Room temperature synthesis of reduced TiO₂ and its application as a support for catalytic hydrogenation†

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Reduced TiO₂ (TiO₂-x) materials have attracted increasing attention due to their large solar absorption and high photo-activity. However, their synthesis procedures usually involve harsh conditions, such as high temperature and/or high pressure. Herein, a facile solid ball-milling method for the synthesis of TiO₂-x under ambient conditions was developed. By using finely dispersed Na/NaCl powders as the reducing agent and TiO₂ (P25, Degussa) as the precursor, a series of TiO₂-x of 20–30 nm with a controllable reduction degree can be successfully synthesized through adjusting the reaction conditions. The surface area of TiO₂-x is much larger than that of pristine TiO₂, showing its great potential as a catalyst support in chemical reactions. Our experimental results show that uniform Ru particles with particle size less than 1 nm can be well dispersed on the surface of the TiO₂-x due to the enhanced surface area and plenty of oxygen vacancies in TiO₂-x. As a result, Ru/TiO₂-x exhibited superior activity upon catalytic hydrogenation of N-methylpyrrole in comparison with Ru/TiO₂.

Introduction

Since Honda and Fujishima discovered hydrogen generation from water by using titanium dioxide (TiO₂) as a photo-electrode,¹ TiO₂ has drawn much attention for its applications in pigments,² sunscreens³ and photocatalysis.⁴ Apart from the applications in photo-catalysis, TiO₂ is also a unique material widely employed as a catalyst support for selective catalytic hydrogenation,⁵ oxidation⁶,⁷ and electrochemical⁸,⁹ reactions.

In 2011, Mao et al. found that the reduction of TiO₂ nanoparticles through hydrogenation was an effective strategy to improve visible light absorption and photo-activity of TiO₂.¹₀ Since then, many studies have been devoted to the synthesis of reduced TiO₂ (denoted as TiO₂-x) that has abundant oxygen vacancies and Ti³⁺ species. Annealing TiO₂ precursors in a reducing gas atmosphere (H₂ or H₂ plasma) under high temperature and/or high pressure is a common synthetic route for the preparation of TiO₂-x.¹₁,¹² Other promising chemical methods including Al vapor,¹³ CaH₂ (ref. 15) and NaBH₄ (ref. 16) reduction etc., can also produce TiO₂-x, where high reaction temperatures were required. Furthermore, electrons, Ar⁺ or other high energy particle bombardments have also been employed to produce TiO₂-x materials.¹⁷ Despite those significant advances in the synthesis of TiO₂-x, it is still highly desirable to develop facile and effective synthetic strategies for the scalable synthesis of TiO₂-x under mild conditions.

As a TiO₂ derived material with unique electronic properties, TiO₂-x may show promises as a new kind of catalyst support, however, only limited investigations have been published so far. A model study of Au on a reduced titanium oxide ordered film has demonstrated that the strength of the interaction between over-layer Au and the support comprised of strong bonding between Au and Ti, yielding an electron-rich Au and exhibiting an exceptional high activity for CO oxidation.¹⁸ Very recently, the hydrogen treated TiO₂ nanotube arrays with more oxygen vacancies and hydroxyl groups was synthesized which served as highly ordered nanostructured electrode supports and were able to significantly improve the electrochemical performance and durability of fuel cells.¹⁹ The unique properties of the TiO₂-x supported catalysts may be derived from (1) the encapsulation of metal particles in TiO₂-x, presumably because of the so-called strong metal support interaction (SMSI)²⁰ between metal and TiO₂-x, (2) the strong bonding between the metal atoms at the interface with surface defects (reduced Ti site) or (3) the electrons transfer between metal particles and TiO₂-x.²¹ It is, therefore, very interesting to investigate the performance of TiO₂-x supported catalysts in related chemical reactions.

In this paper, a solid ball-milling reduction process for the synthesis of nanosized TiO₂-x from crystalline TiO₂ (P25,
were used to disperse Na metal by ball milling. A series of TiO$_2$-$x$ samples with color changing from white to dark blue manifesting the increase in the reduction degree can be facilely prepared. The TiO$_2$-$x$ possesses much higher surface area and visible light absorption than those of the pristine TiO$_2$. Highly dispersed Ru particles supported on TiO$_2$-$x$ were prepared and tested as catalyst for the hydrogenation of N-methylpyrrole, which obviously outperformed that of Ru particles supported on pristine TiO$_2$, evidencing stronger promoting effect of TiO$_2$-$x$ on Ru.

