1. Introduction

One of the most feasible options for the mitigation of carbon dioxide emissions from fossil fuel power stations is post-combustion carbon capture (PCCC) using amines.1–4 Monoethanolamine (MEA) is considered as a benchmark solvent in terms of performance, and has been extensively investigated for use as a PCCC solvent although amines blends are most likely to be used in commercial applications.3,5–7 A range of amines use as a PCCC solvent although amines blends are most likely to be used in commercial applications.3,5–7 A range of amines are being investigated, including other primary amines such as 2-amino-2-methylpropan-1-ol, secondary amines such as N,N-dimethylethanolamine (DMEA) and triethanolamine, and tertiary amines such as N-methyl-2-amino-2-methylpropan-1-ol, secondary amines such as N,N-dimethylethanolamine (MMEA) and triethanolamine (MEA) have been investigated, including other primary amines such as 2-amino-2-methylpropan-1-ol, secondary amines such as piperazine (PZ) and diethanolamine, and tertiary amines such as N-methyl-2-amino-2-methylpropan-1-ol, secondary amines such as N,N-dimethylethanolamine (MMEA) and triethanolamine.8 Other ethanolamines such as N,N-dimethylethanolamine (MMEA) and N,N-dimethylethanolamine (DMEA) have also been studied as PCCC potential solvents, and for our studies are useful model systems to help understand issues associated with the formation of toxic degradation products in the atmosphere.5,6,9,10 In addition, both MMEA and DMEA have been identified as potential products of the in situ degradation of MDEA during natural gas desulphurization and PCCC.6,11

Given the substantial volumes of amines needed for efficient CO2 capture, large scale use of PCCC will potentially result in significant amine emissions to the atmosphere. A PCCC plant using MEA to remove ~ 1 Mt CO2 per year is predicted to emit 40–160 tonnes of MEA per year.12 Previous studies of MEA, PZ and methyamines have shown that, once these amines are released into the atmosphere, their gas-phase processing, primarily initiated by OH radical reactions, will compete with heterogeneous uptake (lifetimes range from 15 min – 1 hour at peak OH concentrations, typically 5 × 106 molecule cm−3).13–18 The atmospheric removal of amines by reaction with Cl atoms in gas-phase has been found to be a minor sink for amines (lifetime of ~ 20 days for typical peak Cl concentrations, around 2 × 103 atom cm−3).19,20 A major concern with PCCC is the impact of carcinogenic nitrosamines (R2N–NO) and nitramines
The title of this paper is "Xenon Onset in Fission Reactor (XOR) Experiments: Experimental Results and Analysis".

The overall rate coefficient $k_{OH}$ for the reaction $\text{OH} + \text{MMEA}$ and the branching ratios in the reactions of $\text{OH}$ with MMEA ($\text{HOCH}_2\text{CH}_2\text{NH}(\text{CH}_3)$) and DMEA ($\text{HOCH}_2\text{CH}_2\text{NH}(\text{CH}_3)_2$) have not previously been studied.

Here we report on the determination of the rate coefficients of $\text{OH} + \text{MMEA}$ and DMEA reactions and their temperature dependence using laser flash photolysis (LFP) for $\text{OH}$ generation and laser induced fluorescence (LIF) for time-resolved $\text{OH}$ monitoring. We find $\text{OH}$ regeneration in the $\text{OH} + \text{DMEA}$ reaction in the presence of oxygen. The $\text{OH}/\text{OD}$ signals obtained by 248 nm photolysis of (CH$_3$)$_2$N(CD$_2$)$_2$OH (d$_4$)-DMEA in the presence of O$_2$ are then used to determine the dominant H-abstraction site for DMEA.

On the millisecond scale of the $\text{OH} + \text{MMEA}/\text{O}_2/\text{NO}$ and $\text{OH} + \text{MMEA}/\text{O}_2/\text{NO}$ LFP/LIF experiments, $\text{HO}_2$ is generated through reaction (R5) and reacts further with NO (reaction (R6)) to give back $\text{OH}$.

$$\text{HOCH}_2\text{CH}_2\text{NH} + \text{NO} \rightarrow \text{HOCH}_2\text{CH}_2\text{NO}$$

The $\text{HO}_2$ yield and the branching ratios in the initial $\text{OH}$ reaction are determined using the method developed previously for $\text{OH}$ reactions with amines such as dimethylaniline (DMA) and ethylamine (EA). Calculations using the MESMER (Master Equation Solver for Multi-Energy Well Reactions) package show that, at the relatively low pressures used in our experiments, the $\text{O}_2$ addition producing a stabilized peroxy species (reaction (R7)) does not compete with reaction (R5) to affect our results. However, the $\text{O}_2$ addition is in competition with $\text{O}_2$ abstraction at atmospheric pressure.

$$\text{HOCH}_2\text{CH}_2\text{NH} + \text{NO} + \text{O}_2 \rightarrow \text{HOCH}_2\text{CH}_2\text{NO} + \text{O}_2$$

The impact of this work on the downwind formation of nitrosamines and nitramines following emission of MMEA from a typical PCCC plant is assessed using an atmospheric chemistry box model. MMEA and NO$_2$ emissions are based on the PCCC pilot plant located at Mongstad, on the west coast of Norway. The model also incorporates typical gas-phase chemistry of background compounds and MMEA, heterogeneous uptake and reactions, and plume dispersion described by a time dependent Gaussian equation.

