



 Cite this: *RSC Adv.*, 2024, 14, 1330

Effect of substitution on the structural, electrical properties, and dielectric relaxor behavior in lead-free BiFeO₃-based ceramics

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BiFeO₃-based ceramics have recently garnered much interest among researchers owing to their valuable and outstanding characteristics. For this reason, the 0.7(Na_{0.5}Bi_{0.5})TiO₃-0.3(Bi_{0.7}Sm_{0.3}FeO₃) ceramic was successfully synthesized by a solid-state route. The central purpose of this research is to investigate the substitution influence of Na, Ti, and Sm on the structural, dielectric, and electric properties of 0.7(Na_{0.5}Bi_{0.5})TiO₃-0.3(Bi_{0.7}Sm_{0.3}FeO₃), as well as to explore its potential applications as it exhibits multiple novel functions. Notably, a structural transition from rhombohedral *R3c* to orthorhombic *P4mm* occurred within this material. In this respect, a suitable equivalent electrical circuit was invested to assess the contributions of grains and grain boundaries to the complex impedance results. Electrical conductivity was attributed to the correlated barrier hopping (CBH) motion of the oxygen vacancies in the sample. The temperature dependence of the dielectric constants revealed the presence of a phase transition. The local disorder provides a dependence of the real part of the permittivity on the frequency which characterizes a relaxor ferroelectric-type behavior of a lead-free material. The modified Curie–Weiss law, in addition to the Vogel Fulcher and Debye law relationships, was utilized to analyze the diffuse transition phase. Furthermore, the studied compound displayed promising electrical properties and chemical stability and proved to be a good relaxor. In this regard, a correlation between dielectric and electric behavior near the ferro-paraelectric phase transition was established.

 Received 12th October 2023
 Accepted 11th December 2023

DOI: 10.1039/d3ra06962a

rsc.li/rsc-advances

1. Introduction

The last decades have witnessed extensive use of perovskite materials in many areas with important applications, such as multilayer capacitors, infrared detectors, shape memories, and electro-mechanical converters.^{1–3} Most applications of perovskite materials depend on their structural and electric properties, which are governed by various fundamental and dominant factors, such as cationic distribution and particle size. Among these structures, there are complex perovskites with two (or more) different atoms occupying the A or B sites at the same time and may contain an ordered arrangement of their positions.^{4,5} The cation ordering can be stabilized if these atoms display sufficient differences in valence, atomic radius, or coordination preference, which are considered the thermal and dynamic driving forces for ordering transition.^{6,7} A high temperature is kinetically required to supply enough energy to overcome the potential barriers of diffusion between different cations. B-site ordering is more common than A-site ordering as

the covalent character is enhanced by the BO₃ framework and can be represented by the general formula, A(B'_xB''_y)O₃, where *x* and *y* correspond to the chemical compositions usually used to describe different types of orders. These complex perovskites undergo a second-order transition from the ordered state to the disordered arrangement above a certain temperature, which is called the order-disorder transition temperature (To-d).^{8–10} It is noteworthy that previous research focused mainly on the effects of 1 : 2 B-site ordering on microwave dielectric properties and proved that B-site ordering can bring significant improvements to microwave dielectric properties, especially the *Q* values.^{11–13} The means of designing the B-site ordering degree and ordered domain structure in 1 : 2 ordered complex perovskites, which is so-called ordered domain engineering, has been further applied to the modification of physical properties and to microwave dielectric properties in recent studies. Furthermore, it has been reported that it improved electrical resistivity, thermal conductivity, and energy storage characteristics.

Recently, bismuth ferrite BiFeO₃ (BFO) was found to have a ferroelectric phase transition at a Curie temperature (*T_c*) of 1103 K, which is much higher than the room temperature.^{14,15} The preparation of phase-pure BFO proved to be difficult owing to the narrow temperature range of the stabilization phase.¹⁶ The presence of such impurities results in a high leakage

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current in the sample, leading to poor ferroelectric behavior. The most common technique developed for the phase formation of pure BFO consists of forming a solid solution of BFO with other ABO₃ types of perovskites.¹⁷ In terms of ion substitution, multiple works have attempted to substitute it with lanthanide rare earth elements. Those enhanced dielectric constants and piezoelectric coefficients proved to be related to the formation of a different structure, namely the rhombohedral phase (*Sm* < 14%), which changes into an orthorhombic structure (*Sm* = 20%).¹⁸ The change from rhombohedral to orthorhombic phase is observed when Fe³⁺ above 33% is substituted with Ti⁴⁺.¹⁹ In addition to ion substitutions, the incorporation of ABO₃ to form solid solutions with BiFeO₃ can dramatically modify the structure and effectively improve the properties of BiFeO₃.²⁰ In addition to BaTiO₃, other types of solid solutions with the same general formula as the latter have also been investigated, such as PbTiO₃,²¹ SrTiO₃,²² or NaNbO₃,²³ which feature different morphotropic phase boundaries (abbreviated MPB). A Ti-doped BiFeO₃ (ref. 24) set forward an electric-optical memory prototype. This device can write in an electric state and read in an optical state. Furthermore, a remarkably high performance was reported which benefited from Ti doping, providing a feasible path to develop next-generation memory devices.

Sm³⁺ and Bi³⁺ proved to have similar radii of 1.032 Å and 1.030 Å, respectively. Thus, the substitution of Bi³⁺ at the A-site of BiFeO₃ with Sm³⁺ is favorable for stabilizing the perovskite phase. On the other side, Ti⁴⁺ was used to substitute the partial Fe³⁺ at the B-site of BiFeO₃ for higher valence cation substitution, which might reduce the charge defects (*e.g.*, oxygen vacancies) of the sample.²⁵ Therefore, better electrical properties were expected in this material. In this research work, the main objective was to examine the mixture and doping effect of BiFeO₃ with the Na, Ti, and Sm elements on the A and B sites and to discuss the results of the structural, electric, and dielectric properties of 0.7(Na_{0.5}Bi_{0.5})TiO₃-0.3(Bi_{0.7}Sm_{0.3})FeO₃, belonging to the ABO₃ oxides group.

1.1. Experimental analysis

The sample with the formula, 0.7(Na_{0.5}Bi_{0.5})TiO₃-0.3Bi_{0.7}Sm_{0.3}FeO₃, was synthesized using the conventional solid-state method. This falls within the framework of obtaining a sample sufficiently homogenous and well crystallized with acceptable and reasonable properties.

