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Oxygen vacancy assisted hydrogen evolution reaction over CeO₂-based solid solutions

Saraswati Roy ^a and Sounak Roy ^{*ab}

Producing sustainable hydrogen through water electrolysis is a promising approach to meet the growing demand for renewable energy storage. Developing affordable and efficient electrocatalysts made from non-precious metals to replace platinum-based catalysts for hydrogen evolution reactions (HERs) continues to be a significant challenge. In the present study, pristine CeO₂ and doped solid solutions Ce_{0.95}Co_{0.05}O₂, Ce_{0.95}Ni_{0.05}O₂ and Ce_{0.95}Cu_{0.05}O₂ were evaluated for the HER, and the electrochemical studies in alkaline medium showed superior HER activity for the doped catalysts, with Ce_{0.95}Ni_{0.05}O₂ achieving the lowest overpotential and highest mass activity, comparable to Pt/C. Tafel slopes and EIS measurements suggested a Volmer–Heyrovsky mechanism facilitated by oxygen vacancies and hydrogen spill-over. Stability tests confirmed the durability of Ce_{0.95}Ni_{0.05}O₂ under prolonged HER conditions. This study highlights aliovalent doping as a viable strategy for engineering oxygen vacancies and enhancing CeO₂-based catalysts for alkaline water electrolysis.

1. Introduction

The continued reliance on fossil fuels not only depletes these finite resources but also contributes to a host of environmental problems. Hydrogen stands out as a clean and efficient energy carrier, representing a significant step forward in building a hydrogen-driven economy. However, it's important to recognize that although using hydrogen fuel does not directly produce CO₂ emissions, many current methods of hydrogen production do contribute to CO₂ emissions. In this regard, water electrolysis with the help of renewable electricity has significant potential to produce affordable, sustainable green hydrogen on a large scale. After its first discovery in 1789, the HER has been one of the most extensively studied processes in electrocatalysis, primarily owing to its straightforward mechanism involving two consecutive proton–electron transfers without side reactions.^{1–5} Although Pt group metals and their alloys are some of the most efficient electrocatalysts for the HER, reducing costs is essential to make water electrolysis a practical method for H₂ production. To achieve this, extensive research is currently directed toward developing affordable, stable, and efficient HER catalysts using non-noble metals.^{6–11}

Conventional alkaline electrolyzers typically rely on electrodes made from late transition metals like Co, Ni, and Cu

because of their affordability and outstanding chemical durability.^{12,13} Oxides of Co, Ni, and Cu are considered promising catalysts for the HER due to their low energy barriers for hydrogen adsorption.^{14–16} However, a significant challenge associated with these cathodes is their low conductivity, which hinders efficient electron transfer during the HER process. One common approach to enhancing conductivity is to introduce highly conductive substrates,^{17–20} while another method is to create oxygen vacancies.^{21–23} Enhanced oxygen vacancy concentrations in electrocatalysts can be achieved by introducing highly reducible oxides as a support for the transition metal active sites. The distinctive ability to reversibly switch between Ce³⁺ and Ce⁴⁺ oxidation states, combined with the presence of abundant oxygen vacancy defects and surface oxygen ion exchange renders CeO₂ a promising support material, capable of forming an active metal–support interface with significant electrochemical activity.^{24,25} For example, in a recent study, Ni-based composites over commercial CeO₂ nanoparticles exhibited a 48% higher current density compared to the conventional Ni-Watts catalyst. The authors attributed this improvement to the interaction between the Fermi level of Ni and CeO₂, which induced band bending and modulated the electrocatalytic activity toward H₂ generation. In another study, 2% Gd doped CeO₂ in 1.0 M KOH showed HER with an overpotential of 99 mV and a Tafel slope of 211 mV dec⁻¹.²⁶ Similarly, Ni-doped CeO₂ supported on N-doped carbon nanofibers (Ni/CeO₂@N-CNFs) exhibited an overpotential of 100 mV and a Tafel slope of 85.7 mV dec⁻¹ in 1.0 M KOH alkaline electrolyte.²⁷ Ni–S alloys on the surface of CeO₂ nanorods and CeO₂ nano flex demonstrated HER overpotential of 200 and 180 mV at a current density of 10 mA cm⁻² in 1.0 M NaOH.²⁸

^a Department of Chemistry, Birla Institute of Technology and Science Pilani, Hyderabad Campus, Hyderabad 500078, India.
E-mail: sounak.roy@hyderabad.bits-pilani.ac.in

^b Materials Center for Sustainable Energy & Environment, Birla Institute of Technology and Science Pilani, Hyderabad Campus, Hyderabad 500078, India



In another investigation, Ni/MoO_{2-x}@NF showed a HER overpotential of 27 mV with a Tafel slope of 26 mV dec⁻¹ in 1.0 M KOH.^{29,30} Furthermore, Ce-doped Ni₂P, with a Ce/Ni atomic ratio of 12.3%, showed overpotentials of 77 mV and 42 mV to achieve a current density of 10 mA cm⁻² in 1 M KOH and 0.5 M H₂SO₄ solutions, respectively.³¹ Notably, in most of the reports, the supported metal catalysts consist of the active metal or metal oxide dispersed on the surface of CeO₂.

We have recently carried out a series of studies with CeO₂ as a promising support to the transition metal active site due to its easily available redox of Ce⁴⁺/Ce³⁺, which is responsible for oxygen vacancy enhancement.³²⁻³⁴ In our studies, the active metal sites (Mⁿ⁺) were ionically substituted in the fluorite CeO₂ matrix to produce a solid solution of Ce_{1-x}M_xO₂. The aliovalent doping of ionic active sites enhanced reducibility and generated oxygen vacancies in CeO₂. We have demonstrated that the formation of the solid solution of Ce_{1-x}M_xO₂ immensely facilitated the reduction of Ce⁴⁺ → Ce³⁺ and subsequent oxidation of Mⁿ⁺ → M⁽ⁿ⁺¹⁾⁺ during anodic oxidation reactions. However, in our previous investigations, the focus was primarily on anodic oxidation reactions. In this study, for the first time, we have evaluated the performances of pristine CeO₂, and doped solid solutions of Ce_{0.95}Co_{0.05}O₂, Ce_{0.95}Ni_{0.05}O₂ and Ce_{0.95}Cu_{0.05}O₂ in the HER. The primary focus is to investigate the role of support reducibility and oxygen vacancies in influencing HER kinetics. This study developed a novel, cost-effective, and efficient material for the HER, and also came up with a meaningful insight into the fundamental mechanisms.

