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Insight into quinoxaline containing D- π -A dyes for dye-sensitized solar cells with cobalt and iodine based electrolytes: the effect of π -bridge on the HOMO energy level and photovoltaic performance†

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Three new quinoxaline-based organic dyes (AQ201, AQ202, and AQ203), containing thiophene, 3,4-ethylenedioxythiophene (EDOT), and cyclopentadithiophene (CPDT) in the π -system, respectively, have been designed and synthesized for dye-sensitized solar cells. Different from the traditional donor- π -bridge-acceptor (D- π -A) type dyes, the dissymmetric π -bridge on both sides of quinoxaline enables great flexibility in fine-tuning the absorption spectra and energy levels of the resultant molecules. By changing the π -bridge between the bulky triphenylamine donor and quinoxaline group, a negative shift was observed regarding the highest occupied molecular orbital (HOMO) levels of AQ201, AQ202, and AQ203 dyes (0.88, 0.79, and 0.72 V vs. NHE, respectively), while the lowest unoccupied molecular orbital (LUMO) levels of these dyes remained the same (-1.19, -1.20, and -1.20 V vs. NHE, respectively), which, in turn, resulted in a gradual shift of the absorption spectra of AQ dyes. The absorption spectra properties of the dyes are also analysed by density functional theory. The calculated results in combination with the experiments indicate that the absorption bands are mainly dominated by charge transfer transitions from the HOMO and HOMO-1 orbital to the LUMO. In all cases, the [Co(bpy)₃]^{2+/3+} redox-shuttle afforded superior solar cell performance compared to I⁻/I₃⁻. More importantly, dye AQ202 shows the highest power conversion efficiency (PCE) of 8.37% with the [Co(bpy)₃]^{2+/3+} based electrolyte by maintaining a balance between the spectral absorption range and driving force for dye regeneration. Transient photocurrent decay experiments as well as electrochemical impedance spectroscopy indicate that the lower HOMO levels lead to higher electron lifetime and dye regeneration efficiency.

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Introduction

Dye-sensitized solar cells (DSSCs) have attracted considerable attention due to their relatively low production cost, transparency and flexibility.¹ It is known that the operating principle of DSSCs usually involves excitation of the dye followed by charge transfer from the dye to the conduction band of TiO₂. Thus the oxidized state of the dye is obtained and then

regenerated by receiving an electron from a redox mediator in the electrolyte.² Accordingly, the highest occupied molecular orbital (HOMO) level of the dye should be sufficiently positive compared to the potential of the redox mediator in the electrolyte for efficient dye regeneration. Therefore, dye regeneration depends strongly on both the potential of the redox mediator and the HOMO level of dyes.³ Typically, iodide/triiodide (I⁻/I₃⁻) and [Co(bpy)₃]^{2+/3+} are most commonly used as redox mediators to obtain highly efficient photovoltaic performances in liquid electrolytes, and their potentials are almost constant.⁴ According to the Marcus theory, a difference between the dye HOMO level and the redox potential of the electrolyte of 0.2–0.3 eV can spur single-electron transfer, resulting in effective dye regeneration.⁵ In that case, in order to achieve efficient dye regeneration, it is essential to modulate the HOMO level of the dye to the potential of the redox mediator.

The general design principle of organic sensitizers consists of a donor- π -acceptor (D- π -A) framework that can efficiently

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facilitate intramolecular charge transfer (ICT) from the ground to excited state and increase light absorption.⁶ In general, one can tune the HOMO level of the dye by alternating each component in the design. However, by doing that, the LUMO level of the dye also changes significantly, leading to difficulty in independently exploring the effects of the HOMO level on the photovoltaic performance.⁷ Recently, an auxiliary electron acceptor has been widely introduced into the π -bridge of the D- π -A framework.⁸ This moiety not only enhances the photovoltaic performance and photostability of sensitizers, but is also beneficial to fine-tune the molecular energy levels.⁹

It is known that quinoxaline is a strong electron-withdrawing unit due to its high electron affinity, which arises from the two symmetric unsaturated nitrogen atoms.¹⁰ Our group recently reported quinoxaline-based sensitizers YA422 (ref. 11) and AQ308,¹² which showed outstanding power conversion efficiencies (PCE) of 10.56% and 9.81%, respectively. Those reports inspired us to do some further investigation into the application of quinoxaline as the auxiliary acceptor in DSSCs. In this regard, herein, we have designed and synthesized three new quinoxaline-based organic sensitizers (AQ201, AQ202, and AQ203) by using the dissymmetric π -bridge on both sides of quinoxaline (Scheme 1). The introduction of a phenyl ring between the quinoxaline moiety and the cyanoacrylic acid group stabilizes the LUMO level of the dye.¹³ And most importantly, an electron-rich unit (thiophene, EDOT, and CPDT, respectively) close to the bulky triphenylamine donor group has been included to tune the HOMO level of the dye.¹⁴ As a result, a negative shift was observed in the HOMO levels of AQ201, AQ202, and AQ203 dyes (0.88, 0.79, and 0.72 V vs. NHE, respectively), while the LUMO levels remained the same (-1.19, -1.20, and -1.20 V vs. NHE, respectively). Hence, a red shift and broadening of the absorption spectra were achieved. Those dyes were employed in DSSCs with [Co(bpy)₃]^{2+/3+} and I⁻/I₃⁻ based electrolytes. Overall, the DSSCs using the three dyes with the cobalt-based electrolyte yielded higher PCE compared to the corresponding DSSCs with the iodide-based electrolyte, which is ascribed to the bulky donor group usually required for devices with the cobalt-based electrolyte.^{11,15} Consequently, the AQ202-based cell with an intermediate HOMO level performed best with a PCE of 8.37% using the [Co(bpy)₃]^{2+/3+} based electrolyte, while the AQ201-based cell with the lowest HOMO level showed the highest PCE

of 6.57% with the I⁻/I₃⁻ based electrolyte. Electrochemical impedance spectroscopy (EIS) and transient absorbance spectroscopy (TAS) analyses were performed to understand the influence of the HOMO energy level on the dye regeneration and recombination kinetics of DSSC devices with the two different electrolytes.

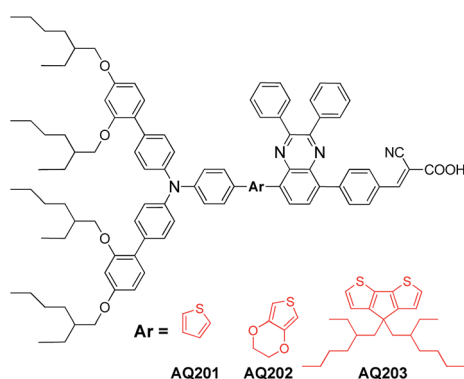
Results and discussion

Design and synthesis

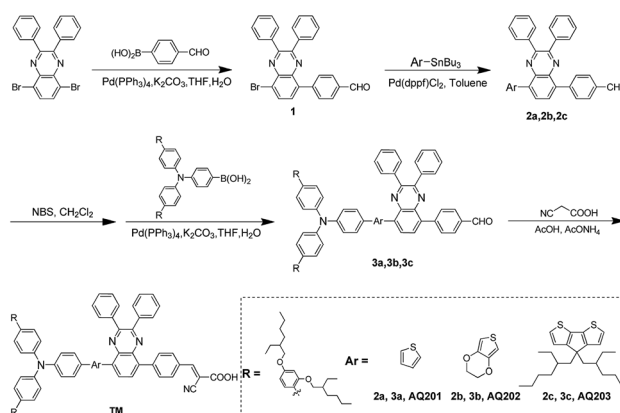
In this context, molecular engineering was focused on tuning the dye HOMO level and extending the absorption range based on the quinoxaline containing the D- π -A structure, and then improving dye regeneration and photovoltaic performances with [Co(bpy)₃]^{2+/3+} and I⁻/I₃⁻ based electrolytes, respectively. The bulky triphenylamine donor offers the possibility to inhibit charge recombination and block molecular aggregation, especially for the cobalt redox system. As shown in Scheme 2, AQ201, AQ202, and AQ203 were prepared following the same synthetic routes. The synthetic work starts from low-cost, commercially available materials such as 5,8-dibromo-2,3-diphenylquinoxaline and (4-formylphenyl)boronic acid. The three dyes were obtained following a sequence of Suzuki coupling, Stille coupling, bromination, and Knoevenagel condensation. Such a synthetic procedure for AQ201, AQ202, and AQ203 was optimized on the basis of reaction selectivity and product yields. The key intermediates and final dyes were fully characterized with ¹H NMR, ¹³C NMR, and HRMS.

