Catalytic decomposition of N₂O over Rh/Zn–Al₂O₃ catalysts†

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Zn–Al₂O₃ supports were prepared by impregnating commercial γ-Al₂O₃ powders with different amounts of Zn(NO₃)₂, followed by calcination in air at 500 or 800 °C. Rh/Zn–Al₂O₃ catalysts were then prepared by impregnating Zn–Al₂O₃ supports with Rh(NO₃)₃ followed by calcination in air at 500 °C. The catalysts and/or supports were characterized by ICP-OES, XRD, N₂ adsorption, Raman spectroscopy, TEM-EDX, XPS, CO₂-TPD, H₂-TPR, and O₂-TPD, and the catalytic performance of supported Rh catalysts in N₂O decomposition was tested. It is concluded that the support can be described as ZnO/Al₂O₃ (ZnO supported on Al₂O₃) when calcining Zn(NO₃)₂/Al₂O₃ at 500 °C, whereas ZnAl₂O₄ spinel forms on the Al₂O₃ surface at 800 °C. Rh/Zn–Al₂O₃ catalysts are much more active than Rh/Al₂O₃ and Rh/ZnO. The best catalyst (Rh/Zn–Al₂O₃-800 with 1 wt% Rh and 1 wt% Zn) has the smallest Rh₂O₃ particle size and can desorb O₂ at lower temperature than other catalysts. Both factors may be important for achieving high activity in N₂O decomposition.

Introduction

The negative environmental impact of N₂O on global warming and ozone layer depletion has raised much concern. The concentration of N₂O in the atmosphere has been increasing at an annual rate of 0.2–0.3% since the industrial revolution.¹ Anthropogenic N₂O emission comes from several chemical processes (e.g., nitric acid and adipic acid production) and fossil fuel/biomass burning. N₂O can be decomposed into N₂ and O₂ over supported noble metal catalysts, bare or supported metal oxides, or zeolite-based catalysts.²–⁵ Rh-Based catalysts often exhibit high activity at relatively low reaction temperatures. Rh species can exist in the form of metallic Rh or RhOₓ, depending on whether the catalysts are reduced or not. Typical supports for loading Rh species include bare metal oxides,⁶–¹⁴ mixed/composite metal oxides,¹⁵–¹⁸ zeolites,¹⁹–²⁳ and metal phosphates/hydroxyapatite.²⁴–²⁷ Al₂O₃ is both a support and a catalyst widely used in industry. Rh/Al₂O₃ shows moderate catalytic activity in N₂O decomposition.³⁰–³⁳ Attempts have been made to improve Al₂O₃-based Rh catalysts for N₂O decomposition. For example, Haber et al. reported that the presence of alkali metal additives on Rh/Al₂O₃ can lead to more active catalysts due to the improved dispersion of Rh species.²⁸ Parres-Esclapez et al. found that Sr can promote the activity of Rh/Al₂O₃ due to the improved dispersion and reducibility of Rh species.²⁹ Zhao et al. reported that Rh/SiO₂–Al₂O₃ shows high activity, because oxygen desorption property is improved and Rh⁰ species is stabilized.³⁰ Kim and co-workers reported that Rh/Ce–Al₂O₃ is more active than Rh/Al₂O₃ due to increased surface area and improved reducibility of Rh species.³¹

It has been reported that spinel phase has a strong interaction with noble metals, resulting in smaller size and better stability of supported noble metal particles,³²,³³ which is beneficial for N₂O decomposition.³²,³³ High temperature calcination of Al₂O₃-supported metal nitrates is a convenient method to form spinel phase.³⁴–³⁶ However, to the best of our knowledge, there has been no work dealing with N₂O decomposition over Rh/M–Al₂O₃ catalysts with spinel phase.

Herein, we prepared Zn–Al₂O₃ supports by impregnating Zn(NO₃)₂ on commercial γ-Al₂O₃ powders followed by calcination. Rh/Zn–Al₂O₃ catalysts were synthesized by impregnating Rh(NO₃)₃ onto Zn–Al₂O₃ supports. The catalyst prepared under optimal conditions was found to be much more active than Rh/Al₂O₃ and Rh/ZnO in N₂O decomposition. The catalysts were characterized in detail, and reasons for the high activity of Rh/Zn–Al₂O₃ were elucidated.

Experimental section

Preparation

A calculated amount of Zn(NO₃)₂·6H₂O was dissolved in 20 mL deionized water in an agate mortar. Then, 1.98 g γ-Al₂O₃ powder (specific surface area = 110 m² g⁻¹) was added and the slurry
was mixed sufficiently using an agate pestle, followed by drying under an infrared lamp. The \( \text{Zn(NO}_3\text{)}_2/\text{Al}_2\text{O}_3 \) precursors were calcined in a muffle furnace at a certain temperature (500, 600, 700, 800, or 900 °C) for 4 h under flowing air. The obtained supports are referred to as x% Zn–Al\(_2\)O\(_3\)-y, where x% represents the wt% Zn in the catalysts and y the calcination temperature in °C. Zn–Al\(_2\)O\(_3\)-y in the text usually has 1 wt% Zn, unless otherwise specified.

