

REVIEW

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Advanced XPS characterization: XPS-based multi-technique analyses for comprehensive understanding of functional materials

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X-ray photoelectron spectroscopy (XPS) has achieved maturity as an analytical technique in that it is a ubiquitous tool in the materials community, however as made apparent by recent reviews highlighting its misuse as a means of chemical deduction, it is a practice which is greatly misunderstood even in its simplest form. Advanced XPS techniques, or a combination of XPS and a complementary surficial probe may elicit auxiliary information outside of the scope of the standard sphere of appreciation. This review aims to bring to the attention of the general materials audience a landscape of some atypical applications of lab-based XPS and combinatorial approaches of related surface analysis, such as ion scattering, ultraviolet photoelectron, electron energy loss and auger emission spectroscopies found on many lab-based instrument set-ups.

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Introduction

Functional materials on the nano- to macro-scale provide solutions for some of the most critical scientific problems of

the current age, from the use of heterogeneous catalysts to improve the environmental footprint and commercial viability of fine chemical synthesis (e.g. pharmaceuticals, food additives, cosmetics)¹ to solid state materials as energy storage materials,² fuel cells³ or photovoltaics⁴ to tackle humanity's dependence on dwindling resources such as fossil fuels. Such is the scope in the application of functional materials that examples may be found across the spectrum of the physical sciences; including drug-delivery systems,⁵ molecular sensors,⁶ tackling drug resistant pathogens,⁷ aerospace components,⁸ electronics⁹

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and many more. In order to design and develop the most effective materials, a thorough understanding of the material properties is often paramount – particularly at the surface of a material which forms the interactive boundary between the functional material and the materials or substances on which they will act.¹⁰

X-ray photoelectron spectroscopy (XPS) has become a well-known technique in the field of materials science, one which

may be used to probe the elemental, electronic and chemical characteristics of a material surface.^{11–14} While the interpretation of XPS data *via* peak deconstruction is commonplace, fundamental understanding of the technique and an appreciation for correct data handling are often missing.¹⁵ Recently, a series of invaluable guides have been prepared from a community effort among the leading practitioners in the field of XPS, with the aim of enabling researchers new to XPS to plan experiments and understand their data to a high level. This series is published under the title 'Practical Guides for X-Ray Photoelectron Spectroscopy', for example 'Practical Guides for X-ray Photoelectron Spectroscopy: First Steps in Planning, Conducting, and Reporting XPS Measurements'¹⁶ and 'Practical guide for curve fitting in X-ray photoelectron spectroscopy'.¹⁷ Furthermore, there exist a number of advanced techniques with which many materials scientists will be unfamiliar. Furthermore, a current focus of several of the XPS manufacturers is high throughput examination of a single analysis spot using an assembly of surface probes and even non-surface specific techniques such as Raman spectroscopy. As the availability of such systems becomes more widespread, the a wider knowledge of the capabilities, strengths and weaknesses of multi-technique surface analysis is required.

This review aims to highlight the benefits of such a combinatorial approach to the analysis of materials utilising lab-based XPS and related surface techniques. Applications of those experimental probes which may be found most regularly on lab-based systems (ion scattering, ultraviolet photoelectron, auger emission and electron energy loss spectroscopies) are covered in detail here, though many other complementary



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2015. His research interests include investigating selectively catalytic hydrogenation using electrochemical methods and CO oxy-chlorination reactions with various surface characterization techniques such as X-ray photoelectron spectroscopy and Surface enhanced Raman spectroscopy.

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mission spectra, and uses a joint experimental and theoretical approach in her research.

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approaches may provide pathways through which to tackle research problems in materials analysis. As mentioned above, Raman spectroscopy has found much interest recently in unison with XPS, since Raman may afford a probe of structural properties including analysis of defect sites, often of paramount importance when dealing with surfaces and interfaces (the study of which XPS excels at).^{18–22} Such is the potential of this approach, selected commercial instruments now afford the option of an ‘off-the-shelf’ Raman module upgrade – permitting users to perform tandem experiments with little to no experience of the modification of UHV chambers.

The sheer scope of combinatorial analysis renders a full discussion of all methods impractical, and hence only a selection (of those deemed most readily available) will be presented here, however for a look into other powerful conjunctive techniques the reader may be directed towards some reviews highlighting the use of XPS/Scanning Tunnelling Microscopy (STM),^{23–25} XPS/Scanning Probe Microscopy (SPM),^{25,26} XPS/Infrared Reflectance Absorption Spectroscopy (IRRAS),^{27–29} XPS/Time-of-Flight Secondary Ion Mass Spectroscopy (TOF-SIMS)^{25,26} and XPS/Atomic Force Microscopy (AFM)^{25,27} though even this is itself by no means an exhaustive list.

Advanced XPS techniques

As well as providing a method by which to qualitatively and quantitatively describe the chemical nature of the surfaces of materials, XPS can also provide insights into the topology, morphology and homogeneity of a sample as a function of depth. Such techniques may facilitate the understanding of new interfaces (*e.g.* in semiconductors, electrodes),^{30,31} solid synthesis pathways,³² catalytic and surface processes^{33–35} and electronically active materials,^{36,37} among others.

Tougaard background analysis

Each individual photoemission peak in an XP spectrum is accompanied by a characteristic stepped background (Fig. 1) which is the result of photoelectrons which have undergone an inelastic scattering process.³⁸ In order to correctly analyse XP data, this inelastic scattering process must be properly accounted for in the form of fitted model background.³⁹

The most widely used model, known as the Shirley background after its creator D. A. Shirley,⁴⁰ uses an iterative process to calculate the number of inelastically scattered electrons based upon the integrated peak area either side of the energy E . It has proven to be a reliable and robust method for accounting for inelastic scatter. However, the physical origin of the Shirley background is rather poorly defined.⁴¹ Detailed theoretical analysis with comparison to experimental data has led to much improved models of electron inelastic scattering⁴² and a new method for quantitatively describing the XPS background named after its creator Sven Tougaard has been developed.⁴³ A discussion of the merits of the various background types for standard XPS analysis is beyond the scope of this review but we will highlight how Tougaard’s approach

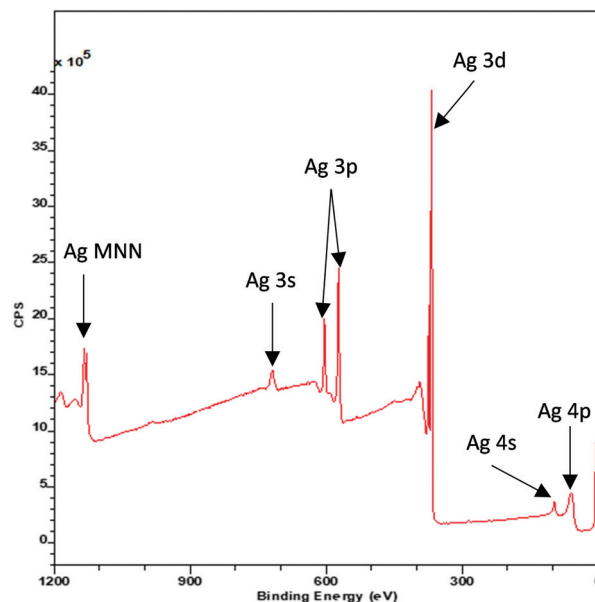


Fig. 1 X-ray photoelectron survey spectrum of a clean silver surface.

provides a non-destructive technique by which depth profile and structural information, including the spatial arrangement of overlayers, islands and sandwich structures can be obtained from the background profiles to complex composite materials.⁴⁴

Nucleation and growth modes in systematic samples can be obtained with an even greater degree of accuracy and consistency than that provided by ellipsometry or Rutherford backscattering spectroscopy.^{45,46} Tougaard background analysis has yielded excellent agreement with atomic force microscopy (AFM) measurements in determining Stransky–Krastanov growth of NiO deposition onto highly ordered pyrolytic graphite (HOPG).⁴⁷ The technique has advanced to the point where analysis may be performed quickly and easily using specialist software (QUASES).^{48,49}

Overlayer thickness analysis

Nanocomposite materials in the form of thin films or core-shell particles often offer unique solutions to critical research problems in areas such as electronics,⁵⁰ medical devices³² and catalysis.⁵¹ Control over the film or shell thickness can be vital to obtain the precise physical and electronic properties required of such composite materials and therefore a robust method is needed to determine said thickness. X-ray reflectivity (XRR)⁵² and electron microscopy⁵³ are often used for this purpose, however they are not without drawbacks; XRR requires highly ordered samples with almost no surface roughness, whilst electron microscopy requires long analysis times to obtain a statistically significant dataset. XPS can provide an alternative approach to film thickness determination involving a large area analysis with few limitations regarding sample chemistry.



For planar materials, such as semiconductor thin films deposited onto a substrate, there are a number of approaches to using XPS to determine film thickness. For simple metal oxide on metal systems such as $\text{Al}_2\text{O}_3/\text{Al}$ for example, one may use the Strohmeyer equation (eqn (1)) based on a single high resolution region spectrum.^{54,55}

$$d = \lambda_{\text{ox}} \sin \theta \ln \left[\left(\frac{N_{\text{m}} \lambda_{\text{m}} I_{\text{ox}}}{N_{\text{ox}} \lambda_{\text{ox}} I_{\text{m}}} + 1 \right) \right] \quad (1)$$

Strohmeyer equation; λ = inelastic mean free path, N = volume density (density/mass), I = intensity and θ = photoelectron take-off angle.

This method assumes a uniform overlayer and that the photoelectrons from both metal and oxide are of similar energy (hence is largely restricted to metal oxide layers of the same element as the underlying metal).

Differences in the inelastic mean free path (IMFP) between oxide and metal permit a facile calculation of path length using only intensity ratios between the two species. An important note for users relatively unfamiliar with XPS, is that between IMFP and the information depth of the sample. IMFP (λ) is defined as the distance between successive inelastic collisions of photoelectrons and the surrounding medium (*i.e.* atoms in the sample as the photoelectron travels to the surface). The information depth (ID), refers to the depth of analysis from which the majority (> 95%) of the signal originates.⁵⁷ This is dependent on both the material IMFP and the energy of the incoming X-rays and is typically found to be approximate to 3λ .

Since the electron escape path length scales with the angle of emission (while X-ray penetration depth remains constant, Fig. 2), angle resolved XPS (ARXPS) provides a simple, effective and non-destructive system by which to measure film thickness and/or produce relative depth profiles for mapping layer structures.^{58–61} A major advantage of ARXPS over XRR and electron microscopy is the ability to not only probe film thicknesses but also reveal elemental and chemical environments at differing depths.^{62–64} This makes ARXPS a particularly useful tool for studying solid interfaces, which play a crucial role in many electronic systems.^{65–68} Application of ARXPS to the valence orbitals is also extensively used to map electronic band structures.⁶⁹

One method with which ultrathin-film thickness may be determined is by monitoring the intensity of the substrate before and after deposition of an overlayer and using known inelastic mean free path (IMFP) lengths to calculate the path length of the attenuating material (overlayer thickness).⁷⁰ This method is not without its restrictions, since it may not account for differences in surface contamination, and surface topography and it also assumes identical experimental conditions and a pure substrate sample – which may not always be possible. It is, however, an excellent probe of uniform flat surfaces such as those prepared *via* deposition.⁷¹ A technique developed by Cumpson known as the “thickogram” facilitates overlayer analysis from a single XPS spectrum *via* graphical means.⁷² The method uses minimal calculation, and has the advantages of being independent of contamination layer and instrument

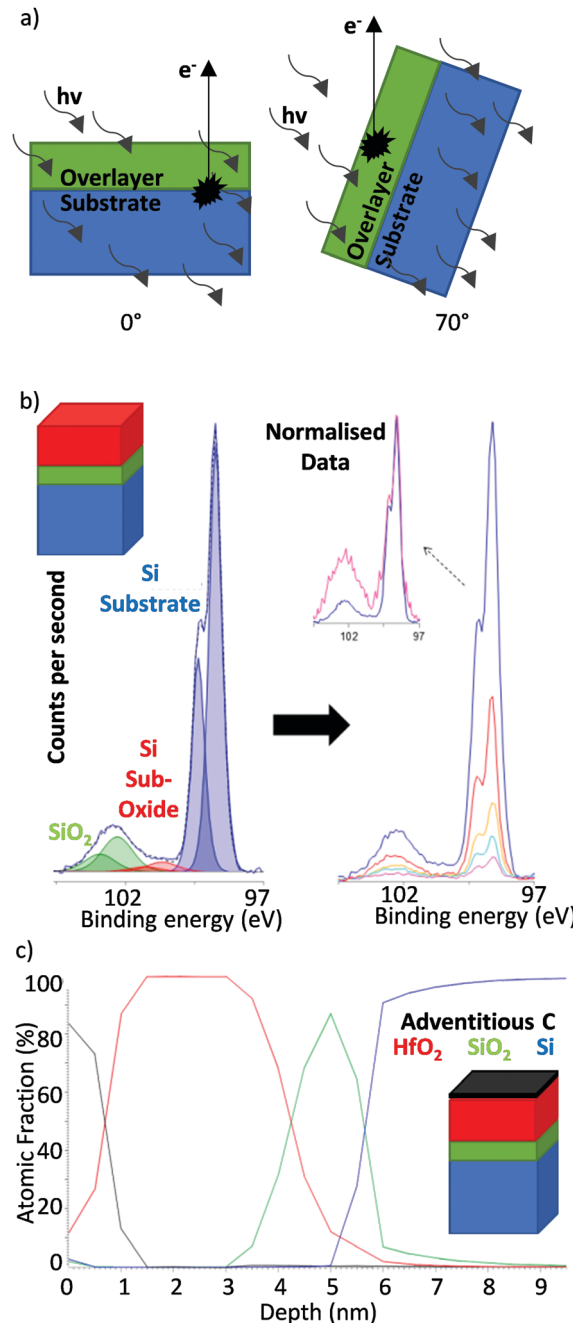


Fig. 2 (a) Angular dependence of the electron path length in ARXPS (vertical black arrows represent photoelectron pathway to the sample surface the escape depth of which, perpendicular to the sample surface, is dependent on IMFP and sample angle), (b) example of ARXPS applied to Hf oxide films on a silicon substrate in which dependence of silicon signal on angle of emission may be clearly observed and (c) relative depth profile calculated from the quantitatively processed spectra in (b). Reprinted (adapted) with permission from ref. 56. Copyright {2019} Kratos Analytical.

