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Oxygen vacancies confined in hierarchically porous CsPbBr₃@Pb-MOF through *in situ* structural transformation for promoting photocatalytic CO₂ reduction†

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All-inorganic perovskite (CsPbBr₃) nanocrystals (NCs) are exceptional candidates for photocatalysis due to their optimal band structure, high molar extinction coefficient, and long charge-carrier diffusion lengths. However, their inevitable instability and low charge transfer efficiency have prevented their widespread applications. A strategy that can improve both stability and charge separation efficiency is desperately required. Herein, by introducing an *in situ* structural transformation strategy, hydrophobic hierarchically porous CsPbBr₃@Pb-MOF is obtained which exhibits high durability in water for more than 7 weeks and appealing thermal stability and resistance to anion intrusion. Meanwhile, abundant oxygen vacancies are detected which efficiently suppress the photogenerated charge recombination. As a result, impressive CO₂ photoreduction activity is achieved with CsPbBr₃@Pb-MOF as a catalyst. Through a gas–solid reaction (without sacrificial agents and photosensitizers), this CsPbBr₃@Pb-MOF-2 composite exhibits a CO yield of 107 μmol g⁻¹ h⁻¹ with 99.2% selectivity under visible-light (λ > 420 nm) irradiation, surpassing most reported CsPbX₃-based photocatalysts under similar conditions.

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1. Introduction

Excessive carbon dioxide (CO₂) emissions have greatly affected the balance of the natural carbon cycle, causing a greenhouse effect.¹ The direct conversion of CO₂ has been considered as a potential strategy for alleviating both global warming and energy shortages.² Lead halide perovskite (CsPbBr₃) nanocrystals (NCs) have emerged as an encouraging photocatalyst for various reactions including CO₂ reduction,^{3–5} pollutant degradation,^{6–8} hydrogen production,⁹ and organic synthesis^{10,11} due to their optimal band structure, high molar extinction coefficient,

and long charge-carrier diffusion lengths.^{12–14} However, pure CsPbBr₃ still faces two major obstacles including the inherent low stability and serious photogenerated electron–hole recombination, which limit its applications in photocatalysis.^{15–17} To overcome instability problems, encapsulating pure CsPbBr₃ into a protective matrix is considered to be a promising strategy. For example, SiO₂,^{18,19} zeolites,²⁰ glass,^{21–23} and polymers^{24,25} have been used as matrixes for the protection of all-inorganic perovskites. Concerning photogenerated charge recombination, the construction of CsPbBr₃-based heterojunctions has been proven to be an efficacious route. Metal oxides,^{26–28} g-C₃N₄,²⁹ and graphene³⁰ have been coupled with CsPbBr₃ to reduce charge recombination. However, a strategy that simultaneously improves the stability and charge separation of pure CsPbBr₃ is still underexplored.

Recently, metal–organic frameworks (MOFs) have emerged as photocatalysts, photocatalytic hosts, and cocatalysts for CO₂ photoreduction, garnering significant attention. This interest stems from their remarkable attributes such as distinctive electronic band structures, exceptional CO₂ adsorption capacities, and customizable light-absorption capabilities.^{31–33} MOFs have recently been established as promising host matrixes to protect and enhance the physicochemical properties of pure CsPbBr₃ NCs because of their tunable porous structure and chemical functionality.^{34–36} As we know, the preparation tech-

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nique largely determines the structures and performance of perovskite@MOF composites. Normally, the preparation methods of perovskite@MOFs can be broadly classified into four categories: (1) “ship-in-bottle” (ship refers to perovskites, bottle refers to MOFs), (2) “bottle-around-ship”, (3) Pb-MOF conversion, and (4) one-pot strategy.³⁵ The most prevalent “ship-in-bottle” strategy involves introducing precursors of a perovskite (e.g., Pb, Cs, Br, ions) into a prepared beforehand MOF inside which the perovskite can assemble subsequently.³⁷ To implement this methodology, Lu *et al.*³⁸ employed pre-synthesized Fe-doped PCN-221 and subjected it to treatment involving a PbI₂ solution, resulting in the formation of PbI₂@PCN-221(Fe_x). In the subsequent step, the PbI₂@PCN-221(Fe_x) composite underwent further treatment, this time utilizing an ethanolic solution of CH₃NH₃I, leading to the creation of the exquisite hybrid MAPbI₃@PCN-221(Fe_x). However, the diffusion of perovskite precursors into MOFs is largely restricted by the microporous structure of MOFs, which often results in the generation of substantial quantities of perovskites on the surface of MOFs rather than within their pores. Concerning the “bottle-around-ship” strategy, the pre-prepared perovskite is distributed in the precursor solution of the MOF. Subsequently, the assembly of the MOF is initiated to embed the perovskite. For example, Su *et al.*³⁹ achieved the growth of a zinc/cobalt zeolite imidazole coating on the surface of CsPbBr₃ quantum dots (QDs). High temperatures and polar solvents are precisely indispensable in the nucleation and building of MOFs but are harmful to the perovskite structure. In contrast, the “one-pot strategy” of premixing both MOFs and perovskite precursors to produce MOFs and perovskites simultaneously or sequentially is a highly desirable strategy due to the time, energy, and cost savings. To prepare MAPbBr₃@ZIF-8, all components required for the synthesis of ZIF-8 (Zn(NO₃)₂ and 2-methylimidazole) and MAPbBr₃ (PbBr₂ and MABr) were mixed and stirred in methanol/DMF solvent at room temperature by Mollick *et al.*⁴⁰ However, with the “one-pot strategy”, the location and size of the perovskite generated in the MOF cannot be easily manipulated. Recently, through “Pb-MOF conversion”, the lead metal nodes of Pb-MOF are picked as the source of Pb for perovskite construction, which facilitates the preparation of composites, minimizes perovskite aggregation, and shortens the electron transportation path.^{41–43} Using this approach, Li *et al.*⁴¹ reported the growth of perovskite nanocrystals (NCs) within the confines of the Pb-MOF structure through a direct conversion reaction between halide salt and the Pb-MOF. Nevertheless, the *in situ* growth of a perovskite in Pb-MOF may occupy and destroy the porous structure and affect the applications requiring porosity. Therefore, the development of a strategy that facilitates the introduction of perovskites into MOFs while preserving the porous structure and rendering the composite functionalities remains the central task.

Herein, *via* an *in situ* structural transformation strategy, hydrophobic hierarchically porous CsPbBr₃@Pb-MOF with oxygen vacancies is successfully prepared based on a lead bromide-based MOF (PbBr-MOF) and the corresponding

cesium bromide salt (CsBr). As expected, the hydrophobic Pb-MOF host significantly increased the stability of pristine CsPbBr₃ NCs. More importantly, the oxygen vacancies confined in CsPbBr₃@Pb-MOF effectively inhibit the photogenerated electron–hole recombination of CsPbBr₃. The produced CsPbBr₃@Pb-MOF is used for CO₂ photocatalytic reduction through the gas–solid mode without sacrificial agents and photosensitizers. This composite exhibits an exceptional CO yield of 107 μmol g⁻¹ h⁻¹ with 99.2% selectivity under visible light irradiation (λ > 420 nm), representing an improvement of 8.9 times over the yield observed with pure CsPbBr₃ (12 μmol g⁻¹ h⁻¹, 88% selectivity). The possible mechanisms responsible for the photoreduction of CO₂ are well studied.