For comparison, commercial \( \gamma\)–Al\(_2\)O\(_3\) was also calcined at 800 °C for 4 h under flowing air. ZnO was prepared by precipitation. 50 mL ammonia solution (2.8 mol L\(^{-1}\)) was added dropwise to 100 mL Zn(NO\(_3\))\(_2\) solution (0.6 mol L\(^{-1}\)) under stirring. The precipitates were isolated by filtration, thoroughly washed by deionized water till pH value of filter liquor reached 7, dried at 100 °C overnight, and calcined at 800 °C for 4 h under flowing air. These two samples are referred to as Al\(_2\)O\(_3\)-800 and ZnO-800, respectively. An additional ZnO sample was prepared by calcining Zn(NO\(_3\))\(_2\)-6H\(_2\)O at 800 °C for 4 h under flowing air.

Rhodium was loaded onto supports by impregnation. 1.98 g \( \text{Rh(NO}_3\text{)}_3\) solution (2 mg mL\(^{-1}\)) in an agate mortar and dried under an infrared lamp (theoretical Rh content is 1 wt%). The obtained powders were calcined at 500 °C overnight, and calcined at 800 °C for 4 h under flowing air. These four catalysts are determined by ICP-OES as 1.1 wt%, 1.1 wt%, respectively, in accordance with the theoretical value.

### Characterization

XRD patterns were recorded on a MSAL XD2 X-ray diffractometer using CuK\(_\alpha\) radiation at a scanning speed of 4° min\(^{-1}\) with voltage of 40 kV and current of 30 mA. BET surface areas were measured on a Micromeritics Tristar 3000 instrument. The samples were treated at 300 °C in vacuum for 3 h, followed by N\(_2\) adsorption at −196 °C. ICP-OES was measured on a PerkinElmer OPTIMA 2100 DV optical emission spectrometer. 0.1 g sample was dissolved in a mixture of 3 mL HNO\(_3\), 9 mL HCl, 1 mL HClO\(_4\), 0.5 mL H\(_2\)O\(_2\), and 3 mL HF reagents, followed by heating at 150 °C for 2–3 h. After that, the sample was mixed with 1 mL HNO\(_3\), 3 mL HCl, 0.5 mL HClO\(_4\), and 1 mL HF again, and the mixture was transferred to Teflon autoclave, heated at 180 °C for 4 h, cooled to the ambient temperature, and then diluted with distilled water for analysis.

Raman spectra were recorded on a HORIBA Jobin Yvon XPolRA spectrometer. In order to avoid fluorescence of ZnO, wavelength of exciting light was selected as 532 nm. TEM data were obtained by an FEI Tecnai G\(^2\) F20 S-TWIN with an EDX instrument. All samples were measured under an accelerating voltage of 200 kV. XPS spectra were recorded on a Shimadzu/Kratos AXIS Ultra DLD spectrometer with AlK\(_\alpha\) radiation as the excitation source. The C\(_1\)s line (284.6 eV) was used as the reference to calibrate the binding energy.

Rh dispersion was determined by CO chemisorption on a Micromeritics AutoChem II instrument, based on 1 : 1 stoichiometry (CO/Rh). 0.15 g Rh catalyst (40–60 mesh) was pretreated in He flow at 400 °C for 1 h, and cooled to 30 °C. Then, a pulse of 5% CO–He mixture was repeatedly injected into the reactor via a six-way valve until the CO signals from the thermal conductivity detector remained constant. He was used as the carrier gas. The volume of CO chemisorbed was determined by summing the fractions of CO consumed in each pulse.

CO\(_2\)-TPD experiments were conducted on a Micromeritics AutoChem II instrument. 0.2 g sample (40–60 mesh) was pretreated in He flow at 500 °C for 1 h, and cooled to 80 °C. The flow was switched to 5% CO\(_2/\text{He}\) (30 mL min\(^{-1}\)) and kept for 1 h, and then swept by \( \text{He}\) (30 mL min\(^{-1}\)) for 1.5 h. Finally, the sample was heated in \( \text{He}\) (30 mL min\(^{-1}\)) to 600 °C at a rate of 10 °C min\(^{-1}\).

H\(_2\)-TPR experiments were conducted on a FINESORB-3010 instrument, equipped with a mass spectrometer (Omnistar\(^\text{TM}\)). 0.15 g catalyst (40–60 mesh) was pretreated in He flow at 400 °C for 1 h, cooled down to room temperature, and further cooled in an ice-water bath. Then, the catalyst was exposed in 4% H\(_2\)-He mixture (30 mL min\(^{-1}\)) at 0 °C for 1 h, and heated by a furnace to 600 °C at a rate of 10 °C min\(^{-1}\). The profile was recorded through the channel of \(m/z = 2\).

O\(_2\)-TPD experiments were conducted on a FINESORB-3010 instrument. 0.2 g catalyst (40–60 mesh) was pretreated in He flow at 500 °C for 1 h, and cooled to 50 °C. The catalyst was exposed to O\(_2\) (10 mL min\(^{-1}\)) at 50 °C for 1 h, swept by \( \text{He}\) (30 mL min\(^{-1}\)) for 3 h, and heated in \( \text{He}\) (30 mL min\(^{-1}\)) to 580 °C at a rate of 10 °C min\(^{-1}\).

