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Synergistic effects of β -NaFeO₂ ferrite nanoparticles for photocatalytic degradation, antibacterial, and antioxidant applications†

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Here, synthesis and thorough characterization of β -NaFeO₂ nanoparticles utilizing a co-precipitation technique is presented. XRD analysis confirmed a hexagonal-phase structure of β -NaFeO₂. SEM revealed well-dispersed spherical nanoparticles with an average diameter of 45 nm. The FTIR spectrum analysis revealed weak adsorption bands at 1054 cm⁻¹ suggested metal-metal bond stretching (Fe-Na). UV-Visible spectroscopy indicates a 4.4 eV optical band gap. Colloidal stability of β -NaFeO₂ was evidenced via Zeta potential (−28.5 mV) and Dynamic Light Scattering (DLS) measurements. BET analysis reveals a substantial 343.27 m² g⁻¹ surface area with mesoporous characteristics. Antioxidant analysis indicates efficacy comparable to standard antioxidants, while concentration-dependent antibacterial effects suggest enhanced efficacy against Gram-positive bacteria, particularly Streptococcus. The Photocatalytic activity of β -NaFeO₂ showed significant pollutant degradation (>90% efficiency), with increased degradation rates at higher nanoparticle concentrations, indicating potential for environmental remediation applications.

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1 Introduction

Ever since the pioneering work of Fujishima and Honda in early 1970s,¹ the use of TiO₂ as a photocatalyst for water splitting under UV light irradiation, photocatalysis gained significant interest for *e.g.*, wastewater treatment. Nowadays, photocatalysis is considered a promising approach to address environmental pollution. Generally, different methods exist for mitigating environmental pollution, including biological and chemical approaches. Biological approaches, *e.g.*, bioremediation, *i.e.*, to use microorganisms to degrade pollutants. However, they endure several challenges, such as the requirement of specific environmental conditions, longer remediation times, and potential sludge generation.^{2–6} Additionally, chemical methods, *e.g.*, ozonation, chlorination, or simple application of chemicals to degrade pollutants are often used. However, these lead to the production of harmful by-products and can be economically and operationally inefficient. Thus, photocatalysis has emerged as a promising frontier in the

restoration of environmental pollution due to its relatively high efficiency, low cost and minimal environmental impact.^{7–12}

Besides the most studied material for photocatalysis, *i.e.*, TiO₂, ferrites gained significant attention in this field. For instance, ferrites are considered superior compared to pure metals as magnetic material due to their high electrical resistivity, low costs, facile preparation and enhanced magnetic properties.¹³ Ferrite nanoparticles have gained tremendous attention because of their extensive exploration in various applications in the fields of electronics and biomedics. Among the different ferrites, delafossites belong to the ferrites of alkali and/or alkaline earth metal with a chemical formula of AFeO₂ (where A = alkali/alkaline earth metals).¹⁴ Several delafossite magnetic nanostructures have been reported, including LaSrMnO₃, NaFeO₂, NiFeO₂, MnFeO₂, CuAlO₂, Fe₂O₃, YFeO₃.^{15–19} Among alkali and alkaline ferrite compounds, sodium ferrite (NaFeO₂) is receiving increasing attention, primarily due to its excellent orthorhombic crystal structure,²⁰ which makes NaFeO₂ an excellent candidate materials for various applications, such as lithium ion batteries,²¹ low magnetization ferrofluids,²² electrochromic displays,²³ and wastewater treatment.²⁴ The compounds of sodium ferrite have been found in two stable states, namely Na₁₀Fe₁₆O₂₉ and NaFeO₂. Sodium ferrate oxide (NaFeO₂) exhibits distinct structural phases denoted as α , β , and γ , which are associated with different temperature ranges.²⁵ The α phase predominates at lower temperatures and this phase demonstrates a rhombohedral crystal structure featuring a twofold symmetry with R3m

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space group number, while the β phase becomes prevalent at intermediate temperatures having space group $Pn2_1a$ and orthorhombic structure, and the γ phase emerges at significantly elevated temperatures it assumes a tetragonal structure and is associated with the space group $P4_12_12$.²⁶ This study focuses on utilizing β -NaFeO₂ as a catalyst due to its distinct structural phases and properties.

β -NaFeO₂ is an advanced material with promising photocatalytic properties for environmental remediation. As a sodium iron oxide, β -NaFeO₂ demonstrates exceptional photocatalytic activity under visible light against pollution, and its unique bandgap structure enables visible and ultraviolet (UV) light absorption, making photocatalysis possible across a wide spectrum.²⁷ Solar-to-initiate redox reactions, in which organic pollutants such as methyl red (MR) are degraded, provides a sustainable and low-cost approach to water treatment that can be powerful against some of the challenges presented by MR and other organic dyes. As such, there is a great need to optimize β -NaFeO₂ properties through tailored composition, morphology, and surface modifications in order to improve pollutant degradation. There is also a critical need to mechanistically, kinetically, and pathwise understand the decomposition of MR in order to understand how best to design these complex photocatalytic processes.

Furthermore, we conducted a comprehensive assessment of the antioxidant and antibacterial properties of β -NaFeO₂ nanoparticles, against three different bacteria's including *Escherichia coli* (*E. coli*), *Staphylococcus aureus* (*S. aureus*), and *Pseudomonas aeruginosa* (*P. aeruginosa*). These bacterial strains were selected because they play important roles as the causal agents of a variety of human illnesses, including respiratory tract, wound, and urinary tract infections. We used the MTT (3-(4,5-dimethylthiazol-2-yl)-2,5-diphenyltetrazolium bromide) assay to determine the antibacterial activity of β -NaFeO₂ nanoparticles and assess their capacity to impede the growth of these harmful bacteria. Additionally, we conducted antioxidant activity assays to assess the capacity of β -NaFeO₂ nanoparticles to neutralize free radicals, comparing their performance against well-established antioxidants such as vitamin C and vitamin E. By systematically analyzing both the antioxidant and antibacterial properties of β -NaFeO₂ nanoparticles, our study aims to provide valuable insights into their safety profile and therapeutic potential in combating bacterial infections and mitigating oxidative stress-related diseases.²⁸

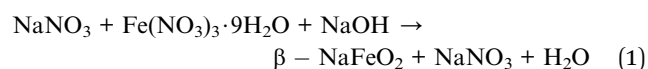
2 Experimental procedure

2.1 Materials

All chemical used for synthesis of β -NaFeO₂ nanoparticles meet with analytical grade requirements. Sodium nitrate (NaNO₃; purity 99%) was procured from Sigma Aldrich, while iron(III) nitrate (Fe(NO₃)₃·9H₂O; purity 98%) and sodium hydroxide (NaOH; purity 95%) were obtained from the Chemistry Lab at Islamia University of Bahawalpur. Hydrochloric acid (HCl; purity 97%) was purchased from Chemical Solutions, Pakistan, and distilled or de-ionized water was sourced from Fiji.

2.2 Synthesis of β -NaFeO₂ nanoparticles

The synthesis of β -NaFeO₂ nanoparticles was conducted *via* the co-precipitation method. Initially, separate solutions of iron nitrate (Fe(NO₃)₃·9H₂O) and sodium nitrate (NaNO₃) were prepared by dissolving 1 M of iron nitrate and 0.5 M of sodium nitrate in 20 mL of distilled water, respectively. Afterwards, the sodium nitrate solution was heated on a hot plate at 80 °C and simultaneously stirred for 30 minutes. Subsequently, the iron nitrate solution was added drop-wise to the sodium nitrate solution with continuous stirring for 30 minutes to ensure thorough mixing and dissolution. Control over the pH was maintained by the 5 mL gradual addition of sodium hydroxide (NaOH) to reach and sustain a pH of 12, crucial for the co-precipitation reaction to stabilize the solvents. This stabilization contributes to the enhancement of nanoparticle properties and serves as a preventive measure against their coagulation when incorporated into a nanofluid.²⁹ The resulting mixture was subsequently heated to 80 °C and kept at this temperature for 3 hours to facilitate the co-precipitation process. The solution gradually turned brown as the nanoparticles formed. After the reaction period, the solution was allowed to cool for 1 hour at room temperature. Subsequently, the precipitate containing the β -NaFeO₂ nanoparticles was separated from the solution *via* centrifugation at 6000 rpm for 30 minutes, followed by decanting of the supernatant. The obtained precipitate was then dried overnight in an oven set to 110 °C to remove any remaining moisture. Finally, the dried powder sample was allowed to cool down to room temperature before further characterization or analysis. Schematic diagram of step-by-step synthesis of β -NaFeO₂ is shown in Fig. 1.



