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# Progress on enhancing the charge separation efficiency of carbon nitride for robust photocatalytic H<sub>2</sub> production

Mengmeng Shao,<sup>a</sup> Yangfan Shao<sup>b</sup> and Hui Pan<sup>c,d</sup>

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Solar-driven H<sub>2</sub> production from water splitting with efficient photocatalysts is a sustainable strategy to meet the clean energy demand and alleviate the approaching environmental issues caused by fossil fuel consumption. Among various semiconductor-based photocatalysts, graphitic carbon nitride (g-C<sub>3</sub>N<sub>4</sub>) has attracted much attention due to its advantages of long term-stability, visible light response, low cost, and easy preparation. However, the intrinsic Coulombic attraction between charge carriers and the interlayer electrostatic barrier of bulk g-C<sub>3</sub>N<sub>4</sub> result in severe charge recombination and low charge separation efficiency. This perspective summarizes the recent progress in the development of g-C<sub>3</sub>N<sub>4</sub> photocatalytic systems, and focuses on three main modification strategies for promoting charge transfer and minimizing charge recombination, including structural modulation, heterojunction construction, and cocatalyst loading. Based on this progress, we provide conclusions regarding the current challenges of further improving photocatalytic efficiency to fulfill commercial requirements, and propose some recommendations for the design of novel and satisfactory g-C<sub>3</sub>N<sub>4</sub> photocatalysts, which is expected to progress the solar-to-hydrogen conversion.

## 1. Introduction

Fuel is an indispensable driving force for the development of human society. In the ancient world, humans bade farewell to a life of devouring raw meat by drilling wood to make fire, and began the era of using firewood as the main energy source. Since the 18th century, with the outbreak of the industrial revolution, fossil fuels, such as coal, oil, and natural gas, have emerged as the main energy source in the world. Nowadays the increasing demand for fossil fuels and various environmental issues (e.g., global warming, acid rain, and smog) caused by their excessive consumption have posed great challenges to the current energy system. Among emerging energy sources, hydrogen has the highest energy density, is eco-friendly, abundant, and carbon-neutral, and is considered a promising alternative to fossil fuels.<sup>1,2</sup>

In recent decades, solar-driven H<sub>2</sub> production from water splitting has received great attention due to the green process only involving sustainable solar energy, water, and photocatalysts.<sup>3–5</sup>

Starting from the pioneering work of Fujishima and Honda about water-splitting on TiO<sub>2</sub>,<sup>6</sup> numerous active semiconductor photocatalysts have been developed, such as CdS, g-C<sub>3</sub>N<sub>4</sub>, BiOBr, Cd<sub>x</sub>Zn<sub>1–x</sub>S, GaN–ZnO solid solution, and ZnIn<sub>2</sub>S<sub>4</sub>.<sup>7–17</sup> The typical process of photocatalytic water-splitting mainly includes three steps (Fig. 1a).<sup>18,19</sup> (1) the generation of photo-induced electron and hole pairs, where the electrons in the valence band (VB) are excited to the conduction band (CB) under light irradiation with higher energy than the band gap of photocatalysts; (2) the separation and transfer of photo-induced charge carriers, which are always accompanied by recombination; (3) the transferred charges participate in the surface catalytic reaction, producing H<sub>2</sub> and O<sub>2</sub> by H<sup>+</sup> reduction and H<sub>2</sub>O oxidation, where the redox potentials of H<sup>+</sup>/H<sub>2</sub> and O<sub>2</sub>/H<sub>2</sub>O should be located between the CB minimum and VB maximum. Based on the above steps, the development of robust photocatalysts is focused on increasing the light-harvesting range, improving charge separation efficiency, and accelerating surface redox reactions. Currently, to promote the practical application of photocatalytic H<sub>2</sub> production, the photocatalysts with the visible light response and high apparent quantum efficiency (AQE) attain significant interest.

As reported, the polymeric semiconductor, graphitic carbon nitride (g-C<sub>3</sub>N<sub>4</sub>), is well known as a photocatalyst for H<sub>2</sub> production because of its advantageous layered structure, long-term stability, ease of synthesis, low cost, and eco-friendly nature. Especially, g-C<sub>3</sub>N<sub>4</sub> with a narrow band gap

<sup>a</sup> School of Materials Science and Engineering, Dongguan University of Technology, Dongguan 523808, China. E-mail: shaomm@dgtu.edu.cn

<sup>b</sup> Institute of Materials Research, Tsinghua Shenzhen International Graduate School, Tsinghua University, Shenzhen 518055, China

<sup>c</sup> Institute of Applied Physics and Materials Engineering, University of Macau, Macao 999078, China. E-mail: huipan@um.edu.mo

<sup>d</sup> Department of Physics and Chemistry, Faculty of Science and Technology, University of Macau, Macao 999078, China

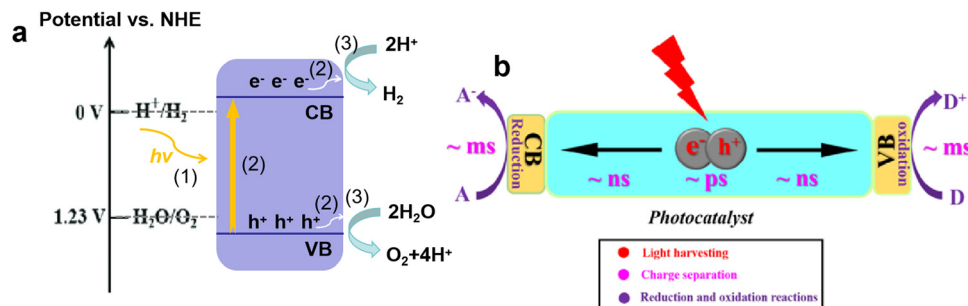


Fig. 1 Photocatalytic H<sub>2</sub> evolution process (a). The kinetic times for separation, transfer, and reaction of photogenerated charge carriers (b). Reproduced with permission from ref. 31. Copyright 2021, Wiley-VCH.

( $\sim 2.7$  eV) is capable of harvesting visible light, and its band edge positions are suitable for water splitting.<sup>20–22</sup> In 2009, Wang *et al.*<sup>23</sup> first showed photocatalytic H<sub>2</sub> production from water using g-C<sub>3</sub>N<sub>4</sub> as the photocatalyst under visible light irradiation. The result of Wang's study motivated the researchers to further explore and improve the photocatalytic H<sub>2</sub> production performance of g-C<sub>3</sub>N<sub>4</sub>. For instance, Zhu and Zhang *et al.*<sup>24</sup> prepared 3D porous g-C<sub>3</sub>N<sub>4</sub> through a bottom-up self-assembly method, achieving photocatalytic overall water splitting with  $\sim 1.4\%$  AQE at 420 nm. Liu and Lu *et al.*<sup>25</sup> designed a ZnIn<sub>2</sub>S<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub> heterojunction composite, which displayed excellent photocatalytic activity with a H<sub>2</sub> production rate of about 87  $\mu\text{mol h}^{-1}$  and AQE of 10.74% at 420 nm. Our group also explored the photocatalytic performance of different cocatalyst loaded g-C<sub>3</sub>N<sub>4</sub>,<sup>26–28</sup> and found the H<sub>2</sub> production rates of VS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> and W<sub>2</sub>C/g-C<sub>3</sub>N<sub>4</sub> both exceeded 80  $\mu\text{mol h}^{-1}$ , while the highest quantum efficiency achieved on VS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> was just 5.5%.

However, although g-C<sub>3</sub>N<sub>4</sub> is a very promising photocatalyst for H<sub>2</sub> production from water splitting, and related studies have made some progress, the quantum efficiency is still too low to meet the solar-to-hydrogen conversion efficiency of 10% required for commercialization. The main reason lies in the high recombination rate of photo-induced charges, leading to less electron involvement in the H<sub>2</sub> evolution reaction.<sup>29,30</sup> For photocatalysis, “no charge, no reaction” is clear, the charge separation and transfer are crucial for photocatalytic performance. As shown in Fig. 1b, due to the strong Coulombic interaction between electron and hole pairs, the charge carriers undergo severe recombination within the bulk and surface phases (several ps  $< t <$  dozens ps), which is faster than the separation of these charges ( $\sim$  ns) and the involved surface reactions ( $\sim$  ms).<sup>31,32</sup> Another challenge for charge separation and transfer is the existence of the interlayer electrostatic barrier in bulk g-C<sub>3</sub>N<sub>4</sub>, which hinders the transfer of electrons from the bulk to surface active sites, and results in severe bulk recombination and poor photocatalytic activity.<sup>33,34</sup> Thus, the efficient charge separation and opposite transfer to surface redox sites are of the utmost importance in photocatalytic reactions.

During the last decade, a series of papers have been published revealing the advances in the field of photocatalytic H<sub>2</sub> production over g-C<sub>3</sub>N<sub>4</sub>.<sup>35–37</sup> As we know, the separation and transfer of charges is a crucial step for photocatalysis.

Currently, various approaches have been developed to modify g-C<sub>3</sub>N<sub>4</sub> and improve charge separation efficiency, such as morphological control, metal/non-metal doping, heterojunction, *etc.*<sup>22</sup> Considering the unique characteristics of each strategy, it is necessary to compare different modification methods and summarize their advantages in a review, which is beneficial for further understanding the mechanism of charge separation and transfer, providing strategies for developing highly active g-C<sub>3</sub>N<sub>4</sub>-based photocatalysts for H<sub>2</sub> production.

Therefore, in this perspective, following the introduction, our focus is to discuss the general strategies for strengthening charge separation and transfer in g-C<sub>3</sub>N<sub>4</sub>, and analyze their respective characteristics. We then extend our discussion to the synergistic effect of combining multiple methods on the improvement of photocatalytic efficiency. Finally, we conclude with some recommendations for the design of benchmark g-C<sub>3</sub>N<sub>4</sub>-based photocatalysts for practical applications.

## 2. Structural modulation of g-C<sub>3</sub>N<sub>4</sub>

As discussed in the introduction, the poor charge separation efficiency of photocatalysts essentially lies in the high Coulombic attraction of photo-induced charges, which makes the dissociation of electron and hole pairs into free charges difficult. For the g-C<sub>3</sub>N<sub>4</sub> photocatalyst, another added problem is the presence of electrostatic barriers between layers, which results in the trapping of charge carriers in the bulk phase and few electrons being transferred to the surface. Thus, to facilitate charge separation and transfer in g-C<sub>3</sub>N<sub>4</sub>, the modulation of intrinsic structure is a commonly adopted strategy. In this section, we focus on discussing several important published studies about structural modulation of g-C<sub>3</sub>N<sub>4</sub> via crystal structure regulation, morphological tuning and doping engineering. A comparison of the photocatalytic H<sub>2</sub> evolution activity of the recent g-C<sub>3</sub>N<sub>4</sub> systems based on structural modulation is given in Table 1.

### 2.1. Crystal structure regulation

g-C<sub>3</sub>N<sub>4</sub> is typically synthesized through the traditional thermal polymerization of organic compounds, such as melamine, dicyandiamide, urea, thiourea, *etc.*<sup>38</sup> The simple one-step polymerization method often causes incomplete polycondensation,

Table 1 Photocatalytic H<sub>2</sub> evolution activity of recent g-C<sub>3</sub>N<sub>4</sub> systems based on structural modulation