factors, and works well over a large range of thicknesses and energies. The ratio of kinetic energies and the ratio of integrated peak areas for a substrate and overlayer peak, are plotted onto a thickogram template. The intersection of a line joining these two points with the third axis on the thickogram provides a value for the coating thickness. Since sample morphology may still



influence the analysis when using this method, a series of correction factors (known as topofactors) have been derived by Shard which may permit the determination of non-uniform sample surfaces and even complex structures such as core shell nanoarchitectures.⁷³

The analysis of smaller particles (< 10 nm) requires particular consideration since the substrate on which the particles are mounted may also contribute to the XP spectrum and in these cases considerable care must be taken with methods such as Tougaard's background analysis. Angle-resolved measurements are typically unsuitable for such systems, since at such small size ranges we have topological considerations rendering the technique inappropriate. Models have been developed which permit calculations to account for the morphological differences of non-planar samples at a wide range of sample dimensions.⁷⁴ Employing a twin energy X-ray source (for example Mg and Al) is one method by which to overcome some limitations regarding experimental conditions, since different energy X-rays may analyse the same region of the same sample but result in photoelectrons with different mean free paths and thus provide more information on the overlayer structures.³² This method has been shown to be in good agreement with electron microscopy for Ag@SiO₂ core shells over a small size range.³²

The overlayer analysis methods described so far have the major advantage of being non-destructive and the superb depth resolution (in the order of Angstroms) of such non-destructive methods makes them ideal for analysing small particles (< 20 nm for energy sources below 5 keV) thin films, overlayers or inhomogeneous samples.⁷⁵ However, all these techniques are limited by the maximum kinetic energy of the chosen X-ray source (typically around the order of 10 nm for a 1486.69 eV Al $K\alpha$ X-ray source). To address this limitation, a range of high energy laboratory X-ray sources such as the liquid gallium jet (9.25 keV),⁷⁶ are being developed which complement the variable energy photons offered by synchrotrons (up to 18 keV)^{75,77,78} and facilitate examination of more deeply lying interfaces. An additional limitation when dealing with HAXPES is the decreased sensitivity due to diminishing photoionisation cross-sections at higher excitation energies^{79–81} (thus also requiring higher powered sources which has the potential to lead to greater X-ray induced sample damage). One result of this is the need for high energy analysers with a large angle of acceptance, creating a further barrier in the development of HAXPES systems.^{82,83} Nevertheless there will always be an upper limit to the depth that non-destructive XPS methods can probe.

Depth profiling

An alternative method for exploring composition gradients in the Z-dimension by XPS is to use a destructive bombardment technique. This typically involves an ion beam (most commonly Ar⁺) in which accelerating voltage, current and raster size may be tuned in order to control the removal of surface atoms *via* collisions with the ions.^{84,85} Obtaining snapshot spectra at regular intervals during the etching process (at which point the ion beam is turned off) permits the construction of a depth

profile in which elemental or chemical state may be charted as a function of total sputter time (Fig. 3); converting the data to a depth profile requires an estimate of the etching rate for the material under study. The major drawbacks in ion bombardment depth profiling are a reduced depth resolution due to a mixture of non-uniform sputtering (*e.g.* cratering),⁸⁶ preferential sputtering (*e.g.* oxygen in metal oxides),⁸⁷ chemical reduction during sputtering⁸⁸ and the need to calibrate etching time with depth.⁸⁴

Until recently another major issue with ion bombardment sputtering was that it was limited when applied to organic materials, since ion etching of polymers results in significant chemical degradation.⁹⁰ However, the advent of gas cluster ion sources/beams (GCIS/GCIB) in which ionised multi-atomic clusters are used in place of single ions has led to more stable etch craters due to multiple secondary collision events⁹¹ and exhibit minimal damage for organic molecules, enabling z-dimension profiling and 3D imaging for the vast majority of organic materials.⁹² Depth resolutions as low as 7 nm have been recorded and may be readily improved through the

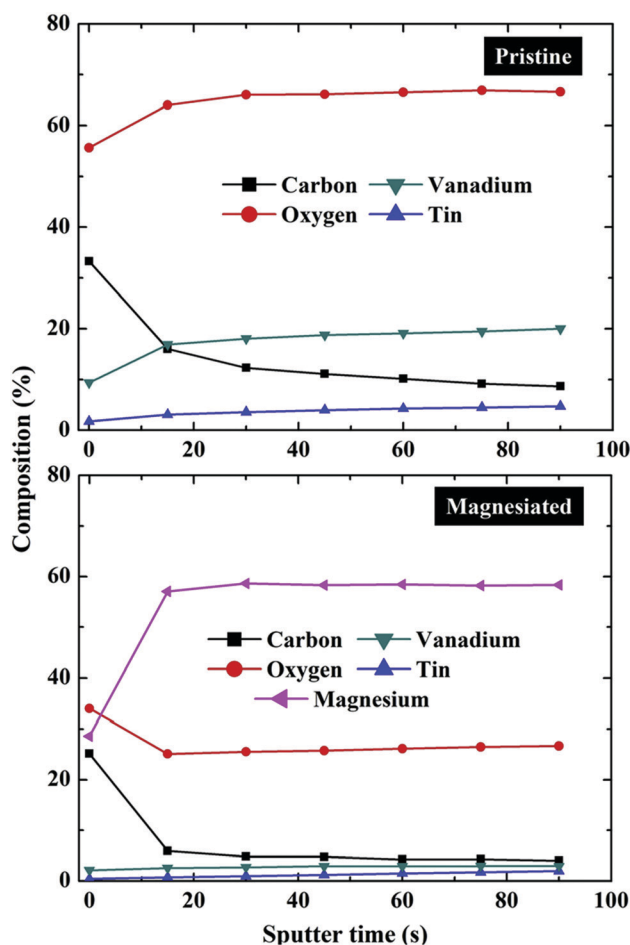


Fig. 3 XPS depth profile composition measurements of pristine and magnesiated V₂O₅ cathodes indicating the relative contents of carbon, oxygen, vanadium, tin and magnesium as a function of overall ion bombardment time (in itself a function of depth from the surface of the sample down into the bulk).⁸⁹



utilisation of a rotating sample stage and a decrease in X-ray flux, since exposure to X-rays can also degrade the steadiness of the etching process.⁹³ Further improvements to depth resolution may come from combining ion bombardment with a technique such as TOF-SIMS.⁹⁴

Advanced sample handling for lab-based XPS systems

While not necessarily an individual technique, the development of technology designed to overcome the imitations of analysing samples within a UHV system has widened the reach of XPS as a tool for probing materials. The advent of near-ambient pressure (NAP)-XPS⁹⁵ is one pertinent example of the science adapting to accommodate samples previously difficult or even impossible to analyse – though this has historically required specialist instrumentation often only found at a synchrotron. Recent years have seen a marked rise in the number of lab-based NAP-XPS systems,⁹⁶ with manufacturers such as Omicron Scienta and SPECS offering bespoke and ‘off-the-shelf’ instruments capable of performing ambient pressure measurements. This review will, however, focus on techniques by which we can use traditional lab-based XPS systems to probe air-sensitive materials or perform *in situ* experiments. For a more detailed discussion regarding the use and application of NAP-XPS to materials analysis, several reviews and articles are available for the reader to explore.^{24,95–100}

Studying electrochemical systems using XPS

One discipline in which such an approach has had great impact is the field of electrochemistry which is currently of great research significance due to its role in the energy sector where electrocatalysis represents a viable solution to hydrogen production¹⁰¹ and electrode materials to that of energy storage.¹⁰² It is generally accepted that an improved understanding of the underlying mechanisms that occur at the interface between the electrolyte and the electrode, and any restructuring of the electrode surface during polarisation, activation and deactivation is extremely important.¹⁰³ *Ex situ* and *in situ* XPS studies of electrode materials,¹⁰⁴ ionic liquids and electrochemical processes including electrochemical double layer (EDL) studies have been employed to provide insights into such mechanisms.

The surface sensitivity of XPS and its UHV requirements make it quite complicated to apply to the electrode/electrolyte interface; liquid electrolytes are obviously not compatible with vacuum conditions, making true *in situ* measurements difficult.

Ex situ electrochemical XPS refers to investigations where the material studied is removed from the original electrochemical cell and transferred to the vacuum chamber, usually this would involve disassembly in a glovebox and transfer using a suitable inert transfer device (Fig. 4). Such transfer modules may involve either the use of a vacuum sealed system (Fig. 4a–c), in which samples may be mounted inside a glovebox and the transfer module evacuated within the antechamber, where a check valve

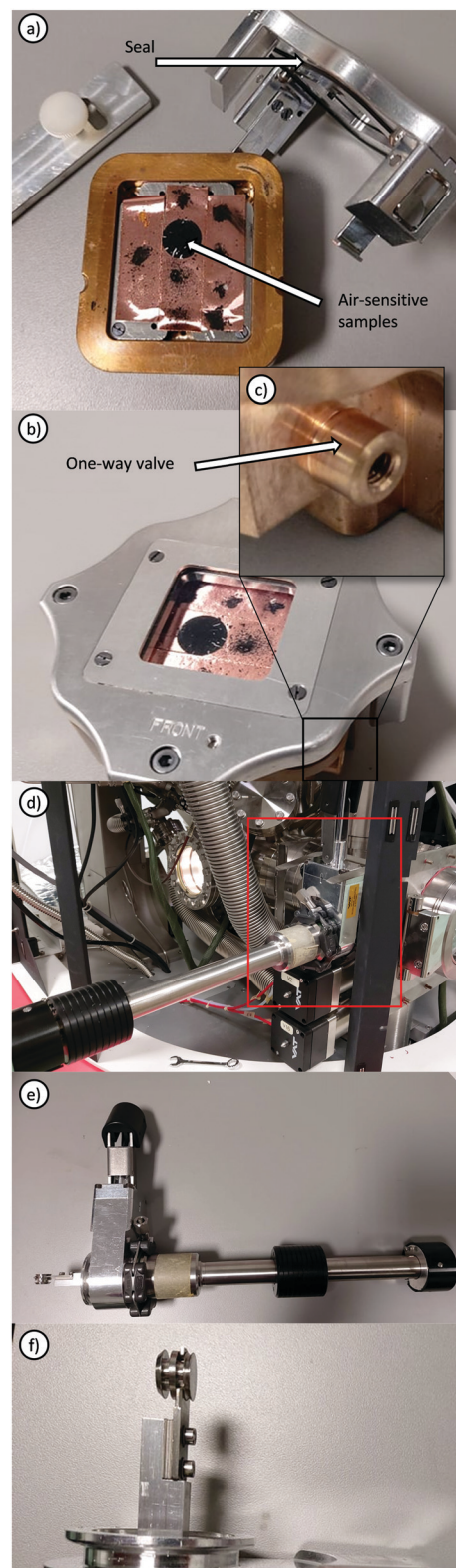


Fig. 4 Typical inert sample handling devices. (a–c) Sample transfer under vacuum from Thermo Scientific™. (a) Disassembled transfer unit, (b) assembled transfer unit and (c) check valve to hold internal vacuum – (d–f) sample transfer under vacuum or inert gas from Kratos Analytical. (d) Transfer arm attached to UHV chamber, (e) transfer arm removed from UHV chamber and (f) sample stub held on manipulator probe.



permits evacuation inside the cell and holds the vacuum until loading inside the instrument. Alternatively, through the use of a sealed manipulator, gate valve and standard KF flange, samples may be held under vacuum or inert gas, transferred from glovebox to instrument and the KF flange directly mounted onto the vacuum chamber, permitting the sample to be moved into the instrument after the gate valve is opened. Such measurements always introduce the possibility that the surfaces may be different from their original states but valuable electrochemical information can still be obtained following such disassembly as demonstrated by XPS studies of the oxygen evolution reaction (OER) on electrodes.¹⁰⁵

A variety of methods have been developed to overcome issues with the *ex situ* approach including elaborate transfer systems to minimise exposure to atmosphere,¹⁰⁶ and the electrode immersion technique.^{107–109} The latter, attempts to retain the original conditions of the samples from the electrochemical cell whilst in the XP spectrometer. Pioneering work by Kolb *et al.*¹¹⁰ showed that it was possible to remove electrodes from the electrolyte during potential control and keep the electrical double layer (EDL) intact. Gold electrodes were immersed from various electrolytes prior to performing XPS analysis, which was used to determine the surface excesses and adsorption isotherms for a number of ions and metal deposits. Later, Hecht *et al.*¹¹¹ examined the EDL on Ag in alkaline NaCl after immersion with various potentials, and studied the concentrations of the cations (Na^+) and anions (Cl^- and OH^-) adsorbed on the silver electrode *via* quantitative XPS analysis. Their analysis revealed that the anions tend to bind Na^+ into the double layer for charge compensation, and that roughly one monolayer of water was found for all potentials studied, in agreement with classical double layer studies.