Photophysical and electrochemical properties

The absorption spectra of AQ201, AQ202, and AQ203 in dichloromethane (CH₂Cl₂) solution and on 4 μ m transparent TiO₂ films are shown in Fig. 1 and their relative intensities are tabulated in Table 1. The variation of the π -bridge of the electron-donating moiety leads to significant changes in the position of the lowest energy absorption band. The intramolecular charge transfer (ICT) absorption peak in the visible region is observed at $\lambda_{\text{max}} = 491, 515, \text{ and } 545 \text{ nm}$ for AQ201, AQ202, and AQ203, respectively. As listed in Table 1, when replacing the



Scheme 1 The structures of AQ201, AQ202, and AQ203.



Scheme 2 Synthetic procedure of the dyes AQ201, AQ202, and AQ203.



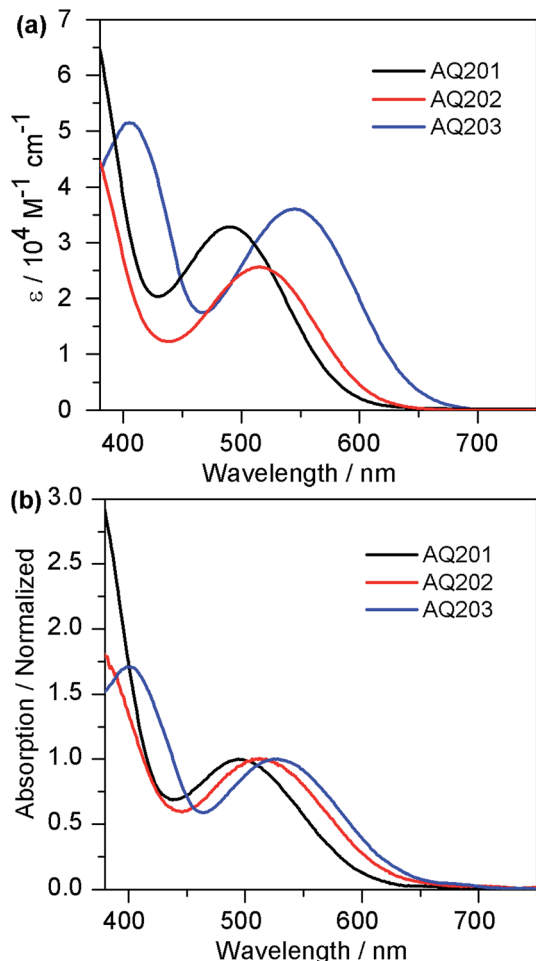


Fig. 1 (a) Absorption spectra of dyes AQ201, AQ202, and AQ203 in CH_2Cl_2 ; (b) normalized absorption spectra of dyes AQ201, AQ202, and AQ203 on $4 \mu\text{m}$ TiO_2 films.

thiophene unit with the EDOT or CPDT unit, the absorption spectra of AQ202 and AQ203 broaden with respect to AQ201 in the visible region, and the absorption peaks (λ_{max}) are red-shifted by 24 and 54 nm, respectively. The red shift in the absorption peak of dye AQ201 sensitized on the $4 \mu\text{m}$ transparent TiO_2 films compared to the absorption spectra in solution might be ascribed to the formation of J-aggregates.¹⁶ In addition, when AQ202 and AQ203 dyes are adsorbed on the TiO_2 films, the absorption peaks (λ_{max}) blue-shifted to $\lambda_{\text{max}} = 513$ and 528 nm,

respectively. This might be ascribed to deprotonation of the carboxylic acid or H-aggregates.¹⁷ In summary, this may be due to the smaller size of the thiophene group than that of EDOT and CPDT groups, leading to different aggregation behavior.

To evaluate the possibility of electron transfer from the excited dye molecule to the conduction band (CB) of TiO_2 , and from the redox couple in the electrolyte to the oxidized dye molecule, cyclic voltammograms of AQ201, AQ202, and AQ203 dyes were recorded in CH_2Cl_2 solution (Fig. 2a) to determine the redox potential and the corresponding values are presented in Table 1. The ground state oxidation potentials of AQ201, AQ202, and AQ203 are 0.88, 0.79 and 0.72 V vs. NHE, respectively, which are sufficiently higher than the redox potential of iodide/triiodide (0.45 V vs. NHE), guaranteeing ample driving force for dye regeneration. However, in the presence of $[\text{Co}(\text{bpy})_3]^{2+/3+}$ (0.57 V vs. NHE) redox couple, the AQ201 and AQ202 dyes show an advantage over the AQ203 dye for dye regeneration according to the Marcus theory.⁵ The structural differences in the π -bridge of the electron-donating part of AQ201, AQ202, and AQ203 have a significant influence on their oxidation potentials.

The ground state oxidation potentials (E_{ox}^0) of AQ202 and AQ203 are decreased by 90 and 160 mV compared to AQ201 via replacing the thiophene with EDOT or CPDT unit. On the other hand, the excited-state oxidation potentials (E_{ox}^*) of AQ201, AQ202, and AQ203, derived from the ground state oxidation potential and optical energy gap ($E_{\text{ox}}^0 - E_{0-0}$), are -1.19 , -1.20 , and -1.20 V, respectively. E_{0-0} is estimated from the absorption threshold from the absorption spectra of dyes in solution. The E_{ox}^* of these dyes are placed sufficiently above the TiO_2 conduction band edge (-0.5 V vs. NHE), ensuring no energetic barriers for electron injection from the excited dyes into the conduction band of TiO_2 . Herein, the π -bridge on the left side of the quinoxaline unit shows little effect on the LUMO level of D- π -A dyes, which is helpful in molecular designing and HOMO energy level tuning for future novel organic dyes.¹⁸

Theoretical approach

Density functional theory (DFT) calculations were carried out using the Gaussian 09 (ref. 19) program to obtain further insight into the electronic distribution of the frontier molecular orbitals and the molecular structures of the dyes AQ201, AQ202, and AQ203. The ground-state structures of those dyes were optimized by DFT at the B3LYP/6-31G(d) level.²⁰ In the calculations, the long alkyl chains were replaced by methyl groups to

Table 1 Photophysical and electrochemical properties of dyes AQ201, AQ202 and AQ203

Sample	$\lambda_{\text{max}}^a/\text{nm}$ (ϵ , $\text{M}^{-1} \text{cm}^{-1}$)	$\lambda_{\text{max}}^b/\text{nm}$ (on TiO_2)	HOMO ^c /V (vs. NHE)	E_{0-0}^d/eV	LUMO ^e /V (vs. NHE)
AQ201	491 (32 898)	494	0.88	2.07	-1.19
AQ202	515 (25 550)	513	0.79	1.99	-1.20
AQ203	545 (35 865)	528	0.72	1.92	-1.20

^a Absorption maximum in CH_2Cl_2 at room temperature. ^b Absorption maximum on $4 \mu\text{m}$ transparent TiO_2 films. ^c HOMO was estimated by the ground state oxidation potential (E_{ox}^0), and all potentials were obtained in CH_2Cl_2 with 0.1 M tetrabutylammonium perchlorate (TBAP) as the electrolyte (working electrode: glassy carbon; reference electrode: Ag/AgCl; counter electrode: Pt; calibrated with ferrocene/ferrocenium (Fc/Fc^+) as an external reference and converted to NHE by the addition of 0.69 V). ^d E_{0-0} was estimated from the absorption thresholds from the absorption spectra of dyes in solution. ^e The LUMO was estimated by subtracting E_{0-0} from the HOMO.