2. Experimental

This work has been carried out in a slow-flow pulsed LFP-LIF system that has been described in several previous publications. The flows of the OH/OD precursor, amine (MEA and DMEA: Sigma-Aldrich, ≥99.5%, MMEA and 2-methoxyethylamine (MeOEA): Alfa Aesar, 99%), O$_2$ (if used, BOC 99.999%), nitrogen monoxide (if used, BOC > 99.9%) and bath gas (N$_2$, BOC oxygen free) were regulated through calibrated mass flow controllers, mixed in a manifold and introduced into a stainless steel reactor. The OH precursor was tertiary butylhydroperoxide (Sigma-Aldrich, 70% in water) and the OD precursor was deuterated acetone (acetone-d$_4$) in the presence of O$_2$ or (CH$_3$)$_2$COOD/N$_2$ or D$_2$O. The total pressure in the cell was controlled via a needle valve on the exhaust line and measured using a capacitance manometer. Temperatures were measured close to the observation region using a K-type thermocouple.
DMEA was prepared as a diluted mixture in nitrogen in a glass bulb. Gaseous mixtures of MEA and MMEA could not be prepared because of the low vapour pressures of these amines at 298 K. Therefore, MEA and MMEA were introduced into the reactor from a glass bubbler by flowing N$_2$ gas over liquid amine samples. The concentrations of MEA and MMEA were determined using the technique developed previously using in situ absorption measurements at 185 nm. Absorptions were converted into concentrations using the absorption cross sections of amines, $\sigma_{185\text{ nm},\ \text{MEA}} = (8.53 \pm 0.24) \times 10^{-18}$ cm$^2$ molecule$^{-1}$ and $\sigma_{185\text{ nm},\ \text{MMEA}} = (9.71 \pm 0.37) \times 10^{-18}$ cm$^2$ molecule$^{-1}$ (this work).

Radicals were generated by excimer laser flash photolysis at 248 nm (Lambda Physik 210, typically 5-15 mJ cm$^{-2}$, 5 Hz repetition rate). OH radicals were probed by off-resonance laser induced fluorescence at an excitation wavelength of ~282 nm generated from a YAG pumped dye laser (Powerlite Precision II 8010, Sirah PRSC-DA-24, operating with Rhodamine 6G dye) introduced perpendicularly to the photolysis laser. OH fluorescence generated from a YAG pumped dye laser (Thorn EMI model 9813 QKB) through an interference filter (Andover, 308 ± 10 nm). The time delay between the photolysis and the OH probe laser was varied using an in-house LabView program to build an entire time-dependent OH profile. Depending on signal to noise ratios, 6-20 laser shots were averaged for each time point.

2.1. OH + amine reactions in the absence of NO

The reactions were carried out under pseudo-first-order conditions using amine concentrations in large excess over the initial OH concentration (~1000:1). Under these conditions, in the absence of NO, the fluorescence intensity, $I(t)$, which is proportional to [OH], decayed according to the single exponential eqn (E1) (see the inset of Fig. 1 and Fig. S1 as examples, ESIF$^\dagger$).

$$I(t) = I(0) \exp(-k_{\text{OH}}t) \quad (E1)$$

where $k_{\text{OH}}' = k_{\text{OH}[amine]} + k_{\text{loss}}$. Here $k_{\text{OH}}$ is the bimolecular rate coefficient for the reaction of OH with amine and $k_{\text{loss}}$ is the pseudo-first-order rate coefficient for OH loss by diffusion and reaction with OH precursor. Fig. 1 shows an example of bimolecular plot for the OH + MMEA reaction. The error in the bimolecular rate coefficient is a combination in quadrature of statistical error at the 2$\sigma$ level and an estimated 10% systematic error.

For the reaction OH + DMEA in the presence of O$_2$, the experiments were performed using sufficiently high O$_2$ concentration (~10$^{-16}$ molecule cm$^{-3}$) to obtain single exponential OH decays. Examples of bimolecular plots for the DMEA + OH reaction in the absence/presence of O$_2$ are shown in Fig. S1 (ESIF$^\dagger$).

2.2. OH + MEA/MMEA and OD + deuterated MEA/MMEA reactions in the presence of O$_2$/NO

Pseudo-first-order conditions were ensured by having amine (0.5-6.0 × 10$^{-14}$ molecule cm$^{-3}$), O$_2$ and NO concentrations in great excess over the initial radical concentration (0.5-2.0 × 10$^{-11}$ molecule cm$^{-3}$). The O$_2$ concentration was higher by a factor of ~100 than the NO concentration and typically [NO] = 0.1-1.0 × 10$^{-15}$ molecule cm$^{-3}$. Biexponential decays of OH were generated under these conditions due to reaction (R6) regenerating OH and were analysed as described previously.

In the reactions of OD with DOCH$_2$CH$_2$ND$_2$ (d$_2$-MEA) and DOCH$_2$CH$_2$ND(CH$_3$)$_2$ (d$_2$-MMEA) (CH$_3$)$_2$COOD was used as OD precursor. The deuteration of amine and radical precursor were achieved by H/D exchange in the delivery tubing of the cell using D$_2$O (99.9 atom% D), as described elsewhere.