2. Experimental

2.1 Materials

Lead bromide (PbBr₂, 99.9%), lead nitrate (Pb(NO₃)₂, 99%), cesium bromide (CsBr, 99%), terephthalic acid (TPA, >99%), trimesic acid (BTC, C₆H₃(CO₂H)₃, 98%), oleic acid (OA, 80%–90%), oleylamine (OM, 90%), triethylamine (C₆H₁₅N, 99.0%), toluene (C₇H₈, 99%), *N,N*-dimethylformamide (DMF, 99.5%), methanol (CH₃OH, 99.5%), fuming nitric acid (HNO₃, ≥95%), ethanol (CH₃CH₂OH, 99.5%), and *n*-hexane (C₆H₁₄, >99%) were purchased from Aladdin Co. Ltd.

2.2 Synthesis

2.2.1 Preparation of PbBr-MOF. The solvothermal approach was used to create PbBr-MOF with a little change to the prior procedure.⁴⁴ For the typical synthesis of PbBr-MOF 0.1835 g PbBr₂ (0.5 mmol), 0.2640 g terephthalic acid (1.60 mmol), 120 μL nitric acid, ethanol (1.5 mL), and DMF (2.5 mL) were mixed introduced into a 20 mL Teflon reactor, kept at 100 °C for 72 h, and then cool down naturally. Finally, the obtained rod-like crystals were rinsed with DMF and ethanol and dried in a natural environment.

2.2.2 Solvothermal preparation of CsPbBr₃@Pb-MOF. To prepare CsPbBr₃@Pb-MOF, PbBr-MOF (20 mg), CsBr (15, 30, 45 mg), methyl alcohol (4 mL), and deionized water (0.1 mL) were loaded into a 25 mL Teflon reactor, reacted at 190 °C for 2 h, then cooled. The precipitates generated were repeatedly rinsed with H₂O. We have labeled the samples with different concentrations of CsPbBr₃ as CsPbBr₃@Pb-MOF-1, CsPbBr₃@Pb-MOF-2, and CsPbBr₃@Pb-MOF-3, respectively.

2.2.3 Preparation of CsPbBr₃@Pb-MOF without oxygen vacancy. CsPbBr₃@Pb-MOF without oxygen vacancies was synthesized using a one-pot ultrasonic method. In detail, Pb(NO₃)₂ (0.1 mmol), BTC (0.1 mmol), CsBr (0.1 mmol), and OM (4 mL) were sequentially added to deionized water (8 mL) and subjected to 30 minutes of sonication. The resulting oil phase solution was separated and washed multiple times with *n*-hexane. Afterward, the obtained precipitate was dried at 80 °C for 4 hours. The resulting sample was named CsPbBr₃@Pb-MOF-WOV.

2.2.4 Preparation of Pb-MOF. To prepare Pb-MOF, PbBr-MOF (20 mg), methyl alcohol (4 mL), and deionized water (0.1 mL) were loaded into a 25 mL Teflon reactor, reacted at 190 °C for 2 h, and then cooled. The precipitates generated were repeatedly rinsed with H₂O.

2.2.5 Preparation of pure CsPbBr₃. CsPbBr₃ NCs were synthesized using a process modified from a reported work.⁴⁵ At room temperature, PbBr₂ (0.2 mmol) and CsBr (0.2 mmol) were dispersed in DMF (8 mL). After complete dissolution, OM (2 mL) and OA (2 mL) were added to the aforementioned solution. To synthesize CsPbBr₃ NCs, the precursor was promptly added to toluene (5 mL). The synthesis of CsPb(BrCl)₃ and CsPb(BrI)₃ was comparable to that of CsPbBr₃. The difference in the preparation was that CsBr₂ was replaced by CsCl₂ or CsI₂.

2.2.6 Synthesis of general Pb-MOF. General Pb-MOF was synthesized using the procedure of a reported work.⁴⁶ The mixture of Pb(NO₃)₂ (0.331 g, 1.0 mmol), terephthalic acid (0.166 g, 1.0 mmol), H₂O (10 mL), and ethanol (2 mL) was placed in a 20 mL Teflon reactor. Then, to achieve a pH of about 4.0 in the solution, triethylamine (0.4 mL) was added. The Teflon reactor was placed in an oven at 160 °C for 4 days and then cooled down to the surrounding ambient temperature. We could obtain the pale-yellow crystals after washing them with H₂O several times.

2.3 Photocatalytic CO₂ reduction experiments

A solid-gas mode was used to test the CO₂ photoreduction performances of the as-prepared samples. Firstly, 10 mg catalysts were dispersed in 1 mL hexane and dropped on a 1 cm × 2 cm glass substrate. Next, to avoid interference with the experimental results, excess organic solvents were removed by vacuum drying. The prepared films and 100 μL water as a reducing agent were put in a 50 mL sealed Pyrex bottle. Before illumination, the reactor atmosphere was replaced with CO₂ (99.99%) for 30 min to remove other gases. As the light source, a 300 W Xe lamp (cut-off 420 nm) was employed and the light intensity was maintained at 120 mW cm⁻². In order to detect CO during the reaction, 0.5 mL of the photocatalyzed mixed gas was fed into a flame ionization detector (FID) for gas chromatography. 1 mL of the gas was injected into a gas chromatography-thermal conductivity detector (GC-TCD) to detect H₂. The isotope examination was performed under identical photocatalytic reaction conditions and ¹³C was determined by GC-MS. Cycling experiments were performed under the same conditions and the samples were directly transferred and reused in the next cycle.

2.4 Electrochemical measurements

A CHI760E electrochemical analyzer with a three-electrode setup was used to conduct each photoelectrochemical measurement. As the working electrode, fluorine-doped tin oxide (FTO) glass (1 × 1 cm²) was covered with a photocatalyst. Ag/AgCl (in saturated KCl) and a Pt slice (1 × 1 cm²) were adopted as reference electrodes and counter electrodes, respectively. The 1-butyl-3-methylimidazolium hexafluorophosphate (TBAPF₆) CH₂Cl₂ solution (0.2 mol) worked as the electrolyte.

2.5 Material characterization

Powder X-ray diffraction (PXRD) patterns were performed on a Rigaku Ultima IV using Cu Kα radiation (λ = 1.54056 Å). A JEM-2010 was used to capture transmission electron microscopy (TEM) and high-resolution transmission electron microscopy (HRTEM) images. Scanning electron microscopy (SEM) and energy-dispersive X-ray spectroscopy (EDS) were conducted using SU8010. N₂ adsorption-desorption isotherms were collected using a MIC-ASAP 2460 instrument at 77 K. Surface wettability was measured using a Phoenix 300. CO₂ adsorption isotherms were collected using a MIC-ASAP 2460 instrument at 298 K. X-ray photoelectron spectroscopy (XPS) was conducted on an Escalab 250 instrument. Using a TU-1900 UV-Vis spectrophotometer, UV-Vis diffuse reflectance spectra (UV-vis DRS) were recorded. The emission lifetime was analyzed using a FLS1000 fluorescence spectrometer. The gas product was collected using a GC-9800 gas chromatograph. ¹³CO₂ and H₂¹⁸O isotope experiments were conducted using an HP 6890GC-5973MSD gas chromatography-mass spectrometer. Using a Bruker EMXplus instrument, electron paramagnetic resonance (EPR) signals were acquired. Inductively coupled plasma mass spectrometry (ICP-MS) was tested using an Agilent 7900 instrument.