Ex situ electrochemical measurements

The importance of the surface sensitivity and chemical state information available from XPS is demonstrated well by recent studies carried out by Andreu *et al.*¹⁰⁴ into lithium ion batteries. The research involved modification of the positive electrode material, a mixed transition metal oxide consisting of $\text{LiNi}_x\text{Mn}_x\text{Co}_{1-2x}\text{O}_2$ where ($0.33 \leq x \leq 0.5$)^{112–116} to improve its electrochemical performance. It has been shown that coating the electrode with a metal oxide coating such as Al_2O_3 or ZrO_2 can improve the electrochemical performance at high voltages.^{117–119} Investigations into how the transition metal nature and Al_2O_3 coating on LiCoO_2 and $\text{LiNi}_{1/3}\text{Mn}_{1/3}\text{Co}_{1/3}\text{O}_2$ (NMC) surfaces affect performance were carried out by Andreu *et al.*¹⁰⁴ studying the adsorption of SO_2 on the surface with XPS. The resulting spectra showed only one adsorption mode present for the LiCoO_2 surface, signified by the sulfate species (Fig. 5a and a') whereas the NMC electrode material exhibited both sulfite and sulfate species (Fig. 5b and b'). In the case of the LiCoO_2 electrode, the coating modifies the surface reactivity, indicated by the appearance of both sulfite and sulfate species, and the authors attribute this to Al/Co substitution shifting the adsorption mode from redox to acid-base. In the case of the NMC electrode, a decrease of sulfite species is seen at

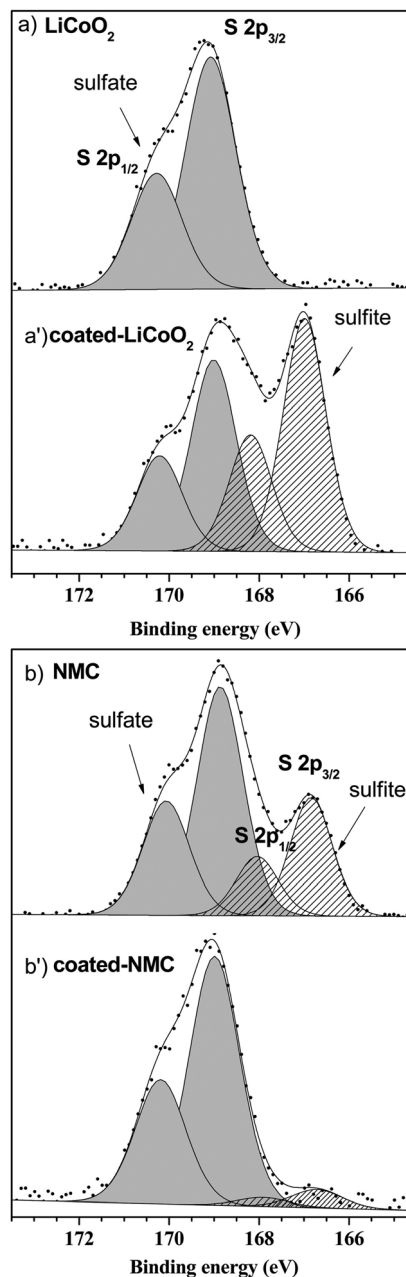


Fig. 5 XPS spectra for the S 2p core peaks after SO_2 adsorption on (a) LiCoO_2 and (a') Al_2O_3 coated LiCoO_2 and XPS spectra for the S 2p core peaks after SO_2 adsorption on (b) NMC and (b') Al_2O_3 coated NMC. Reprinted (adapted) with permission from ref. 104. Copyright {2015} American Chemical Society.

the surface, a phenomenon that can be related to the cation mixing effect in the NMC.¹²⁰

XPS depth profiling has also been important in research into alternative energy storage with pseudocapacitors, which are based on redox reactions of electroactive materials, have a longer life than batteries and a higher energy density than conventional capacitors.¹²¹

A study of the pseudocapacitive performance of an Mn oxide electrode in the ionic liquid (IL) EMI-SCN was carried out by Chang *et al.*¹²² ILs are composed solely of ions and exhibit



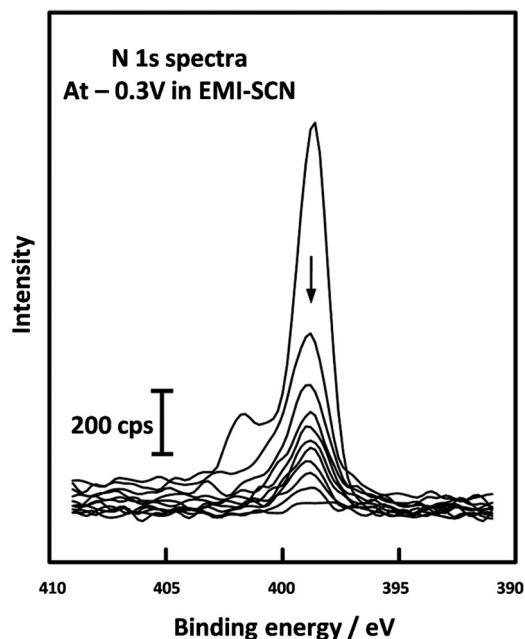


Fig. 6 XPS depth profiling spectra of the N 1s level from Mn oxide electrodes in EMI-SCN, [C₂C₁Im][SCN] polarised at -0.3 V, reported by Chang *et al.* Reprinted (adapted) with permission from ref. 122. Copyright (2009) American Chemical Society.

several desirable properties including high electrical conductivity (up to 100 mS cm^{-1}), a wide electrochemical window (up to 5.8 V), and extended liquid-state temperature range (*ca.* 173 – 523 K).¹²³ Furthermore, their negligible vapour pressure overcomes the issue of incompatible liquid electrolytes in UHV systems. In the case of EMI-SCN, the cation (EMI⁺) and the anion (SCN[−]) both contain nitrogen atoms and so can be conveniently compared using the N 1s region of the XP spectrum, (Fig. 6). Chang *et al.* were able to identify the working species in the pseudocapacitor as the SCN[−] anions by recording sequential N 1s spectra of the Mn oxide electrodes in EMI-SCN during argon ion etching; 10 spectra recorded with the electrode under a potential of -0.3 V are shown superimposed in Fig. 6. The peak at 398.7 eV (due to the EMI⁺ cation) disappears after only one Ar⁺ sputtering cycle, indicating that it is only adsorbed on the surface signal whereas the at 401.6 eV (due to the SCN[−] anion) only gradually decreases with depth, indicating that it has penetrated into the Mn oxide surface.

In situ electrochemical cells

Despite advances brought by *ex situ* characterisation, post-analysis can mean that some metastable species are not analysed since they are only present under reaction conditions,¹²⁴ therefore *in situ*, or *in operando*, analysis is of great value. An *in situ* XPS cell reported by Weingarth *et al.*¹²⁵ demonstrated that XPS analysis can be performed simultaneously whilst carrying out electrochemical measurements, to determine the direct electrochemical shift as a function of applied potential. The system studied was a three-electrode set up involving a counter electrode (CE) and reference electrode (RE) prepared of activated carbon,

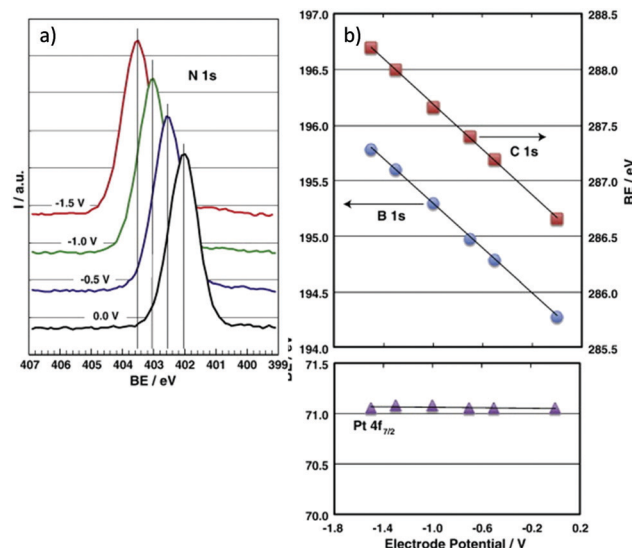


Fig. 7 *In situ* electrochemical XPS measurements as reported by Weingarth *et al.*¹²⁵ for their three electrode *in situ* XPS cell, performed using a VG ESCALAB 220iXL spectrometer equipped with a monochromatic Al K α X-ray source. (a) The N 1s binding energy shift under control of the working electrode (polycrystalline platinum foil) potential (as indicated); (b) the dependence of the binding energies of the B 1s, C 1s and Pt 4f_{7/2} core levels on the applied electrode potential. The system used a [EMIM][BF₄] ionic liquid electrolyte.

and a Pt working electrode (WE) in contact with the IL electrolyte [EMIM][BF₄]. XPS analysis was carried out at the interphase boundary between the IL and Pt WE, with selected associated spectra in Fig. 7. The binding energies (BE) for the electrolyte species core level spectra (B 1s, C 1s, N 1s) exhibited an electrochemical shift of -1.0 eV V^{-1} under the applied voltage because these atoms are in the outer Helmholtz plane (OHP) or the bulk of the electrolyte, and due to a lack of direct contact with the electrode, subsequently experience the full potential drop. The WE was grounded by the XPS system and therefore the Pt 4f BE and peak shape did not change. N 1s spectra were also recorded at very negative potentials, where the appearance of an extra peak at a potential of -1.9 V indicated a degradation product of an electrochemical reaction. Further examples of *in situ* UHV XPS cells include those of Wibowo *et al.*¹²⁶ and Gokturk *et al.*^{127,128}

Developments in XPS techniques that use synchrotron-based radiation sources as well as lab-based differentially pumped analysers and advanced electrostatic lenses now mean that it is possible to make *in situ* electrochemical XPS measurements at near ambient pressure (NAP-XPS), where NAP refers to a working pressure of up to a few millibars.¹²⁹ For now, these remain on the cusp of lab-based surface analysis, but work from multiple independent groups are bringing such cells closer to widespread use.^{130–134} *In situ* electrochemical XPS studies offer huge potential in electrochemistry research and further developments are expected to be made as the field progresses.

Cryo-XPS

Cryo-XPS employs a similar approach to the cryo-TEM technique that won Dubochet, Frank and Henderson the 2017 Nobel prize



for chemistry, and enables the characterisation of solid–liquid interfaces and volatiles. Samples are prepared as suspensions which are supported on a holder, fast frozen and introduced into the vacuum chamber. Usually, the sample is warmed slightly to a stage where multilayers of frozen volatiles sublime and a virtually unaltered liquid monolayer interface can be studied.

Cryo-XPS provides several clear advantages over conventional preparation by avoiding steps such as freeze drying which can damage the surfaces of more sensitive samples while providing ostensibly *in situ* measurements.^{136,137} Furthermore, the presence of a hydrated layer also reduces levels of adventitious carbon because these carbon contaminants tend to be trapped in the first few layers, which are lost during sublimation.^{138,139} Key considerations are the condensation of volatile species from the laboratory atmosphere onto the surface of the cooled sample and potential issues with sample warming on transfer between cooling, transfer and analysis stages. Ideally then an XPS system suitable for cryogenic analysis would have a glove box attached directly to the load dock allowing for sample preparation in an inert atmosphere^{140,141} and LN₂ cooling in both the load dock and analysis chamber, ensuring the sample is maintained at a low temperature during analysis.