### Catalytic tests

Catalytic activity of N\(_2\)O decomposition was tested in a fixed-bed flow microreactor. 0.5 g catalyst (40–60 mesh) was packed in a U-shaped glass tube (7 mm inner diameter) sealed by quartz wool. A gas stream of 0.5% N\(_2\)O (balanced by He) flowed through the catalyst at a rate of 60 mL min\(^{-1}\). The catalyst (first kept at near room temperature) was exposed to the gas stream for 1 h during which the existing stream was periodically analyzed by a gas chromatograph (GC, Agilent 7890A) equipped with columns (Sepupa 6ftQ and Sepupa Q&5A, column temperature: 80 °C) that can separate N\(_2\)O, O\(_2\), and N\(_2\). The reaction temperature was then raised using a furnace and kept at various elevated temperature for 0.5 h in each temperature step. The exhaust was again periodically analyzed by the GC, and the conversion of N\(_2\)O was calculated according to \(X = ([\text{N}_2\text{O}]_\text{in} - [\text{N}_2\text{O}]_\text{out})/([\text{N}_2\text{O}]_\text{in})\), where [N\(_2\)O]\(_\text{in}\) refers to the N\(_2\)O concentration or peak area at room temperature, and [N\(_2\)O]\(_\text{out}\) refers to the N\(_2\)O concentration or peak area at an elevated temperature.

### Results and discussion

#### Structural and physical properties

Rh/Zn–Al\(_2\)O\(_3\)-500 and Rh/Zn–Al\(_2\)O\(_3\)-800 catalysts were prepared by using different supports prepared by calcining Zn(NO\(_3\))\(_2\)/Al\(_2\)O\(_3\) in air at 500 or 800 °C. The Zn and Rh contents are both fixed to be 1 wt%. In addition, Rh/Al\(_2\)O\(_3\)-800 and Rh/ZnO-800 were also prepared for comparison. The Rh contents of these four catalysts are determined by ICP-OES as 1.1 wt%, 1.1 wt%, 1.0 wt%, and 1.0 wt%, respectively. The Zn contents of Rh/Zn–Al\(_2\)O\(_3\)-500 and Rh/Zn–Al\(_2\)O\(_3\)-800 are determined as 1.1 wt% and 1.1 wt%, respectively, in accordance with the theoretical value.

Fig. 1 shows the XRD patterns of supported Rh catalysts. Rh/\(\gamma\)–Al\(_2\)O\(_3\)-500, Rh/\(\gamma\)–Al\(_2\)O\(_3\)-800, and Rh/Al\(_2\)O\(_3\)-800 exhibit identical diffraction peaks at ca. 33, 37, 40, 46, 60, and 67°, in accordance with the diffraction peaks of \(\gamma\)–Al\(_2\)O\(_3\) (PDF#47-1308).
In addition, no Rh or RhO3 can be detected, due to the low Rh content (ca. 1 wt%) and high dispersion, as demonstrated by TEM and Rh dispersion data later. Rh/ZnO-800 shows very sharp peaks corresponding to ZnO, implying that the ZnO support is highly crystalline and has a low surface area. Again, no Rh or RhO3 can be seen on the XRD pattern of Rh/ZnO-800. Fig. S1 in the ESI† depicts the XRD patterns in the 2θ = 30–40° region, highlighting that there is no difference among the XRD patterns of Rh/Zn-Al2O3-500, Rh/Zn-Al2O3-800, and Rh/Al2O3-800. No ZnO or ZnAl2O4 can be observed for Rh/Zn-Al2O3-500 and Rh/Zn-Al2O3-800, due to the low loading of Zn (ca. 1 wt%). According to the literature,34,35 surface spinel (ZnAl2O4) may be formed when calcining ZnO/Al2O3 at 800 °C. The fact that no ZnAl2O4 peaks are observed for Rh/Zn-Al2O3-800 may be because the Zn loading is so low (ca. 1 wt%). To prove this explanation, we additionally prepared Zn-Al2O3 supports with nominal Zn contents of 0.5 wt%, 5 wt%, 10 wt%, and 20 wt%, respectively. These supports were all calcined at 800 °C. All catalysts were accurately scanned in the 2θ = 30–40° range, where the XRD peaks of γ-Al2O3, ZnAl2O4, and ZnO can be distinguished clearly.

As shown in Fig. 2, the diffraction peaks are significantly enhanced with the increase of Zn content, and gradually shift toward 31.2 and 36.8° which represent (220) and (311) planes of ZnAl2O4, respectively. When the Zn content is or exceeds 10 wt%, the ZnAl2O4 peaks become clearer. When the Zn content is 20 wt%, the appearance of peaks at 31.8, 34.4, and 36.3° representing (100), (002) and (101) planes of ZnO respectively proves the formation of ZnO. The data infer that surface ZnAl2O4 may likely form on Zn-Al2O3-800 with a Zn content of 1 wt%, only that the ZnAl2O4 content is very low so the ZnAl2O4 phase can not be detected by XRD.

Fig. 3 presents the Raman spectra of several samples without Rh. ZnAl2O4 (synthesized via citric acid combustion method,37 pure phase proved by XRD) exhibits two Raman peaks at 418 and 659 cm⁻¹, representing E2 and T2g vibration modes of spinel structure, respectively.38 For Zn-Al2O3-800, the peak at 418 cm⁻¹ is obvious, indicating the formation of ZnAl2O4. However, the T2g peak at 659 cm⁻¹ is missing, in accordance with a previous report.39 A possible explanation is that tetrahedron defects are induced into spinel structure by impregnating divalent metal cations onto γ-Al2O3 (a spinel-like structure with stoichiometric ratio of tetrahedron defects).36 For comparison, Zn-Al2O3-500 does not exhibit peaks at 418 or 659 cm⁻¹. Instead, a board and weak peak at 440 cm⁻¹ is observed, close to E2 vibration mode of ZnO at 437 cm⁻¹.39 Therefore, high temperature (800 °C) is necessary for the formation of surface ZnAl2O4.
Fig. S2† shows the Raman spectra of Zn–Al₂O₃-800 with different Zn content. The peaks at 418 and 659 cm⁻¹ representing E₈ and T₂₈ vibration of ZnAl₂O₄ lattice become stronger as the Zn content in the sample increases from 1 wt% to 5 wt% and then to 10 wt%. The relative intensity of T₂₈ vibration to that of E₈ vibration also increases with the Zn content, probably because higher crystallinity of ZnAl₂O₄ restrains the formation of tetrahedron defects. However, when the Zn content of sample is 20 wt%, a new peak at 438 cm⁻¹ representing E₂ vibration of ZnO lattice appears, indicating the formation of ZnO, as also revealed by XRD (Fig. 2).

