



Cite this: *Phys. Chem. Chem. Phys.*,
2024, 26, 23561

Received 12th June 2024,
Accepted 27th August 2024

DOI: 10.1039/d4cp02378a

rsc.li/pccp

Acceleration of xenon kick-out exchange in a cryptophane host explained by *ab initio* simulation†

Rodolphe Pollet, * Jean-Pierre Dognon  and Patrick Berthault 

The degenerate exchange involving xenon and *anti*-cryptophane-222-(OCH₂COOH)₆ in basic water is studied. The reaction consists of the intrusion of a xenon atom into a cavity hosting another xenon and the escape of the latter from the cage to reach the aqueous solution. A series of constrained *ab initio* molecular dynamics simulations were performed according to the Blue Moon ensemble method to reconstruct the free-energy profile for the degenerate exchange reaction at ambient temperature. In addition to the estimation of the free-energy barrier height of the reaction, analysis of the trajectories provides details on its mechanism as well as on the nature of its transition state, where both Xe atoms interact through strong dispersion forces owing to moderate confinement effects.

Introduction

A highly promising molecular imaging approach uses ¹²⁹Xe NMR-based biosensors. Pioneered by the group of Pines,¹ it counteracts the low sensitivity of NMR by using hyperpolarized xenon hosted in dedicated molecular systems that give it a specific spectral signature.² These host systems decorated with ligands are designed to reach the biological receptors of interest and specifically reveal them.

The most powerful method for detecting small quantities of bioprobes *via* ¹²⁹Xe NMR is based on an indirect CEST (chemical exchange saturation transfer) approach³ denoted HYPERCEST,⁴ where the main (free) xenon signal is detected after saturation of the bound xenon frequency. The xenon exchange acts as a pipeline to translate the information on the presence of the bound xenon as a depletion of the main

xenon signal. Spatio-temporal encoding methods further increase the sensitivity of the approach.⁵

Among the most effective xenon host molecules (offering high selectivity for the noble gas, high-frequency separation between the free and bound states, and adequate in-out xenon exchange), cryptophanes have a place of choice. But whatever the host system used, it has been shown that the effectiveness of the approach strongly depends on the xenon in-out exchange rate.^{6,7}

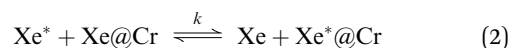
We recently compared the interaction of xenon with two cryptophane congeners, the *syn*- and *anti*-cryptophane-222-(OCH₂COOH)₆. While hyperpolarized ¹²⁹Xe NMR studies found that the former exhibits a faster xenon in-out exchange,⁸ *ab initio* molecular dynamics explained these differences.⁹ Among other structural characteristics, a water molecule blocking the portal of *anti*-cryptophane was shown to slow down the xenon exchange.

In recent years, however, it has been shown that the simple (dissociative) exchange characterized by equilibrium



which had first been assumed is quickly supplanted by a degenerate exchange where the encapsulated xenon is replaced by another xenon atom, as soon as the concentration of dissolved xenon increases. In fact, such a ‘kick-out’ exchange, characterized by a xenon in-out rate depending on the guest/host concentration ratio, was first conjectured¹⁰ and then demonstrated for this kind of molecular system.^{11–13}

The dynamics of this host–guest system is governed by dissociative and degenerate exchange reactions, whose specific mechanisms and rates directly affect the sensitivity of the HYPERCEST method. Specifically, the latter reaction depends on the free xenon concentration because



In this work, the expected faster degenerate exchange involving xenon and *anti*-cryptophane-222-(OCH₂COOH)₆ is

Université Paris-Saclay, CEA, CNRS, NIMBE, 91191, Gif-sur-Yvette, France.