Schukarev *et al.* have carried out several investigations probing water–mineral interfaces using the cryo-XPS approach, delineating the interaction between hydrated surfaces and the chemical speciation of absorbates.^{142,143} Pastes of metal oxides were prepared through the centrifugation of particles suspended in different pH environments. The wet pastes are then placed onto precooled transfer rods at 170 °C and 1×10^{-8} torr before introduction to the analysis chamber.¹⁴⁴ The zwitterionic nature of *N*-(phosphonomethyl)glycine (PMG) absorbed on goethite and manganite was studied using this method and provided information that would not have been possible for traditional XPS analysis. Fig. 8a), shows the stark difference between the N 1s regions recorded at cryogenic and room temperatures, with the absence of the ~402 eV peak due to a protonated nitrogen species at room temperature.^{135,145} Fig. 8b demonstrates the sensitivity of the approach to pH changes in the flash frozen solution with dramatic changes to the relative ratios of NH²⁺ and NH at 402 eV and 399.8 eV respectively with varying pH. The 402 eV decreases in intensity as pH increases. The experiment clearly demonstrates the potential for direct monitoring of the changes to the chemical speciation as a function of pH, which could not be extracted from a dry surface.

Similar studies have been carried out on sulfur rich systems such as pyrites and sphalerite, which until recently could not be accurately characterised in a UHV system due to the loss of sulfur species in vacuum.^{146,147} Ejtemaei *et al.*¹⁴⁸ combined cryo-XPS and depth profiling to study the activation of metal sulfides through the presence of Cu. Polished mineral samples were submerged in a copper sulfate solution for 10 minutes before being precooled to below –160 °C and maintained at that temperature for analysis by XPS. The samples were then etched 1–10 nm by argon ion bombardment. Survey scans of copper activated sphalerite (ZnS) showed copper ions present

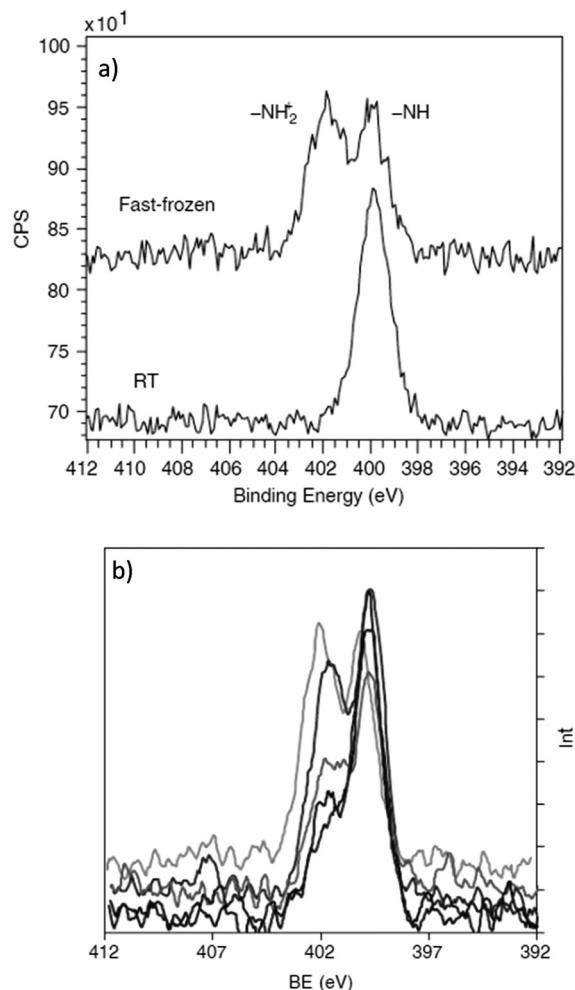


Fig. 8 (a) N 1s spectrum of PMG recorded on a MnOOH surface at room temperature versus fast frozen, evidencing an addition nitrogen species following preparation via cryo-transfer. (b) N 1s spectrum of PMG on MnOOH at pH 3.9, 4.9, 6.0, 7.9 and 8.6 (top to bottom spectral features at around 402 eV). Here we can see the pH dependence on the amount of protonated PMG species on the surface of MnOOH. Copyright © 2004 John Wiley & Sons, Ltd.¹³⁵

up to 10 nm deep (Fig. 9). Furthermore, the Cu and S 2p spectra show a greater proportion of Cu⁺ and S^{2–} compared to Cu²⁺ and S^{2–} species present on the surface following copper oxidation.

These findings were interpreted as evidence for a 2 step mechanism in the copper activation of sphalerite involving: Step 1: Cu ion exchange, Step 2: oxidation of the surface forming a CuS layer.

The study of soft matter including biological samples such as bacteria, is an emerging field within cryo-XPS. Fast freezing of cell walls can reduce the alterations and degradation caused by dehydration and beam damage, allowing for a more natural environment for analysis.¹⁴⁹ The preparation of biological matter is comparable to that of inorganic media with suspensions of cells inoculated and subsequently pelleted onto a sample holder and fast frozen.^{150–152}

Ramstedt *et al.*¹⁴⁹ have used Cryo-XPS to study bacterial cell walls treated in different environments, using a multivariate



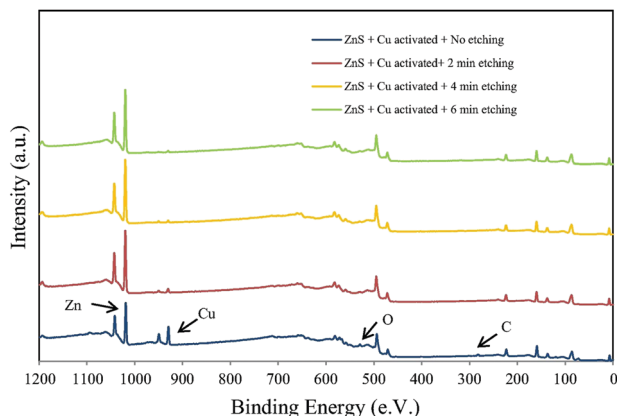


Fig. 9 Survey spectra of copper activated sphalerite after; no etching, 2 minutes, 4 minutes, and 6 minutes of etching corresponding to a depth of 3, 6, 10 nm respectively Reprinted from Ejtemaei *et al.*, Characterisation of sphalerite and pyrite surfaces activated by copper sulfate, *Miner. Eng.*, 2017, **100**, 223–232, Copyright (2007), with permission from Elsevier.¹⁴⁸

curve resolution analysis of carbon spectra to analyse the lipid, protein, and sugar content of the bacteria cell surface layers. The work demonstrates the capability to reveal changes in the cell wall chemistry depending on the solution pH,¹⁵⁰ with the chemical composition of *Bacillus subtilis* being radically different at low pH values compared to high pH, with much higher proportions of protein compared to lipids and carbohydrates. *Bacillus subtilis* tend to survive in environments between 6–9 pH and when exposed to more acidic environments the bacteria form a ‘spore coat’ which is higher in protein and low in sugars and lipids. Recently Shchukarev *et al.*¹⁵³ applied this method to successfully characterise the surface layers of both fungi and viruses, demonstrating that the method can be applied to a large range of biological systems, and illuminating mechanisms such as absorption and anti-microbial interactions.

It is interesting to consider the similarities and differences between cryo-XPS and near ambient pressure XPS, which both offer opportunities to examine surfaces in the presence of high concentrations of adsorbates, albeit in very different physical states. Kjærvi *et al.*¹⁵⁴ used the two techniques in a characterisation of the surface of *Pseudomonas fluorescens*. Fig. 10 shows that both methods give spectra with at least three clearly defined C 1s species at the surface of the bacteria and very little disparity can be immediately discerned between the two sets of data. However, the scattering of photoelectrons by gas phase components in NAP produces much lower signal intensities, and as a result additional S 2p and N 1s species were identified in the cryo-XPS experiments (not shown). Furthermore, there is a slight but significant increase in the 285.0 eV peak of the NAP spectra over that of the cryo-XPS. The authors propose that this difference arises from the presence of adventitious carbon on the surface of the NAP sample which was not present in its cryogenic counterpart because of the sublimation of the condensed water layers.

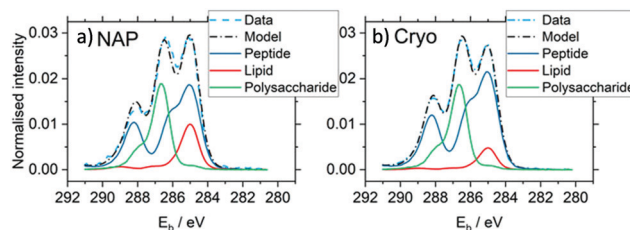


Fig. 10 A comparison of C 1s regions obtained by NAP and cryo-XPS evidencing the benefits cryo-XPS may bring to the analysis of materials in a traditional UHV atmosphere. These graphs show the application of multi-variate curve resolution analysis to the spectra.¹⁵⁴

Ion scattering spectroscopy

Ion scattering spectroscopy (ISS), also known as low energy ion scattering (LEIS), is a surface technique which involves probing the outermost surface layers of a solid using elastic collisions of ions. While ISS is generally used to provide surface identities and ratios, it may also be used in a semi-quantitative manner through comparison with standard samples.¹⁵⁵ Originally involving high energy (MeV) ions (and often referred to as Rutherford Backscattering Spectroscopy or RBS),¹⁵⁶ the ISS technique was refined to utilise lower energy ions to be more readily available in laboratories.^{157–159} It was the discovery that ion-surface collisions could be described kinematically by a simple two body elastic collision model that allowed for facile determination of surface composition:

$$E_1/E_0 = \left[\frac{\cos \theta \pm \left(\left(\frac{m_2}{m_1} \right)^2 - \sin^2 \theta \right)}{1 + \frac{m_2}{m_1}} \right] \quad (\text{provided } m_2 > m_1) \quad (2)$$

The dependence of peak energy position on atomic mass in ISS. E_0 = primary ion E_k , E_1 = primary ion E_k following a scattering process at θ° to initial trajectory, m_1 = mass of primary ion, m_2 = mass of target atom.

The extreme surface sensitivity of ISS allows it to offer additional structural information, complementary to that of pure XPS analysis and it has found use in many fields including catalysis,^{160–194} thin films,^{195–202} electronics^{195,203–206} and sensor materials.^{191,200–202,207,208} While ISS remains a valuable tool in materials analysis, care needs to be taken over sample choice and preparation. For example, ‘rough’ surfaces such as silicas commonly used in heterogeneous catalysis may exhibit broadened peaks in eventual ISS profiles.²⁰⁹ To overcome this, compacting samples into pellets may improve surface suitability, however, care must be taken so as not to alter the material properties.²¹⁰ Furthermore, ion scattering may have limitations depending on sample composition, since preferential sputtering can lead to erroneously low concentrations of easily sputtered components such as sodium.²¹¹



Morphological analysis

Owing to its enhanced surface sensitivity in comparison to XPS, which has an information depth of as much as 9–10 nm (using a typical Al $K\alpha$ X-ray source),²¹² ISS has found much use in extracting qualitative information on elemental distributions in the very uppermost atomic layers of a sample. The ability to probe only a single monolayer allows for a more sophisticated analysis of the atomic arrangement at the surface and near surface.

This information is of critical concern in fields such as heterogeneous catalysis, wherein performance and activity are directed by surface interactions of reactants onto active sites. Ceria based mixed oxides for example, possess unique surface properties which render them highly useful in a variety of applications. Surface segregation of either dopant or ceria, determined by ISS, has enabled a deeper understanding of how different thermal treatments or synthesis methods affect the resultant materials properties as catalysts for CO/alkane oxidations^{164,180,181,213} and dehydrogenation/hydrogen transfer²¹⁴ as well as for oxygen ion conductors/oxygen storage.^{181,215} Furthermore, since this is in effect an ion bombardment technique, the correct experimental parameters may induce sputtering of the surface of a material (even at low accelerating voltages), successive ISS spectra may therefore reveal an elemental gradient, affording further information which with knowledge of sputtering yields (or in comparison with alternative methods of probing the depth distribution of the outer surface) may even be used to prepare an elemental depth profile (Fig. 11).^{215,216}

Thin film development requires methodologies to determine surface coverage and monolayer build-up. ISS is uniquely suited to measuring a relatively large field area to probe full thin film coverage or detect remaining substrate signal.¹⁷⁷ For example, TiCrN thin films for drill coatings, developed using DC magnetron sputtering exhibit different properties when manufactured at different pulse frequencies. Analysis by XPS and ISS enabled the link between surface construction (and in particular, surface segregation of different species) and material properties to be determined.¹⁹⁸ Surface morphology probed by ISS may be advantageous, compared with purely XPS methodologies such as Tougaard's inelastic background modelling since it is free from any contributions from subsurface species (for example, underlying pores containing materials of competing photoemission). The caveat, however, is that multiple samples/measurements are often required for meaningful representation.

