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Dye-sensitised semiconductors modified with molecular catalysts for light-driven H₂ production

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The development of synthetic systems for the conversion of solar energy into chemical fuels is a research goal that continues to attract growing interest owing to its potential to provide renewable and storable energy in the form of a 'solar fuel'. Dye-sensitised photocatalysis (DSP) with molecular catalysts is a relatively new approach to convert sunlight into a fuel such as H₂ and is based on the self-assembly of a molecular dye and electrocatalyst on a semiconductor nanoparticle. DSP systems combine advantages of both homogenous and heterogeneous photocatalysis, with the molecular components providing an excellent platform for tuning activity and understanding performance at defined catalytic sites, whereas the semiconductor bridge ensures favourable multi-electron transfer kinetics between the dye and the fuel-forming electrocatalyst. In this *tutorial review*, strategies and challenges for the assembly of functional molecular DSP systems and experimental techniques for their evaluation are explained. Current understanding of the factors governing electron transfer across inorganic-molecular interfaces is described and future directions and challenges for this field are outlined.

Key learning points

- (1) Key principles of DSP operation in comparison with dye-sensitised solar cells and natural photosynthesis
- (2) Rational design principles for self-assembly of molecular dyes and catalysts on semiconductor nanoparticles
- (3) Methods for assembling and evaluating DSP systems, including bulk performance and spectroscopic techniques
- (4) Applications and current understanding of state-of-the-art DSP
- (5) Outlook into future developments of DSP

1. Introduction

The development of new materials and devices capable of providing non-fossil derived fuels on a worldwide scale is key to ensuring the future sustainability and security of global energy supplies. Of all the available renewable primary energy sources, sunlight is by far the most abundant and the development of systems capable of its capture, conversion and storage remains the most promising means to meet global energy demands. While solar energy to electricity conversion has seen great progress in recent decades, direct conversion of this energy to chemical bonds tackles the challenges of energy storage and transport as well as providing a direct source of renewable fuel for use in established fuel-based technologies.¹

Hydrogen is a strong candidate for a renewable energy

economy since it is a clean burning and energy dense fuel, as well as an important reagent for the chemical industry. Whereas the current primary H₂ production process by steam reformation is highly energy intensive and uses fossil sources, solar-driven electrolysis of water would generate H₂ from inexhaustible resources.¹ Furthermore, extension of the water-splitting process to include CO₂ reduction could open up the possibility of a sustainable economy by generating renewable carbon-based molecules that could be used as fuels or chemical feedstocks.²

Solar water splitting systems use electrons and holes photogenerated by a light-absorber to drive the water oxidation and proton reduction half-reactions at electrocatalyst sites. While many semiconductors and some molecular complexes are capable of acting as both light-absorber and catalyst,³ the majority of the systems reported are multi-component, coupling a light-absorber with a separate electrocatalyst in the form of a metal complex, an enzyme, or inorganic (nano)particles. The two halves of the reaction are often studied separately, with sacrificial electron donors or acceptors closing the catalytic cycle; however,

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practical devices without sacrificial agents will ultimately be required to carry out full water splitting. Several device design strategies for achieving full water splitting have been proposed, including solar cell-driven electrolysis (PV-electrolysers), coupling of complementary photoelectrodes in a photoelectrochemical cell (PEC cell),⁴ and chemically coupling two photocatalysts to form Z-scheme suspension systems.⁵

Photocatalytic systems based on molecular catalysts have been reported for both water oxidation and proton reduction.⁶ We focus herein on photo-reduction systems, which can be broadly divided into two categories: all-molecular (homogeneous) or semiconductor-based (heterogeneous; Fig. 1). In both, electron transfer (ET) from the photosensitiser to the catalyst can be achieved either directly or *via* a molecular mediator ($2M/2M^-$; e.g. methyl viologen) by simple mixing of components in solution; however, these approaches are inherently mass transfer (MT) limited.⁷ Linking the two components together through a covalent bond or another surface interaction (e.g. on a semiconductor) greatly enhances the rate of charge transfer by both increasing the effective local concentration of the components and by creating orbital overlap to facilitate electronic communication between the two species.

In this *tutorial review*, we focus on the emerging approach of dye-sensitised photocatalysis (DSP, Fig. 1). DSP bridges the fields of molecular and materials chemistry by coupling a molecular dye and catalyst through their co-adsorption onto a semiconducting particle, using the particle as both a scaffold and as a solid-state electron mediator. Taking inspiration from dye-sensitised solar cells (DSCs), such systems combine the capacity of semiconductors to accumulate and transport multiple charge carriers with the facile tuning of the system's optical properties by transferring the light absorption function to an interchangeable molecular dye. The immobilisation of dyes and catalysts onto a semiconductor surface can also allow the use of water insoluble molecules in aqueous media. DSP systems are synthetically facile to construct by directed and

rational self-assembly in solution, leading to a 'mix and match' approach, whereby each component can be optimised individually, and the overall activity of the system is highly tuneable. However, to date, the efficiencies of such hybrid systems have been relatively low in comparison to their individual components.⁸ In order to advance this promising field, it is of importance to understand the interfacial interactions and ET processes occurring between the components. Here we aim to provide an introduction to the field of DSP, by first discussing the thermodynamic and kinetic requirements and design principles for such systems, followed by a discussion of DSP self-assembly, characterisation, and the evaluation of performance of such systems for solar fuel generation.

2. DSP Systems

2.1 From 'solar-to-electricity' to 'solar-to-fuels'

The basic concept of integrating molecular dyes with a semiconducting material is well-established in DSC technology for the conversion of sunlight into electricity. In a typical DSC setup (Fig. 2a and b), photogenerated electrons are collected at a TiO_2 photoanode and holes are collected by the reduced species of a redox mediator (M/M^-), which is then re-generated at the cathode to close the redox cycle. The key component of this type of solar cell is the efficient absorption of sunlight by the dye and the almost quantitative electron injection into the conduction band (CB) of the semiconductor.⁹

DSP builds upon the excellent dye-semiconductor interface, but electrons are transferred to an electrocatalyst for fuel production instead of generating electricity (Fig. 2c). Electrons from the photoexcited dye are transferred to a H_2 evolution catalyst (cat_{red}) *via* the conduction band of TiO_2 , while the oxidised dye should be regenerated (directly or indirectly) by a water oxidation catalyst (cat_{ox}) to close the fuel-making cycle.

Although DSC and DSP are conceptually similar, their requirements and operational challenges are markedly different. In order to carry out water splitting, it is necessary to function in water, whereas most DSCs operate in organic solvents and are damaged by moisture. The transition from

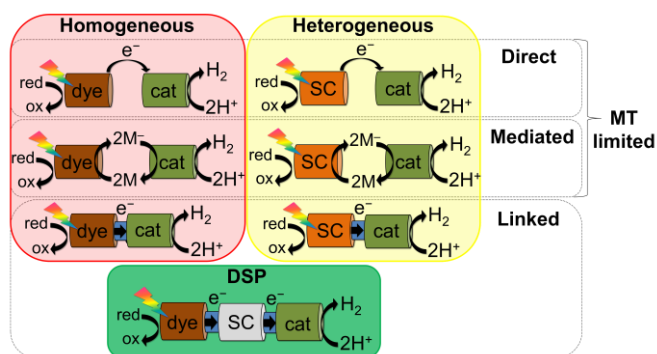


Fig. 1. Schematic representation of approaches to photocatalytic H_2 evolution, demonstrating electron transfer (ET) routes between a light absorbing dye and semiconductor (SC) and an electrocatalyst (cat): direct ET between the two components, mediated ET by a molecular redox pair ($2M/2M^-$), or direct ET when linked together. Without tethering, ET is mass transfer (MT) limited, whereas introduction of a linkage greatly enhances the electronic connection. Reduction of the oxidised dye following H_2 evolution closes the redox cycle. Dye-sensitised photocatalysis (DSP) bridges traditional homogenous and heterogeneous catalysis.

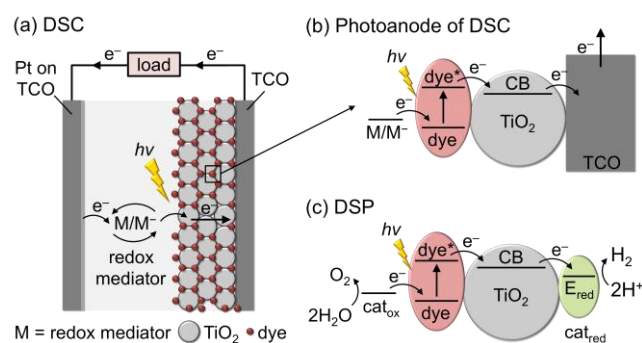


Fig. 2. Schematic representation of (a, b) dye-sensitised solar cells (DSCs) for *solar-to-electricity* conversion and (c) concept of dye-sensitised photocatalysis (DSP) for full water splitting. TCO = transparent conducting oxide on glass; cat_{ox} = water oxidation catalyst, cat_{red} = H_2 evolution catalyst, M/M^- = redox mediator.

organic to aqueous electrolytes is non-trivial, with issues arising from operational stability and changes at the interface between DSP components.

From an energetic perspective, DSP operation requires the energy levels of the photoexcited dye to be suitably placed with respect to the CB of the semiconductor and the redox potential of the mediator to drive electron transfer and to generate a current. This aspect is more challenging in DSP, where the thermodynamics of the water oxidation and proton reduction reactions must also be taken into account. This places upper and lower limits on the absolute energy levels of light-absorbers and mediators: for proton reduction for example, the CB of the semiconductor must be more reducing than the thermodynamic proton reduction potential ($E^\circ < 0$ V vs. the reversible hydrogen electrode, RHE).