2.3 Characterization

X-ray Diffraction (XRD) analysis was conducted using the Philips PW 1050 Bragg-Brentano diffractometer and Mn filtered Cu-K α radiations to elucidate the crystal structure. Scanning Electron Microscopy (SEM) and Energy-Dispersive X-ray Spectroscopy (EDS) were performed with the NOVA NANOSEM 430 FESEM to investigate the morphology and elemental composition. Fourier Transform Infrared Spectroscopy (FTIR) was carried out using the PerkinElmer Premium HATR Flat Top Plate – ZnSe, 45° configuration, providing insights into chemical bonding. Thermogravimetric Analysis (TGA) was executed on the Mettler Toledo AG – TGA/SDTA851e instrument in a low-frequency mode up to 1600 °C, offering thermal stability information. Ultraviolet-Visible Spectroscopy (UV) measurements using the Research Model UVD-3400/UV-3500 revealed optical properties. The Vibrating Sample Magnetometry (VSM) measurements, conducted with the Cryogenic VSM-MAG-2022-I instrument, offered magnetic property insights. Furthermore, Zeta Potential was determined using the Malvern Zetasizer Nano ZSP. Additionally, the research explores the BET analysis for surface area determination, using nitrogen adsorption/desorption and Quantachrome Nova 2200e instrument.



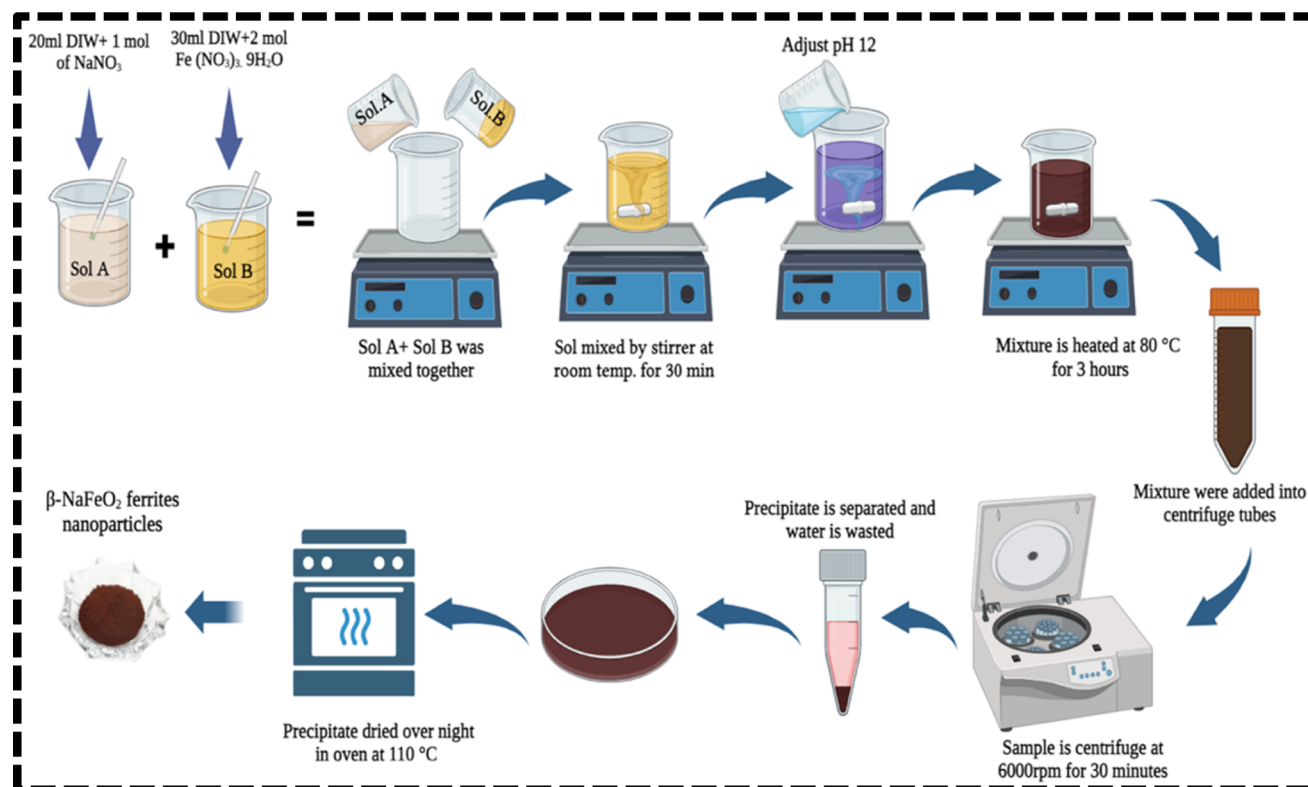


Fig. 1 The schematic illustration of laboratory procedure for synthesizing β -NaFeO₂ ferrite nanoparticles via co-precipitation method.

2.4 Photocatalytic degradation efficiency

The investigation aimed to evaluate the photocatalytic activity of β -NaFeO₂ nanoparticles against methyl red dye. Initially, a solution containing 20 ppm methyl red dye was prepared, and β -NaFeO₂ nanoparticles were added as the photocatalyst. The mixture was kept in the dark for 30 min after proper dispersion to establish adsorption/desorption equilibrium. Subsequently, light irradiation with a wavelength range from 100 nm to 700 nm was applied for a specified period. Samples were withdrawn at regular intervals for analysis using UV-Visible spectroscopy to assess the degradation efficiency,^{30,31} calculated as

$$\% \text{ Degradation efficiency} = \frac{C_0 - C_t}{C_0} \times 100 \quad (2)$$

Here C_0 represents the initial absorbance of the dye solution, and C_t denotes the absorbance of the dye solution at a specified time (t). The investigation will explore the impact of various parameters on the photocatalytic degradation process, including initial dye concentrations (10 to 70 ppm), temperatures (20 °C to 50 °C), pH levels (1 to 14), exposure times (10 to 70 minutes), and photocatalyst dosages.

2.5 Antibacterial activity

The antibacterial efficacy of synthesized β -NaFeO₂ nanoparticles was evaluated against three distinct bacterial groups: *Klebsiella*, *E. coli*, and *Streptococcus*. For the experiment, bacterial suspensions were meticulously prepared using a 0.79% saline solution. Bacterial suspensions were evenly spread on

nutrient agar media within Petri plates to form a dense bacterial lawn. Nutrient agar media served as the cultivation medium for these bacterial strains. Various volumes (10, 15, 20, and 25 $\mu\text{g mL}^{-1}$) of β -NaFeO₂ nanoparticles, with a concentration ratio of 1 : 3, were introduced into the central region of each well. Distilled water served as the control. Following inoculation, the plates underwent a 24 hours incubation period at 35 °C. After this duration, the inhibition zone, indicative of antibacterial effectiveness, was measured using a scale in millimeters or Vernier caliper as shown in Fig. 2.

2.6 Antioxidant activity

The antioxidant activity of β -NaFeO₂ was evaluated using the DPPH (2,2-diphenyl-1-picrylhydrazyl) assay. The β -NaFeO₂ were tested against DPPH at 20, 40, 60, 80 and 100 mg mL^{-1} concentrations; ascorbic acid was used as control. The DPPH solution (1.5 mL) was added to test sample (1.5 mL) of all test tubes, vortex mixed and incubated at room temperature for 40 minutes. The color change from violet to yellow indicated the antioxidant potential. The blank was prepared using methanol (1.5 mL) and DPPH solution (1.5 mL) and their optical density against the colloidal mixture of a sample was measured at 512 nm. The percentage scavenging activity of DPPH was calculated according to the following equation:³²

$$\% \text{ DPPH scavenging activity} = \frac{A_c - A_s}{A_c} \times 100 \quad (3)$$

where A_c is absorbance of control and A_s is the absorbance in the presence of the samples.



