, 149.04, 137.13, 136.77, 122.19, 120.15, 110.67, 41.96, 41.30, 41.07, 32.52, 26.36, 26.32, 21.49; HRMS (m/z):  $[M + H]^+$  calculated for  $C_{15}H_{17}N_3 + H^+$ , 240.1501; found, 240.1494.

Compound 2b was a white solid powder, yield: 82.4%, mp: 150.5–151.5 °C; <sup>1</sup>H NMR (400 MHz, DMSO) δ: 12.61 (s, 1H), 8.94 (d, J = 2.3 Hz, 1H), 8.51 (dd, J = 4.8, 1.6 Hz, 1H), 8.07 (dt, J = 8.0,2.0 Hz, 1H), 7.48 (dd, J = 8.0, 4.8 Hz, 1H), 2.98 (dd, J = 15.8, 3.1 Hz, 1H), 2.91–2.80 (m, 2H), 2.71 (dt, J = 9.3, 5.8 Hz, 1H), 2.31 (dt, J = 5.7, 2.9 Hz, 1H), 1.39 (s, 3H), 1.29 (d, J = 9.3 Hz, 1H), 0.66(s, 3H);  $^{13}$ C NMR (100 MHz, CHCl<sub>3</sub>)  $\delta$ : 157.91, 148.11, 146.80, 137.91, 132.80, 128.07, 123.70, 110.22, 58.08, 41.41, 41.14, 32.37, 26.23, 26.01, 21.46; HRMS (m/z):  $[M + H]^+$  calculated for  $C_{15}H_{17}NO_2 + H^+$ , 240.1504; found, 240.1497.

Compound 2c was a white solid powder, yield: 87.1%, mp: 157.1–158.1 °C; <sup>1</sup>H NMR (400 MHz, DMSO) δ: 12.78 (s, 1H), 8.61 (d, J = 5.1 Hz, 2H), 7.65 (d, J = 5.2 Hz, 2H), 3.00 (dd, J = 15.9),3.1 Hz, 1H, 2.93-2.80 (m, 2H), 2.72 (dt, J = 9.4, 5.8 Hz, 1H), 2.32 (dt, J = 9.4, 5.8 Hz, 2.82 (dt, J = 9.4, 5.8 Hz)(dt, J = 5.7, 2.9 Hz, 1H), 1.40 (s, 3H), 1.28 (d, J = 9.4 Hz, 1H), 0.64(s, 3H);  $^{13}$ C NMR (100 MHz, CDCl<sub>3</sub>)  $\delta$ : 157.96, 150.07, 139.12, 119.83, 111.60, 41.39, 41.33, 41.14, 32.36, 26.25, 26.21, 21.46, 18.39; HRMS (m/z):  $[M + H]^+$  calculated for  $C_{15}H_{17}NO_2 + H^+$ , 240.1498; found, 240.1496.

#### 2.3. Cell culture and imaging

HeLa cells were incubated with probe 2a  $(5.0 \times 10^{-6} \text{ M})$  in Dulbecco's Modified Eagle's Medium (DMEM) for 24 h at 37 °C. The cells were treated with aluminum chloride (5.0  $\times$  10<sup>-5</sup> M). After incubation for 1 h, the cells were washed with phosphatebuffered saline (PBS) three times, and the treated cells were used for cell imaging. The cell images were obtained using a confocal microscope at an excitation wavelength between 300 and 340 nm. Subsequently, in the control experiment, HeLa cells were only incubated with 2a (5.0  $\times$  10<sup>-6</sup> M), and then washed with PBS three times for cell imaging. The cell images were obtained with a confocal microscope at an excitation wavelength between 300 and 340 nm.

#### 3. Results and discussion

#### Synthesis and structural characteristics of probes 2a-2c

The synthetic procedure and molecular structure of probes 2a-2c are shown in Scheme 1. The probes 2a-2c were synthesized by two steps including the Claisen condensation of a-c with the βpinene derivative nopinone catalyzed by NaH in a solution of 1,2dimethoxyethane at reflux to obtain 1a-1c, and 1a-1c was further condensed with 60% N<sub>2</sub>H<sub>4</sub>·H<sub>2</sub>O at reflux to obtain 2a-2c.