Critically, catalysis in DSP is kinetically much more challenging than electron transfer (ET) and transport in DSCs. Whereas photovoltaic systems can operate by single ET processes, chemical fuel synthesis requires the transfer of multiple charges, since two electrons are consumed per molecule of H_2 and four holes are required per molecule of O_2 produced. These chemical reactions occur on timescales that are orders of magnitude slower than the light-induced charge generation. Therefore, in order to couple the multi-electron catalytic processes with the single-electron charge generation in the dye, photocatalytic systems must incorporate strategies to accumulate and prolong the lifetime of the charge carriers.¹⁰

In nature, long charge-carrier lifetimes are achieved by mitigating charge recombination through spatial separation of electrons and holes along electron transfer cascades. For example, in the water oxidation enzyme Photosystem II (PSII)¹¹ and the H_2 evolving enzyme known as hydrogenase,¹² electrons are transferred efficiently along several redox co-factors such as quinones (Fig. 3a) and iron-sulfur clusters from and to a catalytic centre deeply embedded in a protein matrix. To some extent, DSP mimics this cascade arrangement by spatially separating charge generation, transport, and catalysis (Fig. 3b). There is a great wealth of knowledge about principles of charge-separation and ET processes across the dye-semiconductor interface from DSC research; however, as a field, DSP is still in its infancy and the design principles are currently under development.

2.2 Component selection and design principles

The overall performance of multi-component systems such as DSP is determined both by the properties of each of the individual components (dye, semiconductor, catalyst, and sacrificial electron donor), and by the strength of the interactions between the components (anchoring groups) under the experimental conditions (pH, concentration, etc.). Therefore, several factors must be considered when selecting components for DSP.

2.2.1 Thermodynamic requirements. The most fundamental requirement of DSP is for an electron transfer cascade to be created, with the CB of the semiconductor positioned between the photoexcited state oxidation potential of the dye (E_{D^+/D^*}) and the onset of the catalytic reduction potential of the electrocatalyst (E_{red}). To complete the cycle, the potential for dye regeneration ($E_{D^+/D}$) must be further matched with either a sacrificial electron donor or a water oxidation (photo)catalyst. Although the relative alignment of the energy levels of each of the components is important for determining the feasibility of electron transfer, it cannot predict the overall activity of a particular system since many other factors may affect the efficiency such as light absorption, ET processes and catalytic efficiency.¹³

2.2.2 Dye. The required optical and electrochemical properties of the dye employed in DSP are very similar to that for DSCs, in that they must absorb light across a broad range of visible wavelengths with a high molar extinction coefficient, efficiently inject electrons into the conduction band of the semiconductor, and be stable for multiple reduction-oxidation cycles. These properties are well-characterised for a wide range of dyes used in DSCs and have been reviewed extensively.⁹ Due to the similarities of the properties required for the dye in these two systems, the classes of dyes most commonly employed in DSP are those that have shown the highest efficiencies in DSC, sometimes with slight modifications of their molecular structure. These include metal-based dyes, such as ruthenium polypyridine complexes (e.g. $[Ru(bpy)_3]^{2+}$ derivatives, Fig. 4a), porphyrins (Fig. 4b), and organic dyes, such as Eosin Y, donor- π -acceptor structures and chalcogenorhodamine dyes (Fig. 4c–e). More recently, alternative carbon-based materials such as carbon nitride¹⁴ and carbon quantum dots¹⁵ have been explored as sensitisers

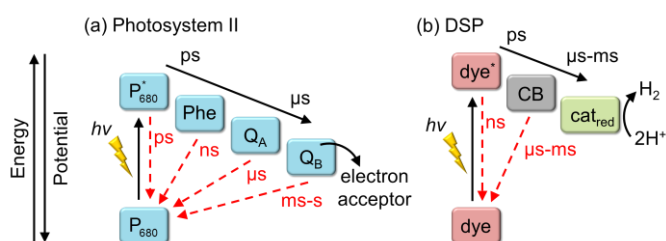


Fig. 3. Electron transfer cascades in (a) photosystem II and (b) DSP with separation of light absorption, electron transfer and catalysis. Forward electron transfer and recombination rates are indicated in black and red arrows, respectively. P_{680} : primary electron donor in the reaction centre of PSII photobleaching at $\lambda = 680$ nm; Phe: pheophytin (primary electron acceptor); Q_A : plastoquinone A; Q_B : plastoquinone B (terminal electron acceptor).

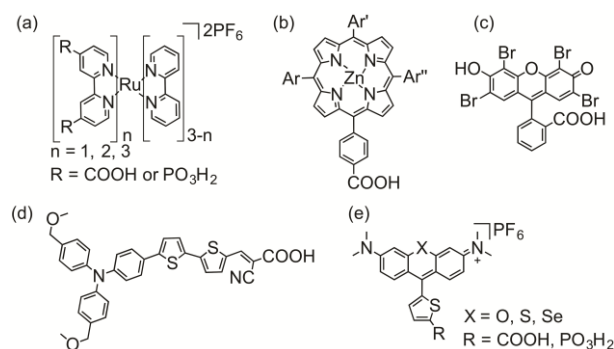


Fig. 4. Chemical structures of dyes employed in DSP (shown in fully protonated form): (a) Ruthenium bipyridine derivatives,¹⁶ (b) porphyrins (Ar = aryl groups),⁹ (c) Eosin-Y,¹⁷ (d) donor- π -acceptor type¹⁸ and (e) chalcogenorhodamine dyes.¹⁹

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for TiO₂ although their photophysical properties have been less well-explored.

Metal-based dyes can absorb light *via* a number of transitions that can either be metal-centred or involve metal-to-ligand or ligand-to-metal charge transfers (MLCT and LMCT, respectively).²⁰ For organic dyes, absorption is often due to π - π^* transitions. In general, organic dyes consist of less expensive, more abundant elements and have higher molar extinction coefficients than metal-based dyes. However, they are usually more prone to surface aggregation under DSP conditions, affecting their electron transfer properties.¹⁹ In addition, metal complexes are generally more photostable, due to the redox flexibility of their metal centre.²⁰

The structure of the dye can play a critical role in the efficiency of electron transfer in dye-sensitised systems. Positioning the lowest unoccupied molecular orbital (LUMO) close to the semiconductor surface allows for good overlap with the orbitals from the semiconductor, enhancing the rate of electron injection.²¹ Similarly, positioning the highest occupied molecular orbital (HOMO) of the resulting dye cation away from the surface can greatly reduce the rate of recombination following electron injection, enhancing the lifetime of the charge-separated state.²¹ This tailored synthetic design is exemplified in bpy-based Ru dyes and the donor- π -acceptor structure of some organic dyes, in which photoexcited electrons are driven towards the semiconductor surface.²²

2.2.3 Semiconductor. The primary role of the semiconductor in DSP is to chemically connect the molecular components and transfer electrons from the excited dye to the electrocatalyst. Therefore, in addition to the chemical composition, accessible surface area and stability of the semiconductor, the position of the CB, charge mobility, electron injection rates and orbital symmetry are all important factors.²³

Although narrow band gap semiconductors that can absorb visible light, such as Cu₂O and metal chalcogenides, have been used as light-absorbers for photocatalysis (without an additional dye),²⁴ these materials often suffer from photo-corrosion in water and their properties are often difficult to tune (Fig. 5, right). The most (photo)chemically stable

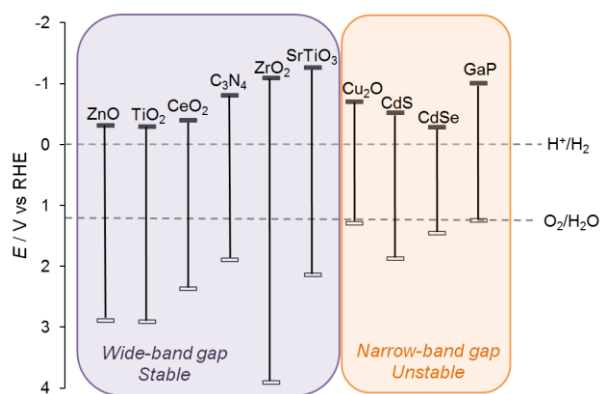


Fig. 5. Conduction band (CB) and valence band (VB) positions of a selection of semiconductors (in V vs. RHE).²⁵ The electrochemical standard potentials for water oxidation and proton reduction are shown for comparison.

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semiconductors are wide band-gap metal oxides that only absorb UV light (Fig. 5, left); whereas their absorption of such a small portion of the solar spectrum is limiting in the absence of a dye, in DSP, light absorption primarily occurs in the dye and therefore the valence band (VB) position is irrelevant. The use of a dye can not only avoid side reactions caused by highly reactive VB holes following direct band gap excitation, but also enables the use of semiconductors with greater stability.

TiO₂ is by far the most studied semiconductor for both DSC and DSP, and Degussa (Evonik) Aeroxide P25 TiO₂ is still one of the most active photocatalysts. P25 typically consists of micrometer sized agglomerates of nanostructured particles (~21 nm) with an 8:2 ratio of anatase:rutile crystal phases. The exact reason for the high performance of P25 over anatase or rutile alone is still under debate, but a heterojunction between the two phases might increase charge separation.²⁶ Although electron transfer can occur across a single particle, it has been proposed that nanoparticle agglomeration is beneficial to photocatalytic activity.²⁷ Particles within agglomerates form a conductive network to transport charges, increasing the spatial separation between the charge generation and catalytic sites, and reducing charge recombination.