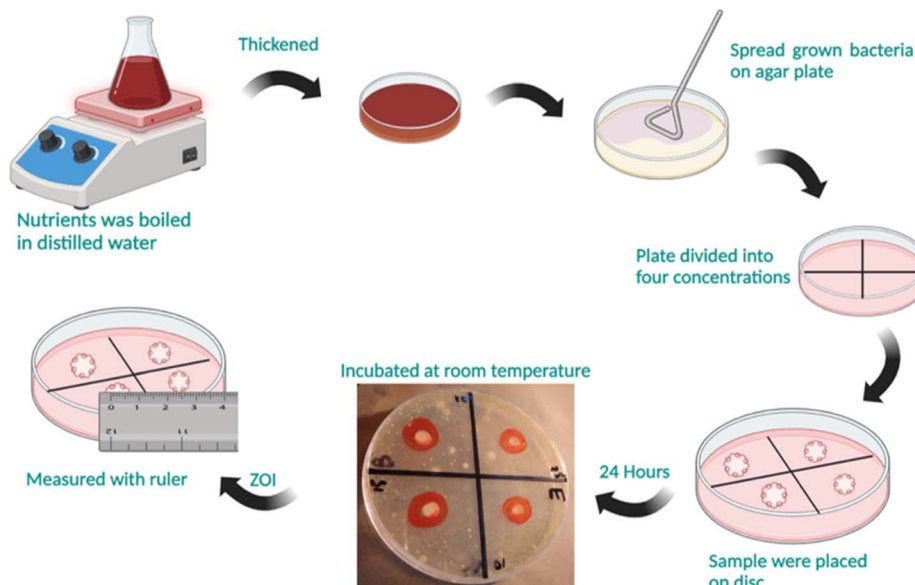


Fig. 2 Steps involved in culturing bacteria on a Petri dish containing agar for assessing the antibacterial activity of β -NaFeO₂.

3 Result and discussion

3.1 Structure and morphology of β -NaFeO₂

The crystallinity of β -NaFeO₂ was assessed using X-ray diffraction (XRD) analysis, as depicted in Fig. 3(a). The diffractions

were plotted within the 2θ angle range of 10° to 80° and matched with their corresponding Miller indices, including (011), (102), (112), (201), (022), (123), (312), (015), (412), and (413), respectively. The presence of a diffraction at approx. 33.5° confirmed the formation of the hexagonal phase, signifying the

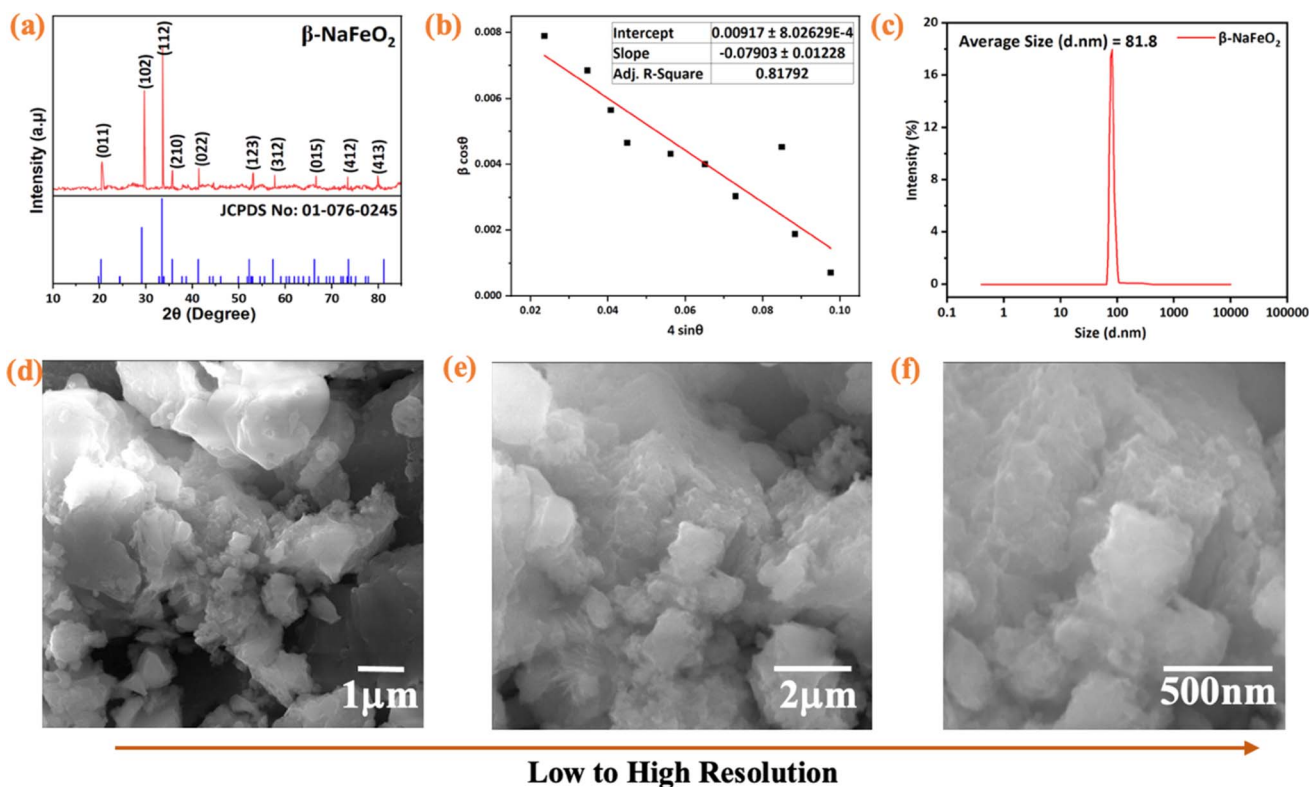


Fig. 3 (a) Analysis of diffraction pattern within 2θ angle range of 10° to 80° by XRD (b) Williamson-Hall plot derived from XRD data of β -NaFeO₂ nanoparticles, yielding an average particle size of approximately 15 nm (c) Dynamic Light Scattering (DLS) analysis of β -NaFeO₂ nanoparticles in solution with an average particle size of approximately 82 nm (d) surface morphology of β -NaFeO₂ nanoparticles by using SEM images at different resolutions (e and f).

orthorhombic structure of β -NaFeO₂. These diffractions correlate with the JCPDS data card number 01-076-0245, with a space group denoted as $Pna2_1$ (Space group number: 33).³³ The high intensity of these diffractions indicates a highly crystalline nature of the prepared material. The average particle size of the prepared β -NaFeO₂ nanoparticles was calculated as approx. 28 nm using Debye–Scherrer's equation $D = K\lambda/\beta \cos \theta$, where D is the crystalline size, K is the shape factor (0.9), λ is the wavelength of Cu-K α , and β is full width at half maximum. Additionally, Williamson–Hall (W–H) equation, expressed as $\beta \cos \theta = K\lambda/D + 4\varepsilon \sin \theta$, was utilized for particle size calculation, where ε represents the strain.³⁴ A plot was constructed with $4 \sin \theta$ on the x-axis and $\beta \cos \theta$ on the y-axis as shown in Fig. 3(b). The crystalline size was obtained from the y-intercept of the plot, which was found to be approx. 15 nm by applying linear regression analysis to the data, while the slope of the linear fit yielded the strain.

Dynamic Light Scattering (DLS) was used to determine the hydrodynamic diameter (HD), which reflects the size of the particles in a solution.³⁵ It is important to note that the sizes obtained by DLS are generally larger than those obtained by XRD and SEM, since the latter techniques provide an estimate of the particle size in the dry state.³⁶ The obtained DLS data of β -NaFeO₂ nanoparticles are summarized in Fig. 3(c), with the average

particle size of approx. 82 nm. This nanoscale dimension is extremely meaningful for a wide variety of applications, as the particle size will have a dramatic impact on biological interactions, as well as systemic distribution in fields such as drug delivery.³⁷ Furthermore, the DLS analysis shows that the particle dispersion is homogeneous and stable providing useful information about the physical characteristics of the nanoparticles.³⁸

The surface morphology and particle size distribution of the β -NaFeO₂ nanoparticles was investigated *via* Scanning Electron Microscopy (SEM). Fig. 3(d–f) shows the SEM images of the nanoparticles at different resolutions, which depicted that the required distribution of particle sizes was relatively uniform, as they appeared as spherical with clear edges, and indicating their porous and homogeneous structure.³⁹ These characteristics leads to large specific surface area and good adsorption performance. Moreover, the SEM in Fig. 2(f) showed that the modified particles were obtained by further deagglomerated and formed a cloud-like structure.