2.3. Synthesis of [1,1,2,2-$^2$H$_4$]-2-dimethylaminoethanol ($d_4$-DMEA)

Ethyl 2-(dimethylamino)-2-oxoacetate (5.00 g, 34.4 mmol) was added slowly over 20 min to a stirred solution of LiAlD$_4$ (3.47 g, 82.6 mmol) in dry THF (200 mL) at 0 °C. The mixture was heated under reflux for 6 h and the reaction cooled to 0 °C. The reaction was quenched by the drop-wise addition of saturated Na$_2$SO$_4$ solution until effervescence ceased and a white solid formed. After stirring for 30 min, the precipitate was collected by filtration and extracted into tetrahydrofuran (600 mL) by Soxhlet extraction for 18 hours to give a pale yellow solution. The solvent was removed in vacuo to give a yellow oil, which was purified by distillation (131-134 °C) to give $d_4$-DMEA (1.98 g, 62%) as a colourless oil; $^1$H NMR (500 MHz, CDCl$_3$); $\delta$ 2.04 (6H, s, CH$_3$); $^{13}$C NMR (125 MHz, CDCl$_3$); $\delta$ 45.2 (CH$_3$); 57.9 (1/COD 21.7, CD$_2$N), 60.0 (1/COD 20.2, CD$_2$OH); m/z (EI$^+$) 93.1.

3. Results and discussion

3.1. Kinetics of OH + MMEA and DMEA

The averages of the measured room temperature rate coefficients, $k_{\text{OH+MMEA}} = (8.26 \pm 0.82) \times 10^{-11}$ cm$^3$ molecule$^{-1}$ s$^{-1}$ and $k_{\text{OH+DMEA}} = (7.29 \pm 0.72) \times 10^{-11}$ cm$^3$ molecule$^{-1}$ s$^{-1}$, are similar to $k_{\text{OH+MMEA}} = (7.61 \pm 0.76) \times 10^{-11}$ cm$^3$ molecule$^{-1}$ s$^{-1}$. The results are in good agreement with $k_{\text{OH}}10^{-11}$ cm$^3$ molecule$^{-1}$ s$^{-1}$ obtained using the SARs. For MMEA and DMEA (see Table S2, ESIF$^\dagger$). This is the first determination of $k_{\text{OH+MMEA}}$.
between the measurement of Harris and Pitts, \((4.7 \pm 1.2) \times 10^{-11} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}\),\(^{37}\) and the measurement of Anderson and Stephens, \((10.3 \pm 2.0) \times 10^{-11} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}\).\(^{38}\)

Both previous studies used the photolysis of water in the VUV \((\lambda \geq 105 \text{ nm})\) and \(\lambda > 160 \text{ nm}\)\(^{38}\) to generate \(\text{OH}\), which was probed by resonance fluorescence using a microwave discharge lamp. The measurement of Anderson and Stephens almost overlaps with our measurement. The finding of \(\text{OH}\) recycling by Harris and Pitts if small amounts of oxygen were present in these experiments.

Fig. 2 shows the temperature dependence of \(k_{\text{OH-MMEA}}\) and \(k_{\text{OH-DMEA}}\). The negative temperature dependence observed for MMEA is similar to the temperature dependence found for MEA.\(^{16}\) The rate coefficient for \(\text{OH} + \text{DMEA}\) reaction decreases more slowly with temperature and at 500 K is about 85\% relative to the value at 298 K. The result is consistent with the only other temperature dependent study for \(\text{OH} + \text{DMEA}\) reaction, which reports no evident temperature dependence over the range 234–364 K.\(^{38}\)

In addition, the reaction between DMEA and OD was studied. As it can be observed in Fig. 2, the rate coefficients for \(\text{OH} + \text{DMEA}\) and OD + DMEA are similar, suggesting a minimal isotope effect. This result is in agreement with our previous finding that the rate coefficients for the reactions of methylamines with \(\text{OH}\) and OD are similar.\(^{13}\)

### 3.2. Kinetics of \(\text{OH} + \text{MEA}, \text{MMEA} \text{ and } \text{DMEA} \text{ reactions in the presence of oxygen}\)

Fig. 3\(^{1}\) shows examples of bimolecular plots for \(\text{OH} + \text{DMEA}\) reaction obtained using \(\text{N}_2\) (100\%) and \(\text{O}_2\) (10\%) + \(\text{N}_2\) (90\%), respectively. It can be clearly seen that the bimolecular rate coefficient in the presence of excess \(\text{O}_2\), \(k_{\text{O}_2}\), is lower than the bimolecular rate coefficient in the absence of \(\text{O}_2\), \(k_{\text{N}_2}\), which is indicative of rapid \(\text{OH}\) regeneration. The OH yield, \(\varphi_{\text{OH}}\), can be calculated from \(k_{\text{O}_2}\) and \(k_{\text{N}_2}\) (eqn \((E2))\), as shown in our previous work.\(^{13}\)

\[
\varphi_{\text{OH}} = 1 - \frac{k_{\text{O}_2}}{k_{\text{N}_2}} \tag{E2}
\]

Our observation of similar \(\text{OH}\) recycling for TMA (trimethylamine) + \(\text{OH}/\text{O}_2\)\(^{13}\) suggests that in the case of DMEA + \(\text{OH}/\text{O}_2\) the hydrogen abstraction occurs predominantly from a position next to the amine group of DMEA. Unlike TMA, not all the \(\pi\)-hydrogens are equivalent, however, the experiments using \(d_4\text{-DMEA}/\text{O}_2\) described in ESI\(^{1}\)\# show that the abstraction from the CH\(_3\) group next to the amine group largely dominates over the abstraction from CH\(_2\) groups. The \(\text{OH}\) recycling is the result of the reaction of \(\text{O}_2\) with \(\text{HOCH}_2\text{CHN(CH}_3\text{)}_2\) radical (reaction (R9) followed by reaction (R10)) generated by the initial abstraction (reaction (R8)).