3. Results and discussion

3.1 Design and fabrication of CsPbBr₃@Pb-MOF

A lead bromide-based MOF (PbBr-MOF) was deliberately selected as the matrix MOF and prepared according to a reported method with adjustments (Fig. S1†).⁴⁴ In the structures, zigzag [Pb₂Br₃]⁺ chains as the secondary building units (SBUs) are coordinated by terephthalic acid (Fig. S2†). CsPbBr₃ encapsulated composites were synthesized using PbBr-MOF and CsBr as raw materials under solvothermal conditions (Fig. 1). In comparison with the method used by Lu *et al.*³⁸ for preparing the composites using the “ship-in-bottle” approach, our method offers a simpler preparation process, while the “ship-in-bottle” approach has advantages in controlling the proportions of each component. Strong emissions were observed under a UV lamp (365 nm) excitation, implying that CsPbBr₃ may be formed (Fig. S3†). While keeping the amount of PbBr-MOF constant, we varied the dosage of CsBr (15, 30, and 45 mg) to obtain samples with different concentrations of CsPbBr₃ within Pb-MOF, that were respectively denoted as CsPbBr₃@Pb-MOF-1, CsPbBr₃@Pb-MOF-2, and CsPbBr₃@Pb-MOF-3. Furthermore, the content of Cs in the CsPbBr₃@Pb-MOF composite materials was determined using inductively coupled plasma mass spectrometry (ICP-MS), yielding values of 0.8% (CsPbBr₃@Pb-MOF-1), 2.7% (CsPbBr₃@Pb-MOF-2), and 3.3% (CsPbBr₃@Pb-MOF-3), respectively.

3.2 Characterization of CsPbBr₃@Pb-MOF

Initially, the crystal structures of CsPbBr₃, Pb-MOF, and CsPbBr₃@Pb-MOF composites were investigated using powder X-ray diffraction (PXRD) analysis (Fig. 2a and Fig. S4†). For

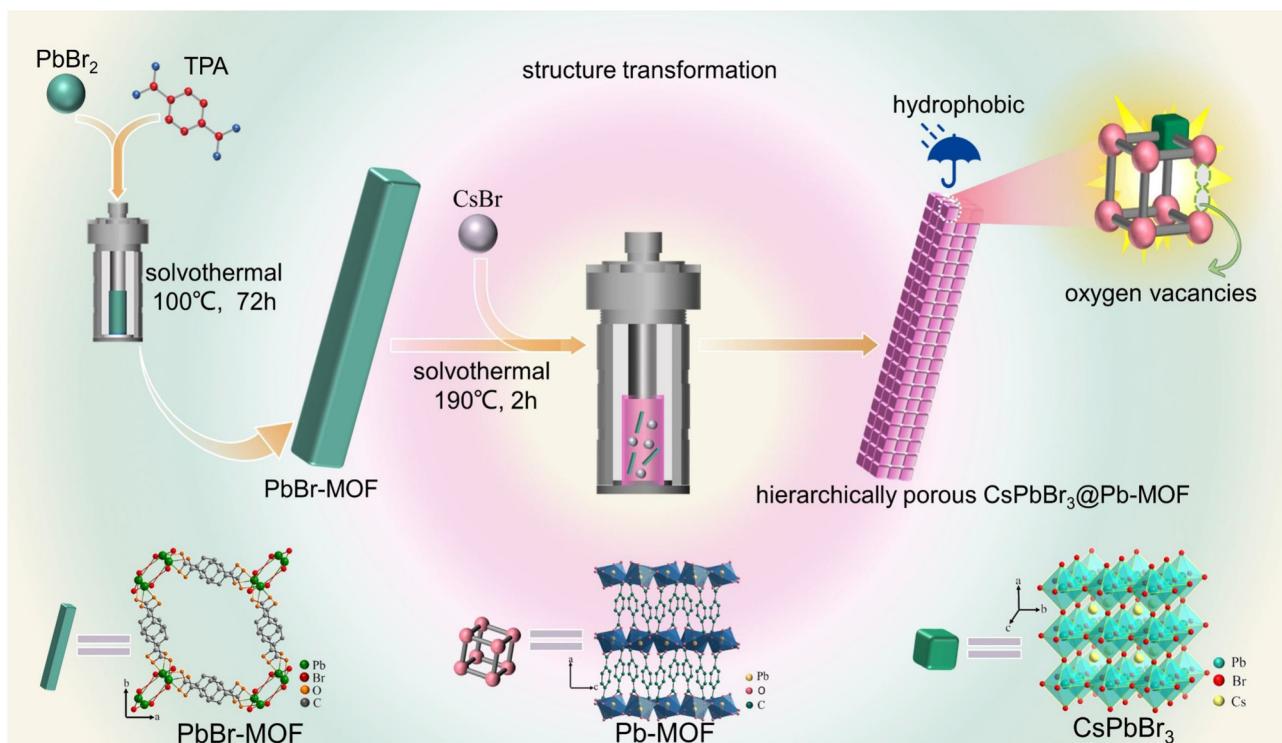


Fig. 1 A schematic diagram for the synthesis of CsPbBr₃@Pb-MOF.

sample CsPbBr₃@Pb-MOF-1, no characteristic peaks corresponding to CsPbBr₃ were observed, due to its low concentration. For both CsPbBr₃@Pb-MOF-2 and CsPbBr₃@Pb-MOF-3, the successful combination of CsPbBr₃ NCs was indicated by powder X-ray diffraction (PXRD), in which three distinctive peaks at 21.4, 26.3, and 43.5° corresponded to the (110), (111), and (220) planes of cubic CsPbBr₃ (PDF#75-0412), respectively.⁴⁷ A notable observation was the significantly stronger intensity of the characteristic peaks of CsPbBr₃ in CsPbBr₃@Pb-MOF-3 compared to CsPbBr₃@Pb-MOF-2, attributed to the higher content of CsPbBr₃ in CsPbBr₃@Pb-MOF-3 compared to CsPbBr₃@Pb-MOF-2. However, to our surprise, the characteristic diffraction peaks belonging to the PbBr-MOF matrix almost disappeared and are replaced by a new set of diffraction peaks. In order to clarify this issue, the parent MOF was subjected to solvothermal conditions. The obtained powder samples were tested by PXRD and it was found that their diffraction peaks basically corresponded to the diffraction peaks in the perovskite composites, suggesting that the structure of PbBr-MOF was transformed under solvothermal synthesis (Fig. S5 and S6†).^{44,46} The elements and their contents in the transformed materials were examined by energy-dispersive X-ray (EDX). In EDX, Pb, C, and O were detected in equal distribution, suggesting that Br ions of PbBr-MOF are shed during the structural transformation (Fig. S7†). Combined with the PXRD results, we boldly speculate that PbBr-MOF is likely to be transformed into Pb-MOF. To confirm the structural stability, we study the PXRD pattern of PbBr-

MOF and Pb-MOF after water immersion (Fig. S8†). The major peaks of PbBr-MOF disappeared after being immersed for 10 min in water, while the peaks of Pb-MOF remained intact, suggesting that Pb-MOF is extremely stable in water. During the test, we found that powdery Pb-MOF can float on the surface of the water. We speculate that the excellent water stability may result from its hydrophobicity. The water contact angle of 140° verifies this prediction (Fig. S9†).⁴⁸ Pb-MOF with a hydrophobic interface makes it a promising candidate for the protection of CsPbBr₃. For comparison, we tried to synthesize this MOF using Pb(NO₃)₂ and terephthalic acid as the precursor. Fortunately, crystals of this MOF (denoted as general Pb-MOF, see more details in Experimental sections) were obtained and its structure was resolved in which Pb ions are octahedrally coordinated by terephthalic acid. By comparing its PXRD patterns, we could find that they matched each other very well, indicating that the transformed structure is Pb-MOF (Fig. S10†). The generated composite was therefore named CsPbBr₃@Pb-MOF. We speculate that the solvothermally induced *in situ* structural transformation provides a prerequisite for the growth of the perovskite in CsPbBr₃@Pb-MOF. The broken Pb-Br bond releases occupied Pb sites, which are crucial for the formation of the CsPbBr₃ perovskite, meanwhile, the detached halogen is also considered an important component of the perovskite.