2.2.4 Catalyst. The key characterisation parameters of a fuel-making catalyst are the onset potential for catalytic reduction (E_{red}) and the catalytic performance parameters such as turnover number (TON) and turnover frequency (TOF) at a given potential.²⁸ The onset potential and catalytic activity are typically probed by electrochemical techniques (without irradiation). Fig. 6a shows the theoretical electrochemical current-voltage (i - V) curves obtained for redox processes at equilibrium.²⁹ The difference between E_{red} and the equilibrium potential, E^0 , is referred to as the overpotential, η_{red} , which, ideally, will be small ("inactive" zone in Fig. 6a). For molecular species, electrochemistry also allows us to examine other redox processes, such as the individual reduction steps of a redox centre, as shown in Fig. 6b for a cobaloxime complex (**1**, Fig. 7). Since H₂ evolution requires the passage of 2 electrons, molecular complexes need to be capable of (and stable to) undergoing two reduction steps. Thermodynamically, a catalyst can only function on a semiconductor when the CB

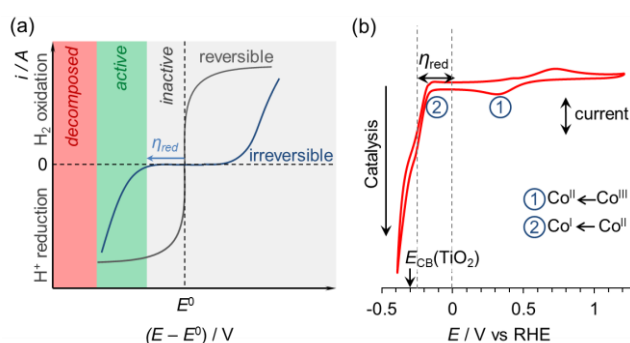


Fig. 6. (a) Theoretical cyclic voltammograms of reversible and irreversible catalysts in the presence of substrates, demonstrating the lack and requirement of a significant overpotential (η) in reversible and irreversible electrocatalysis, respectively. (b) CV of a molecular Co catalyst (**1**, see Fig. 7) demonstrating the key redox processes in pH 7 electrolyte solution.³⁰ The onset of the catalytic proton reduction current is observed following reduction of Co^{II} to Co^I.

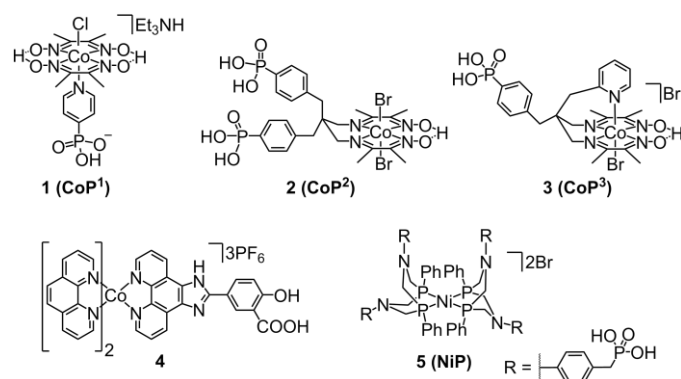


Fig. 7. Chemical structures of molecular cobalt and nickel catalysts employed in DSP for H₂ evolution (shown as isolated).^{8,17,30-32}

potential is more negative than its E_{red} (Fig. 6b, indicated by $E_{\text{CB}}(\text{TiO}_2)$) and a more negative potential (larger driving force) is expected to result in a faster catalytic rate ("active" zone in Fig. 6a). However, care must be taken at very negative potentials as some molecular catalysts are susceptible to decomposition under highly reducing conditions ("decomposed" zone in Fig. 6a).³³

Platinum and hydrogenases are both thermodynamically reversible electrocatalysts able to catalyse proton reduction and H₂ oxidation at a minimal η_{red} . However, hydrogenases are fragile, laborious to isolate and purify, and Pt is expensive and of low abundance, limiting practical scale-up of devices to commercial scale. In comparison, the development of molecular catalysts based on 3d transition metals has received much attention as potentially inexpensive and atom-efficient (all metal centres are involved in catalysis) alternatives. Most importantly, they are model materials to study structure-activity relationships, since their active centre is well-defined ('single-site catalysis') and tuneable.³⁴

Only a limited number of molecular H₂ evolution catalysts suitable for DSP are currently available which display good stability and activity in aqueous solution upon immobilisation on a semiconductor. The synthesis of complexes with suitable anchoring groups is often challenging and the introduction of the anchoring functionality can alter the electronic and steric properties of the catalyst significantly. Surface immobilisation also reduces the probability of interaction between two catalyst molecules and hence catalysts capable of operating *via* a mononuclear proton reduction mechanism are most likely preferred. The most notable examples of molecular catalysts for DSP are cobalt complexes of the cobaloxime family (1–3, Fig. 7),³⁰⁻³² a cobalt phenanthroline catalyst (4),¹⁷ and a nickel bis(diphosphine) complex (5, Fig. 7).⁸ The performance of these catalysts in DSP systems will be discussed in further detail below.

Much less is known about the catalyst-semiconductor interface and charge transfer compared to the dye-semiconductor interface. It has been demonstrated recently that the distance between the catalytic centre and the particle surface affects ET and recombination rates in DSP,³⁵ but more detailed studies are required to fully characterise and optimise these processes. In particular, for cobaloximes (*e.g.* 1),

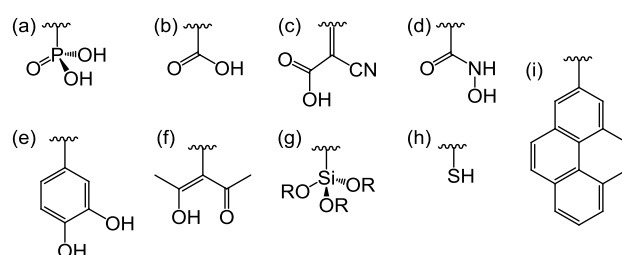


Fig. 8. Binding moieties for linkage of dyes and catalysts to semiconductor surfaces: (a) phosphonic acid, (b) carboxylic acid, (c) cyanoacrylic acid, (d) hydroxamic acid, (e) catechol, (f) acetyl-acetone, (g) silane, (h) thiol and (i) pyrene.

electron-donating axial groups are known to enhance activity,³⁶ but common anchoring groups (see below) are typically electronegative and may currently adversely affect the overall DSP activity.

2.2.5 Anchoring groups. Anchoring groups facilitate ET between the components and linkage of dyes and catalysts to the surface of the semiconductor is a critical requirement for productive DSP.¹³ The primary requirement for the anchor is that it has a strong affinity for the semiconductor surface under the experimental conditions, but the nature of the anchor will also influence the rate of interfacial ET. For metal oxides, oxygen-containing anchors such as phosphonic acids, carboxylic acids, cyanoacrylic acids, hydroxamic acids, catechols, acetylacetone, and silane linkages are commonly used (Fig. 8a–g).²³ In DSCs, carboxylate and cyanoacrylate groups are preferred as dye anchoring groups since they yield the fastest electron injection rates and are stable under moisture-free conditions;³⁷ however, carboxylate groups are easily displaced from a metal oxide surface in water above pH 4, reducing their applicability in DSP.¹⁶ The stability of these anchoring groups can be enhanced by the addition of hydrophobic side chains to the dye, reducing its solubility in water.³⁸ However, in DSP, these hydrophobic chains might also prevent water molecules from reaching the molecular catalyst. Phosphonate groups retain a high affinity for many metal oxide surfaces over a wider range of pH, which compensates for the relatively slower ET rate.^{16,23} Hydroxamate anchoring groups have been proposed as a compromise between phosphonate and carboxylate groups, with improved binding affinity over carboxylates and improved ET rates compared to phosphonates, but they have not been widely adopted.²³ Coupling *via* a silane group generates covalent bonds with the surface which are more resistant to hydrolysis than phosphonate or carboxylate esters; however, the deposition is difficult to control (often generating insulating multi-layers) and does not allow for dynamic self-assembly and disassembly of the supramolecular DSP system.

DSP systems to date have primarily focused on metal oxide semiconductors, but other classes of semiconductor can also be employed in which case other linkers may be more suitable. For example, thiol groups are commonly used as surface modifiers for metal chalcogenides (Fig. 8h). Non-metal semiconductors such as graphene and carbon nitride (C₃N₄) could be modified by pyrene anchors *via* π - π stacking (Fig. 8i).

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2.2.6 Sacrificial electron donor (SED). A wide range of SEDs have been used in photocatalytic proton reduction systems, including tertiary amines (e.g. ethylenediaminetetraacetic acid, triethylamine, triethanolamine), organic acids (e.g. ascorbic acid, lactic acid, oxalic acid, formic acid), alcohols, and inorganic ions such as sulfite.³⁹ SEDs are chosen for their ability to regenerate the oxidised dye and also have a role in buffering the reaction solution at a desired pH value. It is common that SEDs show their optimal performance at a pH close to their pK_a value and may therefore also have a role as proton relay in the photocatalytic scheme.⁴⁰ However, the choice of SED is often based on an initial screening approach as the performance is difficult to predict.

SEDs can also have secondary roles and care must be taken to ensure that the H_2 generated is not a product of SED oxidation,³⁹ and that reactive radicals generated from the SED decomposition do not interfere with the fuel making reaction.¹³ It has also been suggested that non-innocent interactions of the SED with the particle surfaces is a key factor for performance.³⁹ Additionally, the degree of interaction of the SED (and other additives) with the molecular catalyst and DSP system is not well understood.