| Photocatalyst                                          | Light source                        | Reaction conditions                                                              | H <sub>2</sub> evolution rate                    | AQE              | Ref. |
|--------------------------------------------------------|-------------------------------------|----------------------------------------------------------------------------------|--------------------------------------------------|------------------|------|
| Crystalline g-C <sub>3</sub> N <sub>4</sub> nanosheets | 300 W Xe lamp ( $\lambda > 420$ nm) | 25 mg catalyst, TEOA <sup>f</sup> solution, Pt cocatalyst                        | 20.9 mmol h <sup>-1</sup> g <sup>-1</sup>        | 73.6% at 420 nm  | 44   |
| Single-crystalline PTI <sup>a</sup> nanosheets         | 300 W Xe lamp ( $\lambda > 300$ nm) | 100 mg catalyst, pure water, CoO <sub>x</sub> and PtCoO <sub>x</sub> cocatalysts | 6100 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>   | 25% at 365 nm    | 45   |
| CN-TM <sup>b</sup>                                     | 300 W Xe lamp ( $\lambda > 420$ nm) | 50 mg catalyst, TEOA solution, Pt cocatalyst                                     | 3135 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>   | 21.03% at 420 nm | 47   |
| Prism-like PTI                                         | 300 W Xe lamp ( $\lambda > 300$ nm) | 100 mg catalyst, pure water, Pt/Co cocatalysts                                   | 1890 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>   | 8% at 365 nm     | 48   |
| g-C <sub>3</sub> N <sub>4</sub> nanosheets             | 300 W Xe lamp ( $\lambda > 400$ nm) | 50 mg catalyst, TEOA solution, Pt cocatalyst                                     | 10.14 mmol h <sup>-1</sup> g <sup>-1</sup>       | 7.34% at 400 nm  | 49   |
| g-C <sub>3</sub> N <sub>4</sub> nanorods               | 300 W Xe lamp ( $\lambda > 400$ nm) | 20 mg catalyst, TEOA solution, Pt cocatalyst                                     | 2590 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>   | 10.4% at 400 nm  | 50   |
| g-C <sub>3</sub> N <sub>4</sub> nanotubes              | 300 W Xe lamp ( $\lambda > 420$ nm) | 10 mg catalyst, TEOA solution, Pt cocatalyst                                     | 578.5 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>  | —                | 54   |
| g-C <sub>3</sub> N <sub>4</sub> tubes                  | 300 W Xe lamp ( $\lambda > 420$ nm) | 25 mg catalyst, TEOA solution, Pt cocatalyst                                     | 732 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>    | 7.1% at 420 nm   | 56   |
| C <sub>3</sub> N <sub>4</sub> -TDA <sup>c</sup>        | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution, Pt cocatalyst                                     | 4779.8 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup> | —                | 60   |
| g-C <sub>3</sub> N <sub>4</sub>                        | 350 W Xe lamp                       | 30 mg catalyst, TEOA solution, Ni <sub>2</sub> P cocatalyst                      | 19.25 mmol h <sup>-1</sup> g <sup>-1</sup>       | 5.8% at 400 nm   | 62   |
| g-C <sub>3</sub> N <sub>4</sub>                        | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution, Pt cocatalyst                                     | 12.35 mmol h <sup>-1</sup> g <sup>-1</sup>       | 40.5% at 400 nm  | 69   |
| g-C <sub>3</sub> N <sub>4</sub>                        | 350 W Xe lamp                       | 50 mg catalyst, Na <sub>2</sub> S/Na <sub>2</sub> SO <sub>3</sub> solution       | 14.9 mmol h <sup>-1</sup> g <sup>-1</sup>        | 19.9% at 420 nm  | 70   |
| ATZ-C <sub>3</sub> N <sub>4</sub>                      | Xe lamp AM1.5 ( $\lambda > 420$ nm) | 50 mg catalyst, TEOA solution, Pt cocatalyst                                     | 10.8 mmol h <sup>-1</sup> g <sup>-1</sup>        | 10.35% at 420 nm | 71   |
| CN-ATZ-NaK <sup>e</sup>                                | 50 W LED ( $\lambda > 420$ nm)      | 50 mg catalyst, TEOA and K <sub>2</sub> HPO <sub>4</sub> solution, Pt cocatalyst | 13 mmol h <sup>-1</sup> g <sup>-1</sup>          | 65% at 420 nm    | 73   |
| PHI-A <sup>f</sup>                                     | 300 W Xe lamp                       | 50 mg catalyst, TEOA solution, Pt cocatalyst                                     | 4240 $\mu$ mol h <sup>-1</sup> g <sup>-1</sup>   | 12% at 420 nm    | 74   |
| Li-g-C <sub>3</sub> N <sub>4</sub>                     | 350 W Xe lamp                       | 30 mg catalyst, TEOA solution                                                    | 14 mmol h <sup>-1</sup> g <sup>-1</sup>          | 22.9% at 420 nm  | 78   |
| In-g-C <sub>3</sub> N <sub>4</sub>                     | 300 W Xe lamp ( $\lambda > 420$ nm) | 100 mg catalyst, TEOA solution, Pt cocatalyst                                    | 1.35 mmol h <sup>-1</sup> g <sup>-1</sup>        | —                | 80   |
| Cu-g-C <sub>3</sub> N <sub>4</sub>                     | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, methanol solution, Pt cocatalyst                                 | 10.6 mmol h <sup>-1</sup> g <sup>-1</sup>        | 9.2% at 420 nm   | 81   |

<sup>a</sup> Poly(triazine imide) (PTI) based carbon nitride. <sup>b</sup> The precursors of thiourea (T) and melamine (M) for g-C<sub>3</sub>N<sub>4</sub> (CN). <sup>c</sup> Thiophene-2,5-dicarbaldehyde (TDA). <sup>d</sup> Aromatized terminal. <sup>e</sup> 5-aminotetrazole (ATZ) is used to prepare g-C<sub>3</sub>N<sub>4</sub> in NaCl/KCl molten salt. <sup>f</sup> The precursor of 2-aminothiophene-3-carbonitrile (A) for poly-heptazine-imide (PHI) based carbon nitride. <sup>g</sup> Triethanolamine.

a low degree of polymerization, and poor crystallinity of g-C<sub>3</sub>N<sub>4</sub>, which bring about many internal and surface defects, thereby increasing the charge recombination centers and resulting in poor photocatalytic activity (Fig. 2).<sup>39,40</sup> Therefore, it is necessary to prepare highly crystalline g-C<sub>3</sub>N<sub>4</sub> with less defects for enhanced charge separation efficiency and efficient photocatalytic reaction.

Many strategies have been developed to promote the polymerization of precursors and synthesize the highly crystalline g-C<sub>3</sub>N<sub>4</sub>, including salt-assisted methods, microwave synthesis, two-step calcination, *etc.*<sup>41–43</sup> For example, Song *et al.*<sup>44</sup> first prepared ultrathin g-C<sub>3</sub>N<sub>4</sub> nanosheets by air-assisted thermal exfoliation, then the resulting nanosheets were treated in the NaCl/KCl molten salts to obtain the highly crystalline g-C<sub>3</sub>N<sub>4</sub> nanosheets based on the heptazine-triazine units (donor-acceptor motifs, D-A). The high crystallinity and formed D-A structure promoted charge separation and transfer, resulting in a much higher H<sub>2</sub> evolution activity of 20.9 mmol h<sup>-1</sup> g<sup>-1</sup> with an AQE of 73.6% at 420 nm compared with the original g-C<sub>3</sub>N<sub>4</sub> nanosheets (1.0 mmol h<sup>-1</sup> g<sup>-1</sup>). Wang and co-workers synthesized a single-crystalline carbon nitride, poly(triazine imide) nanosheets (PTI), through ternary molten salt treatment (LiCl/KCl/NaCl).<sup>45</sup> Compared to binary molten salt systems (LiCl/KCl or NaCl/LiCl), the ternary eutectic salts could provide a more mild polymerization environment with higher melting points, which was beneficial to the in-plane polymerization and extension of the conjugated structure, leading to the formation of single-crystalline PTI nanosheets with less structure defects and a shorter charge migration distance. Thus, the obtained PTI nanosheets exhibited a high photocatalytic activity with an AQE of 25% at 365 nm for overall water splitting. In addition, compared with the traditional calcination process, the microwave method can quickly increase the temperature and trigger strong collisions between the reactants, thereby promoting the polymerization reaction and resulting in highly crystalline and active g-C<sub>3</sub>N<sub>4</sub>.<sup>46</sup> Chen and Yang *et al.*<sup>47</sup> combined the molten-

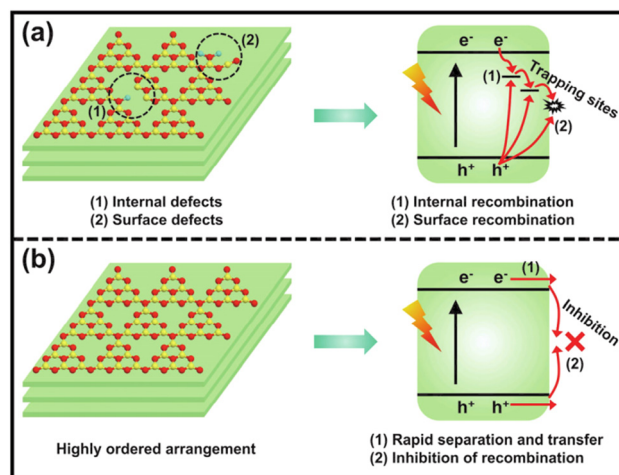


Fig. 2 The charge separation and transfer of conventional poorly-crystalline g-C<sub>3</sub>N<sub>4</sub> (a) and highly ordered g-C<sub>3</sub>N<sub>4</sub> (b). Reproduced with permission from ref. 39. Copyright 2023, Elsevier.

salt strategy and microwave method to prepare highly active triazine-heptazine-based  $g\text{-C}_3\text{N}_4$  with a  $\text{H}_2$  evolution rate of  $3135 \mu\text{mol h}^{-1} \text{g}^{-1}$  (AQE of 21.03% at 420 nm), which was obviously higher than that of  $g\text{-C}_3\text{N}_4$  treated with traditional high-temperature calcination ( $2591 \mu\text{mol h}^{-1} \text{g}^{-1}$ ).

Apart from crystallinity regulation, recently, Wang and co-workers investigated the influence of crystal planes on photocatalytic water splitting.<sup>48</sup> They prepared the highly crystalline polytriazine imides-based carbon nitride (PTI) with a prism-like shape and different facets *via* the KCl/LiCl molten salt-assisted method. The results of *in-situ* photodeposition of the Pt cocatalyst demonstrated that the prismatic planes (1010) were the main reactive facets for catalytic reaction, as the Pt cocatalyst tended to accumulate on the (1010) planes. Furthermore, the calculation results also indicated that the electrons could easily transfer toward the prismatic (1010) planes. Based on this discovery, PTI catalysts with different surface area ratios of the (1010) planes were designed, and it was found that the photocatalytic activity increased with the increasing surface area ratio of the (1010) planes, and the optimal AQE of overall water splitting reached 8% at 365 nm with a  $\text{H}_2$  evolution rate of  $1890 \mu\text{mol h}^{-1} \text{g}^{-1}$ . Thus, the development of the crystalline  $g\text{-C}_3\text{N}_4$ -based system with less internal defects and more reactive facets exposed can improve charge separation efficiency and achieve the desired photocatalytic activity.

## 2.2. Morphological tuning

For the severe charge recombination in bulk  $g\text{-C}_3\text{N}_4$ , the design of ultrathin  $g\text{-C}_3\text{N}_4$  nanosheets with expanded layer spacing is

an effective method to weaken or break the shackle of interlayer electrostatic barriers and promote electron transfer from the bulk to the surface. For example, Yang and Kong *et al.*<sup>49</sup> prepared a series of porous  $g\text{-C}_3\text{N}_4$  nanosheets with different thicknesses (2–20 nm) through a thermal-triggering *in-situ* gas-shocking method. The most exfoliated  $g\text{-C}_3\text{N}_4$  nanosheets presented an ultrathin structure with only 2 nm thickness, resulting in a larger interlayer spacing. The results showed that the ultrathin  $g\text{-C}_3\text{N}_4$  nanosheets ( $\sim 2$  nm) exhibited the highest photocatalytic activity with a  $\text{H}_2$  production rate up to  $10.14 \text{ mmol h}^{-1} \text{g}^{-1}$  (AQE of 7.34% at 400 nm), which was about 57 times that of bulk  $g\text{-C}_3\text{N}_4$ . The obviously enhanced photocatalytic activity was attributed to the fact that the ultrathin structure not only weakens the interlayer electrostatic barrier and promotes charge transfer to the surface, but also shortens the charge migration distance, thereby achieving a high charge separation efficiency. Shi *et al.*<sup>50</sup> also reported the significantly improved photocatalytic  $\text{H}_2$  production activity over  $g\text{-C}_3\text{N}_4$  nanosheets. They developed an available bottom-up acidification method to prepare the ultrathin  $g\text{-C}_3\text{N}_4$  nanosheets with an average thickness of  $\sim 3$  nm (Fig. 3a–d). As a result, the as-prepared  $g\text{-C}_3\text{N}_4$  nanosheets not only showed a high AQE at 400 nm (10.4%) and 420 nm (8.4%), but also exhibited an enhanced photocatalytic  $\text{H}_2$  evolution rate of  $2590 \mu\text{mol h}^{-1} \text{g}^{-1}$ , which was nearly 10-fold higher than that of bulk  $g\text{-C}_3\text{N}_4$ . The combined characterizations of electrochemical impedance spectra (EIS), photocurrent responses, and time-resolved PL spectra clearly demonstrated that the



Fig. 3 TEM images of bulk  $g\text{-C}_3\text{N}_4$  (a) and ultrathin  $g\text{-C}_3\text{N}_4$  nanosheets (b). AFM image of  $g\text{-C}_3\text{N}_4$  nanosheets (c) and the corresponding height curves determined along the red line (d). UV-vis absorption spectra (e) and Tauc plots of the as-prepared  $g\text{-C}_3\text{N}_4$  materials (f), where CN-B and UCNs correspond to bulk  $g\text{-C}_3\text{N}_4$  and ultrathin  $g\text{-C}_3\text{N}_4$  nanosheets, respectively. Reproduced with permission from ref. 50. Copyright 2020, Elsevier.

excellent photocatalytic activity of g-C<sub>3</sub>N<sub>4</sub> nanosheets was due to the improved charge transfer and separation as well as the prolonged charge lifetime.

Although the nanosheet structure can effectively reduce the charge recombination in the bulk phase of g-C<sub>3</sub>N<sub>4</sub>, and shorten the charge migration path to facilitate the subsequent catalytic reaction, the quantum confinement effect of nanoscale g-C<sub>3</sub>N<sub>4</sub> can lead to an increase in its bandgap, which weakens the visible light response ability of g-C<sub>3</sub>N<sub>4</sub> and limits the improvement of photocatalytic performance.<sup>50–52</sup> For instance, the light absorption edges of the g-C<sub>3</sub>N<sub>4</sub> nanosheets prepared by Shi *et al.*<sup>50</sup> displayed a blue shift (Fig. 3e and f), implying the decreased light absorption range and enlarged bandgap of nanoscale g-C<sub>3</sub>N<sub>4</sub>. The corresponding Tauc plots clearly showed that the bandgap of the nanosheet structure (2.67 eV) was larger than that of the bulk structure (2.34 eV).

