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# Atomic-level interface engineering enables efficient and durable acidic hydrogen evolution of osmium at large current densities

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Osmium (Os), the least expensive member of the platinum-group metals, has emerged as a promising alternative to Pt-based catalysts for the hydrogen evolution reaction (HER). However, Os-based electrocatalysts still suffer from poor stability under acidic conditions, despite recent efforts to mitigate H<sup>+</sup> over-adsorption for improved intrinsic activity. Here, we design a porous CeO<sub>2</sub> support that enables the atomic dispersion of Os, forming an Os single-atom catalyst (Os<sub>SA</sub>-CeO<sub>2</sub>). Unlike traditional flat-film supports, the porous CeO<sub>2</sub> architecture prevents Os aggregation and achieves 100% interfacial anchoring of Os atoms. The resulting strong electronic coupling enables tight anchoring of Os and activates the CeO<sub>2</sub> matrix with abundant oxygen vacancies, which facilitate H<sub>2</sub>O dissociation to sustainably supply protons for rapid consumption at large current densities. Also, the generated OH<sup>\*</sup> species are adsorbed by the oxygen vacancies, thus preventing the Os sites from oxidative dissolution. As a result, Os<sub>SA</sub>-CeO<sub>2</sub> exhibits over 500 h of durability at 100 mA cm<sup>-2</sup> without performance decay—surpassing all previously reported Os-based HER catalysts. This work provides a general strategy for achieving complete interfacial anchoring of active metal atoms to enhance catalytic stability without sacrificing activity through support activation.

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## Introduction

The transition to sustainable and non-polluting energy sources has become increasingly urgent amid the global energy, environmental, and geopolitical crises. Among various clean energy technologies, “green hydrogen” has attracted particular attention as a zero-carbon fuel to play a dominant role in the future hydrogen economy.<sup>1–4</sup> Electrocatalytic water splitting powered by renewable energy sources (*e.g.*, solar and wind) represents one of the most promising routes for producing high-purity green hydrogen.<sup>5–7</sup> As the cost of renewable electricity continues to decline, the price, activity, and durability of electrocatalysts for the hydrogen evolution reaction (HER) are becoming the primary factors determining the overall economic feasibility of this process.<sup>6,8–10</sup>

Platinum (Pt)-based catalysts exhibit the highest HER activity due to their optimal hydrogen adsorption free energy. However, the high cost and scarcity of Pt significantly hinder their large-scale application.<sup>6,11–13</sup> Osmium (Os), the least expensive member of the platinum group metals (PGMs), has recently emerged as a potential alternative to Pt-based catalysts.<sup>14–17</sup> A fundamental challenge for Os-based HER catalysts is their tendency to over-adsorb hydrogen intermediates (H<sup>\*</sup>), which results in inherently low HER activity.<sup>18,19</sup> To address this, several

strategies have been developed to tune the electronic structure of Os and modulate its hydrogen adsorption behavior, such as supporting Os nanoparticles on TiO<sub>2</sub>,<sup>20</sup> constructing heterostructures like Os–OsSe<sub>2</sub>,<sup>15,18,19</sup> and introducing anion doping.<sup>21</sup> While these approaches have improved catalytic performance, achieving long-term stability, particularly at high current densities ( $\geq 100$  mA cm<sup>-2</sup>), remains a major challenge.<sup>15,18,19</sup>

The instability of Os-based catalysts primarily originates from nanoparticle aggregation *via* Ostwald ripening and the easily fluctuating valence states of Os under the harsh conditions of the HER.<sup>14,22</sup> At large current densities, oxidative dissolution of Os becomes more pronounced due to the rapid consumption of local H<sup>+</sup> and the accumulation of newly generated OH<sup>-</sup> species derived from H<sub>2</sub>O dissociation.<sup>23</sup> Anchoring Os onto reducible metal oxides, such as CeO<sub>2</sub>, offers a promising pathway to enhance stability. The facile Ce<sup>4+</sup>/Ce<sup>3+</sup> redox cycling and high-order f orbitals of CeO<sub>2</sub> facilitate strong electronic coupling with supported metal atoms.<sup>24–27</sup> Moreover, the oxyphilic nature of Ce and the ease of oxygen vacancy formation endow CeO<sub>2</sub> with a strong affinity for OH<sup>-</sup> species,<sup>23,28–30</sup> which can effectively suppress the adsorption of OH<sup>-</sup> on Os and thus prevent its oxidative dissolution. To achieve these benefits, each Os atom must directly interact with neighboring Ce atoms *i.g.* achieving atomic dispersion of Os on CeO<sub>2</sub>.<sup>31–33</sup> However, the fabrication of this precise architectural structure presents a great challenge and it requires the investigation of the thermodynamic stability properties of isolated Os atoms.

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In this work, we first investigated the structural stability and hydrogen adsorption behavior of atomically dispersed Os on CeO<sub>2</sub> (Os<sub>SA</sub>-CeO<sub>2</sub>) using density functional theory (DFT) calculations. The results reveal that Os atoms can be strongly anchored within the CeO<sub>2</sub> lattice, while their excessive H\* adsorption is effectively mitigated. Guided by these insights, we synthesized Os<sub>SA</sub>-CeO<sub>2</sub> *via* a two-step electrodeposition method. A CeO<sub>2</sub> thin film was first deposited onto a conductive carbon fiber substrate, followed by O<sub>2</sub> annealing to create a porous structure that promotes atomic Os dispersion. This architecture introduces abundant oxygen vacancies and induces lattice compression, which shortens the distance between Os atoms and adjacent vacancies. These features facilitate H<sub>2</sub>O dissociation and the subsequent migration of H\* species to neighboring Os sites, providing a continuous hydrogen source for proton consumption at large current densities. Meanwhile, the generated OH\* intermediates are preferentially adsorbed by the oxygen vacancies rather than the Os sites, effectively suppressing oxidative dissolution. Consequently, Os<sub>SA</sub>-CeO<sub>2</sub> exhibits a remarkably low overpotential of 97 mV and over 500 h of durability at 100 mA cm<sup>-2</sup> with negligible degradation, surpassing nearly all previously reported Os-based HER catalysts.