Energy Dispersive X-ray Spectroscopy (EDS) analysis provides a comprehensive insight into the material's elemental composition as shown in Fig. S1.† EDS revealed a distinct composition, showcasing that iron (Fe) constitutes 47% of the overall composition, oxygen (O) comprises 28%, and sodium (Na) accounts for 25%. This composition confirms the purity of the material,

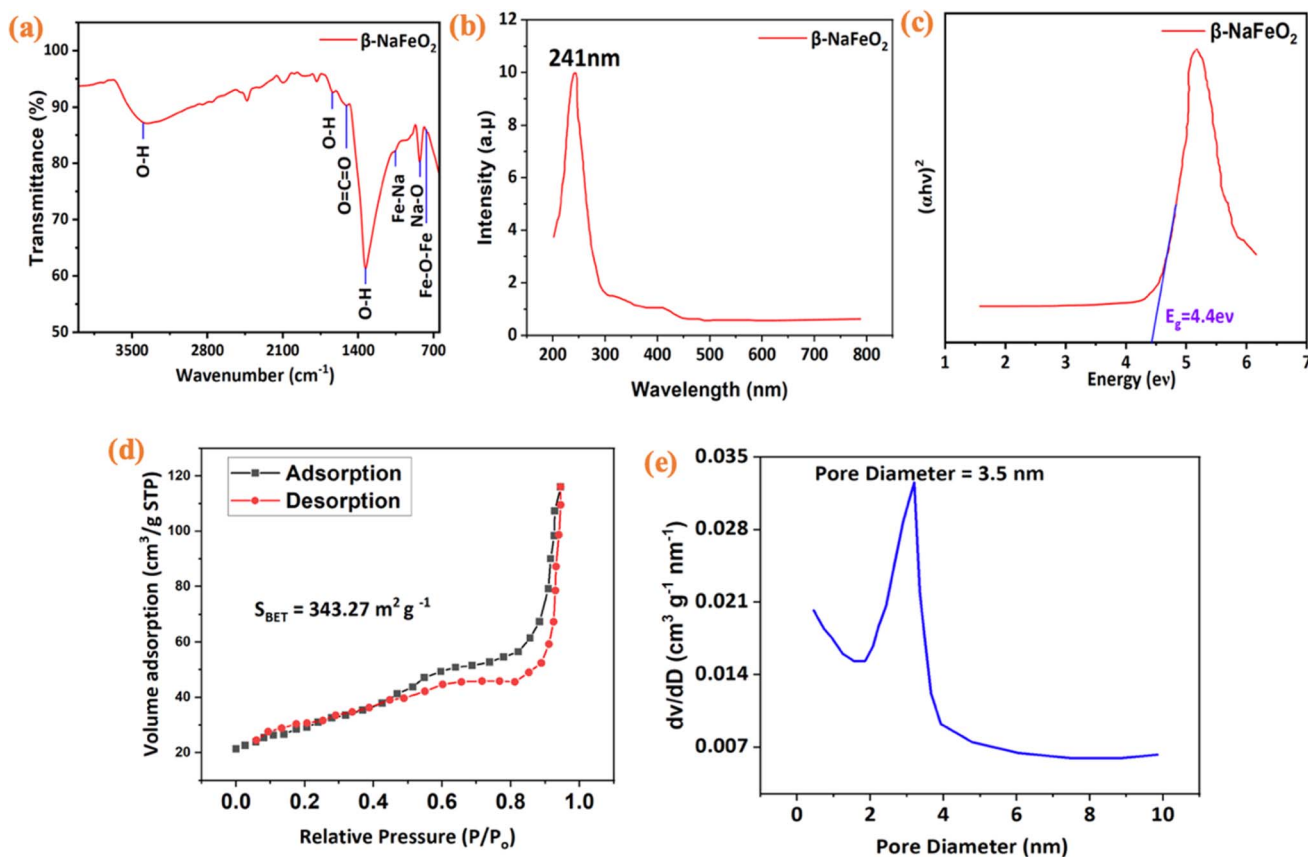


Fig. 4 (a) FTIR analysis of β -NaFeO₂ nanoparticles with characteristic peaks (754 cm⁻¹ for Fe–O–Fe bonds, 834 cm⁻¹ for Na–O vibrations) (b). UV visible analysis of β -NaFeO₂ with the maximum absorbance at 241 nm (c). Optical band gap (E_g) of approximately 4.4 eV (d). Nitrogen adsorption–desorption isothermal curves with the specific surface area of 343.27 m² g⁻¹ (e). Barrett–Joyner–Halenda (BJH) pore diameter distribution having value of 3.5 nm.

specifically demonstrating the presence of pure β -NaFeO₂ without any observable impurities.⁴⁰ The absence of impurities is significant, highlighting the high purity level of the material and its suitability for applications where purity is crucial.

3.2 Comprehensive characterization of β -NaFeO₂

The presence of functional groups on the surface of the synthesized β -NaFeO₂ was analyzed using FTIR. In Fig. 4(a), the FTIR spectrum of β -NaFeO₂ was observed within the span of 640–4000 cm^{−1}. In this spectrum, we observed transmittance peaks at 754 cm^{−1}. This is attributed to the stretching vibrations of Fe–O–Fe bonds.⁴¹ Another peak at about 834 cm^{−1} was attributed to the Na–O vibration. Weak adsorption bands at 1054 cm^{−1} showed the stretching of metal–metal bonds, such as Fe–Na. Finally, the peak at approximately 1502 cm^{−1} is attributed to O=C=O bonds.⁴¹ The peaks at around 1332, 1640, and 3399 cm^{−1} align with the vibrational modes of water's O–H bonds, respectively.⁴²

The changes in composition associated with the calcination process were tracked using Thermogravimetric Analysis (TGA).⁴³ Fig. S2† exhibited distinctive weight loss patterns observed across three temperature regions. In the initial region, from 40 °C to 120 °C, a mass loss of 7.07% indicated the removal of volatile components, likely including water molecules from sources such as ferric nitrate (Fe(NO₃)₃·9H₂O) or citric acid (C₆H₈O₇·H₂O). Transitioning to the second temperature range, spanning 120 °C to 334 °C, a significant mass loss of 43.46% was recorded, suggesting the decomposition of organic constituents such as citric acid and ethylene glycol. In the third temperature region, ranging from 334 °C to 489 °C, a weight loss of 21.88% was observed, likely attributed to the elimination of additional organic residues or the conversion of precursor compounds into the final sodium ferrite (β -NaFeO₂).

At room temperature and within a magnetic field range of ± 5 T, β -NaFeO₂ nanoparticles exhibit superparamagnetic behavior, illustrated by the 'S'-shaped magnetization *versus* applied magnetic field (M – H) curve in Fig. S3.† The coercivity is affected by of porosity, synthesis technique and the size of the particles.^{44–46} The absence of coercive field (H_c) and remanence magnetization (M_r) confirms their superparamagnetic nature.⁴⁷ The saturation magnetization (M_s) at room temperature is measured at 27.11 emu g^{−1}. This change in saturation magnetization (M_s) is attributed to various factors such as surface effects, colloidal stability, nanoparticle volume, interatomic/dipolar interactions, anisotropy energy, dispersed fluid, and environmental factors.⁴⁸ The superparamagnetic nature of these nanoparticles offers several advantages, making them

suitable for applications including magnetic resonance imaging (MRI), targeted drug delivery in cancer treatment, gene therapy, and hyperthermia.⁴⁹ Superparamagnetic nanoparticles possess unique characteristics, including excellent dispersibility, resistance to agglomeration, minimal hysteresis loss, and magnetization retention only in the presence of a magnetic field, holding great promise for diverse applications in advanced electronics, high-performance materials, energy, and transportation industries.⁵⁰

UV-Visible spectroscopy was conducted in the wavelength range of 200–800 nm, as shown in Fig. 4(b). In ultraviolet-visible (UV-Vis) spectroscopy, the measurement of maximum absorbance at 241 nm indicates a significant characteristic of the material.⁵¹ This peak suggests a specific electronic transition or molecular structure within the material that efficiently absorbs light at this wavelength. Additionally, we utilized Tauc's equation ($(\alpha h\nu)^\gamma = A(h\nu - E_g)$) to calculate the optical band gap (E_g) and gain deeper insight into the optical properties. Here, we denote the absorption coefficient as α , Planck's constant as h , the frequency of light as ν , a constant factor as A , and the energy of the band gap as E_g . An evaluated band gap of about 4.4 eV was observed for the synthesized β -NaFeO₂ nanoparticles as displayed in Fig. 4(c). The band gap denotes the energy required for electronic transitions from the valence band to the conduction band. A smaller band gap implies a greater probability of electronic transitions thus render it capable of serving as a potential material for optoelectronic devices and photovoltaic technologies.⁵² The production of β -NaFeO₂ nanoparticles with a maximum absorbance observed at 241 nm along with a band gap of 4.4 eV suggests that this novel material could be suitable for potential applications, *e.g.*, photocatalysis, solar cells, and sensor technology.⁵³

The surface charge of nanoparticles can be determined from the zeta potential, which is an important means by which their colloidal stability can be obtained.⁵⁴ Fig. S4† shows zeta potential of −28.5 mV for β -NaFeO₂ nanoparticles. The zeta potential arises from the charge the particle acquires when it is exposed to a liquid. This negative charge can influence the stability and dispersion of the nanoparticles in a suspension. Generally, higher absolute values (either positive or negative) indicate better stability, as particles with high zeta potential repel each other, preventing aggregation.⁵⁵ The high negative charge on the nanoparticles can enhance their interaction with positively charged drug molecules or cells, facilitating targeted delivery. The repulsion between negatively charged nanoparticles also reduces the likelihood of their undesired aggregation in physiological environments.

