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Defect engineering of two-dimensional Nb-based oxynitrides for visible-light-driven water splitting to produce H₂ and O₂†

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Two-dimensional (2D) Nb-based oxynitrides are promising visible-light-responsive photocatalysts for the water splitting reaction, but their photocatalytic activity is degraded by the formation of reduced Nb⁵⁺ species and O²⁻ vacancies. To understand the influence of nitridation on the formation of crystal defects, this study synthesized a series of Nb-based oxynitrides through the nitridation of LaKNb_{1-x}Ta_xO₅ (x = 0, 0.2, 0.4, 0.6, 0.8, 1.0). During nitridation, K and Na species volatilized, which helped transform the exterior of LaKNb_{1-x}Ta_xO₅ into a lattice-matched oxynitride shell. Ta inhibited defect formation, yielding Nb-based oxynitrides with a tunable bandgap between 1.77 and 2.12 eV, straddling the H₂ and O₂ evolution potentials. After loading with Rh and CoO_x cocatalysts, these oxynitrides exhibited good photocatalytic activity for H₂ and O₂ evolution in visible light (650–750 nm). The nitrided LaKNbTaO₅ and LaKNbNb_{0.8}Ta_{0.2}O₅ delivered the maximum H₂ (19.37 μmol h⁻¹) and O₂ (22.81 μmol h⁻¹) evolution rates, respectively. This work provides a strategy for preparing oxynitrides with low defect densities and demonstrates the promising performance of Nb-based oxynitrides for water splitting.

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Introduction

Photocatalytic water splitting into hydrogen (H₂) and oxygen (O₂) is regarded as one of the most promising strategies for solving the global energy crisis and mitigating environmental problems.^{1–3} In the solar spectrum, UV light accounts for less than 5% of the energy in sunlight, whereas visible light constitutes approximately 54%.^{4–6} In order to achieve high photocatalytic efficiency, semiconductors with narrow band gap energies (E_g) that are capable of absorbing the wide wavelengths of visible light are very attractive.^{7,8} Recently, metal oxynitrides are being reported continuously, whose E_g is significantly narrower than that of conventional metal oxides because N 2p orbitals provide a new valence band (VB) that is more negative than that provided by O 2p orbitals.^{9–11}

Among these oxynitrides, Nb-based oxynitrides exhibit a broad light absorption band in the visible light region, with λ_{\max} values up to approximately 750 nm, which is much higher than their Ta-based counterparts.^{12–15} To date, various Nb-based oxynitrides, such as CaNbO₂N (λ_{\max} = 600 nm), SrNbO₂N (λ_{\max} = 690 nm), BaNbO₂N (λ_{\max} = 740 nm) and LaNbON₂ (λ_{\max} = 750 nm) have been synthesized by nitriding oxide precursors under

an NH₃ flow and applied for photocatalysis. For instance, CaNbO₂N was found to be active for O₂ evolution from a AgNO₃ aqueous solution.¹² The optimal nitridation temperature in that study was determined to be 1023 K. SrNbO₂N prepared by nitriding Sr₂Nb₂O₇ or SrNbO₃ exhibited enhanced activity for photoelectrochemical water splitting, with a current density of 0.77 mA cm⁻² at 1.2 V *versus* the reversible hydrogen electrode (RHE).¹³ Meanwhile, Hisatomi *et al.* improved the crystallinity and uniformity of BaNbO₂N by adding BaCO₃ into the Ba₅Nb₄O₁₅ precursor during the synthesis procedure, and the resulting BaNbO₂N exhibited enhanced O₂ evolution activity by utilizing photons with wavelengths of up to 740 nm.¹⁴ Recently, Wang *et al.* prepared LaNbON₂ with exposed metastable {010} facets by nitriding a plate-like LaKNbO₅ precursor. In combination with a CoO_x co-catalyst, LaNbON₂ exhibited enhanced photocatalytic activity for O₂ evolution, with an apparent quantum yield of 0.82%, whereas conventional LaNbON₂ was almost inactive.¹⁵ However, nitridation is a harsh process. N doping always causes a charge imbalance as a result of aliovalent O²⁻/N³⁻ exchange, yielding numerous crystal defects that are detrimental to the photocatalytic performance of oxynitrides.^{16,17} Like Nb-based oxynitrides, it still exhibits negligible photocatalytic activity, especially for photocatalyzing H₂ evolution, because of the formation of large amounts of reduced Nb⁵⁺ species and O²⁻ vacancies during the nitridation process. These vacancies can act as recombination and trapping centers for photoexcited charges, resulting in a decrease in charge separation efficiency.^{18–20} Moreover, a low concentration

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of the N dopant inevitably can also decrease the absorption coefficient of N-doped oxides in the visible light region.²¹ Therefore, undoped oxynitrides containing a stoichiometric amount of N in their crystal structure are highly desirable, which can guarantee strong visible light absorption through band-to-band electron transitions and the development of Nb-based oxynitrides with low defect densities remains a challenge.

Herein, we synthesized a series of Nb-based oxynitrides through the nitridation of LaKNaNb_{1-x}Ta_xO₅ ($x = 0, 0.2, 0.4, 0.6, 0.8, 1.0$). By introducing Ta, the formation of defects in their crystal structure is effectively inhibited yielding a tunable bandgap. In the meantime, volatilization of K and Na species during nitridation promoted LaKNaNb_{1-x}Ta_xO₅ crystal surfaces to be transformed into lattice-matching oxynitrides, forming a core-shell structure. Considering the above mentioned enhanced properties, these oxynitrides exhibited good photocatalytic activity for H₂ and O₂ evolution in visible light (650–750 nm).