In addition to 2D nanosheet structures, 1D g-C<sub>3</sub>N<sub>4</sub> nanorods or nanotubes have also received extensive attention due to their structural advantages in facilitating charge transfer along the axis.<sup>53–55</sup> Zhu's group developed a network structure assembled by g-C<sub>3</sub>N<sub>4</sub> nanorods for H<sub>2</sub> evolution and phenol degradation.<sup>54</sup> The higher intensity of photocurrent response and smaller arc radius in the Nyquist plots of the g-C<sub>3</sub>N<sub>4</sub> nanorod network determined the obviously enhanced charge separation and transfer efficiency. Thus, compared to bulk g-C<sub>3</sub>N<sub>4</sub>, the g-C<sub>3</sub>N<sub>4</sub> nanorod

system displayed higher photocatalytic activity with a H<sub>2</sub> evolution rate of 578.5  $\mu\text{mol h}^{-1} \text{g}^{-1}$ , a nearly six-fold increase. Liu *et al.*<sup>56</sup> prepared the oxygen doped porous g-C<sub>3</sub>N<sub>4</sub> nanorods, and found that the obtained g-C<sub>3</sub>N<sub>4</sub> nanorods with Pt loading displayed an excellent photocatalytic H<sub>2</sub> evolution rate of 732  $\mu\text{mol h}^{-1} \text{g}^{-1}$ , nearly 6 times higher than the bulk g-C<sub>3</sub>N<sub>4</sub> system. Furthermore, an AQE of 7.1% was achieved at 420 nm in TEOA solution, 0.29% in pure water. Compared with g-C<sub>3</sub>N<sub>4</sub> nanorods, the nanotubes not only have the advantages of a 1D structure, but also promote the directional migration of photo-induced electrons from concave to convex due to the formation of apparent potential difference between the inner and outer surface of the tubular structure.<sup>57–59</sup> Meanwhile, the hollow tubular structure can enhance light scattering and improve the light response of g-C<sub>3</sub>N<sub>4</sub> nanotubes.<sup>60,61</sup> For example, Song *et al.*<sup>62</sup> prepared g-C<sub>3</sub>N<sub>4</sub> square tubes and loaded Ni<sub>2</sub>P as a cocatalyst on the surface of the g-C<sub>3</sub>N<sub>4</sub> tubes (Fig. 4). The calculation and characterization results confirmed the directional electron transfer from the inner to the outer surface of the g-C<sub>3</sub>N<sub>4</sub> tubes, and the increased light response range. Thus, the photocatalytic H<sub>2</sub> evolution activity of the Ni<sub>2</sub>P loaded g-C<sub>3</sub>N<sub>4</sub> tube was much higher than that of pristine g-C<sub>3</sub>N<sub>4</sub> supported with noble metals.

### 2.3. Doping engineering

Doping engineering is another general strategy for regulating the intrinsic structure of g-C<sub>3</sub>N<sub>4</sub>. The incorporation of



Fig. 4 Schematic illustration of spatial separation and utilization of photo-induced electrons in bulk g-C<sub>3</sub>N<sub>4</sub> (a), g-C<sub>3</sub>N<sub>4</sub> tube (b), and Ni<sub>2</sub>P/g-C<sub>3</sub>N<sub>4</sub> tube (c); photocatalytic H<sub>2</sub> evolution activities at different wavelengths for bulk g-C<sub>3</sub>N<sub>4</sub> and g-C<sub>3</sub>N<sub>4</sub> tube systems (d) and (e); the inset of (e) is a FESEM image of the g-C<sub>3</sub>N<sub>4</sub> tube. Reproduced with permission from ref. 42. Copyright 2020, Wiley-VCH.

heteroatoms in g-C<sub>3</sub>N<sub>4</sub> or grafting other delocalization groups would damage the planar  $\pi$ -conjugated structure and lead to differentiation of delocalized  $\pi$ -electron distribution, thereby promoting the directional transfer of photogenerated charges along the potential difference.<sup>63,64</sup> Based on the typical 2D stacking structure of g-C<sub>3</sub>N<sub>4</sub>, there are two main doping configurations, in-plane doping and interlayer doping.

It is known that the high geometrical symmetry of the planar tri-s-triazine structure of g-C<sub>3</sub>N<sub>4</sub> brings about the uniform  $\pi$ -electron distribution,<sup>65</sup> which results in random charge transfer within the plane. Thus, grafting g-C<sub>3</sub>N<sub>4</sub> with diverse delocalization groups by replacing the tri-s-triazine structures or triazine hexagons can induce the in-plane  $\pi$ -electron redistribution and the formation of the intrinsic polarization electric field to favor directional transfer of photo-induced charge carriers. The widely investigated grafting groups include aromatic motifs, carbon nitride allotropes, and C-chains.<sup>66–68</sup> Xu *et al.*<sup>69</sup> constructed a D-A system by incorporating thiophene-2,5-dicarbaldehyde (TDA) into the g-C<sub>3</sub>N<sub>4</sub> network, and found that the modified g-C<sub>3</sub>N<sub>4</sub> displayed an obviously improved photocatalytic performance for bisphenol A degradation and H<sub>2</sub> evolution from water-splitting. The experimental and calculation results (Fig. 5) showed that the grafted TDA molecule could act as an electron acceptor to promote intermolecular charge transfer, and the incorporated aromatic ring from TDA enhanced the  $\pi$ -delocalization as well as the light response, which were responsible for the remarkable improvement of photocatalytic activity. Song *et al.*<sup>70</sup> designed an allotrope of g-C<sub>3</sub>N<sub>4</sub>, g-C<sub>2</sub>N<sub>3</sub>, where the aromatic azide pentagons were incorporated into the triazine hexagons. The theoretical simulations indicated that the positive and negative charge centers demonstrated significant dislocation, which resulted in the formation of the intrinsic polarization electric field to break the Coulomb interaction of electron-hole pairs. Consequently, the designed g-C<sub>2</sub>N<sub>3</sub> exhibited a very high photocatalytic H<sub>2</sub> evolution rate of 14.9 mmol h<sup>−1</sup> g<sup>−1</sup> without any cocatalysts (AQE = 19.9% at 420 nm), which was even 2.6 times that of the g-C<sub>3</sub>N<sub>4</sub> loaded

with Pt. Moreover, Li and Wang *et al.*<sup>71</sup> reported a gas-solid grafting method to prepare an asymmetric structure of g-C<sub>3</sub>N<sub>4</sub> with aromatized terminals. The research demonstrated that the aromatization of the terminals enhanced the asymmetry of  $\pi$ -electron distribution, thereby providing a driving force to achieve the efficient charge transport and separation and improving the photocatalytic activity by 15.4 fold. Recently, Zhang and Wang *et al.*<sup>72</sup> developed the ionothermal method *via* a molten salt mixture to provide a mild liquid environment for polymerization, then obtained the well-organized and highly active carbon nitride. Based on this, they reported the copolymerization of 5-amino-tetrazole with nucleobases in salt melts, and synthesized the allotropes of g-C<sub>3</sub>N<sub>4</sub>, poly-heptazine-imide (PHI) based carbon nitrides.<sup>73</sup> The incorporation of nucleobases induced the formation of the D-A structure, which greatly promoted charge transfer and separation, thus, the PHI-based carbon nitrides showed higher H<sub>2</sub> evolution activity than the melon-based counterparts. Then they further combined the copolymerization and ionothermal method to polymerize urea with a thiophene derivative (2-aminothiophene-3-carbonitrile) and obtained the condensed allotrope of g-C<sub>3</sub>N<sub>4</sub> with separated positive and negative charge centers, which displayed enhanced light absorption and charge separation efficiency. The optimal photocatalytic H<sub>2</sub> evolution activity was about 212.3  $\mu\text{mol h}^{-1}$  with an AQE of 12% at 420 nm, much higher than the pure urea-derived counterpart (81.8  $\mu\text{mol h}^{-1}$ ).<sup>74</sup>

In addition to the modification of molecules or functional groups, the incorporation of metals into the cage structure among heptazine rings is also frequently applied to regulate the electronic structure of g-C<sub>3</sub>N<sub>4</sub>. Currently, the reported doping metals mainly include transition metals, alkali metals, alkaline earth metals, and rare earth metals.<sup>75–77</sup> According to coordination and orbital theory, the lone electron pairs of sp<sup>2</sup>-hybridized N atoms at the cage site can couple with the unoccupied orbital of metals, which can induce the in-plane  $\pi$ -electron redistribution and promote charge directional transfer. For instance, Song *et al.*<sup>78</sup> incorporated Li ions into the cage sites of g-C<sub>3</sub>N<sub>4</sub>



Fig. 5 The possible structure of TDA grafted g-C<sub>3</sub>N<sub>4</sub> (a); the electronic structures of the HOMO and LUMO for (b) g-C<sub>3</sub>N<sub>4</sub> derived from urea (UCN) and (c) TDA grafted g-C<sub>3</sub>N<sub>4</sub> (UCN-xTDA); (d) the HOMO and LUMO levels for UCN and UCN-xTDA. Reproduced with permission from ref. 69. Copyright 2022, Royal Society of Chemistry.

to change the uniform  $\pi$ -electron distribution (Fig. 6a and b). The mulliken charge results showed that the electrons preferentially concentrated around Li and C atoms at the triangular cave, implying the formation of an in-plane coupling electric field, which favored charge directional transfer. As expected, the Li doped  $g\text{-C}_3\text{N}_4$  exhibited a much lower charge recombination rate and higher  $\text{H}_2$  evolution activity of  $0.418 \text{ mmol h}^{-1}$  ( $\sim 14000 \mu\text{mol h}^{-1} \text{ g}^{-1}$ , AQE = 22.9% at 420 nm) compared with pristine  $g\text{-C}_3\text{N}_4$  ( $0.009 \text{ mmol h}^{-1}$ ).

It is noted that the size of the triangular cave among the tri-s-triazine units of  $g\text{-C}_3\text{N}_4$  is about  $0.47 \text{ nm}$ ,<sup>79</sup> which is suitable for accommodating metals with smaller atomic diameters, such as Li and Na, *etc.* The large-sized alkali/alkaline earth metals, such as K, Rb, and Cs, would be incorporated into the interlayers of  $g\text{-C}_3\text{N}_4$ . Meanwhile, the transition metals with different shapes and directions of d orbitals can also be doped between layers instead of the triangular cave, due to the exclusion between the complex d orbitals and N  $2p_z$  orbital at the cave site. Thus, the electronic interaction between the interlayer metals and the layers would induce the formation of an interlayer charge transport channel, which favors charge transfer among the layers of  $g\text{-C}_3\text{N}_4$ . Yang *et al.*<sup>80</sup> reported a series of  $\text{In}^{3+}$  doped  $g\text{-C}_3\text{N}_4$  *via* a facile *in-situ* thermal copolymerization approach, and found that the  $\text{In}^{3+}$  ions were more easily doped into the interlayer of  $g\text{-C}_3\text{N}_4$  due to the lower calculated energy of the out-of-plane configuration (Fig. 6c and d). The differential charge density result showed an obvious difference in electron density between doped In and  $g\text{-C}_3\text{N}_4$ , which facilitated the charge migration in  $g\text{-C}_3\text{N}_4$ . Therefore, the optimal  $\text{In}^{3+}$  doped  $g\text{-C}_3\text{N}_4$  exhibited an excellent photocatalytic  $\text{H}_2$  evolution activity of  $1.35 \text{ mmol h}^{-1} \text{ g}^{-1}$ , 17 fold that of pristine  $g\text{-C}_3\text{N}_4$ . Fu *et al.*<sup>81</sup> incorporated  $g\text{-C}_3\text{N}_4$  with single-atom Cu, which was bonded with compositional N in two ways, in-plane bonding and inter-layer bonding. The characterization results demonstrated that

two groups of  $\text{Cu-N}_x$  as different charge transfer channels could promote the in-plane and interlayer charge transport, respectively, thereby greatly improving photocatalytic performance. Consequently, the Cu decorated  $g\text{-C}_3\text{N}_4$  displayed a robust  $\text{H}_2$  production rate of  $10.6 \text{ mmol h}^{-1} \text{ g}^{-1}$  (AQE = 9.2% at 420 nm), which was about 30 times higher than that of bulk  $g\text{-C}_3\text{N}_4$ .

### 3. Heterojunction construction

With regard to structural modulation of  $g\text{-C}_3\text{N}_4$ , *e.g.* the morphology and doping engineering discussed above, a driving force based on the structural potential difference is introduced to fight against the Coulomb interaction rather than directly weakening the Coulomb field strength. In fact, it is also considered as an efficient strategy to promote the dissociation of excitons (electron-hole pairs) into free charge carriers through weakening the Coulomb effect. According to the Coulomb field equation,  $F_c = kq_e q_h / r^2$ , the Coulombic attraction of electron-hole pairs decreases with increasing the bandgap of semiconductor photocatalysts (*i.e.*, increase in  $r$  value).<sup>82</sup> Meanwhile, the semiconductors with large bandgap usually possess higher CB and lower VB positions, which facilitate the redox reactions. Thus, the wide bandgap property of photocatalysts is desired for weakening the Coulomb field strength and enhancing the redox ability, thereby improving the photocatalytic efficiency. However, a wide bandgap also leads to a weakened light response, and reduces the generation of photo-induced charges. Therefore, it is incompatible for single-component photocatalysts to possess a weakened Coulomb effect, sufficient redox capability, and large light-harvesting range, noticeably which are expected to be achieved by designing heterojunction photocatalysts.

For heterojunction systems, they are typically composed of two semiconductor materials with different band structures, which can be categorized as a straddling gap heterojunction (type I), broken gap heterojunction (type III), and staggered gap heterojunction (type II and Z-scheme).<sup>83,84</sup> In type I heterojunctions (Fig. 7a), the two semiconductors exhibit nested band alignment. The photo-induced electrons and holes on semiconductor A migrate to semiconductor B, resulting in charge accumulation on semiconductor B, which in turn triggers the high recombination of electron-hole pairs. For type III heterojunctions (Fig. 7b), a broken gap alignment is observed, and the photogenerated electrons and holes cannot migrate or transfer to any semiconductor, which also leads to a high charge recombination rate. As for type II heterojunctions (Fig. 7c), the staggered gap between two semiconductors facilitates the spatial separation of electron-hole pairs. As we see, semiconductor A possesses higher CB and VB levels, its photo-induced electrons can transfer to the CB of semiconductor B, while the holes on the VB transfer in the opposite direction. Thus, the transferred electrons and holes accumulate on semiconductor B and semiconductor A, respectively, which realizes the spatial separation of charges. However, in type II heterojunctions, the original electrons and holes would repel the transferred charge



**Fig. 6** Equipotential curve distribution on  $g\text{-C}_3\text{N}_4$  (a) and Li doped  $g\text{-C}_3\text{N}_4$  (b). Reproduced with permission from ref. 78. Copyright 2020, Elsevier. The optimized structure of  $g\text{-C}_3\text{N}_4$  with interlayer In doping (c); differential charge density of In doped  $g\text{-C}_3\text{N}_4$  (d), where the yellow and green isosurfaces represent the gain and loss of electrons, respectively. Reproduced with permission from ref. 80. Copyright 2021, Royal Society of Chemistry.