Although SEDs are useful to establish the activity of the reductive half-reaction without limitations by challenging oxidation chemistry, SED oxidation is an energetically consuming rather than energy storing reaction and is therefore an unsustainable component of photocatalytic schemes. Demonstration of efficient DSP in the presence of a SED is therefore often not an effective predictor of function in full water splitting systems. As such, moving away from SED use and towards non-sacrificial systems is a key for future development in this field.

2.2.7 Solution pH. DSP should be operating in aqueous solution and the solution pH influences virtually all components in DSP; therefore, pH is an important parameter for activity. The pH impacts the redox and photophysical properties of the dye, the CB energy of the semiconductor,²⁵ the activity of the catalyst^{32,40} and the reducing ability of the SED.⁴⁰ In addition, the pH might also affect the affinity of the anchoring groups for the semiconductor¹⁶ and the overall stability of the self-assembled DSP system. The choice of experimental conditions for DSP operation must therefore be a delicate balance between optimisation of each of the individual components and the interactions between them.⁴⁰

3. Assembly and Evaluation of DSP Systems

3.1 Assembly of components

DSP systems can be self-assembled *in situ* or *ex situ*, depending on the solubility of each component and the degree of competition between the molecular components for binding to the semiconductor nanoparticle surface. If all components are water-soluble and bind quantitatively to the semiconductor, the DSP system can be assembled *in situ* in the photoreactor by addition of the molecular catalyst and dye to a suspension of the semiconductor nanoparticles (Fig. 9).^{8,30} A

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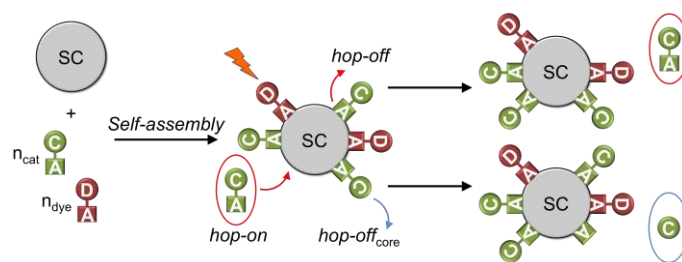


Fig. 9. Self-assembly of DSP system and *hop-on/hop-off* mechanism (shown for catalyst only); *hop-on*: Unbound catalyst molecules occupy free surface sites through self-assembly; *hop-off*: Desorption of catalyst from the particle surface; *hop-off_{core}*: Loss of catalyst core from the particle surface (anchoring group stays on surface). SC = semiconductor; C = catalyst; D = dye; A = anchoring group.³⁰

monolayer or sub-monolayer concentrations of the molecular catalyst and dye should be employed, to avoid displacement of the weaker binding component and to maintain a predictable dye to catalyst ratio on the nanoparticle surface. Water-insoluble components can be pre-immobilised in an organic solvent (*ex situ*), followed by centrifugation and dispersion of the modified semiconductor in the aqueous solution.¹⁷

Information about the loading capacity of a semiconductor and anchoring strength of the molecular species to the semiconductor surface can be obtained by electronic absorption (UV-vis), atomic absorption (AAS) and atomic emission (AES) spectroscopies as well as inductively-coupled plasma mass spectrometry (ICP-MS). The amount of adsorbed molecules can be quantified either by comparison of dye and catalyst solutions before and after exposure to the semiconductor (i.e. absorption in UV-vis or concentration in AAS/AES/ICP-MS) or by deliberate desorption of immobilised molecules from the particle surface.

If the molecules are not quantitatively anchored on the surface during the *in situ* self-assembly, unbound molecules can be in dynamic equilibrium with bound species, where a catalyst or dye molecule can adsorb and desorb from the semiconductor surface (*hop-on/hop-off* mechanism, Fig. 9).³² In this manner, not all catalyst and dye molecules may be in contact with the semiconductor at a given time, but can contribute to photocatalysis in the course of the bulk experiment. This can be verified by removing the unbound molecules from the particle suspension *via* centrifugation.^{8,32} The modified particles are studied for H_2 evolution after re-suspension in fresh media and a reduced activity is expected with reduced loading of the catalyst.

The potential-dependent binding ability of the catalyst can be investigated electrochemically when immobilising it on a conductive metal oxide surface (Fig. 10a and b).^{31,32} Cyclic voltammetry and chrono-amperometry allows to resolve the stability in different redox states and immobilisation on a transparent substrate provides a means to study the catalyst spectroelectrochemically. The latter allows investigation of the molecular integrity and nature of the catalyst in distinct oxidation states through the oxidation-state-dependent absorption profile (Fig. 10c).

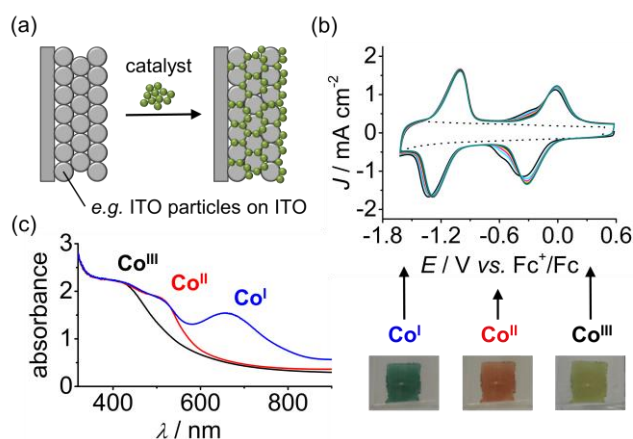


Fig. 10. (a) Schematic representation for integration of a molecular catalyst with a mesoporous indium tin oxide (ITO) electrode. (b) Electrochemistry allows for testing anchoring stability: Consecutive CVs of **2** on mesoporous ITO electrodes (in DMF/TBAPF₄ electrolyte solution). The dotted trace represents the electrode background. (c) Spectroelectrochemistry of **2** on mesoporous ITO allows for evaluation of the electronic absorption spectra of redox states relevant for catalysis. The photographic pictures depict the colour change of the electrode at different redox potentials.^{31,32}

3.2 Experimental setup

The primary means of characterising the overall bulk activity of a photocatalytic system is to irradiate a solution of the photocatalyst and quantify the photo-generated products. Although inexpensive white light sources or light emitting diodes can be used for preliminary experiments, a solar light simulator with an air mass 1.5 (AM 1.5G) filter (the value of 1.5G represents the overall yearly average solar spectrum on the Earth surface, with an incident angle of sunlight of 48.2°) and 1 Sun intensity (100 mW cm⁻²) should ultimately be employed as a reproducible global standard. Additional filters such as a water filter to exclude IR irradiation help to prevent the reactor from overheating, and other filters may be used to avoid direct excitation of the semiconductors or to study visible light activity only ($\lambda > 400$ nm). Monochromatic light irradiation (for example using a bandpass filter or monochromator) is needed for quantum efficiency measurements and to record action spectra.

DSP systems should be kept in the dark during assembly and be deaerated prior to testing, as light and O₂ can degrade the components or interfere with processes in DSP. The photoreactor with the DSP suspension should be placed into a temperature-controlled water bath to maintain a defined temperature (experiments are typically conducted around ambient temperature) and stirred to ensure mixing of all components throughout the experiment. Photoreactor sizes and designs can vary from small vials to large round bottom flasks with illumination from the front, bottom or with samples arranged around the light source.

Gas chromatography (GC) is commonly used to quantify the amount of H₂ in the headspace of the photoreactor. Other quantification methods include mass spectrometry, volumetric detection or pressure monitoring combined with analysis of the gas mixture by GC or Clark-type electrodes for H₂

detection. Regular sampling or *on-line* measurements provide information about the reaction kinetics.

3.3 Key performance parameters and optimisation

The overall amount of H₂ produced and long-term activity as well as the TON (eq. 1), TOF (eq. 2) and (external) quantum efficiency (EQE, eq. 3) are important key performance parameters that need to be assessed in DSP.

$$\text{TON (mol H}_2 \text{ per catalytic site)} = \frac{n(\text{H}_2)}{n(\text{catalytic site})} \quad \text{eq. 1}$$

$$\text{TOF (mol H}_2 \text{ per catalytic site per time)} = \frac{n(\text{H}_2)}{n(\text{catalytic site}) \cdot t_{\text{irr}}} \quad \text{eq. 2}$$

$$\text{EQE (\%)} = \frac{n(\text{H}_2) \cdot N_A \cdot h \cdot c \cdot 2}{\lambda \cdot t_{\text{irr}} \cdot I \cdot A} \quad \text{eq. 3}$$

with $n(\text{H}_2)$ = moles of H₂ [mol] and $n(\text{catalytic site})$ = moles of catalyst or dye [mol] or amount of semiconductor [g], N_A = Avogadro constant [mol⁻¹], h = Planck constant in [kg m² s⁻¹], c = speed of light [m s⁻¹], λ = wavelength [m], t_{irr} = irradiation time [s], I = light intensity [W m⁻²] and A = irradiated area [m²].

The main parameters that can be varied in order to optimise the key performance parameters of a DSP system are temperature, pH, nature and concentration of catalyst, dye, semiconductor or SED, as well as the catalyst/dye, catalyst/semiconductor or dye/semiconductor ratios.

3.4 Characterisation

3.4.1 Establishing the mechanism in DSP. Control experiments are needed to establish the mechanism in the DSP assembly. Experiments in the absence of each of the system's components (*i.e.* without molecular catalyst, dye, semiconductor, light or SED) are essential to confirm the requirement of all components for photocatalytic formation of H₂ and to disregard other possible mechanisms such as electron transfer directly from the dye to the catalyst in homogenous solution or non-light activated processes.