\[
\text{HOCH}_2\text{CH}_2\text{N(CH}_3\text{)}_2\text{ + OH} \rightarrow \text{HOCH}_2\text{CHN(CH}_3\text{)}_2\text{ + H}_2\text{O} \tag{R8}
\]

\[
\text{HOCH}_2\text{CHN(CH}_3\text{)}_2\text{ + O}_2 \rightarrow \text{HOCH}_2\text{CH(O}_2\text{)N(CH}_3\text{)}_2^* \tag{R9}
\]

\[
\text{HOCH}_2\text{CH(O}_2\text{)N(CH}_3\text{)}_2^* \rightarrow \text{OH} + \text{CH}_3\text{NCH}_2 + \text{HOCH}_2\text{CHO} \tag{R10}
\]

DMEA photolysis at 248 nm in the presence of \(\text{O}_2\) generates the same radical as reaction (R8), \(\text{HOCH}_2\text{CHN(CH}_3\text{)}_2\), which explains the fast generation of \(\text{OH}\) observed in the absence of an OH precursor. By analogy with the recycling mechanistic scheme proposed for TMA/O\(_2\),\(^{13}\) we speculate that the by-products of \(\text{OH}\) regeneration route are N-methylamine, \(\text{CH}_3\text{NCH}_2\), and glycolaldehyde, \(\text{HOCH}_2\text{CHO}\). The decomposition of the activated peroxy species, \(\text{HOCH}_2\text{CH(O}_2\text{)N(CH}_3\text{)}_2^*\) (reaction (R10)) is in competition with collisional stabilization (reaction (R11)).

\[
\text{HOCH}_2\text{CH(O}_2\text{)N(CH}_3\text{)}_2^* + \text{M} \rightarrow \text{HOCH}_2\text{CH(O}_2\text{)N(CH}_3\text{)}_2 + \text{M} \tag{R11}
\]

Therefore, the \(\text{OH}\) yield will decrease with increasing pressure, which indeed was confirmed by experiments performed at different pressures at temperatures of 298 and 363 K (Fig. 3). The reciprocal of the \(\text{OH}\) yield is proportional to the total gas concentration, \([\text{M}]\):

\[
\frac{1}{\varphi_{\text{OH}}} = 1 + \frac{k_{\text{stab}}}{k_{\text{regen}}} [\text{M}] \tag{E3}
\]

Fig. 3 shows the Stern–Volmer plots, \(1/\varphi_{\text{OH}}\) vs. \([\text{M}]\), for the DMEA + \(\text{OH}/\text{O}_2\) system at 298 and 363 K. As expected, the Stern–Volmer gradient decreases with temperature showing that the activated peroxy radical, \(\text{HOCH}_2\text{CH(O}_2\text{)N(CH}_3\text{)}_2^*\), is generated with more energy at higher temperatures and hence dissociates in a greater fraction to eliminate \(\text{OH}\).\(^{39}\) The Stern–Volmer plot at 298 K is similar to that obtained for \(\text{OH} + d_4\text{-DMEA}/\text{O}_2\) at room temperature (Fig. S2, ESI\(^{1}\)).
Table 1 shows the Stern–Volmer parameters and the calculated OH yields for DMEA at atmospheric pressure and 298 and 363 K. The results obtained for TMA at room temperature are also included for comparison. At both temperatures OH regeneration by DMEA/O₂ is minor at atmospheric pressure. At 298 K and 1 atm \( \Phi_{\text{OH}(\text{DMEA})} \) is about two times lower than \( \Phi_{\text{OH}(\text{TMA})} \). The more efficient quenching of the OH + DMEA/O₂ is to be expected given the larger size of the system.

The observation of OH regeneration in the OH + DMEA/O₂ reaction, similar to that for the OH + TMA/O₂ reaction, suggests that the direct OH recycling is a common feature of tertiary amines. No OH regeneration was observed in the OH + MEA/MMEA experiments with O₂ over the temperature range 298–500 K. In addition, no OH signal was obtained following MEA and MMEA photolysis in the presence of oxygen. This result is in line with the lack of direct OH recycling in the case of other amines containing N–H bond(s), such as methylamine, DMA and PZ.

### 3.3. OH + MEA/MMEA reactions in the presence of O₂/NO

The observed generation of HO₂ in the OH + MEA/MMEA reactions with O₂ may occur through OH abstraction at either the \( \alpha \) or \( \beta \) position, followed by abstraction at the N–H and O–H sites, respectively, by \( \text{O}_2 \). Therefore, the HO₂ yield for MEA, \( \Phi_{\text{HO}_2(\text{MEA})} \), is equal to the sum of the branching ratios for reactions (R1a) and (R1c), \( r_{1a} + r_{1c} \). Experiments using the OH + 2-methoxyethylamine (MeOEA)/O₂/NO system give a similar fraction of HO₂ generation. In MeOEA, abstraction \( \beta \) to the N atom cannot lead to HO₂ production, suggesting that \( r_{1a} + r_{1c} \approx r_{1a} \) (ESI†), in agreement with previous experimental studies reporting that the \( \alpha \) position is the dominant abstraction site in the OH + MEA reaction. Therefore, \( \Phi_{\text{HO}_2(\text{MEA})} = r_{1a} \). The result is in disagreement with the theoretical calculations of Xie et al., which found that both C–H abstractions are significant for MEA.