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Scheme 1 Synthesis of probe 2a-2c.

# 3.2. Fluorescence properties of 2a-2c toward different metals

The selectivity behavior of probes 2a–2c toward various metal ions, including  $K^+$ ,  $Ba^{2+}$ ,  $Na^+$ ,  $Cr^{2+}$ ,  $Co^{2+}$ ,  $Mg^{2+}$ ,  $Fe^{2+}$ ,  $Fe^{3+}$ ,  $Mn^{2+}$ ,  $Ca^{2+}$ ,  $Pb^{2+}$ ,  $Cu^{2+}$ ,  $Ag^+$ ,  $Pb^{2+}$ ,  $Zn^{2+}$ ,  $Bi^{3+}$ ,  $Ni^{2+}$ ,  $Sn^{2+}$ , and  $Al^{3+}$  was studied by fluorescence spectroscopy. The changes in the fluorescence spectroscopy of 2a–2c before and after the addition of various metal ions are shown in Fig. 1. This shows that compound 2a exhibited the best selectivity and highest fluorescence intensity toward  $Al^{3+}$ , which implies a strong coordination ability between 2a and  $Al^{3+}$ . The changes in the UV-Vis absorption spectra of 2a before and after the addition of various metal ions are shown in Fig. S1.† Upon addition of  $Al^{3+}$  (10 equiv.) to a solution of probe a, the absorption peak at a and a malmost disappears and the peak at a 30 nm is greatly enhanced. These results show that probe a could be used as a fluorescent probe to selectively detect the presence of  $al^{3+}$ .

#### 3.3. Optimization studies of probe toward Al<sup>3+</sup>

The detection conditions for probe **2a** toward  $Al^{3+}$  were optimized by investigating the influence of the concentration of **2a**, pH range, response time, and EtOH/HEPES buffer. The fluorescent intensity can reach a steady state after adding  $Al^{3+}$  into a solution of **2a** for 45 s (Fig. S2C†). Further tests for determining the selectivity and sensitivity of probe **2a** toward  $Al^{3+}$  were performed in aqueous buffer solution (EtOH/HEPES buffer, 10 mM, v/v = 6/4, pH = 7) (Fig. S2†); the concentration

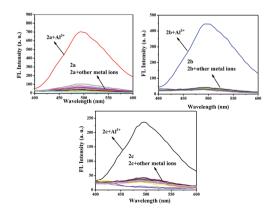


Fig. 1 Fluorescence emission spectra of 2a–2c (5.0  $\times$  10<sup>-6</sup> M) upon the addition of 10 equiv. of various metal ions in C<sub>2</sub>H<sub>5</sub>OH solution.

of probe 2a was  $5.0 \times 10^{-6}$  M (Fig. S2D†). As shown in Fig. S2A,† significant enhancement in fluorescence intensity was observed in the pH range 1–9, which indicated that it is suitable for application in living systems. We studied the photo-stability of the 2a-Al³+ complex in aqueous buffer solution (EtOH/HEPES buffer, 10 mM, v/v = 6/4, pH = 7). After continuous illumination for 60 h, the fluorescence intensity of the 2a-Al³+ complex did not show any obvious change in fluorescence (Fig. S3†). There was good photo-stability of the 2a-Al³+ complex, indicating that probe 2a could be identified as a practical method for Al³+ discrimination.