To confirm the importance of anchoring of the molecular catalyst and dye to the semiconductor for performance, DSP experiments with anchor-free derivatives or in the presence of a reagent that competes for binding sites on the semiconductor (*e.g.* excess phosphate anions to displace phosphonated molecules) can be performed.¹³ The importance of electron transfer through the semiconductor particle can be qualitatively evaluated by using a semiconductor with a CB more negative than the excited state oxidation potential of the dye ($E_{\text{CB}} < E_{\text{D}^+/\text{D}^*}$).^{8,41} It should be noted in this context that the close proximity of catalyst and dye on the semiconductor surface could also have a beneficial effect on catalysis such as avoided diffusion-limited kinetics.⁸

3.4.2 Activity limiting component. Long-term studies can be used to probe the overall stability of the system, and as a method to assay the nature of the stability-limiting component. Often, the photocatalytic activity decreases or even ceases after prolonged illumination. A first indication on the limiting species can be gained from supplying fresh dye,

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catalyst, semiconductor or SED to the deactivated system. Recovery of activity upon addition of an individual component indicates that its role is stability-limiting.⁴¹

3.4.3 Proof of molecular integrity. Molecular catalysts are often prone to decomposition under a range of different conditions and it is therefore important to evaluate if the origin of catalytic activity is from a molecular species or a decomposition product such as metal-containing nanoparticles formed during photocatalysis.³³ Reports on methodologies to identify the nature of the active species in photo- and electrocatalytic H₂ evolving systems containing molecular catalyst have been summarised and discussed recently.⁴²

UV-vis and IR spectroscopic analyses before, during and after photocatalysis can be used to monitor any change of the molecular catalyst. Energy-dispersive X-ray spectroscopy (EDX), transmission and scanning electron microscopy (TEM, SEM) provide some information about elemental composition of the modified particle surface and allows for the detection of metal nanoparticles on the semiconductor surface, respectively.⁴² The oxidation states of elements on the semiconductor surface can be best determined by X-ray photoelectron spectroscopy (XPS).³¹ Other X-ray absorption techniques such as EXAFS (extended X-ray absorption fine structure) and XANES (X-ray absorption near edge structure) provide further insights into the local environment of a single element, *e.g.* to determine the nature and number of neighbouring atoms.

Control experiments with metal salts instead of the molecular catalyst help to rule out the potential formation of active species such as metal-SED complexes or metal-containing deposits *via* the reduction of metal ions.³² In addition to anchoring stability, (spectro)electrochemical studies of catalyst-semiconductor hybrid electrodes also provide information about the molecular integrity under reducing conditions (see above).³¹ Finally, if the activity of a degraded DSP system can be reactivated by supplying fresh catalyst or ligand, the active species is likely to be molecular.⁴¹

4. Electron Transfer Kinetics

Kinetic analyses provide supplementary and quantitative information about charge separation and competing recombination pathways, which could limit the performance of the DSP system. The two main time-resolved spectroscopic techniques used to characterise the ET kinetics of hybrid systems for artificial photosynthesis are transient absorption spectroscopy (TAS) and time correlated-single photon counting (TC-SPC).

4.1 Transient absorption spectroscopy (TAS)

TAS is a technique that provides information about the dynamics of short-lived transient species that absorb in the visible, near-infrared and infrared region. TAS is a pump-probe experiment, in which a sample is excited with a short intense pulse of light (pump) generated by a pulsed laser, and the resulting changes in absorption (characterised as optical density, $\Delta O.D.$) are monitored through the transmission of a

second light source (probe) as a function of time (Fig. 11). Instruments to monitor transient species can be assembled in a range of configurations, to provide information over a wide range of wavelengths and timescales (fs – s).

TAS measurements show two main types of contributions to the change in optical transmission: positive signals correspond to the absorption of transient species formed after photoexcitation, while negative signals are generated upon the disappearance of the ground state absorption of species. It is worth noting that if the excited state is luminescent, the probe may trigger stimulated emission, which is recorded as a negative optical density change.

For the study of photocatalytic systems for H₂ production, it is important to monitor (1) the spectrum of transient species relevant to the catalysis over a broad range of wavelength at a single time delay (Fig. 12a), and (2) the dynamics of these transient species over a broad range of time delays at a single wavelength (Fig. 12b). The former experiment provides information on the absorption spectra of transient species, allowing the identification of the nature of the species involved in the photocatalytic process; and the latter measurement allows the characterisation of the timescale of formation and disappearance of transient species. Transient absorption decays provide information on the ET kinetics that take place in DSP. The most common parameter used to characterise the charge-carrier lifetime is $t_{50\%}$, defined as the time required for the signal amplitude to decrease half of its initial value (Fig. 12b).

TAS has been used to study the relaxation, recombination and ET reactions of the photogenerated charge carriers in a

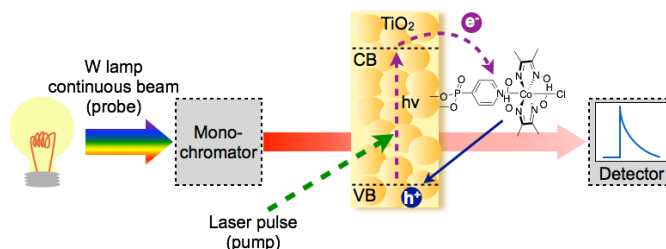


Fig. 11. Basic TAS setup. Dashed purple lines represent the ETs photoinduced by the laser pulse, whereas the solid blue arrow represents the recombination reaction and both are being monitored by the probe.

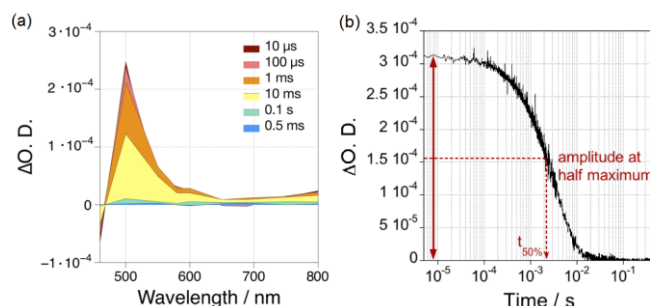


Fig. 12. (a) Example of a transient absorption spectrum measured over a broad range of wavelengths at different time delays (10 μ s to 0.5 ms); (b) Example of a transient absorption decay recorded at a single wavelength over a broad range of time delays. Examples were taken from ref [40] and show the TA spectrum and decay of RuP (4 μ M) in 0.1 M ascorbic acid solution (pH 4.5) after excitation at $\lambda = 355$ nm.⁴⁰

variety of photochemical systems, including natural photosynthetic reaction centres, molecular donor-acceptor complexes⁴³ and molecular-based devices^{8,21,41} for applications in solar cells and solar fuels. In hybrid systems for H₂ evolution, the dye, semiconductor and catalyst can all be studied. The ET kinetics can be monitored at different experimental conditions, such as pH, excitation densities, and dye/catalyst loading ratios. This kinetic information, when correlated with bulk catalytic experiments, can highlight catalytic bottlenecks in DSP and provide insightful mechanistic information.

4.1.1 Dye. Typically, in a dye-sensitised system, the molecular photosensitiser absorbs light and a short-lived molecular excited state is formed (dye*). Electron injection into the semiconductor conduction band can happen in the fs-ps timescale with very high injection yields, with some examples even reaching unity.⁴¹ The formation and lifetime of the resulting dye cation (dye⁺) can also be observed by TAS. For example, RuP* (RuP: n=1, R=PO₃H₂, Fig. 4a) absorbs at λ = 450 – 700 nm and RuP⁺ absorbs at λ ~700 nm.⁴¹

4.1.2 Semiconductor. In a semiconductor, TAS allows the dynamics of charge carriers (electrons and holes) to be determined. In many semiconductors, and in particular in metal oxides such as TiO₂, such charge carriers undergo rapid (fs-ns) relaxation/trapping processes, resulting in the generation of relatively localised 'trapped' electrons and holes.⁴⁴ Such relaxation processes are often in kinetic competition with ultrafast charge carrier recombination. For example, trapped electrons in anatase TiO₂ have been assigned to the reduction of titanium (IV) atoms (Ti⁴⁺ + e⁻_{CB} → Ti³⁺; λ > 700 nm), whereas hole traps have been attributed to O⁻ species (O²⁻ + h⁺_{CB} → O⁻; λ = 460 nm).⁴⁴ The recombination dynamics of such trapped electrons and holes can extend out to the micro- and millisecond timescales, often rate limited by thermally activated detrapping processes.