The structure of the composite was further understood by transmission electron microscopy (TEM), high-resolution transmission electron microscopy (HRTEM), energy dispersive

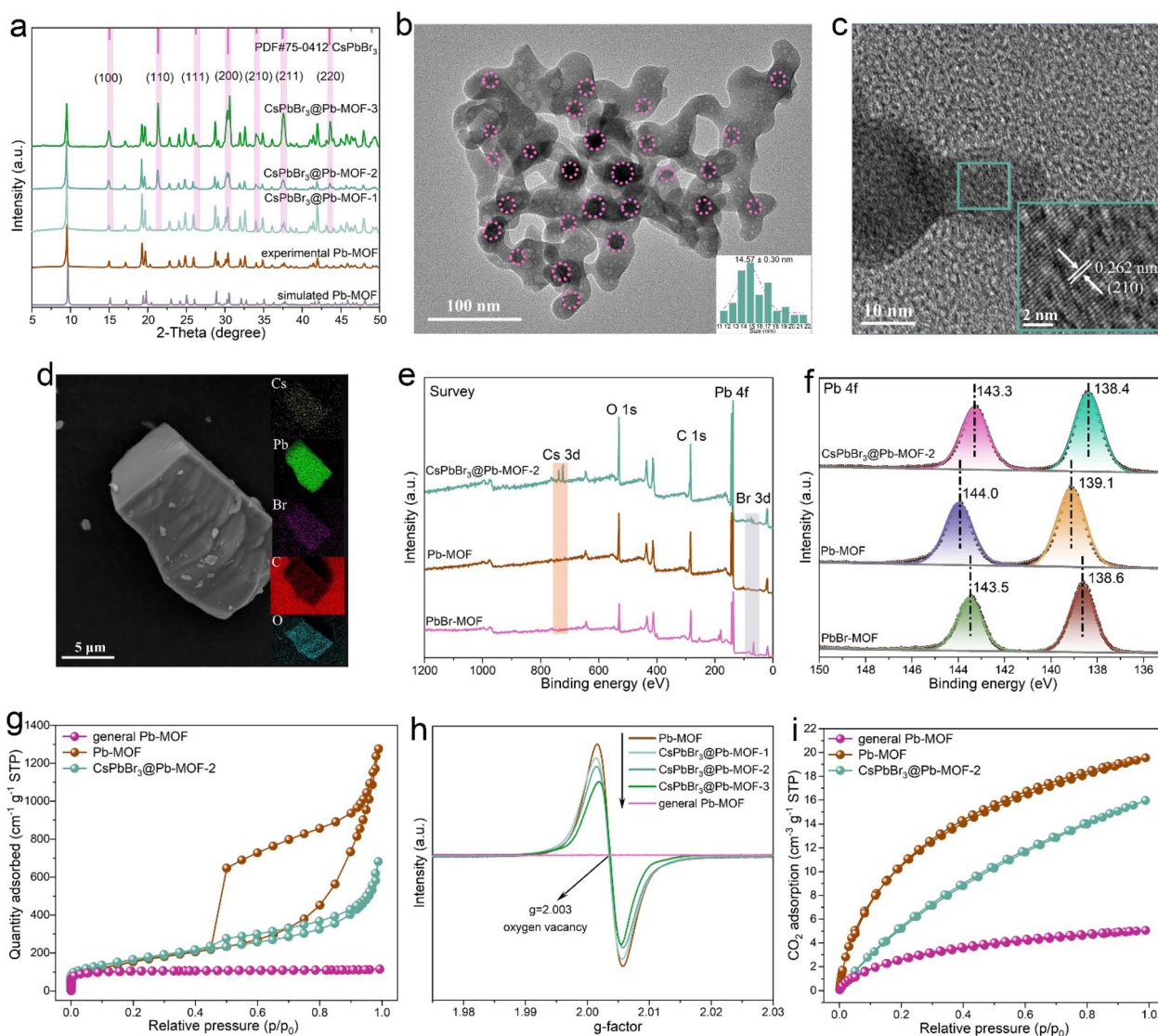


Fig. 2 (a) PXRD patterns of CsPbBr₃@Pb-MOF. (b) TEM image of CsPbBr₃@Pb-MOF-2, dark dots in the figure are CsPbBr₃ NCs. Inset of (b) shows particle size distributions of CsPbBr₃ NCs. (c) The HRTEM image of CsPbBr₃@Pb-MOF. Inset of (c) shows the lattice fringe of CsPbBr₃ NCs. (d) SEM image and EDX mapping of CsPbBr₃@Pb-MOF-2. (e) XPS analysis of the samples: (e) survey spectra and (f) Pb 4f spectra. (g) N₂ adsorption–desorption isotherms of general Pb-MOF, Pb-MOF, and CsPbBr₃@Pb-MOF-2 at 77 K. (h) Electron paramagnetic resonance spectra of general Pb-MOF, Pb-MOF, and CsPbBr₃@Pb-MOF. (i) CO₂ adsorption–desorption isotherms of general Pb-MOF, Pb-MOF, and CsPbBr₃@Pb-MOF-2 at 298 K.

X-ray (EDX) spectroscopy, X-ray photoelectron spectroscopy (XPS), electron paramagnetic resonance (EPR) and gas adsorption. TEM and HRTEM measurements were performed and depicted in Fig. 2b and c. In Fig. 2b, we can see that the CsPbBr₃ NCs marked with circles, indicating a size of 14.75 ± 0.30 nm, are uniformly embedded in the Pb-MOF host without significant aggregations. As shown in Fig. S11,[†] pure CsPbBr₃ NCs exhibit uniform cubic shapes with an average particle size of 10.24 ± 0.07 nm. As illustrated in Fig. 2c, HRTEM images show lattice spaces of 0.262 nm, which are a good match with the (210) plane of cubic CsPbBr₃.⁴⁹ In Fig. 2d, the EDX mapping of CsPbBr₃@Pb-MOF-2 indicates that Cs, Pb, Br, C, and O are uniformly distributed in the matrix. XPS measurements are shown in Fig. 2e and f. In the comparison between

PbBr-MOF and Pb-MOF, the Br 3d signal belonging to PbBr-MOF is not seen for Pb-MOF. In addition, when compared with Pb-MOF, signals for new Br and Cs species are observed in CsPbBr₃@Pb-MOF-2 composites. For a more detailed analysis, high-resolution XPS spectra of Pb 4f are presented in Fig. 2f. For PbBr-MOF, Pb 4f spectra show two prominent peaks at 143.5 (4f_{5/2}) and 138.6 eV (4f_{7/2}).⁵⁰ In addition, the Pb 4f_{5/2} and Pb 4f_{7/2} peaks of Pb-MOF were shifted to 144.0 and 139.1 eV, respectively, due to the change in the coordination environment of the Pb atom during the structural transformation (the breaking of Pb–Br bonds and reorganization of Pb with carboxylic acid ligands). Of note, in comparison with Pb-MOF, the Pb 4f bands of CsPbBr₃@Pb-MOF-2 demonstrates a 0.7 eV shift to the lower binding energy after the growth of

CsPbBr₃ in Pb-MOF. This phenomenon indicates that the Pb node in Pb-MOF is coordinated with Br to produce a new Pb-Br bond belonging to CsPbBr₃.^{41,51}