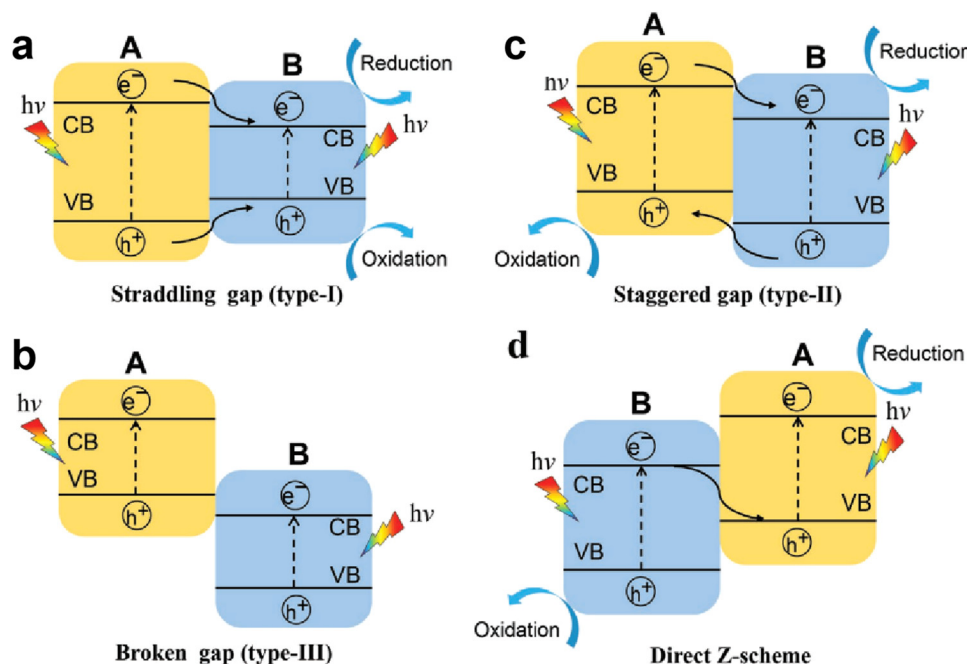


Fig. 7 Schematic illustration of the transfer paths of photo-induced charge carriers in different heterojunction systems: (a) straddling gap heterojunction (type I), (b) broken gap heterojunction (type III), and (c) and (d) staggered gap heterojunction (type II and Z-scheme). Reproduced with permission from ref. 83. Copyright 2020, Wiley-VCH.

carriers, and the work function difference of semiconductors also influences charge transfer, thus, a Z-scheme heterojunction with a band structure similar to type II heterojunctions is proposed as well (Fig. 7d). Typically, a built-in electric field (BIEF) is formed at the interface of semiconductors due to the work function difference, then the useless photogenerated electrons and holes recombine driven by the BIEF.<sup>85</sup> In this way, the remaining electrons and holes are left on semiconductor A and semiconductor B, respectively, achieving the effective separation of charge carriers and favorable redox capability for reaction. In this section, we focus on reviewing the charge separation and transfer processes in two heterojunction systems, Type II heterojunction and Z-scheme heterojunction. A comparison of the photocatalytic H<sub>2</sub> evolution activity of recent g-C<sub>3</sub>N<sub>4</sub>-based heterojunctions is given in Table 2.

### 3.1. Type II heterojunction

Among various g-C<sub>3</sub>N<sub>4</sub> heterojunctions, type II g-C<sub>3</sub>N<sub>4</sub> heterojunctions are the most common system. According to the charge transfer mechanism of the type II system,<sup>86,87</sup> the CB and VB positions of g-C<sub>3</sub>N<sub>4</sub> should be both higher or lower than those of another semiconductor in heterojunctions, thereby forming a staggered band structure that thermodynamically favors the subsequent charge transfer and realizes the spatial separation of electron-hole pairs. As a result, there are numerous studies about developing g-C<sub>3</sub>N<sub>4</sub> heterojunctions with type II alignment to promote charge transfer during photocatalytic reaction, e.g., CuFe<sub>2</sub>O<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>88</sup> ZnIn<sub>2</sub>S<sub>4</sub>/S-doped g-C<sub>3</sub>N<sub>4</sub>,<sup>89</sup> g-C<sub>3</sub>N<sub>4</sub>/r-TiO<sub>2</sub>,<sup>90</sup> CsCu<sub>2</sub>I<sub>3</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>91</sup> β-AsP/g-C<sub>3</sub>N<sub>4</sub>,<sup>92</sup> Cu<sub>2</sub>O/g-C<sub>3</sub>N<sub>4</sub>,<sup>93</sup> B-doped C<sub>3</sub>N<sub>4</sub>/ZnO,<sup>94</sup> CoO<sub>x</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>95</sup> CeO<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>96</sup> CdS/g-C<sub>3</sub>N<sub>4</sub>,<sup>97</sup> Zn<sub>x</sub>-

Cd<sub>1-x</sub>In<sub>2</sub>S<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>98</sup> g-C<sub>3</sub>N<sub>4</sub>/SnS<sub>2</sub>,<sup>99</sup> g-C<sub>3</sub>N<sub>4</sub>/Zn-MOF,<sup>100</sup> and CdSe/WS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub>.<sup>101</sup> Ahmad *et al.*<sup>88</sup> developed a type II heterostructure of g-C<sub>3</sub>N<sub>4</sub> combined with CuFe<sub>2</sub>O<sub>4</sub>, and found the CuFe<sub>2</sub>O<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub> showed a 2.5-fold enhancement in photocatalytic activity and high AQE for H<sub>2</sub> evolution (25% at 450 nm). The obviously improved photocatalytic activity was mainly attributed to the thermodynamically optimal band alignment that facilitated directional charge transfer and obtained spatially separated electrons (on g-C<sub>3</sub>N<sub>4</sub>) and holes (on CuFe<sub>2</sub>O<sub>4</sub>) for efficient catalytic reaction. Xie and Luo *et al.*<sup>89</sup> utilized a simple *in-situ* hydrothermal method to construct S-doped g-C<sub>3</sub>N<sub>4</sub> with ZnIn<sub>2</sub>S<sub>4</sub> as a type II photocatalyst for splitting H<sub>2</sub>O into H<sub>2</sub>. Compared to single S-doped g-C<sub>3</sub>N<sub>4</sub> and ZnIn<sub>2</sub>S<sub>4</sub>, the constructed heterojunction displayed a much higher photocatalytic H<sub>2</sub> production performance of 1.63 mmol h<sup>-1</sup> g<sup>-1</sup> (AQE = 0.90% at 420 nm), which was respectively an 8-fold and 3-fold enhancement in photocatalytic activity for S-doped g-C<sub>3</sub>N<sub>4</sub> and ZnIn<sub>2</sub>S<sub>4</sub>. Additionally, Jiang *et al.*<sup>95</sup> reported a tunable heterojunction photocatalyst of cobalt oxide (CoO<sub>x</sub>) confined in g-C<sub>3</sub>N<sub>4</sub> nanotubes, where the type II heterojunction of CoO/g-C<sub>3</sub>N<sub>4</sub> and type I heterojunction of Co<sub>3</sub>O<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub> were obtained. The results showed that the prepared two heterojunctions of CoO<sub>x</sub>/g-C<sub>3</sub>N<sub>4</sub> both exhibited better photocatalytic performance than the CoO<sub>x</sub> and g-C<sub>3</sub>N<sub>4</sub>. And the CoO/g-C<sub>3</sub>N<sub>4</sub> with type II band alignment exhibited higher photocatalytic H<sub>2</sub> evolution activity than the Co<sub>3</sub>O<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub> counterpart, which was mainly due to the longer lifetime of charge carriers, faster charge transfer, and higher charge separation efficiency in the type II heterojunction of CoO/g-C<sub>3</sub>N<sub>4</sub>. Recently, ternary heterojunctions with band coupling structures of type II/type II were designed to further accelerate the charge separation and transfer, and thus enhance

Table 2 Photocatalytic H<sub>2</sub> evolution activity of recent g-C<sub>3</sub>N<sub>4</sub>-based heterojunctions

| Photocatalyst                                                                     | Type         | Light source               | Reaction conditions                                                     | H <sub>2</sub> evolution rate               | AQE              | Ref. |
|-----------------------------------------------------------------------------------|--------------|----------------------------|-------------------------------------------------------------------------|---------------------------------------------|------------------|------|
| CuFeO <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub>                               | Type II      | 200 W Xe lamp              | 20 mg catalyst, TEOA solution                                           | 700.3 μmol h <sup>-1</sup> g <sup>-1</sup>  | 25% at 450 nm    | 88   |
| ZnIn <sub>2</sub> S <sub>4</sub> /g-C <sub>3</sub> N <sub>4</sub>                 | Type II      | 300 W Xe lamp (λ > 420 nm) | 50 mg catalyst, TEOA solution, Pt cocatalyst                            | 1.63 mmol h <sup>-1</sup> g <sup>-1</sup>   | 0.90% at 420 nm  | 89   |
| CoO/g-C <sub>3</sub> N <sub>4</sub>                                               | Type II      | 300 W Xe lamp (λ > 420 nm) | 40 mg catalyst, TEA <sup>a</sup> solution                               | 263 μmol h <sup>-1</sup> g <sup>-1</sup>    | 1.9% at 420 nm   | 95   |
| CdSe/WS <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub>                             | Type II      | White LED                  | 5 mg catalyst, TEOA solution                                            | 1.29 mmol h <sup>-1</sup> g <sup>-1</sup>   | —                | 101  |
| BaTiO <sub>3</sub> /Au/g-C <sub>3</sub> N <sub>4</sub>                            | Z-scheme     | 300 W Xe lamp (AM 1.5G)    | 20 mg catalyst, TEOA solution                                           | 1769.3 μmol h <sup>-1</sup> g <sup>-1</sup> | —                | 107  |
| g-C <sub>3</sub> N <sub>4</sub> /W <sub>18</sub> O <sub>49</sub>                  | Z-scheme     | 300 W Xe lamp (λ > 400 nm) | 50 mg catalyst, TEOA solution, Pt cocatalyst                            | 5231 μmol h <sup>-1</sup> g <sup>-1</sup>   | 23.1% at 420 nm  | 116  |
| CoSeO <sub>3</sub> /g-C <sub>3</sub> N <sub>4</sub>                               | Z-scheme     | 300 W Xe lamp              | 20 mg catalyst, TEOA solution                                           | 1459.2 μmol h <sup>-1</sup> g <sup>-1</sup> | 39.9% at 584 nm  | 117  |
| ZnIn <sub>2</sub> S <sub>4</sub> /g-C <sub>3</sub> N <sub>4</sub>                 | Z-scheme     | 300 W Xe lamp (λ > 400 nm) | 2 mg catalyst, TEOA solution                                            | 14.8 mmol g <sup>-1</sup> h <sup>-1</sup>   | 0.516% at 400 nm | 119  |
| g-C <sub>3</sub> N <sub>4</sub> /Co-MOF                                           | Z-scheme     | 225 W Xe lamp              | 5 mg catalyst, ascorbic acid solution, Pt cocatalyst                    | 33.17 mmol h <sup>-1</sup> g <sup>-1</sup>  | —                | 120  |
| Ti <sub>3</sub> C <sub>2</sub> /TiO <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub> | S-scheme     | 300 W Xe lamp (λ > 420 nm) | 50 mg catalyst, TEOA solution Ti <sub>3</sub> C <sub>2</sub> cocatalyst | 5540 μmol h <sup>-1</sup> g <sup>-1</sup>   | 5.81% at 420 nm  | 125  |
| CuInS <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub>                               | S-scheme     | 350 W Xe lamp (λ > 365 nm) | 20 mg catalyst, TEOA solution                                           | 102.4 μmol h <sup>-1</sup> g <sup>-1</sup>  | —                | 126  |
| Mn <sub>0.2</sub> Cd <sub>0.8</sub> S/g-C <sub>3</sub> N <sub>4</sub>             | S-scheme     | 300 W Xe lamp (λ > 420 nm) | TEOA solution, Pt cocatalyst                                            | 11.42 mmol h <sup>-1</sup> g <sup>-1</sup>  | —                | 127  |
| CNN/BDCNN <sup>b</sup>                                                            | Homojunction | 300 W Xe lamp (λ > 300 nm) | 40 mg catalyst, pure water, Pt and Co(OH) <sub>2</sub> cocatalyst       | 823.5 μmol h <sup>-1</sup> g <sup>-1</sup>  | 5.95% at 400 nm  | 130  |
| S-C <sub>3</sub> N <sub>4</sub> /C <sub>3</sub> N <sub>4</sub>                    | Homojunction | 300 W Xe lamp (λ > 420 nm) | 20 mg catalyst, TEOA solution, Pt cocatalyst                            | 5548.1 μmol h <sup>-1</sup> g <sup>-1</sup> | 0.43% at 420 nm  | 131  |
| CN-NaK <sup>c</sup> (triazine-heptazine)                                          | Homojunction | 50 W LED (λ > 420 nm)      | 50 mg catalyst, TEOA solution, Pt cocatalyst                            | 11.7 mmol h <sup>-1</sup> g <sup>-1</sup>   | 60% at 420 nm    | 132  |
| PHI/PTI <sup>d</sup>                                                              | Homojunction | 300 W Xe lamp (λ > 420 nm) | 50 mg catalyst, TEOA solution, Pt cocatalyst                            | 6970 μmol h <sup>-1</sup> g <sup>-1</sup>   | 42% at 420 nm    | 133  |
| g-C <sub>3</sub> N <sub>4</sub> /BST <sup>e</sup>                                 | Homojunction | 225 W Xe lamp              | 5 mg catalyst, TEOA solution, Pt cocatalyst                             | 12.47 mmol h <sup>-1</sup> g <sup>-1</sup>  | —                | 134  |

<sup>a</sup> Triethylamine. <sup>b</sup> Boron-doped and nitrogen-deficient g-C<sub>3</sub>N<sub>4</sub> nanosheets. <sup>c</sup> Triazine-heptazine-based carbon nitride. <sup>d</sup> Poly(heptazine-triazine) imides. <sup>e</sup> BST: C-N compound broken by s-triazine units.

the photocatalytic efficiency. For instance, a ternary hybrid photocatalyst of CdSe/WS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> was prepared by Liang and Chen *et al.*<sup>101</sup> According to the band structures of pristine CdSe, WS<sub>2</sub>, and g-C<sub>3</sub>N<sub>4</sub>, the type II/type II heterostructure was concluded to form at the interfaces of CdSe/WS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub>, which could not only expand the light absorption range but also greatly promote charge separation and restrain recombination. As a result, the designed ternary CdSe/WS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> photocatalyst displayed significantly enhanced photocatalytic performance, achieving the maximum H<sub>2</sub> evolution rate of 1.29 mmol h<sup>-1</sup> g<sup>-1</sup>, which was 1.7 and 1.3 times greater than binary WS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> and CdSe/g-C<sub>3</sub>N<sub>4</sub>, respectively.