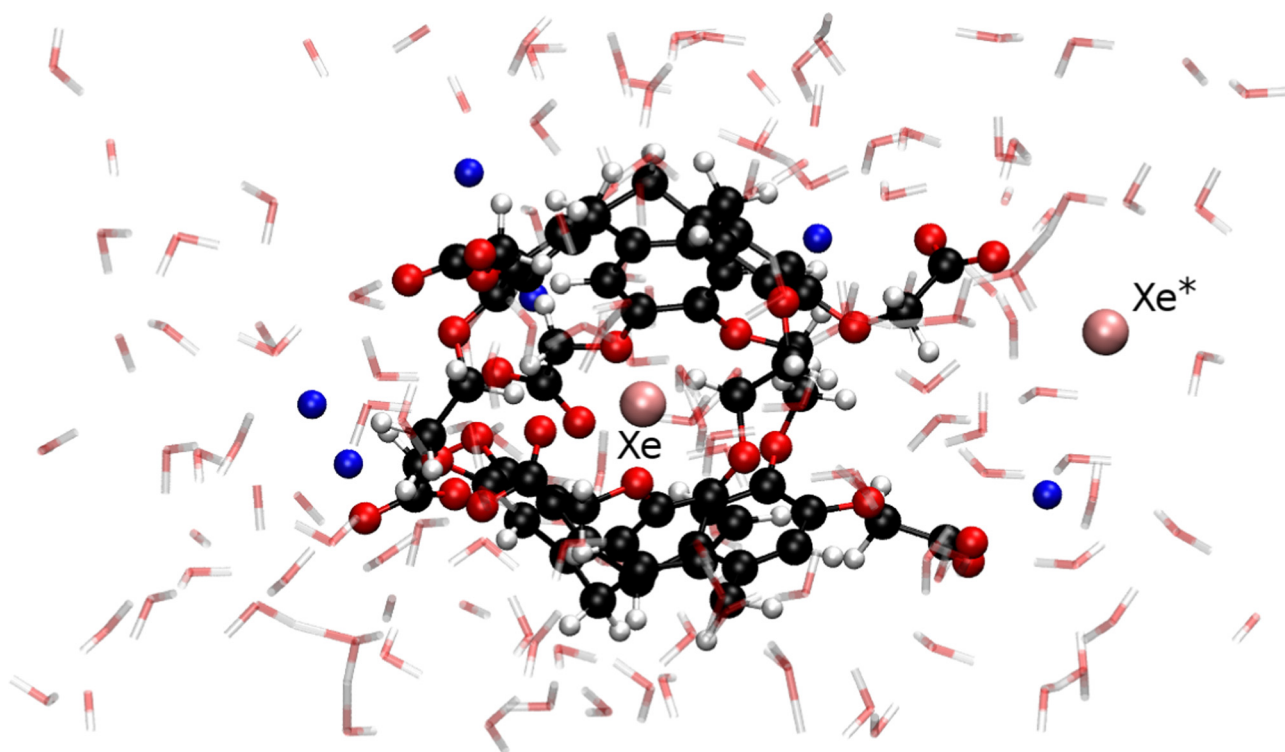
E-mail: rodolphe.pollet@cea.fr

† Electronic supplementary information (ESI) available: Tests of the new Xe pseudopotential, radial distribution function of water molecules, configurations along the degenerate exchange reaction. See DOI: <https://doi.org/10.1039/d4cp02378a>



The Car-Parrinello (CP)¹⁸ simulations were performed at 300 K with an enhanced version of the CPMD code.¹⁹ The calculation relied on the PBE exchange-correlation approximation²⁰ corrected for dispersion effects according to the D3 with zero-damping method.²¹ The fictitious electron mass was set to 700 a.u. and the time step to 6 a.u. The use of deuterium atoms and massive Nosé-Hoover chains (*i.e.*, one for each degree of freedom)²² for the nuclei and electrons helped to maintain the adiabaticity during the CP simulation. The atomic mass of Xe was set at its standard atomic weight of 131.293 a.m.u. The time length of the unconstrained CP trajectory was 9.4 ps.