2. Materials and methods

2.1. Synthesis and characterization of CeO₂, Ce_{0.95}Co_{0.05}O₂, Ce_{0.95}Ni_{0.05}O₂ and Ce_{0.95}Cu_{0.05}O₂

The pristine and the doped materials were synthesized using the solution combustion method. The solution combustion method is advantageous as it is a low temperature initiated single step self-propagating process with a high transient temperature producing huge amounts of gases. The combustion process is fast (instantaneous) and due to the production of huge amounts of gases, the products are homogenous nanocrystalline materials with the desired composition and structure. This process utilizes the highly exothermic redox chemical reactions between metal nitrates and fuels. The details of the synthesis of CeO₂, Ce_{0.95}Co_{0.05}O₂, Ce_{0.95}Ni_{0.05}O₂ and Ce_{0.95}Cu_{0.05}O₂ are provided in our previous study.³²

The structure and crystallinity of the synthesized mixed metal oxides were analysed through X-ray diffraction (XRD) using a Rigaku Ultima IV instrument equipped with Cu K α radiation. The XRD scans were conducted at a scanning rate of 0.4° per minute, with a step size of 0.01°. The surface morphology and compositional analysis were performed using a field emission scanning electron microscope (FESEM) integrated with energy dispersive X-ray spectroscopy (EDS). An acceleration voltage of 20 kV was used to run the device. An Al K α radiation (1486.6 eV) Thermo Scientific K-Alpha surface

analysis spectrometer was used to capture the synthesised catalysts' X-ray photoelectron spectra (XPS). A nonlinear least-squares curve fitting program with a Gaussian-Lorentzian production function was applied to the data profiles, and Avantage software was used for processing. The standard C 1s peak at 284.85 eV was used to calibrate the binding energy (B.E.) of all XPS data. The absorption and steady-state photoluminescence (PL) spectra were recorded using a Fluoro-Log spectrofluorimeter (Horiba Scientific) and a JASCO UV-Vis spectrophotometer (model no. V-650). A Bruker ESR 5000 was used to perform the EPR study. Raman spectroscopy was performed using UniRam Raman spectroscopy.

2.2. Electrochemical studies

A standard three-electrode system was used to perform electrochemical testing on the synthesised compounds. A synthesised catalyst was applied to carbon cloth to create the working electrode. Pt mesh was employed as the counter electrode, while the reference electrode was Hg/HgO saturated with an aqueous KOH solution. 3 mg of the synthesised catalyst and 0.5 mg of powdered carbon black were mixed with one ml of methanol to create the working electrode. Furthermore, the solution was supplemented with 10 μ l of 5% aqueous Nafion (Sigma-Aldrich). To produce a uniform dispersion of the carbon and the synthesised metal oxides, the mixture was sonicated for 60 minutes. At 6:1.5, the catalyst's active mass ratio to carbon black was kept constant. After chopping into 1 \times 0.4 cm² pieces, the carbon cloth electrode was agitated in 3:1 HNO₃ and H₂SO₄ for three hours at 80 °C. It was then cleaned with distilled water and agitated in distilled water for four hours at 80 °C. The produced ink comprising 50 μ l of the catalyst was drop-cast onto the clean carbon cloth electrode with an active mass loading of 0.15 mg cm⁻² after the carbon cloth had been dried for the entire night. At room temperature, the electrode was allowed to dry. Overall, electrocatalytic water splitting was performed using an Origa Flex OGF500 electrochemical workstation. Electrochemical impedance spectroscopy (EIS), linear sweep voltammetry (LSV), and cyclic voltammetry (CV) were measured at a scan rate of 10 mV s⁻¹ in a simple electrolyte medium of 1 M KOH. Between 0 and -0.6 V (*versus* RHE), the CV cycles and the HER polarization curves were performed using LSV traces. The following formula was used to convert the acquired potentials to the reversible hydrogen electrode (RHE): $E_{\text{RHE}} = E_{\text{Hg/HgO}} + 0.059 \text{ pH} + E_{\text{Hg/HgO}}^0$. In this case, $E_{\text{Hg/HgO}}$ is the experimental potential with regard to the Hg/HgO electrode, $E_{\text{Hg/HgO}}^0$ is the standard reduction potential of the Hg/HgO electrode, and E_{RHE} is the final converted potential with regard to RHE. For the water splitting reaction, a two-electrode arrangement was chosen. Chronoamperometric (CA) measurements were carried out in a three-electrode system for 24 hours in 1 M KOH in order to assess material stability. EIS was performed using overpotential *vs.* Hg/HgO in the frequency range of 100 kHz to 10 mHz. A portable gas chromatograph (GC) from Mayura Analytical Pvt. Limited in India was used to measure the gaseous product H₂.



3. Results and discussion

3.1. Structure and oxygen vacancies

The single-step solution combustion approach produced nanocrystalline and porous pure CeO_2 and the solid solutions $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$, and $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$ with uniform dopant distribution.^{35–38} The development of the nano-sized porous particles is confirmed by the FE-SEM micrographs of pristine CeO_2 and the doped solid solutions $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$, and $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$ shown in Fig. 1. The particle size distribution was within around 100 nm. A constant doping concentration with the theoretically computed values was

validated by the elemental atomic percentages obtained from EDX examination.

All of the synthesized materials' powder XRD patterns are shown in Fig. 2(a). The materials crystallized in a phase-pure cubic fluorite structure with the space group $Fm\bar{3}m$. No extra diffraction peaks corresponding to secondary phases like Co_3O_4 , NiO , or CuO were observed in the doped oxides $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$, and $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$. This absence of extraneous peaks confirms the successful formation of phase-pure solid solutions. The synthesized materials displayed relatively broad diffraction peaks, indicative of their nanocrystalline nature, which is attributed to the specific characteristics of the solution



Fig. 1 FE-SEM and EDX of CeO_2 and doped $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$, and $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$.