Structural changes following reaction

ISS may also be used to determine changes to the outermost surface following chemical or physical processing and has been used to study systems undergoing many common processes such as thermal treatments,^{217–219} hydrogen reduction^{161,220} or exposure to gases or vapours such as O₂ and H₂O.²²¹

ISS has also been used to study sintering, an important consideration in the field of catalysis, both in terms of the stability and reusability of the catalytic materials but also in the

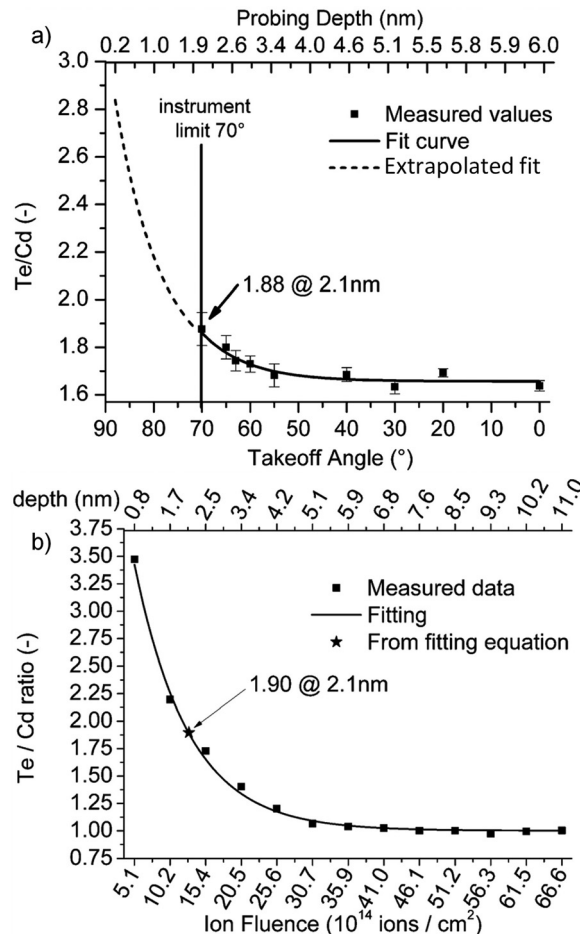


Fig. 11 (a) Te:Cd ratio from a sequence of ARXPS measurements compared with (b) a profile of sequential 5 keV Ne⁺ ISS measurements as a comparison probe of the two methods efficacy in the analysis of ultra-thin films.²¹⁶

preparation of the materials. Vapour deposition,²²² wet impregnation²²³ and even cluster deposition²²⁴ are often followed by a thermal treatment step to nucleate particles or remove organics, counterions or capping agents. Changes to surface coverage may give an insight into sintering processes which are difficult to observe by alternative means such as XRD or STM.¹⁷⁰ This principle applies to monitoring synthesis procedures to optimise conditions. For example, determining the content of transition metal sites within a PtCo electrocatalyst/polymer electrolyte membrane fuel cell. Marked performance changes are observed depending on the number of transition metal sites within the surface layer. Thermal processing and ISS/XPS was used to optimise preparation methods and monitor surface properties in combination with catalytic studies.²¹⁷

ISS is also a viable method for monitoring the surface of a system during exposure to working conditions. For example, alloy oxidation is a key parameter to optimal material properties, ISS/XPS confirmed a Grimley-modified Cabrera–Mott model for the oxidation of a multitude of mixed metal oxides during plasma oxidation, as well as using ISS etching to profile the distribution of elements.²⁰⁵ Similarly, in catalysis, Ag foils



exposed to NO/O₂ mixtures to model the catalytic process of NO_x removal showed structural reorganisation, forming oxidic and nitric sandwich layers.¹⁹³ Pt based automotive three-way catalysts were studied by ISS to monitor coking under different conditions the enhanced surface sensitivity of ISS in comparison to XPS enabling the analysis of catalysts with active component loadings below the XPS limit of ~0.5 wt%.¹⁶⁷

Temperature dependant ISS (TD-ISS)

ISS has found use in several sectors in conjunction with variable heating stages in order to utilise the unique information available to ISS as a probe for surface changes resulting from an increase in system temperature.

In one example of this approach, size selective Ir clusters deposited onto TiO₂ surfaces for model catalytic studies made use of a very simplistic form of TD-ISS in which model surfaces were analysed by XPS and ISS before and after a the temperature programmed desorption (TPD) of CO in order to assess surface restructuring.¹⁶⁰ It was observed that following a TPD procedure, the surface Ir intensity (ISS) greatly diminished.

While XP data indicated that Ir desorption was not the likely cause of this attenuation, some attenuation suggested that perhaps the composite was restructuring either *via* sintering or encapsulation of the Ir centres. A second TPD was performed and revealed markedly differing behaviour, which combined with the ISS and XPS data suggested the Ir centres were undergoing encapsulation by the reducible support (a strong-metal support interaction which has been observed for a number of metal-titania systems).^{166,225,226}

Monitoring dynamic changes to ISS spectra as an *in situ* method of determining temperature-structure relationships, rather than an *ex situ* pre- and post-process snapshot, can offer even greater insights. For example, following exposure to CO, small (Pd₂) and large (Pd₂₀) clusters on TiO₂ were analysed by ISS whilst heating from 100 K to 500 K (Fig. 12).

Comparisons of these changing ISS ratios with CO TPDs led to the conclusion that low temperature CO desorption in the TPD was due to CO adsorbed onto Pd-TiO₂ peripheral sites, rather than directly on-top of Pd clusters. A similar experimental procedure has been used to study other Pd cluster/metal oxide systems^{172,173} as well as the relationships between Pt cluster sizes and ethylene²²⁷ and diborane²²⁸ surface reactions.

Modelling growth modes

When considering nanoparticle or thin-film growth over a substrate (of non-matching identities), knowledge of nucleation and growth modes permits sophisticated and precise control over nanomaterial synthesis. ISS analysis combined with XPS affords a simple methodology for modelling the growth based on the calculated surface coverage or peak intensity relative ratios.¹⁸⁹ It is important to note, that the related higher energy technique, RBS, provides beneficial complementary here in that it may probe film thicknesses up to ~200 nm due to the deeper probing depth compared with ISS.²²⁹ The efficacy of RBS, however, diminishes with smaller overlayer thicknesses (<3 nm) however, and as such ISS is much more suited for

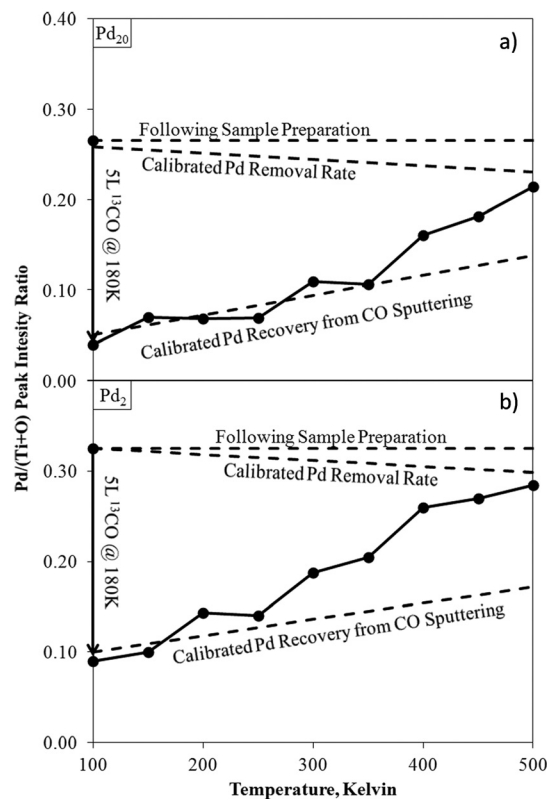


Fig. 12 Pd/(Ti + O) ISS intensity ratios for Pd₂₀/TiO₂ (a) and Pd₂/TiO₂ (b) as-deposited, after exposure to CO at 180 K, and during temperature ramping. Calibrated rates for recovery of Pd from ISS sputtering.¹⁷¹ Calibration lines account for unintended sputtering effects from the ISS experimental conditions selected for this analysis. Deviations from this calibration indicate a changing surface structure (desorption of CO).

applications such as modelling growth modes on the atomic scale.

For example, surface coverages and XPS peak ratios enabled González-Elipe *et al.* to study a range of metal oxide-metal oxide interfaces, model the growth of the different oxide combinations and relate said models to binding energy and modified Auger parameter effects in the XPS spectra.^{200–202,230} ISS was also used to shed light on the nucleation and growth of iron oxides on Au(111), in which the rate of change of surface coverage with calculated monolayer coverage provided insights into a change in growth mode at 0.5 ML from layer-by-layer to island growth.¹⁹⁶ STM was used alongside the XPS and ISS to confirm these observations.

More complex systems, involving mixed oxide interfacial layers of multiple metal site systems may make use of the combined power of ISS analysis and subsequent modelling, with the chemical information gleaned from XPS (particularly angle resolved XPS) to construct sophisticated models of changing interfacial species and growth modes. When modelling the deposition of titanium onto Si(100) Palacio and Arranz measured the Ti and Si ISS signals at regular time intervals during the deposition process and used them to rule out Volmer-Weber growth during the initial stages.¹⁹⁵ Combining ISS with XPS, they were also able to determine that the process involved a period of



titanium silicate growth (up to 4 ML) followed by Stranski-Krastanov growth of metallic Ti.

Similarly, Liu *et al.* analysed platinum deposition on nickel using ISS and XPS (plus the application of overlayer models to the XPS data) to develop a model for the alloying and growth of the adsorbate layer.²³¹

Cu/CeO₂ catalysts are widely used for oxidations, water-gas shift and methanol synthesis. Measuring the ISS spectra of a range of depositions of discrete amounts of Cu onto a Ce(111) surface and normalising the signal to that of bulk (>40 ML) Cu/CeO₂ resulted in the profiles observed in Fig. 13.¹⁶⁹ Theoretical profiles for layer-by-layer (Frank-van der Merwe) and hemispherical cap growth (island growth of uniform

hemispherical particles)²³² are superimposed onto the measured datapoints to illustrate the correct growth formation. Combination of the surface coverage from ISS with the XPS data and knowledge of the growth mode (and parameters such as IMFP) allows for high sophisticated calculations of particle thickness/radius.²³²

Measuring surface potential

Ion-induced secondary electron (IISE) emission is an important material property to be considered in the production of coatings for scientific equipment, such as non-evaporable getters for distributed pumping at CERN.²³³ ISS has been used as an effective probe for determining the surface potential of thin films under ion irradiation. Nagatomi *et al.* derived an experimental approach for the study of MgO films on silica (at the time an important material composite for plasma screen displays) in which irradiation of the surface with 950 eV He⁺ ions with an electron flood gun was used to determine the resulting surface potential.^{234,235}

Ultraviolet photoelectron spectroscopy (UPS)

The process giving rise to UPS is in essence identical to that of XPS. The only difference is the energy of the incoming photons and thus the source of the photoelectrons that are emitted. Typically, UPS is performed using a He(I) or He(II) lamp emissions (21.22 and 40.8 eV respectively), relating to 1s 2p → 1s² relaxations from He atoms (He(I)) and ions (He(II)).²³⁶ These lower energy sources limit this technique to the analysis of valence shell electrons and molecular orbitals, or density-of-states (DOS) structure, and also results in a shallower sampling depth (due to the reduced kinetic energy of emitted photoelectrons). The technique has the potential to provide powerful spectroscopic representation of molecular orbitals and can reveal information on their bonding or antibonding character, how the bonding is localised in a molecule and also changes to molecular geometry.^{236–238} It is particularly powerful in the study of molecules in the gas phase where multiple emission states may arise as a result of vibrational lines on top of purely energetic transitions but UPS is also a powerful probe of solid and composite materials, occasionally as a simple fingerprint technique or through more detailed understanding of the emissions. UPS has found use in the analysis of polymers,^{239,240} self-assembled monolayers (SAMs),²⁴¹ catalysts^{242–244} and semiconductors/electronic materials.^{245,246}

Valence band structure

A simple application of UPS is as a tool to fingerprint valence structure to obtain information either not available using X-rays or to glean information from closer to the surface. In the case of metal-semi/non-conductor systems (prevalent in the field of catalysis), emissions near the Fermi energy allow monitoring of the presence of metallic bonding. For example, Roberts *et al.* used the presence of photoemission intensity near *E_F* to discern the amount of metallic bonding in the