Fig. 4 shows the TEM graphs of supported Rh catalysts. The size and morphology of Zn–Al₂O₃ supports (Fig. 4a and b) are identical to those of Al₂O₃ (Fig. 4c). This conclusion is in line with the XRD patterns (Fig. 1) and specific surface area data (Table 1). The surface area of Al₂O₃-800 is 110 m² g⁻¹, and those of Zn–Al₂O₃-500 and Zn–Al₂O₃-800 are 108 and 105 m² g⁻¹, respectively. For comparison, ZnO-800 exhibits as much bigger spherical particles, in accordance with its low surface area (4 m² g⁻¹, Table 1). Rh₂O₃ particles are dispersed on these supports. The homogeneous distribution of Rh and Zn on Rh/Zn–Al₂O₃-800 is shown by EDX-mapping (Fig. S3†).

Fig. S4† shows the TEM graphs with larger size. The graphs show that Rh₂O₃ are well dispersed on the supports. The size distributions of Rh₂O₃ particles were determined by analyzing 300 particles from over 5 TEM graphs for each sample. As shown in Fig. S5,† most Rh₂O₃ particles are in the range of 0.5 and 1.0 nm over Al₂O₃-based catalysts. However, due to the low surface area of ZnO support (4 m² g⁻¹), Rh₂O₃ particles on ZnO are relatively big. The average sizes of Rh₂O₃ particles on Zn–

Table 1: Some structural and physical properties of Rh/Zn–Al₂O₃ and reference samples

<table>
<thead>
<tr>
<th>Sample</th>
<th>BET surface area a (m² g⁻¹)</th>
<th>Average size of Rh₂O₃ particles b (nm)</th>
<th>Rh dispersion (%)</th>
<th>Amount of basic sites a (µmol g⁻¹)</th>
<th>Starting desorption temperature of O₂ (°C)</th>
<th>Amount of desorbed O₂ (a.u.)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Rh/Zn–Al₂O₃-800</td>
<td>105</td>
<td>0.70 ± 0.20</td>
<td>71</td>
<td>14.6</td>
<td>234</td>
<td>496</td>
</tr>
<tr>
<td>Rh/Zn–Al₂O₃-500</td>
<td>108</td>
<td>0.78 ± 0.19</td>
<td>65</td>
<td>13.5</td>
<td>293</td>
<td>324</td>
</tr>
<tr>
<td>Rh/Al₂O₃-800</td>
<td>110</td>
<td>0.87 ± 0.25</td>
<td>59</td>
<td>14.3</td>
<td>313</td>
<td>286</td>
</tr>
<tr>
<td>Rh/ZnO-800</td>
<td>4</td>
<td>2.15 ± 0.90</td>
<td>7</td>
<td>4.4</td>
<td>—</td>
<td>0</td>
</tr>
</tbody>
</table>

a Bare supports (without Rh). b Obtained by analyzing 300 Rh₂O₃ particles in TEM graphs.
Fig. 5 Rh 3d XPS spectra of Rh/Zn-\(\text{Al}_2\text{O}_3\)-800, Rh/Zn-\(\text{Al}_2\text{O}_3\)-500, Rh/\(\text{Al}_2\text{O}_3\)-800, and Rh/ZnO-800.

Activity measurement

Fig. 6 shows the \(\text{N}_2\text{O}\) conversions on supported Rh catalysts as a function of reaction temperature. The catalytic activity on these catalysts follows the sequence of Rh/Zn-\(\text{Al}_2\text{O}_3\)-800 > Rh/Zn-\(\text{Al}_2\text{O}_3\)-500 > Rh/\(\text{Al}_2\text{O}_3\)-800 > Rh/ZnO-800. The \(\text{N}_2\text{O}\) conversions over these catalysts at 275 °C are 98.7%, 54.4%, 27.2%, and 22.3%, respectively. The specific rates of these catalysts (expressed as moles of \(\text{N}_2\text{O}\) converted per mole of Rh per minute) at 275 °C are calculated to be 0.226, 0.125, 0.069, and 0.056 min\(^{-1}\), respectively. The \(T_{50}\) (temperature required for 50% conversion) values of these catalysts are 251, 273, 289, and 300 °C, respectively. Note that ZnO referred to above was prepared by precipitation. An additional Rh/ZnO-800 catalyst was prepared using ZnO obtained by calcining \(\text{Zn(NO}_3\text{)_2}\cdot 6\text{H}_2\text{O}\) at 800 °C for 4 h. That catalyst is less active than Rh/ZnO-800 mentioned above (Fig. S7†). In addition, \(\text{Zn-Al}_2\text{O}_3\)-800 (without Rh) is much less active (17.0% \(\text{N}_2\text{O}\) conversion at 400 °C) than Rh/Zn-\(\text{Al}_2\text{O}_3\), indicating that \(\text{Rh}_2\text{O}_3\) act as the main active sites for \(\text{N}_2\text{O}\) decomposition.