Results and discussion

Fig. 1 illustrates the crystal structure of LaKNaNb_{1-x}Ta_xO₅, which is similar to that of LaKNaNbO₅ (Fig. S1†) and has LaNaNbO₅⁻ layers along the [001] direction with K⁺ ions occupying the spaces between the layers. Each LaNaNbO₅ layer is arranged in a “checkerboard” pattern, where NaO₅ and NbO₅ square pyramids are arranged in alternating rows.^{23,24} In LaKNaNb_{1-x}Ta_xO₅, a fraction of the Nb atoms are substituted with Ta atoms. Accordingly, LaNbON₂ has an orthorhombic perovskite structure with the space group *Pnma* (Fig. 1). The unit cell parameters of LaNbON₂ are $a = 5.7221 \text{ \AA}$, $b = 8.0684 \text{ \AA}$, and $c = 5.7428 \text{ \AA}$.²⁵ As the amount of Ta⁵⁺ increases, Nb_xTa_{1-x}O₅ square pyramids in LaNb_{1-x}Ta_xON₂ become distorted because Ta–(O, N) bonds (2.056 Å) are shorter than Nb–(O, N) bonds (2.064 Å). As a result, pure LaTaON₂ demonstrated a monoclinic perovskite structure, indicating the successful substitution of Nb by Ta in the oxynitride lattice with angles from 90° to 135° between the *a*- and *c*-axes (Fig. S2†).

Plate-like LaKNaNb_{1-x}Ta_xO₅ was prepared by a NaOH/KOH flux method as previously reported.^{26,27} Specifically, 5 mmol of La₂O₃ and 5 mmol in total of Ta₂O₅ and Nb₂O₅ at various molar ratios (0 : 5, 1 : 4, 2 : 3, 3 : 2, 4 : 1, and 5 : 0) were combined and finely ground in a mortar, after which the mixture was

transferred to an alumina crucible containing an excess of KOH (10 g) and NaOH (5 g). Subsequently, the mixture was calcined at 873 K for 3 h and aged at 773 K for 15 h. In this synthesis, KOH and NaOH were in excess during synthesis to serve as major flux components for the crystal growth of LaKNaNb_{1-x}Ta_xO₅ oxides, resulting in large, highly crystalline products. A scheme of the synthetic procedure is shown in Fig. S3.†

The XRD patterns of the LaKNaNb_{1-x}Ta_xO₅ precursor ($x = 0, 0.2, 0.4, 0.6, 0.8, \text{ and } 1$) were nearly identical to those of the tetragonal LaKNaNbO₅ phase with a space group of *P4/nmm*, indicating the phase purity of the flux-grown oxide crystals (Fig. 2a). Notably, the intensity of the diffraction peak at 10° for LaKNaNbO₅ was enhanced, indicating the exposure of the {001} facets on the surface. Interestingly, with the increasing Ta content, the position of the diffraction peak of LaKNaNb_{1-x}Ta_xO₅ shifted to a slightly higher angle than that of the LaKNaNbO₅ crystals, indicating that the substitution of Nb with Ta caused a decrease in the crystal size and an overall contraction of the lattice parameters in the LaKNaNb_{1-x}Ta_xO₅ samples.^{28,29} After nitridation at 1173 K for 4 h, peaks attributable to the LaNbON₂ phase were observed together with intense LaKNaNbO₅ peaks, which could be ascribed to the transformation of the surface of the LaKNaNbO₅ precursor into LaNbON₂ during nitridation, while the crystal interior remained an oxide (Fig. 2b). Interestingly, the intensity of the diffraction peaks attributed to the LaNb_{1-x}Ta_xON₂ phase became weaker with the increasing Ta content. This finding indicates that introducing Ta species can slow down the transformation from the LaKNaNb_{1-x}Ta_xO₅ precursor into its lattice-matched LaNb_{1-x}Ta_xON₂ owing to the lower electronegativity of the Ta

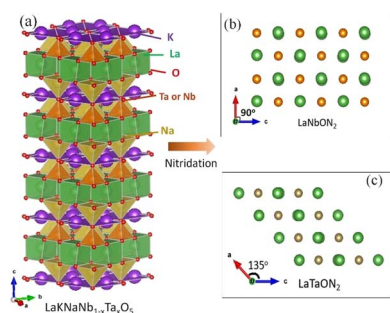


Fig. 1 (a) Crystal structure of layered LaKNaNb_{1-x}Ta_xO₅. (b and c) Views of the (010) facets of (b) LaNbON₂, and (c) LaTaON₂. The crystal structures were visualized using the VESTA suite of programs.²²

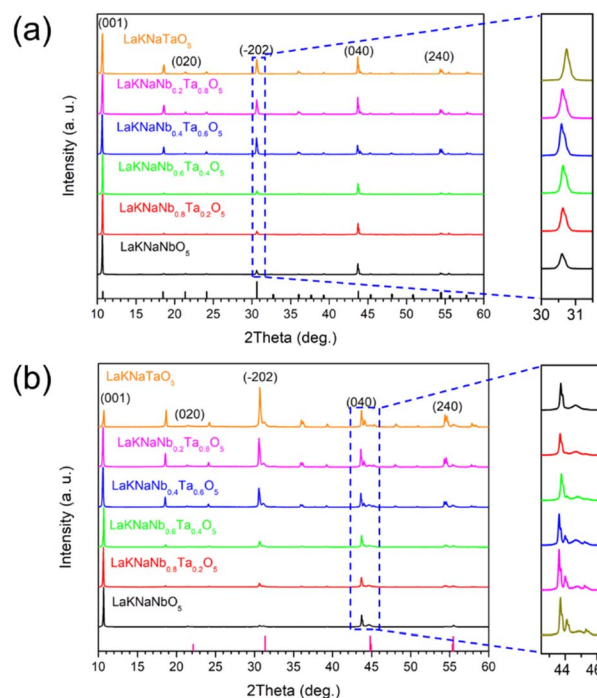


Fig. 2 XRD patterns of the (a) layered LaKNaNb_{1-x}Ta_xO₅ oxide precursors and (b) nitrided LaKNaNb_{1-x}Ta_xO₅ ($x = 0, 0.2, 0.4, 0.6, 0.8, 1.0$).



species. Consequently, this transformation process can be finely tuned by varying the degree of substitution.