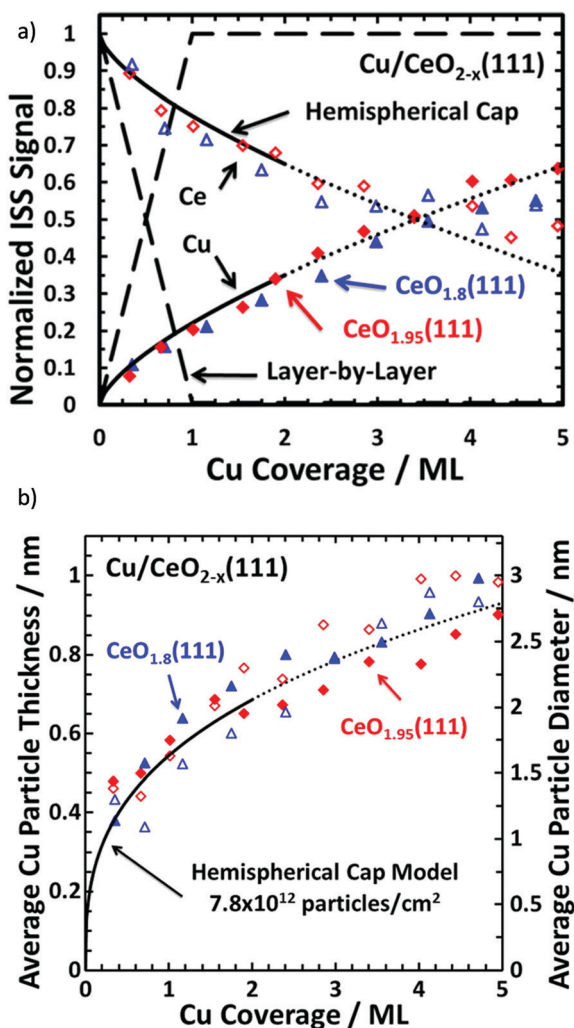


Fig. 13 (a) Normalised ISS intensities for Cu on CeO₂(111) with model growth modes for comparison and (b) average Cu thickness versus Cu surface coverage. The dashed line represents the normalised ISS signal to be observed if Cu grew in a layer-by-layer fashion, whereas the solid line represents Cu growing as hemispherical caps with a fixed radius and a fixed particle density of 7.8×10^{12} particles per cm². This model is only reasonable up to ~35% surface coverage (above this, potential particulate overlaps render the model unreliable) and the dotted lines are purely for visualisation. Reprinted (adapted) with permission from ref. 169. Copyright (2015) American Chemical Society¹⁶⁹



deposition of size selective clusters of Pd on TiO₂ and correlate high catalytic activity with higher Pd 4d occupancy (*vs.* Pd 5s).²⁴⁷ Spectral intensity (or lack thereof) around the Fermi energy of MoO₃ based catalysts for alkane/ene cracking was used to monitor changes to the molybdenum centres following hydrogen reduction.¹⁶¹ Prior to reduction, no structure near the Fermi level could be observed, however following thermal treatment under H₂ at 673 K, π and σ bands from MoO₂ could be observed in the DOS as well as a clear emission around the Fermi energy. While the XPS did indeed show a marked reduction from MoO₃ to a mixed MoO₃–MoO₂ system, no metallic Mo could be observed. The UPS however, indicated that the system was largely MoO₂ with metallic sites present, suggesting these species may be found predominantly at the surface. UPS and XPS were further used to determine the degree of reduction following doping with Pt and K where it was revealed that K doping enables the full reduction of MoO₃ to MoO₂. Such utilisation of UPS is often hindered by strong overlapping signals in mixed or complex systems, however small windows in the spectra can provide critical insights in many structures.

Electronic structure

Possibly the most widely used application of UPS is the analysis of electronic states (for example in semiconductors). This requires slightly more detailed analysis and understanding of the photoemission in a system than the simple fingerprint approach.

Since UPS concerns itself with the energy region associated with valence orbitals, it is primarily analysing the electrons concerned when discussing DOS. XPS may also be used to record these of course, but with the penalty of reduced resolution due to the greater line width of the exciting photons sources.²⁴⁸ While not often a requirement for pure metals, the enhanced resolution allows for more detailed analysis of composite materials where determination of orbital characteristics in the DOS may be more complicated.²⁴⁹

There are a number of terms which are vital to understanding molecular electronics including, but not limited to: the vacuum level (E_{VAC} – the energy level at which an electron is at rest a few nm outside of the atom), the Fermi energy (E_{F} – in metals this is the boundary between occupied and unoccupied states in the continuum, in semiconductors this is a statistical state within the band gap) and work function (the energy required to take an electron from E_{F} to E_{VAC}).²⁵⁰ Understanding the work function is key to improving our knowledge of semiconductors and developing improved materials. While it is possible to measure work function using XPS, UPS is traditionally favoured due to its superior line width and flux.²⁵¹

Sample work function may be determined using the spectral width (this is the energy difference between the secondary electron cut-off and the Fermi energy),²⁵² though while this process may be simple for conductors, it is less straightforward for semiconductors because E_{F} lies within the band gap. Analysis of the high kinetic energy edge permits determination of the Fermi edge of a material, while the low kinetic energy

edge provides the position of the secondary electron cut-off. Sample biases may also be applied in order to decouple emissions arising from secondary electrons associated with the spectrometer.²⁵³ It should also be noted, that while this review does not intend to dive deeply into (N)AP-XPS, powerful methodologies have been determined for the determination of material work functions through the analysis of gas phase molecules in the vicinity of the material surface.²⁵⁴ While this analysis requires the use of a (N)AP-XPS system – it serves as an additional probe and may be a powerful combinatorial approach with UPS, enabling the more facile determination of the valence band maximum (VBM) of semiconductors and assisting with the interpretation of the UP spectrum.

Using this data in combination with a band gap measurement (*e.g.* from UV-Vis), one may construct a full band energy diagram – including the positions of the Fermi level, conduction band minimum (CBM) and VBM.^{255–257}

TiO₂ represents a major class of semiconducting solids in which the understanding of the electronic structure of the pure material, as well as modified, doped and mixed oxide systems is important for the development of advanced functional materials in fields such as photocatalysis,²⁵⁸ solar cells²⁵⁹ and gas sensors.²⁶⁰

As a result it has been studied extensively using UPS with particular interest in the calculation of absolute values of the CBM and VBM when for example, Tauc-plots.²⁶¹ generated from UV-Vis spectroscopy permits only relative values. Experimentally, recording UPS spectra from non- or semi-conducting samples represents a challenge due to a more demanding need for sample homogeneity and electrical contact with the system. Maheu *et al.* compared two preparation methods which have shown some promise as means of acquiring UPS spectra from granular semiconducting materials: electrophoretic deposition (EPD) and drop casting a suspension of TiO₂ in alkanes.²⁶² The objective was to study anatase, rutile and P-25 (a mixture of rutile and anatase often used as a benchmark material for photocatalytic applications due to its beneficial electronic properties) in an effort to understand the nature of the unusually high photoactivity of P-25 when compared with other forms of titania.

While EPD prepared a sample in which electrical contact was sufficient to overcome any potential charging, drop casting required the determination of the absolute position of the CBM and VBM energies using the signal from the Ag foil substrate. Both methods were able to determine the E_{VB} of anatase and rutile from which an energy level diagram and representation of DOS was developed (Fig. 14), correlating well with theoretical predictions.²⁶³

The inclusion of heating stages in many modern spectrometers raises the possibility of combining procedures such as those demonstrated by Maheu *et al.* with annealing processes capable of producing oxygen defects in TiO₂ with the ability to effectively neutralise charge build up and further augment the spectral acquisition process.²⁶⁴ Note this method makes use of a modified background developed by Li *et al.*, designed and optimised for UPS spectra.²⁶⁵



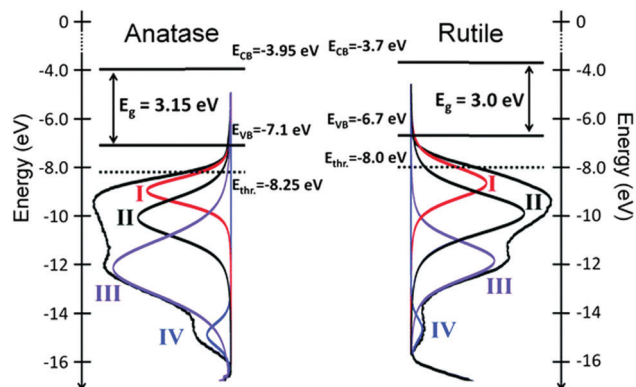


Fig. 14 He(I) UPS profiles, fit with an established empirical model and an overlaid electronic structure of the two TiO₂ polymorphs anatase and rutile relative to the vacuum level. E_{thr} corresponds to the deepest state that can be excited in a photocatalytic test ($E_{\text{photons}} \leq 4.3$ eV).²⁶²

Whilst the ability to calculate the work function of a material is vital to semiconductors and electronics,²⁶⁶ its also very important to chemical performance in many other areas including catalysis,^{267,268} photochemistry^{269,270} and chemical sensors.^{271,272} For example, work function changes were correlated with molecular dipole moment measurements of partially fluorinated alkyl and aromatic phosphonate SAMs on a Zn-terminated ZnO(0001) surface to create a tuneable ZnO work function that could be tailored to improve the efficiency of organic–inorganic electronic devices (Fig. 15).²⁷³

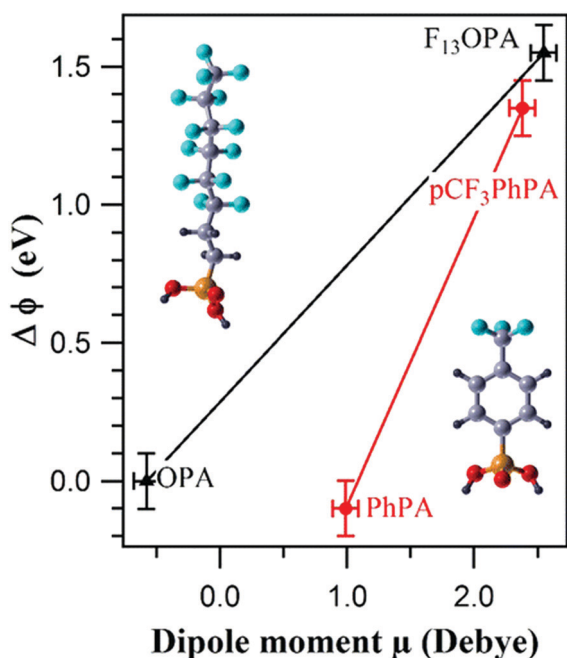


Fig. 15 Correlation between the experimental change in work function ($\Delta\phi$) and the molecular dipole moment μ of the PA molecules for 3,3,4,4,5,5,6,6,7,7,8,8,8-tridecafluorooctyl phosphonic acid-((F₁₃)OPA, left molecule in figure) and *p*-(trifluoromethyl)phenylphosphonic acid-((pCF₃)PhPA, right molecule in figure) modified ZnO surfaces, along with corresponding non-fluorinated analogues.²⁷³

In some cases, work function measurements can also reveal insights into growth modes. Monitoring the work function of Ag deposited onto MgO for example, revealed a valley-like gradient which was ascribed to Ag atoms binding to surface defects and forming sites of nucleation for further growth.²⁷⁴

Organic semiconductors have gained popularity in the 21st century, wherein widely conjugated π -systems take the role of the solid system.²⁷⁵ AR-UPS has been able to evidence the electronic structure and charge carrier ability of such systems, though experimental methods are often not so straightforward, requiring organic single crystals which may need extra consideration for charge compensation and supplementary excitation sources.^{276,277} The result of the utilisation of AR-UPS, however, is that a picture of the electronic structure of the HOMO-derived valence band may be derived, which is fundamental to understanding the nature of positive charge carriers in organic semiconductors.²⁷⁷

Adsorbate analysis

UPS analysis of molecular adsorbates has for a long time been the preserve mainly of the heterogeneous catalysis community, but the advent of hybrid organic–inorganic systems has resulted in a much wider interest in the interaction between substrate and adsorbate as materials designed for electronics, sensors and optics benefit from an increased understanding of the bonding and interfaces involved.²⁷⁸

CO adsorption was an early model for this field, since metal–CO interactions are prevalent among a multitude of chemical catalytic transformations. CO adsorbed onto Pd(111) produced 2 of the three spectral emissions associated with the gas phase spectrum of CO (namely, the 1π and 5σ features).²⁷⁹ The 5σ is primarily situated on the carbon, the 1π on the oxygen and the 4σ an oxygen lone pair. Due to final-state screening even in the absence of any bonding^{280,281} adsorbate binding energies will decrease by 2–3 eV which led to the proposed model that the 1π and 5σ orbitals are non-bonding and the 4σ orbital increases in energy (due to the interaction with the metal) to the point where it is no longer observable on the UPS spectrum. This was observed to be true for a number of systems²⁸² and led to the explanation that the CO 5σ orbital is responsible for the bonding of this system *via* the carbon atom, and donates electron density to the vacant d-orbitals of the metal and subsequent back-donation from filled d-orbitals is directed into unoccupied CO $2\pi^*$ antibonding orbitals, a model which would come to be known as the Dewar–Chatt–Duncanson (DCD) model.