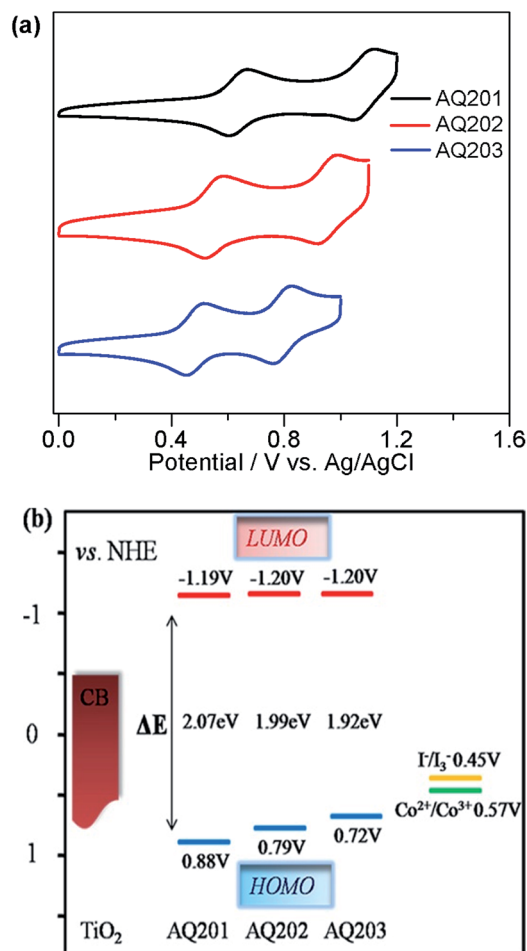


Fig. 2 (a) Cyclic voltammograms of dyes AQ201, AQ202, and AQ203; (b) schematic diagram of the energy levels of the TiO₂ conduction band, dyes, and [Co(bpy)₃]^{2+/3+} and I⁻/I₃⁻ redox couples.

reduce computational costs without affecting the nature of frontier molecular orbitals. The optimized structures were then used to calculate the 20 lowest singlet electronic transitions using time-dependent DFT calculations (TD-DFT) at the CAM-B3LYP/TZVP level.²¹ A polarizable continuum model (PCM) was applied using CH₂Cl₂ as the solvent.²² The TD-DFT results and frontier molecular orbitals of AQ201, AQ202, and AQ203 are shown in Tables 2 and 3, respectively.

Table 2 TD-DFT calculated energies, oscillator strengths (*f*), and compositions in terms of molecular orbital contributions of AQ201, AQ202, and AQ203

Dye	Composition ^a	<i>E</i> (nm)	<i>f</i>
AQ201	42% H → L 23% H-1 → L	463	1.5749
AQ202	44% H → L 21% H-1 → L	475	1.5401
AQ203	45% H → L 16% H-1 → L	507	1.9105

^a H = HOMO, L = LUMO, H-1 = HOMO-1.

Table 3 Frontier molecular orbitals of the dyes AQ201, AQ202, and AQ203 calculated at the B3LYP/6-31G level of theory

Dye	LUMO	HOMO	HOMO-1
AQ201			
AQ202			
AQ203			

Table 3 summarizes the orbital distributions of HOMO-1, HOMO, and LUMO. These orbitals were chosen here for discussion because they participate in the lowest singlet electronic transition according to TD-DFT. The three dyes share similar features in the spatial distribution of these frontier molecular orbitals. The HOMO is mainly located on the donor moieties and the first π -bridge on the left side of quinoxaline (Scheme 1), while HOMO-1 extends to the quinoxaline moiety. The LUMO is mainly located on the quinoxaline, phenyl ring and cyanoacrylic acid with a little contribution from the first π -bridge.

The electronic transitions shown by the TD-DFT calculations indicate that the absorption bands of AQ201, AQ202, and AQ203 with the maximum absorption peaks calculated at $\lambda = 463, 475,$ and 507 nm are dominated by HOMO → LUMO transitions with some contributions from HOMO-1 → LUMO. They agree qualitatively well with experimental results, although the energies were overestimated. The insertion of the EDOT and CPDT moiety significantly extends the π -conjugation of the donor in AQ202 and AQ203, which increases the HOMO energy and facilitates its oxidation. Consequently, they exhibit a lower HOMO → LUMO energy gap compared to AQ201, which, in turn, corresponds well with the red shift of the absorption spectra (Fig. 1a).

Photovoltaic performance

The DSSC performances of all the dyes with two different redox electrolytes were tested under AM1.5 G irradiation (100 mW cm^{-2}). Fig. 3a shows the photocurrent density voltage curves ($J-V$) of DSSCs based on AQ201, AQ202, and AQ203 dyes with the [Co(bpy)₃]^{2+/3+} redox electrolyte, and the photocurrent density voltage curves ($J-V$) of DSSCs based on the I⁻/I₃⁻ redox electrolyte are shown in Fig. 3b. The detailed photovoltaic parameters are collected in Table 4. Hereby, the influence of altering the π -bridge on the left side of the quinoxaline unit on the DSSC characteristics is apparent. For the devices based on the [Co(bpy)₃]^{2+/3+} redox electrolyte, the AQ202-based cell shows both higher power conversion efficiency (PCE) and short-circuit



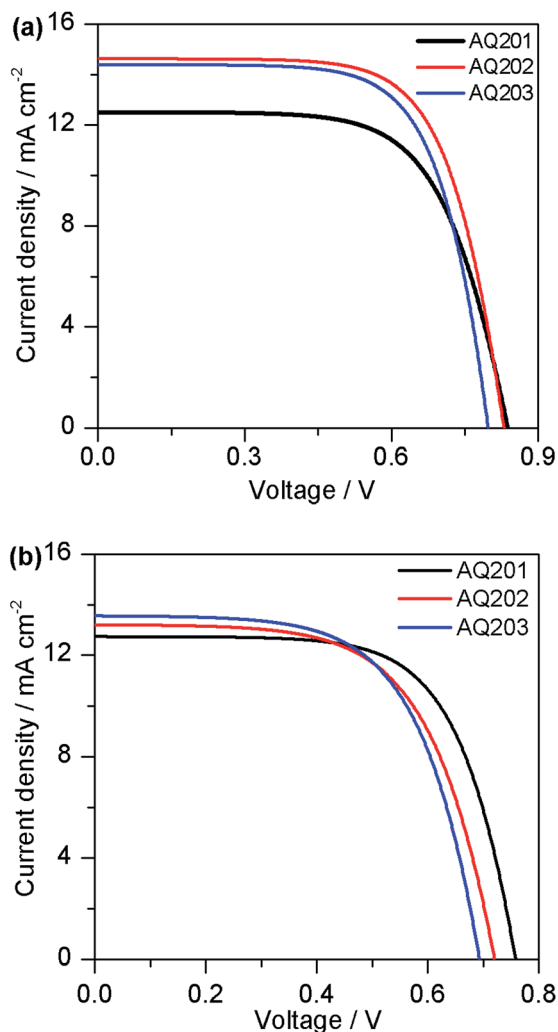


Fig. 3 J - V curves of DSSCs based on dyes AQ201, AQ202, and AQ203 with (a) the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox electrolyte and (b) the I^-/I_3^- redox electrolyte.

Table 4 Photovoltaic performance of the DSSCs based on AQ201, AQ202 and AQ203 under irradiation of AM 1.5 simulated solar light (100 mW cm^{-2})^a

Dye	Electrolyte	J_{sc} (mA cm^{-2})	V_{oc} (mV)	FF	η (%)
AQ201	$[\text{Co}(\text{bpy})_3]^{2+/3+}$	12.52	838	67.68	7.10
AQ202	$[\text{Co}(\text{bpy})_3]^{2+/3+}$	14.61	830	69.01	8.37
AQ203	$[\text{Co}(\text{bpy})_3]^{2+/3+}$	14.39	797	70.27	8.06
AQ201	I^-/I_3^-	12.75	756	68.16	6.57
AQ202	I^-/I_3^-	13.21	720	68.08	6.48
AQ203	I^-/I_3^-	13.56	692	68.08	6.39

^a The cobalt electrolyte contains 0.22 M $[\text{Co}(\text{II}) (\text{bpy})_3](\text{TFSI})_2$, 0.06 M $[\text{Co}(\text{III}) (\text{bpy})_3](\text{TFSI})_3$, 0.1 M LiClO_4 , and 0.5 M 4-*tert*-butylpyridine in acetonitrile, (bpy = 2,2'-bipyridine, and TFSI = bis(trifluoromethane) sulfonamide); the iodide electrolyte contains 0.1 M LiI , 0.03 M I_2 , 0.5 M 1-butyl-3-methyl imidazolium iodide (BMII), 0.1 M guanidine isothiocyanate, and 0.5 M 4-*tert*-butylpyridine (TBP) in acetonitrile : valeronitrile (85 : 15, v/v).

current density (J_{sc}), because the AQ202 dye has a wider absorption spectrum as well as matches the Marcus theory requirements of the gap between the HOMO level and $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox potential, thus effectively promoting dye regeneration. Clearly, the AQ201-based cell shows the lowest J_{sc} due to the relatively narrow absorption spectrum range of the three dyes. Although the AQ203 dye exhibits the widest absorption spectrum and highest molar extinction coefficient compared with the two other dyes, the J_{sc} of the AQ203-based cell with the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox electrolyte remains lower than that of the AQ202-based cell which is attributed to the smallest driving force for dye regeneration.