Fig. 2 (a) XRD patterns, (b) zoomed in version, (c) the refined cell volume of CeO_2 and doped $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$, and $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$.



combustion synthesis. Interestingly, upon doping in CeO_2 , a noticeable shift of the diffraction peaks towards higher angles was observed (Fig. 2(b)), indicating a reduction in the unit cell volume. The Shannon ionic radius of eight-coordinated Ce^{4+} in the fluorite lattice is 0.97 Å, whereas the ionic radii of divalent Co, Ni, and Cu range from 0.65 to 0.75 Å. The diffraction peaks may shift to higher angles and the unit cell volume may decrease when Ce^{4+} is substituted with smaller cations like Co^{2+} , Ni^{2+} , and Cu^{2+} . The experimentally obtained diffraction patterns were subjected to Rietveld refinement in order to validate these structural alterations. During the refinement, Co, Ni, and Cu cations were modelled at the 4a Wyckoff sites typically occupied by Ce. Details are provided elsewhere.³² The cell volumes derived from the refinement, plotted in Fig. 2(c), further confirm the unit cell contraction, corroborating the formation of substitutional solid solutions.

We subsequently looked at the surface characteristics of the produced materials. In the N_2 adsorption–desorption isotherms, both the pristine and doped solid solutions displayed a type II isotherm with a small hysteresis, indicating the surface's

nanoporous nature. The average specific surface area, as determined by the BET equation, was around $20 \text{ m}^2 \text{ g}^{-1}$ (Fig. 3(a)). Fig. 3(b) and (c) show the non-faradaic region CV and C_{dl} plots, respectively. The typical ECSA values were calculated from the observed C_{dl} values. Fig. 3(d) shows the 3-fold increased ECSA of $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ compared to pristine CeO_2 .

Doping with aliovalent substitution would cause extrinsic oxygen vacancies to emerge in the CeO_2 matrix. O 1s core-level XPS data were collected in order to conduct a detailed analysis of the oxygen vacancies (Fig. 4(a)). In the O 1s XPS spectra, the peak at 529.2 eV corresponds to lattice oxygen (O_L), while the peak at 531.7 eV is attributed to adsorbed oxygen species associated with surface defect sites, particularly oxygen vacancies (O_V). CeO_2 , $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, and $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ were found to have surface oxygen vacancy concentrations of 10%, 17%, 38%, and 57%, respectively. The highest concentration of surface oxygen vacancies was observed in $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$. Under 310 nm excitation, PL investigations were conducted to further evaluate the existence of oxygen vacancies. Fig. 4(b) depicts the highest intensity of the PL spectrum for pristine CeO_2 . The radiative



Fig. 3 (a) N_2 adsorption–desorption isotherms of the synthesized catalysts; (b) CV curves of $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ in the non-faradaic region (-0.05 to 0.1 V vs. RHE) at scan rates of 10, 20, 30, 40, 50 mV s^{-1} ; (c) linear regression between the current density differences in the 0.02 V potential window of CV vs. scan rate and (d) ECSA of all four catalysts.





Fig. 4 (a) O1s core level XPS spectra; (b) PL; (c) Raman and (d) EPR spectra of CeO₂ and doped Ce_{0.95}Co_{0.05}O₂, Ce_{0.95}Ni_{0.05}O₂ and Ce_{0.95}Cu_{0.05}O₂.

recombination of holes trapped in reduced Ce⁴⁺ states with electrons located in intrinsic oxygen vacancies (F-centres) produces the strong PL emission band seen in pristine CeO₂.^{39,40} Upon doping, a notable reduction in PL intensity was observed, with Ce_{0.95}Ni_{0.05}O₂ exhibiting the lowest intensity. The extrinsic oxygen vacancies generated due to the aliovalent doping can form intermediate electronic states within the bandgap.⁴¹ These states may facilitate non-radiative recombination pathways, thereby suppressing radiative electron-hole recombination and reducing PL intensity. The PL studies verified that the doped solid solutions have more oxygen vacancies than pure CeO₂, with Ce_{0.95}Ni_{0.05}O₂ having the highest concentration of oxygen vacancies. Additionally, Raman spectroscopy was used to verify that oxygen vacancies were present. Fig. 4(c) shows a Gaussian peak at around 558 cm⁻¹ attributed to the A_{1g} mode, which is suggestive of oxygen vacancies, and a distinctive Lorentzian peak at about 448 cm⁻¹, which corresponds to the F_{2g} vibrational mode of the CeO₂ lattice.⁴² The largest oxygen vacancy concentration in Ce_{0.95}Ni_{0.05}O₂ was further supported by the maximum intensity of the A_{1g} mode over Ce_{0.95}Ni_{0.05}O₂. To look into the existence of oxygen vacancies further, EPR measurements were also carried

out. In agreement with the signal ascribed to unpaired electrons close to oxygen vacancies, the *g*-value was found to be approximately 2. The peak intensity at about 336 mT systematically increased with doping, as seen in Fig. 4(d). This enhancement in intensity is attributed to the presence of surface oxygen vacancies.⁴³ Electrons originally localized in the O 2p orbitals adjacent to these vacancies become delocalized, contributing to the observed EPR signal. Comprehensive analyses using XPS, PL, Raman, and EPR conclusively demonstrated an enhancement in oxygen vacancies resulting from substitutional doping, with Ce_{0.95}Ni_{0.05}O₂ exhibiting the highest concentration of oxygen vacancies.