## Results and discussion

### DFT calculations for material design

To assess the feasibility of Os incorporation into CeO<sub>2</sub>, we calculated the formation energies (FEs) of Os substituting Ce

(Os<sub>Ce</sub>) and Os<sub>Ce</sub> accompanied by an oxygen vacancy (Os<sub>Ce</sub> + V<sub>O</sub>). The FE of Os<sub>Ce</sub> on the CeO<sub>2</sub> (111) surface is 0.64 eV (Fig. S1), whereas that of Os<sub>Ce</sub> + V<sub>O</sub> is -0.81 eV (Fig. 1a). The significantly lower FEs of Os<sub>Ce</sub> + V<sub>O</sub> compared to Os<sub>Ce</sub> indicate that oxygen vacancies form readily in Os-doped CeO<sub>2</sub>. Compared to other investigated platinum-group metals, Os<sub>Ce</sub> + V<sub>O</sub> exhibits the lowest FE, demonstrating the high structural stability of atomically dispersed Os in CeO<sub>2</sub> (Os<sub>SA</sub>-CeO<sub>2</sub>).

The strong anchoring of Os atoms originates from pronounced interfacial electronic coupling with the CeO<sub>2</sub> lattice, as revealed by DFT calculations. For metallic Os nanoparticles (Os<sub>NP</sub>), the projected density of states (PDOS) displays delocalized Os 5d orbitals with occupied states at the Fermi level (Fig. 1b, left), showing metallic behavior. Upon Os doping into CeO<sub>2</sub>, the 5d orbitals become localized, and unoccupied states above the Fermi level nearly vanish (Fig. 1b, right). This reflects charge redistribution and strong hybridization between Os and CeO<sub>2</sub>. The modified electronic structure of Os in Os<sub>SA</sub>-CeO<sub>2</sub> brings in different hydrogen adsorption behavior (Fig. 1c). The energy range of its hybridization with the H 1s orbital is approximately -6.3 to -2.3 eV for Os<sub>SA</sub>-CeO<sub>2</sub> compared to -8 to -4 eV for Os<sub>NP</sub>, indicating the weaker interaction between Os and H in Os<sub>SA</sub>-CeO<sub>2</sub>. The Gibbs free energy diagram for the HER (Fig. 1d) further supports this result. The calculated  $\Delta G_{H^*}$  values for the Ce site in Os<sub>SA</sub>-CeO<sub>2</sub> and Os site in Os<sub>NP</sub> are -0.23 eV and -0.19 eV, respectively, implying facile H adsorption. However, this hinders the desorption of H\*, thereby limiting the formation of H<sub>2</sub>. In contrast, the Os site in Os<sub>SA</sub>-CeO<sub>2</sub> exhibits a moderate  $\Delta G_{H^*}$  of -0.08 eV and a shortened H-H