It is informative to compare the results of the devices based on the I^-/I_3^- redox electrolyte with those based on the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox electrolyte. The AQ203-based cell shows higher J_{sc} with the I^-/I_3^- redox electrolyte, because the driving force for dye regeneration is sufficient in iodine-based electrolytes. Meanwhile, the J_{sc} of the devices with the I^-/I_3^- redox electrolyte decreases in the trend of AQ203 > AQ202 > AQ201, which are positively correlated with their absorption spectra range. It should be noted that the AQ201-based cell shows the highest PCE with the I^-/I_3^- redox electrolyte due to the highest open-circuit voltage (V_{oc}) compared with the cells of the other two dyes. Furthermore, the effect of varying the π -bridge is highlighted in terms of the changes in V_{oc} , it is apparent that the V_{oc} of the devices with $[\text{Co}(\text{bpy})_3]^{2+/3+}$ or I^-/I_3^- redox electrolyte increases in the order of AQ201 > AQ202 > AQ203. This may be caused by the lowest HOMO level of AQ201 and then the highest dye regeneration efficiency, which reduces the probability of electron recombination from the conduction band of TiO_2 /excited state dyes to oxidised dyes and improves the V_{oc} .

Fig. 4 shows the IPCE curves of the AQ dye based devices with two different electrolytes. Notably, the IPCE onsets are well consistent with the corresponding electron donating ability, that is, thiophene (AQ201) < EDOT (AQ202) < CPDT (AQ203). In particular, all the AQ dye based devices exhibit small notches around 450 nm, which is attributed to the weak absorbance in this region according to the UV-visible absorption spectra. However, when compared with cobalt-based devices, iodide-based devices displayed a poorer IPCE response with a plateau of around 75% in the visible region, resulting in the lower J_{sc} . This may be ascribed to the bulky donor group being better suited for the devices with the cobalt-based electrolyte. As a result, the devices using the three dyes with the cobalt-based electrolyte realized higher performances compared to those with the iodide-based electrolyte.

Electrochemical impedance spectroscopy (EIS)

Generally, V_{oc} is influenced by the conduction band (E_{cb}) position and the electron recombination process at the TiO_2 /dye/electrolyte interface. To obtain a complete analysis of the difference of V_{oc} among these dyes with two different redox electrolytes, EIS was utilized to investigate the interfacial charge transfer and recombination processes of the devices. The EIS



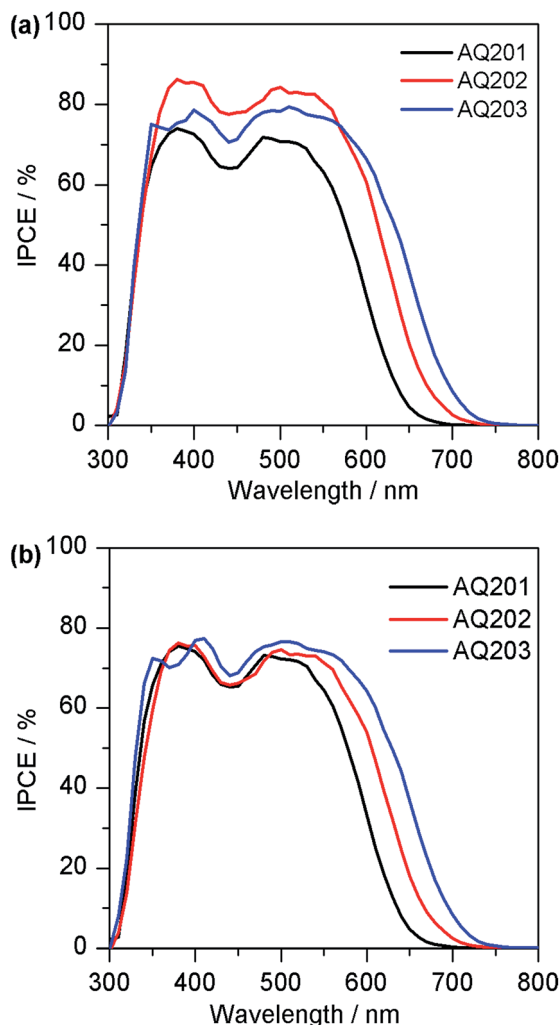


Fig. 4 IPCE curves of DSSCs based on dyes AQ201, AQ202, and AQ203 with (a) the $[\text{Co}(\text{bpy})_3]^{3+/2+}$ redox electrolyte and (b) the I^-/I_3^- redox electrolyte.

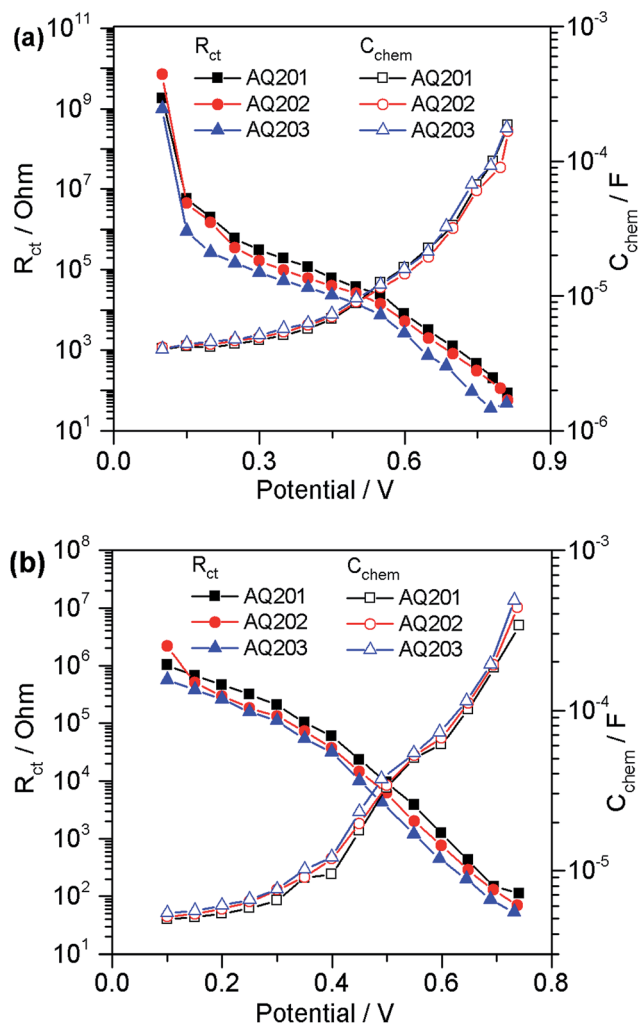


Fig. 5 Charge transfer resistance (R_{ct}) and chemical capacitance (C_{chem}) plotted against the IR drop corrected potential for DSSCs based on dyes AQ201, AQ202, and AQ203 with (a) the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox electrolyte and (b) the I^-/I_3^- redox electrolyte.

measurements were performed in the dark and data were fitted by using the transmission line model.²³ Fig. 5a depicts the charge transfer resistance (R_{ct}) (between the TiO_2 films and the electrolyte) and chemical capacitance (C_{chem}) of the devices based on AQ201, AQ202, and AQ203 dyes with the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox electrolyte as a function of the IR (internal resistance) drop corrected potential, and the R_{ct} and C_{chem} of the devices with the I^-/I_3^- redox electrolyte as a function of the IR drop corrected potential are shown in Fig. 5b. There is no difference in the R_{ct} value trend (AQ201 > AQ202 > AQ203) of the devices when applying either $[\text{Co}(\text{bpy})_3]^{2+/3+}$ or I^-/I_3^- as the redox electrolyte. We may attribute these small differences to the small differences in dye structures leading to slightly different effectiveness as a blocking layer.