Table 1 Comparison of properties of β -NaFeO₂ synthesized in this study with literature values for similar materials

Property	This study (β -NaFeO ₂)	Literature comparison
Particle size (XRD)	~28 nm	258 nm, ²⁵ 45 nm, ⁵⁷
Crystalline size (WH)	~15 nm	22 nm, ⁵⁰ 77 nm ²⁵
Optical band gap (UV-Vis)	4.4 eV	2.35 eV ⁵⁰
Surface charge (zeta potential)	−28.5 mV	−43.51 mV ⁴⁷
Specific surface area (BET)	343.27 m ² g ^{−1}	—



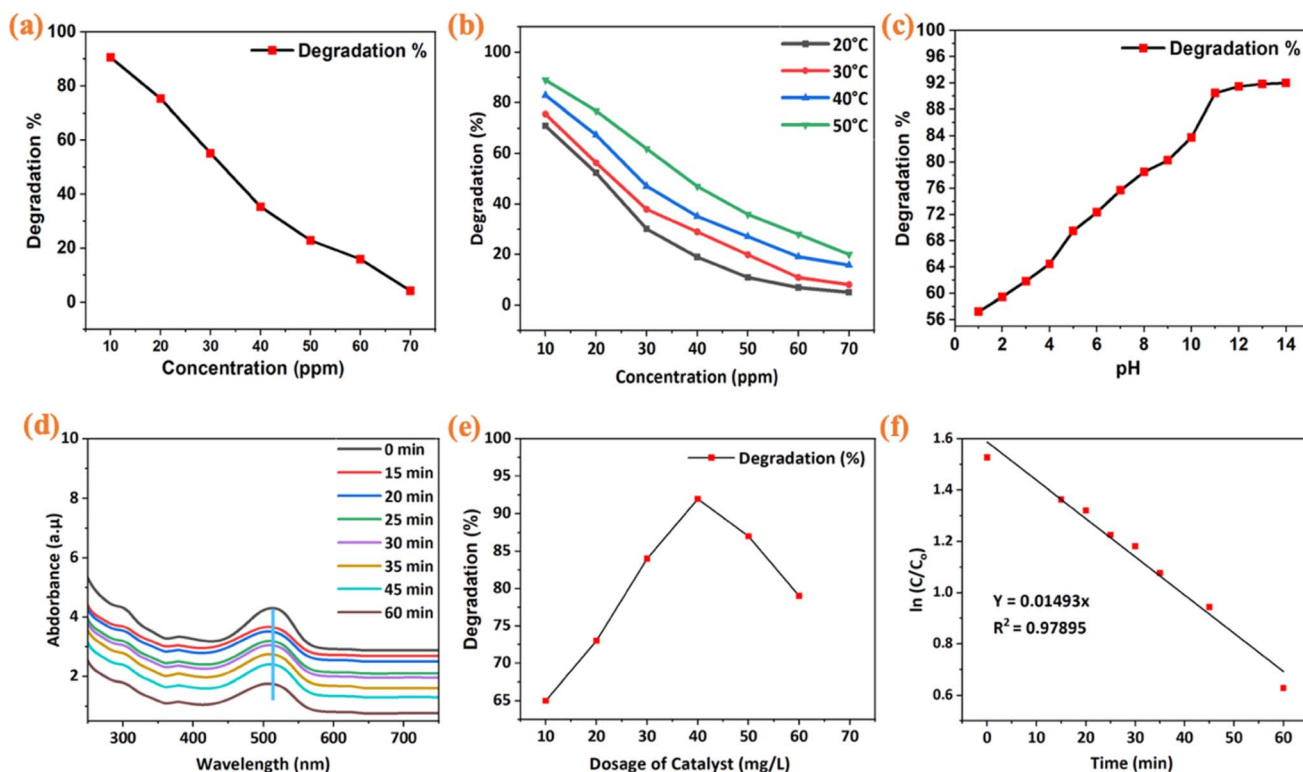


Fig. 5 (a) Impact of varying initial dye concentrations on photocatalytic degradation of methyl red (MR) dye using β -NaFeO₂ (b). Effect of different temperatures ranges (20 °C, 30 °C, 40 °C, 50 °C) on photocatalytic degradation of MR dye with the indication of temperature-dependent enhancement of photocatalytic activity (c). Unveiling the pH-dependent performance dynamics of a promising β -NaFeO₂ catalyst (d). Impact of contact time on MR dye degradation with the optimal degradation between 35 to 60 minutes (e). Effect of catalyst dosage on MR dye degradation, with maximum efficiency observed at 40 mg L⁻¹ dosage (f). Pseudo-first-order linear aggression.

The nitrogen adsorption-desorption isotherm (conducted *via* BET) is presented in Fig. 4(d). The isotherm exhibits a Type II behavior, indicative of non-rigid aggregates of plate-like particles and slit-shaped pores. The hysteresis loop, denoted as H3, suggests the presence of mesopores in the sample. The specific surface area (S_{BET}) calculated from the BET analysis is determined to be 343.27 m² g⁻¹, highlighting a substantial surface area available for nitrogen adsorption. The pore diameter distribution, depicted in Fig. 4(e), indicates a dominant mesopore size at 3.55 nm. This finding is further supported by the identification of a mesopore region, suggesting the existence of pores within this size range.⁵⁶ The calculated pore volume is reported as 0.015 cm³ g⁻¹, indicating the volume of nitrogen gas adsorbed at standard temperature and pressure conditions. Table 1 provides a comparison of the properties of β -NaFeO₂ as investigated in this study with relevant data from the literature.

3.3 Photocatalytic performance of β -NaFeO₂

Fig. 5(a) depicts the impact of varying initial dye concentrations on the photocatalytic activity of β -NaFeO₂ in methyl red (MR) dye degradation. A clear inverse relationship is observed, with higher initial concentrations leading to decreased photocatalytic degradation efficiency. At 10 ppm, degradation percentage was notably high at 91%, but decreased consistently with increasing concentrations, reaching 4% at 70 ppm. This

trend suggests saturation of active sites on β -NaFeO₂ surface with higher concentrations, reducing adsorption and thus decreasing the overall photocatalytic activity.⁵⁸ Fig. 5(b) showcases the effect of temperature on photocatalytic degradation of MR dye using β -NaFeO₂. Results indicate a temperature-dependent enhancement in degradation efficiency. For instance, at 20 °C and 10 ppm concentration, degradation percentage is 71%, rising substantially to 89% at 50 °C. This correlation between temperature elevation and enhanced photocatalytic activity underscores the sensitivity of the process to temperature variations. The impact of pH on the photocatalytic degradation of methyl red (MR) dye using β -NaFeO₂ nanoparticles is depicted in Fig. 5(c). The results reveal a strong correlation between pH levels and degradation efficiency. Lower pH values, *e.g.*, 2 and 3, result in photocatalytic degradation percentages of 57% and 59%, respectively, while pH values ranging from 4 to 14 show a progressive increase in degradation, peaking at 92% at pH 14. However, degradation slows notably beyond a pH of 11, indicating diminishing enhancement. This reduction in efficiency at lower pH is attributed to the positively charged surface of β -NaFeO₂ nanoparticles in acidic conditions, causing electrostatic repulsion with anionic groups of methyl red dye, thus limiting dye adsorption. As pH increases, the surface charge becomes less positive, enhancing electrostatic attraction with dye molecules and promoting adsorption, leading to increased degradation.⁵⁹ However,

Table 2 Coefficient of determination and rate constants for dye concentration, temperature, pH, catalytic dosage, and pseudo first order and Langmuir isotherm

Factors	Coefficient of determination R^2	Rate constant (min^{-1})
Initial dye concentration	0.91732	1.4651
Temperature	0.8722	1.09921
pH	0.9872	2.97919
Photocatalytic dosage	0.28129	0.34286
Pseudo first order	0.97895	0.01493
Langmuir isotherm	0.90963	0.00065

degradation slows notably beyond pH 11, possibly due to the formation of anions from $\text{OH}^\cdot$ radicals reacting with HCl, limiting degradation efficiency. The observed pH variations are crucial for optimizing conditions to maximize the photocatalytic efficacy of $\beta\text{-NaFeO}_2$ nanoparticles in methyl red dye degradation.⁶⁰ The synthesized $\beta\text{-NaFeO}_2$ exhibits superior photocatalytic degradation as compared to other sodium composites and detailed results can be found in Table 3.