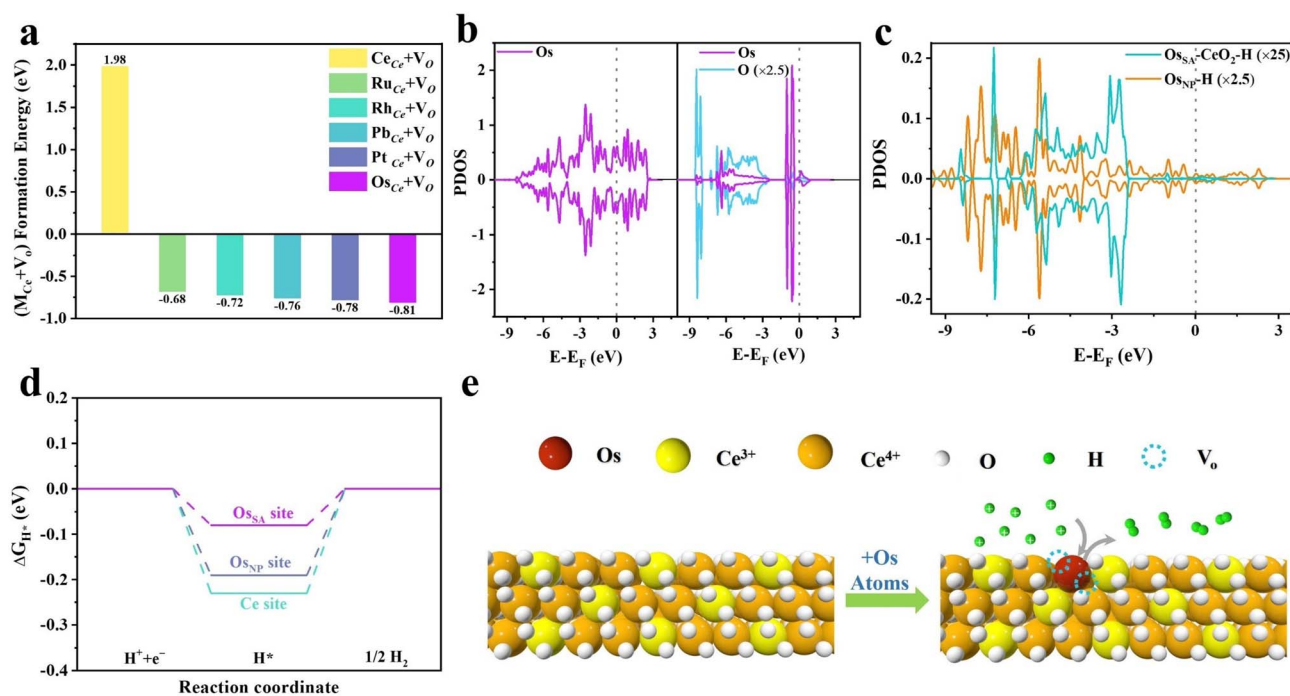


Fig. 1 Theoretical calculations illustrating the strong interfacial coupling of Os SAs in CeO<sub>2</sub>. (a) Calculated formation energies of M<sub>Ce</sub> + V<sub>O</sub> (M stands for the substituted atom for Ce). PDOS for (b) Os 5d orbital of Os<sub>NP</sub> (left) and Os<sub>SA</sub>-CeO<sub>2</sub> (right) and (c) H 1s orbital of adsorbed H. (d) Gibbs free energy diagram for the HER in acid media on Os<sub>NP</sub> (blue line) and Os<sub>SA</sub>-CeO<sub>2</sub> (purple and cyan lines). (e) Schematic mechanism of proton adsorption on the Os site leading to H<sub>2</sub> generation.



coupling distance (1.60 Å vs. 1.97 Å for Os<sub>NP</sub>, Fig. S2), suggesting optimized hydrogen binding and improved HER kinetics. Overall, atomic Os incorporation into CeO<sub>2</sub> yields a stable configuration with balanced H\* adsorption energy and robust structural anchoring (Fig. 1e).

### Material synthesis and structural characterization

Guided by the theoretical predictions, we synthesized the Os<sub>SA</sub>-CeO<sub>2</sub> sample consisting of Os SAs dispersed in the CeO<sub>2</sub> matrix

(Fig. 2a). Compared with the dense film structure of CeO<sub>2</sub> (Fig. S3), the O<sub>2</sub> post-treatment induced CO<sub>2</sub> release from the carbon substrate, generating a porous structure of CeO<sub>2</sub>-O<sub>2</sub> (Fig. S4) that was preserved in Os<sub>SA</sub>-CeO<sub>2</sub> (Fig. 2b). The aberration-corrected high-angle annular dark field scanning transmission electron microscopy (HAADF-STEM) image (Fig. 2c) reveals abundant dark regions within the sheet-like aggregates, corresponding to the voids seen in the scanning electron microscopy (SEM) image (Fig. 2b). Elemental mapping (Fig. 2d