Chemical capacitance, as the capacitive response of the TiO_2 photoanode, depends on the density of electronic state distribution below the E_{cb} , which can be used as an indirect measure of the Fermi level (E_{Fn}). Overall, there is a minor difference in the C_{chem} of the devices based on AQ201, AQ202, and AQ203

dyes with either $[\text{Co}(\text{bpy})_3]^{2+/3+}$ or I^-/I_3^- redox electrolyte, indicating only a small shift of the TiO_2 E_{Fn} position. Accordingly, with both $[\text{Co}(\text{bpy})_3]^{2+/3+}$ and I^-/I_3^- as redox electrolytes, the increase in the V_{oc} mainly results from an increase in the electron lifetime (τ_{n}), because τ_{n} , positively correlated with V_{oc} , is calculated by the equation of $\tau_{\text{n}} = R_{\text{ct}} \times C_{\text{chem}}$ (Fig. 6). Clearly, there is also no difference in the τ_{n} value trend (AQ201 > AQ202 > AQ203) of the devices using either $[\text{Co}(\text{bpy})_3]^{3+/2+}$ or I_2/I_3^- as the redox electrolyte, which is in good agreement with the observed V_{oc} trend of the devices. In conclusion, the differences in V_{oc} arise only from the structural difference in the π -bridge between the bulky triphenylamine donor group and quinoxaline, demonstrating that, similar to the π -bridge, the thiophene unit is beneficial for achieving higher V_{oc} , while the CPDT unit contributes negatively in this system.

Photoinduced charge-transfer dynamics

The above results show that the three dyes experience enough driving force for dye regeneration in the iodide-based



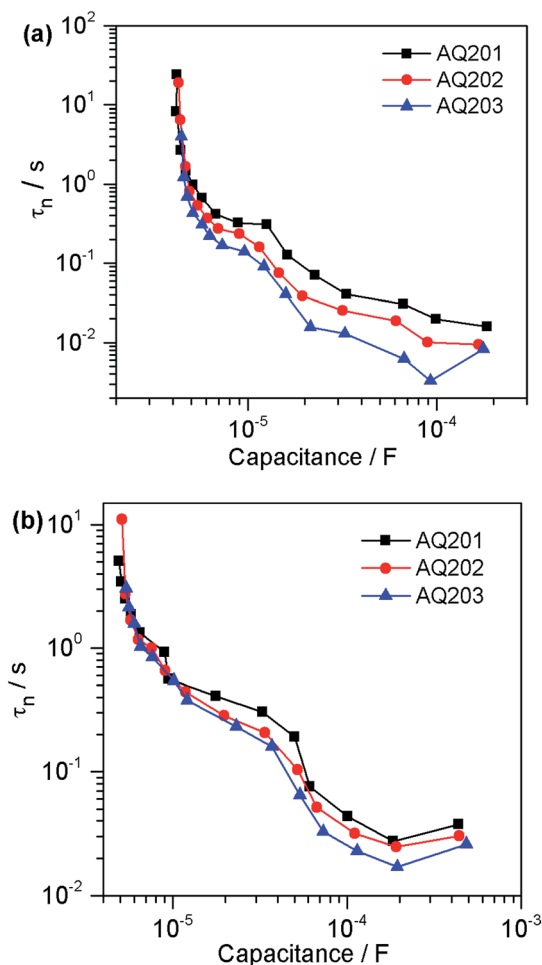


Fig. 6 Electron lifetime (τ_n) plotted against chemical capacitance (C_{chem}) of DSSCs based on dyes AQ201, AQ202, and AQ203 with (a) the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox electrolyte and (b) the I^-/I_3^- redox electrolyte.

electrolyte, while the devices with the cobalt-based electrolyte result in higher PCE. Accordingly, micro-to-millisecond transient absorption spectroscopy (TAS) measurements were focused on the cobalt-based electrolyte to gain insight into the influence of the π -bridge on the oxidized dye lifetime and its regeneration in the $[\text{Co}(\text{bpy})_3]^{2+/3+}$ redox system. Photoinduced absorption spectroscopy (PIA) is used to locate spectral features corresponding to the oxidized dye absorbance. All dyes exhibit comparable spectral features in their oxidized states, which are measured under 510 nm excitation and absorbance changes are reported in Fig. 7. An isosbestic point between the absorbance of the dye ground and oxidized states is found for these dyes at around 700 nm. It is observed that replacing the thiophene with EDOT or CPDT group not only affects the ground state electronic devices, but also decreases the oxidized state absorbance, with a modification in its absorbance shape. For all the dyes, we observe a positive transient absorbance of around 1000–1100 nm. Hence, following the decay kinetics of the transient absorption signal measured at 1050 nm allowed for the analysis of the charge recombination and dye regeneration rate constants.

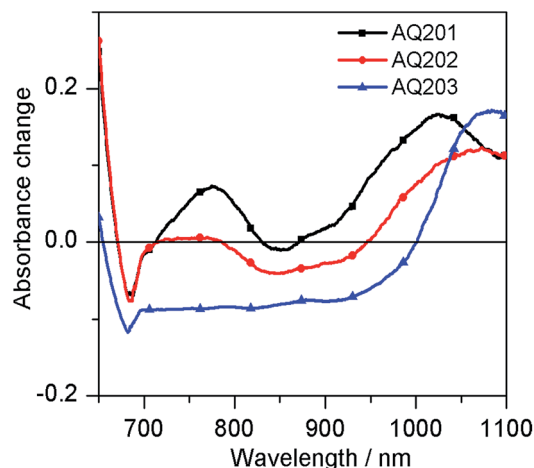


Fig. 7 PIA spectra measured upon laser excitation ($\lambda_{\text{exc}} = 510$ nm) of AQ201, AQ202 and AQ203 dyes adsorbed on $4 \mu\text{m}$ mesoporous TiO_2 films in the absence of redox electrolytes.

The electron transfer and dye regeneration lifetimes of AQ201 and AQ202 were extracted by fitting the spectra with mono-exponential decays (acetonitrile) and bi-exponential decays (cobalt). However, the kinetic fit of AQ203 was extracted by fitting the spectra with tri-exponential decays (acetonitrile and cobalt). This may be due to the larger size and more donating nature of the CPDT group than those of thiophene and EDOT groups, or even to different aggregation behavior. It is likely that not only triphenylamine but also the CPDT unit is acting as a charge donor. And then, AQ203 showed kinetically complicated behavior under the same conditions, which probably refers to several complicated recombination processes. Fig. 8 shows the TAS signals of AQ201, AQ202, and AQ203 in the absence and presence of the cobalt redox electrolyte under similar conditions, and their fitting curves. In the absence of a redox mediator in pure acetonitrile solvent, the decay of the absorption signal presents the recombination process of the oxidized dye and the electrons injected in TiO_2 . Oxidized AQ201 has a lifetime of 478 μs . Elevating the HOMO level leads to a decrease in the lifetime of AQ202 and AQ203 to 316 and 262 μs , respectively, confirming hole stabilization by the lower HOMO levels. In the presence of the cobalt redox couple, regeneration of the oxidized dye occurs and significantly accelerates the decay of the oxidized dye signal. The lifetime is 13.4 μs for AQ201, 67.4 μs for AQ202 and 81.3 μs for AQ203. Assuming a first-order rate for both dye regeneration and back electron transfer reactions, we can estimate regeneration yields for all the dyes according to $\eta_{\text{reg}} = k_{\text{reg}}/(k_{\text{reg}} + k_{\text{rec}})$, $k_i = 1/\tau_i$. The summary of the lifetime and the dye regeneration yields (η_{reg}) is shown in Table 5. We conclude from these results that regeneration of the three dyes will be the limiting step in the operation of devices with the cobalt electrolyte. However, dye regeneration is certainly governed by the chemical structures of the dyes, especially with respect to its impact on the HOMO energy level.