Fig. 5(d) presents the effect of contact time on MR dye degradation. Optimal degradation occurs between 35 to 60 minutes, beyond which no further degradation is observed. This cessation suggests a possible saturation point in the photocatalytic degradation process, wherein the available active sites on the catalyst surface might have been fully utilized or the concentration of reactive species may have reached a plateau. Moreover, the maximum wavelength recorded during this period was identified at 514 nm, emphasizing a specific wavelength at which the photocatalytic activity of $\beta\text{-NaFeO}_2$ nanoparticles in degrading methyl red dye is most pronounced. The findings highlight the critical role of contact time in achieving optimal photocatalytic efficiency. Fig. 5(e) examines the effect of catalyst dosage on MR dye degradation. Maximum absorption occurs at 40 mg L^{-1} dosage, with degradation efficiency peaking at 92. Further increases in dosage lead to decreased efficiency, indicating an optimal dosage threshold. Additionally, we carried out radical scavenging experiment by adding scavenger (0.01 M, 2 mL) to the dye solution. The effect of catalyst such as TEMPO (2,2,6,6-tetramethyl-1-piperidinyloxy) on the photo-degradation of methyl red dye was extracted by using the irradiation time of 90 min depicted in Fig. S5.† The degradation % of scavenging TEMPO concentration was about 58% after 15 minutes and then increase to 75% after 60 minutes. This intensified scavenging capability contributes to a more effective

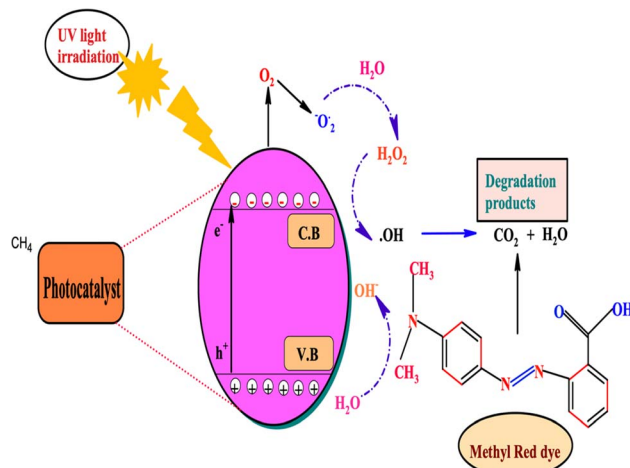


Fig. 6 Proposed charge transfer mechanism and photocatalytic pathway of methyl red dye using $\beta\text{-NaFeO}_2$ as a catalyst.

radical suppression inhibiting the degradation pathways, thereby increasing the percentage of dye adsorption to a higher level. After 60 minutes, no change in degradation was observed.⁶¹ Fig. 5(f) demonstrates the pseudo-first-order kinetics analysis for methyl red dye degradation using $\beta\text{-NaFeO}_2$ nanoparticles. This shows absorbance changes over time, while graph depicts linear regression for the pseudo-first-order reaction. The high coefficient of determination ($R^2 = 0.97895$) indicates strong correlation between experimental and theoretical data, supporting the model's suitability. The calculated rate constant ($k = 0.01493 \text{ min}^{-1}$) suggests moderate degradation speed. These findings provide a robust understanding of the degradation kinetics. In Fig. S6.†, the Langmuir-Hinshelwood model is applied to analyze methyl red dye degradation. $R^2 = 0.90$ indicates a strong correlation with the experimental data, validating the model's representation of reaction kinetics. The calculated rate constant ($k = 0.000613 \text{ min}^{-1}$) signifies the efficiency of $\beta\text{-NaFeO}_2$ nanoparticles in degrading methyl red dye. This comprehensive analysis confirms the applicability of pseudo first-order to describe the degradation process. The rate constant of different factors such as dye concentration, temperature, pH, catalytic dosage, and contact time is presented in Table 2.

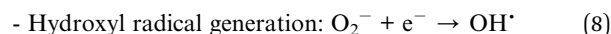
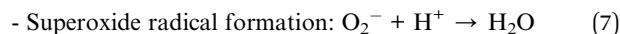
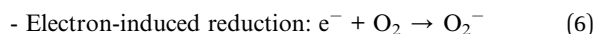
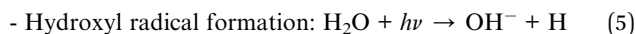
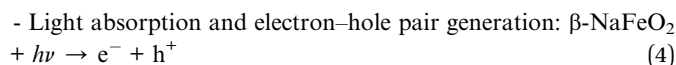
For a more profound understanding, a charge transfer mechanism is suggested in Fig. 6. The degradation mechanism is intricately linked to the pH of dye molecules in the aqueous solution and the surface charge characteristics of nanoparticles.

Table 3 Comparison of photocatalytic activity of $\beta\text{-NaFeO}_2$ with some other Na composites

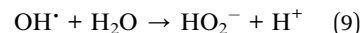
Compounds	Synthesis method	Organic dye	Time (minutes)	Degradation %	Reference
$\text{Bi}_{0.65}\text{Na}_{0.2}\text{Ba}_{0.15}\text{FeO}_3$	Sol-gel	Methylene blue	120	75	64
$\text{Bi}_{0.96}\text{Na}_{0.04}\text{FeO}_3$	Sol-gel	Methylene blue	100	94%	65
$\text{Ag}_{10}\text{-CN-Na}_2\text{SO}_4$	One-pot	Rhodamine B	50	96	66
$\text{Bi}_{0.94}\text{Na}_{0.06}\text{Fe}_{0.94}\text{Co}_{0.06}\text{O}_3$	Sol-gel	Methyl blue	160	99.46	67
$\beta\text{-NaFeO}_2$	Co-precipitation	Methyl red	45	90.56	This study



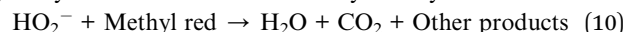
The isoelectric point (IEP) plays a pivotal role, delineating a shift between positive and negative charges on the surface of nanoparticles in acidic and basic solution media, respectively.⁶² This transition in surface charge, influenced by the solution's pH, significantly impacts the interaction between dye molecules and nanoparticles. Upon exposure to sunlight, the surface electrons of the photocatalyst undergo excitation, moving from the valence band to the conduction band and leaving behind positive holes (h^+). These holes then react with water, generating hydroxyl radicals. These radicals, particularly hydroxyl radicals and h^+ , instigate the degradation of dyes. In depth, the electrons present in the conduction band ($C\cdot B$) react with oxygen to form $O_2^{\cdot-}$, while the holes (h^+) generate hydrogen peroxide (H_2O_2) on reaction with water. This hydrogen peroxide then reacts with electrons (e^-) and H^+ ions to yield $\cdot OH$ radicals. Further, the holes also directly react with water to produce hydroxyl radicals. A concatenation of aforementioned radicals, h^+ and OH radicals in tandem attack and degrade the pollutants, converting them to mineral acids such as H_2O and CO_2 .⁶³ The photocatalytic degradation mechanism of methyl dyes using β -NaFeO₂ follows several sequential steps:



- Hydroxyl radical interaction with water:



- Hydroxyl radical reaction with methyl red dye:



3.4 Antibacterial activity of β -NaFeO₂

Fig. 7 shows inhibition zones produced through the action of antibiotics applied against three strains of bacteria: *Klebsiella*, *E. coli*, and *Streptococcus*. The *Streptococcus* shows the widest zone, which indicates high susceptibility. The *Klebsiella* and the *E. coli* showed narrower zones which indicates reduced efficiency of the antibiotics. The concentration-dependent effect is evident, with larger antibiotic disks leading to bigger inhibition zones, suggesting higher antibacterial activity.⁶⁸ Varied zone sizes imply potential differences in the antibiotic's mechanisms against distinct bacterial strains. The outcomes highlight a concentration-dependent decrease in β -NaFeO₂ nano-composite antibacterial activity, manifested by prolonged bacterial detection times and enhanced efficacy against *Streptococcus* (Table 4). This variation is linked to the nano-composite's interaction with bacterial membranes, causing