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A series of 22 constrained molecular dynamics simulations were performed according to the Blue Moon prescription¹⁴ to calculate the free-energy profile for the degenerate exchange reaction (2). The reaction coordinate ξ was the difference between the distances from each of the two xenon atoms and the center of mass (CM) of the cavity, namely $r_{\text{Xe}^*-\text{CM}} - r_{\text{Xe}-\text{CM}}$. The length of each constrained trajectory was 9.4 ps, which included a 3 ps period of relaxation, in order to obtain well-converged average Lagrange multipliers. The change in free energy along ξ was eventually obtained by numerically integrating the time-averaged force of constraint along the reaction path according to

$$\Delta F = - \int_{\xi_0}^{\xi_1} f_{\xi'} d\xi' \quad (3)$$

with

$$f_{\xi'} = \frac{\langle Z^{-1/2} [\lambda - k_B T G] \rangle_{\xi'}}{\langle Z^{-1/2} \rangle_{\xi'}} \quad (4)$$

where λ is the Lagrange multiplier, T the temperature, k_B the Boltzmann constant, and

$$Z = \sum_{i=1}^N \frac{1}{m_i} \left(\frac{\partial \xi}{\partial \mathbf{r}_i} \right)^2 \quad (5)$$

$$G = \frac{1}{Z^2} \sum_{i=1}^N \sum_{j=1}^N \frac{1}{m_i} \frac{1}{m_j} \frac{\partial \xi}{\partial \mathbf{r}_i} \cdot \frac{\partial^2 \xi}{\partial \mathbf{r}_i \partial \mathbf{r}_j} \cdot \frac{\partial \xi}{\partial \mathbf{r}_j} \quad (6)$$

which vanishes in the case of the difference of distances $r_{\text{Xe}^*-\text{CM}} - r_{\text{Xe}-\text{CM}}$, whereas the weight factor is given by

$$Z = \frac{2}{m_{\text{Xe}}} + \frac{2}{m_{\text{CM}}} \left(1 - \frac{\mathbf{r}_{\text{Xe}-\text{CM}}}{r_{\text{Xe}-\text{CM}}} \cdot \frac{\mathbf{r}_{\text{Xe}^*-\text{CM}}}{r_{\text{Xe}^*-\text{CM}}} \right) \quad (7)$$

Results and discussion

First the structure of the $\text{Xe}^* + \text{Xe}@\text{Cr}$ system is briefly described, then the degenerate exchange reaction is analyzed. In the initial CP trajectory of 9.4 ps, the free Xe^* atom was surrounded by approximately 20 water molecules (between 2.85 and 5.45 Å with a maximum at 4.05 Å, see Fig. S1, ESI†), at 3.6 Å from a carboxylate group and 10.3 Å from the encapsulated Xe. The geometry of the filled cryptophane was barely affected by the presence of this external xenon atom, with OCCO dihedral angles (179, 72 and 62°) and diameters (8.7, 9.3 and 9.6 Å), defined as distances between corresponding CH_2 carbon atoms of the cyclotrimeratylene bowls, close to the values reported in our study of the simple – dissociative – $\text{Xe}@\text{Cr}$ exchange (1).⁹ Prior to the exchange, the three ethylenedioxy bridges therefore exhibited a *tgg* conformation.

Free-energy profile for the xenon degenerate exchange as obtained from eqn 3 is shown in Fig. 2, and free-energy differences between A–E configurations (see Fig. S2 (ESI†) for representative pictures) are reported in Table 1 (A is the initial and E the final states). The local minimum B where Xe^* is near

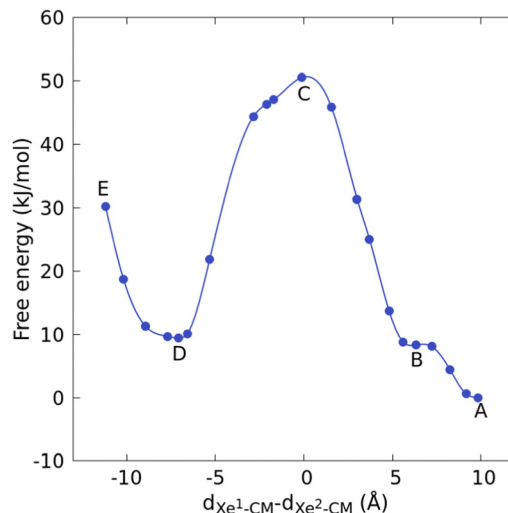


Fig. 2 Free-energy profile for the xenon degenerate exchange as obtained by the Blue Moon method. Natural cubic splines was used for interpolation.