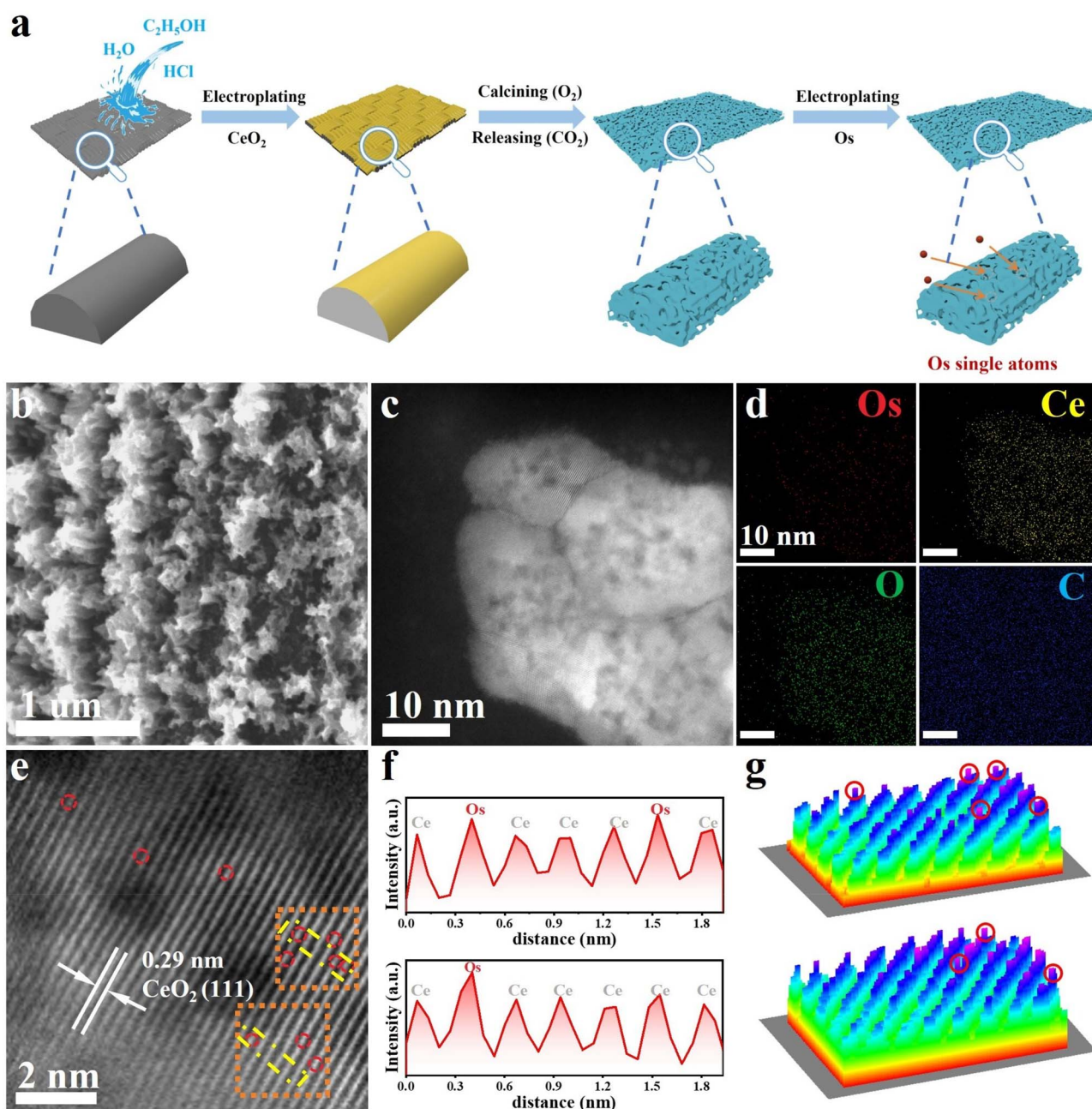


Fig. 2 Synthesis and morphological characterization of Os SAs anchored on porous CeO<sub>2</sub> (Os<sub>SA</sub>-CeO<sub>2</sub>). (a) Schematic illustration of the synthesis route for Os SAs dispersed on porous CeO<sub>2</sub>. (b) SEM image of Os<sub>SA</sub>-CeO<sub>2</sub>. (c) HAADF-STEM image and (d) corresponding elemental mapping images of Os<sub>SA</sub>-CeO<sub>2</sub>. (e) High-resolution HAADF-STEM image of Os<sub>SA</sub>-CeO<sub>2</sub>. (f) Two-dimensional atomic distribution map corresponding to the yellow region highlighted in (e). (g) Three-dimensional atomic distribution map corresponding to the orange region highlighted in (e).



and Table S1) and high-resolution STEM imaging (Fig. 2e) confirm the sheet-like  $\text{CeO}_2$  structure with atomically dispersed Os. The reduced lattice spacing (0.29 nm vs. 0.31 nm for pristine  $\text{CeO}_2$ ) indicates lattice compression due to Os incorporation. Fig. 2f shows the 2D atom distribution map of the yellow area framed in Fig. 2e. The atoms with higher intensity are Os single atoms, and those with weaker intensity are Ce atoms. The 3D atom distribution map of the orange area in Fig. 2e also demonstrates the successful preparation of Os single atoms.

Reference samples ( $\text{Os}_{\text{NP}}/\text{CeO}_2$  and  $\text{Os}_{\text{NP}}$ ) were prepared *via* the same electrodeposition route for Os (Fig. S5–S6). Compared to  $\text{Os}_{\text{SA}}-\text{CeO}_2$ , the electrodeposited  $\text{CeO}_2$  film was annealed in Ar instead of  $\text{O}_2$  in  $\text{Os}_{\text{NP}}/\text{CeO}_2$ , and the morphology remains flat for  $\text{CeO}_2$ -Ar (Fig. S7). For  $\text{Os}_{\text{NP}}$ , the carbon fiber without  $\text{CeO}_2$  was directly used as the support for the deposition of Os. In contrast to  $\text{Os}_{\text{SA}}-\text{CeO}_2$ , Os aggregated into nanoparticles in both  $\text{Os}_{\text{NP}}/\text{CeO}_2$  (Fig. S8) and  $\text{Os}_{\text{NP}}$  (Fig. S9). This indicates that the porous structure of the  $\text{CeO}_2$  matrix is crucial for obtaining atomically dispersed Os atoms due to its much larger exposed surface area (Fig. S10). In addition, the Os loading amount is also necessary for the successful formation of Os single atoms (Fig. S11).

The electronic structure of Os in samples was investigated by X-ray photoelectron spectroscopy (XPS) and X-ray absorption

near-edge structure (XANES) spectroscopy. In the Os 4f XPS spectra, the  $\text{Os}^{2+}$  fraction in  $\text{Os}_{\text{NP}}/\text{CeO}_2$  is approximately 37 at%, which is a bit higher than the 36 at% for  $\text{Os}_{\text{NP}}$  (Fig. 3a). This is probably due to the electronic transfer between  $\text{Os}_{\text{NP}}$  and the  $\text{CeO}_2$  support.<sup>31,32</sup> In contrast,  $\text{Os}_{\text{SA}}-\text{CeO}_2$  exhibits a substantially increased  $\text{Os}^{2+}$  ratio of 46 at%. As revealed in Os  $L_3$  edge XANES spectra, Os in  $\text{Os}_{\text{SA}}-\text{CeO}_2$  possesses an oxidation state between those of metallic Os and  $\text{OsO}_2$  (Fig. 3b). The much higher valence state of Os in  $\text{Os}_{\text{SA}}-\text{CeO}_2$  further evidences the successful doping of Os into  $\text{CeO}_2$  that enables strong electronic interactions between them. Also, the relatively lower valence of Os compared to  $\text{Ce}^{4+}$  will induce the generation of oxygen vacancies to keep the electronic balance,<sup>34</sup> consistent with the DFT results.

We studied the extended X-ray absorption fine structure (EXAFS) spectra to determine the coordination environments of Os in  $\text{Os}_{\text{SA}}-\text{CeO}_2$ . The Os  $L$ -edge R-space spectrum of  $\text{Os}_{\text{SA}}-\text{CeO}_2$  (Fig. 3c) shows a dominant peak at  $\sim 1.72$  Å corresponding to the Os–O band, and no peak of Os–Os bond is observed. The small peak at  $\sim 2.83$  Å is probably ascribed to the Os–Ce bond,<sup>24,35</sup> as illustrated by the Fourier transform (FT) EXAFS fitting (Fig. S12 and Table S2). These verify the atomic dispersion of Os in  $\text{CeO}_2$ , consistent with the STEM results. Wavelet transform (WT)-EXAFS analysis (Fig. 3d and e) further confirms