ET transfer from a photoexcited dye to the semiconductor CB can also be studied by monitoring the CB electron signal of the semiconductor instead of the dye species. However, in some cases the dye species might mask the signals from the semiconductor charge carrier.⁴⁵ To study ET from a semiconductor to a catalyst, direct band-gap excitation is typically used. For reductive catalysts, a short-lived transient absorption decay of the electronic signal in the presence of a hole scavenger is then indicative of a fast and quantitative ET to the catalyst (Fig. 13).^{35,41,45}

4.1.3 Catalyst. We can characterise two main types of optical signal of molecular catalysts using TAS. (1) The kinetics of disappearance of the ground state absorption, resulting from reduction of the catalyst, will appear as a negative signal (bleach) if the catalyst absorbs in the monitored region.⁴⁰ (2) The kinetics of formation and reactivity of catalytically active species that absorb in the visible region, such as different oxidation states (e.g. Co catalysts)⁴⁶ might appear as positive transient absorption bands. In addition to this, IR-active species can be monitored using TAS in order to study reaction mechanisms. However for water splitting, this technique has

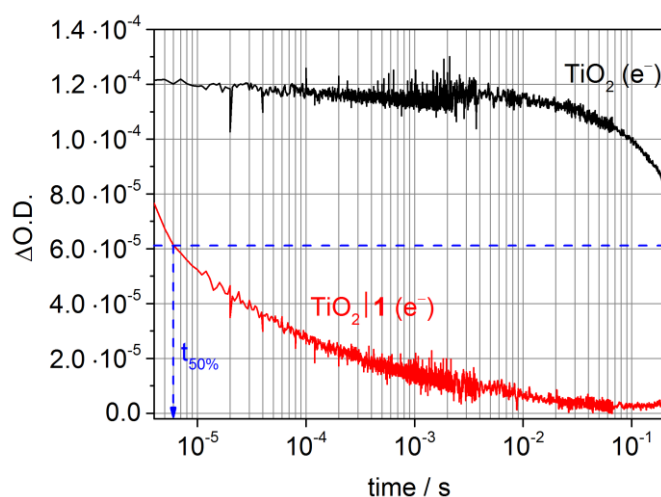


Fig. 13. Transient absorption decay of photoexcited TiO₂ CB electrons after excitation at λ_{ex} = 355 nm (λ_{probe} = 900 nm) of a bare TiO₂ film (black trace) and when modified with the cobalt catalyst **1** (red trace; see Fig. 7) in the presence of TEOA (0.1 M, pH 7) as sacrificial donor. The blue arrows indicate t_{50%} for the first reduction of the catalyst (Co^{III}/Co^{II}).^{35,41}

an added difficulty due to the broad absorption of water molecules in the IR.

4.2 Time correlated-single photon counting (TC-SPC)

TC-SPC is an optical technique that is mainly used to determine the lifetime of molecular excited states, as well as the kinetics of charge/energy transfer in molecular homogeneous or heterogeneous systems, where the excited state of the chromophore is luminescent.⁸ The operational basis is the detection of single emitted photons as a function of time, typically in the ps – μs timescale. With this technique, the lifetime of the photosensitiser can be measured both in solution or anchored onto the surface of mesoporous nanocrystalline film such as Al₂O₃, ZrO₂ or TiO₂.^{8,41}

In an inert environment, the luminescence of a molecular dye typically decays within nanoseconds. However, when it is in solution with an easily reducible molecule, or is attached to the surface of a semiconductor with a sufficiently low CB, the luminescence decay of the dye can be quenched due to electron or energy transfer. Therefore, TC-SPC provides information of the timescale at which the electron or energy transfer takes place. In addition, a quantitative measurement of the electron/energy transfer yield can be obtained by integration of the area under the decay, and compared against steady state photoluminescence quenching measurements.⁴¹

4.3 Electron transfer kinetics in DSP

Fig. 14 summarises the typical values of electron transfer kinetics of DSP.^{8,41,45} The overall kinetic requirement in the design of these photocatalytic systems is the matching of the lifetime of the photogenerated charge carriers with the rate of catalytic proton reduction. This challenge can only be achieved by promoting fast and quantitative ETs (solid blue arrows), and by minimising and slowing down the recombination reactions (dashed red arrows). ET across the dye-semiconductor

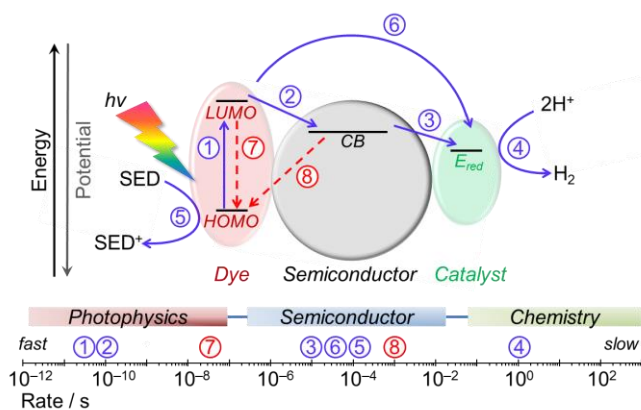


Fig. 14. Main kinetic processes considered in the design of DSP systems. Photocatalysis-promoting (forward) electron transfer processes are shown in blue (solid arrows) and recombination pathways in red (dashed arrows). Typical rates determined for **RuP**|TiO₂ modified with molecular catalysts are indicated on the scale bar.^{8,35,41,45} 1: photoexcitation of dye; 2: electron transfer from photoexcited dye to the semiconductor-CB; 3: electron transfer from semiconductor to the catalyst (rate indicates $Co^{III} \rightarrow Co^{II}$ reduction process); 4: catalytic reduction of protons to H₂; 5: regeneration of the oxidised dye by the sacrificial electron donor; 6: direct electron transfer from photoexcited dye to the catalyst; 7: lifetime of photoexcited state of the dye; 8: charge recombination between semiconductor-CB and the oxidised dye.

interface has been found to increase charge carrier lifetimes up to the hundreds of microseconds timescale.

Molecular and device design can be used to control the rate of electron transfer (k_{ET}) at the dye-semiconductor and semiconductor-catalyst interfaces. According to Marcus theory, the rate of non-adiabatic ET is given by:

$$k_{ET} = \frac{2\pi}{\hbar} \frac{H_0^2 e^{-\beta r}}{\sqrt{4\pi\lambda k_B T}} e^{-\left(\frac{\Delta G^0 + \lambda}{4\lambda k_B T}\right)^2} \quad \text{eq. 4}$$

where $H_0^2 e^{-\beta r}$ corresponds to the electron tunnelling through a potential barrier and is dependent on the spatial separation of the semiconductor and the molecular component (r). λ is the reorganization energy, mainly dependant on the polarity of the solvent, and ΔG^0 is the reaction free energy defined as the difference between the molecule's redox potential ($E_{red/ox}$) and the quasi-Fermi Level (E_F) of the semiconductor. Control of these parameters, especially ΔG^0 and r , through rational design of dyes and catalyst can lead to enhanced DSP efficiencies.³⁵ The optimisation of the ΔG^0 in DSP, where two ETs are required to produce H₂, is known to be critical. For most molecular reduction catalysts, the second ET has a lower driving force compared to the first one. This lower value of ΔG^0 for the second reduction has been found to be the main limiting step, highlighting the importance of synthetic control of the redox potential in improving catalytic yields. It is worth noting that recombination reactions are slower than the forward ETs, despite having a larger driving force. This is explained by their values of ΔG^0 lying in the Marcus' normal (forward ET) and inverted (recombination) regions.

5. DSP for H₂ Evolution

5.1 Thermodynamically reversible catalysts

Platinum¹⁶ and hydrogenases⁴⁷ were among the first catalysts applied in DSP for H₂ evolution, demonstrating the feasibility of the concept and establishing performance benchmarks.

5.1.1 Platinum. Pt deposited on TiO₂ in combination with Ru(bpy)₃²⁺ dyes (Fig. 4a) containing different anchoring groups was studied for H₂ evolution activity.¹⁶ The best performance was obtained with the doubly phosphonated Ru(bpy)₃-dye **RuP** ($n=1$, $R=PO_3H_2$, Fig. 4a) achieving a TOF_{cat} of 104 h⁻¹ at pH 3 with ethylenediaminetetraacetic acid as SED (Table 1). Pt is often used in such model studies due to the excellent ability to easily accumulate multiple electrons and its high activity towards proton reduction. Since the noble metal catalyst is directly deposited on the semiconductor surface, forward ET is fast and efficient and can be eliminated as a performance limiting factor.

5.1.2 Hydrogenases. A range of hydrogenases featuring different H₂ evolution activities and O₂ tolerances have been studied in DSP.⁴⁷ A [NiFeSe]-hydrogenase from *Desulfomicrobium baculatum* on **RuP**-sensitised TiO₂ with triethanolamine (TEOA) as SED was identified as the best system giving a TON_{cat} > 600,000 and TOF_{cat} = 180,000 h⁻¹ under visible light irradiation (Table 1). This particular enzyme is known as a 'titanophilic' hydrogenase and is strongly bound to TiO₂ in an electroactive state – presumably with a high concentration of protein surface-exposed carboxylates from aspartate and glutamate residues – facilitating efficient electron transfer from the TiO₂-CB to the active centre of the enzyme.⁴⁷ Recently, a DSP system consisting of [NiFeSe]-hydrogenase and carbon nitride-sensitised TiO₂ was reported. A TON_{catalyst} of > 580,000 was achieved under full solar spectrum irradiation setting a new benchmark for noble and toxic metal-free DSP for H₂ evolution (Table 1).¹⁴

5.2 Molecular catalysts

During the last few years, the concept of DSP has been extended from reversible 'model' catalysts to synthetic molecular H₂ evolution catalysts. The first DSP system with an anchored molecular catalyst was reported in 2011 and consists of TiO₂ modified with the phosphonated dye **RuP** and cobaloxime catalyst **1** (Fig. 7).³⁰ Subsequently, several other cobalt and nickel complexes were employed in DSP (**2–5**, Fig. 7)^{8,17,31,32} and the activity and mechanism of these systems will be discussed below. Other examples of molecular catalysts integrated with light absorbing SCs (in the absence of dyes) have been reviewed elsewhere.⁴⁸

5.2.1 Molecular cobalt catalysts in DSP. Linkage of the cobaloxime catalyst **1** (Fig. 7) to TiO₂ was achieved *via* introduction of a phosphonic acid group on the axial pyridine ligand of the complex. Since both dye and catalyst are soluble under the experimental conditions, the DSP system could be easily self-assembled *in situ*. H₂ evolution by **RuP**|TiO₂|**1** was observed under visible light irradiation ($\lambda > 420$ nm) from pH neutral aqueous solution containing TEOA as SED (black trace,