Our result for MEA and the finding that the abstraction from DMEA by OH occurs predominantly from the \( \alpha \) methylene group (vide supra) suggest that the abstraction from the \( \alpha \) CH₂
Table 2 Yields of HO2/DO2 and branching ratio for abstraction at αC–H site for OH + MEA/MMEA and OD + d1-MEA/d2-MMEA reactions at 298 K

<table>
<thead>
<tr>
<th>p (Torr)</th>
<th>MEA</th>
<th>MMEA</th>
<th>d1-MEA</th>
<th>d2-MMEA</th>
</tr>
</thead>
<tbody>
<tr>
<td>15</td>
<td>—</td>
<td>—</td>
<td>0.71 ± 0.05 &amp; 0.57 ± 0.09</td>
<td></td>
</tr>
<tr>
<td>20</td>
<td>0.64 ± 0.04 &amp; 0.48 ± 0.03</td>
<td>0.75 ± 0.07 &amp; 0.54 ± 0.06</td>
<td></td>
<td></td>
</tr>
<tr>
<td>30</td>
<td>—</td>
<td>—</td>
<td>0.61 ± 0.06 &amp; 0.51 ± 0.04</td>
<td></td>
</tr>
<tr>
<td>40</td>
<td>0.61 ± 0.06 &amp; 0.45 ± 0.06</td>
<td>0.62 ± 0.07 &amp; —</td>
<td></td>
<td></td>
</tr>
<tr>
<td>60</td>
<td>0.61 ± 0.06 &amp; 0.51 ± 0.04</td>
<td>0.62 ± 0.07 &amp; —</td>
<td></td>
<td></td>
</tr>
<tr>
<td>150</td>
<td>—</td>
<td>—</td>
<td>—      &amp; —</td>
<td></td>
</tr>
</tbody>
</table>

* HO2/DO2 yields was determined as described previously and additionally presented in ESI. For MEA rαC–H = r1α. Combination in quadrature of statistical errors at the 2σ level and an estimated 5% systematic error.

The branching ratios, the HO2 yield for OH + MEA/MMEA and the DO2 yield for OD + d1-MEA/d2-MMEA were assigned by the method developed in previous work. Table 2 shows the HO2/DO2 yields at room temperature and several total pressures. The values in Table 2 represent the means of a series of measurements where either [NO] or [O3] was varied (typically by one order of magnitude).

The slightly higher value of rαC–H for d2-MEA and d2-MMEA compared to rαC–H for MEA and MMEA suggests that the NH2 deuterium results in less abstraction from the amine group due to a small kinetic isotope effect (KIE) and in agreement with the results for undeuterated and deuterated DMA and EA.

Our result, r1α = 0.62 ± 0.06 for MEA, is in between of the value obtained by the EUPHORE studies of Karl et al., r1α = 0.80, and the value predicted using the SARs (Table S2, ESI†), r1α = 0.45. Our work used a direct method for determination of branching ratios, while EUPHORE studies calculated the branching ratios by modelling the observed time profiles of end-products. The maximum branching ratio for N–H abstraction is rN–H = 1 − rαC–H and represents an upper limit for the carboxylic compound production. For MMEA rN–H = 0.52 ± 0.06, which is 30% lower than the value obtained using the SARs, rN–H(MMEA) = 0.74, and is higher than the values found for the carbon capture amines MEA and PZ, rN–H(MEA) = 0.38 ± 0.06 and rN–H(PZ) = 0.09 ± 0.06.18

The result suggests that the PCCC emission of MMEA has a higher potential to form toxic compounds than the emission of MEA and PZ, although all secondary amines are of particular concern.

4. Atmospheric modelling

4.1. Model description

Simulations were performed for the Mongstad PCCC plant emissions using two chemical box models. The first model contained typical gas-phase concentrations of volatile organic compounds (VOCs) observed at a European coastal background site - Mace Head, Ireland. The background concentrations of NOx, NO2 and O3 were constrained to the hourly averaged diurnal cycles measured for the north-westerly wind sector at the Weybourne Atmospheric Observatory located on the Norfolk coast of England in September 2002 (Fig. S7–S9, ESI†). The Weybourne data are representative for summer conditions and have been chosen as their daily averages: [O3] = 39 ppbv, [NO] = 0.3 ppbv and [NO2] = 3 ppbv best describe the moderately polluted background conditions at Mongstad. The background model was based on detailed chemistry from the Master Chemical
Mechanism (MCM v3.2), with the extracted model chemical mechanism containing 2255 species participating in 6920 gas-phase reactions. The model was run for four days re-initialising the background concentrations after each day, in order to generate model intermediate species in steady state. Fig. S10 (ESI†) shows the diurnal cycle of OH radical as an example. The concentrations at the end of four days were input in a second box model (the plume model), which contained both the background reactions and the OH initiated oxidation of the emitted MEA (vide infra), which was run for a further day. The time steps in the numerical simulations were 15 seconds in the clear day scenario and 1 second in the cloudy day scenario (vide infra). During these simulations the concentrations of NO2 and O3 were constrained to the one minute averages measured at Weybourne in September 2002 (ESI†). As the PCCC emits NO (133 tonnes per annum), the [NO] in the plume initially exceeds [NO] outside the plume and hence it is diluted to ambient values by the plume dispersion (vide infra).