SEM was used to investigate the size and morphology of the resulting products. As shown in Fig. 3a and b, the oxide precursors that were not thermally aged at 773 K for 15 h exhibited an incomplete plate morphology with a rough surface. In contrast, the aged $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ showed a plate-like shape with a smooth surface, indicating that the long-term flux process facilitates the synthesis of oxide precursors with controllable morphology and high crystallinity (Fig. S4†). After nitridation, the prepared sample maintained its plate-like shape but became porous, as illustrated in Fig. 3c and S5.† The corresponding cross-sectional SEM image in Fig. 3d shows the formation of a core-shell structure based on a porous oxynitride shell on a dense oxide core. The core-shell structure has the advantage of generating as many oxynitrides on the oxide surface as possible during the mild nitridation process while minimizing the number of crystal defects of oxynitrides as much as possible.³⁰

The exposed facets of oxynitrides after nitridation were confirmed using TEM and high-resolution TEM (HRTEM) images (Fig. 4a and b, respectively), where the interplanar distances associated with the (200) and (002) facets of $\text{LaNb}_{1-x}\text{Ta}_x\text{ON}_2$ were 0.228 and 0.202 nm, respectively, with an angle of approximately 105° . This finding indicates that the La and Nb atoms are located in the expected positions in the orthorhombic structure viewed in the [010] direction, which energetically favors the {010} facet on the surface. The energy dispersive spectroscopy (EDS) mapping images shown in Fig. 4c show that the resulting product is composed of La, K, Na, Nb, Ta, O, and N. Comparing the atomic ratios obtained from the EDS spectrum after nitridation with the known composition before nitridation reveals that the amounts of K and Na decreased dramatically during the nitridation process (Fig. S6†).

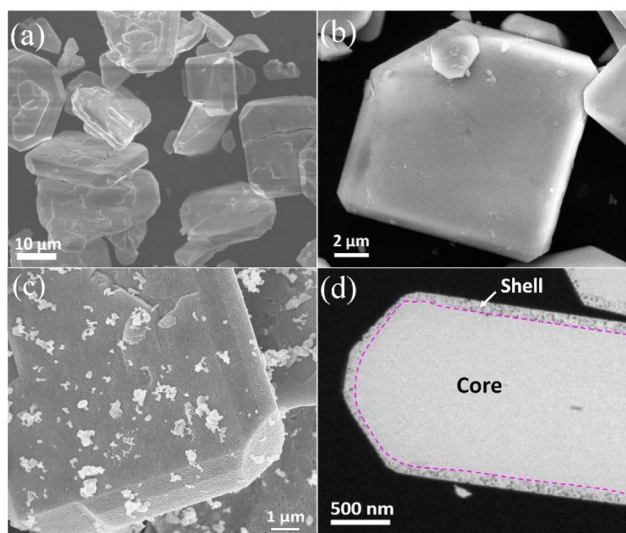


Fig. 3 SEM images of layered $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ (a) calcined at 873 K for 3 h and (b) then aged at 773 K for 15 h (c) surface and (d) cross-sectional SEM images of $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ powder nitridated at 1123 K for 4 h, revealing the porous surface.

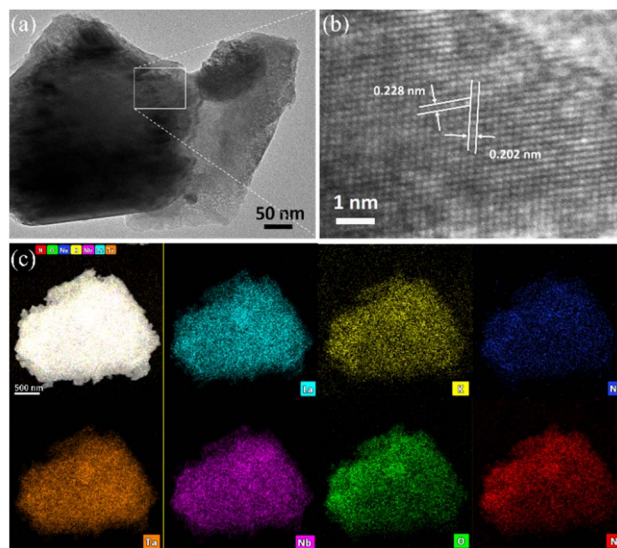


Fig. 4 (a) TEM and (b) HRTEM images and (c) EDS elemental maps of plate-like $\text{LaKNaNb}_{0.8}\text{Ta}_{0.2}\text{O}_5$ nitridated at 1123 K for 4 h. The atomic ratios were approximately $\text{La}/\text{Nb} \approx 1.3$, $\text{K}/\text{Nb} < 0.05$, $\text{Na}/\text{Nb} < 0.05$, $\text{O}/\text{Nb} \approx 1.3$, and $\text{N}/\text{Nb} \approx 2.5$.