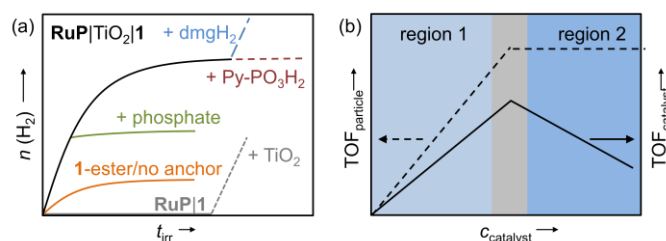


Fig. 15. (a) Representation of the activity of **1** on RuP-sensitised TiO₂ (black trace) and a range of important control experiments; (b) 'per particle' (TOF_{particle}) and 'per catalyst' activity (TOF_{catalyst}) as a function of catalyst concentration; Region 1: catalyst-dependent zone; Region 2: light-dependent zone.^{13,30,41}

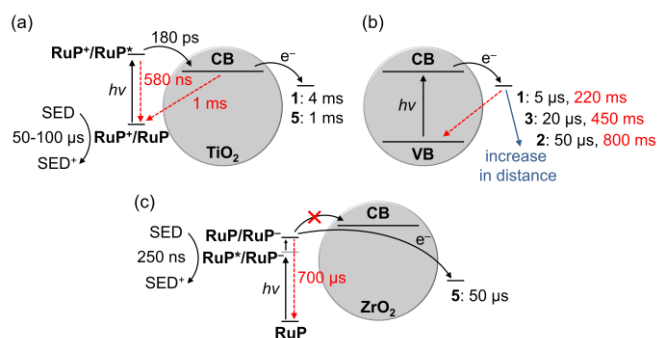


Fig. 16. Electron transfer (black arrows) and recombination (red arrows) rates in DSP systems: (a) RuP-sensitised TiO₂ in combination with catalyst **1** or **5**; (b) Distance dependent kinetics determined for catalysts **1–3** on photoexcited TiO₂; (c) RuP-sensitised ZrO₂ with **5** as catalyst. Electron transfer and recombination rates are given as $t_{50\%}$ and were determined for the first reduction of the catalysts only (**1–3**: Co^{III}/Co^{II}; **5**: Ni^{II}/Ni^I).^{8,35,41,45} It should be noted that electron transfer rates from the TiO₂-CB to the catalyst depend on the electron density in the TiO₂-CB, resulting in faster rates in (b) when TiO₂ is directly excited across the bandgap compared to (a) where electrons are supplied by the excited dye.⁴⁵

Fig. 15a) and an optimised TON_{cat} of 108 ± 9 and TOF_{cat} of $15.0 \pm 0.8 \text{ h}^{-1}$ were achieved after 10 h of irradiation (Table 1).

Surprisingly, the homogeneous multicomponent system RuP|**1** was found to be inactive (Table 1), even in the presence of the redox mediator methyl viologen but could be activated by the addition of TiO₂ (Fig. 15a, grey trace).¹³ Activity of the DSP system decreased after several hours of irradiation and could be fully restored by providing new dimethyl glyoxime (dmgH₂) ligand (Fig. 15a, blue trace). In this work, degradation of the core ligand was found to be the major deactivation pathway limiting the long-term performance of this DSP system to a few hours. In addition, it also provides further evidence that a molecular species instead of a cobalt-containing deposit was the active H₂ evolution catalyst.^{13,41} Photocatalytic H₂ evolution ceased in the presence of phosphate ions (Fig. 15a, green trace) and was significantly lower when the ester or anchor-free derivatives of **1** were used as catalysts (Table 1 and Fig. 15a, orange trace), demonstrating that linkage of the complex to the TiO₂ particle and electronic communication *via* the phosphonic acid moiety was vital for good DSP performance.^{13,41} When using ZrO₂ as semiconductor no H₂ was observed, ruling out direct electron transfer from the excited dye to the catalyst on the particle

(note that electron injection from RuP* into the ZrO₂-CB is not possible due to a too negative CB-edge).⁴¹

When optimising for the maximum H₂ evolution rate (TOF) of RuP|TiO₂|**1**, two distinct regions were identified (Fig. 15b).¹³ At low catalyst concentrations, the 'per catalyst' and 'per particle' TOF (TOF_{catalyst} and TOF_{particle}, respectively) increased with the concentration of **1** and the overall DSP performance was limited by H₂ evolution catalysis rather than light absorption (region 1, catalysis-dependent zone). At a higher concentration of **1**, light absorption by the RuP-sensitised TiO₂ became the limiting factor (region 2: light-dependent zone) resulting in a constant TOF_{particle} and decrease of TOF_{catalyst}.

Kinetic investigations of RuP|TiO₂|**1** under DSP conditions revealed a quantitative and ultrafast electron injection from RuP* to the TiO₂-CB (Fig. 16a).⁴¹ Charge recombination occurred on the ms timescale which is significantly slower than in analogous homogeneous systems (ms vs. ps/ns timescale).⁴³ In the presence of a SED, the lifetime of TiO₂-CB electrons was further extended to about 0.5 to 1 s (Fig. 13). These long-lived electrons are capable of reducing **1** on the timescale of a few milliseconds (Co^{III} → Co^{II}, Fig. 16a).⁴¹

Despite the notable faradaic efficiency of **1** (> 60 %) at low overpotential (−0.7 V vs. NHE at pH 7, $\eta = 290 \text{ mV}$),³⁶ the overall EQE of RuP|TiO₂|**1** is comparably low ($1 \pm 0.2 \%$).⁴¹ This was mainly attributed to slow kinetics for the second reduction (Co^{II} → Co^I) of **1**, being several orders of magnitude slower than the first reduction (Co^{III} → Co^{II}).⁴⁵

It is well known that the axial pyridine ligand in cobaloximes becomes labile after initial reduction and dissociates from the Co(dmgH₂)₂ core.⁴³ Similarly, this dissociation can occur in the particle hybrid (Fig. 9, *hop-off*_{core}). Even though the core can readily re-coordinate to the anchored ligand in the particle suspension, a more permanent attachment would in general be preferable.

The molecular cobalt catalyst **2** (Fig. 7) was designed to provide a more stable linkage to the TiO₂ surface *via* the introduction of phosphonic acid anchoring groups to the more robust equatorial diimine-dioxime ligand.³¹ Spectro-electrochemical studies of **2** immobilised on mesoporous ITO electrodes proved a stronger binding to the metal oxide surface than **1** and high stability of the complex under reducing conditions.^{31,32} However, no H₂ was observed for **2** on RuP|TiO₂ most likely due to a mismatch of the TiO₂-CB edge ($E_{\text{CB}} = -0.7 \text{ V vs. NHE at pH 7}$) and the required potential for proton reduction E_{red} of **2** in pH 7 TEOA buffer ($E_{\text{red}} = -0.72 \text{ V vs. NHE}$).³² Introduction of an axial pyridine ligand into the coordination sphere (**3**, Fig. 7) lowered the required η_{red} for H₂ evolution by about 70 mV and an optimised TON_{cat} of 22.0 ± 1.5 and TOF_{cat} of 9.7 ± 0.4 were achieved with RuP|TiO₂|**3** in pH 7 TEOA buffer (Table 1). The molecular catalyst **3** was identified as the lifetime limiting component in this DSP system, with photocatalysis ceasing after about 4 h of irradiation.

In a comparative kinetic study of catalyst **1–3** on TiO₂, ET and recombination rates between the TiO₂-CB and the catalyst

Table 1. Summary of colloidal DSP systems for light-driven H₂ evolution.

Entry	Catalyst	Photosensitiser	λ^a / nm	Electron donor ^b	pH	TON _{cat} ^c (t _{irr})	TOF _{cat} ^d / h ⁻¹ (1 h)	Ref.
1	Pt	RuP TiO ₂	> 420 ^g	EDTA	3	≈ 135 ⁱ (3 h)	≈ 104 ⁱ	[16]
2	[NiFeSe]-hydrogenase	RuP TiO ₂	> 420 ^h	TEOA	7	(6 ± 0.5)·10 ⁵ (4 h)	(1.8 ± 0.1)·10 ⁵	[47]
3	[NiFeSe]-hydrogenase	C ₃ N ₄ TiO ₂	> 300 ⁱ	EDTA	6	> (5.8 ± 0.6)·10 ⁵ (72 h)	(2.8 ± 0.3)·10 ⁴	[14]
4	1	RuP TiO ₂	> 420 ^j	TEOA	7	108 ± 9 (10 h)	15.0 ± 0.8	[13]
5	1-ester ^e	RuP TiO ₂	> 420 ^j	TEOA	7	13.4 ± 0.7 (4 h)	4.6 ± 0.2	[13]
6	1 w/o anchor ^f	RuP TiO ₂	> 420 ^j	TEOA	7	11.4 ± 1.3 (4 h)	3.4 ± 0.2	[13]
7	1	RuP	> 420 ^j	TEOA	7	– ^m (4 h)	– ^m	[13]
8	1	RuP ZrO ₂	> 420 ^j	TEOA	7	– ^m (4 h)	– ^m	[41]
9	2	RuP TiO ₂	> 420 ^j	TEOA	7	– ^m (4 h)	– ^m	[32]
10	3	RuP TiO ₂	> 420 ^j	TEOA	7	22.0 ± 1.5 (4 h)	9.7 ± 0.4	[32]
11	4	Eosin-Y TiO ₂	> 420 ^k	TEA	11	90 (3 h)	≈ 22	[17]
12	4	Eosin-Y	> 420 ^k	TEA	11	8 (6 h)	4	[17]
13	5	RuP TiO ₂	> 420 ^j	AA	4.5	278 ± 19 (30 h)	72 ± 5	[8]
14	5	RuP ZrO ₂	> 420 ^j	AA	4.5	524 ± 36 (30 h)	92 ± 26	[8]