The amine chemistry scheme used in the plume model contained 20 species involved in 24 gas-phase reactions, reversible partitioning of MEA and the nitrosamine and nitramine products between the gas-phase and aqueous phase particles, and reactions of MEA, R2N–NO and R2N–NO2 with OH in the liquid phase (ESI†). Heterogeneous uptake was represented by a parameterization of a first-order loss process to the aerosol surface:46

$$k_i' = \left( \frac{r}{D_i} + \frac{4}{\gamma_i c_i^f} \right)^{-1} A$$

(E4)

where $k_i'$ is the first-order rate coefficient for heterogeneous loss, $r$ is the aerosol particle effective radius, $D_i$ is the gas-phase diffusion coefficient of the species $i$, $\gamma_i$ is the uptake coefficient for the gas $i$, $c_i^f$ is the mean molecular speed of the particle and $A$ is the aerosol surface area per unit volume. More details on the parameters used in the calculation of $k_i'$ are given in ESI†. The first-order rate coefficient describing the transfer from aerosol to gas-phase, $k_b'$, was calculated using the Henry’s law coefficient, $H_i$:44

$$\frac{k_i'}{k_b'} = RTH_i$$

(E5)

Here $R$ is the gas constant, $T$ is the temperature in Kelvin and $H_i = \frac{[i]_{aq}}{p_i}$, where $[i]_{aq}$ is the molar concentration of compound $i$ in the aqueous particles and $p_i$ is the partial pressure of gas $i$ in the plume emitted by the PCCC plant.

Two scenarios were studied: a clear day where $r = 0.5 \mu m$ and atmospheric liquid water content, $g_w = 10^{-4} g m^{-3}$47,48 and a cloudy day where $r = 10 \mu m$ and $g_w = 0.5 g m^{-3}$,19,47 In the cloudy day scenario an attenuation of the actinic flux by a factor of 0.5 relative to the actinic flux under clear sky conditions was applied.49 Under the clear sky conditions emissions were set at 10:00, 14:00 and 24:00, while under the cloudy sky conditions the emissions were set at 10:00.

Emissions were set to 1.27 g s−1 of MEA (40 tonnes per year) and 4.22 g s−1 of NO34 in mid-July and $T = 288.15 K$. Species concentrations within the PCCC plume were modelled within a well-mixed box transported in downwind direction. Numerical simulations were started at 10 min after emission, when the plume was considered to have filled the boundary layer (set to 1000 m depth) by vertical mixing. Therefore, the box had a fixed height of 1000 m. The width of the plume, $y(t)$, has been expanding via horizontal (cross-flow) mixing with the ambient air starting from $y(0) = 3816 m$ at day time and from $y(0) = 1658 m$ at night (ESI†):50,51

$$y(t) = \left[ y(0)^2 + 8K_y t \right]^{1/2}$$

(E6)

Here $K_y$ is the horizontal diffusion coefficient, 3034 m2 s−1 for day time emissions and 573 m2 s−1 for the midnight emission (ESI†).

The emitted amine, NO and the targeted products nitrosamine (R2N–NO) and nitramine (R2N–NO2), were diluted with ambient air according to a Gaussian equation:

$$\frac{d[i]_g}{dt} = \frac{4k_i}{\gamma(0)^2 + 8K_y} \left( [i]_g - [i]_g^b \right)$$

(E7)

where $[i]_g$ is the gas-phase concentration of compound $i$ within the plume and $[i]_g^b$ is the concentration of $i$ outside the plume. In order to describe the time evolution of NO, a separate model was used to simulate the dispersion of the emitted NO to ambient values (eqn (E7)). Ambient NO was constrained to the one minute averaged diurnal cycle of NO measured at Weybourne.42 During each box model simulation incorporating the amine emission, NO was constrained to follow the profile generated by the dispersion model (see Fig. S12 as an example, ESI†), while NO2 and O3 were constrained to Weybourne data (vide supra).

4.2. Results of the simulations

4.2.1. Day time emissions in the clear day and the cloudy day. The MEA losses under clear sky conditions are: 48% reaction with OH and 52% dispersion for emission at 10:00 and 51% reaction with OH and 49% dispersion for emission at 14:00, with heterogeneous uptake negligible in both cases. In the cloudy day simulation, emission at 10:00, MEA losses are: 93% uptake, 4% gas-phase reaction and 3% dispersion. Therefore, under cloudy conditions, only a relatively small fraction of MEA is available for reaction with OH in gas-phase. Fig. 5 and 6 show the rates of the MEA losses emitted at 10:00 under clear and cloudy sky conditions, respectively. The dominant clear sky losses, dispersion and reaction with OH, become minor in the cloudy day scenario, where the initial rate of uptake is two orders of magnitude faster than the rates of dispersion and reaction. MEA is almost completely removed in the first 2–3 hours after emission in the clear day and in about one minute in the cloudy day.

Fig. 7 and 8 show the concentrations of nitramine and nitrosamine in the PCCC plume vs. time and downwind distance from the emission point in the cloud free and cloudy sky scenarios, respectively. The time was converted into distance by using a wind speed of 3 m s−1. Fig. 7a shows the profile of nitramine concentration for emissions at 10:00 and 14:00. Emissions at 10:00 and 14:00 amplifies gas phase chemistry as at 10:00 NO...
and NO\textsubscript{2} reach their maximum, while at 14:00 O\textsubscript{3} peaks, maximising OH concentration (ESI\textsuperscript{†}). The maximum of [R\textsubscript{2}N–NO\textsubscript{2}] is \textasciitilde30\% higher for the 10:00 emission than in the 14:00 scenario and is three times above the current safe guideline level of 0.3 ng m\textsuperscript{-3} (annual average), as established by the Norwegian Institute for Public Health (NIPH).\textsuperscript{21} Owing to a lack of toxicity data on nitramines, the NIPH decided to use the cancer risk estimated for N\textsuperscript{-}nitrosodimethylamine (DMA–NO) for the sum of the concentrations of nitramines and nitrosamines in air. The restriction of the air level of total nitrosamines and nitramines to 0.3 ng m\textsuperscript{-3} may be regarded as a conservative limit as DMA–NO was found to have \textasciitilde6 times higher cancer potency than N\textsuperscript{-}nitrodimethylamine, DMA–NO\textsubscript{2}.\textsuperscript{21} However, recent studies reported that DMA–NO\textsubscript{3} is less toxic than 2-nitroaninoethanol and emphasized the need for additional investigations of the cancer potency of nitramines.\textsuperscript{52}