3.2. HER

LSV polarization curves in a 1.0 M KOH solution utilizing a three-electrode setup operating between 0 and -0.6 V (vs. RHE) were used to assess the cathodic HER activity of the produced materials. The alkaline electrolyte was chosen herein as alkaline water electrolysis is a highly significant technology that is widely used in the industry, especially for controlling alkaline effluents from processes like water-alkali and chlor-alkali





Fig. 5 (a) HER polarisation curve of the synthesized four catalysts with commercial Pt/C; (b) overpotentials; (c) intrinsic mass activity of the synthesized catalysts, and (d) intrinsic activity of the synthesized catalyst.

electrolyzers.⁴⁴ As illustrated in Fig. 5(a), the solid solution catalysts demonstrated significantly higher HER activity compared to pristine CeO₂, with Ce_{0.95}Ni_{0.05}O₂ exhibiting the lowest overpotential among all synthesized catalysts. A commercial 5% Pt/C sample was also examined under the same circumstances as the benchmark. Only a limited number of HER electrocatalysts have been reported with commercial Pt/C in alkaline electrolyte in the literature compared to that of acidic medium. The overpotential of Ce_{0.95}Ni_{0.05}O₂ was remarkably close to that of 5% Pt/C. The required overpotential to reach a current density of 10 mA cm⁻² is a commonly accepted benchmark standard for the production of solar fuel. The overpotentials for the synthesized materials at current densities of 10 and 50 mA cm⁻² are shown in Fig. 5(b). The catalysts mass activity derived by normalizing the current density to the amount of catalyst was also ascertained to aid in global comparisons (Fig. 5(c)). The overpotentials as well as the mass activity data show that Ce_{0.95}Ni_{0.05}O₂ as a significantly better catalyst than the pure and other doped oxides. Intrinsic activity was also determined for the catalysts (Fig. 5(d)). Intrinsic activity assessments help identify catalysts capable of delivering high current densities, which is critical for efficient and scalable hydrogen

production.⁴⁵ The intrinsic activity analysis revealed that Ce_{0.95}Ni_{0.05}O₂ exhibits superior reactivity compared to the other catalysts studied (Fig. 5(d)).

The electrochemical stability of the most active catalyst, Ce_{0.95}Ni_{0.05}O₂ was assessed *via* chronoamperometry in 1.0 M KOH for 48 hours at its corresponding HER overpotential and -0.25 V. As shown in Fig. 6(a), the catalyst maintained a remarkably stable current density. A stepwise CA experiment was conducted over 30 hours using six different potentials, each applied for a 5-hour interval (Fig. 6(b)). The results in Fig. 6(b) demonstrate that the Ce_{0.95}Ni_{0.05}O₂ catalyst remains highly stable under both low and high applied potentials. The study reveals a sharp change in the HER potential upon switching the current density, indicating efficient charge and mass transport facilitated by effective electrolyte diffusion within the porous and vertically aligned electrode structure. Furthermore, the electrodes maintain a dimensionally stable architecture, which is advantageous for sustained high-rate hydrogen production.⁴⁶

The superior HER performance of Ce_{0.95}Ni_{0.05}O₂ was further validated by the amount of H₂ evolved during the reaction, measured using gas chromatography. Representative chromatograms



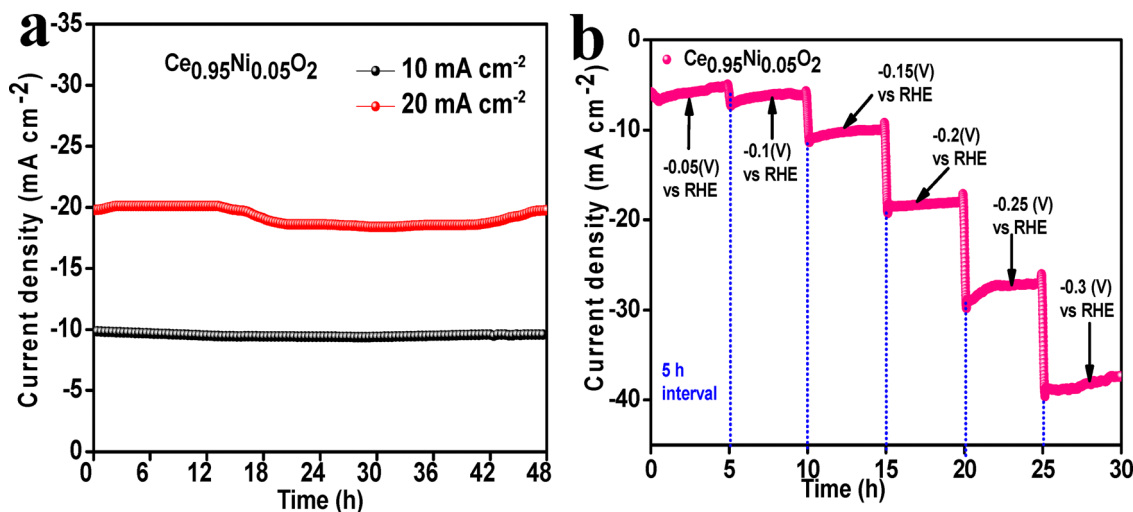


Fig. 6 (a) Stability studies with Ce_{0.95}Ni_{0.05}O₂ at 10 and 20 mA cm⁻² and (b) stepwise CA study of Ce_{0.95}Ni_{0.05}O₂.

over Ce_{0.95}Ni_{0.05}O₂ in different potentials are provided in Fig. 7(a). The faradaic efficiency for H₂ evolution (FE_{H₂}%) was calculated and is presented in Fig. 7(b). Notably, Ce_{0.95}Ni_{0.05}O₂ exhibited an impressive FE_{H₂}% of approximately 88%, surpassing the values obtained for the other pristine and doped catalysts. This high faradaic efficiency indicates that a significant proportion of the electrons supplied during the reaction were effectively utilized for hydrogen production, highlighting the material's exceptional selectivity and efficiency for the HER. The relationship of HER activity on scan rate and activation energy was thoroughly assessed in order to clarify the kinetics of the HER for the produced catalysts. The scan rate dependency tests were carried out by altering the scan rate from 10 to 50 mV s⁻¹ in the presence of 1.0 M KOH electrolyte. The CV at different scan rates of Ce_{0.95}Ni_{0.05}O₂ are shown in Fig. 7(c). A diffusion-controlled electrocatalytic process on the catalysts was suggested as the HER current density increased linearly with the square root of the scan rate in Fig. 7(d). The activation energy was calculated using an Arrhenius plot of current density (ln*j*) vs. temperature (*T*⁻¹) at the HER potential of the four catalysts. The LSV traces of Ce_{0.95}Ni_{0.05}O₂ at different temperatures are shown in Fig. 7(e). Fig. 7(f) shows the activation energy of different catalysts. While Cu, Co, and Ni doped CeO₂ showed activation energies of 12.97, 10.26, and 9.27 kJ mol⁻¹, respectively, pristine CeO₂ had an activation energy of 13.36 kJ mol⁻¹. The fastest HER kinetics over Ce_{0.95}Ni_{0.05}O_{2-δ} were indicated by the lowest activation energy.

