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# Influence of the catalyst surface chemistry on the electrochemical self-coupling of biomass-derived benzaldehyde into hydrobenzoin†

Li Gong,<sup>ab</sup> Shiling Zhao,<sup>id c</sup> Jing Yu,<sup>ad</sup> Junshan Li,<sup>e</sup> Jordi Arbiol,<sup>id dh</sup>  
Tanja Kallio,<sup>id f</sup> Mariano Calcabrini,<sup>id g</sup> Paulina R. Martínez-Alanis,<sup>\*a</sup>  
Maria Ibáñez<sup>id g</sup> and Andreu Cabot<sup>id \*ah</sup>

The electroreduction of biomass-derived benzaldehyde (BZH) provides a potentially cost-effective route to produce benzyl alcohol (BA). This reaction competes with the electrochemical self-coupling of BZH to hydrobenzoin (HDB), which holds significance as a biofuel. Herein, we demonstrate the selectivity towards one or the other product strongly depends on the surface chemistry of the catalyst, specifically on its ability to adsorb hydrogen, as showcased with Cu<sub>2</sub>S electrocatalysts. We particularly analyze the effect of surface ligands, oleylamine (OAm), on the selective conversion of BZH to BA or HDB. The effect of the electrode potential, electrolyte pH, and temperature are studied. Results indicate that bare Cu<sub>2</sub>S exhibits higher selectivity towards BA, while OAm-capped Cu<sub>2</sub>S promotes HDB formation. This difference is explained by the competing adsorption of protons and BZH. During the BZH electrochemical conversion, electrons first transfer to the C in the C=O group to form a ketyl radical. Then the radical either couples with surrounding H<sup>+</sup> to form BA or self-couple to produce HDB, depending on the H<sup>+</sup> availability that is affected by the electrocatalyst surface properties. The presence of OAm inhibits the H adsorption on the electrode surface therefore reducing the formation of high-energy state H<sub>ad</sub> and its combination with ketyl radicals to form BA. Instead, the presence of OAm promotes the outer sphere reaction for obtaining HDB.

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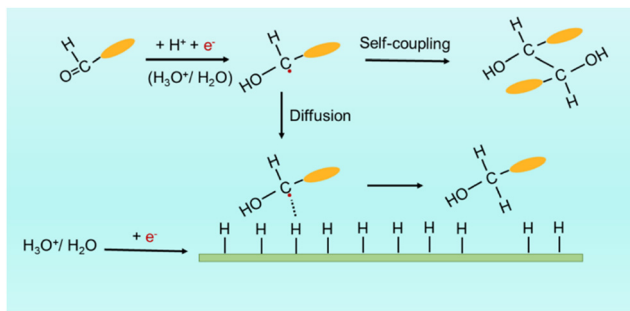
## 1. Introduction

Addressing the environmental crisis requires a transition to clean and zero-carbon energies in both supply and consumption, as well as the adoption of sustainable and efficient chemical synthesis methods.<sup>1,2</sup> Electrocatalysis plays a crucial role by allowing precise control over reaction conditions and promoting the use of mild, sustainable, and energy-efficient processes.<sup>3–10</sup> This approach enables the fine-tuning of reaction

products through adjustable parameters, such as the electrode potential and current density, facilitating the eco-friendly synthesis of complex molecules with unprecedented efficiency.<sup>11–13</sup> In particular, the electrocatalytic reduction of organic compounds is a field of significant interest with broad application potential.

Despite the high economic potential and clean characteristics of electrocatalytic organic reduction processes, their industrial application in the valorization of biomass-derived feedstocks is not widely extended. This is due to their notable complexity and challenging optimization, which arises from the intricate nature of organic compounds and the involvement of multi-step electron transfer mechanisms and various intermediates. As an example in this direction, the electroreduction of biomass-derived benzaldehyde (BZH) provides a fascinating route for the selective production of benzyl alcohol (BA), a versatile compound widely utilized in the pharmaceutical and fine chemical industries.<sup>14–16</sup> Moreover, the electrochemical self-coupling of BZH to hydrobenzoin (HDB) holds significance as a potential biofuel production route.<sup>17,18</sup> Both processes, the production of BA and HDB occurring at around 0 V vs. RHE, compete with the hydrogen evolution reaction (HER, 0 V vs. RHE). As shown in Scheme 1, the first elementary step in both

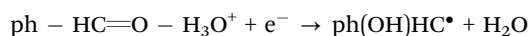
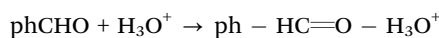
<sup>a</sup> Catalonia Institute for Energy Research – IREC Sant Adrià de Besòs, Barcelona 08930, Spain. E-mail: pmartinez@irec.cat, acabot@irec.cat<sup>b</sup> University of Barcelona, Barcelona 08028, Spain<sup>c</sup> School of Petrochemical Engineering, Lanzhou University of Technology, Lanzhou 730050, China<sup>d</sup> Catalan Institute of Nanoscience and Nanotechnology (ICN2), CSIC and BIST, Campus UAB, Bellaterra, 08193 Barcelona, Spain<sup>e</sup> Institute of Advanced Study, Chengdu University, Chengdu 610106, China<sup>f</sup> Department of Chemistry and Materials Science, Aalto University School of Chemical Engineering, P.O. Box 16100, FI-00076 Aalto, Finland<sup>g</sup> IST Austria, Am Campus 1, Klosterneuburg 3400, Austria<sup>h</sup> ICREA, Pg. Lluís Companys 23, 08010 Barcelona, Spain† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4ya00334a>