Table 1 Absolute free-energy differences (in kJ mol^{-1}) between configurations A to E of Fig. 2

	A–B	B–C	C–D	D–E
ΔF	08.31	42.21	41.12	20.84

the portal is approximately 8 kJ mol^{-1} higher than when the atom is farther in bulk water (A). The free-energy barrier, between A and C, is estimated to approximately 51 kJ mol^{-1} . For comparison, the free energy difference corresponding to the dissociative escape of a xenon atom from the center of this cage was estimated to 68 kJ mol^{-1} ,⁹ namely 26 kJ mol^{-1} greater than $\Delta F_{\text{B–C}}$. Configurations D and B, where one xenon atom is located outside a portal and the other is inside the cage are close in free energy, with a difference of only 1 kJ mol^{-1} . The final state (that is, bound Xe^* and free Xe) is, however, approximately 30 kJ mol^{-1} higher in free-energy than the initial state (that is, bound Xe and free Xe^*). Although the statistical errors are larger for this final state owing to the Blue Moon protocol, this difference can be caused by a conformational change in the cryptophane, as explained in the detailed analysis that follows.

Degenerate exchange begins with the intrusion of Xe^* into the cage. The atom crossed an available portal between a *gauche* and a *trans* conformations of ethylenedioxy bridges despite the presence of a blocking water molecule that formed hydrogen bonds between its hydrogen atoms and oxygen atoms of two OCH_2COO^- arms. These features were also observed during the escape of xenon in the case of the dissociative exchange (1) and are responsible for the higher free-energy barrier of the present *anti* isomer compared to the *syn* isomer, which experimentally exhibits a faster xenon in-out rate.⁹ The second event of the degenerate exchange was the escape of Xe, which was kicked out by the intruder. The available reaction path crossed a portal



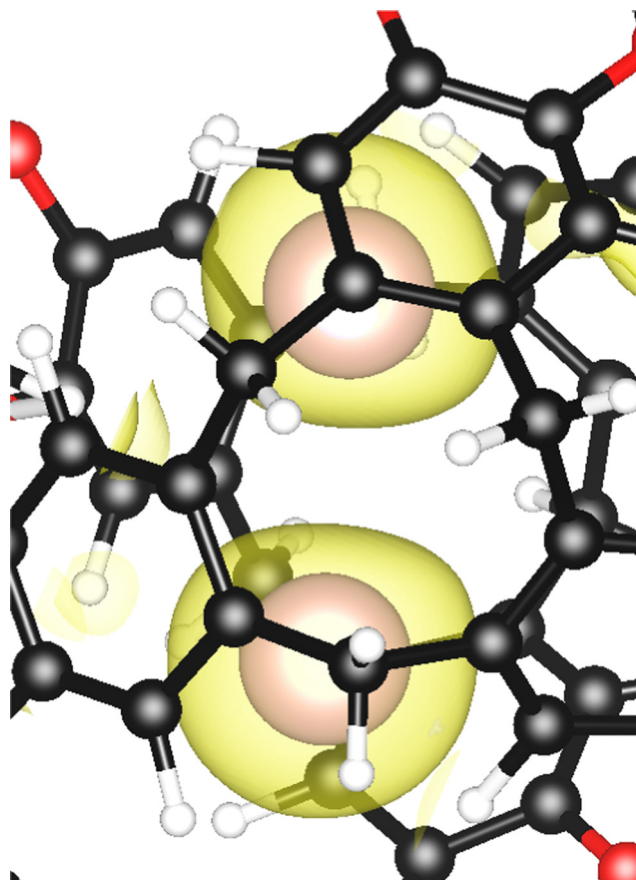


Fig. 3 Zoomed-in view of the electron localization function (ELF), plotted as a yellow transparent isosurface at 0.7, calculated at the transition state geometry with a Goedecker pseudopotential for Xe and a plane-wave cutoff energy of 120 Ry.