The starting precursors, Na₂CO₃, Bi₂O₃, TiO₂, Sm₂O₃, and Fe₂O₃, with a purity of 99% were dried and then weighed in stoichiometric proportions. In the next step, the mixed powder was ground neatly in an agate mortar for 2 h to reduce the size of the grain. The powder was pulverized and pressed into pellets with a diameter of 13 mm and a thickness of 10 mm, under a pressure of 5 tons per cm². Subsequently, the resulting mixture was calcinated at different temperatures (*T*₁ = 750 °C and *T*₂ = 850 °C) for 2 h to eliminate volatile compounds (CO₂). Once the reaction was complete, the obtained powder was ground for 1 h and pressed into a circular disc of 8 mm in diameter and 1 mm in thickness. To obtain high-density

ceramics with perfect crystallization, the pellet was sintered at 1180 °C for 1 h in an electric muffle furnace and slowly cooled to room temperature.

The phase analysis, as well as quantification of the samples, was performed using a Siefert X-ray diffractometer applying CuK_{α1} radiation with a wavelength of 1.5405 Å. The Fullprof package was used to analyze the XRD patterns of the studied samples, while using the Rietveld refinement technique.²⁶ Impedance spectroscopy measurements were undertaken over a wide frequency range from 100 to 10⁶ Hz at temperatures between 623 K and 803 K, using a Solartron SI 1260 Impedance/Gain-Phase Analyzer coupled with a Linkam LTS420 temperature control system and a Solartron 1296 dielectric interface. These measurements were carried out on a cylindrical pellet characterized by a diameter of 8 mm and a thickness of 1.1 mm by installing silver electrodes as contacts on both sides and then mounting them in a temperature control system.

2. Results and discussion

2.1. Structural study

All these results revealed that the substituted content in BiFeO₃ ceramics was favorable for forming a polycrystalline perovskite structure and good crystallization. Fig. 1 traces the evolution of X-ray diffraction data in the solid solution 0.7(Na_{0.5}Bi_{0.5})TiO₃-0.3(Bi_{0.7}Sm_{0.3})FeO₃. All the peaks could be indexed using the orthorhombic (*Pmmm*) space group, which is different from pure BiFeO₃. A structural transition from a rhombohedral phase to an orthorhombic phase was recorded in certain studies of rare earth-doped BiFeO₃, such as Sm, Nd, Ho, and Er.²⁷ Another type of substitution at the A-site was doped by alkaline earth ions, such as Ca, Sr, and Ba, which affected the BiFeO₃ structure.²⁸ On the other hand, when the ionic size and electronegativity of the dopants were like that of Fe³⁺, the dopants tended more to occupy the B-site replacing Fe³⁺, such as doping at B-site with Mn, Ti.²⁹ Moreover, no impurity-phase Fe (55.847 amu)³⁰ was reported.

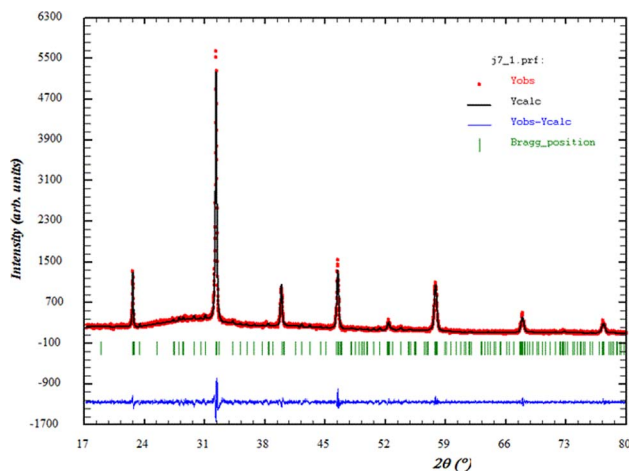


Fig. 1 X-ray diffraction data and Rietveld refinements for 0.7(Na_{0.5}Bi_{0.5})TiO₃-0.3Bi_{0.7}Sm_{0.3}FeO₃.



A typical Rietveld refinement analysis of the sample is portrayed in Fig. 1. The results of the refinement of the structural parameters of our sample are as follows: $a = 2.224 \text{ \AA}$, $b = 3.142 \text{ \AA}$, $c = 5.147 \text{ \AA}$.

Using the X-ray diffraction (XRD) pattern, the size of the crystallites was determined in the case of Scherrer and that of Williamson–Hall (W–H). The average crystallite size was calculated from the full width at the half maximum (FWHM) of the most intense peak by using the Debye–Scherrer equation:³¹

$$D_{\text{SC}} = \frac{k\lambda}{\beta \cos \theta} \quad (1)$$

where λ is the wavelength of the used X-ray radiation; θ is Bragg's angle of the most intense peak, and β is its angular full width at half maximum of the most intense peak.

The Williamson–Hall equation takes into consideration all diffraction peaks and assumes their broadening, which is related to the crystalline size and the lattice strain. This is not considered in the Scherrer equation. Thus, X-ray line broadening is the sum of contributions of small crystallite size (β_{size}) and the broadening results from lattice strain (β_{strain}) in the system:

$$\beta = \beta_{\text{size}} + \beta_{\text{strain}} \quad (2)$$

Where $\beta_{\text{size}} = \frac{K\lambda}{\beta \cos \theta}$ (Scherrer formula) and $\beta_{\text{strain}} = 4\epsilon \tan \theta$.

We obtain the following equation:

$$\beta \cos \theta = \frac{k\lambda}{D_{\text{W}}} + 4\epsilon \sin \theta \quad (3)$$

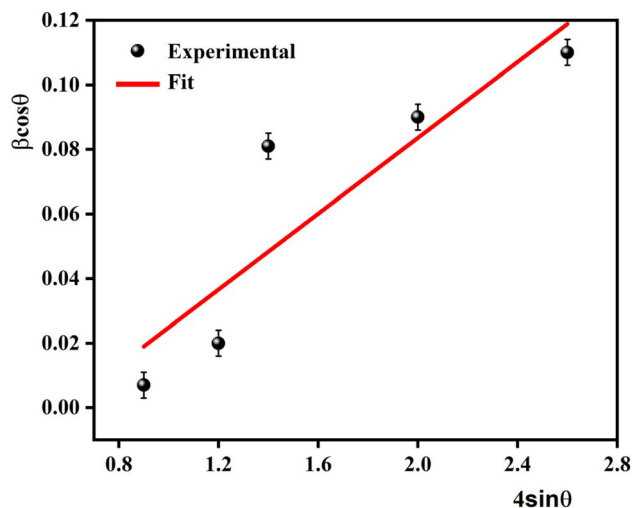


Fig. 2 The Williamson–Hall plots for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.

