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Enhanced photocatalytic performance of SnS₂ under visible light irradiation: strategies and future perspectives

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Tin(II) sulfide (SnS₂) has emerged as a promising candidate for visible light photocatalytic materials. As a member of the transition metal dichalcogenides (TMDs) family, SnS₂ features a band gap of approximately 2.20 eV and a layered structure, rendering it suitable for visible light activation with a high specific surface area. However, the application of SnS₂ as a visible light photocatalyst still requires improvement, particularly in addressing the high recombination of electrons and holes, as well as the poor selectivity inherent in its perfect crystal structure. Therefore, ongoing research focuses on strategies to enhance the photocatalytic performance of SnS₂. In this comprehensive review, we analyze recent advances and promising strategies for improving the photocatalytic performance of SnS₂. Various successful approaches have been reported, including controlling the reactive facets of SnS₂, inducing defects in the crystal structure, manipulating morphologies, depositing noble metals, and forming heterostructures. We provide a detailed understanding of these phenomena and the preparation techniques involved, as well as future considerations for exploring new science in SnS₂ photocatalysis and optimizing performance.

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1 Introduction

The family of transition metal dichalcogenides (TMDs) has garnered significant attention in photocatalytic applications owing to its visible light band gap and layered structure, which provides a high specific surface area.^{1,2} However, the traditional utilization of primary TMD materials such as MoS₂, MoSe₂ and WS₂ has predominantly been as supporting photocatalysts alongside other materials like ZnO,^{3–6} TiO₂,^{7–10} C₃N₄¹¹ and others, due to their properties akin to graphene.^{12,13} In contrast, unlike MoS₂, MoSe₂ and WS₂, SnS₂ has emerged as a promising photocatalyst under visible light irradiation, attributable to its suitable band gap energy for visible light.^{14,15}

SnS₂ belongs to the family of layered structure transition metal dichalcogenides (TMDs), characterized as an n-type semiconductor with a band gap ranging from 2.18 to 2.44 eV.¹⁶ In practical applications, SnS₂ has found widespread use

across various fields such as batteries,^{17,18} sensors,^{19,20} and photocatalysis due to its low cost, low toxicity, excellent stability, and abundant natural reserves.²¹ SnS₂ can be synthesized using various techniques, including hydrothermal,^{22,23} solvothermal,^{24,25} CVD,^{26,27} and pulse laser deposition,^{28–30} among others. Additionally, SnS₂ can exhibit diverse morphologies, such as flower-like structures, plate-like forms, nanoparticles, and quantum dots, rendering it particularly intriguing for both study and application purposes.

Various reports have shown that SnS₂ can be utilized in numerous photocatalytic applications, such as hydrogen production,^{31–34} wastewater treatment,^{35–39} antibacterial and antifungal activity,^{40–43} and air purification.⁴⁴ For example, Yu *et al.*⁴⁵ first reported the photocatalytic hydrogen production of SnS₂ nanosheets under solar light simulation with H₂ evolution activity of 1.06 mmol h⁻¹ g⁻¹. In the case of waste water remediation, there are numerous reports for different kinds of waste water such as heavy metals,^{46–48} organic dyes,^{49,50} and pharmaceutical waste.^{51–53} For antibacterial activity, SnS₂ has been applied to inactivate *S. aureus* and *E. coli* as reported by Fakhri *et al.*,^{40,41} and for air purification SnS₂ has been reported to be able to remove various VOC gasses such as trichloroethylene⁵⁴ and formaldehyde⁵⁵ as well as NO_x gas.⁴⁴ Other applications have also been reported by researchers such as CO₂ reduction and nitrogen fixation which make SnS₂ a promising photocatalyst candidate. (For detailed summaries see sub-section 2.2.)

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As a photocatalyst, some characteristics of SnS₂ have become key for an excellent performance. These include low recombination and hole rates,^{56–59} a high specific surface area,^{60,61} high-reactivity facets,^{62,63} and good stability.^{64,65} Despite the confirmed photocatalytic performance of SnS₂, there are still limitations that hinder its efficiency, such as a high recombination of electrons and holes,^{66,67} and low specific surface area. Consequently, many researchers have focused on modifying SnS₂ to meet the requirements.

Several strategies have been widely adopted to enhance photocatalytic performance, including morphological modification,^{68,69} structural engineering,^{70,71} and heterostructure formation.^{72–74} Each of these strategies has unique advantages. For instance, morphological modification can alter the surface structure of SnS₂, influencing parameters such as surface area and facet arrangement, which are beneficial for improving photocatalytic performance. Structural engineering involves disturbing the crystal structure of SnS₂ through various treatments, including vacancy generation, doping addition, and strain engineering. These treatments effectively decrease electron and hole recombination, thereby enhancing photocatalytic performance. Noble metal doping, such as Au, Pt, and Ag, is a well-known method for improving photocatalytic performance due to its ability to act as an electron acceptor, slowing down the recombination of electrons and holes.⁷⁵ Meanwhile, heterostructure formation can generate a synergistic effect between one sample and another, effectively improving the photocatalytic performance.⁷⁶

Presently, there are several reports that have reviewed the photocatalytic performance of SnS₂. Sharma *et al.*⁷⁷ reviewed the photocatalytic performance of SnS₂ related to energy conversion and a heterostructure perspective approach, while Guo *et al.*⁷⁸ reviewed the advancement of the photocatalytic performance of SnS₂ by comparing pure SnS₂ and SnS₂-based composite. While these two reviews have addressed some

important aspects of improvement strategies, the detailed explanation and fundamental mechanism of the photocatalytic enhancement of SnS₂ have not been thoroughly explored. In this study, we systematically investigate the current advancement strategies for enhancing photocatalytic performance, which can be classified into three categories: structural modification, morphology engineering, and heterostructure formation. We will elucidate the factors from each strategy that have the most significant impact on the photocatalysis of SnS₂. Furthermore, we will propose future perspectives for the development of photocatalytic studies on SnS₂ based on the existing literature. Fig. 1 illustrates three distinct strategies for enhancing photocatalytic performance across various applications, discussed in this review. The first strategy involves structural modification, encompassing vacancy generation and doping processes. The second strategy focuses on morphological manipulation, featuring a range of morphologies such as nanoflowers, nanoparticles, nanosheets, and quantum dots. The final strategy concerns heterojunction formation, encompassing Schottky junctions, heterojunctions types I and II, p–n junctions, as well as Z-scheme and S-scheme photocatalysts. Additionally, we address current challenges and offer insights into future developments.