between two *gauche* conformations, which was not favorable, so that the approach of Xe caused one conformational change from *g* to *t*. This *gt* configuration seems therefore required for any xenon atom to enter or leave the cage although the *gauche* ethylenedioxy bridges provide a more stable complex. In other words, “breathing” of the cage is required for in-out motions of xenon. Interestingly, such self-organization in response to a stimulus (*i.e.*, the degenerate exchange) is an example of adaptive chemistry.²³ During the escape of Xe, the portal also opened, as shown by the 2 Å increase of the separation between two carboxylate arms. The resulting “gating” phenomenon is therefore a combination of the sliding and French door mechanisms.²⁴ To get an indication of how long this new conformation can last with Xe outside of the cage, one of the Blue moon trajectory (approximately 12 kJ mol^{−1} below E and with a xenon–xenon separation of 10.6 Å) was extended up to 19 ps and the cryptophane remained in this *gtt* conformation.

Between these two events, namely intrusion and escape, in the transition state (C), the two xenon atoms formed a dimer located in the center of the cage, which preserved its conformation *ggt* and whose dimensions only increased weakly (by a few percent). Both atoms were equidistant from the center of mass of the cryptophane, facing two portals, so the average angle of



Fig. 4 Electron localization function (ELF) profile along the Xe–Xe* axis calculated at the transition state geometry with a Goedecker pseudopotential for Xe and a plane-wave cutoff energy of 120 Ry.

Xe–CM–Xe* was 142°. The average Xe–Xe* distance was 3.87 Å, which is 0.61 Å shorter than our estimate *in vacuo* (see ESI†). Such confinement effects have been observed for noble gases in fullerenes.^{25–27} Especially for Xe₂@C₆₀, a genuine chemical bond has been found with a length of 2.49 Å.²⁸ In the Cr222(OCH₂COO[−])₆ cryptophane, the electron localization function (ELF)^{29–31} did not show any sign of such a shared electron bond (see Fig. 3). Recently, ELF analysis has been conducted not only for chemical bonding but also for physical binding, such as ionic binding, hydrogen binding, or van der Waals binding.³² Here, the apparent contraction of the ELF basins, the nonzero ELF value (0.064) between the two nuclei (see Fig. 4), and the distance which is shorter than the sum of van der Waals radii denoted a strong dispersion interaction. We infer that the confinement effects are therefore moderate for Xe₂@Cr.

Conclusions

The degenerate exchange exhibits a much shorter reaction time (approximately a tenth of milliseconds at the studied concentration in free xenon according to Eyring equation) than the dissociative mechanism. The reaction crosses a transition state where both xenon atoms interact through strong dispersion forces because of moderate confinement effects. When the first xenon atom is kicked out by the second one, the gate of the cryptophane cage simultaneously undergoes an opening and a deformation, namely a conformational change of an ethylenedioxy linker from *gauche* to *trans* with a higher free energy.

According to this simulation, the xenon in-out exchange rate could depend on the relative proportion of *gtt* and *ggt* conformations for the ethylenedioxy linkers of the cryptophanes. The limiting step being the passage from one conformation to the other one, with increasing xenon density, a more important



number of cryptophanes should remain in *gtt* (or *ttt*) conformation, facilitating the xenon exchange. This spectacular case of molecular adaptation has important consequences on the efficacy of the ^{129}Xe NMR bioprobes based on *anti*-cryptophanes.

Data availability

The data that support the findings of this study are available in the ESI,[†] of this article.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was granted access to the HPC resources of TGCC under the AD010811468R1 and AD010811468R2 allocations made by GENCI. The authors warmly thank Thierry Brodin and his team at ENS Lyon for giving them access to these water-soluble cryptophanes.