So

$$D_{\text{W}} = \frac{k\lambda}{\beta \cos \theta} + 4\epsilon \tan \theta \quad (4)$$

β is the FWHM intensity of the diffraction peak; θ is the diffraction angle; k is the shape factor taken equal to 0.9, and λ is the wavelength of Cu-K α radiation (1.540 \AA). In the present case, the instrumental broadening was corrected by considering the FWHM of the standard sample. The crystallite size of the samples was estimated by using the Williamson–Hall equation.

The Williamson–Hall (W–H) plots are displayed in Fig. 2. This plot provides a linear relationship between $\beta \cos \theta$ and $4 \sin \theta$, whereas the slope provides the induced strain ($\epsilon = 1.25 \times 10^{-3}$). The size of the crystallites ($D_{\text{W–H}}$) was examined departing from the intercept of the line with the y-axis. It is noticeable that the straight-line intercepts all the points, indicating a uniform particle size distribution of the crystallites and a very low deformation. The calculated values of average crystallite size D_{SC} and $D_{\text{W–H}}$ are outlined in Table 1. The $D_{\text{W–H}}$ value calculated using the Williamson–Hall method is higher than the one obtained using the Scherrer equation D_{SC} .

$\rho_{\text{X-ray}}$ is the X-ray density that was calculated taking into consideration that a basic unit cell of the perovskite structure contains 8 ions according to the following formula:³²

$$\rho_{\text{X-ray}} = \frac{ZM}{VN_{\text{A}}} \quad (5)$$

where $Z = 1$ is the number of patterns in a unit cell of perovskite lattice; M is the molecular weight of each compound; V is the calculated volume from the lattice parameter, and N_{A} is Avogadro's number detected in these powders. This decline in experimental density with the substitution can be attributed to the smaller weight of the dopant as compared to the bulk density (ρ_{th}) estimated using the following relation:³³

$$\rho_{\text{exp}} = \frac{m}{\pi r^2 h} \quad (6)$$

Where h , r , and m are the thickness, the radius, and the mass of the pellet, respectively.

The compacity was specified according to the formula:³⁴

$$C = \frac{\rho_{\text{exp}}}{\rho_{\text{X-ray}}} \quad (7)$$

The notion of porosity is attributed to the difference between one and the compacity according to the following formula:³⁵

$$P(\%) = (1 - C) \times 100 \quad (8)$$

The different parameters, such as bulk density ρ_{exp} , X-ray density $\rho_{\text{X-ray}}$, C compacity, and porosity P values, were

Table 1 The average crystallite size (nm) (with the different method) for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$

Average crystallite size (nm) (Scherrer method)	D_{SC}	36
Average of crystallite size (nm) (Williamson–Hall method)	$D_{\text{W–H}}$	54
Particle size (μm)	D_{SEM}	1.25



calculated to obtain 3.653 g cm^{-3} , 3.987 g cm^{-3} , 0.913 , and 8.7 (%), respectively.

Assuming that all the particles are spherical, the specific surface area was determined using the following relation:³⁶

$$S = \frac{6000}{D \cdot \rho_{\text{X-ray}}} \quad (9)$$

where D is the diameter of the particle; " $\rho_{\text{X-ray}}$ " is the X-ray density, and the constant 6000 is a form factor for spherical particles. The specific surface area value is $32.543 \text{ m}^2 \text{ g}^{-1}$. These values of specific surface area were relatively high, indicating that our sample of nanoparticles is a potential candidate for multiple applications.³⁷

2.2. Morphological study

A scanning electron microscopy (SEM) analysis was performed to examine the surface morphology of this material. In Fig. 3, the SEM images exhibited in Fig. 3 reveal a microstructure with an almost homogeneous granular distribution and indicate that the morphology changes with doping.³⁸ One can observe that the undoped sample (BiFeO_3) is more compact in terms of agglomeration.³⁹

The micrograph displayed in Fig. 3 illustrates an interesting surface morphology of $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$. Scanning electron microscopy reveals a high melting or accumulation microstructure with polygonal and polyhedron-shaped grains. The average grain size of the particles (D_{SEM}) is deduced by fitting the histogram (Fig. 4) by employing the image analyzer software ImageJ. D_{SEM} which is $\sim 1.25 \mu\text{m}$. This suggests that the grains have a polycrystalline structure, compared to the low crystallite size obtained from the XRD study.

2.3. Electrical study

The impedance plots of the ceramics can have one or more semicircles. Notably, in ceramics, the grain boundaries are highly resistive owing to the presence of a higher density of

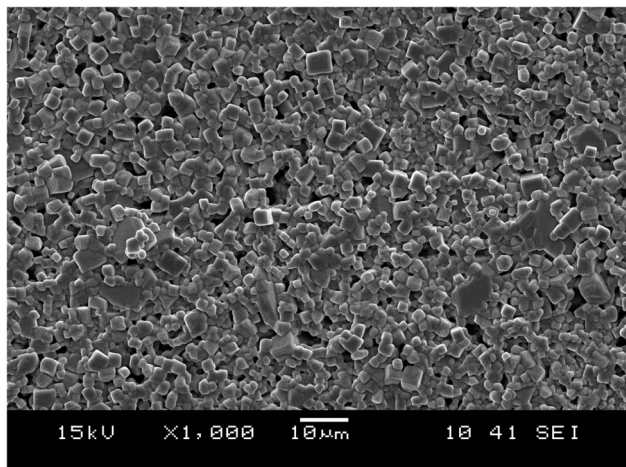


Fig. 3 The SEM image for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.