In the photographs of the nitrided $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ powder samples (Fig. 5a), the color of the samples clearly changes from dark brown to orange with the increasing amount of Ta^{5+} . Meanwhile, in the UV-vis diffraction reflectance spectra (DRS) of the nitrided $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples presented in Fig. 5b, the absorption edges of the nitrided $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples gradually blue-shifted from 720 to 600 nm with the increasing Ta content, indicating that the bandgap energies of these materials were greater than those of pure LaNbON_2 ($\lambda \approx 740$ nm) during mild nitridation.¹⁵ The corresponding Tauc plots of the nitrided $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples (Fig. 5c) based on the UV-vis DRS spectra reveal that the bandgap energy increased monotonically from 1.77 to 2.12 eV with increasing substitution with Ta^{5+} . Moreover, their bandgap positions were estimated based on the VB XPS data (Fig. S7 and Table S1†)

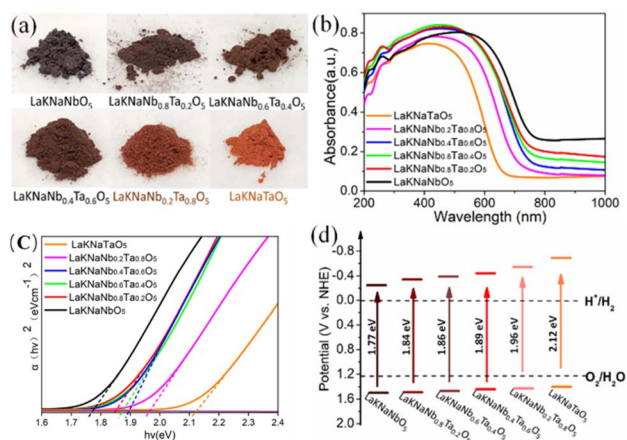


Fig. 5 (a) Photographs, (b) UV-vis DRS, (c) corresponding Tauc plots, and (d) band structure diagrams of the $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples ($x = 0, 0.2, 0.4, 0.6, 0.8, 1.0$) nitridated at 1123 K for 4 h.



along with the Tauc plots. As shown in Fig. 5d, the VB of the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples remained almost unchanged with the increasing amount of Ta^{5+} , and the enlarged bandgaps of these samples mainly resulted from the upshift of the conduction band (CB), demonstrating that the Ta CB orbitals mainly contributed to the Ta 5d orbitals.

Notably, the nitrated LaKNaNbO_5 samples exhibited obvious background absorption at wavelengths >650 nm, which could be ascribed to the formation of a large amount of reduced Nb^{5+} species as well as O^{2-} vacancies.^{18,19} Reduced metallic species are known not to exhibit photocatalytic activity owing to bandgap excitation, and such defect states usually generated by prolonged exposure to NH_3 during nitridation at high temperatures can act as recombination and trapping centers for photoexcited charges. In particular, the background absorption became much weaker as the amount of Ta^{5+} increased, indicating that fewer crystal defects formed during the nitridation process. This is because Ta is less electronegative than Nb, which slowed down the nitridation process, thus inhibiting the typical formation of crystal defects in the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples.

The chemical compositions of the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples were analyzed using high-resolution XPS. The survey spectra are shown in Fig. S8a.† They confirm the presence of Ta, Nb, N and O in the as-prepared nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples. By comparing with the unnitrated sample (Fig. S8b†), the presence of N on the surface of nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ can be confirmed. As shown in Fig. 6a, the

Nb valence species at the surface of the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ sample were evaluated to investigate the effect of adding Ta^{5+} on the reduction of Nb^{5+} during nitridation. The Nb 3d XPS spectra were deconvoluted into peaks for three different Nb valence species,³¹ Nb^{5+} , Nb^{4+} , and Nb^{3+} , based on the tailing of the peaks toward lower binding energies, demonstrating that more than one chemical state exists in the samples. Here, the presence of Nb^{4+} and Nb^{3+} indicates the reduction of the Nb^{5+} species derived from NH_3 nitridation. These reduced Nb^{4+} and Nb^{3+} species can act as charge trapping and recombination centers for the photoexcited electrons and holes. Table S2† summarizes the proportions of surface Nb species in the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ sample with different amounts of Ta^{5+} . With increasing Ta^{5+} , the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples exhibited the decreased presence of reduced Nb species (both Nb^{3+} and Nb^{4+}) from 82.06% for nitrated LaKNaNbO_5 without Ta to 48.61% for nitrated $\text{LaKNaNb}_{0.2}\text{Ta}_{0.8}\text{O}_5$. These results indicate that substitution with Ta^{5+} species can effectively inhibit the reduction of Nb.

A similar phenomenon was also observed for the Ta^{5+} species in the deconvoluted Ta 4f XPS spectra of the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples, which can be deconvoluted into two different reduced Ta species,^{32,33} Ta^{5+} and Ta^{4+} , based on the binding energies (Fig. 6b).³⁰ The increasing Ta^{5+} content was found to inhibit the reduction of Ta^{5+} species, from 70.72% for the nitrated $\text{LaKNaNb}_{0.2}\text{Ta}_{0.8}\text{O}_5$ sample to 25.84% for the nitrated LaKNaNbO_5 sample (Table S3†). Therefore, according to the band diagrams and XPS data, the replacement of Nb^{5+} with Ta^{5+} evidently hindered the reduction of both Nb and Ta during nitridation, thus tuning the band structure of the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ sample while maintaining the visible-light-driven photocatalytic activity.

Moreover, the O 1s state of the oxide precursor could be decomposed into two overlapping peaks centered at 529.6 and 532.5 eV (Fig. S9†). These two peaks could be attributed to lattice oxygen in the oxide (O_L) and hydroxyl groups (O-H) that adsorbed on the sample surfaces during sample handling, respectively.^{34,35} As illustrated in Fig. S10,† a new peak centered at 531.4 eV appeared after nitridation, the intensity of which is partly related to the concentration of O^{2-} vacancies. Notably, the intensity of the O_v peak decreased from 43.27% for LaKNaNbO_5 (no Ta^{5+}) to 20.75% for LaKNaNbO_5 in which Nb^{5+} was completely replaced with Ta^{5+} , indicating that fewer O^{2-} vacancies formed in the nitrated samples (Table S4†).