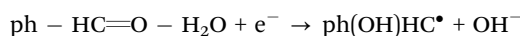
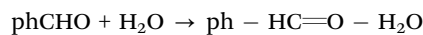
**Scheme 1** Two pathways for the BZH electroreduction, to BA or HDB. The orange ellipsoid represents the benzene ring. Red spots represent electrons provided by the electrode that is represented here by a green bar.

reactions involves the protonation of C=O in BZH, which facilitates the posterior acquisition of one electron from the electrode to form a radical intermediate.<sup>18–20</sup>

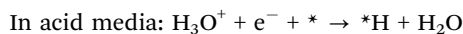
In acid media:



In alkaline media:



After the protonation and electron transfer steps, either the dimerization of two radicals in solution to produce HDB or the combination of the radical with an adsorbed hydrogen ( $\text{H}_{\text{ad}}$ ) in a Langmuir–Hinshelwood (L–H) scheme to obtain BA takes place. While the electroreduction of BZH to HDB is an outer-sphere reaction, *i.e.* does not involve surface adsorption,<sup>21–25</sup> BA production depends on the  $\text{H}_{\text{ad}}$  availability on the catalyst surface:



where \* represents an adsorption site. Both in acid and basic conditions, the selective production of HDB requires reducing the amount of  $\text{H}_{\text{ad}}$  to reduce the formation of BA.

To maximize HDB selectivity, current catalyst design strategies primarily focus on materials with high HER overpotentials such as Hg, Pb, In, and Cd.<sup>22,26–29</sup> However, beyond concerns of metal toxicity and cost, this strategy has intrinsic limitations. Increasing the negative voltage or temperature to accelerate the reaction raises the Fermi level on the catalyst surface, resulting in more adsorbed hydrogen protons concentrating on the solution side of the electrical double layer (EDL). This, in turn, promotes the hydrogenation of ketyl radicals leading to the formation of BA.

Beyond material selection, the rational design and engineering of the catalyst surface chemistry is a powerful strategy to promote selectivity. More specifically, the modification of the electronic energy levels of the catalyst surface using functional groups such as –OH and –SOOH,<sup>30,31</sup> as well as larger organic groups such as aromatic thiols, have been shown to favor specific reaction mechanisms.<sup>32–34</sup> Besides modifying the electronic energy levels, the presence of surface species can also adjust selectivity by tailoring the surface adsorption properties. As an example, lipophilic species can enhance the gathering of organic substances at the catalyst surface. In this direction, Baik's group demonstrated an increase in selectivity by promoting the Suzuki–Miyaura cross-coupling reaction with organic monolayers self-assembled at the electrode surface.<sup>35</sup> Besides, Li's group tailored the interfacial environments of Pb electrodes using quaternary alkyl ammonium cations to control the selectivity of BZH to BA or HDB. They showed that the longer the chain length and the higher the cation concentration, the higher the selectivity towards HDB.<sup>26</sup>

Herein, we modify the surface of  $\text{Cu}_2\text{S}$  nanoparticles (NPs) by adding a monolayer of unsaturated fatty amines, specifically oleylamine (OAm), and explore its effect on the electrocatalytic conversion of BZH to either BA or HDB. Cu was selected taking into account its exceptional ability for radical generation.<sup>27</sup> We investigate the effect of voltage and temperature on the reaction selectivity at pH 5.2 and pH 9.0. We further explore the involved electroreduction reaction pathways aiming to unveil the rate-determining steps and gain insight into the parameters determining the reaction selectivity. Our final goal is to enable a more rational design of electrocatalyst for the selective electroreduction of biomass-derived BZH.

## 2. Experimental

### 2.1 Materials

Copper(II) acetylacetonate ( $\text{Cu}(\text{acac})_2$ , Sigma Aldrich, 97%), oleylamine (OAm, Aladdin Reagent, 70%), dodecanethiol (DDT, Sigma Aldrich, 98%), *N,N*-dimethylformamide (DMF, Sigma Aldrich, anhydrous, 99.8%), methanol ( $\text{CH}_3\text{OH}$ , Letslab, 96%), ethanol ( $\text{CH}_3\text{CH}_2\text{OH}$ , Letslab, 96%), benzaldehyde ( $\text{C}_7\text{H}_6\text{O}$ , Sigma Aldrich), benzyl alcohol ( $\text{C}_7\text{H}_8\text{O}$ , Sigma Aldrich), benzoic acid ( $\text{C}_7\text{H}_6\text{O}_2$ , Sigma Aldrich), and hydrobenzoin ( $\text{C}_{14}\text{H}_{14}\text{O}_2$ , Sigma Aldrich) were used as purchased, without further purification.