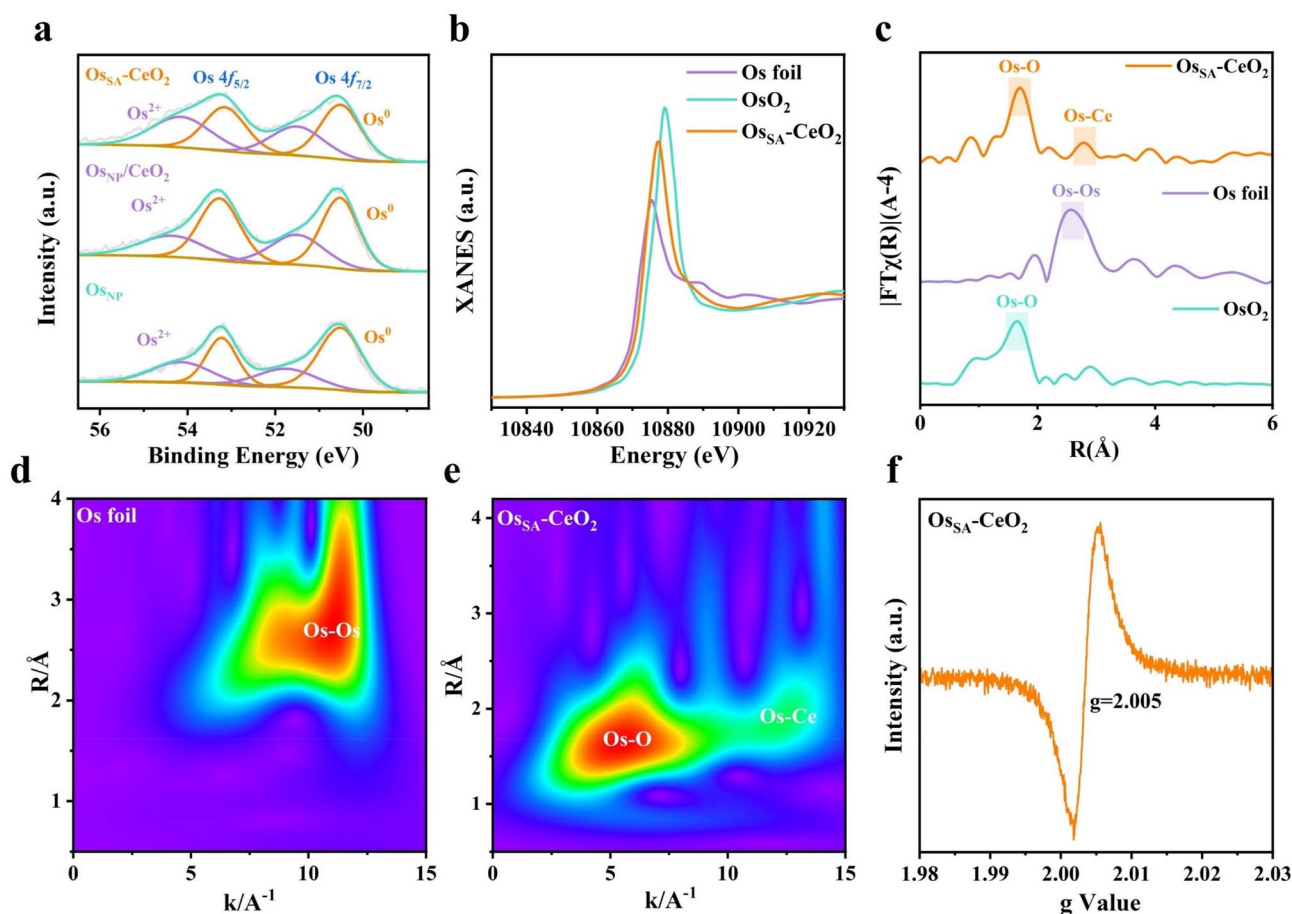


Fig. 3 Structural characterization of  $\text{Os}_{\text{SA}}-\text{CeO}_2$ . (a) High-resolution XPS spectra of Os 4f. (b) Os  $L_3$ -edge XANES spectra. (c) FT-EXAFS spectra of Os. WT of the  $k^4$ -weighted EXAFS spectra of (d)  $\text{Os}_{\text{SA}}-\text{CeO}_2$  and (e) Os foil. (f) EPR of  $\text{Os}_{\text{SA}}-\text{CeO}_2$ .



the appearance of Os–O and Os–Ce bands in Os<sub>SA</sub>–CeO<sub>2</sub>. The Os–O coordination number (3.2, Table S2) is significantly lower than the Ce–O coordination (7.7) in CeO<sub>2</sub>, indicating abundant oxygen vacancies, as confirmed by the electron paramagnetic resonance (EPR, Fig. 3f) and O 1s spectrum (Fig. S13) of Os<sub>SA</sub>–CeO<sub>2</sub>. The shorter Os–Ce bond (2.83 Å vs. Ce–Ce 3.30 Å) reflects lattice compression, consistent with the STEM results (Fig. 2e). Thus, we can conclude that the porous CeO<sub>2</sub> support enables the successful fabrication of atomically dispersed Os SAs,

accompanied by the generation of numerous oxygen vacancies within the CeO<sub>2</sub> matrix.

### Electrochemical performances

HER measurements were conducted in 0.5 M H<sub>2</sub>SO<sub>4</sub> using a three-electrode setup. As shown in Fig. 4a, Os<sub>SA</sub>–CeO<sub>2</sub> exhibits markedly superior HER activity compared to Os<sub>NP</sub>/CeO<sub>2</sub> and Os<sub>NP</sub>, approaching the performance of commercial 20% Pt/C.