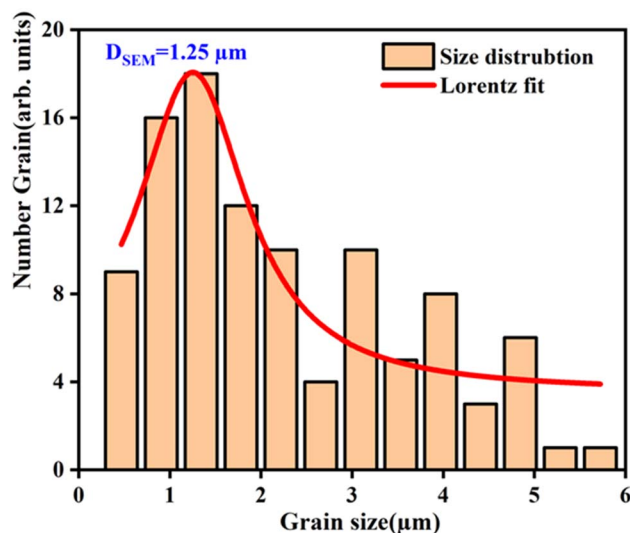


Fig. 4 The histograms of the grain size distribution.

structural and chemical defects compared to grain interiors. Therefore, at high frequencies, the appearance of a smaller diameter semicircle is assigned to grain interiors, and a larger diameter semicircle is ascribed to grain boundaries.^{40,41} Nyquist plots as a function of frequency at different temperatures, of the studied compound are illustrated in Fig. 5. The Cole-Cole plots are characterized by the appearance of semicircles whose diameter decreases with the increase of temperature. The diameter of the semi-circle is related to the resistance of the compound. The increase in temperature reduces the diameter of the semicircle, indicating that the conduction process is thermally activated and therefore the sample is a semiconductor.⁴² The presence of a single semicircular arc implies a single relaxation phenomenon. The Cole-Cole plots are adjusted using the Z-view software,⁴³ with combinations of the

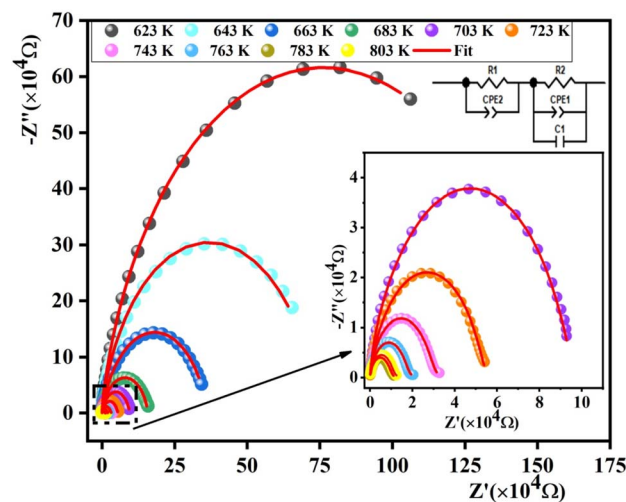


Fig. 5 Samples of the impedance plan plots. The inset shows the electrical equivalent circuit and the spectra of impedance from 723 to 803 K.



Table 2 Electrical parameters obtained from the fit of Nyquist plots using the Zview software for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$

T(K)	R1 (Ω)	Q_1	α_1	R2	Q_2	α_2	C_2
623	28 326	$1,0973 \times 10^{-8}$	0.84914	1.5798×10^6	3.863×10^{-9}	0.62872	3.897×10^{-10}
643	25 284	5.6958×10^{-9}	0.86702	7.2857×10^5	2.8531×10^{-9}	0.68305	3.99668×10^{-10}
663	22 047	5.58×10^{-9}	0.86876	3.4297×10^5	4.408×10^{-9}	0.6799	4.989×10^{-10}
683	15 007	1.069×10^{-8}	0.82985	1.4339×10^6	3.2752×10^{-9}	0.75795	3.499×10^{-10}
703	11 578	5.094×10^{-9}	0.8562	85 890	2.599×10^{-8}	0.49974	5.799×10^{-10}
723	8968	6.5927×10^{-9}	0.81999	87 076	7.5908×10^{-8}	0.4298	6.309×10^{-10}
743	7163	5.223×10^{-9}	0.80822	26 286	4.9993×10^{-7}	0.29344	3.442×10^{-10}
763	5091	7.032×10^{-9}	0.7794	15 769	1.2459×10^{-6}	0.2602	3.1×10^{-10}
783	1517	8.782×10^{-8}	0.67279	10 529	1.293×10^{-6}	0.29291	3.338×10^{-10}
803	2280	3.0762×10^{-8}	0.73221	11 315	2.2064×10^{-6}	0.24041	3.8791×10^{-10}

equivalent circuit models presented in Fig. 5. The appropriate fit was achieved through the use of an equivalent circuit, consisting of a combination series of two circuits: grain (R1-C1) and grain boundaries (R2-C2-CPE2), where R is the resistance and CPE is the constant phase element referring to the complex element.⁴⁴ The presence of both relaxation phenomena (two circuits) in the Nyquist impedance plots can be attributed to the inhomogeneous distribution of oxygen. The values of all fitted parameters are presented in Table 2. In this respect, the good agreement between the theoretical background and experimental findings confirms the contribution of the grains and the grain boundary effects.

On the other hand, we notice that the values of R2 are larger than those of R1, which can be attributed to a disorder of the atomic arrangement near the region of the grain boundaries, and accordingly, it causes an increase in the diffusion of the electrons.

The activation energy can be computed using the Arrhenius law:⁴⁵

$$R_{g,gb} = R_0 e^{-\frac{E_a}{k_B T}} \quad (10)$$

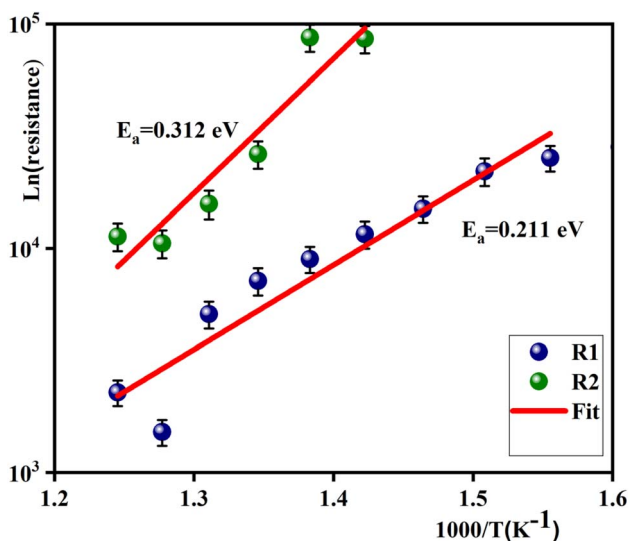


Fig. 6 Arrhenius plots showing the dependence resistance of grain boundaries (R1) and grains (R2) versus $1000/T$ for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.

where R_0 represents the pre-exponential constant or characteristic resistance, k_B is the Boltzmann constant, and E_a is the activation energy. The logarithmic variation of R1 and R2 as a function of the inverse of temperature is conveniently plotted in Fig. 6. The value of the activation energy of the grain