### 2.2 $\text{Cu}_2\text{S}$ preparation

$\text{Cu}_2\text{S}$  NPs were produced following a modified version of a previously reported procedure.<sup>35</sup> Briefly, in a three-neck flask equipped with a reflux condenser, a mixture of  $\text{Cu}(\text{acac})_2$  (1.56 g, 0.6 mmol), DDT (20 mL), and OAm (40 mL) was deaerated with nitrogen at 120 °C for 20 min. After the nitrogen purging step, the mixture exhibited a light brownish-yellow color. Then, the solution was heated to 200 °C at a constant heating rate of 8 °C  $\text{min}^{-1}$  and kept at this temperature for 60 min. During this time, the solution becomes turbid, and the color deepens to a brown-black indicating the occurrence of the



chemical reaction. Then, the mixture was cooled down to room temperature. The material was washed with hexane/methanol, separated by centrifugation (9000 rpm, 10 min), and finally dispersed in hexane.

### 2.3 Ligands removal

The ligands used to control the Cu<sub>2</sub>S NP growth were removed using an annealing step.<sup>36</sup> The dried Cu<sub>2</sub>S NP powder was annealed for 2 h under an Ar flow inside a tube furnace at 400 °C, reached using a heating rate of 10 °C min<sup>-1</sup>.

### 2.4 Materials characterization

High-resolution transmission electron microscopy (HRTEM) and high-angle annular dark field (HAADF) images were acquired in a Tecnai F20 microscope operated at 200 kV. Energy dispersive spectroscopy (EDS) spectra were obtained in a double-corrected and monochromated Thermo Fisher Spectra 300 microscope operated at 200 kV. Powder X-ray diffraction (XRD) patterns were collected at 40 kV and 40 mA with Cu K $\alpha$  radiation ( $\lambda = 1.5406 \text{ \AA}$ ). Fourier-transformed infrared (FT-IR) spectra of the unsupported samples were recorded on an Alpha Bruker spectrometer. X-Ray photoelectron spectroscopy (XPS) was conducted on a ThermoFisher (ESCALAB 250Xi) system with Mono Al K $\alpha$  radiation. The vacuum of the analysis chamber was approximately  $2 \times 10^{-9}$  mbar. The energy, voltage, and beam current were 1486.6 eV, 16 kV, and 15 mA, respectively. XPS spectra were corrected to the position of the C 1s peak at 284.6 eV.

### 2.5 Product identification and conversion and selectivity quantification

High-performance liquid chromatography (HPLC) analysis of the organics was carried out on an Agilent 1200 series apparatus at 25 °C. The HPLC was equipped with an ultraviolet-visible detector and a 4.6 mm  $\times$  150 mm Shim-pack GWS 5  $\mu$ m C18 column. The eluting solvents of A: 5 mM ammonium formate aqueous solution; and solution B: acetonitrile was used. The eluent was 60% solution B and 40% solution A for 11 min. The flow rate was 0.5 mL s<sup>-1</sup>. The injection amount was 1  $\mu$ L. Calibration curves of standard chemicals were employed to determine and quantify the products. The total conversion and selectivities were calculated using the following equations:

$$\text{Conv. \%} = \frac{\text{moles of BZH consumed}}{\text{initial moles of BZH}} \times 100\% \quad (1)$$

$$\text{Sel}_{\text{BA}} \% = \frac{\text{ratio of BA}}{\text{conversion ratio amount BZH}} \times 100\% \quad (2)$$

$$\text{Sel}_{\text{HDB}} \% = \frac{2 \times \text{ratio of HDB}}{\text{conversion ratio amount of BZH}} \times 100\% \quad (3)$$

### 2.6 Electrocatalytic tests

The electrochemical performance of the electrocatalyst was evaluated on a Corrtest CS2350 EIS bipotentiostat (2-channel, Wuhan Corrtest Instrument Corp. Ltd) using a 3-electrode

system in an H-cell reactor divided by a Nafion 117 proton exchange membrane. The working electrode was prepared as follows: 9 mg powder was dispersed into 500  $\mu$ L H<sub>2</sub>O, 420  $\mu$ L methanol, and 80  $\mu$ L 5% Nafion solution. For activity, conversion, and selectivity determination, the catalyst was supported on a carbon cloth used as working electrodes. 1000  $\mu$ L ink solution was dropped on  $2 \times 3 \text{ cm}^2$  carbon cloth ( $\sim 1.5 \text{ mg cm}^{-2}$ ) that was then cut into  $1.5 \times 1 \text{ cm}^2$  pieces. When testing, an electrode area of  $1 \times 1 \text{ cm}^2$  was immersed into the solution. To determine activation energies, 6  $\mu$ L ink solution was dropped on a glassy carbon working electrode ( $d = 0.15 \text{ cm}$ ,  $\sim 0.7 \text{ mg cm}^{-2}$ ) and let dry. Platinum mesh and Ag/AgCl (KCl saturated) were used as counter and reference electrodes, respectively. Cyclic voltammetry (CV) and linear sweep voltammetry (LSV) results were calibrated with Nernst eqn (4) to the reference hydrogen electrode (RHE) without iR correction. The temperature effects were neglected when setting the parameters for  $I$ - $t$  curves.

$$E_{\text{RHE}} = E + 0.0591 \times \text{pH} + E_{\text{Ag/AgCl}}^{\theta} \quad (4)$$

where  $T$  represents temperature in °C, and

$$E_{\text{Ag/AgCl(KCl saturated)}}^{\theta} = [0.197 - 0.0007 (T - 25)] \text{ V} \quad (5)$$

Two electrolytes were tested to explore the influence of the pH. A sodium acetate-acetic acid solution (pH = 5.2), and a potassium carbonate-potassium bicarbonate solution (pH = 9.0). LSV was tested in the range 0.5 V to -1.0 V (vs. RHE). Electrochemical impedance spectroscopy (EIS) was tested at -0.7 V vs. RHE. For the  $I$ - $t$  curve, -0.4 V, -0.8 V, -1.0 V (vs. RHE) was applied as a constant voltage.