References

- M. M. Spence, S. M. Rubin, I. E. Dimitrov, E. J. Ruiz, D. E. Wemmer, A. Pines, S. Q. Yao, F. Tian and P. G. Schultz, *Proc. Natl. Acad. Sci. U. S. A.*, 2001, **98**, 10654–10657.
- P. Berthault, G. Huber and H. Desvaux, *Prog. Nucl. Magn. Reson. Spectrosc.*, 2009, **55**, 35–60.
- S. Forsén and R. A. Hoffman, *J. Chem. Phys.*, 1963, **39**, 2892–2901.
- L. Schröder, T. J. Lowery, C. Hilty, D. E. Wemmer and A. Pines, *Science*, 2006, **314**, 446–449.
- C. Boutin, E. Léonce, T. Brodin, A. Jerschow and P. Berthault, *J. Phys. Chem. Lett.*, 2013, **4**, 4172–4176.
- M. Zaiss, M. Schnurr and P. Bachert, *J. Chem. Phys.*, 2012, **136**, 144106.
- M. Kunth, C. Witte, A. Hennig and L. Schröder, *Chem. Sci.*, 2015, **6**, 6069–6075.
- E. Léonce, T. Brodin and P. Berthault, *Phys. Chem. Chem. Phys.*, 2022, **24**, 24793–24799.
- R. Pollet, J.-P. Dognon and P. Berthault, *Chem. Phys. Chem.*, 2023, **25**, e202300509.
- K. Bartik, M. Luhmer, J.-P. Dutasta, A. Collet and J. Reisse, *J. Am. Chem. Soc.*, 1998, **120**, 784–791.
- G. Huber, L. Beguin, H. Desvaux, T. Brodin, H. A. Fogarty, J.-P. Dutasta and P. Berthault, *J. Phys. Chem. A*, 2008, **112**, 11363–11372.
- S. Korchak, W. Kilian and L. Mitschang, *Chem. Commun.*, 2015, **51**, 1721–1724.
- M. Kunth, C. Witte and L. Schröder, *J. Chem. Phys.*, 2014, **141**, 194202.
- M. Sprik and G. Ciccotti, *J. Chem. Phys.*, 1998, **109**, 7737–7744.
- D. Marx and J. Hutter, *Ab initio Molecular Dynamics: Basic Theory and Advanced Methods*, Cambridge University Press, Cambridge, 2009.
- G. Huber, T. Brodin, L. Dubois, H. Desvaux, J.-P. Dutasta and P. Berthault, *J. Am. Chem. Soc.*, 2006, **128**, 6239–6246.
- E. Léonce, J.-P. Dognon, D. Pitrat, J.-C. Mulatier, T. Brodin and P. Berthault, *Chem. – Eur. J.*, 2018, **24**, 6534–6537.
- R. Car and M. Parrinello, *Phys. Rev. Lett.*, 1985, **55**, 2471–2474.
- T. Klöffel, G. Mathias and B. Meyer, *Comput. Phys. Commun.*, 2021, **260**, 107745.
- J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- G. J. Martyna, M. L. Klein and M. Tuckerman, *J. Chem. Phys.*, 1992, **97**, 2635–2643.
- J.-M. Lehn, *Angew. Chem., Int. Ed.*, 2013, **52**, 2836–2850.
- K. Houk, K. Nakamura, C. Sheu and A. E. Keating, *Science*, 1996, **273**, 627–629.
- D. Paul, H. Dua and U. Sarkar, *Front. Chem.*, 2020, **8**, 621.
- I. Fernández, M. Solà and F. M. Bickelhaupt, *J. Chem. Theory Comput.*, 2014, **10**, 3863–3870.
- M. Khatua, S. Pan and P. K. Chattaraj, *J. Chem. Phys.*, 2014, **140**, 164306.
- A. Krapp and G. Frenking, *Chem. – Eur. J.*, 2007, **13**, 8256–8270.
- A. D. Becke and K. E. Edgecombe, *J. Chem. Phys.*, 1990, **92**, 5397–5403.
- B. Silvi and A. Savin, *Nature*, 1994, **371**, 683–686.
- D. Marx and A. Savin, *Angew. Chem., Int. Ed. Engl.*, 1997, **36**, 2077–2080.
- K. Koumpouras and J. A. Larsson, *J. Phys.: Condens. Matter*, 2020, **32**, 315502.