The photocatalytic performance of the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples toward the evolution of H_2 and O_2 under visible light was examined in an aqueous methanol solution loaded with a Rh co-catalyst and an aqueous AgNO_3 solution loaded with the CoO_x co-catalyst. Fig. 7a shows the H_2 evolution as a function of time. Evidently, all the nitrated $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples showed photocatalytic activity for H_2 evolution. However, the nitrated sample without Ta^{5+} showed the lowest activity ($0.2635 \mu\text{mol h}^{-1}$, Table S5†), which could be ascribed to the formation of large amounts of reduced Nb^{5+} species and O^{2-} vacancies in the LaNbON_2 shells, as discussed in the analyses of the DRS and XPS data (Fig. 5b and 6). In contrast, the activity of the nitrated LaKNaNbO_5 sample was

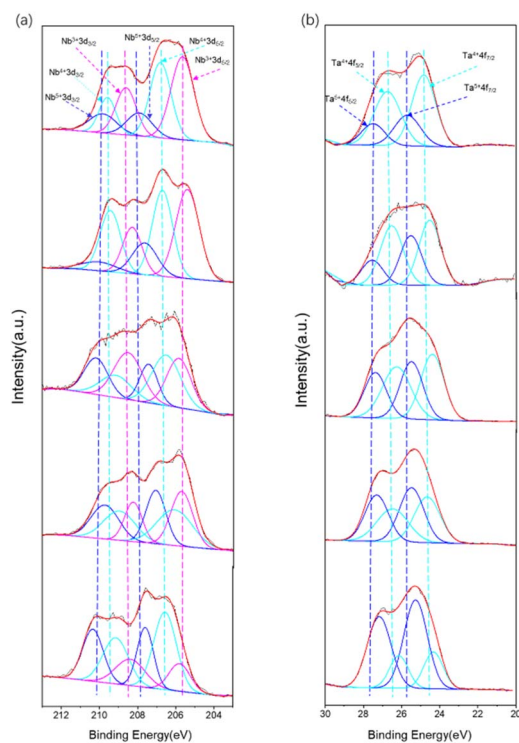


Fig. 6 High-resolution XPS spectra of the $\text{LaKNaNb}_{1-x}\text{Ta}_x\text{O}_5$ samples nitrated at 1123 K for 4 h: (a) Nb 3d spectra for $x = 0, 0.2, 0.4, 0.6$ and 0.8 from up to down; (b) Ta 4f spectra for $x = 0.2, 0.4, 0.6, 0.8$ and 1.0 (no Nb) from up to down.



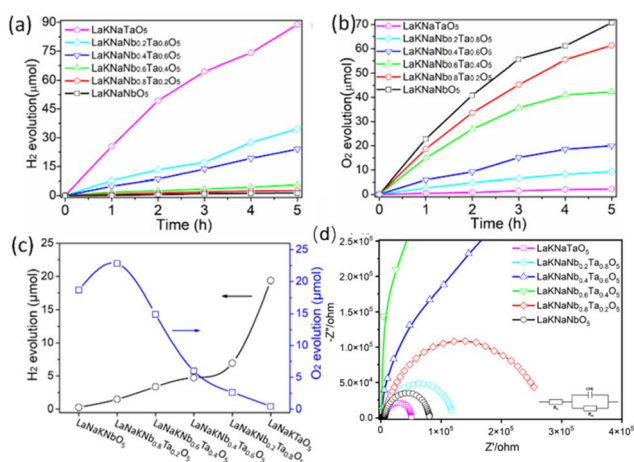


Fig. 7 Photocatalytic performance of LaKNaNb_{1-x}Ta_xO₅ nitrided at 1123 K for 4 h ($x = 0, 0.2, 0.4, 0.6, 0.8, 1.0$): (a) H₂ and (b) O₂ evolution as a function of time, (c) H₂ and O₂ evolution rates, and (d) EIS spectra recorded under visible light illumination. Reaction conditions for H₂ evolution: 150 mL of aqueous methanol (20 vol%), photocatalyst: 300 mg, co-catalyst: 0.5 wt% Rh, loading method: impregnation with H₂ reduction, light source: a 300 W xenon lamp ($\lambda \geq 420$ nm), reaction cell: a top-irradiation cell with a Pyrex window. Reaction conditions for O₂ evolution: 150 mL of an aqueous 50 mM AgNO₃ solution containing 0.20 g of La₂O₃, photocatalyst: 300 mg, co-catalyst: 0.5 wt% CoO_x, light source: a 300 W xenon lamp ($\lambda \geq 420$ nm), reaction cell: a top-irradiation cell with a Pyrex window.

enhanced by the increased substitution of Nb with Ta⁵⁺. Thus, introducing Ta⁵⁺ species effectively inhibited the reduction of Nb species, thereby decreasing the possibility of photo-generated charges recombining at defects and enhancing the photocatalytic activity toward H₂ evolution.^{15,18,30} After complete Ta⁵⁺ substitution, the nitrided LaKNaTaO₅ sample showed the highest photocatalytic H₂ evolution activity (19.37 $\mu\text{mol h}^{-1}$) because of the presence of a low-defect-density LaTaON₂ shell on the lattice-matched LaKNaTaO₅ core.