3.3. Mechanism and surface-reactivity relationship

It is well-established that water electrolysis follows different mechanisms depending on the electrolyte. For instance, the abundance of protons in acidic electrolytes facilitates HER catalysis, whereas achieving the HER at low overpotentials is comparatively more challenging in alkaline media.⁴⁷ In alkaline solutions, additional energy is required to generate protons through water molecule dissociation, which subsequently affects the overall reaction rate. The evolution of electrocatalytic

hydrogen in alkaline medium is a multi-step process. The Volmer step is the initial step in an alkaline medium (eqn (1)):



Subsequent to this step, the HER pathway in alkaline medium follows the Volmer–Heyrovsky path (eqn (2)) or Volmer–Tafel path (eqn (3)) to produce H₂.



All the above listed reaction phenomenon strongly depend on the inherent surface chemistry leading to charge transfer capabilities of the catalyst. The charge transfer capabilities *vis à vis* Tafel slopes determined from the LSV curves reveal the catalytic kinetics information of the HER. For the HER, a theoretical Tafel slope of 120 mV dec⁻¹ corresponds to the Volmer step, whereas 40 and 30 mV dec⁻¹ represent the Volmer–Heyrovsky path and Volmer–Tafel path, respectively.⁴⁸ Therefore, the Tafel slopes were calculated over the combustion synthesized solid solutions along with the commercial Pt/C catalyst (Fig. 8(a)). It is conventionally believed that Pt-based electrocatalysts go through the Volmer–Tafel path at least in low-potential regions because of the optimal Gibbs free energy of Pt–H* ($\Delta G_{\text{H}^*} = 0$).²² The calculated Tafel slope in our experimental conditions over commercial Pt/C was found to be 35.6 mV dec⁻¹ indicating the Volmer–Tafel path in accordance with the literature. However, the Tafel slope calculated over doped Ce_{0.95}Ni_{0.05}O₂ was 48 mV dec⁻¹ suggesting the Volmer–Heyrovsky mechanistic path of the HER. Ce_{0.95}Ni_{0.05}O₂ showed the lowest Tafel slope among the doped solid solutions, and it was reduced by 28% compared with that of pristine CeO₂. An effective electrocatalyst with significant charge transfer capabilities should exhibit a low Tafel slope. Ni in Ce_{0.95}Ni_{0.05}O₂ is believed to initially adsorb the H* effectively. The substantial presence of oxygen vacancies in Ce_{0.95}Ni_{0.05}O₂, as established from the detailed spectroscopic studies, facilitates the migration of H* from the Ni site to the