N₂ adsorption-desorption isotherms of general Pb-MOF, Pb-MOF, and CsPbBr₃@Pb-MOF-2 were collected at 77 K to evaluate porosity and the specific surface area, as illustrated in Fig. 2g. The curves for general Pb-MOF are a type I isotherm, signifying the presence of microporosity.⁵² Moreover, the N₂ adsorption-desorption isotherms of Pb-MOF reveal a combination mode of type I and type IV, implying the simultaneous existence of micropores and mesopores.^{52,53} On the other hand, intrinsic micropores (<2 nm) are maintained, while mesoporous features (concentrated at 4 nm) appear in Pb-MOF in the pore size distribution curve (Fig. S12†). The Brunauer-Emmett-Teller (BET) surface area of Pb-MOF is determined to be 904.58 m² g⁻¹, which is bigger than the general Pb-MOF (398.43 m² g⁻¹). The rich mesoporous and larger surface area of Pb-MOF are expected to favor mass transport and enhance the utilization of active sites.^{47,54} In Fig. 2g, a combination mode of type I and type IV of the N₂ adsorption-desorption isotherm is observed for CsPbBr₃@Pb-MOF-2.⁵² The result confirms the presence of micropores and mesopores in CsPbBr₃@Pb-MOF-2.^{52,53} Likewise, the pore size distributions of CsPbBr₃@Pb-MOF-2 agree with this conclusion (Fig. S13†). Meanwhile, compared with Pb-MOF (904.58 m² g⁻¹), CsPbBr₃@Pb-MOF (587.68 m² g⁻¹) shows lower BET surface areas. Furthermore, apparent mesopores were observed in Pb-MOF on comparing the TEM of general Pb-MOF (Fig. S14†). The generation of mesopores may be ascribed to the *in situ* structural transformation of MOFs which leads to the missing linker defects (Fig. S15†).^{53,55,56}

Subsequently, electron paramagnetic resonance (EPR) was used to detect oxygen vacancy, as shown in Fig. 2h. Compared with general Pb-MOF, a symmetrical signal of the oxygen vacancy ($g = 2.003$) is exhibited in the EPR spectra of Pb-MOF.⁵⁷ These results unambiguously suggest the existence of oxygen vacancies in Pb-MOF together with coordinatively unsaturated Pb atoms. Obviously, the symmetric signal belonging to the oxygen vacancy is also revealed in the CsPbBr₃@Pb-MOF composite.⁵⁷ However, as the content of CsPbBr₃ increased in Pb-MOF, a downward trend was observed in the intensity of oxygen vacancies. Such oxygen vacancies can act as electron capture traps that can facilitate photogenerated charge separation.⁵⁸ CO₂ adsorption is a prerequisite for photocatalytic CO₂ reduction. CO₂ capture capacities of general Pb-MOF, Pb-MOF, and CsPbBr₃@Pb-MOF-2 were evaluated from CO₂ adsorption isotherms at 297 K, shown in Fig. 2i. The general Pb-MOF and Pb-MOF display CO₂ uptakes of 5 and 19.5 cm³ g⁻¹, respectively. We ascribed the enhanced CO₂ adsorption to a higher specific surface area and mesoporous channel in Pb-MOF.^{59,60} Correspondingly, the CO₂ adsorption amount of CsPbBr₃@Pb-MOF-2 (15.9 cm³ g⁻¹) is between general Pb-MOF and Pb-MOF. The good water stability, abundant oxygen vacancy, mesoporous structure, and higher specific surface area make Pb-MOF a promising host to assist CsPbBr₃ for photocatalytic CO₂ reduction.^{48,54,61}

3.3 Optical properties

Along with the structural characteristics mentioned above, the optical characteristics of CsPbBr₃@Pb-MOF-2 composites, pure CsPbBr₃ NCs and Pb-MOF were also studied, as illustrated in Fig. 3. As shown in Fig. 3a, with ultraviolet-visible (UV-Vis) diffuse reflectance spectroscopy (DRS) the light absorption properties of pure CsPbBr₃, Pb-MOF and CsPbBr₃@Pb-MOF-2 composites were evaluated. The absorption band edges of Pb-MOF (425 nm), pure CsPbBr₃ (520 nm) and CsPbBr₃@Pb-MOF-2 (543 nm) samples are in the visible range. However, on coupling CsPbBr₃ with Pb-MOF, the absorption edge of CsPbBr₃@Pb-MOF-2 was markedly red-shifted, resulting in an apparent enhancement of the sample's light-absorbing capacity. This is of significant importance for the utilization of photocatalysts in a practical setting. The band gap of the samples was calculated using the Kubelka-Munk function,²⁸ as depicted in Fig. 3b. These calculated values are consistent with previous literature results.^{14,20} Based on the reaction thermodynamics, the energy band structure is also critical in photocatalytic performance. Herein, Mott-Schottky (MS) analyses of CsPbBr₃ and Pb-MOF were performed and depicted in Fig. 3c and d. Firstly, as shown in Fig. 3b, the bandgaps ($E_{g,s}$) of CsPbBr₃ and Pb-MOF were evaluated to be 2.42 eV and 3.08 eV, respectively. In the MS curve, the slope of CsPbBr₃ and Pb-MOF demonstrates a positive value, suggesting the samples are n-type characteristic photocatalysts. As known, the MS curve is assessed as a flat band potential and the flat band potential of an n-type semiconductor is 0.1 V lower than its valence band (E_{CB}).⁶² According to the MS curve, the flat band potential of CsPbBr₃ is about -1.10 V vs. Ag/AgCl (≈ -0.90 V vs. NHE) (flat band potential of Pb-MOF is about -0.17 V vs. NHE). Therefore, the E_{CB} of CsPbBr₃ is -1.00 V vs. NHE (E_{CB} of Pb-MOF is -0.27 V vs. NHE) and the valence band (E_{VB}) of CsPbBr₃ was calculated as 1.42 V vs. NHE (E_{VB} of Pb-MOF is 2.81 V vs. NHE), based on the formula of $E_{CB} = E_{VB} - E_g$. Analyzing the energy level, as illustrated in Fig. 3e, the E_{CB} of CsPbBr₃ (-1.00 V vs. NHE) is more negative than CO₂/CO (-0.52 V vs. NHE) and O₂/O₂⁻ (-0.33 V vs. NHE) and the E_{VB} of CsPbBr₃ (0.85 V vs. NHE) is more positive than H₂O/O₂ (0.81 V).¹² As for Pb-MOF, the E_{CB} of Pb-MOF (-0.27 V vs. NHE) is not more negative than CO₂/CO (-0.52 V vs. NHE) and O₂/O₂⁻ (-0.33 V vs. NHE). But the E_{VB} of Pb-MOF (2.81 V vs. NHE) is more positive than H₂O/O₂ (0.81 V), H₂O/OH (1.99 V) and OH⁻/OH (2.34 V). To further verify the energy band structure of CsPbBr₃@Pb-MOF-2, CsPbBr₃ and Pb-MOF, electron paramagnetic resonance (EPR) analysis was executed. For EPR studies, the sample was dissolved in 5,5-dimethyl-1-pyrroline *N*-oxide (DMPO, capture agent of the spin electron) to identify [•]O₂⁻ and [•]OH, respectively.^{26,63} In Fig. 3f, it can be seen that under visible-light irradiation conditions, Pb-MOF did not exhibit the characteristic peaks of DMPO-[•]O₂⁻. In comparison, the characteristic signals of DMPO-[•]O₂⁻ are successfully detected in CsPbBr₃@Pb-MOF-2 and CsPbBr₃ NCs. In addition, the peak intensity of CsPbBr₃@Pb-MOF-2 is much stronger than