The 5σ peak can actually be found underneath the 1π peak using angle resolved UPS,²⁸³ while inverse photoemission spectroscopy was used to identify the partially filled $2\pi^*$ antibonding orbital.²⁸⁴

UPS was also used to study the effect of potassium and caesium doping of Pt(111) (plus O₂ preadsorption) on the adsorption of ethylene.²⁸⁵ Firstly, ethylene was adsorbed (π -bonded – mild disruption to double bond) on the surface of the crystal at 37 K so as to avoid any double-bond dissociation before heating to 57 and 300 K which resulted in the formation



of a di- σ species (major disruption to the double bond followed by the formation of an ethynyl molecule on the Pt surface. Multilayer formation at 37 K was observed in the presence of alkali-dopants, with minimal distortions to the adsorbed molecular orbitals, however the introduction of halides, resulted in a more weakly bound species with UPS revealing that with increasing alkali coverage there were fewer di- σ species formed at higher temperatures, to the point where π -ethylene is the only adsorbate (desorbing at 130 K).^{286,287}

This analysis is not limited to observing changes in the adsorbate itself, however, as electronic changes induced by the process of adsorption may also produce simultaneous spectral features which provide insight into surface reaction mechanics. One example of this is measurement of the work function of a metal surface before and after adsorption.^{285–290} For a clean Pt(111) surface, ethylene adsorption reduces the work function (ϕ) by 1.05 eV, consistent with the adsorbate UPS and vibration spectroscopy measurements in observing a large charge transfer from ethylene to Pt(111) upon forming the di- σ species.²⁸⁷ Adsorption of potassium onto Pt(111) reduces ϕ by 4.8 eV, after which adsorption of ethylene induces very little change in ϕ , again consistent with the observation from UPS/TPD that alkali doping of Pt(111) inhibits σ -bonding of ethylene to the metal surface. In contrast, CO adsorbed onto Pt(111) increases the work function by 1.5 eV,²⁹¹ consistent with metal to adsorbate charge transfer in the form of π -back bonding. The combination of UPS spectral features and substrate analysis, particularly when used with complementary techniques such as HREELS and TPD, has been used in this way to provide insight into a bounty of systems, studying the combination and reaction of free molecule and active site and facilitating a greater understanding of catalytic processes.^{292–297}

Monolayer orientation

UPS has been used to reveal information regarding the orientation and organization of self-assembled alkyl chains adsorbed on to metal films, often in conjunction with XPS in order to probe monolayer thickness.²⁹⁸ Chain tilt angles for alkanethiols on metal facets (e.g. Ag(111) vs. Au(111)) differ significantly, and this difference was noted in the UPS spectra.^{299,300} It was observed that not only can UPS identify adsorbates in different orientations (e.g. standing vs. lying chains), but it can also identify them independently, unlike techniques such as Fourier transformed infrared (FT-IR) and near-edge X-ray absorption (NEXAFS) spectroscopies which report an average tilt-angle.³⁰¹ Measurements recorded at differing take off angles (TOA) may also yield information regarding the degree of homogeneity in the orientation of the chains,³⁰² where spectral features which do not change as a function of TOA are indicative of randomly oriented chains while TOA dependence suggests a good molecular orientation.³⁰³ Furthermore, the intensity of the Fermi edge as a function of TOA may yield additional insights into the order of the adsorbates.³⁰⁴ This technique has found use with more complex surface molecules, such as 2- and 4-chlorobenzylmercaptan, adsorbed on gold films. Identification of spectral features from model gas phase UPS spectra permitted the identification of both

the *ortho*- and *para*-positions of the benzene ring as well as the orientation of the molecules, although this technique is often performed using a related technique called metastable induced emission spectroscopy (MIES), due to its superior surface sensitivity.³⁰⁵

Auger electron spectroscopy

Auger electron spectroscopy (AES) is a type of secondary electron emission spectroscopy, resulting from the bombardment of a surface by an electron beam. AES complements XPS and ISS through its highly surface sensitive probing (~ 5 atomic layers depending on energy) and is capable of extremely specific elemental identification.³⁰⁶ Furthermore, unlike ISS, it is capable of determining chemical or valence states in certain cases.³⁰⁷ In the case of carbon, for example, analysis of the KLL Auger peak (involving valence electrons) may provide more chemical information than even XPS.³⁰⁷

Auger emission is a widely known process and occurs as an inter-atomic response to the photoemission process caused from the irradiation by X-rays or electrons. Following emission of an electron, the atom is left in an ionised, energetically unfavourable state with a core-hole formed in a low-lying orbital. Relaxation for this state comes as an electron in a higher orbital may drop down to fill the core-hole, releasing excess energy which may then facilitate the emission of a secondary electron, the “Auger” electron, with a kinetic energy that is independent of the exciting radiation/electron.

Analysis of films and overlayers

The enhanced surface sensitivity of AES (when applied to lab-based sources) results in a lower window for film thickness analysis than for XPS for example, with its suitability limited to ultrathin layers only (though it may be possible to analyse even up to 100 Å in some cases depending on the element).³⁰⁸ Much as with XPS, this surface sensitivity may be used to produce depth profiles in combination with a destructive ion bombardment technique – though with highly surface sensitive measurements, care must be taken to avoid topological changes which can affect the resulting profile.³⁰⁹

An example of Auger spectroscopy being used in this way is for the determination of graphene layer thickness. Multilayer graphene thickness is commonly determined by Raman spectroscopy,³¹⁰ however, there are major drawbacks with using Raman from complications arising from graphene on substrates that negatively affect the Raman signal (e.g. Ni³¹¹ and GaAs³¹²). The use of AES avoids these issues. Importantly, it is also possible to determine the thickness or number of graphene layers from a single, non-destructive spectrum – providing an advantage over most XPS methods and may also determine crucial sample properties such as electronic coupling between graphene and substrate, observe structural defects and quantify dopants or chemical impurities.³¹³

For adsorbates, monolayers and multilayers the analysis of AES peak maxima of substrate (I_s) and adsorbate (I_a), can



readily probe adsorbate growth modes and determine monolayer coverage even if sticking probability changes with surface coverage.³¹⁴ Such a technique enables the calibration of monolayer coverage and subsequent calculation of a coefficient of attenuation, which may be used to determine the surface coverage of an unknown sample (e.g. carbon monolayer coverage on a spent catalyst).³¹⁵ Furthermore, a simple graphical model of I_s/I_a allows for the facile distinction between Volmer–Weber, Stranski–Krastanov and Frank–van der Merwe growth of films.³¹⁶

Surface mapping

Since AES utilises an exciting beam of electrons, rather than X-rays, there exists a much finer spatial resolution on the area of illumination than for traditional XPS. This enhancement permits much more detailed surface imaging than XPS imaging, with image resolutions capable of reaching 10 s of nm, rather than 1000 s.³¹⁷

Modern spectrometers often combine the traditional XPS configuration with a field emission gun and provide both AES map scanning and scanning emission microscopy into one unit.³¹⁸ This combination facilitates the simultaneous acquisition of a morphological/topological image with that of a chemical surface map, which may be superimposed upon one another (Fig. 16).

X-ray generation of auger electrons

The scope of this review has so far focused on AES using an electron gun as the exciting beam, however X-rays are capable of producing augers too and the generation of X-ray induced auger electrons represents an invaluable tool for the chemical analysis of several key systems. Auger peaks often provide a fingerprint method by which to identify chemical compounds by which XPS may not be possible or straightforward (such as some first row transition metal compounds).³¹⁹ The classic example of the importance of auger electrons in XPS is the speciation of copper. Copper metal and copper(I) are virtually indistinguishable upon analysis of the Cu 2p photoemission, however the Cu LMM auger produces a difference of around 2 eV, facilitating the chemical distinction of the sample.³²⁰ This may be taken further with the use of the modified auger parameter (α'),^{321,322} which is defined as the product of the kinetic energy of the auger emission and the binding energy of the primary photoelectron peak (e.g. Cu LMM E_k + Cu 2p E_B). This parameter is advantageous, since it is independent of both exciting energy and any calibration factors applied to non-conductive material data. When working with the modified auger parameter, it is often beneficial to refer to, or produce, a Wagner plot^{322,323} – a graphical means by which to visualise trends in α' across a series of compounds.

In the analysis of nanomaterials, α' may provide even more insight into the surface chemistry and inter-particulate interactions.^{324,325} Particle size effects and metal–support interactions between supported metal nanoparticles and oxidic supports have been observed to influence the resultant α' . Since α' depends on initial and final state effects, an estimate

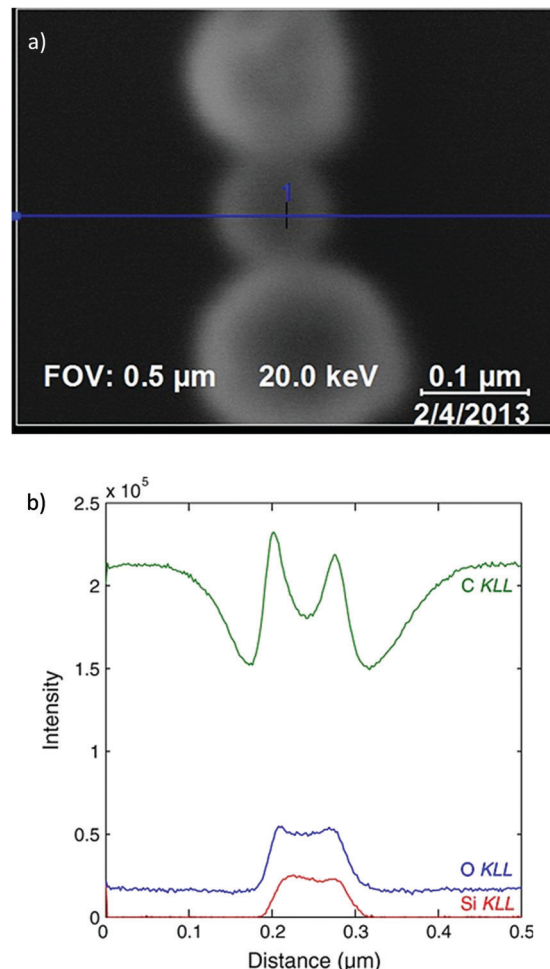


Fig. 16 (a) Scanning electron microscopy (SEM) micrograph indicating region of analysis for corresponding AES spectrum (blue line), (b) resultant AES profiles providing chemical analysis along the aforementioned spatial dimension.³²⁶

of the relaxation energy or degree of screening in the presence of core holes may be extracted from changes to α' .³²⁷ An increase to α' is associated with a higher relaxation energy/screening energy – which may be correlated with increasing particle sizes (and subsequently an increase in the average number of neighbouring atoms capable of screening the core-hole).³²⁴ Metal–support interactions may also influence this parameter, with more polarizable supports enhancing the resultant α' .³²⁸

EELS & REELS

Electron energy loss spectroscopy (EELS) involves a beam of monoenergetic electrons, (typically with kinetic energies < 2 keV), incident with a sample of interest. Inelastic scattering collisions occur from which information about the electronic state of the surface can be deduced.

In reflection electron energy loss spectroscopy (REELS), both elastically and inelastically scattered electrons are collected. The elastically scattered electrons mostly reflect directly from



the surface without losing any energy, forming the most intense part of the REELS spectrum which defines the 0 eV position on the energy loss scale. The low-energy loss part of the spectrum (< 50 eV) can then be used to determine information about the band structure and dielectric properties of the sample.

The high-energy loss region (generally > 50 eV) contains the ionisation edges due to inner shell ionisation, whose features are equivalent of an absorption edge in X-ray absorption spectroscopy (XAS). They are characteristic of the elements present in the sample surface (*ca.* 15 nm) and their local bonding environment. Therefore, the information from EELS is complementary to XPS and other surface sensitive techniques.

Probing surface valence and conduction band electronic properties

EELS is particularly sensitive to the valence region of a solid surface promoting electrons from the occupied valence band to the unoccupied conduction band and as a result EELS is capable of identifying chemical states that may not be determined through other methods.

For instance, SnO and SnO₂ have similar binding energies in XP spectra. Neither XPS nor AES can unambiguously distinguish between the various forms of tin oxides.^{329–331} However, Hoflund *et al.* were able to use EELS to distinguish between metallic Sn, SnO, and even suboxides believed to be Sn₃O₄ and SnO₂.^{332–334} Non-destructive depth profile information may also be obtained by the simple process of tuning the primary electron beam energy. Such information is crucial for understanding the mechanisms of various surface modification processes. In the example of EELS analysis of the oxidation of tin surfaces, it was revealed that the oxidation process underwent a three-dimensional growth mechanism within the subsurface layers, while the more oxidic species were formed at the metal–oxide interface.^{333,334}

A major drawback in the use of EELS is that, for many of the most electronically relevant solids (*e.g.* mixed oxides/perovskites), the EELS spectra are too complex to interpret. Typically, the intensities and the energy positions of EELS features from interband transitions are determined by the dipole selection rules for single-particle excitations. It also depends on the joint density of filled and unfilled states of the solid.³³⁵ However, with the presence of delocalised electrons in a polycrystalline sample, a combination of dipole and non-dipole active transitions in the EELS spectra may occur. For accurate interpretation, EELS then requires a detailed knowledge of the electronic structure within the solid, prior to unravelling the various components of the resultant spectra.