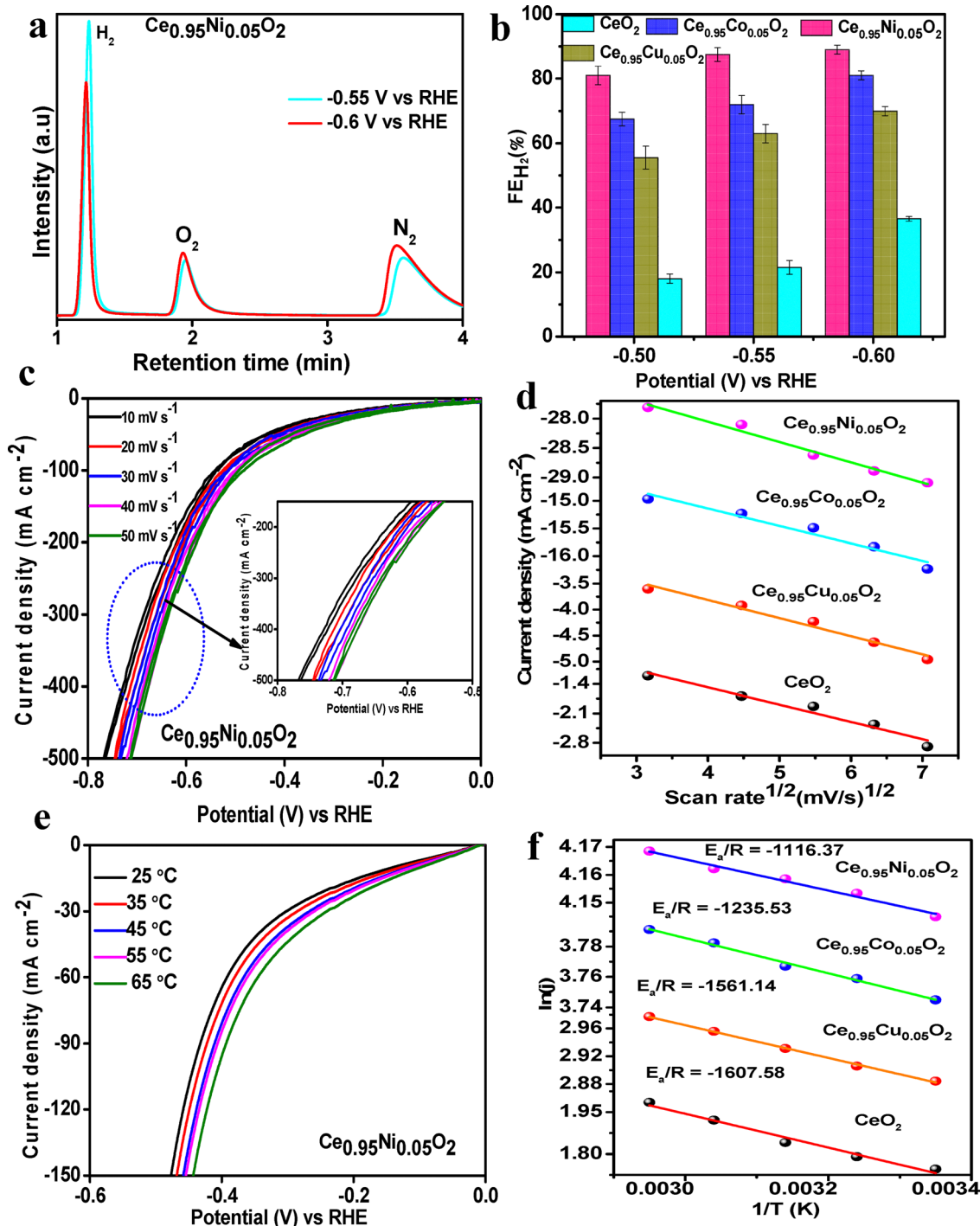


Fig. 7 (a) Representative GC chromatograms of H_2 from $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ after 1 h of chronoamperometric study; (b) FE_{H_2} (%) of the four catalysts at varied potentials; (c) CV of $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ with different scan rates; (d) scan rate-dependent study of the four catalysts; (e) LSV of $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ at various temperatures and (f) activation energy plot of the hydrogen evolution reaction over the four synthesized materials.

neighbouring vacancy, a well-known hydrogen spill-over phenomenon (eqn (4)).⁴⁹ The spilled over H^* reacts through the Volmer–Heyrovsky mechanistic path with another water molecule in reducing cathodic potential to produce H_2 (eqn (5)).



The Tafel slopes obtained for the other synthesized catalysts suggest that this specific mechanistic pathway occurs more sluggishly over other solid-solution catalysts in comparison to



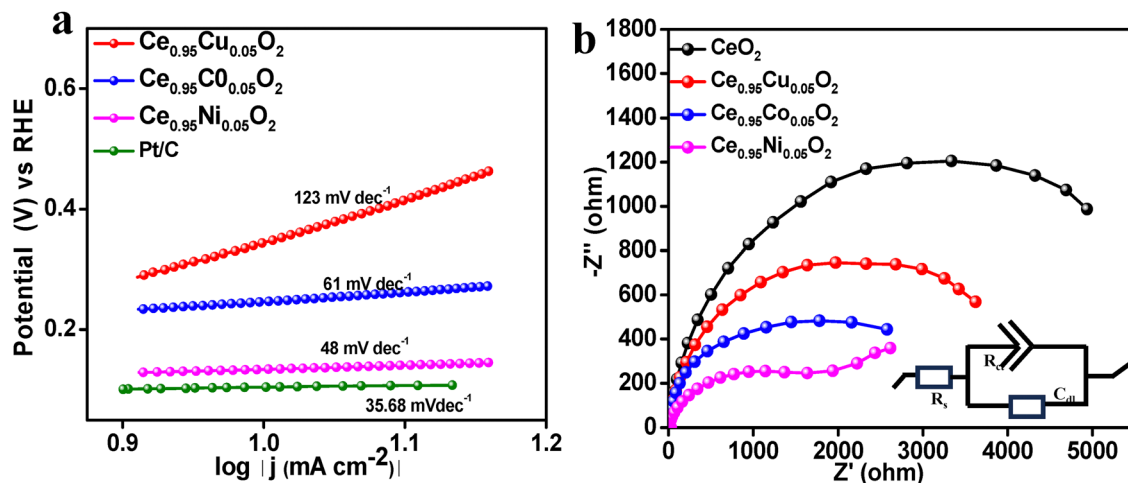


Fig. 8 (a) Tafel plots and (b) EIS studies of the synthesized catalysts.

Ce_{0.95}Ni_{0.05}O₂. The highest Tafel slope observed for CeO₂ indicates the slowest HER kinetics for the pristine material. The reduced Tafel slope over Ce_{0.95}Ni_{0.05}O₂ also signifies the enhanced charge transfer between the electrode and electrolyte. Therefore, the EIS studies were conducted under the HER conditions and the data are shown in Fig. 8(b). The fitted circuit is provided in the inset of Fig. 8(b). Ce_{0.95}Ni_{0.05}O₂ showed the lowest semi-circle suggesting the fastest shuttling of charges corroborating the superiority of the material. The significant presence of oxygen vacancies in Ce_{0.95}Ni_{0.05}O₂ likely accounts for its low charge transfer resistance. The R_{ct} values were found to be 4.9×10^3 , 3.6×10^3 , 2.5×10^3 and $1.3 \times 10^3 \Omega$ for CeO₂, Ce_{0.95}Cu_{0.05}O₂, Ce_{0.95}Co_{0.05}O₂ and Ce_{0.95}Ni_{0.05}O₂, respectively.

The above Volmer–Heyrovsky mechanistic path mediated through the oxygen vacancy assisted H spill-over phenomenon should result in reduction of active sites of ionically doped transition metals as well as the support CeO₂. Therefore, a thorough XPS study was conducted to probe the surface elemental composition and oxidation states of the freshly prepared and exhausted catalysts. Fig. 9(a) represents the Ce 3d core level deconvoluted spectra (3d_{5/2} is labelled as v, and 3d_{3/2} is labelled as u) from the as-synthesized catalysts. The spectra exhibited a total of 5 spin orbit coupled peaks, in which v–u, v'–u'', and v'''–u'''' belong to Ce⁴⁺, and v⁰–u⁰ and v'–u' belong to Ce³⁺. The Ce⁴⁺ doublets designated as v–u (882.6–901.3 eV), v'–u'' (889.2–907.6 eV) and v'''–u'''' (897.6–917.0 eV) are associated with Ce 3d⁹4f²O2p⁴, Ce 3d⁹4f¹O2p⁵ and Ce 3d⁹4f⁰O2p⁶ final states, respectively.^{50–52} The other two doublets of Ce³⁺ positioned at v⁰–u⁰ (880–898.5 eV) and v'–u' (885.4–903.5) correspond to the Ce 3d⁹4f²O2p⁵ and Ce 3d⁹4f¹O2p⁶ final states. The deconvoluted Ce 3d spectra indicated the coexistence of Ce in both the 4+ and 3+ oxidation states in all four as-prepared catalysts. The amount of Ce⁴⁺ present in CeO₂ and doped Ce_{0.95}Cu_{0.05}O₂, Ce_{0.95}Co_{0.05}O₂ and Ce_{0.95}Ni_{0.05}O₂ was 88, 80, 70 and 66%, respectively. The Cu 2p core level spectra of Ce_{0.95}Cu_{0.05}O₂ in Fig. 9(b) show the deconvoluted Cu 2p_{3/2} and Cu 2p_{1/2} peaks associated with Cu²⁺ as well as Cu¹⁺ ions. Peaks at 935.7–955.1 eV appeared for Cu²⁺ 2p_{3/2}–2p_{1/2} and peaks