#### 3.4. Competitive selectivity of the 2a toward Al<sup>3+</sup>

To further investigate the selectivity of **2a** as a fluorescence probe for Al<sup>3+</sup>, competition experiments were carried out in the presence of Al<sup>3+</sup> mixed with other metal ions, such as K<sup>+</sup>, Ba<sup>2+</sup>, Na<sup>+</sup>, Cr<sup>2+</sup>, Co<sup>2+</sup>, Mg<sup>2+</sup>, Fe<sup>2+</sup>, Fe<sup>3+</sup>, Mn<sup>2+</sup>, Ca<sup>2+</sup>, Pb<sup>2+</sup>, Cu<sup>2+</sup>, Ag<sup>+</sup>, Pb<sup>2+</sup>, Zn<sup>2+</sup>, Bi<sup>3+</sup>, Ni<sup>2+</sup>, and Sn<sup>2+</sup>. As shown in Fig. 2, other metal ions had very little influence on the fluorescence intensity of the **2a**–Al<sup>3+</sup> complex. Combined with the data in Fig. 1, this demonstrates that probe **2a** has very high selectivity toward Al<sup>3+</sup>.

#### 3.5. Sensitivity behavior of 2a toward Al<sup>3+</sup>

The sensitivity of probe 2a toward Al3+ was examined with a fluorescence titration method, and the fluorescence titration spectra of probe 2a toward Al<sup>3+</sup> are shown in Fig. 3. As shown in Fig. 4B, the  $5.0 \times 10^{-6}$  M probe 2a solution (EtOH/HEPES buffer, v/v = 6/4, 10 mM HEPES, pH = 7.4) exhibited nonfluorescence. Upon the addition of Al3+ into the probe 2a solution (5.0  $\times$  10<sup>-6</sup> M), a green fluorescence dramatically appeared. Fig. 4A reveals that the fluorescence intensities at 495 nm increased linearly between the fluorescence intensity and the low Al<sup>3+</sup> concentration in the range  $0-1.2 \times 10^{-5}$  M, y =63.43x + 50.82,  $R^2 = 0.9908$  (fluorescence quantum yield  $\Phi =$ 0.49, when the concentration of Al $^{3+}$  was  $1.5 \times 10^{-5}$  M). The detection limit (LOD) for Al<sup>3+</sup> was found to be  $8.1 \times 10^{-8}$  M by using DL =  $3\sigma/k$  (where DL is the detection limit,  $\sigma$  is the standard deviation of the blank solution and k is the slope of the calibration plot). The association constant  $(K_B)$  of probe 2a with  $Al^{3+}$  was determined to be 1.89  $\times$  10<sup>3</sup> M<sup>-1</sup> via the Benesi-Hildebrand equation<sup>37-39</sup> (see Fig. 4C). Table S1† summarizes the

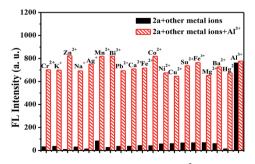


Fig. 2 Fluorescence intensity of **2a** (5.0  $\times$   $10^{-6}$  M) in buffer solution (EtOH/HEPES buffer, v/v = 6/4, 10 mM HEPES, pH = 7.4) and its complexes with Al $^{3+}$  (5.0  $\times$   $10^{-5}$  M) in the presence of various metal ions (5.0  $\times$   $10^{-5}$  M),  $\lambda_{\rm ex} = 330$  nm.

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900 800 700 400 450 500 500 500 600 Wavelength (nm)

Fig. 3 Fluorescence spectral changes in 2a (5.0  $\times$  10<sup>-6</sup> M) upon addition of Al<sup>3+</sup> (0–1.5  $\times$  10<sup>-5</sup> M) in solution (EtOH/HEPES buffer, v/v = 6/4, 10 mM HEPES, pH = 7.4),  $\lambda_{ex}=330$  nm.

detection limits of recently reported Al<sup>3+</sup> fluorescent sensors and highlights their applications in food samples.<sup>40-46</sup> The detection limit of **2a** toward Al<sup>3+</sup> is the lowest among these reported probes, implying that probe **2a** can straightforwardly detect the concentration of Al<sup>3+</sup> in water samples and food samples.