EELS can provide a useful insight into the interaction between probe molecules and a model surface and, like many surface science techniques has found much use, within the field of catalysis as a result. One system to have been investigated thoroughly is that of Ni and NiO which are strongly catalytically active (yet cheap and abundant) materials; EELS studies have been made of the adsorption of various gases, such as O₂,^{336–339} CO,^{336,340–342} H₂,^{340,341,343} H₂O,³⁴⁴ N₂,^{345–347} and NO,³⁴⁸ on clean Ni and on Ni alloyed with, or following the deposition of, other

metals, such as Al,^{349,350} Cu,^{351,352} Fe,^{353,354} Mo,³⁵⁵ and the alkali metals (Na, and Cs).^{356,357} In the case of NiO, both EELS studies of single crystalline^{336,358–364} and polycrystalline NiO^{365–367} have been carried out and EELS has also been used to monitor changes to other material properties of NiO, such as temperature dependence^{368,369} and optical behaviours.^{370,371}

However, it becomes apparent from a survey of the available data that the literature lacks consistency of peak assignments and interpretations for the Ni and NiO EELS spectra. Part of this could be due to the lack of detailed systematic investigation but can also be attributed to the need, as mentioned above, for an accurate knowledge of the electronic structure. Hoflund *et al.* have addressed that issue with an extensive study of the electronic structure of Ni and NiO both experimentally and theoretically.^{372,373} With the aid of this relevant information, Hoflund *et al.* were able to carry out the peak assignments of EELS spectra of Ni and NiO obtained from different depths of the samples with systematic alteration of the primary electron energy from 100 to 1150 eV.³³⁵ In this approach, the surface and bulk loss peaks can be more clearly distinguished, since an enhancement in the surface sensitivity is more noticeable than the surface loss features, which leads to the improvement of the EELS peak assignments.

Finally, of note is the use of EELS in the probing of plasmons, screened and unscreened final states – of great importance when confronted with the complex multiplet structures associated with transition metal photoemission spectra. As previously mentioned, EELS was used to study the electronic structure of Ni metal, however as well as purely evaluating EELS data of the metal, the combined use of EELS with XPS has resulted in the determination of peaks in the XPS spectrum too.³⁷⁴ Nickel metal, as well as producing a typically asymmetric photoemission, produces two additional broad peaks at +3.7 and +6.0 eV. While there had been some debate on the origin of the major emission at 6.0 eV, Grovesnor *et al.* used REELS to determine both surface and bulk plasmons as the primary origin of these emissions (though small shake-up features could not be discounted as contributing), and further observed a similar process within Co metal.³⁷⁵ Similar insights are available to more complex systems, such as metal oxides^{376,377} and semiconductors.³⁷⁸ For example, REELS has been used in combination with XPS and theoretical calculations to probe OsO₂ where again, REELS was able to identify the presence and energy of a plasmon. This measurement contributed to the complete understanding of the electronic configuration through theoretical modelling of the density-of-states applied to the valence region of the photoelectron spectrum.³⁷⁶

Analysis of inelastic scattering of electrons

As discussed at the beginning of this review, upon exiting a solid material, photoelectrons may undergo scattering processes which reduce their kinetic energies, contributing to asymmetric peaks, spectral backgrounds and satellite peaks.³⁸ We have covered the importance of the analysis of overlayers and introduced several non-destructive techniques by which we can probe these such as the Tougaard background analysis



method.⁴⁴ One parameter which is of paramount importance in this analysis is an understanding of the inelastic background produced by the overlayer. While the QUASES software provides some elegant solutions to this, one opportunity provided by REELS is the ability to probe the inelastic scattering *via* electron beam. This may then be modelled to not only build a background model, but also provide additional insights into the surface chemistry. For example, TiO₂ overlayer formation on Ti metal was analysed using REELS to produce an inelastic background,³⁷⁹ which was then applied to the Ti 2p region and the resultant background subtraction revealed weak suboxides which may otherwise have not been observed.

Band gap determination for semiconductor materials

Modern device construction often requires the controlled deposition of monolayers with desired properties one of which, that is often crucial for optimised device performance, is the band gap of the materials involved. However, the usual optical means (*i.e.* a Kubelka–Munk function on UV-Vis spectra)²⁶¹ often fail when a layer is thin or grown on a different substrate. For samples with a well-defined band gap, the REELS technique provides a convenient method of measurement in these cases since REELS' electrons must give up at least the energy equivalent to the band gap. Furthermore, while it is already a surface sensitive technique, decreasing the energy of the incoming electron beam can augment the surface sensitivity and hence, the band gap of thin films may be more accurately determined.

Vos *et al.* have used REELS under controlled conditions to extensively study a series of semiconductors with well-established band gap values.³⁸⁰ They discovered that, without surface preparation, it was possible to get a meaningful value of the band gap if a relatively high incoming energy of ~ 5 keV was used. When the energy was below 1 keV however, the intensity of the REELS spectrum inside the gap is insignificant. This usually caused a fading of the onset, and resulted in an inability to record the band gap energy.

Recently, REELS has also been successfully used to investigate ultrathin gate oxide materials applied in complementary metal-oxide-semiconductor (CMOS).^{381–384} Transition metal oxides such as HfO₂, ZrO₂, and their compounds deposited on SiO₂ or Al₂O₃ have been exploited as alternative high-*k* gate dielectrics^{385–387} to avoid quantum tunnelling³⁸⁸ during further scale reduction process of semiconductors. Kang and Tougaard and their colleagues obtained the band gap of HfZrO₄ gate oxide thin film and inelastic mean free path (IMFP) values through a quantitative analysis of REELS spectra.³⁸³

Investigation of conjugation or aromaticity in organic materials

For materials with conjugated carbon bonding systems, REELS can provide complementary information to that obtained *via* XPS. An example is shown in Fig. 17, involving the organic light emitting diode (OLED) material poly(9,9-di-*n*-octylfluorenyl-2,7-diyl), known as PFO, which was probed using XPS and REELS. The C 1s XPS spectrum shows two very weak peaks at a binding energy position slightly greater than the main peak.

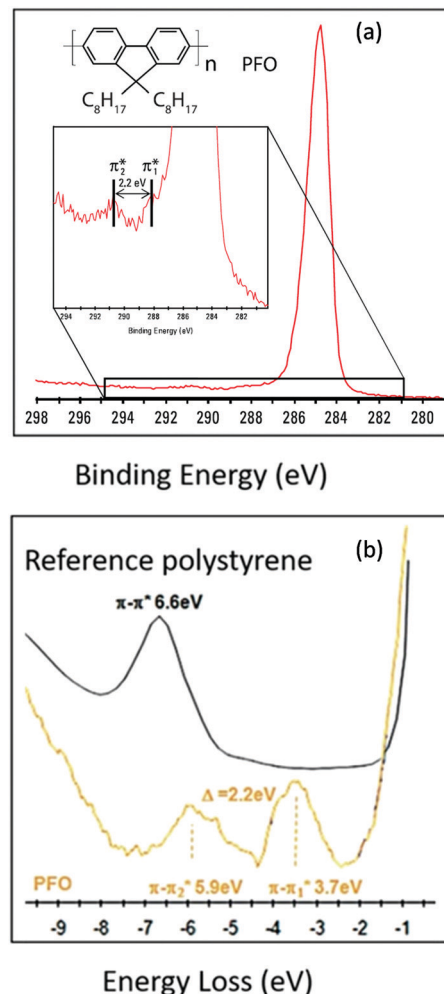


Fig. 17 (a) Photoelectron spectrum of poly (9,9-dioctylfluorene) (PFO), including observation of weak shake-up peaks caused from inelastic collisions between emitted photoelectrons and π/π^* HOMO/LUMO (inset) and (b) REELS profile from PFO in which π/π^* shake-up satellites may be clearly observed and energies recorded. Polystyrene is included as a reference for position of phenyl $\pi-\pi^*$ transition. Reprinted (adapted) with permission from ref. 389. Copyright {2013} Thermo Scientific.

These are assigned as shake-up satellites promoting an electronic transition from HOMO (π) into LUMO (π^*) of PFO. In the REELS spectrum the same two peaks are much more intense and can therefore be more easily identified. Quantification of sp^2/sp^3 hybridisation of carbons may also be achieved by REELS, in cases where it is not always possible *via* other techniques such as XPS/AES.³⁹⁰

Comparison of relative hydrogen content in polymer materials

XPS has great difficulty detecting hydrogen (H 1s) because of its very low photoionisation cross section, its only really feasible with a synchrotron source. REELS however has the capability to detect and quantify hydrogen because of a process that resembles that of the ISS phenomenon. When the REELS electron strikes a hydrogen atom, the atom recoils because it is so much lighter than other elements and results in a discrete energy loss for the electron. This is demonstrated in Fig. 18 in



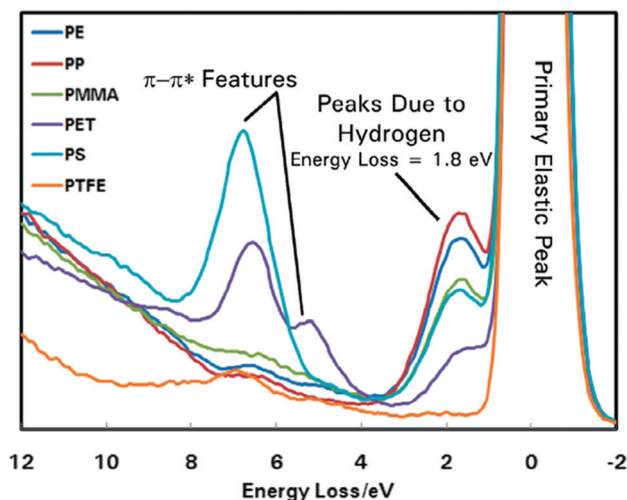


Fig. 18 REELS spectra of polyethylene, polypropylene, poly(methyl methacrylate), polyethylene terephthalate, polystyrene and poly tetrafluoroethylene to evidence the changing peak intensity at ~ 1.8 eV with differing H concentration. Reprinted (adapted) with permission from ref. 391. Copyright (2013) Thermo Scientific.

which REELS with a 1 keV beam energy was used to analyse a series of polymers. Strong features relating to conjugated carbon bonding are readily observed, however another set of peaks at around 1.8 eV from the primary elastic peak are also present and these are assigned to hydrogen, an assignment that is supported by the REELS spectrum for PTFE (which has no hydrogens), where this peak is not visible.

The relative intensity of the hydrogen peak to the primary elastic peak is proportional to the amount of hydrogen relative to other atoms (e.g. C, O and N). A simple peak deconvolution of the two peaks can therefore be used to quantify the hydrogen present.

Summary

In this review, we hope to have emphasised the great importance of utilising a multi-technique approach to materials analysis, in order to extract the maximum amount of information and enable a greater understanding of materials properties at play. Surface analysis of materials is of enormous importance in almost every aspect of science and technology and combinations of techniques can provide much more information than is often apparent through the utilisation of a single technique alone. Whilst XPS is arguably the most widely appreciated technique in this area, we have highlighted some of the many ways in which it may be used in conjunction with a broad range of complementary analyses to elucidate additional sample information, far more than simply the chemical nature of the material surface as is often the extent of the information gleaned in the general analysis of materials. Electronic, 2-dimensional, topological and structural information are readily obtained through the combination of multiple techniques associated with photoelectron spectroscopy. Manufacturers have developed

multi-technique systems for many years now, largely focused on users in the surface science field and relying on knowledge and experience of working with complex UHV chambers and systems. The advent of current gen systems such as the Kratos Analytical Ltd AXIS Supra⁺ and Thermo ScientificTM NEXSATM G2 has signalled a pronounced shift towards the development of multi-functional XPS based instruments aimed at a more general user base. Particularly the introduction of automated sample handling, more robust software and data handling and easy switching between experimental modes represents a great opportunity for such techniques to really enter mainstream use across research laboratories of many fields. Whilst the relatively large investment will restrict the uptake somewhat, the reduced requirement for specialised skills and knowledge will certainly broaden the general userbase and increase the number of researchers looking to utilise XPS and multi-technique surface analysis. The field is even starting to see more niche general use systems, such as the SPECS GmbH EnviroESCATM and Scienta Omicron GmbH HAXPES Lab – capable of user friendly NAP-XPS and HAXPES experiments respectively. While some manufacturers will remain focused on one or two applications, others seem to be moving in the direction of adding more and more functionality to individual instruments. As the demand for such systems grows, it is entirely possible we see an increase in the number of available add-ons to standard commercial instruments, for example Diffuse Reflectance UV-Vis (DRUVS), Attenuated Total Reflectance Infrared (ATR) and even traditionally synchrotron exclusive techniques such as Near-edge X-ray Absorption Fine Structure (NEXAFS)³⁹² as improvements to radiation sources and detectors are developed. The combination of these, in particular with the *in operando* capabilities of systems such as the lab-based NAP-XPS instruments – represents the potential for incredibly powerful analytical investigations to operate within the confines of even a small scale laboratory.

Author contributions

Conceptualisation: MAI. Funding acquisition: PRD and RP. Investigation: JDD, MAI, RL, SG and DJM. Resources: PRD and RP. Writing – original draft: JDD, MAI, RL, SG and DJM. Writing – reviewing and editing: PRD, MAI, DJM, RP.

Conflicts of interest

There are no conflicts to declare.

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