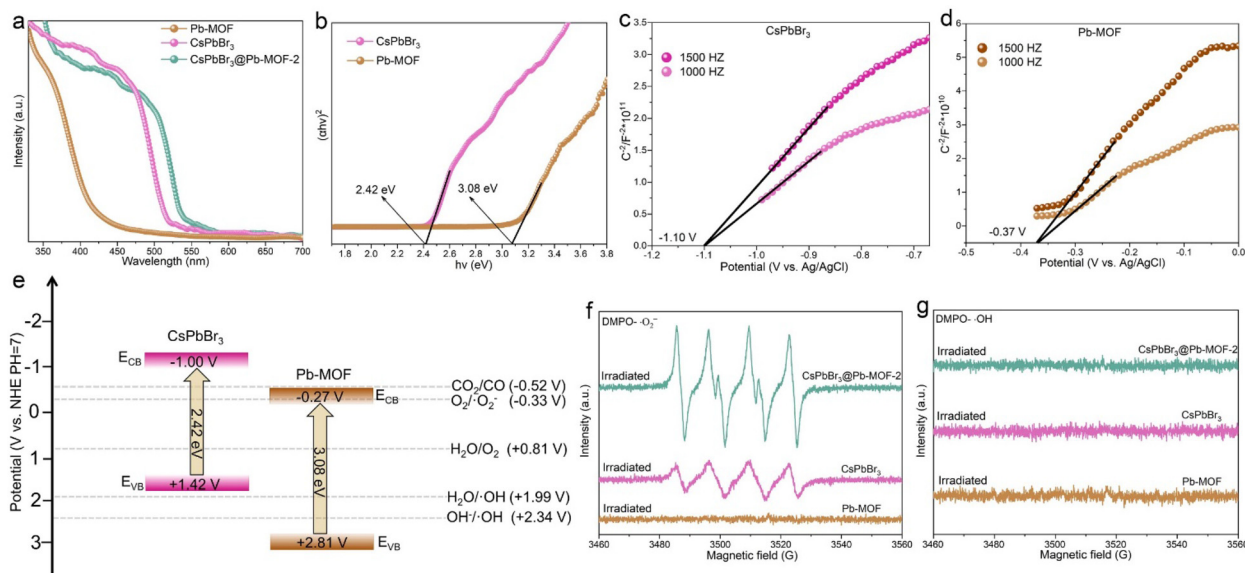


Fig. 3 (a) UV-vis diffuse reflectance spectra of Pb-MOF, CsPbBr₃ and CsPbBr₃@Pb-MOF-2. (b) The band gap spectra of CsPbBr₃ and Pb-MOF. Mott-Schottky plots of (c) CsPbBr₃ and (d) Pb-MOF. (e) Schematic illustration of the potential energy diagram for CsPbBr₃ and Pb-MOF. Electron paramagnetic resonance spectra of (f) DMPO- $\cdot\text{O}_2^-$ and (g) DMPO- $\cdot\text{OH}$ in the presence of CsPbBr₃@Pb-MOF-2, CsPbBr₃ and Pb-MOF.

CsPbBr₃ NCs, suggesting CsPbBr₃@Pb-MOF-2 possesses a higher electron transfer efficiency. As shown in Fig. 3g, on irradiating Pb-MOF, CsPbBr₃ NCs and CsPbBr₃@Pb-MOF-2, no typical peak of DMPO- $\cdot\text{OH}$ is detected, indicating the possible formation of a Z-scheme heterojunction between CsPbBr₃ and Pb-MOF.⁶⁴ This observation highlights the essential role of oxygen vacancies in Pb-MOF in facilitating the transfer of photogenerated electrons between CsPbBr₃ and Pb-MOF. Meanwhile, the band positions of CsPbBr₃@Pb-MOF and CsPbBr₃ satisfy the thermodynamic requirements for reduction and oxidation reactions, indicating it might be an appropriate catalyst for CO₂ reduction.

3.4 Water stability, thermal stability, and preventing anion intrusion studies

To check the potential of CsPbBr₃@Pb-MOF for photocatalytic applications, we systematically investigated the stability of CsPbBr₃@Pb-MOF-2 including water stability, thermal stability, and preventing anion intrusion (Fig. 4). Being ionic crystals, CsPbBr₃ NCs readily dissolve in the presence of water, which quenches their luminescence. The PL intensities of CsPbBr₃@Pb-MOF-2 remained consistent, displaying approximately 82.6% of their initial strength even after 49 days of immersion in water, as depicted in Fig. 4a (left side) and Fig. S16.† In sharp contrast, when 1 ml of water was added to pure CsPbBr₃ NCs dispersed in hexane for 20 min, the fluorescence value of the CsPbBr₃ sample was only 1% of the initial value (Fig. 4a, right side). This is attributed to the excellent hydrophobicity exhibited by CsPbBr₃@Pb-MOF-2 (water contact angle 140.2°), (Fig. S17†). Next, the thermal stability of CsPbBr₃@Pb-MOF-2 was investigated during heating-cooling in the range of 20–120 °C (Fig. 3b). After heating at 120 °C for

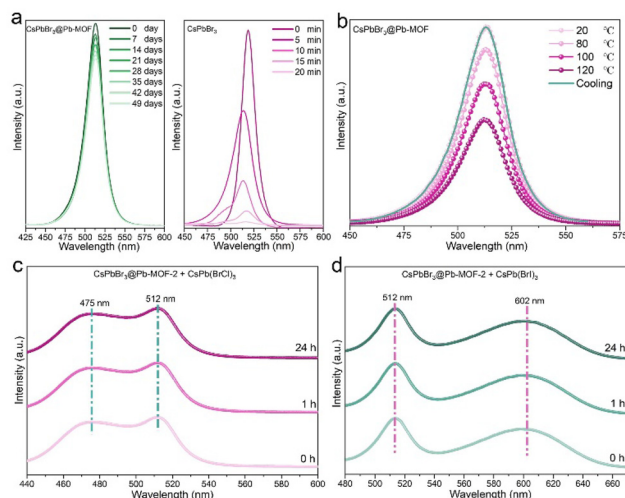


Fig. 4 (a) PL intensity changes of CsPbBr₃@Pb-MOF-2 and CsPbBr₃ with time. (b) The PL spectra of CsPbBr₃@Pb-MOF-2 as a function of temperature. PL spectra of (c) CsPbBr₃@Pb-MOF-2 and CsPb(BrCl)₃ and (d) CsPbBr₃@Pb-MOF-2 and CsPb(BrI)₃ mixed powders at different time intervals.

30 min, CsPbBr₃@Pb-MOF-2 could retain about 52% of its original PL intensity. Impressively, the PL emission value progressively recovered after cooling (green line), eventually reaching the original value. PXRD of CsPbBr₃@Pb-MOF-2 after immersion in water and after heating remains intact further proving their high stability (Fig. S18†). The anion intrusion issue of CsPbBr₃ is another shortcoming in the application.⁶⁵ To illustrate the behavior of CsPbBr₃@Pb-MOF-2, samples containing single and two halogens including CsPbBr₃@Pb-MOF-2 and CsPb(BrCl)₃, and CsPbBr₃@Pb-MOF-2 and CsPb

(BrI)₃, were mixed and subjected to the environment. As shown in Fig. 4c and d, the position and intensity of the fluorescence peaks of the samples did not change after 24 hours. The excellent water stability, thermal stability, and prevention of anion exchange of this CsPbBr₃@Pb-MOF-2 composite can be attributed to the hydrophobic Pb-MOF walls that protect CsPbBr₃ from various chemical conditions.⁶⁶