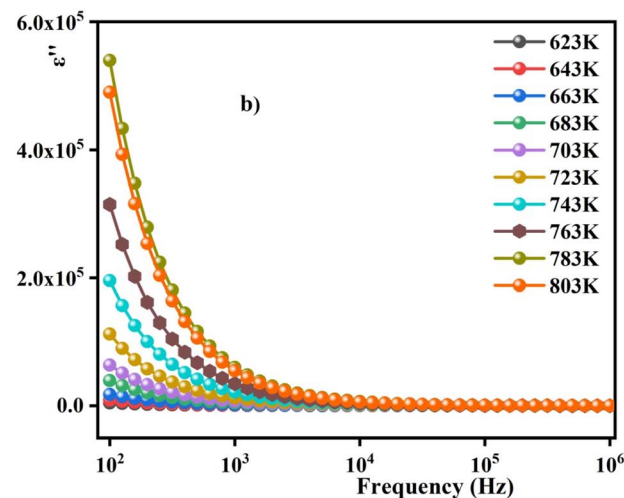
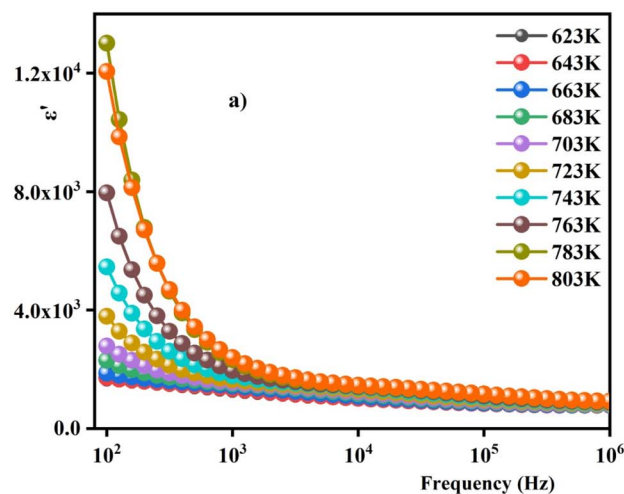


Fig. 7 (a) Frequency-dependent real part (ϵ') and (b) imaginary part (ϵ'') of the dielectric constant measured at different temperatures for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.



boundaries is higher than that of the grains, which indicates a higher resistive behavior than that of the grains, where there is an absence of defects inferred as the material is pure and dense.

Fig. 7(a) and (b) illustrate the variation of the real and imaginary parts $\epsilon'(\omega)$ and $\epsilon''(\omega)$ with frequency on a double logarithmic scale at different temperatures. It can be detected from Fig. 7 that the ϵ' exponentially decreases with increasing frequency and reaches almost a constant value at higher frequencies; which can be associated with the polar nature of the host matrix. There is adequate time for the dipoles to position themselves in the electric field at low frequencies before changing their direction; as they rise, the dielectric loss value.⁴⁶ Conversely, when the frequency becomes high, the dipoles cannot follow up the too-rapid variations of the electric field, which reduces the dielectric loss value. This result corroborates the lossless nature of the samples at higher frequencies. The higher value of ϵ' in the low-frequency region may be ascribed to the presence of mobile charges inside the ceramics backbone.⁴⁷

Fig. 8 foregrounds the complex modulus spectrum (M'' vs. M') of the sample at different temperatures. The appearance of an arc in the spectrum confirms the single-phase character of the material. This goes in good accordance with the observations drawn from XRD, which may be attributed to the presence of electrical relaxation phenomena.

The variation of the imaginary part of modulus (M'') as a function of frequency at various temperatures is plotted in Fig. 9. Departing from the graph, it is clear that (M'') typically exhibits a well-defined maximum to which a characteristic relaxation rate can be associated and also displays asymmetric frequency dependence. The charge carriers may be considered mobile over long distances at lower frequencies lying to the left of the peaks. The frequency region below the maximum peak determines the range in which charge carriers are mobile on long-range distances, and above the maximum peak, the carriers are confined to potential wells being mobile on short

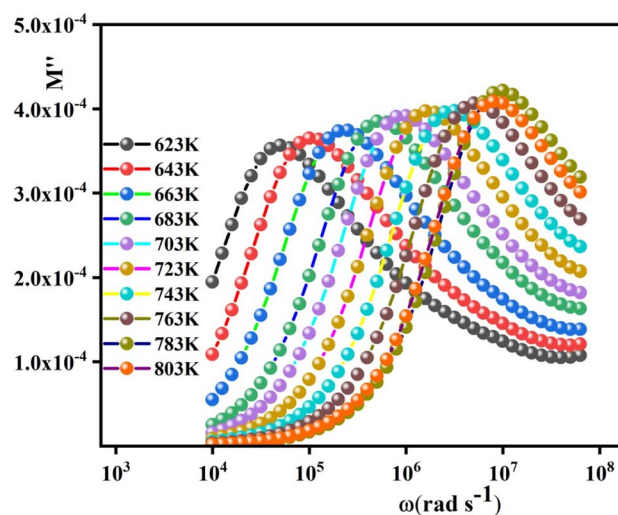


Fig. 9 The imaginary part of the electric modulus as a function of frequency for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.

distances. This region of the peak is therefore indicative of the transition from long-range dominated to short-range dominated mobility.⁴⁸ In addition, the shifting of the peak frequencies in the higher frequency region with temperature implies that the relaxation time decreases with increasing temperature.

On the other hand, to determine the activation energy E_a of the relaxation process in the studied sample, the thermal evolution of the relaxation time τ is traced in Fig. 10. The latter is related to the relaxation frequency f_{max} of the maximum of each peak, by $2\pi\tau f_{\text{max}} = 1$. The obtained plot abides by the following Arrhenius law:⁴⁹

$$\tau = \tau_0 \left(\frac{E_a}{e k_B T} \right) \quad (11)$$

where τ_0 is a constant and k_B is the Boltzmann constant. The activation energy is calculated departing from the slope of the

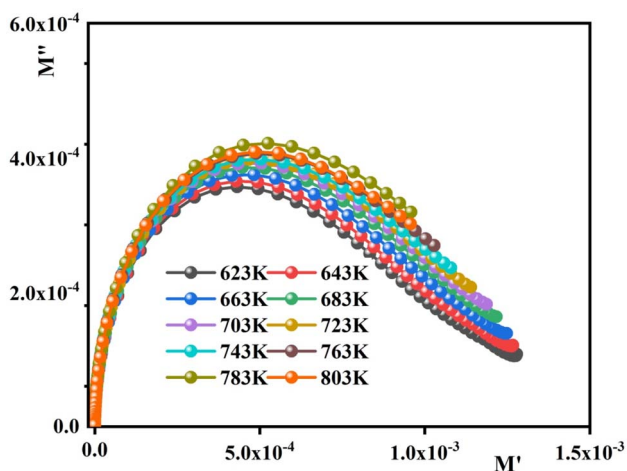


Fig. 8 Complex modulus spectrum of the material at different temperatures.