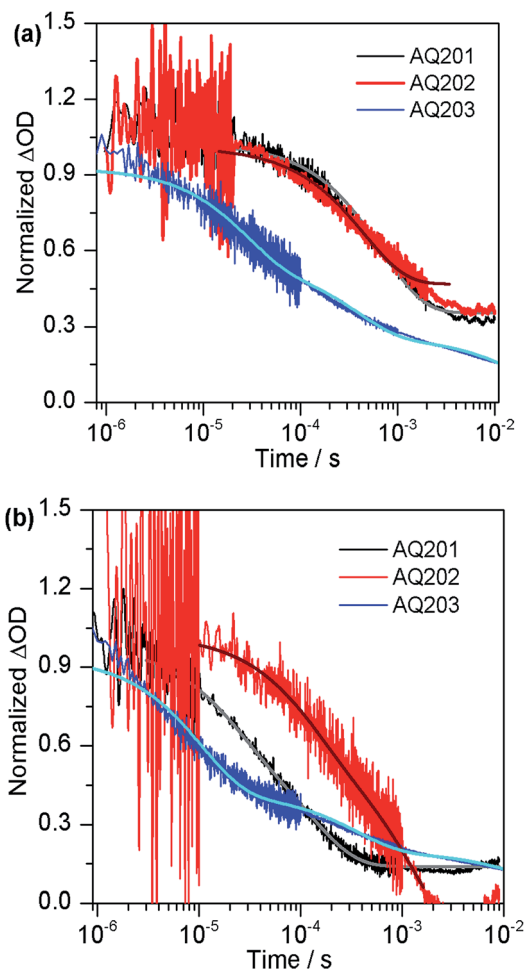


Fig. 8 Microsecond normalized transient absorption decays of AQ201, AQ202 and AQ203 dyes at 1050 nm (a) in the absence (no iodide and cobalt electrolytes added) and (b) in the presence of the cobalt redox electrolyte under similar conditions.

Table 5 Time constants extracted from the fits of transient absorption and calculated regeneration yields

Dye	τ_{rec} (μs)	$\tau_{\text{reg cobalt}}$ (μs)	$\eta_{\text{reg cobalt}}$
AQ201	478	13.4	0.97
AQ202	316	67.4	0.83
AQ203	262	81.3	0.76

Conclusion

In summary, we have synthesized three new quinoxaline-based organic dyes and studied the effect of the different π -bridges on the HOMO levels of the dyes and the performance of DSSCs with $[\text{Co}(\text{bpy})_3]^{2+/3+}$ and Γ^-/I_3^- based electrolytes. The cyclic voltammograms show that the HOMO levels of the dyes increase in the trend of AQ203 > AQ202 > AQ201, while the LUMO levels of these dyes remained the same. It is interesting to note that the AQ201-based DSSC with the lowest HOMO level showed the poorest PCE of 7.10% with the cobalt-based electrolyte, but the

best PCE of 6.57% with the iodide-based electrolyte. Moreover, the AQ202-based cell shows the highest PCE of 8.37% with the cobalt-based electrolyte, due to a wider absorption spectrum as well as a higher driving force for dye regeneration. Last but not least, the AQ203-based cells are not the best with either $[\text{Co}(\text{bpy})_3]^{2+/3+}$ or Γ^-/I_3^- based electrolyte despite the widest absorption spectrum and highest molar extinction coefficient, due to the smallest driving force for dye regeneration. Electrochemical impedance spectroscopy (EIS) and transient absorbance spectroscopy (TAS) indicate that the lower HOMO level leads to higher electron lifetime and dye regeneration efficiency. This work emphasizes the significance of the HOMO level of the dye in the trade-off between the spectral absorption range and energy level matching, which are responsible for determining the device PCE. Designing rational metal-free organic dyes will be focused in further work towards molecular energy level tuning, giving possibilities to increase the device performance.

Experimental

Materials

All chemicals and reagents were from commercial sources without purification. Solvents for chemical synthesis were purified by distillation. All chemical reactions were carried out under an argon atmosphere. The starting materials [4-bis(2',4'-bis((2-ethylhexyl)oxy)biphenyl-4-yl)amino]phenylborate^{14b} and 5,8-dibromo-2,3-diphenylquinoxaline¹¹ were synthesized according to the published procedures.

Characterization

¹H NMR and ¹³C NMR spectra were recorded on Bruker AM-400 MHz instruments with tetramethylsilane (TMS) as the internal standard. HRMS was performed by using a Waters LCT Premier XE spectrometer. The absorption spectra of the sensitizer dyes in solution and those adsorbed on TiO₂ films were measured with a Varian Cary 500 spectrophotometer. The stability and redox behaviors of dye molecules were analyzed by cyclic voltammetry (CV) in CH₂Cl₂ solutions without deoxygenation. The cyclic voltammograms (CV) were determined by using a CHI660C electrochemical workstation (Chenhua Co. Ltd, Shanghai, China) in a three-electrode cell. The working electrode was glassy carbon, used in conjunction with a Pt auxiliary electrode and an Ag/AgCl wire pseudo-reference electrode. All the experiments were performed in CH₂Cl₂ solutions containing 0.1 M tetrabutylammonium perchlorate (TBAP) at a scan rate of 100 mV s⁻¹. The ferrocenium/ferrocene (Fc/Fc⁺) redox couple was used as an external potential reference.

Fabrication and photovoltaic measurements

4 or 8 μm TiO₂ films composed of 20 nm nanoparticles, α -terpineol and ethyl cellulose, and a 4 μm thick scattering layer (60% nanoparticles of 20 nm and 40% large particles of 100 nm) was prepared according to the established method.¹⁰ These nanocrystal titania films were treated with 40 mM TiCl₄ aqueous solution at room temperature overnight, following by



sintering at 450 °C for 30 min. After cooling down to room temperature, these films were dipped into a 3×10^{-4} M solution of dyes in chloroform/ethanol (v/v: 3/7) mixed solvent with 5 mM CDCA for 3 h. After this, the electrodes were rinsed with CH_2Cl_2 and ethanol and then dried. The size of the electrodes was 0.12 cm^2 . The counter electrode (Pt-coated TCO) and the working electrode (dye-loaded TiO_2 film) were sealed together by a hot-melt 30 μm thick spacer under heating. The Pt-counter electrodes were prepared following the literature.¹⁰ The iodide electrolyte contained 0.1 M LiI, 0.03 M I_2 , 0.5 M 1-butyl-3-methylimidazolium iodide (BMII), 0.1 M guanidine isothiocyanate, and 0.5 M 4-*tert*-butylpyridine (TBP) in acetonitrile : valeronitrile (85 : 15, v/v), and the cobalt electrolyte contained 0.22 M $[\text{Co}(\text{II}) (\text{bpy})_3](\text{TFSI})_2$, 0.06 M $[\text{Co}(\text{III}) (\text{bpy})_3](\text{TFSI})_3$, 0.1 M LiClO_4 , and 0.5 M 4-*tert*-butylpyridine in acetonitrile, (bpy = 2,2'-bipyridine, and TFSI = bis(trifluoromethane)sulfonamide).

The DSSCs were evaluated by recording the J - V curves with a Keithley 2400 source meter under the illumination of Air Mass 1.5 G simulated solar light coming from a solar simulator (Oriental 91160 equipped with a 300 W Xe lamp). Incident monochromatic IPCE spectra were obtained with an Oriental-74125 system and the intensity of incident monochromatic light was measured with a Si detector (Oriental-71640). Electrical impedance experiments (EIS) were performed in the dark. The spectra were scanned in a frequency range from 0.1 Hz to 100 kHz at room temperature with the applied potential set at open circuit. The alternate potential amplitude was set at 10 mV.

Photoinduced charge-transfer dynamics

The nanosecond laser flash photolysis technique was applied to the dye-sensitized, 4 μm thick, transparent TiO_2 mesoporous films deposited on normal flint glass. The samples were excited by pulses produced at a repetition rate of 30 Hz by a broad-band optical parametric oscillator (GWU OPO-355) pumped by a Powerlite 7030 frequency-tripled Q-switched Nd : YAG laser (Continuum, Santa Clara, California, USA). The output excitation wavelength was tuned at $\lambda_{\text{exc}} = 530$ nm with a pulse width of 7 ns (fwhm). The laser beam was expanded by a planoconcave lens to irradiate a large cross-section of the sample, whose surface was kept at a 45° angle to the excitation beam. The probe light, produced by a continuous wave xenon arc lamp, was first passed through a monochromator tuned in 20 nm steps throughout the 420–740 nm range of the VIS light spectrum. The probe light passed various optical elements, the sample, and then a second monochromator, before being detected by a fast photomultiplier tube (Hamamatsu, R9110). Data waves were recorded on a DSA 602A digital signal analyzer (Tektronix, Beaverton, Oregon, USA). Satisfactory signal-to-noise ratios were typically obtained by averaging over 1500 laser shots. The transient absorption measurements were performed on the previously described dye-sensitized, 4 μm -thick, transparent TiO_2 films either immersed in $[\text{Co}(\text{bpy})_3]^{2+/3+}$ or I^-/I_3^- based electrolyte.