3.5 Photoinduced charge separation and transfer dynamics

As known, the photogenerated charge transfer and separation efficiency of the sample have a crucial effect on the photocatalytic efficiency.^{12–17} On account of this point, photoelectrochemical, electrochemical, and photoluminescence experiments were performed, as presented in Fig. 5. Firstly, the photocurrent response of CsPbBr₃@Pb-MOF-2 and pure CsPbBr₃ NCs were measured under turn-on-off cycles of visible-light ($\lambda > 420$ nm) irradiation. As demonstrated in Fig. 5a, the CsPbBr₃@Pb-MOF-2 sample exhibits a significantly higher photocurrent density compared to CsPbBr₃ NCs. This result indicates that CsPbBr₃@Pb-MOF-2 has a more efficient transfer of photogenerated charge. In other words, oxygen vacancies in Pb-MOF act as a reservoir to receive the photogenerated charge generated by CsPbBr₃. Furthermore, the charge transport behaviors of CsPbBr₃@Pb-MOF-2 and CsPbBr₃ were probed using electrochemical impedance spectroscopy (EIS), as depicted in Fig. 5b. In the Nyquist plot, CsPbBr₃@Pb-MOF-2 exhibits smaller semicircle radii owing to small charge transfer resistance.²⁷ In comparison, pure CsPbBr₃ NCs show a larger charge transfer resistance, demonstrating the superb carrier mobility between CsPbBr₃ NCs and Pb-MOF. Furthermore, the charge transfer dynamics of

CsPbBr₃@Pb-MOF and CsPbBr₃ NCs are evaluated by steady-state photoluminescence (PL). The fluorescence intensity of CsPbBr₃@Pb-MOF-2 was dramatically quenched compared to CsPbBr₃ NCs (Fig. 5c), verifying that the recombination of photogenerated charges of CsPbBr₃@Pb-MOF-2 was effectively inhibited. Furthermore, the time-resolved PL spectra of CsPbBr₃@Pb-MOF-2 and CsPbBr₃ NCs were recorded (Fig. 5d). The average lifetime of the samples was observed to decline significantly, from 66.8 ns for pure CsPbBr₃ NCs to 12.3 ns for CsPbBr₃@Pb-MOF (Table S1†).^{30,39} From all the above experiments, we believe that the enhanced photogenerated charge separation in CsPbBr₃@Pb-MOF-2 may be due to the introduced oxygen vacancies acting as charge capture centers.

3.6 Photocatalyst activity tests

This as-synthesized CsPbBr₃@Pb-MOF, CsPbBr₃@Pb-MOF-WOV (Fig. S19†), pure CsPbBr₃ NCs, and pure Pb-MOF were further explored by CO₂ photoreduction in the gas-solid mode with H₂O as the reductant, without using any sacrificial reagent or photosensitizers (Fig. S20†). Under visible-light ($\lambda > 420$ nm) irradiation, for CsPbBr₃@Pb-MOF, CsPbBr₃@Pb-MOF-WOV and CsPbBr₃ NCs, CO is the main reduction product along with a trace of H₂. CsPbBr₃@Pb-MOF exhibits $\geq 99.2\%$ selectivity towards CO production, while pure CsPbBr₃ NCs and CsPbBr₃@Pb-MOF-WOV only show 88% and 90.0 selectivities (Table S2†). The high selectivity of CsPbBr₃@Pb-MOF may be attributed to its hydrophobic surface depressing H₂ production (Fig. S21†).⁶⁷ In Fig. 6a, CsPbBr₃@Pb-MOF-2 displays the highest CO production rate of 107 $\mu\text{mol g}^{-1} \text{h}^{-1}$, which is 8.9 times higher than that of pure CsPbBr₃ NCs (12 $\mu\text{mol g}^{-1} \text{h}^{-1}$). Simultaneously, the CO production of CsPbBr₃@Pb-MOF-2 was about 3.7 and 1.2 times higher than that of CsPbBr₃@Pb-MOF-1 (29 $\mu\text{mol g}^{-1} \text{h}^{-1}$) and CsPbBr₃@Pb-MOF-3 (89 $\mu\text{mol g}^{-1} \text{h}^{-1}$),

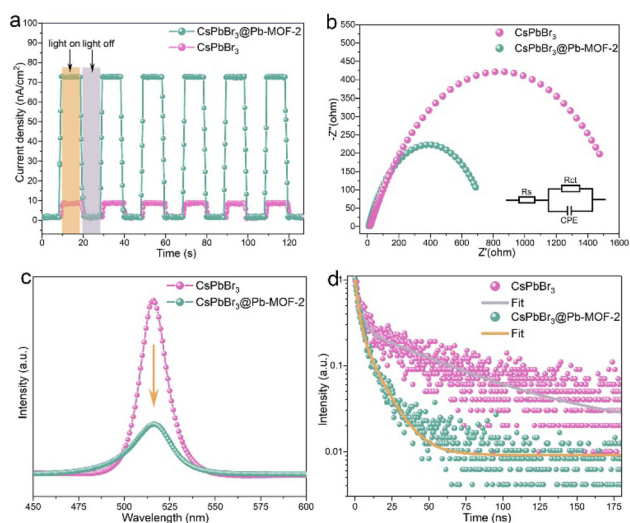


Fig. 5 (a) Photocurrent responses of the samples at -0.4 V vs. Ag/AgCl under visible light irradiation ($\lambda > 420$ nm). (b) Electrochemical impedance spectra of the samples. The inset of (b) shows the equivalent circuit diagram, R_s : solution resistance, R_{ct} : interface charge transfer resistance, CPE: constant phase element. (c) Steady-state PL spectra with a 365 nm excitation wavelength. (d) Time-resolved PL decay plots with a 365 nm excitation wavelength and a 516 nm emission wavelength.

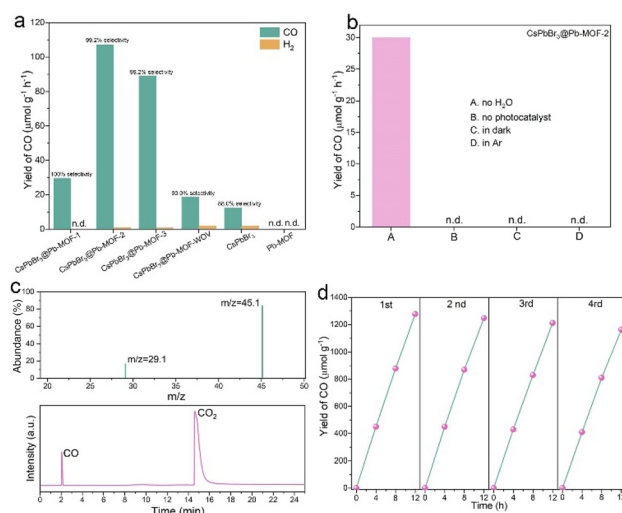
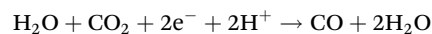


Fig. 6 (a) Photoreduction CO₂ performances of the samples. (b) The photocatalytic performance of CsPbBr₃@Pb-MOF-2 under different conditions. (c) The GC chromatogram and GC-mass result of an isotopic experiment using ¹³CO₂. (d) 48 h recycle photocatalytic CO₂ reduction using CsPbBr₃@Pb-MOF-2.