## 3. Results and discussion

Cu<sub>2</sub>S and Cu<sub>2</sub>S-OAm electrocatalysts were prepared at 200 °C using a one-pot synthetic reaction followed by post-treatment (see details in the Experimental section, Fig. 1a). DDT was employed as the source of S and solvent and OAm was used as an additional solvent and surface ligand. After 1 h reaction at 200 °C, Cu<sub>2</sub>S-OAm was obtained by centrifugation. Cu<sub>2</sub>S was obtained by annealing the Cu<sub>2</sub>S-OAm powder at 400 °C under an Ar atmosphere.

Fig. 1b shows the Cu<sub>2</sub>S-OAm NPs to have spherical geometry and an average size of  $10 \pm 2 \text{ nm}$ . HRTEM micrographs of Cu<sub>2</sub>S-OAm reveal their hexagonal crystal phase, space group  $P6_3mmC$ , with  $a = b = 3.95 \text{ \AA}$  and  $c = 6.75 \text{ \AA}$  (Fig. 1c), which was consistent with XRD analysis results (PDF#00-02-1294, Fig. 1d). HAADF intensity profiles and STEM-EDS chemical composition maps show a homogeneous dispersion of Cu and S within the Cu<sub>2</sub>S-OAm particles that are surrounded by a carbon-based shell (Fig. 1e and Fig. S1, ESI<sup>†</sup>). The carbon shell is discerned by the stronger carbon signal detected at the border of the particle and it is associated with the presence of OAm at the NP surface.

To remove OAm ligands from the NP surface, the material was annealed at 400 °C under an Ar atmosphere. The removal of the stabilizing ligands resulted in increased surface energy that, combined with the relatively high processing temperature,





**Fig. 1** (a) Scheme of the procedure used to prepare  $\text{Cu}_2\text{S-OAm}$  and  $\text{Cu}_2\text{S}$ . (b) TEM and (c) HRTEM images and indexed power spectrum of  $\text{Cu}_2\text{S-OAm}$  NPs. From the crystalline domain, the  $\text{Cu}_2\text{S}$  lattice fringe distances were measured to be 0.1902 nm, 0.3057 nm, at  $80.46^\circ$  which is interpreted as the hexagonal  $\text{Cu}_2\text{S}$  phase, visualized along its [010] zone axis. (d) STEM-EDS chemical composition maps of  $\text{Cu}_2\text{S-OAm}$  NPs; (e) XRD patterns of  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S-OAm}$ . (f) FT-IR spectra of  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S-OAm}$ .

led to NP aggregation and additional growth into  $\text{Cu}_2\text{S}$  particles with irregular morphologies (Fig. S2a, ESI<sup>†</sup>). Yet, the hexagonal crystal phase of  $\text{Cu}_2\text{S}$  was maintained as shown by XRD characterization (Fig. 1d). EDS analysis confirmed the  $\text{Cu}_2\text{S}$  composition to be close to the stoichiometric Cu:S = 2:1 (Fig. S2e, ESI<sup>†</sup>). Besides, HAADF intensity profiles and STEM-EDS elemental maps showed no evidence of an organic layer surrounding the  $\text{Cu}_2\text{S}$  particles, confirming the surface ligand removal (Fig. S2, ESI<sup>†</sup>). The complete ligand removal was further verified using FT-IR analysis (Fig. 1f). The FT-IR spectrum of  $\text{Cu}_2\text{S-OAm}$  displays the signal corresponding to the  $-\text{NH}$  bending vibration at  $1619\text{ cm}^{-1}$  and  $-\text{NH}_2$  wagging vibration at  $721\text{ cm}^{-1}$ , which indicates the presence of OAm molecules on the  $\text{Cu}_2\text{S}$  surface bond to the surface Cu by the amine polar groups to form OAm/Cu complex. Besides, the FT-IR spectrum of  $\text{Cu}_2\text{S-OAm}$  does not show the band at  $2577\text{ cm}^{-1}$  associated with the stretching vibration of S-H. This result indicates the lack of free alkylthiol and Cu thiolate complexes, *i.e.* the absence of DDT on the surface of the  $\text{Cu}_2\text{S-OAm}$  NPs.<sup>35,37</sup> Peak signals at around  $3000\text{ cm}^{-1}$  corresponds to the stretching vibration of both saturated and unsaturated C-H. On the other hand, the FT-IR spectrum of  $\text{Cu}_2\text{S}$  displayed no fingerprint associated with the presence of OAm or DDT, proving the successful ligand removal during the annealing process.