Moreover, the photocatalytic activity of nitrided LaKNaNb_{1-x}Ta_xO₅ for O₂ evolution was evaluated by modifying the CoO_x co-catalyst (Fig. 7b). The treatment temperature of the CoO_x co-catalyst modified on the sample surface was first optimized (Fig. S11†). A low treatment temperature resulted in low photocatalytic activity for O₂ evolution owing to weak interfacial bonding between the CoO_x co-catalyst and LaNb_{1-x}Ta_xON₂. On the other hand, a high treatment temperature above 873 K dramatically decreased the photocatalytic activity for O₂ evolution, which could be ascribed to the severe aggregation of the CoO_x co-catalyst at high temperature. Therefore, the sample that was treated at ~ 773 K showed the highest photocatalytic activity for O₂ evolution owing to its well-integrated interface between the CoO_x co-catalyst and LaNb_{1-x}Ta_xON₂.³⁶ Although the nitrided LaKNaNbO₅ sample without Ta possessed a large number of crystal defects, it showed high photocatalytic activity for O₂ evolution (18.65 mol h⁻¹), which could be ascribed to the exposure of metastable {010} facets on the surface of LaNbON₂ crystals. These facets enabled photogenerated electrons to transfer from the interior of the oxynitride to its surface along the [010] direction, thus improving the charge separation efficiency.¹⁵ The

photocatalytic O₂ evolution activity of nitrided LaKNaNb_{0.8}Ta_{0.2}O₅ (*i.e.*, substituted with 20% Ta⁵⁺) was the highest (22.81 $\mu\text{mol h}^{-1}$). However, higher Ta⁵⁺ contents decreased the photocatalytic O₂ evolution activity of the photocatalysts, even though their defect densities were lower in the formed oxynitride. This could be attributed to the fact that a greater Ta content decreased the transformation of LaKNaNb_{1-x}Ta_xO₅ into LaNb_{1-x}Ta_xON₂. This phenomenon was confirmed using the XRD results shown in Fig. 2b.

Overall, the nitrided LaKNaNb_{1-x}Ta_xO₅ samples showed photocatalytic activity for both H₂ and O₂ evolution. This activity can also be finely tuned by varying the Ta⁵⁺ substitution content, which indicates that these oxynitrides have a bandgap straddling the H₂ and O₂ evolution potentials and thus show promise for use in the water splitting reactions (Fig. 7c). Furthermore, electrochemical impedance spectra (EIS) were recorded to further confirm the charge separation ability of these nitrided LaKNaNb_{1-x}Ta_xO₅ samples. As illustrated in Fig. 7d, nitrided LaKNaTaO₅ exhibited a smaller arc radius than nitrided LaKNaNbO₅, demonstrating a smaller resistance to charge transfer, which could be related to the lower amount of crystal defects in the former.^{37,38} Interestingly, pure nitrided LaKNaNbO₅ and LaKNaTaO₅ both exhibited a lower resistance to charge transfer than the nitrided LaKNaNb_{1-x}Ta_xO₅ samples, suggesting that the Ta substitution distorted the lattice in the LaKNaNb_{1-x}Ta_xO₅ precursor. Moreover, regarding the stability performance of the nitrided LaKNaNb_{1-x}Ta_xO₅ samples, we carried out XPS analysis on nitrided LaKNaNb_{0.8}Ta_{0.2}O₅ before and after the photocatalytic H₂ production reaction, which shows no obvious difference between fresh and used catalysts (Fig. S12 and S13†). The detailed peak contents are provided in Table S6.† The ratio of N 1s/Nb 3d decreased slightly from 0.33 to 0.28, and the ratio of N 1s/Ta 4f decreased from 0.86 to 0.80. This result indicates a small decrease in the amount of N on the surface of the sample, which suggests that the nitrided LaKNaNb_{1-x}Ta_xO₅ samples remained stable in this photocatalytic system. These findings provide the basis for further studies on optimizing the nitridation temperature and heat treatment time of oxide precursors to improve the photocatalytic activity of oxynitrides for H₂ and O₂ evolution.

Conclusion

We investigated the nitridation of layered LaKNaNb_{1-x}Ta_xO₅ oxides and its effect on both crystal defects and the bandgap energy. During nitridation, the volatilization of K and Na species promoted the transformation of the surface of LaKNaNb_{1-x}Ta_xO₅ crystals into lattice-matched oxynitrides, thus forming a core-shell structure. Moreover, introducing Ta species decreased the defect densities, yielding Nb-based oxynitrides with a tunable bandgap from 1.77 to 2.12 eV, which straddled the H₂ and O₂ evolution potentials. As a result, these oxynitrides exhibited good photocatalytic activities for H₂ (19.37 $\mu\text{mol h}^{-1}$) and O₂ (22.81 $\mu\text{mol h}^{-1}$) evolution, which utilized visible light wavelengths between 650 and 750 nm. This work provides a novel strategy for preparing oxynitrides with a tunable bandgap and low defect densities, thus uncovering



the prospects of Nb-based oxynitrides for the overall water splitting reaction.

Experimental section

Materials

La₂O₃, Ta₂O₅, and Nb₂O₅ precursors and NaOH and KOH fluxing agents were all obtained from Macklin. Co(NO₃)₂·6H₂O and RhCl₃·3H₂O co-catalyst precursors were purchased from Sigma-Aldrich.