As discussed above, the type II band alignment can indeed accelerate the migration and transfer of photogenerated charge carriers, and thus enhance the photocatalytic efficiency. However, as indicated in Fig. 7c, the spatial separation of charge carriers in type II heterojunctions is at the expense of their redox abilities when the electrons and holes are accumulated in more positive and negative energy levels, respectively, resulting in a weakened driving force for the following redox reaction.

### 3.2. Z-scheme heterojunction

Compared to type II heterojunctions, the Z-scheme heterojunction not only favors the spatial separation of charge carriers and inhibits charge recombination, but also retains a high catalytic activity with the photogenerated electrons and holes at higher chemical potentials of CB and VB, respectively.<sup>102</sup> Typically, in a Z-scheme photocatalyst (Fig. 7d), the close contact of two semiconductors induces the recombination of ineffective charges at the interface, resulting in an enlarged distance for the active charge carriers, which effectively weakens the Coulombic attraction of electron-hole pairs and in turns achieves a high charge separation efficiency. Therefore, among different heterojunction photocatalysts, Z-scheme heterojunctions have received more attention and research due to their combined advantages of enhanced charge separation efficiency, weakened Coulomb field strength, and favorable redox ability.

The first generation of Z-scheme heterojunctions based on mimicking the natural photosynthesis process was proposed by Bard in 1979,<sup>103</sup> where the electron acceptor/donor ion pairs (A/D) were used as the redox mediators (Fig. 8a), *e.g.*, I<sup>-</sup>/IO<sub>3</sub><sup>-</sup>, Fe<sup>2+</sup>/Fe<sup>3+</sup>, and Co(bpy)<sub>3</sub><sup>2+</sup>/Co(bpy)<sub>3</sub><sup>3+</sup>.<sup>104,105</sup> Although the introduction of the redox couples is beneficial for charge separation, it also brings several issues that the reversible redox couples would compete with the reactants for reaction and their application is limited to the liquid phase systems. In order to address the drawbacks of the conventional Z-scheme heterojunctions, an all-solid-state Z-scheme heterojunction was developed by Tada *et al.*,<sup>106</sup> where the redox ion pairs were substituted by solid electron mediators (Fig. 8b), such as noble metals, carbon nanotubes/nanofibers, graphene, and other conductive materials. Li *et al.*<sup>107</sup> constructed an all-solid-state Z-scheme photocatalyst by combining BaTiO<sub>3</sub> and g-C<sub>3</sub>N<sub>4</sub> as the main catalytic components, using Au nanoparticles as the electron mediator. The prepared Z-scheme BaTiO<sub>3</sub>/Au/g-C<sub>3</sub>N<sub>4</sub> was able to show enhanced charge separation efficiency while retaining the sufficient redox capability for reaction. Compared

to g-C<sub>3</sub>N<sub>4</sub>, the BaTiO<sub>3</sub>/Au/g-C<sub>3</sub>N<sub>4</sub> exhibited higher photocatalytic activity toward H<sub>2</sub> production and RhB degradation, particularly, an 18-fold enhancement in H<sub>2</sub> evolution rate was achieved. Wang and Li *et al.*<sup>108</sup> also designed an all-solid-state ternary Z-scheme photocatalyst by coupling g-C<sub>3</sub>N<sub>4</sub>, carbon dots (CDs) and TiO<sub>2</sub>, where the CDs were adopted as an electron mediator. They found that the g-C<sub>3</sub>N<sub>4</sub>/CDs/TiO<sub>2</sub> exhibited a high H<sub>2</sub> evolution activity of 580 μmol h<sup>-1</sup> g<sup>-1</sup>, and surpassed the binary g-C<sub>3</sub>N<sub>4</sub>/TiO<sub>2</sub> by 1.6 times, which was attributed to the effective spatial separation of charge carriers and their fast transfer assisted by the CDs mediator.

Obviously, as the second-generation Z-scheme heterojunctions, all-solid-state photocatalysts can effectively solve the problems of the ion pair-mediated Z-scheme photocatalytic system. However, the loading of solid electron mediators on the surface of semiconductors can trigger a shielding effect, reducing the light absorption and utilization, and the widely used noble metal mediators would also increase the costs and limit the wide application. Subsequently, a mediator-free Z-scheme heterojunction of ZnO/CdS was designed to further optimize the charge transfer process in photocatalytic reaction,<sup>109</sup> and then Yu's group proposed the third-generation Z-scheme heterojunction without any mediators (Fig. 8c), namely the direct Z-scheme heterojunction or S-scheme heterojunction.<sup>110,111</sup> In a direct Z-scheme photocatalytic system, the shortcomings of the above two Z-scheme heterojunctions can be overcome, such as the light-shielding effect induced by solid mediators, high cost caused by the loading of mediators, and the backward reactions due to the presence of redox ion pairs.

We can see that the band structure of the direct Z-scheme heterojunction is similar to that of type II heterojunctions, whereas the migration and transfer paths of charge carriers are different (Fig. 7c and 8c). In direct Z-scheme heterojunction photocatalysts, the charge reorganization would occur at the interface of two closely contacted semiconductors due to their differences in work function or Fermi level, that is, electrons can autonomously move from one semiconductor with a lower work function to another with a higher work function, forming a built-in electric field (BIEF) at the interface.<sup>112</sup> Subsequently, under light irradiation, the BIEF drives the separation and transfer of photogenerated electrons and holes, resulting in the recombination of useless charge carriers and leaving the

active charge carriers at more favorable redox potentials.<sup>113,114</sup> Therefore, the direct Z-scheme heterojunction with a spontaneously formed BIEF at the interface exhibits a high charge separation efficiency and redox capability, while avoiding the defects of the first and second generation Z-scheme heterojunction photocatalysts (*i.e.*, electron mediator-coupled systems).

Due to these advantages, the development of g-C<sub>3</sub>N<sub>4</sub>-based direct Z-scheme heterojunctions has received more attention than those traditional Z-scheme systems. Lately, numerous direct Z-scheme heterojunctions of g-C<sub>3</sub>N<sub>4</sub>-coupled composites with robust photocatalytic activities have been reported, such as g-C<sub>3</sub>N<sub>4</sub>/Bi<sub>4</sub>NbO<sub>8</sub>Cl,<sup>115</sup> g-C<sub>3</sub>N<sub>4</sub>/W<sub>18</sub>O<sub>49</sub>,<sup>116</sup> CoSeO<sub>3</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>117</sup> Bi<sub>4</sub>O<sub>5</sub>Br<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>118</sup> ZnIn<sub>2</sub>S<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>119</sup> g-C<sub>3</sub>N<sub>4</sub>/Co-MOF,<sup>120</sup> CoS/g-C<sub>3</sub>N<sub>4</sub>,<sup>121</sup> LaCoO<sub>3</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>122</sup> hBN/g-C<sub>3</sub>N<sub>4</sub>,<sup>123</sup> g-C<sub>3</sub>N<sub>4</sub>/Bi<sub>2</sub>MoO<sub>6</sub>,<sup>124</sup> Ti<sub>3</sub>C<sub>2</sub>/TiO<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>125</sup> CuInS<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub>,<sup>126</sup> and Mn<sub>0.2</sub>Cd<sub>0.8</sub>S/g-C<sub>3</sub>N<sub>4</sub>.<sup>127</sup> Typically, Zou *et al.*<sup>116</sup> constructed a C<sub>3</sub>N<sub>4</sub>/W<sub>18</sub>O<sub>49</sub> composite and found that the Fermi level of g-C<sub>3</sub>N<sub>4</sub> could be uplifted through the adsorption of triethanolamine (TEOA), thereby transforming the composite from type II to direct Z-scheme (Fig. 9a and b). The C<sub>3</sub>N<sub>4</sub>/W<sub>18</sub>O<sub>49</sub> composite with efficient Z-scheme charge transfer exhibited much higher photocatalytic activity compared with pure g-C<sub>3</sub>N<sub>4</sub> and type-II C<sub>3</sub>N<sub>4</sub>/W<sub>18</sub>O<sub>49</sub>, resulting in a H<sub>2</sub> evolution rate of 5231 μmol h<sup>-1</sup> g<sup>-1</sup> and AQE of 23.1% at 420 nm (Fig. 9c). Dong and Cao *et al.*<sup>117</sup> reported a novel CoSeO<sub>3</sub> assisted g-C<sub>3</sub>N<sub>4</sub> composite with a Z-scheme structure. After combining with CoSeO<sub>3</sub>, the CoSeO<sub>3</sub>/g-C<sub>3</sub>N<sub>4</sub> displayed an expanded light absorption range, reduced charge transfer resistance and accelerated charge separation, and the optimal photocatalytic H<sub>2</sub> evolution activity of CoSeO<sub>3</sub>/g-C<sub>3</sub>N<sub>4</sub> was about 65 times that of pristine g-C<sub>3</sub>N<sub>4</sub>. Similarly, Dai and Zhang *et al.*<sup>127</sup> designed an S-scheme heterojunction photocatalyst of Mn<sub>0.2</sub>Cd<sub>0.8</sub>S/g-C<sub>3</sub>N<sub>4</sub> through an *in-situ* hydrothermal growth process. The characterization results showed that the Mn<sub>0.2</sub>Cd<sub>0.8</sub>S/g-C<sub>3</sub>N<sub>4</sub> hybrid possessed an intimate interface and favorable S-scheme band alignment, which could promote the charge transfer at the interface and increase the lifetime of reactive charge carriers, as well as retain the high redox capacity. The photocatalytic H<sub>2</sub> production activity of Mn<sub>0.2</sub>Cd<sub>0.8</sub>S/g-C<sub>3</sub>N<sub>4</sub> reached 11.42 mmol h<sup>-1</sup> g<sup>-1</sup>, which was much higher than that of the original g-C<sub>3</sub>N<sub>4</sub> and Mn<sub>0.2</sub>Cd<sub>0.8</sub>S.

Besides these, recently, the isotype heterojunctions of g-C<sub>3</sub>N<sub>4</sub>-based composites have also been discussed and investigated due to their better lattice matching and compatibility,

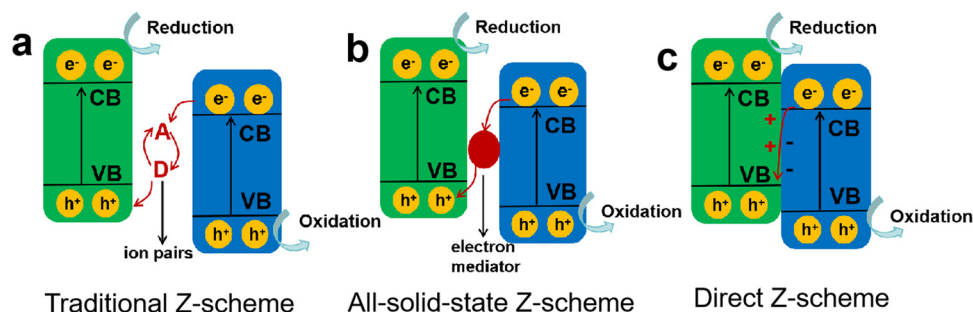
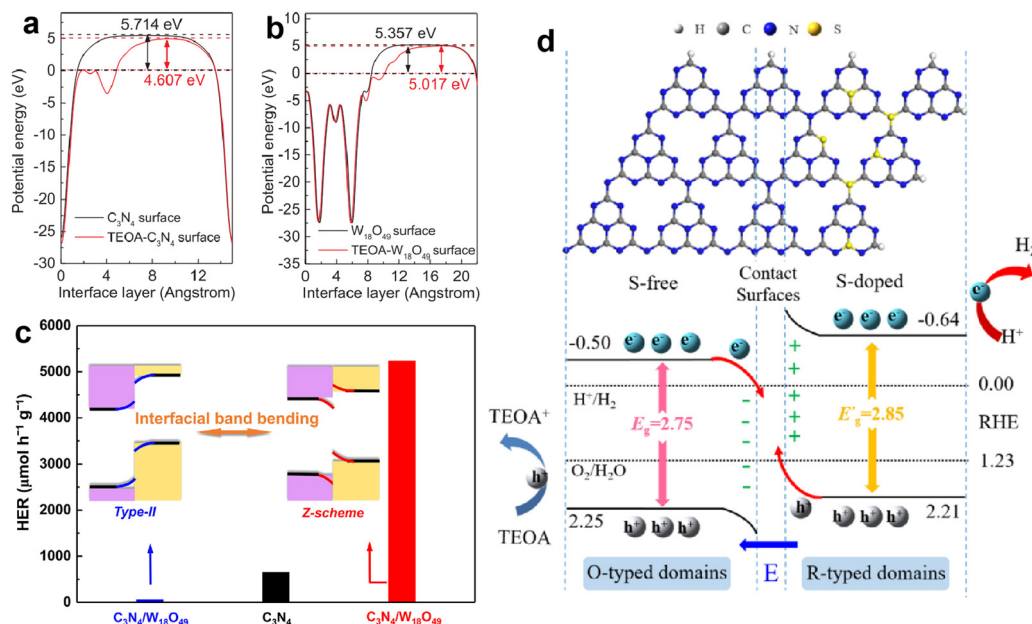


Fig. 8 Different types of Z-scheme heterojunctions: (a) traditional Z-scheme, (b) all-solid-state Z-scheme, (c) direct Z-scheme.