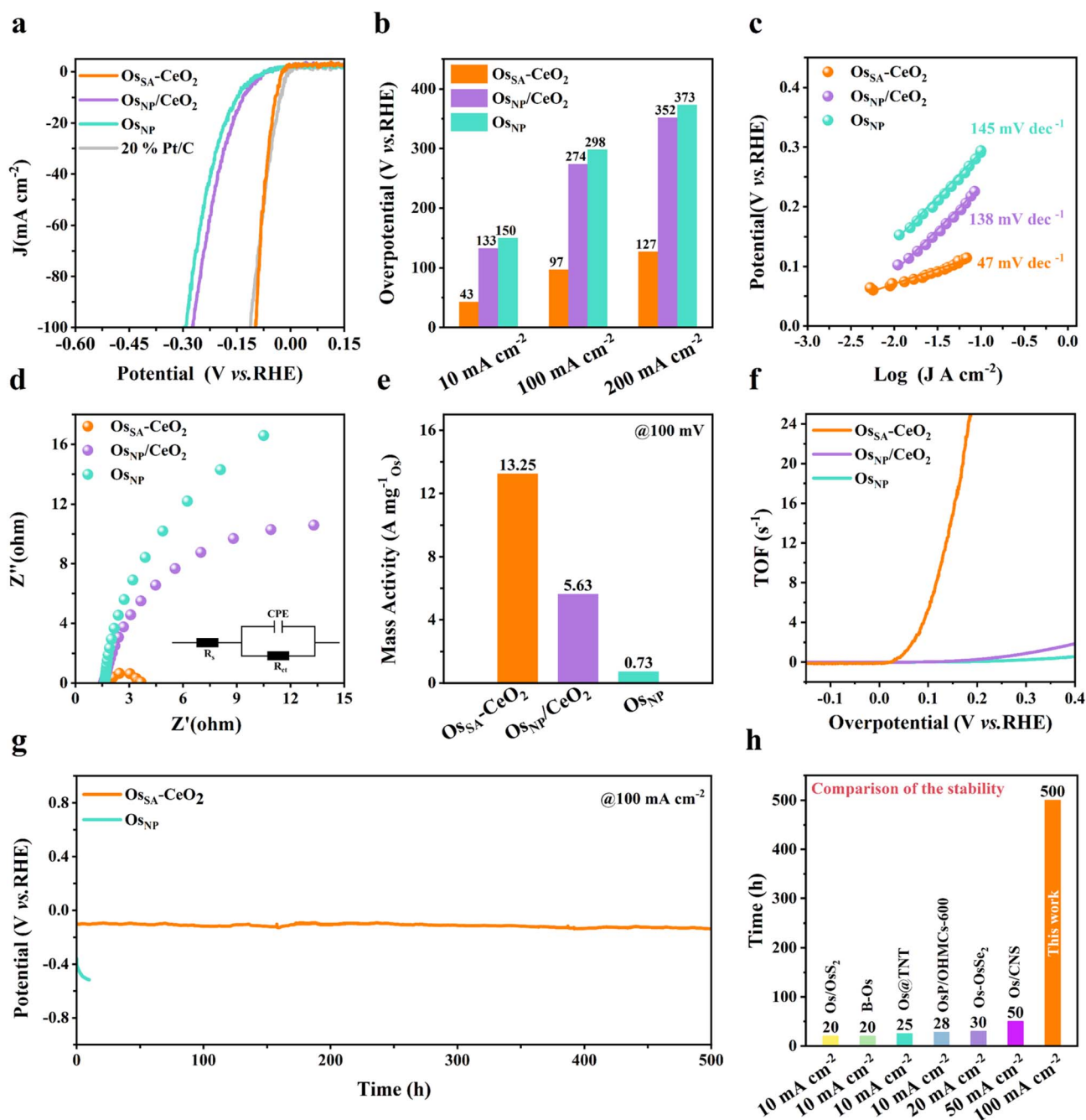


Fig. 4 HER performance of Os<sub>SA</sub>–CeO<sub>2</sub> in acidic media. (a) HER polarization curves of various catalysts in 0.5 M H<sub>2</sub>SO<sub>4</sub>. (b) Comparison of overpotentials at 10, 100 and 200 mA cm<sup>-2</sup>. (c) Tafel plots. (d) Nyquist plots; the inset shows the corresponding equivalent electrical circuit used for fitting. (e) Mass activity at an overpotential of 100 mV. (f) Relationship between TOF values and overpotentials. (g) Stability tests of Os<sub>SA</sub>–CeO<sub>2</sub> and Os<sub>NP</sub>. (h) Comparison of the stability of Os<sub>SA</sub>–CeO<sub>2</sub> with previously reported Os-based HER catalysts in acidic media.<sup>14,15,18–21</sup>



The overpotentials of  $\text{Os}_{\text{SA}}\text{-CeO}_2$  are only 43 mV and 97 mV at current densities of 10 and 100  $\text{mA cm}^{-2}$ , respectively (Fig. 4b). The Tafel slope (47  $\text{mV dec}^{-1}$ , Fig. 4c) is significantly lower than that of  $\text{Os}_{\text{NP}}/\text{CeO}_2$  (139  $\text{mV dec}^{-1}$ ) and  $\text{Os}_{\text{NP}}$  (145  $\text{mV dec}^{-1}$ ), indicating faster kinetics and a favorable HER pathway.<sup>36–38</sup> Electrochemical impedance spectroscopy (Fig. 4d) shows that  $\text{Os}_{\text{SA}}\text{-CeO}_2$  possesses the lowest charge transfer resistance, as summarized in Table S3. These factors contribute to its exceptionally high mass activity (Fig. 4e) and TOF (Fig. 4f). The mass activity of  $\text{Os}_{\text{SA}}\text{-CeO}_2$  is 13.25  $\text{A mg}_{\text{Os}}^{-1}$  at 100 mV overpotential (Fig. 4e), 2.4 times higher than that of  $\text{Os}_{\text{NP}}/\text{CeO}_2$  and 18.2 times higher than that of  $\text{Os}_{\text{NP}}$  (Table S4). The effect of Os loading on HER performance was further investigated by varying the deposition cycles (Fig. S14). Os-50 CV ( $\text{Os}_{\text{SA}}\text{-CeO}_2$ ) exhibits markedly higher activity than Os-10 CV, while showing slightly lower activity than Os-90 CV. However, excessive Os loading in Os-90 CV leads to the formation of large aggregated nanoparticles (Fig. S11c), which is detrimental to achieving atomically dispersed active sites. Therefore, Os-50 CV, which balances high catalytic activity with atomic dispersion, was selected as the representative sample for further investigation.