Synthesis

Synthesis of 2a. Under an argon atmosphere, compound 1 (439 mg, 1.0 mmol), K_2CO_3 (5 mL, 2 M aqueous solution) and

$\text{Pd}(\text{PPh}_3)_4$ (42 mg, 0.036 mmol) were dissolved in 20 mL of THF. After stirring for half an hour, the mixture was heated to 50 °C and a solution of (4-formylphenyl)boronic acid (1.0 mmol) in THF (5 mL) was added slowly. The mixture was then refluxed for 12 h before cooling to room temperature and extracted with CH_2Cl_2 . The organic layers were washed with brine water and dried over anhydrous Na_2SO_4 . After removing the solvent under reduced pressure, the residue was chromatographed on a silica gel column with PE/ CH_2Cl_2 (3/1–1/1, v/v) as eluent to give compound 2 (260 mg, 56% yield). ^1H NMR (400 MHz, CDCl_3 , ppm): $\delta = 10.15$ (s, 1H), 8.18 (d, $J = 8.4$ Hz, 1H), 8.06 (d, $J = 8.4$ Hz, 2H), 7.99 (d, $J = 8.8$ Hz, 2H), 7.75–7.71 (m, 3H), 7.56 (d, $J = 8.8$ Hz, 2H), and 7.44–7.32 (m, 6H). ^{13}C NMR (100 MHz, CDCl_3 , ppm): $\delta = 192.19, 153.26, 153.01, 143.81, 139.03, 138.74, 138.68, 138.29, 138.23, 135.47, 132.92, 131.44, 130.34, 130.19, 130.00, 129.46, 129.37, 129.33, 128.38, 128.29, \text{ and } 124.51$. HRMS (m/z): $[\text{M} + \text{H}^+]$ calcd for $\text{C}_{27}\text{H}_{18}\text{N}_2\text{OBr}$, 465.0602; found: 465.0538.

Synthesis of 3a. Under an argon atmosphere, 2-(tributylstannyl)thiophene (193 mg, 0.52 mmol), compound 2 (200 mg, 0.43 mmol) and $\text{Pd}(\text{dppf})\text{Cl}_2$ (7.3 mg, 0.01 mmol) were dissolved in anhydrous toluene (20 mL) and refluxed for 12 h. The organic phase was extracted with CH_2Cl_2 , washed with HCl (1 M), dried over Na_2SO_4 and concentrated under reduced pressure. The crude solid was purified by chromatography on silica gel using PE/ CH_2Cl_2 (3/1–1/1, v/v) as eluent to give a yellow solid 3a (180 mg, 89% yield). ^1H NMR (400 MHz, CDCl_3 , ppm): $\delta = 10.13$ (s, 1H), 8.23 (d, $J = 8.8$ Hz, 1H), 8.04 (s, 4H), 7.93 (d, $J = 4.4$ Hz, 1H), 7.89 (d, $J = 8.4$ Hz, 1H), 7.76–7.74 (m, 2H), 7.59–7.56 (m, 3H), 7.41–7.29 (m, 6H), and 7.22 (t, $J = 4.2$ Hz, 1H). ^{13}C NMR (100 MHz, CDCl_3 , ppm): $\delta = 192.27, 152.05, 151.79, 144.45, 138.69, 138.63, 138.53, 138.51, 137.21, 137.10, 135.27, 133.07, 131.48, 130.44, 130.27, 130.03, 129.75, 129.36, 129.14, 128.88, 128.31, 128.28, 126.90, 126.84, \text{ and } 126.72$. HRMS (m/z): $[\text{M} + \text{H}^+]$ calcd for $\text{C}_{31}\text{H}_{21}\text{N}_2\text{OS}$, 469.1375; found: 469.1388.

Synthesis of 4a. Compound 3a (160 mg, 0.34 mmol) and *N*-bromosuccinimide (NBS) (73 mg, 0.41 mmol) were dissolved in anhydrous CH_2Cl_2 in an ice bath for 2 h. Then, a Suzuki coupling reaction was performed with [4-bis(2',4'-bis((2-ethylhexyl)oxy)bi-phenyl-4-yl)amino]phenylborate by using a similar procedure to that of compound 2 to give an orange solid 4a (210 mg, 45% yield). ^1H NMR (400 MHz, CDCl_3 , ppm): $\delta = 10.12$ (s, 1H), 8.23 (d, $J = 8.8$ Hz, 1H), 8.04 (s, 4H), 7.83 (d, $J = 4.4$ Hz, 1H), 7.89 (d, $J = 8.4$ Hz, 1H), 7.80–7.78 (m, 2H), 7.61–7.59 (m, 4H), 7.45 (d, $J = 8.2$ Hz, 4H), 7.41–7.39 (m, 3H), 7.36–7.27 (m, 6H), 7.22–7.18 (m, 6H), 6.57–6.55 (m, 4H), 3.89–3.86 (m, 8H), 1.76–1.68 (m, 4H), 1.45–1.26 (m, 32H), and 0.97–0.87 (m, 24H). ^{13}C NMR (100 MHz, CDCl_3 , ppm): $\delta = 192.26, 159.93, 157.27, 152.02, 151.56, 147.80, 147.67, 145.61, 144.51, 138.75, 138.66, 138.59, 137.04, 136.71, 135.23, 133.46, 133.12, 131.48, 130.77, 130.55, 130.48, 130.34, 130.08, 129.37, 129.21, 129.17, 128.30, 128.12, 126.42, 126.08, 123.96, 123.44, 123.17, 122.03, 105.22, 100.35, 70.75, 70.56, 39.54, 39.44, 30.70, 30.64, 29.20, 29.07, 24.13, 23.97, 23.16, 23.14, 14.21, 11.27, \text{ and } 11.25$. HRMS (m/z): $[\text{M} + \text{H}^+]$ calcd for $\text{C}_{93}\text{H}_{106}\text{N}_3\text{O}_5\text{S}$, 1376.7853; found: 1376.7836.

Synthesis of AQ201. Under an argon atmosphere, a mixture of 4a (170 mg, 0.12 mmol), cyanoacetic acid (85 mg, 1 mmol),



and ammonium acetate (120 mg) in acetic acid (25 mL) and THF (5 mL) was heated at reflux for 12 h. After cooling down, the mixture was poured into water. The precipitates were collected by filtration and washed with water. The residue was purified by column chromatography on silica gel with CH₂Cl₂/ethanol (10 : 1, v/v) to yield a red solid **AQ201** (120 mg, 69% yield). ¹H NMR (400 MHz, acetone-*d*₆, ppm): 8.45–8.43 (m, 2H), 8.26 (d, *J* = 8.8 Hz, 2H), 8.14 (d, *J* = 8.2 Hz, 2H), 8.08–8.04 (m, 2H), 7.83 (d, *J* = 8.4 Hz, 2H), 7.71 (d, *J* = 8.4 Hz, 2H), 7.66 (d, *J* = 8.8 Hz, 2H), 7.51–7.49 (m, 5H), 7.45–7.44 (m, 3H), 7.40–7.35 (m, 3H), 7.28 (d, *J* = 8.4 Hz, 2H), 7.17 (d, *J* = 8.8 Hz, 6H), 6.69 (s, 2H), 7.64 (d, *J* = 8.2 Hz, 2H), 3.96–3.94 (m, 8H), 1.76–1.68 (m, 4H), 1.50–1.30 (m, 32H), and 0.98–0.86 (m, 24H). ¹³C NMR (100 MHz, acetone-*d*₆, ppm): δ = 161.01, 158.17, 149.46, 146.41, 143.98, 134.76, 132.50, 131.43, 131.35, 130.92, 129.07, 129.02, 127.16, 124.71, 124.13, 123.68, 123.05, 106.71, 101.08, 71.35, 71.04, 40.43, 40.35, 31.51, 31.34, 24.61, 23.78, 23.76, 14.46, 14.39, 11.64, 11.48, and 5.45. HRMS (*m/z*): [M + H⁺] calcd for C₉₆H₁₀₇N₄O₆S, 1443.7911; found: 1443.7953.