Synthesis of LaKNaNb_{1-x}Ta_xO₅. 2D layered LaKNaNb_{1-x}Ta_xO₅ oxide precursors were synthesized using the NaOH/KOH flux method, as reported previously.^{26,27} In this synthesis, Ta₂O₅ and Nb₂O₅ (5 mmol in total) at a certain molar ratio (0 : 5, 1 : 4, 2 : 3, 3 : 2, 4 : 1, and 5 : 0) were mixed with La₂O₃ (5 mmol) and finely ground in a mortar. Subsequently, the mixture was added to an alumina crucible containing excess KOH (10 g) and NaOH (5 g). The crucible was then placed in a furnace heated to 873 K at a rate of 10 K min⁻¹. This temperature was maintained for 3 h and then held at 773 K for 15 h. The furnace was subsequently allowed to cool naturally to room temperature. Afterward, the product was washed with ultrapure water at 343 K for 2 h and centrifuged twice to remove any residual KOH and NaOH. The powdered product was completely dried by heating at 313 K under vacuum overnight.

Synthesis of Nb-based oxynitrides. The as-prepared LaKNaNb_{1-x}Ta_xO₅ was subjected to a mild nitridation process in a tube furnace under an NH₃ atmosphere. For each sample, 0.4 g of LaKNaNb_{1-x}Ta_xO₅ was transferred into an alumina tube and then nitrided at various temperatures ranging from 1123 K for 4 h under a flow of gaseous NH₃ at 200 mL min⁻¹. The powder was then collected for further use.

Deposition of Co-catalysts. To load the photocatalyst with the Rh co-catalyst, 0.3 g of the powder was dispersed in a specific volume of an aqueous RhCl₃·3H₂O solution. The solution was completely evaporated by heating in a water bath at 353 K under stirring, after which the powder was collected and reduced at 473 K for 1 h under a flow of 5% H₂/N₂ (200 mL min⁻¹). To load the photocatalyst with the CoO_x co-catalyst, 0.3 g of the powder was added to 1.5 mL of an aqueous Co(NO₃)₂·6H₂O solution (1 mg mL⁻¹) and dispersed using sonication for 5 min. The solution was evaporated by heating in a water bath at 353 K, after which the powder was collected and heated at 773 K for 1 h under a flow of NH₃ (200 mL min⁻¹).

Characterization

XRD patterns of the prepared samples were recorded using a MiniFlex 300, Rigaku, using Cu- K α radiation at 40 kV and 15 mA over a 2 θ range of 10° to 60° with a step size of 0.017° and counting time of 80 s per step. SEM was performed with a JEOL JSM-7600F microscope and SEM-EDX mappings were acquired by employing a FESEM TESCAN S9000G. TEM and HRTEM images were obtained using an FEI Tecnai G2 F30 Spirit microscope operated at an accelerating voltage of 300 kV. UV-vis DRS (V-670, JASCO) was conducted over the range of 200–1000 nm using BaSO₄ as the reflectance standard. XPS was

conducted using a Thermo Scientific ESCALAB 250Xi spectrometer and the decumulation of the peaks was carried out using the software of Avantage. EIS measurements were performed on an electrochemical analyser (CHI660C instruments) in a standard three-electrode system utilizing the synthesized samples as the working electrodes, Ag/AgCl (saturated KCl) as a reference electrode, and a Pt wire as the counter electrode. 0.2 M Na₂S and 0.04 M Na₂SO₃ mixed aqueous solution was used as the electrolyte.

Photocatalytic reactions

The photocatalytic reactions were performed in a top-irradiation-type reaction vessel with a Pyrex window using a Perfect 6A reaction system. For both H₂ and O₂ evolution, 0.3 g of the photocatalyst was dispersed in 150 mL of a 20 vol% aqueous methanol solution. Before photoirradiation, the reaction system was evacuated to remove all air and then irradiated from the top using a 300 W xenon lamp equipped with a dichroic mirror and a cut-off filter ($\lambda \geq 420$ nm). The reactant solution was maintained at 288 K using a cooling-water system during the reaction. The reaction was carried out for 5 h, and the gas production was analyzed every 30 min using an online analysis system with a gas chromatograph (GC-9790II) equipped with a thermal conductivity detector with N₂ as the carrier gas.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 Y. Qi, J. Zhang, Y. Kong, Y. Zhao, S. Chen, D. Li, W. Liu, Y. Chen, T. Xie, J. Cui, K. Domen and F. Zhang, *Nat. Commun.*, 2022, **13**, 484.
- 2 T. Takata, J. Jiang, Y. Sakata, M. Nakabayashi, N. Shibata, V. Nandal, K. Seki, T. Hisatomi and K. Domen, *Nature*, 2020, **581**, 411–414.
- 3 Y. H. Hong, Y.-M. Lee, W. Nam and S. Fukuzumi, *J. Am. Chem. Soc.*, 2022, **144**(2), 695–700.
- 4 K. Sivula and R. van de Krol, *Nat. Rev. Mater.*, 2016, **1**, 15010.
- 5 Y. Tachibana, I. Vayssieres and J. R. Durrant, *Nat. Photonics*, 2012, **6**, 511–518.
- 6 Q. Wang, M. Nakabayashi, T. Hisatomi, S. Sun, S. Akiyama, Z. Wang, Z. Pan, X. Xiao, T. Watanabe, T. Yamada, N. Shibata, T. Takata and K. Domen, *Nat. Mater.*, 2019, **18**, 827–832.
- 7 C. Pan, T. Takata, M. Nakabayashi, T. Matsumoto, N. Shibata, Y. Ikuhara and K. Domen, *Angew. Chem., Int. Ed.*, 2015, **54**, 2955–2959.