at 933.7–953.4 eV appeared for Cu¹⁺ 2p_{3/2}–2p_{1/2}.^{32,53–55} The amount of Cu¹⁺ was found to be 76% in the as-prepared Ce_{0.95}Cu_{0.05}O₂.⁵⁶ Similarly, Fig. 9(c) shows the Co 2p core level deconvoluted spectrum of Ce_{0.95}Co_{0.05}O₂. Peaks at 781.8–796.2 eV appeared for Co²⁺ 2p_{3/2}–2p_{1/2} and peaks at 779.7–794.5 eV were attributed to Co³⁺ 2p_{3/2}–2p_{1/2}.^{57–59} The amount of Co²⁺ was found to be 67% in the as-prepared Ce_{0.95}Cu_{0.05}O₂. Furthermore, the Ni 2p core level spectrum investigated for Ce_{0.95}Ni_{0.05}O₂ in Fig. 9(d) shows the Ni 2p_{3/2} and 2p_{1/2} peaks associated with Ni²⁺ as well as Ni³⁺ ions. The Ni³⁺ peak appeared at higher binding energy (2p_{3/2}–2p_{1/2} at 856.7–874 eV) compared to Ni²⁺ (2p_{3/2}–2p_{1/2} at 855–872.8 eV).^{60,61} The amount of Ni²⁺ was found to be 68% in the as-prepared Ce_{0.95}Ni_{0.05}O₂. It is noteworthy that the doped transition metal ions in the synthesized oxides predominantly exhibited lower-valent oxidation states.

XRD analysis was performed on the spent Ce_{0.95}Ni_{0.05}O₂ catalyst (Fig. 10(a)), revealing that the crystal structure remained stable even after long-term stability testing. The absence of any additional peaks in the post-reaction XRD pattern confirmed that no structural changes had occurred. FE-SEM (inset, Fig. 10(a)) and EDX (Fig. 10(b)) analyses showed minor agglomeration of particles following the reaction.

The core-level deconvoluted XPS spectrum of Ce 3d for the four exhausted catalysts is presented in Fig. 11(a). The spectra displayed five spin–orbit coupled peaks, attributed to the Ce⁴⁺ and Ce³⁺ oxidation states, which agreed with those observed in the fresh catalysts. Fig. 11(a) illustrates the characteristic change of v⁰ peak from the fresh catalysts. Following the HER, the concentration of Ce⁴⁺ in the exhausted catalyst was reduced to 77, 66, 56 and 45% for CeO₂, Ce_{0.95}Cu_{0.05}O₂, Ce_{0.95}Co_{0.05}O₂, and Ce_{0.95}Ni_{0.05}O₂, respectively. The reduction of the formal charge of Ce⁴⁺ supports the H spill-over phenomenon from the metal to the support. Furthermore, we have conducted the core level XPS of Cu, Co and Ni from exhausted catalysts Ce_{0.95}Cu_{0.05}O₂ (Fig. 11(b)), Ce_{0.95}Co_{0.05}O₂ (Fig. 11(c)), and Ce_{0.95}Ni_{0.05}O₂ (Fig. 11(d)). Cu and Co were found in their reduced oxidation states after the HER. The Ni 2p spectra for the exhausted Ce_{0.95}Ni_{0.05}O₂, shown in Fig. 11(d), reveal that 73% of