The effect of calcination temperature of \(\text{Zn-Al}_2\text{O}_3\) on the activity of the resulting Rh/Zn-\(\text{Al}_2\text{O}_3\) catalysts was studied. As shown in Fig. S8,† the \(\text{N}_2\text{O}\) conversions over Rh/Zn-\(\text{Al}_2\text{O}_3\)-500, Rh/Zn-\(\text{Al}_2\text{O}_3\)-600, Rh/Zn-\(\text{Al}_2\text{O}_3\)-700, Rh/Zn-\(\text{Al}_2\text{O}_3\)-800, and Rh/\(\text{Zn-Al}_2\text{O}_3\)-900 at 275 °C are 54.4%, 55.5%, 68.0%, 98.7%, and 47.1%, respectively. The \(T_{50}\) values of these catalysts are 273, 272, 266, 251, and 277 °C, respectively. The activities of these catalysts follow the sequence of Rh/Zn-\(\text{Al}_2\text{O}_3\)-900 ~ Rh/Zn-\(\text{Al}_2\text{O}_3\)-600 < Rh/Zn-\(\text{Al}_2\text{O}_3\)-500 < Rh/Zn-\(\text{Al}_2\text{O}_3\)-800 < Rh/Zn-\(\text{Al}_2\text{O}_3\)-700, i.e., Rh/Zn-\(\text{Al}_2\text{O}_3\)-800 is the most active.

The effect of Zn content on the performance of Rh/Zn-\(\text{Al}_2\text{O}_3\) catalysts was also studied. As shown in Fig. S9,† the \(\text{N}_2\text{O}\) conversions over Rh/\(\text{Al}_2\text{O}_3\)-800, Rh/0.5% \(\text{Zn-Al}_2\text{O}_3\), Rh/1% \(\text{Zn-Al}_2\text{O}_3\), Rh/10% \(\text{Zn-Al}_2\text{O}_3\), and Rh/20% \(\text{Zn-Al}_2\text{O}_3\) at 275 °C are 27.2%, 89.7%, 98.7%, and 98.7%, respectively. The \(T_{50}\) values of these catalysts are 289, 258, 251, 261, and 265 °C, respectively. The activities of these catalysts follow the sequence of Rh/\(\text{Al}_2\text{O}_3\)-800 < Rh/0.5% \(\text{Zn-Al}_2\text{O}_3\) < Rh/1% \(\text{Zn-Al}_2\text{O}_3\) < Rh/10% \(\text{Zn-Al}_2\text{O}_3\) < Rh/20% \(\text{Zn-Al}_2\text{O}_3\), i.e., Rh/1% \(\text{Zn-Al}_2\text{O}_3\)-800 is the most active.

Fig. 6 Conversion of \(\text{N}_2\text{O}\) on the catalysts as a function of reaction temperature.
Fig. S10† shows the influence of different GHSV values (7420, 14,840, and 37,100 h⁻¹), achieved by adjusting the catalyst dosage (0.5, 0.25, and 0.1 g, respectively), on the performance of the optimal catalyst Rh/1% Zn–Al₂O₃-800. In general, a decrease of N₂O conversion is seen at a high GHSV at the same reaction temperature. The T₅₀ values are 251, 266, and 291 °C at GHSV of 7420, 14,840, and 37,100 h⁻¹, respectively. However, the N₂O conversion still reaches 87.5% at 325 °C when the GHSV is as high as 37,100 h⁻¹.

Fig. S11† shows the influence of Rh/M–Al₂O₃-800 (M = Zn, Mg, Co, Ni, Cu) catalysts with different divalent metal cations on N₂O decomposition. The M content is fixed to be 1 wt%. All Rh/M–Al₂O₃-800 catalysts exhibit superior performance than Rh/Al₂O₃-800. The N₂O conversions on Rh/Zn–Al₂O₃-800, Rh/Mg–Al₂O₃-800, Rh/Co–Al₂O₃-800, Rh/Ni–Al₂O₃-800, Rh/Cu–Al₂O₃-800, and Rh/Al₂O₃-800 at 275 °C are 98.7%, 64.1%, 76.0%, 72.9%, 82.5%, and 27.2% respectively. Rh/Zn–Al₂O₃-800 is still the most active among these catalysts.

Fig. 7 shows N₂O conversion over Rh/Zn–Al₂O₃-800 and Rh/Al₂O₃-800 in the absence or presence of H₂O, O₂, or CO₂. According to the literature, H₂O and O₂ cause competitive adsorption with N₂O on active sites, leading to severe inhibition of activity on Rh-based catalysts. The T₅₀ value of Rh/Zn–Al₂O₃-800 in the absence of H₂O, O₂, or CO₂ is 251 °C. The T₅₀ value increases to 272 °C when 5% O₂ alone is added to the reaction mixture, and it increases more obviously to 321 °C when 2% H₂O alone is added. When 5% O₂ and 2% H₂O are co-fed to the reaction mixture, the T₅₀ value increases further to 333 °C. The addition of 2% CO₂ alone makes the T₅₀ value become 284 °C. The presence of H₂O, O₂, or CO₂ also exerts similar inhibiting effects on Rh/Al₂O₃-800, only that Rh/Al₂O₃-800 is always much less active than Rh/Zn–Al₂O₃-800 under respective condition for comparison. The stability of both Rh/Zn–Al₂O₃-800 and Rh/Al₂O₃-800 as a function of time on stream is good, either in the absence or presence of 2% H₂O and 5% O₂ (Fig. S12†).