For both 10:00 and 14:00 cases the concentration of R\textsubscript{2}N–NO\textsubscript{2} reaches its maximum after ca. 11 km (~1 hour), where R\textsubscript{2}N–NO\textsubscript{2} formation becomes less significant than R\textsubscript{2}N–NO\textsubscript{2} loss (OH reaction and dispersion). [R\textsubscript{2}N–NO\textsubscript{2}] goes below the health limit after \textasciitilde130 km downwind of the emission. The concentration of nitrosamine peaks early (after \textasciitilde12 min or 2 km from source) at a value about 20 times lower than the safety limit (0.3 ng m\textsuperscript{-3}), after which photolytic degradation and dispersion start to dominate.

Assuming that nitrosamine isomerises as soon as it is formed leading to an imine;\textsuperscript{23,25,30} for emission at 10:00 under clear sky conditions, the obtained concentration of R\textsubscript{2}N–NO\textsubscript{2} is practically the same as [R\textsubscript{2}N–NO\textsubscript{2}] obtained considering that nitrosamine does not isomerise but photolyses to generate back R\textsubscript{2}N radical. Note that, assuming nitrosamine formation, the rate of nitramine formation is 10 times higher than the rate of nitrosamine formation.

Fig. 8 shows that, under cloudy conditions, the maximum value of [R\textsubscript{2}N–NO\textsubscript{2}] + [R\textsubscript{2}N–NO] is \textasciitilde75 times lower than the safety level. After a brief increase (over the first 12 s of the numerical simulation), both R\textsubscript{2}N–NO\textsubscript{2} and R\textsubscript{2}N–NO are almost completely removed from the gas-phase within 90 s by heterogeneous uptake.

4.2.2. Midnight emission under clear sky conditions.
Numerical simulations performed using the background model
shows that the concentration of NO₃ radicals (the dominate night time oxidant)²³ peaks at ~10⁸ molecule cm⁻³ at midnight. Similarly to OH, NO₃ abstracts from both C–H and N–H sites of amines. Abstraction at the N–H site generates a nitrogen-centred radical which produces nitrosamine and nitramine by reacting with NO (reaction (R2)) and NO₂ (reaction (R3)).

The rate coefficient of the reaction of MEA with NO₃, \( k_{\text{MEA-NO}_3} \), has not been determined experimentally. Karl et al. calculated \( k_{\text{MEA-NO}_3} = 1.5 \times 10^{-13} \) cm³ molecule⁻¹ s⁻¹ using a structure–property relationship,²³ in good agreement with the SAR estimate \( k_{\text{MEA-NO}_3} = 1.4 \times 10^{-13} \) cm³ molecule⁻¹ s⁻¹ found by Carter.²⁴ However, assuming that the linear correlation between the reactions of OH and NO₃ with the same substrate²³,²⁵ is followed by amines, a value of \( 1 \times 10^{-12} \) cm³ molecule⁻¹ s⁻¹ is obtained for \( k_{\text{MEA-NO}_3} \). As shown below, using the higher estimates for the rate coefficient of NO₃ with MEA leads to significant nitramine and nitrosamine formation. In addition, there are no data available for branching in the MEA + NO₃ reaction. Here we consider the worst case scenario where 100% of the MEA + NO₃ reaction proceeds to nitrogen-centred radical.

Fig. 9a and b show the time and the downwind distance dependences of [R₂N–NO₂] and [R₂N–NO], respectively in three cases: the MEA + NO₃ reaction is negligible and hence not considered in the model (the base case), \( k_{\text{MEA-NO}_3} = 1.5 \times 10^{-13} \) cm³ molecule⁻¹ s⁻¹ and \( k_{\text{MEA-NO}_3} = 1 \times 10^{-12} \) cm³ molecule⁻¹ s⁻¹. By using \( k_{\text{MEA-NO}_3} = 1.5 \times 10^{-13} \) cm³ molecule⁻¹ s⁻¹ an 8% increase over the base case in the total nitramine formed in the first 24 hours after emission is produced, while the value of \( 1 \times 10^{-12} \) cm³ molecule⁻¹ s⁻¹ produces a 50% increase in the total nitramine formed in the first 24 hours. The change in the total nitrosamine generated over one day is more significant due to the lack of photolytic destruction at night: a 20% increase using \( 1.5 \times 10^{-13} \) cm³ molecule⁻¹ s⁻¹ and a 100% increase using \( 1 \times 10^{-12} \) cm³ molecule⁻¹ s⁻¹. In all three midnight scenarios the total amount of carcinogenic compounds formed over 24 hours is significant, several times larger than the one produced by the day time emissions (vide supra).

In all three midnight scenarios, dispersion is the main loss of MEA within the first five hours. As the dispersion rate decreases fast in time (eqn (E7)) and in the morning OH concentration starts to increase, between 05:00–10:00 the rate of the MEA + OH reaction increases, dominating over the dispersion rate (Fig. S13, ESI†) and leading to more nitramine formation. After reaching a maximum of about 3–4 times above the safety level at ~10:00, nitramine is consumed by reaction with OH (Fig. 9a). Nitrosamine formed during the night is 15–30 times lower than the tolerance limit and photolyses rapidly upon sunrise (Fig. 9b). A small increase in R₂N–NO occurs between 06:30–08:00 due to the increase in both OH and NO levels in the morning.