The survey and high-resolution XPS spectra are displayed in Fig. S3 (ESI<sup>†</sup>) and Fig. 2, respectively. The Cu 2p XPS spectrum of  $\text{Cu}_2\text{S-OAm}$  displays a unique doublet, at  $932.1\text{ eV}$  ( $\text{Cu } 2p_{3/2}$ ), associated with Cu(I) within the  $\text{Cu}_2\text{S}$  lattice.<sup>35</sup> In contrast, the Cu 2p XPS spectrum of  $\text{Cu}_2\text{S}$  displays two doublets and

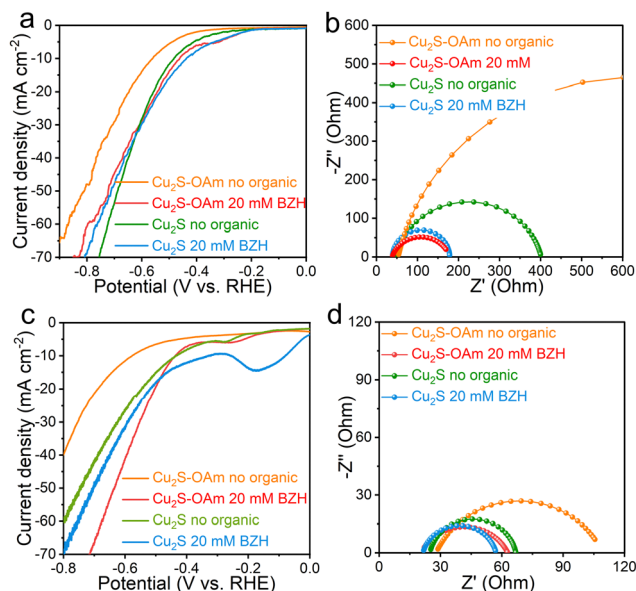


**Fig. 2** High-resolution Cu 2p (a) and S 2p (b) XPS spectra of  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S-OAm}$ .

additional satellite peaks. The first doublet, at  $932.1\text{ eV}$  ( $\text{Cu } 2p_{3/2}$ ), is associated with Cu(I) within the  $\text{Cu}_2\text{S}$  lattice. The second doublet, at  $934.3\text{ eV}$  ( $\text{Cu } 2p_{3/2}$ ), and the satellites are related to the presence of a significant amount of Cu(II) at the  $\text{Cu}_2\text{S}$  surface. This second component is associated with the oxidation of Cu(I) after ligand removal. The S 2p XPS spectrum of  $\text{Cu}_2\text{S-OAm}$  shows one doublet at  $161.5\text{ eV}$  ( $\text{S } 2p_{3/2}$ ) which is assigned to S ions coordinated to Cu within the  $\text{Cu}_2\text{S}$  lattice. A second minor component is observed at lower binding energies,  $160.9\text{ eV}$  ( $\text{S } 2p_{3/2}$ ), which we associate with sulfur next to OAm-bond metal ions. After the ligand removal, the S 2p XPS spectrum of  $\text{Cu}_2\text{S}$  shows just one doublet associated with S ions within the  $\text{Cu}_2\text{S}$  lattice. This component is redshifted by around  $0.4\text{ eV}$ , which indicates a higher electron concentration around the surface S within the  $\text{Cu}_2\text{S}$  sample.

The electrochemical properties of  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S-OAm}$  particles supported on glassy carbon were initially tested in acidic conditions, within a 1 M sodium acetate-acetic acid electrolyte ( $\text{pH} = 5.2$ ) with and without 20 mM BZH. Lower pH values were not considered as pH values below 5.2 might cause C=O cleavage of BZH to toluene or other side products. Fig. 3a displays the LSV curves obtained in a three-electrodes H-cell reactor. In acidic conditions, without BZH, electrons are only transferred to  $\text{H}^+$  to form  $\text{H}_{\text{ad}}$ , thus the measured current densities in the absence of BZH correspond to the HER. We observe that the  $\text{Cu}_2\text{S-OAm}$  electrode displays a significantly lower HER activity, with a higher overpotential and lower current densities at all applied voltages than the  $\text{Cu}_2\text{S}$  electrode. This result is consistent with the presence of OAm partially inhibiting  $\text{H}_{\text{ad}}$  formation and HER. In the presence of BZH, electrons can be also transferred to C in  $\text{ph-HC=O-H}_3\text{O}^+$  to form a ketyl radical. The subsequent step, after the electron transfer process, involves the combination of the ketyl radical and  $\text{H}_{\text{ad}}$  into BA or the combination of ketyl radicals into BZH, involving no additional charge transfer. For both electrodes, the addition of BZH in the electrolyte results in a current increase at low negative potentials, below the HER onset potential. This implies that ketyl radicals are readily formed at lower negative voltages than  $\text{H}_{\text{ad}}$  and  $\text{H}_2$ , at *ca.*  $-0.3\text{ V vs. RHE}$  in both electrodes. For the  $\text{Cu}_2\text{S-OAm}$  electrode, the current density in the presence of BZH is systematically higher than in its absence, at all potentials tested, which reflects a promoted ketyl radical generation over the partially inhibited  $\text{H}_{\text{ad}}$  formation and  $\text{H}_2$  evolution. As for the





**Fig. 3** (a) LSV curves and (b) Nyquist plots at  $-0.7$  V vs. RHE of  $\text{Cu}_2\text{S-OAm}$  and  $\text{Cu}_2\text{S}$  in 1 M sodium acetate-acetic acid electrolyte ( $\text{pH} = 5.2$ ) with and without BZH at room temperature; (c) LSV curves and (d) Nyquist plot at  $-0.7$  V vs. RHE of  $\text{Cu}_2\text{S-OAm}$  and  $\text{Cu}_2\text{S}$  in 1 M potassium carbonate-potassium bicarbonate electrolyte ( $\text{pH} = 9$ ) with or without BZH at room temperature.