Gas-switching experiments were conducted to know whether the inhibiting effect of O₂, and H₂O is reversible. Fig. 8 clearly shows that N₂O conversion decreases almost immediately when 5% O₂ is co-fed into the reactor, but the conversion can be restored when stopping feeding O₂. 2% H₂O has a more obviously inhibiting effect, but such an inhibiting effect is also reversible.¹¹,²⁶

Fig. 9 shows the influence of O₂ concentration on the activities of Rh/Zn–Al₂O₃-800, Rh/Zn–Al₂O₃-500, Rh/Al₂O₃-800, and Rh/ZnO-800 in N₂O decomposition.
O$_2$ reaction order was measured by changing O$_2$ concentration while keeping N$_2$O concentration as 0.5%, to further explore the inhibition effect of oxygen. To eliminate diffusion factor, catalysts usage (0.128, 0.204, 0.256, and 0.306 g for Rh/Zn-Al$_2$O$_3$-800, Rh/Zn-Al$_2$O$_3$-500, Rh/Al$_2$O$_3$-800 and Rh/ZnO-800, respectively) and reaction temperature (230, 250, 260, and 290 °C for Rh/Zn-Al$_2$O$_3$-800, Rh/Zn-Al$_2$O$_3$-500, Rh/Al$_2$O$_3$-800 and Rh/ZnO-800, respectively) were chosen to control the N$_2$O conversion below 15%. As shown in Fig. 9, a good linearity between ln(rate) versus ln(O$_2$ concentration) enables us to derive O$_2$ reaction orders as −0.352, −0.540, −0.598, and −0.707 for Rh/Zn-Al$_2$O$_3$-800, Rh/Zn-Al$_2$O$_3$-500, Rh/Al$_2$O$_3$-800, and Rh/ZnO-800, respectively. The data indicate that the inhibition effect of O$_2$ on Rh/Zn-Al$_2$O$_3$-800 is the weakest, whereas that on Rh/ZnO-800 is the strongest.

The influence of 4% H$_2$ pretreatment on Rh/Zn-Al$_2$O$_3$-800 was investigated. As shown in Fig. S13, the H$_2$ pretreatment could improve the activity. The N$_2$O conversion reaches 64.5% at 250 °C (in comparison to the 48.9% conversion achieved on the catalyst without being pretreated in H$_2$). It has been reported that metallic Rh is more favorable for N$_2$O decomposition on non-reducible support, thus the formation of some metallic Rh after H$_2$ reduction (the proportion of Rh$^0$ increases from 9.1% to 26.0% upon H$_2$ pretreatment, Fig. S6†) should be the reason of the promoted activity although Rh$_2$O$_3$ can also catalyze this reaction.

**Chemical properties of Rh/Zn-Al$_2$O$_3$ and reference samples**

Fig. S14† shows the CO$_2$-TPD profiles of Zn-Al$_2$O$_3$-800, Zn-Al$_2$O$_3$-500, Al$_2$O$_3$-800, and ZnO-800, and the values of calculated basicity are listed in Table 1. It was reported that ZnAl$_2$O$_4$ exhibits higher basic site density than Al$_2$O$_3$. However, here Zn-Al$_2$O$_3$-800, Zn-Al$_2$O$_3$-500, and Al$_2$O$_3$-800 show similar CO$_2$-TPD profiles. The amounts of basic sites of these supports are 14.6, 13.5, and 14.3 μmol g$^{-1}$, respectively. Low Zn loading should be the reason for the similarity, because only a small amount of ZnAl$_2$O$_4$ phase is formed on Al$_2$O$_3$ surface. On the other hand, ZnO-800 with a low surface area (4 m$^2$ g$^{-1}$) has the smallest amount of basic sites (4.4 μmol g$^{-1}$), in accordance with a previous report.$^{45}$

Fig. 10 shows the H$_2$-TPR profiles of Rh/Zn-Al$_2$O$_3$-800, Rh/Zn-Al$_2$O$_3$-500, Rh/Al$_2$O$_3$-800, and Rh/ZnO-800. All catalysts exhibit reduction peaks below 300 °C, assigned to the reduction of Rh species.$^{17,30,42}$ The reduction temperature follows the sequence of Rh/Zn-Al$_2$O$_3$-800 (56 °C) < Rh/Zn-Al$_2$O$_3$-500 (76 °C) < Rh/Al$_2$O$_3$ (86 °C) < Rh/ZnO-800 (217 °C). The reduction temperature of Rh/ZnO-800 is much higher than those of other catalysts, probably due to larger Rh particles (TEM, Fig. 4 and S4†) with a distinctive electronic property (XPS, Fig. 5). In addition, a peak located at 233 °C is exhibited (the result is reproducible) in Rh/Zn-Al$_2$O$_3$-800, may due to the strong interaction between the support and Rh, resulting in the formation of other Rh-containing species instead of Rh$_2$O$_3$.