**Experimental section**

**Chemicals and materials**

TiO$_2$ were purchased from Degussa. Na, NaBH$_4$, NaCl, and NaOH were purchased and used directly without further purification. Tetrahydrofuran (THF) was purchased from Merck, and dried by molecular sieve before usage.

**Preparation of Na/NaCl fine powders**

Due to the soft and ductile nature of Na metal, NaCl powders were used to disperse Na metal by ball milling. In a typical experiment, Na metal and NaCl powders with a weight ratio of 1/10 were mechanically milled under argon atmosphere using a Retsch PM400 planetary ball milling. All the manipulations were conducted inside a glove box filled with purified argon. The ball milling was carried out at a milling rate of 150 rpm for 4 h at room temperature, and then black Na/NaCl fine powders can be obtained.

**Synthesis of TiO$_2$-$x$**

The Na/NaCl fine powder, composed of small Na particles dispersed by NaCl, is expected to be an effective reducing agent for the reduction of crystalline TiO$_2$ to TiO$_2$-$x$ (Scheme 1). In a typical experiment, crystalline TiO$_2$ was milled with Na/NaCl fine powders with a weight ratio of 1/n (n = 1–4) under argon atmosphere using a Retsch PM400 planetary ball milling, which was carried out at a series of milling rates, such as 80, 120, 150 and 180 rpm at room temperature. The samples milled for 0.25 to 4 hours were collected and washed with deionized water for several times to remove the Na and NaCl. Finally, the obtained TiO$_2$-$x$ products were dispersed in a small amount of deionized water and then vacuum-dried at room temperature to get TiO$_2$-$x$ powders. The synthesized TiO$_2$-$x$ samples are marked as TiO$_2$-$n$-$v$-$t$, where n, t and v stand for weight ratio between Na/NaCl and TiO$_2$, ball milling rate and reaction time, respectively. For example, TiO-4-80-1 means that the obtained TiO$_2$-$x$ was synthesized at a milling rate of 80 rpm for 1 h, and the weight ratio of Na/NaCl fine powders to P25 is 4.

**Preparations of catalysts and characterizations**

5% Ru/TiO$_2$ and 5% Ru/TiO$_2$-$x$ catalysts were both prepared by a deposition method using NaBH$_4$ as the reducing agent (Scheme 1). The support TiO$_2$ or TiO$_2$-$x$, was added to a certain concentration of RuCl$_3$ aqueous solutions and stirred for 6 h. Then a NaBH$_4$ solution was added slowly to reduce the Ru$^{3+}$ cations with intensive stirring. Finally the powders were filtered, washed with deionized water and dried under vacuum overnight.

Powder X-ray diffraction (XRD) patterns were recorded on an X’Pert Pro [PANalytical] diffractometer with Cu K$_\alpha$ radiation at 40 kV and 40 mA. Raman spectra were recorded with a Renishaw Raman spectrometer equipped with a He/Ne laser with a wavelength of 514 nm. Transmission electron microscopy (TEM) images were obtained on a JEOL 2000EX electronic microscope operating at 120 kV. High-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) images were collected on a JEM-2100F instrument equipped with STEM dark-field (DF) detector operating at 200 kV. The specific surface area was measured on Autosorb-1 system (Quantachrome, USA) by N$_2$ adsorption isotherm through BET method. The X-ray photoelectron spectroscopy (XPS) measurements were performed using an Escalab 250 Xi X-ray photoelectron spectrometer (Thermo Scientific) with non monochromatic Al K$_\alpha$ radiation (photon energy, 1486.6 eV). Due to the overlapped signal of C 1s and Ru 3d, all the samples were mixed with a certain amount of silicon that is used for calibration (Si 2p at 98.4 eV). The UV-Vis absorption spectra were measured on a Shimadzu UV 2600 UV/Vis spectrophotometer. The Ru loadings of the catalysts were determined by inductively coupled plasma spectrometry (ICP-OES, optima 7300DV, Perkin-Elmer, USA).