$\text{Cu}_2\text{S}$  electrode, while at low negative potentials, the current density is higher in the presence of BZH, as the applied potential increases, the HER-related current in the absence of BZH overcomes the combination of the current density associated with the BZH reaction and the HER in the presence of BZH. Thus the presence of BZH partially inhibits the HER in this material, potentially as a result of the ketyl radical reaction with  $\text{H}_{\text{ad}}$ , reducing its concentration, or through the adsorption of  $\text{ph-HC=O-H}_3\text{O}^+$  or the ketyl radical on the  $\text{Cu}_2\text{S}$  surface thus occupying  $\text{H}_{\text{ad}}$  adsorption sites.

LSV curves depend on charge transfer frequency. They are influenced by mass transport in the electrode and solution, the intrinsic catalytic properties and the number of available adsorption and reaction sites. Thus processes that do not involve charge transfer but potentially blocking/releasing reaction sites such as the ketyl radical conversion to BA through reaction with a  $\text{H}_{\text{ad}}$  play also an important role. EIS spectra were used for a more precise characterization of the charge transfer mechanisms, without the involvement of site density and occupancy.

EIS spectra of the two electrodes at  $-0.7$  V vs. RHE in the presence and absence of BZH show one semicircle associated with the impedance of the charge transfer from the electrode surface to the electrolyte species ( $R_{\text{ct}}$ ). The  $R_{\text{ct}}$  of  $\text{Cu}_2\text{S-OAm}$  in the absence of BZH is significantly larger than that of  $\text{Cu}_2\text{S}$ , which is consistent with the inhibition of the  $\text{H}^+$  adsorption and  $\text{H}_2$  evolution being related not to a lower site density but to a larger barrier for electron transfer to  $\text{H}^+$  in solution probably related to the long OAm chains. We observe the series resistance ( $R_s$ ) to slightly decrease in the presence of BZH, which is

related to an increase of the electrolyte conductivity in the presence of the generated  $\text{ph-HC=O-H}_3\text{O}^+$  and ketyl radicals. Besides, both,  $\text{Cu}_2\text{S}$  and especially  $\text{Cu}_2\text{S-OAm}$  electrodes show much smaller  $R_{\text{ct}}$  values in the presence of BZH. This is consistent with LSV data for  $\text{Cu}_2\text{S-OAm}$  but it is in contrast with LSV data for  $\text{Cu}_2\text{S}$ . The smaller  $R_{\text{ct}}$  is associated with the easier electron transfer from the electrode to BZH and formation of ketyl radical than to  $\text{H}^+$  to form  $\text{H}_{\text{ad}}$ . This result, combined with LSV data, shows the slower kinetics detected in the presence of BZH to be related to the BZH radical conversion to BA instead of the ketyl radical formation. The charge transfer resistance of  $\text{Cu}_2\text{S-OAm}$  in the presence of BZH is slightly smaller than that of  $\text{Cu}_2\text{S}$ , which is surprising taking into account the presence of OAm hindering charge transfer to species in solution. We associate these results with the hydrophobic surface of  $\text{Cu}_2\text{S-OAm}$  that leads to an enrichment of BZH, facilitating electron transfer to  $\text{ph-HC=O-H}_3\text{O}^+$ . Also, the possibility of partial electron transfer to OAm in  $\text{Cu}_2\text{S-OAm}$  cannot be ruled out.

In an acidic solution, under an applied voltage, the OAm of the  $\text{Cu}_2\text{S-OAm}$  surface may accept electrons and could be also replaced by  $\text{H}^+$ .<sup>38</sup> Therefore, the stability of the  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S-OAm}$  electrode in 1 M sodium acetate-acetic acid electrolyte ( $\text{pH} = 5.2$ ) was further analyzed (Fig. S4, ESI<sup>†</sup>). The repetitive CV curves showed  $\text{Cu}_2\text{S}$  to have good stability both with and without BZH. In contrast, the current density obtained from the  $\text{Cu}_2\text{S-OAm}$  electrode was not constant. It increased in an electrolyte containing 20 mM BZH and decreased in the pure electrolyte. These results indicate that the OAm ligand undergoes slight reaction/substitution. Still, after more than 20 cycles, the effect of OAm on the  $\text{Cu}_2\text{S-OAm}$  electrode is clearly noticed (Fig. S5, ESI<sup>†</sup>).

Results obtained in alkaline conditions, using glassy carbon with 1 M potassium carbonate-potassium bicarbonate electrolyte ( $\text{pH} = 9.0$ ), are displayed in Fig. 3d and f. In these conditions, due to the significant suppression of HER, the concentration of ketyl radicals is expected to increase, resulting in an enhancement of their self-coupling. Indeed, in the absence of BZH, a clear decrease in the HER-related current densities is observed for both electrode materials. LSV curves show the current densities to significantly increase when adding BZH into the system for  $\text{Cu}_2\text{S-OAm}$  (Fig. 3d), displaying a similar onset potential as in acidic conditions, but providing overall lower current densities at higher negative voltages due to the lower HER contribution. For  $\text{Cu}_2\text{S}$ , current densities show a much slighter variation in the presence of BZH, as in the acidic case.