^a Irradiation wavelength; ^b EDTA = ethylenediaminetetraacetic acid, TEOA = triethanolamine, TEA = triethylamine, AA = ascorbic acid; ^c Turnover number (TON_{cat}) is given as mol H₂ per mol catalyst; ^d Turnover frequency (TOF_{cat}) is given as TON_{cat} h⁻¹ based on the first hour of irradiation; ^e Phosphonate ester derivative of **1**: [CoCl(dimethylglyoximate)₂(diethylpyridyl-4-phosphonate)]; ^f Anchor-free derivative of **1**: [CoCl(dimethylglyoximate)₂(pyridine)]; ^g 450 W; ^h 45 mW cm⁻²; ⁱ 100 mW cm⁻²; ^k 300 W; ^j TON_{cat} and TOF_{cat} were calculated assuming an active site consists of at least two Pt atoms; ^m No detectable amounts of H₂ being produced.

and singly reduced species (Co^{II}) and the TiO₂-VB, respectively, were determined. Rates correlated with the distance of the catalytic core from the semiconductor surface, resulting in slower rates for forward ET and recombination with increasing distance (Fig. 16b).³⁵

Finally, a cobalt phenanthroline catalyst (**4**, Fig. 7) with a salicylic acid anchoring moiety was recently used in DSP. **4** showed higher H₂ evolution activity on TiO₂ compared to the homogeneous system with Eosin-Y as light-absorber. A TON_{cat} of 90 and TOF_{cat} of about 20 were achieved in pH 11 aqueous solution with triethylamine as SED (Table 1). In this work, the lifetime of the DSP system was claimed to be limited by dye stability.¹⁷ Information about kinetics were not provided for this system.

5.2.2 Molecular nickel catalysts in DSP. The nickel bis(diphosphine) complex **5** (Fig. 7) has been the only molecular nickel catalyst studied in DSP in water. Similarly to the cobalt catalysts, **5** features phosphonic acids groups in the outer coordination sphere of the complex for anchoring to the metal oxide surface.⁸ On RuP-sensitised TiO₂, an optimised TON_{cat} of 278 ± 19 and TOF_{cat} of 72 ± 5 h⁻¹ was achieved after 30 h of visible light irradiation (Table 1), setting the benchmark for molecular DSP. However in this work, ascorbic acid (AA) was used as sacrificial electron donor, able to reductively quench RuP* to form RuP⁻ (Fig. 16c, t_{50%} = 250 ns). The highly reducing RuP⁻ species can in turn reduce **5** with t_{50%} = 50 μs when immobilised on ZrO₂. The *on particle* reduction can also occur in the TiO₂ systems and was found to be ~2 orders of magnitude faster than the reduction by TiO₂-CB electrons (t_{50%} = 1 ms, Fig. 16a and c). However, oxidative quenching of RuP* by the TiO₂-CB was 10³ times faster than reductive quenching by AA (180 ps vs. 250 ns), most likely favouring a *through particle* pathway on TiO₂ at certain dye and catalyst concentrations. Both, catalyst and dye, were further identified as life-time limiting species in this system.

6. Challenges and Future Directions

The concept of DSP for light-driven H₂ evolution offers several potential advantages over conventional homogeneous approaches: a straightforward self-assembly of the modular system with a rational strategy to optimise performance, and accumulation of multiple, long-lived and high-energy electrons in the conduction band of the semiconductor for the catalyst to collect.

Despite this improvement, DSP systems still display rather low efficiencies, low long-term stability, and rely on sacrificial electron donors. To improve performance, in-depth studies on the molecule-semiconductor interface are required to reveal the major pathways accounting for efficiency losses. Only few structure-activity relationships are currently available,³⁵ and a better understanding of the electron transfer reactions across the component interfaces will help guide the design of more efficient DSP systems.

The primary source of long-term instability in most reported DSP systems has been the fragility of the molecular photosensitiser and catalyst architecture under photo-catalysis conditions. The development of more robust catalysts and strategies to stabilise catalysts on semiconductor surfaces is required in order to produce a step-change in the performance of these systems.

Avoiding sacrificial agents should be a priority in the future development of DSP systems. Ultimately, H₂ evolution must be coupled to sustainable oxidation chemistry such as water oxidation in order to become commercially viable. Coupling the proton reduction with the water oxidation half-reaction *via* a redox mediator (Fig. 17a) or an external circuit in a photoelectrochemical cell (Fig. 17b) are attractive strategies to remove sacrificial components.^{5,49,50} Challenges lie in choosing a suitable redox mediator and selecting materials that allow transfer of DSP onto an electrode. An alternative approach is

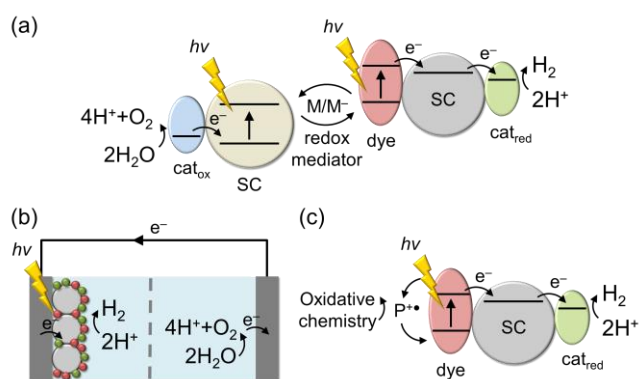


Fig. 17. Concepts for non-sacrificial DSP systems: (a) mediated Z-scheme system for full water splitting; (b) immobilisation of a DSP systems on a cathode, which is wired to an anode for water oxidation via an external circuit; (c) dye regeneration in DSP via oxidation of an organic substrate to a higher value chemical.

to couple the H_2 evolution reaction to other useful one- or two-electron oxidation reactions instead of the 4-electron water oxidation reaction, e.g. oxidation of organic substrates (Fig. 17c). While it is rather straightforward to adjust the thermodynamics of the DSP system to match the required redox potentials of each half-reaction, matching the kinetics of charge separation/recombination and the redox reactions remains highly challenging.

Ultimately, the DSP approach can be easily applied to other chemical processes, such as the reduction of CO_2 to fuels. An early example for the selective photo-conversion of CO_2 to CO are phosphonated $[\text{Re}(\text{bpy})(\text{CO})_3\text{X}]$ (X = leaving group) derivatives, which have recently been used in DSP schemes and exhibited improved long-term stability and activity when anchored to TiO_2 in comparison to the corresponding homogeneous systems.^{3,18} In the future, it can also be anticipated that self-assembly of dye and catalyst on a semiconductor will enable us to synthesise high value chemicals in DSP through organic photocatalysis.

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Notes and references

‡ The molecular catalysts **1** (CoP^1 , CAS: 27-3015) and **5** (NiP , CAS: 28-1720) are commercially available via Strem Chemicals. The photosensitiser **RuP** (CAS: 239472-59-4) is available from Dyanamo AB.

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Left to Right: Katherine Orchard, Janina Willkomm and Erwin Reisner.

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Katherine Orchard obtained her PhD in chemistry at Imperial College London under the supervision of Prof. Charlotte Williams and Prof. Milo Shaffer, where she worked on nanomaterials synthesis. She spent two years in industry developing quantum dots for display technologies before

joining the group of Dr Reisner as a Cambridge-AIMR Joint Research Centre Scientist to work on nanomaterial/molecular catalyst hybrid photocatalysts and photoelectrodes.

Janina Willkomm obtained her master's degree in chemistry at the Leibniz University in Hanover under supervision of Prof. Peter Behrens, working on graphite oxide composite materials. She joined the group of Dr Reisner in 2013 as a PhD student in the Christian Doppler Laboratory for Sustainable SynGas Chemistry. Her research focuses on molecular H₂ evolution catalysts and their integration with semiconducting materials to assemble hybrid systems.



Anna Reynal obtained her PhD in nanoscience and nanotechnology at the Institute of Chemical Research of Catalonia (ICIQ). Her work, supervised by Dr. Emilio Palomares, was focused on the development of dye sensitised solar cells. She then joined Prof. James Durrant's group (Imperial College London), to study the kinetics of hybrid photocatalytic systems for proton and CO₂ reduction. She is now a teaching fellow at Newcastle University, and her research focuses on studying the photocatalytic activity of solar-to-fuel conversion systems.



James R. Durrant (picture right) is Professor of photochemistry in the Department of Chemistry at Imperial College London and is leading the Ser Cymry Solar Initiative at Swansea University. His PhD, supervised by Prof. Lord Porter and Prof. James Barber (Royal Institution), was focused on studying the primary processes of natural photosynthesis. Since then, his research has focused upon chemical approaches for solar energy conversion into electricity (photovoltaics) or molecular fuels. He currently holds an ERC advanced grant and his research has recently been awarded with the Tilden Prize by the Royal Society of Chemistry (2012).

Ernest Pastor (picture left) obtained his degree in Chemistry in the University of Valencia. He joined the group of Prof. James Durrant in 2012 at Imperial College London as a PhD student. His work focuses on the characterisation of materials for solar energy conversion using various electrochemical and laser spectroscopic techniques

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Recent progress and basic understanding in the emerging field of dye-sensitised photocatalysis for light-driven hydrogen production is reviewed.