Synthesis of 3b. Compound **3b** was obtained by using a similar procedure to that of **3a** (150 mg, 75% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ = 10.05 (s, 1H), 8.63 (d, *J* = 8.8 Hz, 1H), 7.96 (s, 4H), 7.82 (d, *J* = 4.4 Hz, 1H), 7.69 (d, *J* = 8.4 Hz, 2H), 7.50 (d, *J* = 8.4 Hz, 2H), 7.32–7.23 (m, 6H), and 7.55 (s, 1H). ¹³C NMR (100 MHz, CDCl₃, ppm): δ = 191.25, 150.66, 150.06, 143.61, 140.35, 139.74, 137.71, 137.46, 137.25, 136.09, 134.83, 134.07, 130.63, 130.41, 129.48, 129.37, 128.98, 128.28, 127.98, 127.19, 126.64, 111.86, 102.97, 64.02, and 63.27. HRMS (*m/z*): [M + H⁺] calcd for C₃₃H₂₃N₂O₃S, 527.1429; found: 527.1428.

Synthesis of 4b. Compound **4b** was obtained by using a similar procedure to that of **4a** (190 mg, 80% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ = 10.11 (s, 1H), 8.74 (d, *J* = 8.8 Hz, 1H), 8.03 (s, 4H), 7.88 (d, *J* = 8.4 Hz, 1H), 7.84–7.82 (m, 2H), 7.75 (d, *J* = 8.4 Hz, 2H), 7.60 (d, *J* = 8.4 Hz, 2H), 7.45 (d, *J* = 8.8 Hz, 4H), 7.39–7.35 (m, 4H), 7.33–7.23 (m, 6H), 7.20–7.18 (m, 4H), 6.57–6.56 (m, 4H), 4.48–4.47 (m, 2H), 4.42–4.41 (m, 2H), 3.90–3.87 (m, 8H), 1.77–1.70 (m, 4H), 1.45–1.27 (m, 32H), and 0.98–0.87 (m, 24H). ¹³C NMR (100 MHz, CDCl₃, ppm): δ = 192.31, 159.86, 157.23, 151.67, 150.82, 138.81, 138.49, 135.78, 135.08, 131.46, 130.67, 130.34, 130.06, 129.36, 129.14, 129.05, 128.26, 128.18, 125.55, 123.73, 123.19, 121.37, 105.18, 100.31, 70.72, 70.55, 64.94, 64.39, 39.50, 39.40, 34.25, 30.64, 30.59, 30.35, 29.73, 29.47, 29.15, 29.02, 24.08, 23.92, 23.11, 23.08, 14.15, 11.20, and 11.19. HRMS (*m/z*): [M + H⁺] calcd for C₉₅H₁₀₈N₃O₇S, 1434.7908; found: 1437.7910.

Synthesis of AQ202. Compound **AQ202** was obtained by using a similar procedure to that of **AQ201** (90 mg, 78% yield). ¹H NMR (400 MHz, acetone-*d*₆, ppm): δ = 8.73 (s, 1H), δ = 8.37 (s, 1H), 8.18 (d, *J* = 8.8 Hz, 1H), 8.06 (d, *J* = 8.8 Hz, 1H), 7.88–7.72 (m, 5H), 7.59 (d, *J* = 8.4 Hz, 2H), 7.45 (d, *J* = 4.2 Hz, 4H), 7.36–7.28 (m, 6H), 7.23 (d, *J* = 8.4 Hz, 2H), 7.12–7.18 (m, 6H), 6.66–6.56 (m, 4H), 4.43–4.36 (m, 4H), 3.92–3.91 (m, 8H), 1.73–1.66 (m, 4H), 1.44–1.28 (m, 32H), and 0.97–0.84 (m, 24H). ¹³C NMR (100 MHz, acetone-*d*₆, ppm): δ = 160.92, 158.14, 155.10, 146.51, 139.87, 139.47, 134.39, 132.43, 131.58, 131.45, 131.32, 131.25, 130.97, 129.74, 128.93, 127.75, 124.41, 123.74, 116.75, 106.65, 103.72, 101.13, 71.38, 71.04, 65.16, 40.40, 40.35, 31.50,

31.34, 24.84, 24.61, 23.77, 14.49, 14.42, 11.65, and 11.50. HRMS (*m/z*): [M + H⁺] calcd for C₁₁₄H₁₄₀N₃O₅S₂, 1501.7966; found: 1501.7733.

Synthesis of 3c. Compound **3c** was obtained by using a similar procedure to that of **3a** (320 mg, 47% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ = 10.12 (s, 1H), 8.22–8.20 (m, 1H), 8.03 (s, 4H), 7.88–7.86 (m, 2H), 7.79–7.77 (m, 2H), 7.58 (d, *J* = 4.4 Hz, 2H), 7.45–7.44 (m, 3H), 7.35–7.29 (m, 3H), 7.20 (d, *J* = 4.2 Hz, 1H), 6.99–6.97 (m, 1H), 1.97–1.93 (m, 4H), 1.29–1.22 (m, 2H), 1.01–0.88 (m, 16H), and 0.78–0.60 (m, 12H).

Synthesis of 4c. Compound **4c** was obtained by using a similar procedure to that of **4a** (250 mg, 65% yield). ¹H NMR (400 MHz, CDCl₃, ppm): δ = 10.05 (s, 1H), 8.13 (s, 1H), 7.97 (s, 4H), 7.80 (d, *J* = 8.4 Hz, 2H), 7.72–7.70 (m, 2H), 7.51 (d, *J* = 8.8 Hz, 2H), 7.42–7.35 (m, 9H), 7.28–7.21 (m, 5H), 7.10 (d, *J* = 8.2 Hz, 7H), 6.49–6.48 (m, 4H), 3.82–3.78 (m, 8H), 1.94–1.86 (m, 4H), 1.69–1.61 (m, 4H), 1.35–1.18 (m, 34H), 0.95–0.80 (m, 40H), and 0.68–0.55 (m, 12H). ¹³C NMR (100 MHz, CDCl₃): δ = 192.29, 159.86, 157.21, 138.79, 138.65, 135.12, 133.28, 131.42, 130.69, 130.49, 130.38, 130.05, 129.39, 129.07, 128.37, 128.26, 125.92, 123.73, 123.15, 105.15, 100.27, 70.69, 70.52, 43.43, 39.48, 39.39, 35.21, 34.19, 30.63, 30.57, 29.14, 29.01, 28.61, 27.45, 24.04, 23.90, 23.10, 23.07, 22.84, 14.14, 11.17, and 10.70. HRMS (*m/z*): [M + H⁺] calcd for C₁₁₄H₁₄₀N₃O₅S₂, 1695.0234; found: 1695.0217.

Synthesis of AQ203. Compound **AQ203** was obtained by using a similar procedure to that of **AQ201** (85 mg, 63% yield). ¹H NMR (400 MHz, acetone-*d*₆, ppm): δ = 8.35 (d, *J* = 8.4 Hz, 1H), 8.29 (s, 1H), 8.16–8.12 (m, 3H), 8.03 (d, *J* = 8.8 Hz, 2H), 7.93 (d, *J* = 8.8 Hz, 1H), 7.69–7.67 (m, 2H), 7.54–7.50 (m, 4H), 7.37–7.35 (m, 7H), 7.24–7.22 (m, 3H), 7.15–7.13 (m, 3H), 7.04–7.00 (m, 6H), 6.56 (s, 2H), 6.48 (d, *J* = 8.8 Hz, 2H), 3.83–3.81 (m, 8H), 1.64–1.53 (m, 4H), 1.24–1.16 (m, 34H), 0.85–0.74 (m, 40H), and 0.60–0.51 (m, 12H). ¹³C NMR (100 MHz, acetone-*d*₆) δ = 160.98, 158.16, 146.42, 139.96, 134.61, 132.43, 131.43, 131.35, 130.92, 129.13, 129.00, 124.54, 123.69, 106.69, 101.09, 71.35, 71.05, 54.79, 43.91, 40.43, 40.34, 36.11, 35.02, 32.64, 31.51, 31.34, 24.82, 24.61, 23.78, 23.76, 23.57, 23.51, 23.34, 14.49, 14.41, 11.65, 11.50, and 11.18. HRMS (*m/z*): [M + H⁺] calcd for C₁₁₇H₁₄₁N₄O₆S₂, 1762.0293; found: 1762.0281.

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