# 3.6. Binding ratio and detection mechanism between probe 2a and Al<sup>3+</sup>

A Job plot experiment was carried out to determine the stoichiometry between 2a and  $Al^{3^+}$ . As shown in Fig. S4,† the stoichiometry ratio of 2a to  $Al^{3^+}$  was found to be 1:1. The binding mode of probe 2a toward  $Al^{3^+}$  was confirmed by  $^1H$  NMR experiments in DMSO, as shown in Fig. 5. In the presence of 1.0 equiv. of  $Al^{3^+}$ , the proton signal of pyrazole ( $H_1$ ) disappeared and the proton signal of the pyridine moiety shifted upfield. So, the sensing mechanism of probe 2a toward  $Al^{3^+}$  could be the result of the synergistic complexation of the N atom in pyrazole and pyridine rings to  $Al^{3^+}$  with a 1:1 stoichiometry. From the HRMS spectra (Fig. S5†), the mass peak at m/z 359.1464 corresponds to  $[2a + Al^{3^+} + 2Cl^- + Na]^+$  (calculated at 359.2094). The proposed coordination mechanism is shown in Scheme 2. Furthermore, the energies of both the HOMO and LUMO of 2a

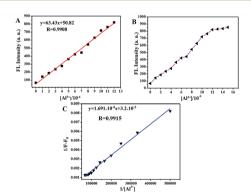


Fig. 4 (A) A linear increase in intensity at 495 nm of probe 2a ( $5.0 \times 10^{-6}$  M) with increasing concentrations of Al<sup>3+</sup> from 0 to  $1.2 \times 10^{-5}$  M,  $\lambda_{\rm ex} = 330$  nm. (B) A plot of the fluorescence intensity *versus* the concentrations of Al<sup>3+</sup> (0 to  $1.5 \times 10^{-5}$  M) in buffer solution (EtOH/HEPES buffer, v/v = 6/4, pH = 7.4). (C) Benesi–Hildebrand analysis of the emission changes for the complexation between 2a and Al<sup>3+</sup>.

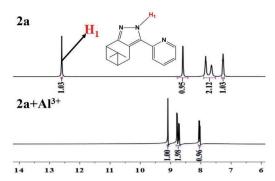


Fig. 5 <sup>1</sup>H NMR spectra changes in **2a** with the addition of Al<sup>3+</sup>

and the 2a-Al<sup>3+</sup> complex were calculated (Fig. S6†). The decrement in energy band gap confirms that there are obvious intramolecular charge-transfer (ICT) phenomena in 2a-Al<sup>3+</sup> complexes. Therefore, the calculated results were in good agreement with the emission wavelengths of 2a and 2a-Al<sup>3+</sup>.

#### 3.7. Preparation of the test strips

As shown in Fig. S7,† the test strips showed no fluorescence under 365 nm UV-lamp when they were prepared by soaking filter papers in an ethanol solution of  $2a~(2.0\times10^{-4}~M)$  and dried in air. When immersed in an aqueous solution of  $Al^{3+}~(2.0\times10^{-4}~M)$ , the test strips showed green fluorescence. Therefore, the 2a-based test strips show promising application for the detection of  $Al^{3+}$  in water by fluorimetric changes.

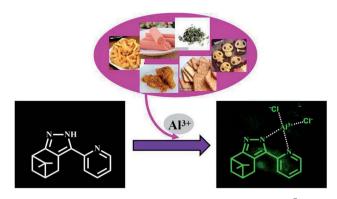
#### 3.8. Determination in different water samples

Novel probe 2a (5  $\times$  10<sup>-5</sup> M) was used for the detection of the concentration of Al<sup>3+</sup> in tap water, distilled water, and lake water samples. All the water samples were collected and simply filtered. As shown in Fig. S8,† a good linear relationship was obtained between the fluorescence intensity at 495 nm and the concentration of Al<sup>3+</sup> (0, 2, 5, 10, 15  $\times$  10<sup>-6</sup> M) in various water samples. The results listed in Table 1 show that in all the water samples recovery was higher than 95%. Therefore, the novel Al<sup>3+</sup> fluorescent probe can be used for detection of the concentration of Al<sup>3+</sup> in real water samples.