- 8 B. Tian, B. Tian, B. Smith, M. C. Scott, R. Hua, Q. Lei and Y. Tian, *Nat. Commun.*, 2018, **9**, 1397.
- 9 K. Maeda, Y. Tokunaga, K. Hibino, K. Fujii, H. Nakaki, T. Uchiyama, M. Eguchi, D. Lu, S. Ida, Y. Uchimoto and M. Yashima, *ACS Appl. Energy Mater.*, 2018, **1**, 1734–1741.
- 10 Y. Tang, K. Kato, T. Oshima, H. Mogi, A. Miyoshi, K. Fujii, K. Yanagisawa, K. Kimoto, A. Yamakata, M. Yashima and K. Maeda, *Inorg. Chem.*, 2020, **59**, 11122–11128.
- 11 B. Siritanaratkul, K. Maeda, T. Hisatomi and K. Domen, *ChemSusChem*, 2011, **4**, 74–78.
- 12 B. Siritanaratkul, K. Maeda, T. Hisatomi and K. Domen, *ChemSusChem*, 2011, **4**, 74–78.
- 13 M. Kodera, Y. Moriya, M. Katayama, T. Hisatomi, T. Minegishi and K. Domen, *Sci. Rep.*, 2018, **8**, 15849.
- 14 T. Hisatomi, C. Katayama, Y. Moriya, T. Minegishi, M. Katayama, H. Nishiyama, T. Yamada and K. Domen, *Energy Environ. Sci.*, 2013, **6**, 3595–3599.
- 15 X. Wang, T. Hisatomi, J. Liang, Z. Wang, Y. Xiang, Y. Zhao, X. Dai, T. Takata and K. Domen, *J. Mater. Chem. A*, 2020, **8**, 11743–11751.
- 16 S. Ida, Y. Okamoto, M. Matsuka, H. Hagiwara and T. Ishihara, *J. Am. Chem. Soc.*, 2012, **134**, 15773–15782.
- 17 K. Maeda, Y. Tokunaga, K. Hibino, K. Fujii, H. Nakaki, T. Uchiyama, M. Eguchi, D. Lu, S. Ida, Y. Uchimoto and M. Yashima, *ACS Appl. Energy Mater.*, 2018, **1**, 1734–1741.
- 18 J. Seo, T. Hisatomi, M. Nakabayashi, N. Shibata, T. Minegishi, M. Katayama and K. Domen, *Adv. Energy Mater.*, 2018, **8**, 1800094.
- 19 T. Hryniewicz, K. Rokosz and H. R. Z. Sandim, *Appl. Surf. Sci.*, 2012, **263**, 357–361.
- 20 Y.-I. Kim, P. M. Woodward, K. Z. Baba-Kishi and C. W. Tai, *Chem. Mater.*, 2004, **16**, 1267–1276.
- 21 Y. Tang, K. Kato, T. Oshima, H. Mogi, A. Miyoshi, K. Fujii, K. Yanagisawa, K. Kimoto, A. Yamakata, M. Yashima and K. Maeda, *Inorg. Chem.*, 2020, **59**, 11122–11128.
- 22 K. Momma and F. Izumi, *J. Appl. Crystallogr.*, 2011, **44**, 1272.
- 23 I. P. Roof, T.-C. Jagau, W. G. Zeier, M. D. Smith and H.-C. Loye, *Chem. Mater.*, 2009, **21**, 1955–1961.
- 24 J.-H. Liao and M.-C. Tsai, *Cryst. Growth Des.*, 2002, **2**, 83–85.
- 25 L. Wan, F.-Q. Xiong, B. Zhang, R. Che, Y. Li and M. Yang, *J. Energy Chem.*, 2018, **27**, 367–371.
- 26 Y. Du, L. Yuan, K. Huang and S. Feng, *J. Lumin.*, 2013, **135**, 196–200.
- 27 K. Korzeniewski and M. Sobczyk, *Solid State Commun.*, 2018, **273**, 30–33.
- 28 V. R. Reddy, D. W. Hwang and J. S. Lee, *Catal. Lett.*, 2003, **90**, 39–43.
- 29 M. Hojamberdiev, E. Zahedi, E. Nurlaela, K. Kawashima, K. Yubuta, M. Nakayama, H. Wagata, T. Minegishi, K. Domen and K. Teshima, *J. Mater. Chem. A*, 2016, **4**, 12807–12817.
- 30 X. Wang, T. Hisatomi, Z. Wang, J. Song, J. Qu, T. Takata and K. Domen, *Angew. Chem., Int. Ed.*, 2019, **58**, 10666–10670.
- 31 J. Seo, S. Jeong and S. Kim, *ACS Appl. Energy Mater.*, 2021, **4**, 3141–3150.
- 32 V. Khanal, N. O. Balayeva, C. Günnemann, Z. Mamiyev, R. Dillert, D. W. Bahnemann and V. (Ravi) Subramanian, *Appl. Catal., B*, 2021, **291**, 119974.
- 33 Y. Okamoto, S. Ida, J. Hyodo, H. Hagiwara and T. Ishihara, *J. Am. Chem. Soc.*, 2011, **133**, 18034.
- 34 H. Zhang, S. Wei and X. Xu, *J. Catal.*, 2020, **383**, 135–143.
- 35 G. R. Dillip, A. N. Banerjee, V. C. Anitha, B. D. P. Raju, S. W. Joo and B. K. Min, *ACS Appl. Mater. Interfaces*, 2016, **8**, 5025–5039.
- 36 S. Chen, S. Shen, G. Liu, Y. Qi, F. Zhang and C. Li, *Angew. Chem., Int. Ed.*, 2015, **54**, 3047–3051.
- 37 B. Dong, J. Cui, T. Liu, Y. Gao, Y. Qi, D. Li, F. Xiong, F. Zhang and C. Li, *Adv. Energy Mater.*, 2018, **8**, 1801660.
- 38 C. Peng, T. Zhou, P. Wei, H. Ai, B. Zhou, H. Pan, W. Xu, J. Jia, K. Zhang, H. Wang and H. Yu, *Chem. Eng. J.*, 2022, **439**, 135685.