Post-HER XPS analysis (Fig. S15) reveals minimal changes in Os binding energy for  $\text{Os}_{\text{SA}}\text{-CeO}_2$ , whereas  $\text{Os}_{\text{NP}}/\text{CeO}_2$  and  $\text{Os}_{\text{NP}}$

exhibit positive shifts, indicating Os oxidation. Due to the strong interactions between Os and  $\text{CeO}_2$  in  $\text{Os}_{\text{SA}}\text{-CeO}_2$ , the electron-buffering effect of  $\text{CeO}_2$  regulates the electronic structure of Os and prevents its oxidation dissolution.<sup>33,39</sup> Controlled-current water electrolysis (Fig. 4g) was further done to test the long-term performance and stability.  $\text{Os}_{\text{SA}}\text{-CeO}_2$  maintains stable operation for over 500 h at 100  $\text{mA cm}^{-2}$ , far exceeding  $\text{Os}_{\text{NP}}$  and previously reported Os-based HER catalysts (Fig. 4h).

### Insights into the origin of enhanced stability at large current densities

pH-dependent experiments (Fig. S16 and 5a) were performed to simulate the insufficient proton supply conditions that occur during high current operation. The overpotential at 100  $\text{mA cm}^{-2}$  increases sharply with pH for  $\text{Os}_{\text{NP}}$  (slope = 171), indicating strong proton dependence, whereas  $\text{Os}_{\text{SA}}\text{-CeO}_2$  shows a much smaller slope (16), suggesting additional water-splitting participation. Kinetic isotope effect (KIE) experiments (Fig. 5b) confirm this: HER performance in 0.5 M  $\text{H}_2\text{SO}_4$  with  $\text{D}_2\text{O}$  is significantly lower than in 0.5 M  $\text{H}_2\text{SO}_4$  with  $\text{H}_2\text{O}$ , due to slower  $\text{D}^+$  migration and higher D–D bond energy, implying that  $\text{H}_2\text{O}$  dissociation contributes to  $\text{H}^*$  generation on  $\text{Os}_{\text{SA}}\text{-CeO}_2$ . *Operando* infrared spectroscopy (Fig. 5c) further supports this