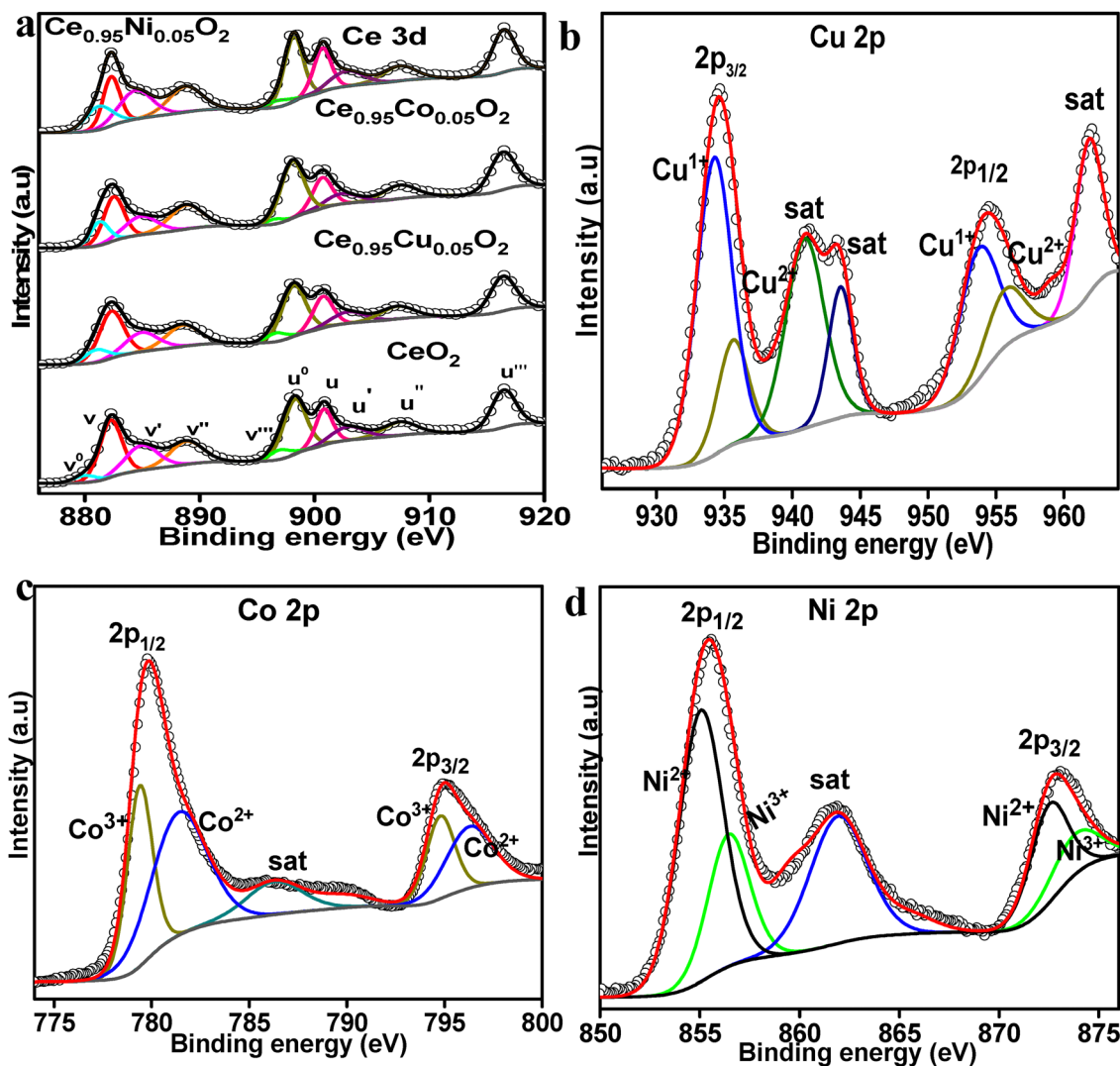


Fig. 9 (a) Ce 3d, (b) Cu 2p, (c) Co 2p and (d) Ni 2p core level spectra of the freshly synthesized pristine and doped solid solution oxides.

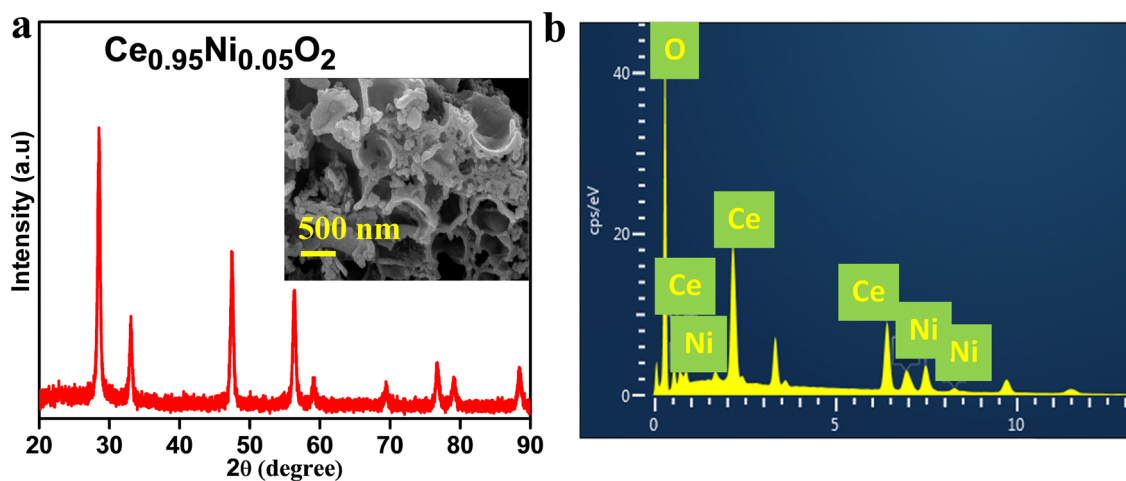


Fig. 10 (a) XRD and SEM image and (b) EDX of the exhausted $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ catalyst.





Fig. 11 (a) Ce 3d, (b) Cu 2p, (c) Co 2p, (d) Ni 2p and (e) O 1s core level spectra of the exhausted catalysts.

the Ni $2p_{3/2}$ peak corresponds to Ni^{2+} ions, indicating the maximum reduction of Ni during the HER process. Furthermore, we conducted the O 1s core level spectra of the exhausted catalysts (Fig. 11(e)). Interestingly, the vacancies increased significantly in the exhausted catalysts (vide O1s XPS of freshly prepared catalyst) after the HER, and the order of oxygen vacancies was in the following order: $\text{CeO}_2 < \text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2 < \text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2 < \text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$. The substantial reduction of the formal oxidation states of Ni and Ce in $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ compared to the other transition metals and support helped to create the highest oxygen

vacancies in the material. The higher extent of reduction of Ni as well as Ce and the substantial presence of oxygen vacancies during the HER facilitated the H spill-over and consequently faster kinetics over $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$.

4. Conclusion

A one-step solution combustion approach was used to synthesize the pristine CeO_2 and doped $\text{Ce}_{0.95}\text{Co}_{0.05}\text{O}_2$, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$,



and $\text{Ce}_{0.95}\text{Cu}_{0.05}\text{O}_2$ oxides. The production of phase-pure fluorite crystals with aliovalent ionic substitution of transition metals was validated by structural studies. The resulting porous nanocrystalline materials exhibited high oxygen vacancy concentrations, with $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ demonstrating the highest values, as conclusively revealed through detailed spectroscopic analysis. Among the synthesized materials, $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$ displayed superior HER activity with enhanced kinetics. The HER overpotential (η_{10}) was measured at 139 mV, accompanied by an impressive mass activity of 95 mA mg^{-1} and remarkable stability over 24 hours. Detailed Tafel analysis identified the Volmer–Heyrovsky pathway as the operative HER mechanism for $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$. Further mechanistic investigations, including in-depth surface analysis, suggested a H spill-over phenomenon from the metal to the support, which facilitated faster HER kinetics in $\text{Ce}_{0.95}\text{Ni}_{0.05}\text{O}_2$.

Data availability

The data for this article are included herein.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

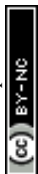
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