**Fig. 9** The calculated work functions of g-C<sub>3</sub>N<sub>4</sub> (a) and W<sub>18</sub>O<sub>49</sub> (b) surfaces with and without TEOA; photocatalytic H<sub>2</sub> evolution activities of pure g-C<sub>3</sub>N<sub>4</sub>, type II g-C<sub>3</sub>N<sub>4</sub>/W<sub>18</sub>O<sub>49</sub>, and Z-scheme g-C<sub>3</sub>N<sub>4</sub>/W<sub>18</sub>O<sub>49</sub> (c). Reproduced with permission from ref. 116. Copyright 2017, Elsevier. The structural configuration and charge transfer mechanism of S-doped g-C<sub>3</sub>N<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub> isotype heterojunction (d). Reproduced with permission from ref. 131. Copyright 2022, Elsevier.

which are beneficial to the separation and transfer of photo-induced charge carriers at the interface.<sup>128–131</sup> Shen *et al.*<sup>130</sup> prepared a series of g-C<sub>3</sub>N<sub>4</sub> nanosheets with different contents of B dopants and N defects, resulting in the formation of g-C<sub>3</sub>N<sub>4</sub> with controllable band structures, which could be applied as an O<sub>2</sub>-evolving component and combined with the original g-C<sub>3</sub>N<sub>4</sub> as a H<sub>2</sub>-evolving component to construct an isotype Z-scheme heterojunction for overall water splitting. The obtained g-C<sub>3</sub>N<sub>4</sub> self-based heterojunction displayed excellent photocatalytic activity, and the optimal H<sub>2</sub> and O<sub>2</sub> production rates achieved stoichiometric values of 32.94 and 16.42 μmol h<sup>−1</sup> (AQE of 5.95% at 400 nm). The superior photocatalytic performance was attributed to the strong interfacial interaction, favorable redox ability, and efficient charge transfer through the interfacial Z-scheme route. Jiang *et al.*<sup>131</sup> also reported a novel isotype heterostructure of S-doped g-C<sub>3</sub>N<sub>4</sub>/g-C<sub>3</sub>N<sub>4</sub> with S-scheme band alignment (Fig. 9d), which exhibited high photocatalytic H<sub>2</sub> evolution activity of 5548.1 μmol h<sup>−1</sup> g<sup>−1</sup> (AQE of 0.43% at 420 nm) and remarkable durability for 24 h, nearly 49 times that of pure g-C<sub>3</sub>N<sub>4</sub>. Zhang and Wang *et al.*<sup>132</sup> prepared a triazine-heptazine based copolymer with the internal donor-acceptor heterostructure, which greatly accelerated the interface charge transfer and achieved high photocatalytic H<sub>2</sub> evolution activity with an AQE of 60% at 420 nm. Afterwards, they also fabricated binary poly(heptazine-triazine) imides with semi-coherent interfaces. The fine lattice compatibility and intimate contact interface induced the formation of the built-in electric field, which facilitated rapid charge separation and transfer, resulting in a high AQE of 42% at 420 nm, even 64% by using K<sub>2</sub>HPO<sub>4</sub> as the charge mediator.<sup>133</sup> Additionally, Shi *et al.*<sup>134</sup> also developed a g-C<sub>3</sub>N<sub>4</sub> homojunction by decorating

the broken *s*-triazine unit C–N compound, which was tightly anchored on the g-C<sub>3</sub>N<sub>4</sub> and could extract photogenerated electrons from g-C<sub>3</sub>N<sub>4</sub>. Furthermore, the broken *s*-triazine units provided active sites to grab H<sup>+</sup>, thereby promoting the H<sup>+</sup> reduction and achieving a fourfold improvement in H<sub>2</sub> evolution activity.

## 4. Cocatalyst loading

In addition to the heterojunction of coupling two semiconductors, loading metal-based cocatalysts on the surface of photocatalysts can also greatly promote the separation and transfer of photo-induced charge carriers, and then improve photocatalytic H<sub>2</sub> production activity, such as the reported Pt loaded g-C<sub>3</sub>N<sub>4</sub>, CdS and TiO<sub>2</sub>, and they all exhibited better photocatalytic performance than their original counterparts.<sup>135–137</sup> The role of cocatalyst in improving photocatalytic activity is as follows:<sup>138</sup> on the one hand, the cocatalyst can serve as an electron trap to capture photogenerated electrons from the CB level of the semiconductor photocatalyst and retain the holes on the semiconductor, which achieve spatial separation of charge carriers and effectively suppress the recombination; on the other, the cocatalyst can provide abundant catalytic sites and reduce the activation energy of the H<sub>2</sub> evolution reaction. Fig. 10a shows a schematic diagram of the role of cocatalysts in photocatalytic water splitting. In order to achieve state-of-the-art photocatalytic activity, the cocatalysts should display suitable band alignment with the semiconductor, strong conductivity, and high catalytic activity toward H<sub>2</sub> evolution.<sup>139,140</sup> According to the requirements of cocatalysts, H<sub>2</sub>-evolution electrocatalysts with excellent catalytic activity are commonly considered as promising



Fig. 10 The role of cocatalyst in photocatalytic water splitting (a). Reproduced with permission from ref. 139. Copyright 2020, Wiley-VCH. The band alignment between metallic cocatalysts and n-type g-C<sub>3</sub>N<sub>4</sub> after intimate contact (b). Reproduced with permission from ref. 140. Copyright 2021, Wiley-VCH.

cocatalysts for photocatalytic H<sub>2</sub> production.<sup>141–143</sup> Most electrocatalysts have lower Fermi levels (larger work functions), which is conducive to capturing photogenerated electrons, and their high catalytic activity can also boost the kinetics of surface H<sub>2</sub> evolution reaction. We previously prepared the VS<sub>2</sub> electrocatalyst with excellent catalytic activity toward H<sub>2</sub> evolution, and loaded it on g-C<sub>3</sub>N<sub>4</sub> as a cocatalyst.<sup>27,144</sup> Compared to pure g-C<sub>3</sub>N<sub>4</sub>, the VS<sub>2</sub> supported g-C<sub>3</sub>N<sub>4</sub> displayed much higher photocatalytic H<sub>2</sub> evolution activity, a nearly 30 times increase. Our characterization and calculation results clearly revealed that the loading of VS<sub>2</sub> obviously improved the charge separation efficiency and reduced the energy barrier for H<sub>2</sub> evolution.

For the n-type g-C<sub>3</sub>N<sub>4</sub> semiconductor, its Fermi level is close to the CB and generally higher than that of the metal-based cocatalysts, that is, the smaller work function of g-C<sub>3</sub>N<sub>4</sub> ( $\phi_{\text{g-C}_3\text{N}_4} < \phi_{\text{cocatalyst}}$ ). When the g-C<sub>3</sub>N<sub>4</sub> surface is in intimate contact with the cocatalyst, the interfacial electrons transfer and lead to the formation of a Helmholtz double layer at the cocatalyst-g-C<sub>3</sub>N<sub>4</sub> junction, where the g-C<sub>3</sub>N<sub>4</sub> is positively charged and the cocatalyst is negatively charged. The established electric field between g-C<sub>3</sub>N<sub>4</sub> and cocatalyst causes the band edges of g-C<sub>3</sub>N<sub>4</sub> to bend upwards, forming an energy barrier called the Schottky barrier.<sup>145,146</sup> Although the presence of a Schottky barrier at the interface would affect the charge transfer, the photogenerated electrons with sufficient energy can still flow from g-C<sub>3</sub>N<sub>4</sub> to the cocatalyst through the Schottky barrier. Meanwhile, the rectification characteristic of the Schottky barrier suppresses the back recombination of charge carriers, ensuring the directional charge transfer and improving charge separation efficiency.<sup>147,148</sup> Therefore, a perfect match of g-C<sub>3</sub>N<sub>4</sub> and cocatalyst and the appropriate Schottky barrier are necessary to facilitate the interfacial charge transfer, prolong the lifetime of charge carriers, and improve photocatalytic H<sub>2</sub> production activity.

In this section, typical noble metal-based cocatalysts and transition metal-derived cocatalysts are discussed, with a focus on the role of cocatalysts in charge separation and transfer. A comparison of the photocatalytic H<sub>2</sub> evolution activity of recent g-C<sub>3</sub>N<sub>4</sub> loaded with different cocatalyst is given in Table 3.

#### 4.1. Noble metal-based cocatalysts

Noble metals such as Pt, Pd, Ru, Ir, Rh, and Au have often been utilized as the cocatalysts for photocatalytic H<sub>2</sub> production, due

to their strong metallicity, high work function for electron transfer, and robust catalytic activity for H<sub>2</sub> evolution.<sup>149–151</sup> Among them, Pt can form an ideal Pt-H bond with appropriate strength and then significantly reduce the activation energy of H<sub>2</sub> evolution, so that Pt shows the highest H<sub>2</sub> evolution activity.<sup>152</sup> Consequently, Pt is the most commonly used cocatalyst for improving photocatalytic H<sub>2</sub> production performance. Wang and Domen *et al.*<sup>153</sup> compared the photocatalytic H<sub>2</sub> production activities of g-C<sub>3</sub>N<sub>4</sub> loaded with various noble metal cocatalysts, including Ru, Rh, Pd, Ir, Pt, and Au, and found the Pt loaded g-C<sub>3</sub>N<sub>4</sub> exhibited the highest photocatalytic activity toward H<sub>2</sub> evolution. Moreover, nanoscale noble metals with a strong plasmonic effect, such as Au, Ag, and Pt, can improve the absorption of visible light and provide many high-energy hot electrons for reaction, thus further enhancing the photocatalytic performance.<sup>154,155</sup> Wang *et al.*<sup>154</sup> deposited Au nanoparticles on the surface of g-C<sub>3</sub>N<sub>4</sub> and obtained the obviously enhanced photocatalytic H<sub>2</sub> evolution activity. The characterization revealed that the loaded Au nanoparticles played a dual role in the photocatalytic reaction. On one hand, the Au nanoparticles could improve the charge separation efficiency through capturing electrons from g-C<sub>3</sub>N<sub>4</sub>; on the other hand, the Au plasma could provide additional hot electrons for H<sup>+</sup> reduction. Hence, the accelerated charge transfer and increased number of electrons contributed to the enhanced photocatalytic activity of Au/g-C<sub>3</sub>N<sub>4</sub>. Yang and Wang *et al.*<sup>156</sup> also investigated the effect of different sizes and shapes of Au nanocrystals on the photocatalytic performance of g-C<sub>3</sub>N<sub>4</sub>. Among the prepared Au/g-C<sub>3</sub>N<sub>4</sub>, the 18-nm-sized Au nanosphere loaded g-C<sub>3</sub>N<sub>4</sub> showed an excellent photocatalytic H<sub>2</sub> production activity of 540  $\mu\text{mol h}^{-1} \text{g}^{-1}$  (AQE = 1.2% at 380 nm), far exceeding the larger-sized Au nanosphere and Au nanorod loaded g-C<sub>3</sub>N<sub>4</sub> system. The superior H<sub>2</sub> evolution activity was attributed to the optimized Au nanocrystals in promoting charge transfer and extending the light-harvesting range. Although noble metals can effectively promote proton reduction, they are also active sites for the backward reaction of H<sub>2</sub> and O<sub>2</sub>. Thus, a noble metal/metal oxide cocatalyst with a core-shell structure was proposed to suppress backward reaction.<sup>157,158</sup> For example, Zhang and Wang *et al.*<sup>159</sup> decorated poly(triazine imide)-based carbon nitride with a Rh/Cr<sub>2</sub>O<sub>3</sub> core-shell cocatalyst for H<sub>2</sub> evolution, and CoO<sub>x</sub> as the O<sub>2</sub>