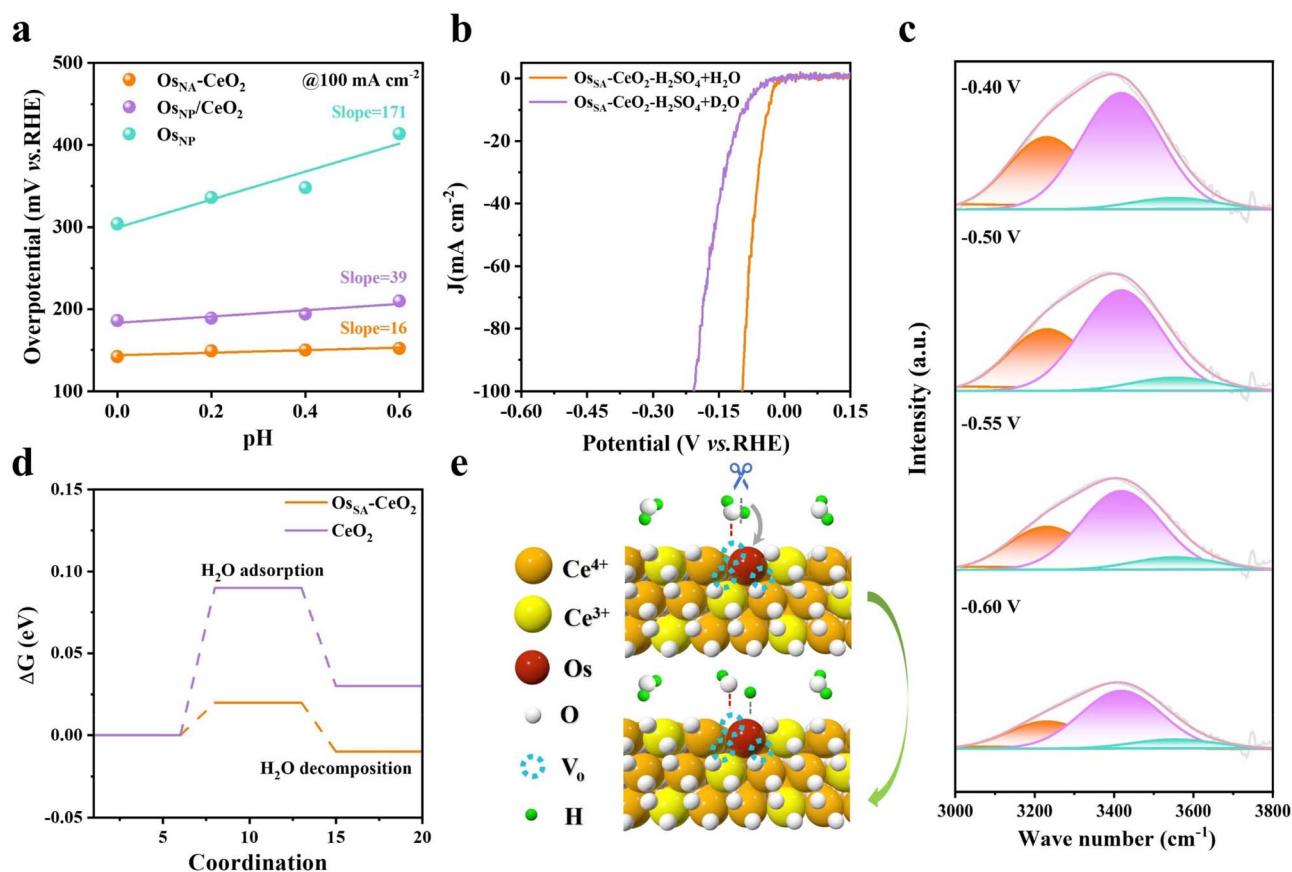


Fig. 5 Mechanistic insights into the  $\text{H}_2\text{O}$  splitting ability of  $\text{Os}_{\text{SA}}\text{-CeO}_2$ . (a) Overpotentials as a function of pH at 100  $\text{mA cm}^{-2}$  for  $\text{Os}_{\text{SA}}\text{-CeO}_2$ ,  $\text{Os}_{\text{NP}}/\text{CeO}_2$ , and  $\text{Os}_{\text{NP}}$ . (b) LSV curves of  $\text{Os}_{\text{SA}}\text{-CeO}_2$  measured in 0.5 M  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O}$  and 0.5 M  $\text{H}_2\text{SO}_4 + \text{D}_2\text{O}$ . (c) *In situ* FTIR spectroscopy of  $\text{Os}_{\text{SA}}\text{-CeO}_2$  at different applied potentials. (d) Gibbs free energy diagrams for  $\text{H}_2\text{O}$  adsorption and decomposition on  $\text{Os}_{\text{SA}}\text{-CeO}_2$  and  $\text{CeO}_2$ . (e) Schematic illustration of the HER reaction pathway, showing  $\text{H}_2\text{O}$  as the supplied hydrogen source.



conclusion. The broad band from 3000–3800  $\text{cm}^{-1}$  corresponds to interfacial water species, including 4-hydrogen bonding (HB)- $\text{H}_2\text{O}$  (3231  $\text{cm}^{-1}$ ), 2-HB- $\text{H}_2\text{O}$  (3418  $\text{cm}^{-1}$ ), and free  $\text{H}_2\text{O}$  (3553  $\text{cm}^{-1}$ ). With increasing negative potential, the proportion of strongly bound 4-HB- $\text{H}_2\text{O}$  decreases, transforming into free  $\text{H}_2\text{O}$  (Table S5), indicating active water participation during the HER.<sup>40</sup>

The facile participation of  $\text{H}_2\text{O}$  for the HER on  $\text{Os}_{\text{SA}}\text{-CeO}_2$  is probably due to the abundant oxygen vacancies. The oxygen vacancies can act as the strong adsorption sites for the oxygen-containing species like  $\text{H}_2\text{O}$ , which is confirmed by the DFT calculations (Fig. 5d). The Gibbs free energy barrier ( $\Delta G$ ) for  $\text{H}_2\text{O}$  adsorption decreases from 0.09 eV on  $\text{CeO}_2$  to 0.02 eV on  $\text{Os}_{\text{SA}}\text{-CeO}_2$ , greatly facilitating  $\text{H}_2\text{O}$  dissociation. As summarized in Fig. 5e, Os doping induces oxygen vacancies that promote  $\text{H}_2\text{O}$  adsorption and splitting. The generated  $\text{H}^*$  species migrate to adjacent Os sites, serving as a hydrogen source to support the rapid proton consumption at large current densities, thereby enabling stable HER operation. Also, the generated  $\text{OH}^*$  species are adsorbed by the oxygen vacancies, thus preventing the Os sites from oxidative dissolution.

## Conclusions

In summary, we have successfully synthesized an  $\text{Os}_{\text{SA}}\text{-CeO}_2$  catalyst composed of atomically dispersed Os single atoms strongly anchored on a porous  $\text{CeO}_2$  support. The porous  $\text{CeO}_2$  architecture provides abundant anchoring sites that stabilize Os atoms and strengthen the Os– $\text{CeO}_2$  interfacial coupling. This interaction optimizes the electronic structure of Os, moderates  $\text{H}^*$  binding strength, and induces abundant oxygen vacancies that activate  $\text{CeO}_2$  toward enhanced  $\text{H}_2\text{O}$  dissociation. Benefiting from these synergistic effects, the  $\text{Os}_{\text{SA}}\text{-CeO}_2$  catalyst exhibits outstanding HER activity in acidic electrolytes, surpassing commercial Pt/C at high current densities. It delivers a remarkable mass activity of 13.25  $\text{A mg}^{-1}$  at 100 mV overpotential and maintains stable operation for over 500 h at 100  $\text{mA cm}^{-2}$  with negligible degradation. This work provides a new design strategy to enhance the stability of noble-metal catalysts without sacrificing activity by engineering strong metal–support interfacial interactions.

## Author contributions

J. Y. conceptualized the project. J. Y. and L. Z. supervised the work. Q. L. performed catalyst synthesis and general characterization. Q. L., M. L., W. L., and Y. H. contributed to experimental design and data analysis. Q. L. and J. Y. wrote the original manuscript draft. All authors contributed to discussions on the data and to the development of the manuscript.

## Conflicts of interest

There are no conflicts to declare.

## Data availability

The data supporting this article have been included as part of the supplementary information (SI). Supplementary information: experimental methods, additional DFT calculations, structure models for DFT calculations; SEM, TEM, schematic illustration of the synthesis route, EXAFS, XPS and ICP-OES; additional electrochemical data and summarized table for situ FTIR spectroscopy. See DOI: <https://doi.org/10.1039/d5sc09741j>.

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