Fig. 11 shows the O$_2$-TPD profiles of Rh/Zn-Al$_2$O$_3$-800, Rh/Zn-Al$_2$O$_3$-500, Rh/Al$_2$O$_3$-800, and Rh/ZnO-800. Rh/ZnO-800 does not show any oxygen desorption peak, whereas other three catalysts all show an oxygen desorption peak below 580 °C. The starting desorption temperature follows the order of Rh/Zn-Al$_2$O$_3$-800 (233 °C) < Rh/Zn-Al$_2$O$_3$-500 (298 °C) < Rh/Al$_2$O$_3$-800 (313 °C), whereas the amounts of oxygen desorption follow the sequence of Rh/Zn-Al$_2$O$_3$-800 > Rh/Zn-Al$_2$O$_3$-500 > Rh/Al$_2$O$_3$-800 (Table 1).

**Discussion**

The N$_2$O decomposition activities of four catalysts follow the sequence of Rh/Zn-Al$_2$O$_3$-800 > Rh/Zn-Al$_2$O$_3$-500 > Rh/Al$_2$O$_3$-800 > Rh/ZnO-800 (Fig. 6). The catalytic performance of the best catalyst (Rh/Zn-Al$_2$O$_3$-800) in this work is then compared with the performance of other catalysts. Table S1† lists the $T_{50}$ and $T_{90}$ values of various catalysts, together with the reaction conditions. Although the Rh/Zn-Al$_2$O$_3$-800 catalyst seems to be more active than most of the catalysts, such as Cu-, Fe-, and Co-based catalysts,$^{51-56}$ as well as Pd-, Pt,$^{8,10,57}$ and Ir-based$^{18}$
catalysts (judging from the \( T_{50} \) values), it is not appropriate to claim that because of the difference in reaction conditions. Hence, the activity of Rh/Zn-Al\(_2\)O\(_4\)-800 is compared with Rh/SiO\(_2\), Rh/Al\(_2\)O\(_3\), Rh/TiO\(_2\), and Rh/CoO\(_2\) (all with ca. 1 wt% Rh) tested under the same reaction condition in our laboratory. As shown in Table S1, the \( T_{50} \) values of Rh/Zn-Al\(_2\)O\(_4\)-800, Rh/SiO\(_2\), Rh/Al\(_2\)O\(_3\), Rh/TiO\(_2\), and Rh/CoO\(_2\) are 251, 324, 289, 310, and 223 °C, respectively, indicating the good performance of Rh/Zn-Al\(_2\)O\(_4\)-800 among these supported Rh catalysts. The activity sequence of catalysts is Rh/CoO\(_2\) > Rh/Zn-Al\(_2\)O\(_4\)-800 > Rh/Al\(_2\)O\(_3\) > Rh/TiO\(_2\) > Rh/SiO\(_2\). Rh/Zn-Al\(_2\)O\(_4\)-800 (\( T_{50} = 251 \) °C) is less active than Rh/HAP-PEG-200 (ca. 1 wt% Rh supported on hydroxyapatite nanorods synthesized through hydrothermal method, and tested under the same reaction condition) that shows a \( T_{50} \) value of 223 °C.\(^\text{27} \) However, the synthesis of Rh/Zn-Al\(_2\)O\(_4\)-800 is more convenient and of low cost, and Rh/Zn-Al\(_2\)O\(_4\)-800 shows superior activity than Rh/HAP-PEG-200 in the presence of 2% H\(_2\)O and 5% O\(_2\). \( T_{50} \) values of these two catalysts under such a condition are 333 and 344 °C, respectively.

After comparing the catalytic activity, we then attempt to discuss why Rh/Zn-Al\(_2\)O\(_4\)-800 is more active than Rh/Zn-Al\(_2\)O\(_4\)-500, Rh/Al\(_2\)O\(_3\)-800, and Rh/ZnO-800. The structural properties of \( \gamma \)-Al\(_2\)O\(_3\) are not significantly altered by the incorporation of 1 wt% Zn, as seen from the micro-morphology (TEM, Fig. 4) and specific surface areas (Table 1). The average size of Rh\(_2\)O\(_3\) particles follows the sequence of Rh/Zn-Al\(_2\)O\(_4\)-800 < Rh/Zn-Al\(_2\)O\(_4\)-500 < Rh/Al\(_2\)O\(_3\)-800 < Rh/ZnO-800 (also proved by Rh dispersion, Table 1), inversely correlating to the order of catalytic activity, in the sense that a catalyst with smaller Rh\(_2\)O\(_3\) particles has higher activity in N\(_2\)O decomposition.\(^\text{25,26,28} \) The data thus infer that the formation of surface ZnAl\(_2\)O\(_4\) spinels on Al\(_2\)O\(_3\) support may be beneficial for the high dispersion of Rh\(_2\)O\(_3\) particles.