### 4.2.3. Comparison with the previous atmospheric modelling studies

Previous studies of the formation of nitrosamines and nitramines by the atmospheric oxidation of carbon capture amines have addressed the MEA emissions from the Mongstad PCCC pilot plant using chemical transport models,²⁴,²⁵ incorporating much more simplified atmospheric chemistry schemes than used in this study, as well as dispersion and wet deposition. In addition, Karl et al. included dry deposition. The present study uses the MEA and
NO emissions reported by Karl et al. In the study of Karl et al., the yearly average sum of the concentrations of $R_2N-NO$ and $R_2N-NO_2$ at ground level air is 1 to 3 orders of magnitude lower than the safety limit of 0.3 ng m$^{-3}$. These results are explained by a relatively small contribution of the reaction with OH to the total yearly MEA removal. In the baseline case used by Karl et al., the yearly average losses of MEA in a 200 km $\times$ 200 km area around the PCCC were: ~50% transport by diffusion and advection, 29% dry deposition, 17% wet deposition and 5% reaction with OH.

The previous investigations used a branching ratio for OH abstraction at NH$_2$ group of 0.08 while here we considered that the abstraction at N–H site to be ~5 times greater ($r_{SH} = 0.38$). Karl et al. identified that the branching ratios in the initial amine + OH reaction is one of the main uncertainties affecting their results. In line with these authors we have found that the sum of air concentrations of $R_2N-NO$ and $R_2N-NO_2$ increases linearly with the branching ratio for the OH abstraction at the N–H site. Koeijer et al. found that $[R_2N-NO] + [R_2N-NO_2]$ is 30–180 times lower than the defined safety level for emissions of MEA 10–60 times lower than in the present work and Karl et al.

Our work treats the atmospheric chemistry in detail, while using a simple representation of the dispersion of amines and their toxic degradation products. The results show that, under cloud free conditions, for both day and night PCCC emissions, nitramine can reach levels significantly above the generally adopted tolerance limit of 0.3 ng m$^{-3}$. The peak concentration for our scenario (horizontal wind speed of 3 m s$^{-1}$) is at distance of ca. 10 km (~1 hour) from the source for the daytime emissions and ca. 100 km (~10 hours) from the emission point for the midnight scenario. For both day and night emissions the safety limit is still not met after 100 km away from the point of maximum health risk. The results obtained for the midnight scenario are more uncertain as $k_{MEAN-NO_2}$ and the H atom abstraction branching ratios in the MEA + NO$_3$ reaction are not well known. Results of numerical simulations using estimates of $k_{MEAN-NO_3}$ are very sensitive to the value of this parameter. An order of magnitude increase of $k_{MEAN-NO_3}$ generates an important increase in the total amount of gas-phase toxic products formed over 24 hours after emission, by 50% for nitramine and by 100% for nitrosamine. It is therefore important that more research is carried out in order to constrain the rate constants and H-abstraction branching ratios in the reaction of NO$_3$ with ethanol amines.

**5. Conclusions**

The rate coefficients for the OH radical reactions with N-methylethanolamine (MMEA) and N,N-dimethylethanolamine (DMEA) have been determined as a function of temperature using the laser flash photolysis – laser induced fluorescence technique. The result for the OH + DMEA reaction is similar to previous results reported for a series of OH + amine reactions. This is the first kinetic study of the OH + MMEA reaction and shows that the room temperature rate coefficient $k_{OH-MMEA}$ and its temperature dependence are close to the results obtained for the OH + MEA reaction. The rate coefficients for both MMEA and DMEA at 298 K are in good agreement with the calculations using the structure–activity relationships (SARs). Similarly to the OH + trimethylamine (TMA)/O$_2$ reaction, the OH + DMEA/O$_2$ reaction has been found to recycle OH. The result suggests that OH recycling through a chemical activated peroxy radical is a common signature for tertiary amines.

The dominant abstraction site in the OH reaction with mono-ethanolamine (MEA), MMEA and DMEA has been found to be the methylene group next to the amine group. The branching ratio for abstraction at the $\alpha$C–H site of MEA is in between the value obtained by the EUPHORE atmospheric chamber study of Karl et al. and the value calculated using SARs (ESI$^+$$^1$). Abstraction at the $\alpha$C–H site generates a HOCH$_2$CHNR radical, which reacts further with O$_2$. Calculations using MESMER shows that at atmospheric pressure the HOCH$_2$CHNR + O$_2$ reaction occurs predominantly through the abstraction route, producing HO$_2$ and 2-imino-ethanol. The result is in agreement with previous findings for O$_2$ reactions with RCHNHR (R = CH$_3$ or H) radicals.

The production of carcinogenic nitramines and nitrosamines depends on the branching ratio for abstraction at the N–H site, $r_{SN-H}$. As $r_{SN-H}$ is higher for MMEA than for the standard post-combustion carbon capture (PCCC) solvent MEA, MMEA has a higher potential to form toxic compounds by atmospheric oxidation than MEA.

The health risks due to the formation of carcinogenic compounds following a typical MEA emission into the troposphere are assessed using a modified box model focusing on the detailed chemistry. Model plume emissions are based on the Mongstad PCCC pilot plant, located on the west coast of Norway. The results show that on clear, cloud free days concentrations of carcinogenic nitramine formed by the atmospheric photo-oxidation of the benchmark PCCC solvent MEA, exceed the recommended annual average value of 0.3 ng m$^{-3}$. The studies also show that heterogeneous uptake significantly reduces atmospheric concentrations and so the annual average concentration will be very dependent on local meteorology.

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