#### 3.9. Application in food samples

It is well known that  $Al^{3+}$  is widely used in food products. But, excessive ingested  $Al^{3+}$  may cause several diseases. Therefore, it is practically important to detect  $Al^{3+}$  in food products using this novel fluorescent probe. Some food samples containing  $Al^{3+}$ , such as chips, fried chicken, tea, sausage, biscuit and baby biscuit were chosen to examine the application of probe 2a (5 ×  $10^{-5}$  M) in food samples. These food samples were first crushed and 20% (v/v) HCl aqueous solution was added. Then it was stirred for one day until the solution became turbid. The mixture was filtered to obtain the  $Al^{3+}$ -containing filtrate. The fluorescence intensity at 495 nm displayed a good linear relationship with the concentration of  $Al^{3+}$  (0, 2, 5, 10, 15 ×  $10^{-6}$  M) ( $R^2 = 0.992$ , Fig. S9†). The results listed in Table 2 show that probe 2a can detect the concentration of  $Al^{3+}$  in different food

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Scheme 2 Proposed coordination mechanism of 2a with Al<sup>3+</sup>

Table 1 Application of 2a in the determination of Al<sup>3+</sup> in various water samples

Samples	Add (1 $\times$ 10 <sup>-6</sup> M)	Detected (1 $\times$ 10 <sup>-6</sup> M)	Recovery (%)
Tap water	0	0.62	0
	2	2.04	102.80
	5	5.15	103.20
	10	9.85	98.46
	15	14.93	99.54
Distilled water	0	0.17	0.00
	2	1.95	97.50
	5	5.11	102.20
	10	10.13	101.30
	15	14.69	97.96
Lake water	0	0.65	0.00
	2	1.92	96
	5	4.91	98.2
	10	10.06	100.64
	15	14.31	95.4

solutions with good recovery, ranging from 96 to 103%. Therefore, the novel Al<sup>3+</sup> probe 2a can be applied as a simple method to detect the concentration of Al<sup>3+</sup> in various food samples.

#### 3.10. Cellular imaging

Fluorescence imaging experiments of 2a were performed to study the utility of probe 2a in living cells. Cytotoxicity assays

Table 2 Results for the determination of Al<sup>3+</sup> in various food samples

	$Al^{3+}$	Added	Found	
Samples	$(1\times10^{-6}\text{ M})$	$(1\times10^{-6}\text{ M})$	$(1 \times 10^{-6} \text{ M})$	Recovery (%)
Chips	6.15	3	8.99	98.25
		6	12.22	100.57
Fried chicken	5.13	3	8.32	102.33
		6	10.96	98.47
Sausage	2.89	3	5.78	98.13
		6	8.57	96.40
Tea	1.47	3	4.29	95.97
		6	7.18	96.12
Biscuit	0.55	3	3.49	98.30
		6	6.33	96.64
Baby biscuit	0	3	3.11	103.67
		6	6.12	102

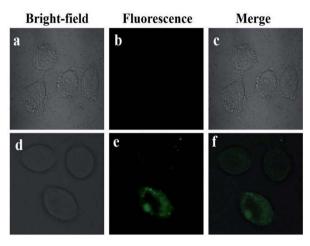


Fig. 6 (a) Fluorescent image of HeLa cells treated with probe 2a (5.0  $\times$  $10^{-6}$  M) in the absence of Al $^{3+}$ ; (b) microscope image of HeLa cells treated with probe 2a (5.0  $\times$  10<sup>-6</sup> M) in the absence of Al<sup>3+</sup>; (c) merged image of frames (a) and (b); (d) microscope image of HeLa cells treated with Al<sup>3+</sup> (5.0  $\times$  10<sup>-5</sup> M) and probe 2a (5.0  $\times$  10<sup>-6</sup> M); (e) fluorescence image of HeLa cells treated with Al $^{3+}$  (5.0 imes 10 $^{-5}$  M) and probe **2a** (5.0  $\times$  10<sup>-6</sup> M); (f) merged image of frames (d) and (e).