**Hydrogenation of N-methylpyrrole**

5% Ru/TiO$_2$ and 5% Ru/TiO$_2$-$x$ catalysts were employed for catalytic hydrogenation of N-methylpyrrole (C$_5$H$_7$N), which is a commercial chemical with large annual global production. The reactions were carried out in the autoclave reactor (PARR®5500 series compact reactor). 100 mg catalyst (Ru/TiO$_2$ or Ru/TiO$_2$-$x$), 425 µL N-methylpyrrole and 30 mL THF (as the solvent) were put into the autoclave filled with Ar. The temperature programmer began to heat the reactor with stirring speed of 500 rpm. When the temperature of the reactor was stable at the value we set, 30 atm hydrogen was filled into the
The results and discussion of the experiment show the following:

**Results and discussion**

A highly efficient and economic viable reducing agent is needed to synthesize scalable TiO$_2$-$x$ under mild condition. Previous study reported that TiO$_2$ could be reduced by Al vapor at high temperatures,$^{21}$ showing the strong reducing potential of metals. To perform the reduction of TiO$_2$ at room temperature, a more powerful reductant and a better contact between TiO$_2$ and reductant are needed. Recently, Na metal dispersed in inert medium (for instance NaCl) has received attention for its use as a very strong reducing agent in the synthesis of NaB$_3$H$_6$.$^{21}$ The well dispersed Na metal in Na/NaCl may function as a promising reductant for the reduction of TiO$_2$. Thus, we chose a commercially available crystalline TiO$_2$ (P25, Degussa), consisting of mixed phases of anatase and rutile, as precursor for the synthesis of TiO$_2$-$x$. Ball milling of P25 with Na/NaCl fine powders at room temperature and different mechano-chemical conditions led to reduction of P25 into TiO$_2$-$x$ of different oxygen vacancies content as evidenced by the visible color changes after reduction reaction (Fig. 1). The degree of reduction can be facilely controlled by varying the reaction conditions, *i.e.*, weight ratio of P25 nanocrystals to Na/NaCl fine powders, milling rate and reaction time. The synthesized TiO$_2$-$x$ samples are marked as TiO-$n$-$t$-$v$, where $n$, $t$, and $v$ stand for weight ratio between Na/NaCl and TiO$_2$, ball milling rate and reaction time, respectively. The color of post-reduced P25 samples ranges from white to light blue and finally to dark blue with the increase of ball milling speed, reaction time and weight ratio of Na/NaCl and P25 (Scheme 1). The corresponding UV-Vis diffuse reflectance spectra clearly show the intensity of the absorption in the visible light region (400–800 nm) gradually increases with the increases of reduction degree, which is consistent with the color changes of the TiO$_2$-$x$ samples (Fig. S1†).

TEM images (Fig. 2) were collected to show the morphology and particle size of the P25 and TiO-4-180-4. The particle size of TiO-4-180-4 sample is similar to that of P25 nanocrystals (20–30 nm), which indicates that such a solid reduction treatment has little influence on the particle size of the reduced TiO$_2$ samples. Different from that of the pristine P25, disordered layer can be observed on the surface of TiO-4-180-4 particles, which is probably resulted from ball-milling treatment and/or surface reaction between TiO$_2$ and Na. As a consequence, a remarkable increase of BET surface area was obtained, *i.e.*, the TiO-4-180-4 (113 m$^2$ g$^{-1}$) has a BET surface area that is *ca.* 2.5 times of the pristine P25 (45 m$^2$ g$^{-1}$).

The crystal structure of TiO$_2$-$x$ is characterized by XRD and compared with that of P25 nanocrystals. As shown in Fig. S2,† a mild reduction treatment with ball milling speed of 80 rpm has little influence on the crystallinity of TiO$_2$-$x$. However, obvious changes occurred under harsh solid ball milling conditions (longer ball milling time with high Na/TiO$_2$ ratio or at a high ball milling speed, Fig. 3). Compared with that of the pristine TiO$_2$, there is no obvious change in the intensities of diffraction peaks corresponding to the rutile phase in the TiO$_2$-$x$ samples treated with different ball milling speeds. However, the intensities of peaks corresponding to the anatase phase significantly weaken with the increase of ball milling speed. The maintenance of rutile phase in the synthesized TiO$_2$-$x$ may be due to its chemical stability compared with that of anatase.

As shown in Fig. 4, P25 nanocrystals display the typical anatase Raman active modes with frequencies at 144, 197, 399, 515, 519 (superimposed with the 515 cm$^{-1}$ band), and 639 cm$^{-1}$ together with modes at 447 and 612 (should peak) cm$^{-1}$ corresponding to the rutile phase.$^{24}$ The relatively low intensity of rutile mode may originate from the low Raman response and low content of rutile in the P25.$^{25}$ For TiO$_2$-$x$ samples, intensities of peaks corresponding to the rutile phase did not show any obvious change. However, intensities of the peaks corresponding to the anatase phase decreased obviously (Fig. 4e). More

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**Fig. 1** Photographs of P25 nanocrystals and TiO$_2$-$x$. (a) P25 nanocrystals; (b) TiO-1-80-0.5; (c) TiO-1-80-1; (d) TiO-1-120-4; (e) TiO-1-150-4; (f) TiO-1-180-4; (g) TiO-2-180-4; (h) TiO-3-180-4 and (i) TiO-4-180-4.