In our ongoing research, bulk ZnAl\(_2\)O\(_4\) support is also found to be able to lead to high dispersion of Rh\(_2\)O\(_3\) species, thus promoting the catalytic activity (data not shown). There seems to be a stronger interaction between Rh\(_2\)O\(_3\) and ZnAl\(_2\)O\(_4\) phase. Similar conclusions can be found in the literature. As revealed by theoretical calculation reveals that MgAl\(_2\)O\(_4\) (spinel structure) has a stronger interaction with supported Rh and Ir particles than \( \gamma \)-Al\(_2\)O\(_3\), promoting the dispersion of Rh and Ir as well as the stability of the catalysts in methane steam reforming.\(^\text{22} \) In addition, smaller Pd nanoparticles could be immobilized on ZnAl\(_2\)O\(_4\) surface, leading to higher activity and good stability in Suzuki–Miyaura coupling reaction.\(^\text{23} \) It was also reported that Rh prefers to be located on the spinel phase of metal modified-Al\(_2\)O\(_3\) support.\(^\text{29} \) Therefore, the better Rh dispersion on Rh/Zn-Al\(_2\)O\(_4\)-800 should be caused by the formation of ZnAl\(_2\)O\(_4\) phase on Al\(_2\)O\(_3\) support, as proved by Raman (Fig. 3). XRD data also provide indirect evidence because Zn-Al\(_2\)O\(_4\)-800 supports with low Zn contents do not show ZnAl\(_2\)O\(_4\) peaks due to the low concentration of Zn, whereas Zn-Al\(_2\)O\(_4\)-800 supports with relatively high Zn contents show ZnAl\(_2\)O\(_4\) peaks clearly (Fig. 2). Another piece of evidence showing the strong interaction between Rh\(_2\)O\(_3\) and ZnAl\(_2\)O\(_4\) comes from the H\(_2\)-TPR data, as an additional reduction peak shows up at a higher temperature of 233 °C (Fig. 10).

As shown in XPS data (Fig. 5), the binding energies of Rh 3d\(_{5/2}\) follow the sequence of Rh/Zn-Al\(_2\)O\(_4\)-800 (309.3 eV) < Rh/Zn-Al\(_2\)O\(_4\)-500 (309.5 eV) < Rh/Al\(_2\)O\(_3\)-800 (309.6 eV), indicating the electronic properties of Rh species are altered by the supports. Rh/Zn-Al\(_2\)O\(_4\)-800 exhibits the lowest Rh 3d\(_{5/2}\) binding energy among these catalysts, indicating the strongest shielding effect caused by the highest outer electron density. It has been reported that metal-oxygen bonding energy is weaker in metal oxides with higher electron density, because the electron could fill into the empty orbital of oxygen.\(^\text{66,67} \) As a result, Rh\(_2\)O\(_3\) on Zn-Al\(_2\)O\(_4\)-800 is easier to be reduced due to the weaker bonding between Rh and oxygen (H\(_2\)-TPR, Fig. 10). In the literature, the higher activity of catalysts in N\(_2\)O decomposition has been correlated to the reduction behaviour of catalysts (more specifically, the active components) in some cases.\(^\text{54,55} \) That is not to say that these catalysts have to be reduced prior to the reaction testing. Rather, the enhanced reducibility means the weakening of the M–O bonds, which is expected to be important for N\(_2\)O decomposition, i.e., the adsorbed oxygen can be desorbed from the catalyst more easily.

In another aspect, oxygen desorption is a key step in N\(_2\)O decomposition, even rate-determining step in many cases.\(^\text{26,27,70,71} \) If oxygen can not be smoothly desorbed from catalysts surface, the active sites will be occupied by oxygen, blocking the catalytic circle. As shown in Fig. 9, the starting O\(_2\)-desorption temperature follows the order of Rh/Zn-Al\(_2\)O\(_4\)-800 (233 °C) < Rh/Al\(_2\)O\(_3\)-500 (298 °C) < Rh/Al\(_2\)O\(_3\)-800 (313 °C), in line with the notion that easy desorption of O\(_2\) is beneficial for catalytic N\(_2\)O decomposition.\(^\text{26,27,70,71} \) In addition, the inhibition effect of O\(_2\) (as seen from the oxygen reaction order, Fig. 8) on catalytic activity follows the sequence of Rh/Zn-Al\(_2\)O\(_4\)-800 < Rh/Zn-Al\(_2\)O\(_4\)-500 < Rh/Al\(_2\)O\(_3\)-800 < Rh/ZnO-800, inversely correlating with the activity sequence of these catalysts.

**Conclusions**

Zn(NO\(_3\))\(_2\) was impregnated onto a commercial \( \gamma \)-Al\(_2\)O\(_4\) support. The composite was then calcined in air at 500 or 800 °C. The resulting Zn-Al\(_2\)O\(_4\)-500 and Zn-Al\(_2\)O\(_4\)-800 supports were used to load Rh via impregnation with Rh(NO\(_3\))\(_3\) and calcination at 500 °C. While the Zn species in Rh/Zn-Al\(_2\)O\(_4\)-500 is ZnO, the Zn species in Rh/Zn-Al\(_2\)O\(_4\)-800 is ZnAl\(_2\)O\(_4\). The activities of four catalysts in N\(_2\)O decomposition follow the sequence of Rh/Zn-Al\(_2\)O\(_4\)-800 > Rh/Zn-Al\(_2\)O\(_4\)-500 > Rh/Al\(_2\)O\(_3\)-800 > Rh/ZnO-800, correlating with the size of Rh\(_2\)O\(_3\) particles, the reducibility of Rh\(_2\)O\(_3\), the O\(_2\)-desorption property, and the oxygen reaction order. The most active Rh/Zn-Al\(_2\)O\(_4\)-800 has the smallest Rh\(_2\)O\(_3\) particles (i.e., the highest Rh dispersion), the lowest reduction temperature of Rh\(_2\)O\(_3\), the lowest O\(_2\) desorption temperature, and its activity is influenced by the presence of O\(_2\) least obviously. The formation of ZnAl\(_2\)O\(_4\) on Al\(_2\)O\(_3\) support is beneficial for the stabilization of Rh\(_2\)O\(_3\) particles. This work demonstrates a convenient way for preparing active Rh catalysts based on commercial \( \gamma \)-Al\(_2\)O\(_4\) support.

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Notes and references