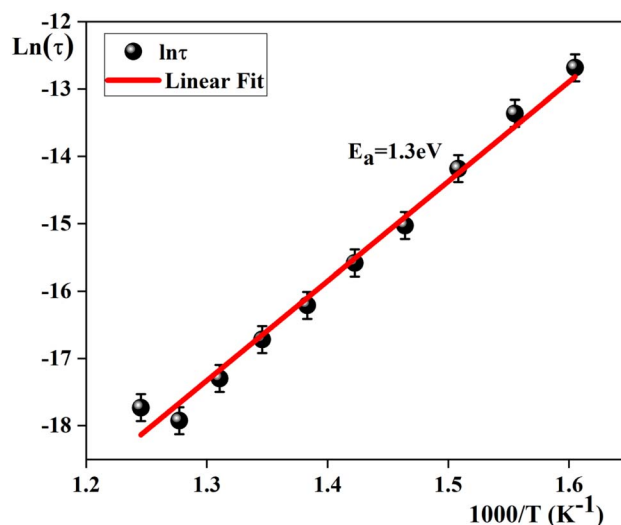


Fig. 10 Variation in $\ln(\tau)$ versus $1000/T$.



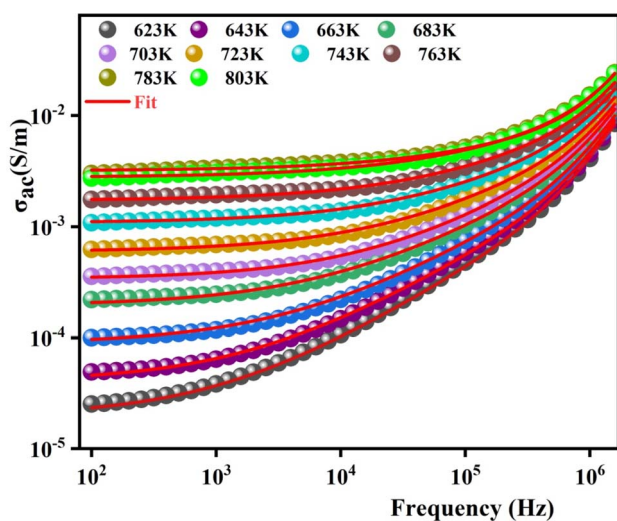


Fig. 11 The variation in the electrical conductivity versus the frequency at different temperatures for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.

linear fit of the plot and its value is 0.85 eV. Such a high value is typical of the ionic origin of the transport mechanism.⁵⁰ With increasing temperature, the sample exhibits a rapid variation in the hopping rate. Hence, the ions need high energy to jump from one site to another.

To gain a deeper insight into electrical behavior and determine the nature of conduction as well as the parameters controlling this process, the variation of the ac conductivity as a function of frequency is represented in Fig. 11. These spectra demonstrate that the conductivity increases with the increase in temperature. At the lower frequencies (10^2 – 10^4 Hz), the conductivity spectra display nearly constant values (σ_{dc}) as the random distribution of the charge carriers gives rise to frequency-independent conductivity. DC conductivity is thermally activated, which can exhibit the semiconductor behavior by the leap of the localized load carrier.⁵¹ The semiconductor nature indicates that this compound can be promising in certain applications such as optoelectronics, photodetectors, and photovoltaics.^{52–54} With increasing frequencies, the σ_{ac} shows a dispersion that moves towards the higher frequencies with the increase in temperature. The conductivity spectra are identified by Jonscher's universal power law in the form of:⁵⁵

$$\sigma(\omega) = \sigma_{\text{dc}} + \sigma_{\text{ac}} = \sigma_{\text{dc}} + A\omega^S \quad (12)$$

In this expression, σ_{dc} and σ_{ac} describe the conductivity in continuous and alternative currents, respectively. A is a constant depending on temperature.

The results of fitting are summarized in Fig. 12. A good consistency between experimental and theoretical data is noted.

The logarithmic variation of dc conductivity as a function of the inverse of temperature is depicted in Fig. 13. This variation can be described by the Arrhenius equation as:⁵⁶

$$\sigma_{\text{dc}} \cdot T = \sigma_0 e^{-\frac{E_a}{k_B T}} \quad (13)$$

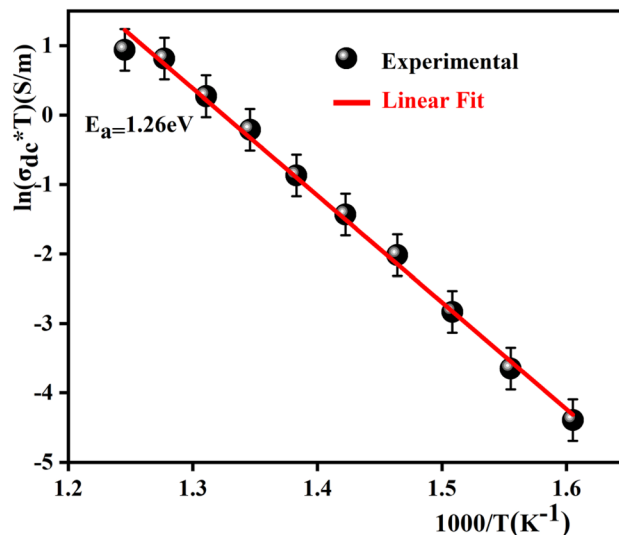


Fig. 12 The variation in $\ln(\sigma_{\text{dc}} \times T)$ versus $1000/T$ for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.