EIS spectra of the two electrodes in basic conditions at  $-0.7$  V vs. RHE are displayed in Fig. 3f. A notable decrease of  $R_{\text{ct}}$  is noticed in alkaline conditions compared with acidic ones in the absence of BZH, associated with the enhancement of radical-radical coupling. At low pH, the hydrogen binding strength is lower due to the surrounding hydronium ions,<sup>39</sup> which facilitates the formation of  $\text{H}_{\text{ad}}$  that combines with the ketyl radicals to produce benzyl alcohol. As the pH increases, the formation of  $\text{H}_{\text{ad}}$  requires overcoming a higher energy barrier. Thus radicals



are more inclined to undergo self-coupling, which further accelerates the charge transfer and radical formation. With the introduction of BZH, both electrodes display similar tendencies as in acidic conditions, with a notable decrease of the  $R_{ct}$  attributed to the formation of ketyl radicals.

As observed from the CV curves (Fig. S6, ESI<sup>†</sup>), the performance of the  $\text{Cu}_2\text{S}$  electrode was stable both with and without BZH (Fig. S6a and b, ESI<sup>†</sup>). In contrast, the current density increased for  $\text{Cu}_2\text{S}$ -OAm electrodes in the pure electrolyte, indicating moderate stability in these conditions, potentially due to ligand displacement with  $-\text{OH}$ .<sup>38</sup> This instability was not observed in the presence of BZH (Fig. S7, ESI<sup>†</sup>), which indicates a smooth and efficient transfer of charge to BZH.

To quantify the reaction products, additional electrocatalytic measurements were carried out with the materials supported on CC using a constant reaction time of 20 min. Different pH, voltage, and temperature conditions were tested to evaluate the effect of these parameters on the BZH conversion and product selectivity.

Fig. 4 shows the effect of the applied potential on the BZH conversion and selectivity. Under acidic conditions (Fig. 4a), at a low voltage of  $-0.4$  V vs. RHE, both  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S}$ -OAm catalysts exhibit relatively low BZH conversion ( $\sim 5\%$ ), and notable selectivity towards BA, especially  $\text{Cu}_2\text{S}$ . This result is attributed to the hydrogen bonding to the catalyst surface, forming Cu-H bonds at low negative voltages. Then, this H is combined with ketyl radical.<sup>40</sup> As the negative voltage increases, the conversion increases. Additionally, as increasing the voltage, the selectivity of the two electrodes towards the conversion of BZH into BA first decreases and then increases again. The voltage increase elevates the electrochemical potential at the electrode-solution interface, augmenting the formation of  $\text{H}_{ad}$ , thereby bolstering BA selectivity. As expected,  $\text{Cu}_2\text{S}$ -OAm exhibited higher selectivity to HDB at all voltages due to the inhibited formation of  $\text{H}_{ad}$ . The slight decrease in HDB selectivity obtained for  $\text{Cu}_2\text{S}$ -OAm at high voltage may be related to the partial OAm reaction/replacement from the  $\text{Cu}_2\text{S}$  surface.

Under basic conditions (Fig. 4b), a substantial increase in BZH conversion is measured, but the selectivity trends with the applied voltage remain unchanged. Notably, the selectivity of

$\text{Cu}_2\text{S}$ -OAm towards the conversion of BZH into HDB increases with the voltage, which is consistent with good catalyst stability.

The reactant concentration effect was also analyzed. The results are summarized in Table S1 (ESI<sup>†</sup>). It can be observed that the BZH conversion is slightly increased with higher reactant concentration within the same reaction time on different pH and the same electrode materials, while the selectivity towards BA or HDB shows no significant change on the same electrode materials.

A temperature increase also results in an elevation of the material surface Fermi level thus leading to a decrease of the charge transfer resistance and the formation of additional  $\text{H}_{ad}$ . Therefore, a similar dependence to that obtained for the voltage is expected. The difference lies in that the temperature also accelerates the mass transfer efficiency in the solution. At  $-0.8$  V vs. RHE, the  $\text{Cu}_2\text{S}$  selectivity towards BA increased with temperature both in acid and alkaline conditions owing to the higher  $\text{H}_{ad}$  formed. The  $\text{Cu}_2\text{S}$  selectivity towards BA is higher under acidic conditions compared to alkaline in the whole temperature range tested. For  $\text{Cu}_2\text{S}$ -OAm, the selectivity towards HDB decreased with temperature in acidic conditions, owing to the  $\text{H}_{ad}$  increase. However, the HDB selectivity increased with temperature under alkaline conditions (Fig. 5).

To further understand the effect of temperature on mass transfer and the acceleration of the reaction rate, EIS measurements at different temperatures were carried out on the catalysts supported on glassy carbon. From these results, the corresponding activation energy barriers were calculated (Fig. 6 and Fig. S8, ESI<sup>†</sup>). As pH increased, the reaction energy barrier in  $\text{Cu}_2\text{S}$  decreased from  $70$  kJ mol<sup>-1</sup> to  $44$  kJ mol<sup>-1</sup>. In contrast, the energy barrier in  $\text{Cu}_2\text{S}$ -OAm increased from  $46$  kJ mol<sup>-1</sup> to  $61$  kJ mol<sup>-1</sup>. These results imply the higher sterical barriers for the formation of  $\text{H}_{ad}$  to combine with the hydroxyl carbon of the intermediate ketyl radical in alkaline than in acidic conditions, while BZH can combine with  $\text{H}^+$  in acidic conditions to form ketyl radical more easily than in alkaline.