**Fig. 2** TEM images of (a) P25 nanocrystals and (b) TiO-4-180-4.
Fig. 3 XRD patterns of P25 nanocrystals and TiO$_2$–x (a) P25 nanocrystals; (b) TiO–1–120–4; (c) TiO–1–150–4; (d) TiO–1–180–4; (e) TiO–2–180–4; (f) TiO–3–180–4 and (g) TiO–4–180–4.

Fig. 4 Raman spectra of P25 nanocrystals and TiO$_2$–x (a) P25 nanocrystals; (b) TiO–1–120–4; (c) TiO–1–150–4; (d) TiO–1–180–4; (e) TiO–4–180–4. A: anatase, R: rutile.

importantly, a blue shift of the strongest $E_g$ mode at 144 cm$^{-1}$ for anatase can be detected. As reported in previous research, this peak shift was mainly caused by the stoichiometry defects in TiO$_2$.

Fig. 5 (a) Ti 2p and (b) O 1s XPS of P25 nanocrystals and TiO–4–180–4.
Ru/TiO₂ has not been well investigated. As shown in Fig. 7a, about 80% of 363 molar ratio is 1 : 100. (b) The Arrhenius plots in the temperature range of 363–393 K.

Fig. 6 (a) TEM image of 5 wt% Ru/TiO₂. (b) HAADF-STEM image of 5 wt% Ru/TiO₂.

Fig. 7 (a) Hydrogenation of N-methylpyrrole with catalysts Ru/TiO₂ or Ru/TiO₂-x at 30 atm H₂ pressure, 90 °C and 100 °C, catalyst/substrate molar ratio is 1 : 100. (b) The Arrhenius plots in the temperature range of 363–393 K.

3d₅/₂ signal shows that the binding energy of Ru supported on TiO₂-x down shifts 0.5 eV as compared with that on TiO₂ (Fig. S6†), indicating an electron rich state of Ru on TiO₂-x, which also suggests a stronger interaction between support and metal particles. Although Ru is known to catalyze the hydrogenation of N-methylpyrrole well, the effect of catalyst support has not been well investigated. As shown in Fig. 7a, about 80% and 88% N-methylpyrrole can be hydrogenated to N-methylpyrrolidine in 60 min at 90 °C and 100 °C by using Ru/TiO₂ catalyst, respectively. Ru/TiO₂-x, on the other hand, shows superior catalytic performance compared to that of Ru/TiO₂, i.e., about 91% and 95% N-methylpyrrole can be hydrogenated under the same condition. Furthermore, the initial hydrogenation rate on Ru/TiO₂ at 100 °C is about twofold as that with Ru/TiO₂ (Table S1†), showing encouraging promotion effect of TiO₂-x as support. Calculated from the Arrhenius plots (Fig. 7b), the activation energies for the hydrogenation reaction are 50.9 kJ mol⁻¹ and 50.0 kJ mol⁻¹ for Ru/TiO₂ and Ru/TiO₂-x, respectively, suggesting similar hydrogenation mechanism for both catalysts. Therefore, the improved catalytic activity of Ru/TiO₂-x can be probably attributed to the better dispersion of Ru on the support, which is confirmed by the larger pre-exponential factor (A) as shown in Table S1.† We suggest that in addition to the larger specific surface area of TiO₂-x (113 m² g⁻¹), oxygen vacancies or Ti³⁺ in TiO₂-x may intensify the interaction between Ru particles and TiO₂-x, and therefore, enhancing the better dispersion of Ru particles by inhibiting Ru migration and agglomeration. However, the strong interaction between TiO₂-x and metal particles is an interesting subject that needs to be further investigated and elucidated over reactions that are sensitive to the electronic state of transition metals.

Conclusions

In summary, we have developed a room temperature solid reduction approach for the synthesis of nanosized TiO₂-x from TiO₂ crystals. A series of TiO₂-x with controllable reduction degree have been successfully synthesized by ball-milling of TiO₂ crystal with finely dispersed Na/NaCl powders. The obtained TiO₂-x with high surface area can be employed as an effective support for Ru particles and the Ru/TiO₂-x catalyst exhibited superior activity in catalytic hydrogenation of N-methylpyrrole, a commercial available heterocycle with large annual global production. We believe that this highly efficient room temperature reduction approach for the production of TiO₂-x offers a promising opportunity for the practical applications of TiO₂-x in different areas.

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Notes and references