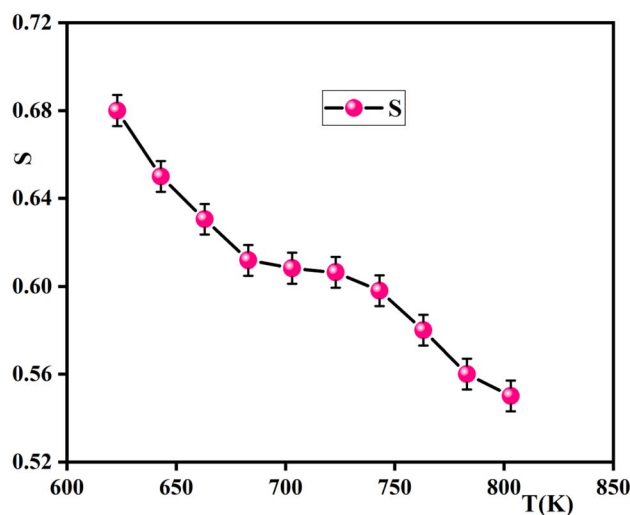


Fig. 13 Temperature dependence of the exponent s for the sample.

where σ_0 denotes the pre-exponential factor, k_B is the Boltzmann constant, and E_a is the activation energy. Relying on the slopes of the linear fit of $\ln(\sigma_{\text{dc}} \times T)$ as a function of $1000/T$, we can estimate the values of the activation energies associated with dc conduction. This value proved to be 1.28.

The exponent “ S ,” which represents the degree of interaction between ions and their surroundings, represents an outstanding source of information about the model for the conduction mechanism in the material from which it is clear that s declines with the decrease of temperature (Fig. 13). This indicates that the Correlated Barrier Hopping (CBH) model is suitable to explain the conduction process. Based on the CBH model, the frequency exponent s is computed as:⁵⁷

$$S = 1 - \frac{6k_B T}{W_m - k_B \ln\left(\frac{1}{\omega\tau_0}\right)} \quad (14)$$



Here, W_m is the height of maximum barrier (or binding energy) which is defined as the energy required to remove polaron completely from one site to another. It is calculated in experiments starting from the linear line slope of (1 s) as a function of temperature. τ_0 is the characteristic relaxation time which is in the order of an atom's vibrational period ($\tau_0 \sim 10^{-12}$ s) and k_B is Boltzmann's constant.

2.4. Dielectric study

Dielectric measurements were carried out as a function of temperature and frequency to investigate their dielectric behavior by determining the ferroelectric phase transition temperature and the nature of the relaxing ferroelectric behavior of the perovskite.

Fig. 14(a) illustrates the variation of the real part of dielectric permittivity (ϵ') as a function of temperature at different frequencies for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$ ceramics. The ϵ' values increase gradually with the increase in temperature and frequency, respectively. This refers basically to the intrinsic behavior of BiFeO_3 's bulk. Up to 320 K, the dielectric permittivity gradually increases as the temperature rises. This

behavior can be assigned to orientational polarization where dipoles align themselves along the electric field direction and fully contribute to the total polarization.⁵⁸ As the temperature rises above 350 K, permittivity starts to increase abnormally referring to the ferroelectric–paraelectric transition. The fluctuation of the dielectric constant as a function of temperature suggests the existence of a peak at a specific temperature (T_{max}), confirming that the dielectric constant has a relaxor thermal evolution. The relaxor behavior is triggered by homovalent and heterovalent substitutions at sites A and B for the studied ceramics and the composition fluctuation gives rise to the relaxing behavior, favoring the strong heterogeneity. Furthermore, the creation of the Coulomb interactions at long distances prevents the formation of ferroelectric microdomains in favor of the formation of the polar nanoregions PNRs.⁵⁹ Similar observations were reported with other ceramics.⁶⁰

Moreover, the values of ϵ' for the doped system proved to be higher than BiFeO_3 , which can be ascribed to the homovalent and heterovalent substitutions at sites A and B for the studied ceramics. These substitutions enhance the polarisability effect and consequently increase the polarisation in the doped BiFeO_3 materials.⁶¹ In comparison with other similar ceramics, we notice that the insertion of the dopant ions entails an increase in permittivity of the ceramics (Table 3) owing to the enhancement of polarisation in the doped materials.^{62–64}

The thermal variation of $1/\epsilon'$ of this compound is presented in Fig. 14(b), demonstrating a deviation from the Curie–Weiss law provided by the equation:⁶⁵

$$\epsilon'_r = \frac{C}{T - T_0} T > T_C \quad (15)$$

Further theoretical fitting reveals that the transition is diffuse in this case and that the behavior is of the relaxor type. This is a second-order transition, where C is the Curie–Weiss constant and T_C is the Curie–Weiss temperature. Since ion diffusion occurs with increasing temperature, the movement of T_C (i.e., the transition from the antiferroelectric to the paraelectric state) at high temperatures is most likely due to space charge polarization.⁶⁶

It is noteworthy that the evolution of the maximum peak temperature of permittivity depending on frequency was analyzed using two theoretical models, namely the Debye law and Vogel–Fulcher (V-F). Dipoles can rotate and are thermally activated in Debye relaxation. The moments of the dipoles are the same, and there is no interaction between them. The Debye model is expressed in terms of:⁶⁷

$$f = f_0 \exp \left[- \frac{E_a}{k_B T_m} \right] \quad (16)$$

where E_a is the activation energy for the jump, f_0 is the pre-exponential factor, T_m is the temperature of $\epsilon'(T)$ maximum, and k_B is the Boltzmann constant. Departing from the fit (Fig. 15(a)), the unphysical higher values $f_0 = 1.6 \times 10^5$ Hz and $E_a = 1.49$ eV were determined.⁶⁸ In the solid state, the attempted frequency proved to cover a broad range of values between 10^6 and 10^7 Hz. The value of f_0 was higher than the typical value for non-