Based on the above results, it is clear that a high selectivity for obtaining HDB by BZH electrochemical reduction can be achieved by adjusting the pH, applied voltages, catalyst

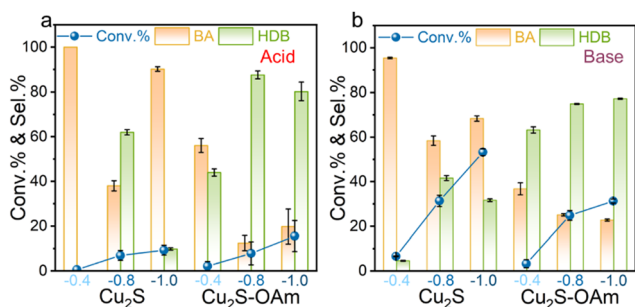


Fig. 4 BZH conversion and selectivity toward BA and HDB after 20 min reaction on  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S}$ -OAm electrodes at 0.4,  $-0.8$ ,  $-1.0$  V vs. RHE,  $25$  °C in (a) 1 M sodium acetate-acetic acid solution ( $\text{pH} = 5.2$ ) with 20 mM BZH; (b) 1 M potassium carbonate-potassium bicarbonate solution ( $\text{pH} = 9.0$ ) with 20 mM BZH.

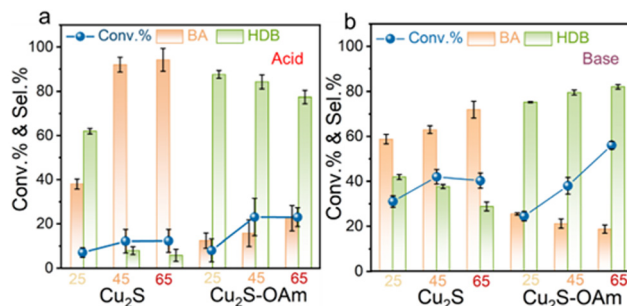


Fig. 5 BZH conversion and selectivity toward BA and HDB after 20 min reaction on  $\text{Cu}_2\text{S}$  and  $\text{Cu}_2\text{S}$ -OAm electrodes at  $-0.8$  V vs. RHE,  $25$  °C,  $45$  °C or  $65$  °C in (a) 1 M sodium acetate-acetic acid solution ( $\text{pH} = 5.2$ ) with 20 mM BZH; (b) 1 M potassium carbonate-potassium bicarbonate solution ( $\text{pH} = 9.0$ ) with 20 mM BZH.



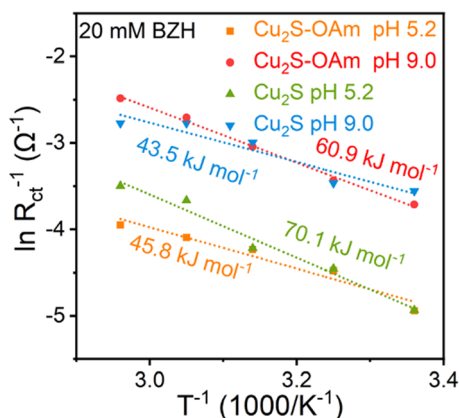


Fig. 6 Activation energies calculated from the Nyquist plot.

components and catalyst surface properties. The energy barrier for the formation of the high-energy intermediate species  $H_{ad}$  increases as the pH increases, leading to a decrease in the concentration of  $H_{ad}$  on the catalyst surface, thereby increasing the chances of radical dimerization for the formation of HDB. For catalysts with exposed active sites, applying a higher voltage leads to surface electron accumulation, causing enrichment of  $H^+$  on the EDL solution side, and increasing the possibility of forming  $H_{ad}$  and BA. Thus, blocking the active sites of the catalyst surface with organic ligands to prevent the formation of  $H_{ad}$  can improve the selectivity of HDB independently of pH.

## 4. Conclusions

In summary,  $Cu_2S$ -OAm and  $Cu_2S$  were synthesized using a one-pot synthetic approach and a post-treatment process. By comparing their electrocatalytic performance in the electrochemical reduction of BZH, we conclude that OAm prevents the organic substrate and the H proton from reaching the  $Cu_2S$  active sites. Consequently, hydrogen cannot become activated to combine with ketyl radicals to form BA. At the same time, the organic BZH and its ketyl radicals concentrate near the OAm ligands on the catalyst surface due to their similar polarity, thus promoting selectivity and conversion towards forming HDB.

## Author contributions

L. Gong: conceptualization and writing – original draft; Shiling Zhao and Tanja Kallio: investigation; Jing Yu, Jordi Arbiol and Junshan Li: formal analysis; Paulina R. Martínez-Alanis: supervision and funding acquisition and writing – review & editing; Mariano Calcabrini and Maria Ibáñez: methodology; Andreu Cabot: supervision and writing – review & editing.

## Data availability

The data supporting this article have been included as part of the ESI.†

## Conflicts of interest

There are no conflicts to declare.

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