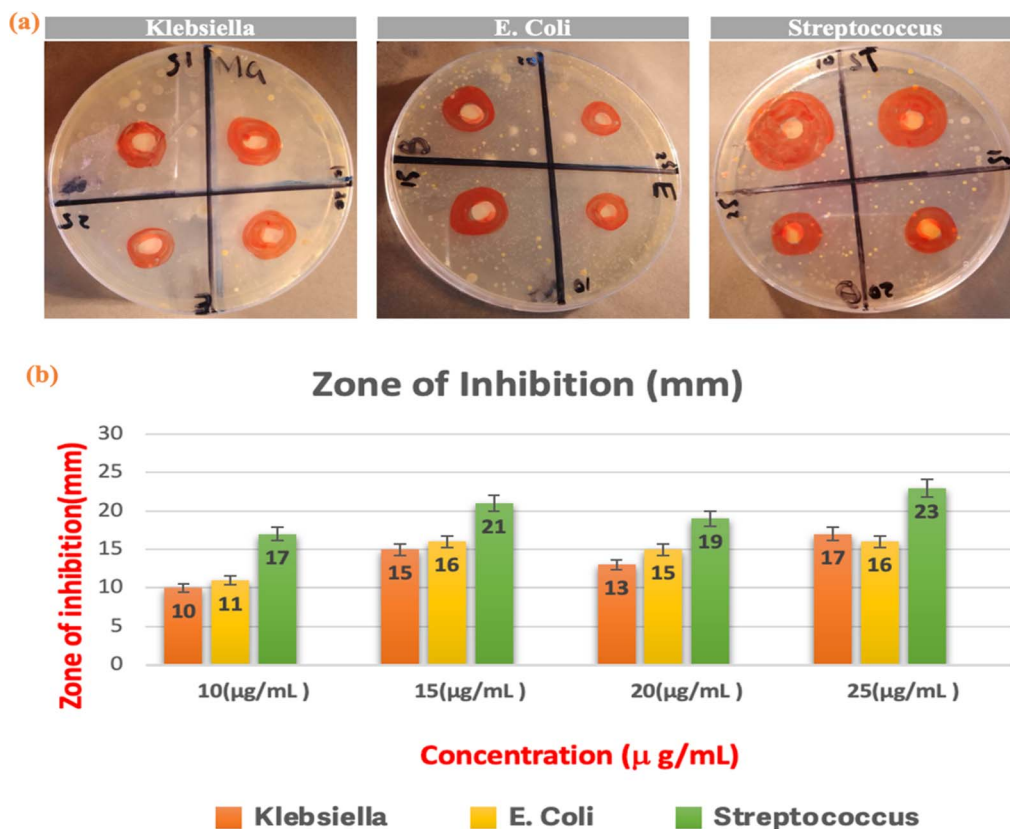


Fig. 7 Inhibition zones against *Klebsiella*, *E. coli*, and *Streptococcus* show varying susceptibility, with *Streptococcus* displaying the widest zone.



Table 4 Results of synthesized β -NaFeO₂ nanoparticles antibacterial activity: zone of inhibition (mm) against various microorganisms

Concentration	Inhibition zone (mm) at different concentrations			
	10 ($\mu\text{g mL}^{-1}$)	15 ($\mu\text{g mL}^{-1}$)	20 ($\mu\text{g mL}^{-1}$)	25 ($\mu\text{g mL}^{-1}$)
<i>Klebsiella</i>	10	15	13	17
<i>E. coli</i>	11	16	15	16
<i>Streptococcus</i>	17	21	19	23

membrane death and showing specificity against Gram-positive bacteria like *Streptococcus*.⁶⁹

The toxicity of β -NaFeO₂ measured in mice for the assessment of its safety to environmental species. The mice being treated with dosages of β -NaFeO₂ via oral or IV way were regularly checked, and the overall health and the survival of the mice were recorded at different time points. Our experimental data showed that no significant symptoms and no death were observed even at high dose level of β -NaFeO₂ in mice for acute toxicity testing. Meanwhile, the necessity of additional research to discover any potential long-term consequences, such as the positioning in tissues and subacute or chronic toxicity that come with long-term exposure to mice organism of β -NaFeO₂ still remains crucial. From the results, β -NaFeO₂ seems to demonstrate a good safety picture in mice under treatment of applied conditions suggesting it's good safety for biomedical applications.⁷⁰

The ability of β -NaFeO₂ to selectively and rapidly kill bacteria implies that it has multiple mechanisms that target essential bacterial cellular processes. It initially disrupts bacterial cell

walls, thereby compromising the structural integrity of the bacterial cell walls, leading to cell lysis and eventual cell death. It also causes massive protein damage as demonstrated by the Cloud Zonal Pellet observed in early time points in the protein profiles in the supplement, which likely arises from the ability of β -NaFeO₂ to damage bacterial ribosomes, thus leading to the inhibition of protein synthesis and consequent disruption of vital cellular processes. β -NaFeO₂ also induces damage to the DNA of the bacterium by inhibiting both its replication and transcription processes, leading to the instability and consequently death of the bacterium. In addition, the compound also disrupts the bacterial ribosomes inducing Destabilization of the ribosomes thereby halting the synthesis of proteins at the level of polypeptide elongation. Finally, β -NaFeO₂ also induces damage to the mitochondria, thus affecting energy metabolism and cellular respiration. These events ultimately lead to cellular collapse and death, emphasizing the diverse and effective antibacterial actions of β -NaFeO₂. Fig. 8 represents the detailed schematic of the mechanism of antibacterial activity associated with β -NaFeO₂.

3.5 Antioxidant activity

Fig. 9(a) presents an antioxidant analysis of β -NaFeO₂, comparing its efficacy with standard antioxidants such as vitamin E, vitamin C, and ascorbic acid. The graph displays the percentage inhibition of free radicals at different concentrations (20 $\mu\text{g mL}^{-1}$, 40 $\mu\text{g mL}^{-1}$, 60 $\mu\text{g mL}^{-1}$, 80 $\mu\text{g mL}^{-1}$, and 100 $\mu\text{g mL}^{-1}$) for each antioxidant. Higher percentages indicate greater effectiveness in neutralizing free radicals. Vitamin E and vitamin C exhibit increased antioxidant activity with higher

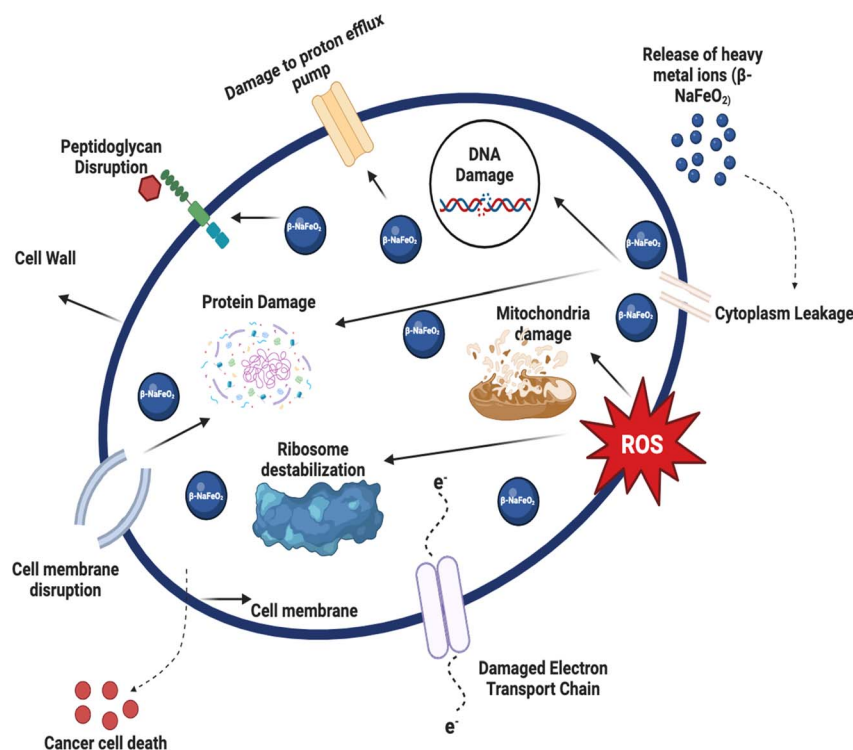
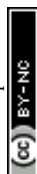


Fig. 8 Illustration of schematic antibacterial activity mechanism of β -NaFeO₂ by targeting essential cellular processes.