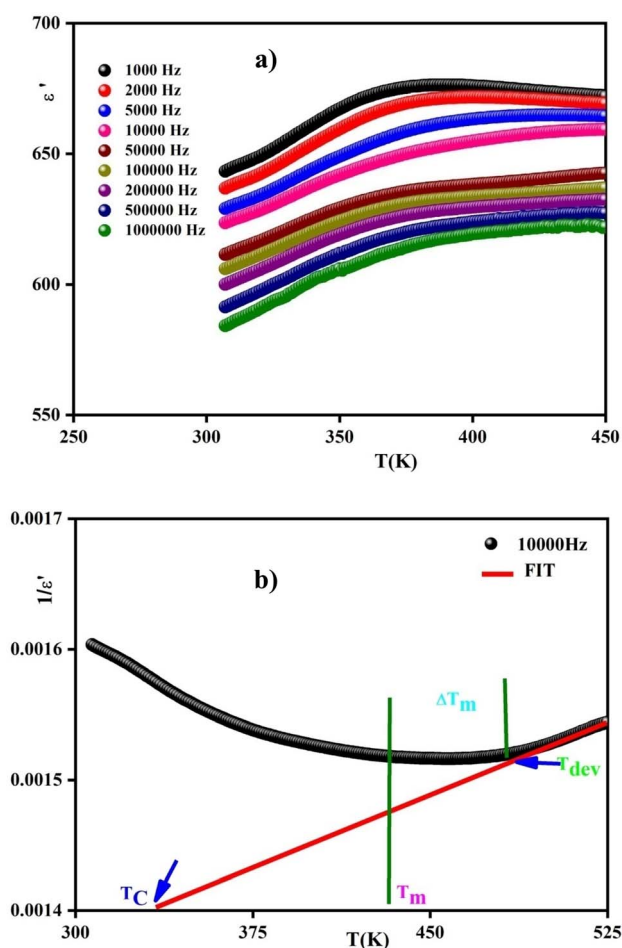


Fig. 14 (a) Thermal variation in permittivity ϵ'_r for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$ (b) Thermal variation in the inverse of the real part of the permittivity $1/\epsilon'_r$ at 1 kHz, for $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$.



Table 3 A comparison of the dielectric and electric parameters

	Our material $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3\text{Bi}_{0.7}\text{Sm}_{0.3}\text{FeO}_3$	(Pb, Ti)-doped BiFeO_3 (ref. 62)	Gd-doped BiFeO_3 (ref. 63 and 64)
ϵ'_{rmax}	675	600	530
T_C	350 K	560 K	460 K

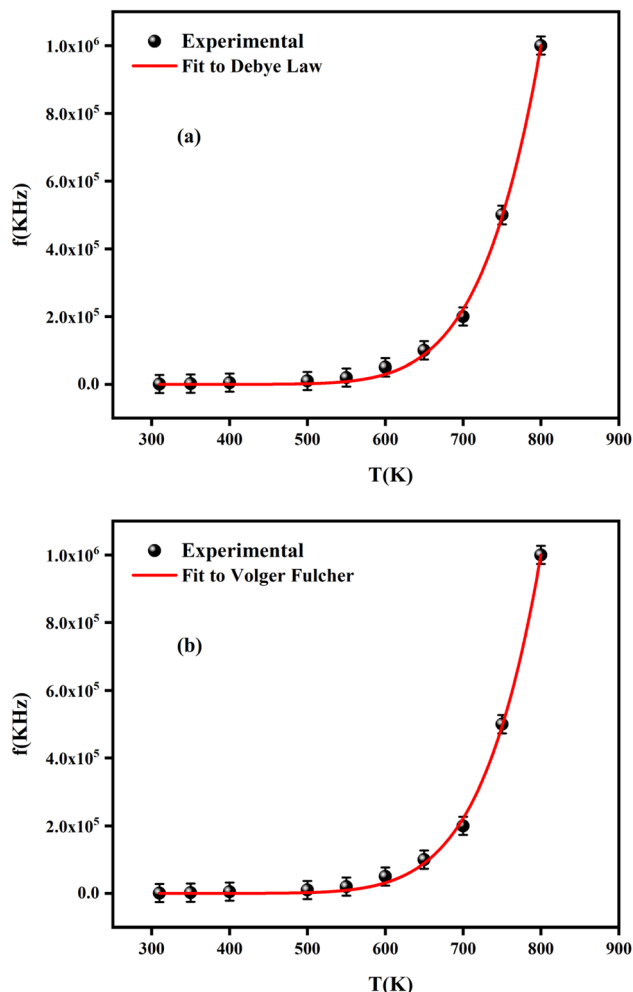


Fig. 15 Fitting the dielectric relaxation data of the sample of (a) Debye and (b) Vogel–Fulcher model.

interacting dipoles. The Debye model cannot explain the dynamics of these dipoles, revealing the presence of non-negligible dipole–dipole interactions. For this reason, further analysis of the relaxation features was undertaken. The experimental curves were fitted according to the Vogel–Fulcher formula, which is described by a relation similar to that recognized for cluster glasses. It can be estimated by:⁶⁹

$$f = f_0 \left[-\frac{E_a}{k_B(T_m - T_{VF})} \right] \quad (17)$$

where T_{VF} is the freezing temperature. The obtained parameters from the best fit to the data (Fig. 15(b)) are $E_a = 0.27$ eV, $T_{VF} = 316$ K, and $f_0 = 3.2 \times 10^5$ Hz. About phonon-mediated charge carrier hopping, the activation energy is comparable to the

activation energy required for the oscillator relaxation.^{70,71} The parameters (T_{VF} and E_a) are compatible with thermally activated polarization fluctuations and are compared to the thermally induced spin glass freezing process, in which cluster flipping and interparticle interaction mechanisms control the freezing process. This suggests a diffuse phase transition (DPT) observed in relaxor ferroelectrics near the ferro–paraelectric transition for the bulk sample.

3. Conclusion

The effect of Na, Ti, and Sm on the structural, dielectric, and electric properties and the temperature of the phase transition of BiFeO_3 was explored, investigated, and discussed.

Unit cell parameters at room temperature for all studied compounds were determined. It was demonstrated that Na, Ti, and Sm substitutions, tend to stabilize the orthorhombic structure.

SEM micrographs revealed the appearance of particles of unequal size. The ac conductivity proved to meet the universal power law. The Z'' and M'' spectra indicated the scaling behavior. The thermally activated process variation with temperature and frequency was explained based on the CBH model. The behavior of the dielectric constant observed in this sample proved to be a typical ferroelectric–paraelectric phase transition with relaxor-like behavior. As a final note, it is noteworthy that the obtained results can be regarded as valuable and worthwhile as they suggest that bismuth-doped nanocrystalline $0.7(\text{Na}_{0.5}\text{Bi}_{0.5})\text{TiO}_3-0.3(\text{Bi}_{0.7}\text{Sm}_{0.3})\text{FeO}_3$ can be a promising room-temperature multiferroic material for memory devices. Indeed, the sample has a high dielectric constant, a low loss tangent, and a high T_C , and can therefore be considered an outstanding material for multifunctional devices.

Conflicts of interest

The authors declare that they have no conflict of interests.

Acknowledgements

The researchers would like to acknowledge the Deanship of Scientific Research, Taif University for funding this work.

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