Table 3 Photocatalytic H<sub>2</sub> evolution activity of recent g-C<sub>3</sub>N<sub>4</sub> loaded with different cocatalyst

| Photocatalyst                                                         | Cocatalyst                                          | Light source                        | Reaction conditions                                                        | H <sub>2</sub> evolution rate              | AQE              | Ref. |
|-----------------------------------------------------------------------|-----------------------------------------------------|-------------------------------------|----------------------------------------------------------------------------|--------------------------------------------|------------------|------|
| Au/g-C <sub>3</sub> N <sub>4</sub>                                    | Au                                                  | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution                                              | 159.9 $\mu\text{mol h}^{-1} \text{g}^{-1}$ | —                | 154  |
| Au/g-C <sub>3</sub> N <sub>4</sub>                                    | Au nanospheres                                      | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution                                              | 540 $\mu\text{mol h}^{-1} \text{g}^{-1}$   | 1.2% at 380 nm   | 156  |
| Rh/Cr <sub>2</sub> O <sub>3</sub> -CoO <sub>x</sub> /PTI <sup>a</sup> | Rh/Cr <sub>2</sub> O <sub>3</sub> -CoO <sub>x</sub> | 300 W Xe lamp                       | 100 mg catalyst, pure water                                                | 4500 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 20.2% at 365 nm  | 159  |
| Pt-CoO <sub>x</sub> /PTI                                              | Pt-CoO <sub>x</sub>                                 | 300 W Xe lamp                       | 100 mg catalyst, pure water                                                | 2730 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 12% at 365 nm    | 160  |
| TiO <sub>2</sub> @Pt/C <sub>3</sub> N <sub>4</sub>                    | TiO <sub>2</sub> @Pt                                | 300 W Xe lamp                       | 100 mg catalyst, pure water                                                | 7.63 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 0.11% at 365 nm  | 161  |
| PrAu/g-C <sub>3</sub> N <sub>4</sub>                                  | PrAu                                                | 300 W Xe lamp                       | 50 mg catalyst, Na <sub>2</sub> S/Na <sub>2</sub> SO <sub>3</sub> solution | 1009 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 0.45% at 420 nm  | 164  |
| PtNi/g-C <sub>3</sub> N <sub>4</sub>                                  | PtNi                                                | 300 W Xe lamp ( $\lambda > 420$ nm) | 25 mg catalyst, TEOA solution                                              | 9528 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 10.6% at 370 nm  | 165  |
| Ni/TiO <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub>                  | Ni                                                  | 300 W Xe lamp                       | 50 mg catalyst, TEOA solution                                              | 134 $\mu\text{mol h}^{-1} \text{g}^{-1}$   | 15% at 420 nm    | 170  |
| NiMo/g-C <sub>3</sub> N <sub>4</sub>                                  | NiMo                                                | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution                                              | 1785 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | ~0.25% at 365 nm | 172  |
| NiCu/g-C <sub>3</sub> N <sub>4</sub>                                  | NiCu                                                | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution                                              | 2088 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 6.83% at 475 nm  | 173  |
| Mo <sub>2</sub> C@C/g-C <sub>3</sub> N <sub>4</sub>                   | Mo <sub>2</sub> C                                   | 5 W LED                             | 10 mg catalyst, TEOA solution                                              | 651 $\mu\text{mol h}^{-1} \text{g}^{-1}$   | 1.44% at 400 nm  | 178  |
| WC/g-C <sub>3</sub> N <sub>4</sub>                                    | WC                                                  | 300 W Xe lamp ( $\lambda > 420$ nm) | 50 mg catalyst, TEOA solution                                              | 146.1 $\mu\text{mol h}^{-1} \text{g}^{-1}$ | —                | 180  |
| Ni <sub>2</sub> P/rGO/g-C <sub>3</sub> N <sub>4</sub>                 | Ni <sub>2</sub> P                                   | 300 W Xe lamp ( $\lambda > 420$ nm) | 10 mg catalyst, TEOA solution                                              | 2922 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 5.6% at 420 nm   | 187  |
| Cu <sub>3</sub> P/g-C <sub>3</sub> N <sub>4</sub>                     | Cu <sub>3</sub> P                                   | 300 W Xe lamp ( $\lambda > 420$ nm) | 20 mg catalyst, TEOA solution                                              | 343 $\mu\text{mol h}^{-1} \text{g}^{-1}$   | —                | 188  |
| MoP/g-C <sub>3</sub> N <sub>4</sub>                                   | MoP                                                 | 300 W Xe lamp                       | 20 mg catalyst, TEOA solution                                              | 3868 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 21.6% at 405 nm  | 189  |
| NiCoP/g-C <sub>3</sub> N <sub>4</sub>                                 | NiCoP                                               | 300 W Xe lamp                       | 20 mg catalyst, TEOA solution                                              | 5162 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 18.5% at 400 nm  | 190  |
| Cu-Ni(OH) <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub>               | Cu-Ni(OH) <sub>2</sub>                              | 300 W Xe lamp (AM 1.5 G)            | 20 mg catalyst, TEOA solution                                              | 2037 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | —                | 194  |
| Pt/Ni(OH) <sub>2</sub> /g-C <sub>3</sub> N <sub>4</sub>               | Pt/Ni(OH) <sub>2</sub>                              | 300 W Xe lamp                       | 10 mg catalyst, TEOA solution                                              | 3005 $\mu\text{mol h}^{-1} \text{g}^{-1}$  | 11.2% at 420 nm  | 196  |
| Pt/NVC/g-C <sub>3</sub> N <sub>4</sub>                                | Pt/NVC                                              | 300 W Xe lamp ( $\lambda > 400$ nm) | 50 mg catalyst, TEOA solution                                              | 800 $\mu\text{mol h}^{-1} \text{g}^{-1}$   | 0.20% at 420 nm  | 197  |

<sup>a</sup> PTI: poly(triazine imide) based carbon nitride.

evolution cocatalyst. The presence of a Cr<sub>2</sub>O<sub>3</sub> layer prevented the O<sub>2</sub> transfer but facilitated the proton penetration, thereby hindering the backward reaction over Rh and achieving an AQE of 20.2% at 365 nm for overall water splitting, which was much higher than the Pt and CoO<sub>x</sub> modified one (AQE = 12%).<sup>160</sup> Su *et al.*<sup>161</sup> modified Pt/C<sub>3</sub>N<sub>4</sub> with a TiO<sub>2</sub> thin layer, which effectively prevented the reverse reaction and photocorrosion. Compared with naked Pt/C<sub>3</sub>N<sub>4</sub>, TiO<sub>2</sub>@Pt/C<sub>3</sub>N<sub>4</sub> exhibited 130% enhanced overall efficiency of water splitting (AQE = 0.11% at 365 nm). In addition to pure noble metals, noble metal-based bimetallic alloys have also been extensively studied as cocatalysts to improve photocatalytic performance.<sup>162,163</sup> The advantages of bimetallic alloys are related to the synergistic effect of two metals in tuning the electronic structure, work function and catalytic activity, which lead to efficient charge transfer and photoreduction reaction. For example, Pradhan *et al.*<sup>164</sup> synthesized pure metals (Pt, Au) and PtAu alloy loaded g-C<sub>3</sub>N<sub>4</sub>, respectively, and investigated their photocatalytic H<sub>2</sub> production activities. The PtAu alloy modified g-C<sub>3</sub>N<sub>4</sub> exhibited better photocatalytic activity than the monometallic counterparts. A similar phenomenon of the improved photocatalytic performance was also observed in PtNi alloy loaded g-C<sub>3</sub>N<sub>4</sub>.<sup>165</sup>

Although the loading of noble metal-based cocatalysts can significantly boost the photocatalytic activity of H<sub>2</sub> production, the high cost and scarcity of noble metal materials limit their large-scale applications. Therefore, the development of low-cost, abundant, and highly active non-noble metal-based cocatalysts is highly desired for photocatalytic H<sub>2</sub> production. During the next part, various transition metal-based cocatalysts are discussed.

#### 4.2. Transition metal-based cocatalysts

Currently, there are many studies on transition metals as cocatalysts for photocatalytic H<sub>2</sub> production, such as Ni, Cu, and Co, *etc.*<sup>166–168</sup> Among various transition metals, Ni has been widely studied as a cocatalyst for g-C<sub>3</sub>N<sub>4</sub> because its work function ( $\phi = 5.4$  eV) is close to that of Pt ( $\phi = 5.7$  eV) and its Fermi level is more positive than the CB of g-C<sub>3</sub>N<sub>4</sub>, which facilitate the capture of photogenerated electrons from g-C<sub>3</sub>N<sub>4</sub>.<sup>169</sup> In the meantime, the Fermi level of Ni is more negative than the reduction potential of H<sup>+</sup>/H<sub>2</sub>, which promotes the transferred electrons to participate in the reduction reaction. Wu *et al.*<sup>170</sup> deposited the single-atom Ni on TiO<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> as a cocatalyst, which could capture photogenerated electrons from g-C<sub>3</sub>N<sub>4</sub> and accelerate the H<sup>+</sup> reduction, resulting in higher H<sub>2</sub> production activity (AQE of 15% at 420 nm) than the TiO<sub>2</sub>/g-C<sub>3</sub>N<sub>4</sub> without Ni cocatalyst. Do *et al.*<sup>171</sup> also reported the obviously improved photocatalytic H<sub>2</sub> evolution activity of g-C<sub>3</sub>N<sub>4</sub> and S-doped g-C<sub>3</sub>N<sub>4</sub> when Ni was applied as the cocatalyst. Besides individual transition metals, transition metal-based alloys have also received widespread attention as cocatalysts for photocatalytic reaction. For example, Zhang and Wang *et al.*<sup>172</sup> showed the drastically enhanced photocatalytic performance of g-C<sub>3</sub>N<sub>4</sub> with NiMo alloy as the cocatalyst. The optimal H<sub>2</sub> production rate of NiMo alloy modified g-C<sub>3</sub>N<sub>4</sub> was found to be 1785  $\mu\text{mol h}^{-1} \text{g}^{-1}$  (AQE of ~0.25% at 365 nm),

which was even comparable to the Pt/g-C<sub>3</sub>N<sub>4</sub> (2250  $\mu\text{mol h}^{-1} \text{g}^{-1}$ ). A similar enhancement in photocatalytic H<sub>2</sub> production performance was also observed in Ni-Cu bimetallic cocatalyst loaded g-C<sub>3</sub>N<sub>4</sub>, which exhibited excellent photocatalytic activity, about 24 and 2 times that of g-C<sub>3</sub>N<sub>4</sub> and Ni/g-C<sub>3</sub>N<sub>4</sub>, respectively.<sup>173</sup> As mentioned above, the photocatalytic activity of g-C<sub>3</sub>N<sub>4</sub> can be obviously improved by loading transition metals and their alloys. However, compared with the benchmark cocatalyst Pt, although most transition metals can effectively capture photogenerated electrons and improve charge separation efficiency, their catalytic activities toward H<sub>2</sub> evolution are lower than Pt, resulting in the lower photocatalytic performance of transition metal cocatalyst modified g-C<sub>3</sub>N<sub>4</sub> than that of Pt loaded g-C<sub>3</sub>N<sub>4</sub>.

As we know, the intrinsic activity of catalysts is closely related to their electronic structures; exploring transition metal-derived cocatalysts with Pt-like electronic properties is necessary for further boosting photocatalytic H<sub>2</sub> evolution. Levy *et al.*<sup>174</sup> reported that carbonized tungsten (WC) exhibited Pt-like catalytic activity in hydrogenolysis and isomerization reactions. Subsequently, Bennett and coworkers carried out theoretical calculations to investigate the WC and found that the introduction of carbon into the tungsten changed its electronic distribution, resulting in the density of electronic states near the Fermi level close to Pt, which contributed to the Pt-like catalytic activity of WC.<sup>175</sup> Therefore, the transition metal-derived compounds as the substitute for Pt cocatalyst have received increasing attention in the field of photocatalytic H<sub>2</sub> production, including transition metal carbides, nitrides, phosphides, sulfides, *etc.*<sup>176–179</sup> Li *et al.*<sup>180</sup> loaded the WC onto g-C<sub>3</sub>N<sub>4</sub> and studied the effect of WC on photocatalytic H<sub>2</sub> production activity. It was demonstrated that the loaded WC as an efficient cocatalyst not only enhanced the surface H<sub>2</sub>-evolution kinetics, but also served as an electron transfer channel to promote interfacial charge transfer. We previously studied the role of tungsten carbide (W<sub>2</sub>C), tungsten sulfide (WS<sub>2</sub>), and tungsten nitride (W<sub>2</sub>N) in photocatalytic H<sub>2</sub> production and compared their catalytic activities.<sup>28</sup> We found that the W<sub>2</sub>C loaded g-C<sub>3</sub>N<sub>4</sub> exhibited the highest H<sub>2</sub> evolution activity. Based on the above study, we also designed a carbonized MoS<sub>2</sub> (MoS<sub>2</sub>/Mo<sub>2</sub>C) as the cocatalyst for g-C<sub>3</sub>N<sub>4</sub>, and demonstrated its superior photocatalytic activity to that of the pure MoS<sub>2</sub> loaded one, nearly 6 times higher in the rate of H<sub>2</sub> evolution.<sup>181</sup> Thus, transition metal carbides as the alternative for noble metal Pt have prominent advantages and potential application value.