results showed that compound 2a had low cytotoxicity to HeLa cells (Fig. S10†). HeLa cells were incubated with 2a (5.0  $\times$  10<sup>-6</sup> M) at 37 °C for 24 h. And no obvious fluorescence was observed. However, after HeLa cells were incubated with AlCl<sub>3</sub> for 1 h, remarkable fluorescence enhancement can clearly be detected (Fig. 6). The fluorescence imaging experiments show that the novel Al<sup>3+</sup> fluorescent probe 2a has potential application in living cells.

#### Conclusions

In summary, we developed a β-pinene-based fluorescent probe 2a from natural  $\beta$ -pinene derivative nopinone for the detection of Al<sup>3+</sup>. The new fluorescent probe 2a exhibits advantages, such as clearer change in fluorescence, wider pH range, higher sensitivity, better selectivity, lower detection limit, and simpler synthetic procedures. The sensing mechanism of 2a with Al<sup>3+</sup> was studied by <sup>1</sup>H NMR, HRMS, and DFT analysis. Fluorescence probe 2a can be used as fluorescent sensing tool for the realtime detection of Al3+ in aqueous media, food samples, and living cells.

#### Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The research was supported by the National Natural Science Foundation of China (No. 31470592), Jiangsu Provincial Key Lab for the Chemistry and Utilization of Agro-Forest Biomass, the open Fund of Jiangsu Key Laboratory for Biomass Energy and Material (JSBEM201702), and Priority Academic Program Development of Jiangsu Higher Education Institutions, China.

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#### Notes and references

- 1 V. K. Gupta, A. K. Singh and N. Mergu, *Electrochim. Acta*, 2014, **117**, 405–412.
- 2 S. Sen, T. Mukherjee and B. Chattopadhyay, *Analyst*, 2012, 137, 3975–3981.
- 3 W. S. Miller, L. Zhuang and J. Bottema, *Mater. Sci. Eng.*, *A*, 2000, **280**, 37–49.
- 4 N. W. Baylor, W. Egan and P. Richman, *Vaccine*, 2002, 20, S18–S23.
- 5 D. Maity and T. Govindaraju, *Chem. Commun.*, 2010, **46**, 4499–4501.
- 6 S. Sooksin, V. Promarak and S. Ittisanronnachai, *Sens. Actuators, B*, 2018, **262**, 720–732.
- 7 J. L. Yang, L. Zhang and L. I. Ying, *Ann. Bot.*, 2006, **97**, 579–584.
- 8 J. Lee, H. Kim and S. Kim, Dyes Pigm., 2013, 96, 590-594.
- 9 S. Kim, J. Y. Noh, K. Y. Kim and J. H. Kim, *Inorg. Chem.*, 2012, 51, 3597–3602.
- 10 P. D. Darbre, J. Inorg. Biochem., 2005, 99, 1912-1919.
- 11 S. R. Paik, J. H. Lee and D. H. Kim, *Arch. Biochem. Biophys.*, 1997, **344**, 325–334.
- 12 J. Barceló and C. Poschenrieder, Environ. Exp. Bot., 2002, 48, 75–92.
- 13 G. X. Zhang, R. X. Ji and X. Y. Kong, RSC Adv., 2019, 9, 1147–1150
- 14 Y. Li, C. Y. Liao and S. S. Huang, *RSC Adv.*, 2016, **6**, 25420–25426.
- 15 T. L. Cui, S. Z. Yu and Z. J. Chen, RSC Adv., 2018, 8, 12276–12281.
- 16 D. P. Singh, R. Dwivedi and A. K. Singh, *Sens. Actuators, B*, 2017, 238, 128–137.
- 17 X. L. Yue, Z. Q. Wang and C. R. Li, *Spectrochim. Acta, Part A*, 2017, **193**, 415–421.
- 18 F. Wang, Y. Xu and S. O. Aderinto, *J. Photochem. Photobiol.*, *A*, 2017, 332, 273–282.
- 19 S. R. Gupta, P. Singh and B. Koch, *J. Photochem. Photobiol.*, *A*, 2017, 348, 246–254.
- 20 Y. Wang, Z. Y. Ma and D. L. Zhang, *Spectrochim. Acta, Part A*, 2018, **195**, 157–164.
- 21 H. Xie, Y. Wu and J. Huang, Talanta, 2016, 151, 8-13.
- 22 J. C. Qin, X. Y. Cheng and R. Fang, *Spectrochim. Acta, Part A*, 2016, **152**, 352–357.