indicating that an appropriate concentration of oxygen vacancies and perovskite content can optimize CO yield. In contrast, the CO production of CsPbBr₃@Pb-MOF-WOV was only 18 μmol g⁻¹ h⁻¹, highlighting the crucial role of oxygen vacancies in enhancing the CO yield. As for Pb-MOF, no CO or H₂ could be observed. The CO yield of CsPbBr₃@Pb-MOF-2 is also comparable to the reported highest perovskite-based photocatalysts in gas-solid phase and liquid-solid phase photocatalysis (Table S3†). To confirm the actual roles of each component during the photocatalytic reduction of CO₂, control experiments were conducted, as shown in Fig. 6b. When H₂O was removed from the system, the CO yield decreased to 30 μmol g⁻¹ h⁻¹, indicating a significant contribution of H₂O as the electron/proton source. The absence of light and photocatalyst resulted in no detection of CO, affirming the crucial role of both light and photocatalyst in the photoreduction of CO₂. When replacing CO₂ with Ar, no CO was observed, indicating CO did not originate from the organic matter in the MOF. To demonstrate this perspective, isotope labeling tests were carried out (Fig. 6c). On replacing CO₂ with ¹³CO₂ (*m/z* = 45) as the feedstock, an obvious peak of ¹³CO (*m/z* = 29) was detected. Therefore, CO was confirmed to be obtained from the photoreduction of CO₂. Moreover, the photocatalytic reaction was also investigated using H₂¹⁸O as a replacement for H₂O, and the MS results are shown in Fig. S22.† The corresponding signals at *m/z* = 28 and *m/z* = 30 were observed, indicating the formation of C¹⁶O and C¹⁸O, respectively. These findings suggest that the O atom in the produced CO is not solely derived from CO₂, but can also originate from H₂O. This observation is consistent with the find-

ings reported by Zhang *et al.*⁶⁸ Based on the experimental results and the findings of Zhang *et al.*, two possible pathways (Fig. S23 and S24†) for photocatalytic CO₂ reduction to produce CO can be presented as follows.



Given the significance of the stability in the practical application of photocatalysts, a cycling test was performed on CsPbBr₃@Pb-MOF-2. As shown in Fig. 6d, only a 9% loss of photocatalytic activity is incurred after 48 h of recycling of the reaction. Moreover, after a 48 h stability test, CsPbBr₃@Pb-MOF-2 was subjected to characterization through PXRD, XPS, TEM, and UV-Vis (Fig. S25 and S28†). The results demonstrate the exceptional stability of CsPbBr₃@Pb-MOF-2.

3.7 The possible photocatalytic mechanism

A possible mechanism for CO₂ photoreduction over CsPbBr₃@Pb-MOF is proposed based on the above results and analyses, as described in Fig. 7. Firstly, the opportune specific surface area and mesoporous structure of Pb-MOF can benefit from the adsorption and concentration of CO₂ and supply abundant active sites, thereby promoting the subsequent photoreduction reaction. Secondly, under light irradiation, CsPbBr₃ was excited and released photogenerated electrons (e⁻) and photogenerated holes (h⁺). Later, the oxygen vacancies in the Pb-MOF function act as a photogenerated electron acceptor from CsPbBr₃. The improved charge transfer between

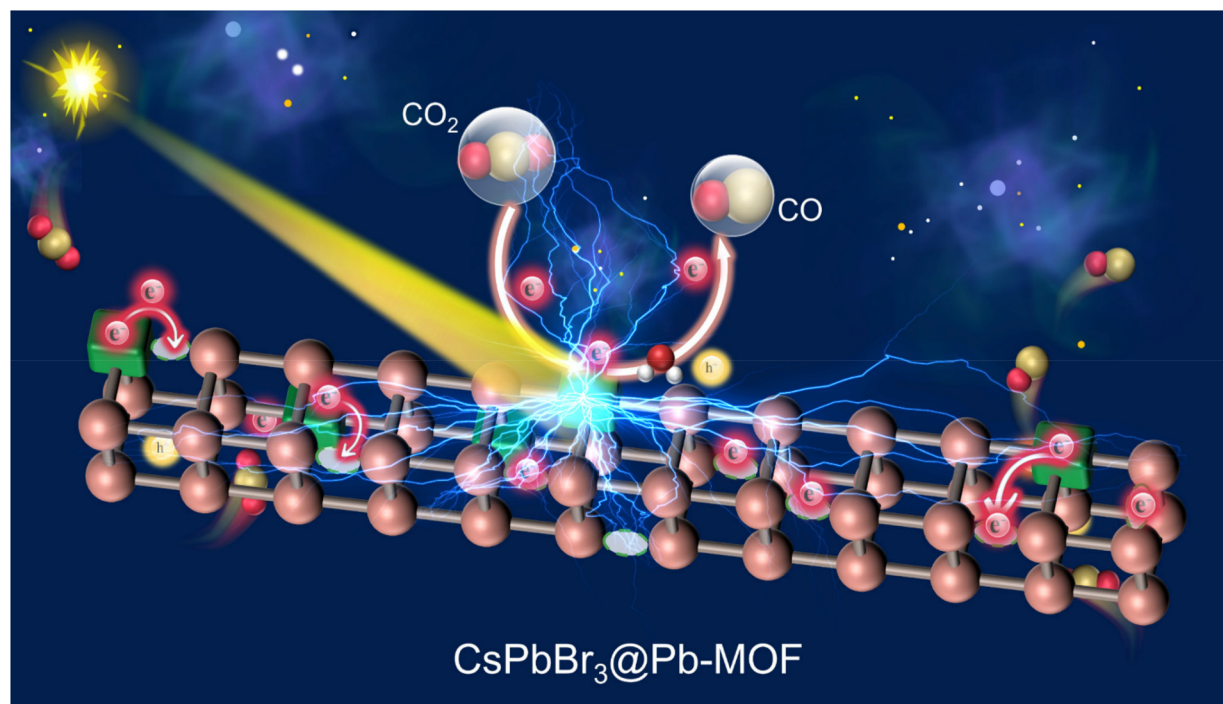


Fig. 7 The possible mechanism of CO₂ photoreduction for CsPbBr₃@Pb-MOF.

CsPbBr₃ and Pb-MOF inhibits the radiative recombination of the photogenerated electrons and holes. Thirdly, the enriched electrons on the surface of Pb-MOF undergo a 2-electron reaction to reduce CO₂ to CO. Meanwhile, a Pb-MOF hydrophobic surface environment with a high concentration of CO₂ boosts the selectivity for CO₂ photoreduction, suppressing the generation of the H₂ byproduct. Finally, the performance of CsPbBr₃ NCs is improved, owing to the synergistic effect of Pb-MOF.

4. Conclusions

In summary, we developed a facile and low-cost *in situ* structural transformation strategy to introduce both CsPbBr₃ and oxygen vacancies into hydrophobic hierarchically porous Pb-MOF by reacting PbBr-MOF with CsBr. During the structural transformation from PbBr-MOF to Pb-MOF, the exposure of Pb metal nodes in Pb-MOF is involved in the construction of CsPbBr₃, which facilitates CsPbBr₃@Pb-MOF preparation, minimizes CsPbBr₃ aggregation, and shortens the photogenerated electron transportation path. Meanwhile, the organic linker omission in structural transformation results in a hierarchically porous structure and abundant oxygen vacancies which enables fast mass transport and accelerates the photo-generated charge separation. The enhanced stability and outstanding electron transfer efficiency permit CsPbBr₃@Pb-MOF-2 to exhibit a significantly impressive photocatalytic performance for CO₂ reduction to CO (107 μmol g⁻¹ h⁻¹, 99.2% selectivity) which is an 8.9-fold improvement over that of pure CsPbBr₃ (12 μmol g⁻¹ h⁻¹, 88% selectivity) under visible-light (λ > 420 nm). Our research provides new insights and offers novel perspectives for the development of efficient lead-based perovskite MOF composite photocatalysts.

Author contributions

The manuscript was written through the contributions of all authors.

Conflicts of interest

There are no conflicts of interest to declare.

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