Similar to metal carbides, many metal phosphides display high electrocatalytic activity toward H<sub>2</sub> evolution, such as Ni<sub>2</sub>P, CoP, MoP, Cu<sub>3</sub>P, *etc.*<sup>182</sup> And the use of metal phosphides as cocatalysts can effectively improve the surface H<sub>2</sub> evolution kinetics and charge separation efficiency, thus achieving efficient photocatalytic H<sub>2</sub> production activity.<sup>183–186</sup> For example, Huang and coworkers compared the effect of different Ni-based compounds as cocatalysts on rGO/g-C<sub>3</sub>N<sub>4</sub>, and revealed that the Ni<sub>2</sub>P loaded rGO/g-C<sub>3</sub>N<sub>4</sub> displayed better photocatalytic H<sub>2</sub> evolution activity than the NiS, Ni<sub>3</sub>C, and Ni<sub>3</sub>N loaded system.<sup>187</sup> Among these photocatalysts, the suitable band alignment between the Ni<sub>2</sub>P and rGO/g-C<sub>3</sub>N<sub>4</sub> obviously promoted

electron transfer from g-C<sub>3</sub>N<sub>4</sub> to Ni<sub>2</sub>P *via* the rGO medium, then the H<sub>2</sub> was produced on the Ni<sub>2</sub>P reaction site. This work clearly demonstrated the advantages of metal phosphide cocatalysts in suppressing charge recombination and accelerating H<sub>2</sub> evolution. Furthermore, Liu *et al.*<sup>188</sup> synthesized two types of copper phosphide loaded g-C<sub>3</sub>N<sub>4</sub>, Cu<sub>3</sub>P/g-C<sub>3</sub>N<sub>4</sub> and Cu<sub>97</sub>P<sub>3</sub>/g-C<sub>3</sub>N<sub>4</sub>. The loading of copper phosphide significantly boosted the separation and transfer of photogenerated charge carriers, leading to an improved photocatalytic activity of Cu<sub>3</sub>P and Cu<sub>97</sub>P<sub>3</sub> loaded g-C<sub>3</sub>N<sub>4</sub>. Shi *et al.*<sup>189</sup> deposited the MoP as a cocatalyst on g-C<sub>3</sub>N<sub>4</sub>, the close interface induced the bonding state of Mo-N, which promoted photogenerated electron transfer to MoP and greatly improved the charge separation efficiency. Meanwhile, the loading of MoP also reduced the energy barrier for H<sub>2</sub> production. The designed MoP/g-C<sub>3</sub>N<sub>4</sub> displayed a H<sub>2</sub> evolution activity of 3868  $\mu\text{mol h}^{-1} \text{g}^{-1}$  with an AQE of 21.6% at 405 nm. Moreover, Liu and Fan *et al.*<sup>190</sup> also designed a bimetallic phosphide NiCoP as the cocatalyst for g-C<sub>3</sub>N<sub>4</sub> to further improve the photocatalytic performance. It was demonstrated that the rate of photocatalytic H<sub>2</sub> evolution over NiCoP/g-C<sub>3</sub>N<sub>4</sub> was obviously higher than that over Ni<sub>2</sub>P/g-C<sub>3</sub>N<sub>4</sub> and Co<sub>2</sub>P/g-C<sub>3</sub>N<sub>4</sub>, even better than that over Pt anchored g-C<sub>3</sub>N<sub>4</sub>. The improved activity of g-C<sub>3</sub>N<sub>4</sub> was attributed to the formation of an appropriate Schottky barrier at the interface of NiCoP/g-C<sub>3</sub>N<sub>4</sub>, which could promote the directional migration of photo-excited electrons to the NiCoP cocatalyst and inhibit the back recombination. Meanwhile, the synergistic effect of bimetals in NiCoP could further reduce the activation energy of H<sub>2</sub> evolution, thus accelerating the H<sup>+</sup> reduction.

Recently, on the basis of single-component cocatalysts, dual cocatalysts and composite cocatalysts have gradually been developed.<sup>191–193</sup> In dual cocatalyst decorated g-C<sub>3</sub>N<sub>4</sub>, the spatially separated dual cocatalysts serve as the reduction site and oxidation site to trap the electrons and holes from g-C<sub>3</sub>N<sub>4</sub>, respectively, thus greatly minimizing charge recombination. For example, the use of Cu and Ni(OH)<sub>2</sub> as dual cocatalysts for g-C<sub>3</sub>N<sub>4</sub> showed better photocatalytic activity with the H<sub>2</sub> evolution amount exceeding the sum of single Cu and Ni(OH)<sub>2</sub> counterparts.<sup>194</sup> The improvement in photocatalytic activity was mainly attributed to the coupling of Cu and Ni(OH)<sub>2</sub> cocatalysts, which could drive the transfer of photogenerated electrons and holes to Cu and Ni(OH)<sub>2</sub> respectively, achieving the spatial separation of charge carriers. A combination of Pt and MnO<sub>x</sub> was also reported as the dual cocatalysts to promote the separation, transfer, and utilization of photo-induced electron-hole pairs in CdS/g-C<sub>3</sub>N<sub>4</sub>.<sup>195</sup> It is known that loading cocatalysts with good metallicity and matched Fermi level can indeed favor the directional transfer of charge carriers from g-C<sub>3</sub>N<sub>4</sub> to cocatalysts. In the subsequent catalytic reaction, most cocatalysts can display high catalytic activity toward the reduction of H<sup>+</sup> to H<sub>2</sub>, but confront an obstacle regarding the large energy barrier for water dissociation, which is the first step in H<sub>2</sub> evolution reaction under neutral or alkaline conditions. Thus, the composite cocatalysts capable of promoting charge separation and improving reaction kinetics of proton reduction and water dissociation are developed. Zou and

coworkers designed Pt cluster decorated  $\text{Ni}(\text{OH})_2$  as a composite cocatalyst for  $\text{g-C}_3\text{N}_4$ , where the electron sink (Pt) and hole sink ( $\text{Ni}(\text{OH})_2$ ) were included in one cocatalyst.<sup>196</sup> It was found that the  $\text{g-C}_3\text{N}_4$  with spatially separated Pt- $\text{Ni}(\text{OH})_2$  cocatalyst exhibited a lower charge recombination rate than the composite counterpart, which was attributed to the effective spatial separation of electron-hole pairs. While in terms of the redox reaction, the close contact of Pt and  $\text{Ni}(\text{OH})_2$  in the composite cocatalyst could synergistically promote  $\text{H}_2\text{O}$  dissociation and  $\text{H}^+$  reduction, which could not be achieved in the spatially separated Pt- $\text{Ni}(\text{OH})_2$  system. As a result, the  $\text{g-C}_3\text{N}_4$  with Pt/ $\text{Ni}(\text{OH})_2$  composite cocatalyst displayed higher photocatalytic activity than the Pt,  $\text{Ni}(\text{OH})_2$ , and spatially separated Pt- $\text{Ni}(\text{OH})_2$  loaded  $\text{g-C}_3\text{N}_4$ . We also investigated a composite cocatalyst of Pt/N doped VC (Pt/NVC) and found an improvement in the photocatalytic  $\text{H}_2$  production activity of Pt/NVC/ $\text{g-C}_3\text{N}_4$  compared to NVC/ $\text{g-C}_3\text{N}_4$  and Pt/ $\text{g-C}_3\text{N}_4$ .<sup>197</sup> Therefore, the highly efficient cocatalysts are expected to achieve a high charge separation efficiency and robust catalytic activity simultaneously, and the multifunctional composite cocatalyst with strong synergistic effect is a promising cocatalyst for photocatalytic reactions.

## 5. Summary and outlook

### 5.1. Summary

The easy preparation, low cost, visible-light response, suitable band structure, and high stability of metal-free  $\text{g-C}_3\text{N}_4$  make it one of the most promising photocatalysts for  $\text{H}_2$  production from water reduction reactions. However, the inherent Coulomb interaction of photogenerated electron-hole pairs and the existence of an interlayer potential barrier in  $\text{g-C}_3\text{N}_4$  induce the severe bulk recombination of charge carriers, thus greatly reducing the number of active electrons involved in the surface reduction reaction. To oppose the charge recombination and improve the charge separation efficiency, various approaches have been adopted to promote the separation and transfer of photoexcited charge carriers through weakening their Coulomb attraction.

In this perspective, we summarized several common and effective strategies for assisting charge separation in the bulk and surface of  $\text{g-C}_3\text{N}_4$ , including structural modulation, heterojunction construction, and cocatalyst loading. In terms of structural modulation, we discussed the effects of crystal structure, morphology and doping on charge separation and migration. The  $\text{g-C}_3\text{N}_4$  prepared by traditional high-temperature calcination shows poor crystallinity and many defects, resulting in a high charge recombination rate. Hence, many strategies have been developed to promote the polymerization of precursors and synthesize the highly crystalline  $\text{g-C}_3\text{N}_4$ , such as salt-assisted methods, microwave synthesis, and two-step calcination. Meanwhile, for bulk  $\text{g-C}_3\text{N}_4$ , the electrostatic barrier between layers can hinder charge transfer to the surface active sites and induce severe bulk recombination. Thus, many works about morphological tuning of  $\text{g-C}_3\text{N}_4$  have been carried out to change the

original structure of layer-by-layer stacking and facilitate the charge transfer across the interlayers. Designing ultrathin  $\text{g-C}_3\text{N}_4$  nanosheets effectively enlarges the distance between layers, thereby breaking the constraint of interlayer electrostatic barrier and reducing the bulk recombination of photogenerated charge carriers. Otherwise, the tubular morphology of  $\text{g-C}_3\text{N}_4$  also favors charge transfer and separation within the bulk phase. When the 2D  $\text{g-C}_3\text{N}_4$  bends to form the hollow tubular structure, an apparent potential difference is formed between the inner and outer surfaces of tubular  $\text{g-C}_3\text{N}_4$ , which could drive the electron transfer from the inner to the outer surface and greatly improve the bulk charge separation efficiency. Besides the morphological tuning, the doping strategy has also been widely used to regulate the structure of  $\text{g-C}_3\text{N}_4$ , thus facilitating the charge migration and achieving a high charge separation efficiency. In particular, the incorporation of metals and delocalization groups would break the planar  $\pi$ -conjugated structure of  $\text{g-C}_3\text{N}_4$  and induce the redistribution of in-plane delocalized charges, which result in the formation of an intrinsic polarization electric field to promote the in-plane charge transfer along the potential difference. Therefore, the appropriate structural modulation of bulk  $\text{g-C}_3\text{N}_4$  *via* crystallinity regulation, morphology and doping engineering can obviously suppress the charge recombination within the bulk and surface and then improve the charge separation efficiency.

Additionally, the design of  $\text{g-C}_3\text{N}_4$ -based composites, including semiconductor combination and cocatalyst loading, is another effective approach to improve surface charge separation efficiency by promoting charge migration and transfer at the interface of composites. Among various  $\text{g-C}_3\text{N}_4$  composites, the Z-scheme heterojunction has attracted much attention due to its advantage of the directional charge transfer driven by a built-in electric field. In a Z-scheme  $\text{g-C}_3\text{N}_4$  heterojunction, the intimate contact of two semiconductors with different work functions induces interfacial charge reorganization and the formation of a built-in electric field at the interface, which can drive the photogenerated charges across the interface and consume ineffective charges, thus achieving the spatial separation of charge carriers and significantly minimizing the surface recombination. Similarly, for the cocatalyst loaded  $\text{g-C}_3\text{N}_4$  system, the difference in work function also causes the charge rearrangement at the interface and results in the formation of a Schottky barrier, which controls the flow of the photogenerated electrons to the cocatalyst and suppresses the back recombination of charge carriers.

### 5.2. Outlook

This progress can indeed improve the separation efficiency of photogenerated electron-hole pairs and achieve robust photocatalytic  $\text{H}_2$  evolution activity. However, the solar-to-hydrogen conversion efficiency of the  $\text{g-C}_3\text{N}_4$  system still struggles to meet the commercial level, due to the difficulty in synergistically regulating the separation and transfer of whole charge carriers, including bulk phase separation and surface separation. For the above discussed studies, the crystallinity regulation, doping engineering, heterojunction design, and cocatalyst loading can greatly promote the charge transfer within the plane or at the interface, which effectively suppress the surface

recombination of charge carriers. While the morphological tuning mainly favors the charges to transfer from the bulk to surface, its role in promoting the surface charge separation is not so prominent. Therefore, although the above modification approaches have achieved many positive results, there are still certain limitations in improving the photocatalytic performance of g-C<sub>3</sub>N<sub>4</sub> based on a single strategy since each strategy has its own advantages and weaknesses. To facilitate the charge transfer and separation both in the bulk phase and surface and achieve an excellent whole charge separation efficiency, it is essential to combine multiple methods to synergistically modify the pristine g-C<sub>3</sub>N<sub>4</sub>.

Furthermore, there is another issue that is often overlooked in the charge transfer process, the randomness of charge transfer paths. Although the introduction of a potential difference or built-in electric field in g-C<sub>3</sub>N<sub>4</sub> can induce directional transfer of charge carriers from one component to another, the charge movement and transfer paths are random, which can easily trigger charge recombination during the charge migration. Thus, constructing charge transport channels to achieve ordered charge migration is another key to further improving charge separation efficiency and photocatalytic activity.

In the meantime, the transfer mechanism of photogenerated charge carriers and photocatalytic mechanism still need detailed and in-depth understanding and research, which directly affect the structural design of the g-C<sub>3</sub>N<sub>4</sub> system. For instance, the type II and direct Z-scheme heterojunction systems possess similar band structures; precisely probing interfacial charge transfer is crucial to distinguishing the heterojunction types. Hence, it is necessary to combine various characterization techniques to investigate the behavior of photogenerated charge carriers in photocatalytic reactions, including the surface photovoltage technique, transient photocurrent response, electrochemical impedance spectroscopy (EIS), photoluminescence (PL), X-ray photoelectron spectroscopy (XPS), *etc.* In particular, the advanced characterizations of time-resolved technology and in-operando technology need to be developed to explore the dynamics and specific migration pathways of photogenerated charge carriers.

In summary, continuous efforts are required to develop novel and robust g-C<sub>3</sub>N<sub>4</sub> photocatalytic systems for practical application. There are some recommendations for the design of g-C<sub>3</sub>N<sub>4</sub> photocatalysts: (1) combining the advantages of multiple methods to modify g-C<sub>3</sub>N<sub>4</sub> and improve the overall charge separation efficiency, including bulk phase and surface separation; (2) constructing efficient in-plane and interlayer joint charge transport channels to achieve ordered migration of overall photogenerated charge carriers; (3) developing various and advanced characterizations to clearly reveal the charge transfer and separation mechanism, thereby constructing the desired g-C<sub>3</sub>N<sub>4</sub> system.

## Author contributions

Mengmeng Shao: investigation, writing – original draft. Yangfan Shao: data curation, writing – review & editing. Hui Pan: conceptualization.

## Conflicts of interest

There are no conflicts to declare.

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