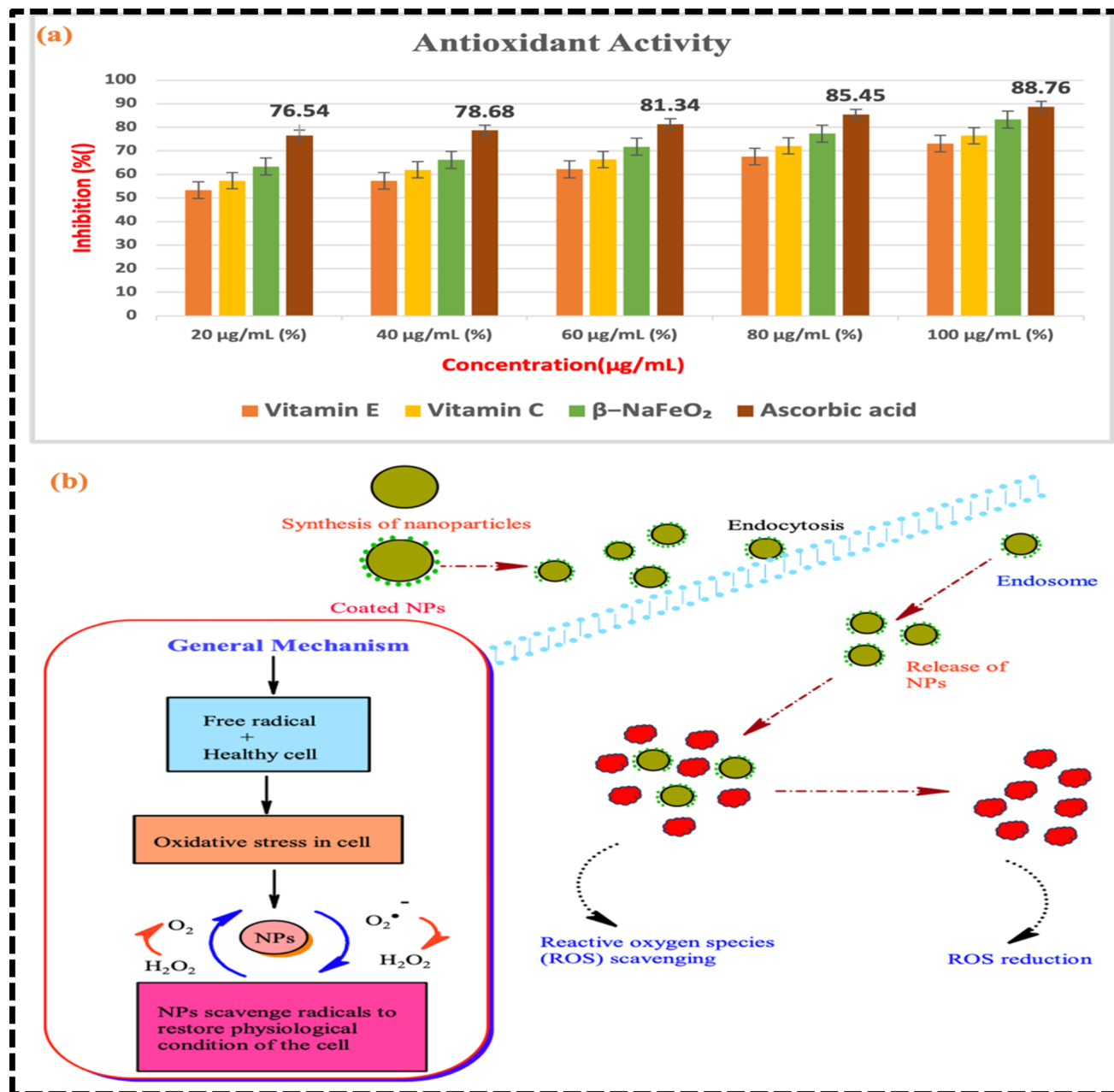


Fig. 9 (a) An antioxidant analysis of β -NaFeO₂ compared to standard antioxidants (vitamin E, vitamin C, and ascorbic acid) at varying concentrations (20 $\mu\text{g mL}^{-1}$ to 100 $\mu\text{g mL}^{-1}$) (b) proposed mechanism of β -NaFeO₂ nanoparticles in combating oxidative stress and protecting cells from free radical-induced damage in biological systems.

concentrations, suggesting efficient scavenging of free radicals. Similarly, β -NaFeO₂ demonstrates notable antioxidant effects, with activity rising as concentration increases, indicating efficient neutralization of free radicals.⁷¹ Ascorbic acid also shows strong antioxidant effects, with a significant increase in activity at higher concentrations. Overall, the analysis underscores β -NaFeO₂ potential as an effective antioxidant, comparable to or even surpassing standard antioxidants like vitamin E and vitamin C. Its concentration-dependent increase in antioxidant activity suggests a promising role in combating oxidative stress and neutralizing free radicals in biological systems (Table 5).

For a deeper understanding, a mechanism illustrating antioxidant activity is proposed in Fig. 8(b). Antioxidants are substances that play a crucial role in protecting cells from damage caused by free radicals. Free radicals are highly reactive molecules or ions that can cause oxidative stress in the body, leading to cell damage and contributing to various diseases, including cancer and cardiovascular diseases. As β -NaFeO₂ has boron, sodium, and iron ions in its structure; it shows quite versatile antioxidant activities. Iron ions react with reactive oxygen species (ROS), *e.g.*, hydroxyl radical ($\cdot\text{OH}$) or superoxide radical ($\text{O}_2^{\cdot-}$), by undergoing redox reactions. Therefore, it scavenges these radicals by donating its electrons. On the other

Table 5 Results of synthesized β -NaFeO₂ ferrite NPs antioxidant activity

Compounds	20 $\mu\text{g mL}^{-1}$ (%)	40 $\mu\text{g mL}^{-1}$ (%)	60 $\mu\text{g mL}^{-1}$ (%)	80 $\mu\text{g mL}^{-1}$ (%)	100 $\mu\text{g mL}^{-1}$ (%)
Vitamin E	53.34	57.2	62.16	67.65	73.02
Vitamin C	57.28	61.91	66.35	72.12	76.43
β-NaFeO₂	63.34	66.16	71.76	77.32	83.34
Ascorbic acid	76.54	78.68	81.34	85.45	88.76

hand, boron ions form very stable complexes with ROS and prevent them to be reactive and damage cellular components by oxidation. β -NaFeO₂ also neutralizes reactive oxygen species by participating hydrogen atom transfer reactions. It donates the hydrogen atoms to ROS and turns ROS into less harmful species. Moreover, it has a quite good ability to chelate the transition metal ions like copper and nickel. These metal ions can catalyze the generation of reactive oxygen species *via* Fenton-like reactions and β -NaFeO₂ prevents this to happen.

4 Conclusions

In conclusion, the comprehensive analysis of synthesized β -NaFeO₂ nanoparticles through various characterization techniques has provided a thorough understanding of their structural, morphological, compositional, magnetic, optical, and catalytic properties. XRD analysis confirmed the highly crystalline nature of the nanoparticles, exhibiting a hexagonal phase with an orthorhombic structure, while DLS and SEM elucidated their nanoscale dimensions, homogeneous dispersion, and porous morphology. EDS analysis confirmed the purity of the material, crucial for its suitability in diverse applications. Furthermore, the superparamagnetic behavior observed through VSM highlights their potential in magnetic resonance imaging (MRI) and targeted drug delivery. UV-Visible spectroscopy revealed significant optical properties, while zeta potential analysis indicated colloidal stability. Photocatalytic studies unveiled their efficacy in degrading methyl red dye, with factors such as initial dye concentration, temperature, pH, contact time, and catalyst dosage influencing degradation efficiency. Antibacterial testing demonstrated concentration-dependent inhibition zones, particularly notable against *Streptococcus* strains, underscoring their potential in biomedical applications. Antibacterial testing revealed concentration-dependent inhibitory zones, most notably against *Streptococcus* strains, highlighting their potential in biomedical applications. Furthermore, antioxidant analysis demonstrated that they were as effective as normal antioxidants, indicating a function in oxidative stress reduction. Overall, the synthesized β -NaFeO₂ nanoparticles exhibit multifaceted capabilities, positioning them as promising candidates for a wide range of scientific and technological applications, with future research endeavors focused on optimizing synthesis techniques and elucidating underlying mechanisms.

Data availability

Data will be made available on request.

Author contributions

Tahira Jabeen: conceptualization; data curation; formal analysis; funding acquisition; investigation; methodology; project administration; resources; software; validation; visualization; roles/writing – original draft; and writing – review & editing. Shahid Khan: conceptualization; data curation; formal analysis; funding acquisition; investigation; methodology; project administration; resources; software; supervision; validation; visualization; roles/writing – original draft; and writing – review & editing. Sana Javid: data curation; formal analysis; funding acquisition; investigation; methodology. Rabia Ayoub: data curation; formal analysis; funding acquisition; investigation; methodology. Waqar Azeem: data curation; formal analysis; funding acquisition; investigation; methodology. Motola Martin: conceptualization; data curation; formal analysis; funding acquisition; investigation; methodology; project administration; resources; software; validation; visualization; roles/writing – original draft; and writing – review & editing.

Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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