- 23 Y. Zhang, Y. Fang and N. Z. Xu, *Chin. Chem. Lett.*, 2016, 11, 1673–1678.
- 24 P. G. Ding, J. H. Wang and J. Y. Cheng, *New J. Chem.*, 2015, **39**, 342–348.
- 25 X. Y. Li, M. M. Yu and F. L. Yang, New J. Chem., 2013, 37, 2257–2260.
- 26 X. Zhang, P. Sun and F. Li, *Sens. Actuators, B*, 2017, **255**, 366–373.
- 27 M. Tajbakhsh, G. B. Chalmardi and A. Bekhradnia, *Spectrochim. Acta, Part A*, 2018, **189**, 22–31.
- 28 S. Chemate and N. Sekar, Sens. Actuators, B, 2015, 220, 1196– 1204.
- 29 B. J. Pang, C. R. Li and Z. Y. Yang, *Spectrochim. Acta, Part A*, 2018, **204**, 641–647.
- 30 R. Lu, S. Cui and S. Li, Tetrahedron, 2017, 73, 915-922.
- 31 Q. Diao, P. Ma and L. Lv, Sens. Actuators, B, 2016, 229, 138-144.
- 32 C. Lim, H. Seo and J. H. Choi, *J. Photochem. Photobiol.*, *A*, 2018, **356**, 312–320.
- 33 C. Sun, J. Sun and F. Z. Qiu, Spectrochim. Acta, Part A, 2018, 188, 1–7.
- 34 H. Liang, Z. Li and D. Wu, Sens. Actuators, B, 2018, 269, 62-
- 35 Y. Tang, X. Kong and A. Xu, *Angew. Chem., Int. Ed.*, 2016, 55, 3356–3359.
- 36 Z. l. Wang, Y. Zhang and J. Song, *Dyes Pigm.*, 2019, **161**, 172–181.
- 37 K. Huang, X. Jiao and C. Liu, Dyes Pigm., 2017, 142, 437-446.
- 38 J. M. Jung, J. H. Kang and J. Han, Sens. Actuators, B, 2018, 267, 58-69.
- 39 Z. Y. Li, H. K. Su and K. Zhou, *Dyes Pigm.*, 2018, **149**, 921–926.
- 40 E. Feng, C. Fan and N. Wang, Dyes Pigm., 2018, 151, 22–27.
- 41 E. Feng, R. Lu and C. Fan, *Tetrahedron Lett.*, 2017, **58**, 1390–1394.
- 42 Q. Wang, X. Wen and Z. Fan, J. Photochem. Photobiol., A, 2018, 358, 92–99.
- 43 D. Wang, X. Fan and S. Sun, *Sens. Actuators*, *B*, 2018, **264**, 304–311.
- 44 X. Gan, W. Li and C. Li, Sens. Actuators, B, 2017, 239, 642-651.
- 45 Y. Wang, Z. Y. Ma and D. L. Zhang, *Spectrochim. Acta, Part A*, 2018, **195**, 157–164.
- 46 L. Li, H. Li and G. Liu, *J. Photochem. Photobiol.*, *A*, 2017, **338**, 192–200.