2 SnS₂ properties and photocatalytic applications

2.1 Physical properties of SnS₂

Tin disulfide (SnS₂) belongs to the IV–VI chalcogenide materials and exhibits a golden-yellowish color, resembling the configuration of transition metal dichalcogenide (TMD) materials.⁷⁷ The atomic arrangement of SnS₂ consists of a Sn atom sandwiched between two hexagonally close-packed sulfur atoms. In layered SnS₂, planar threefold layers (TLs) are



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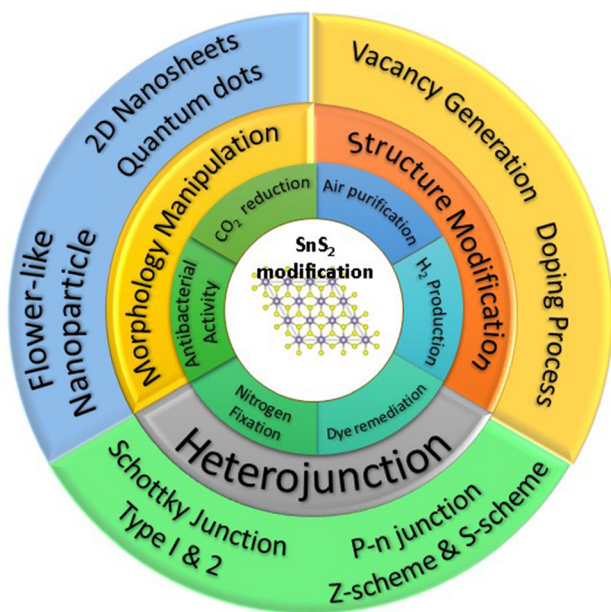


Fig. 1 Modification strategy of SnS₂ for improving photocatalytic performance.

formed, where strong covalent bonding exists in a plane, while weak van der Waals interactions dominate out of plane, as shown in Fig. 2(a).⁷⁹ The SnS₂ structure belongs to the covalent bonding group of $P\bar{3}m1$ (164) with lattice parameters $a = b = 0.3643$ nm and $c = 1.7683$ nm.⁷⁹ Like the other members of the TMD family, SnS₂ has a layered structure with several interesting properties such as a high optical absorption coefficient, tunable bandgap, low toxicity, and natural abundance, which makes it very promising for 2D material applications such as photocatalysis, photoanodes,⁸⁰ and gas sensors.²⁰ SnS₂ corresponds to an n-type semiconductor that possesses a narrow energy band gap of 2.0–2.4 eV. The band positions of SnS₂



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were determined by linear extrapolation of the band edges and secondary electron cut-off, to the baseline, and yielded an ionization potential ($E_{\text{vac}} - E_{\text{VBM}}$) of 6.44 ± 0.07 eV and an electron affinity ($E_{\text{vac}} - E_{\text{CBM}}$) of 4.16 ± 0.17 eV. A bandgap ($E_{\text{CBM}} - E_{\text{VBM}}$) of 2.28 ± 0.15 eV is thus found that corresponds to the indirect band gap as shown in Fig. 2(b). SnS₂ also has higher electron mobility (~ 230 cm² V⁻¹ s⁻¹), higher capacitance (645 mA h g⁻¹), and lower environmental noxiousness.⁸¹ The photocatalytic reaction process occurs when the band positions of the photocatalyst align with the redox potentials required for specific chemical reactions. For instance, the conversion of H⁺ to H₂ can occur at 0 V vs. NHE energy, where electrons in the conduction band with an energy lower than 0 V can reduce H⁺ to H₂. Conversely, the transformation of OH⁻ to OH[•] requires 1.93 V vs. NHE, necessitating holes with an energy higher than 1.93 V vs. NHE to oxidize OH⁻ to OH[•]. In this context, the band positions of SnS₂ fulfill both criteria, with the conduction band at -1.5 V vs. NHE and the valence band at 2.13 V vs. NHE,⁸² as illustrated in Fig. 2(c). Various chemical reactions may yield different potential energy requirements, thus influencing the applicability of SnS₂ in photocatalysis.

2.2 Photocatalytic applications of SnS₂

The utilization of SnS₂ as a photocatalyst was initially documented in 2007 by He *et al.*⁸³ who explored the formation of a heterostructure between SnS₂ and TiO₂ for methyl orange degradation. However, after this initial investigation, research on SnS₂ photocatalysis remained limited until 2017, when the number of publications on the subject notably increased, as illustrated in Fig. 3(a). Presently, the total number of publications on SnS₂ as a photocatalyst remains at less than 500, significantly lower compared with conventional photocatalysts such as TiO₂ or ZnO, and even other transition metal dichalcogenide (TMD) materials like MoS₂ and WS₂ (based on Scopus data). As a 2D photocatalyst material, SnS₂ has found applications across various domains, including air purification,^{44,55} wastewater treatment,^{47,84–87} hydrogen production,^{88–97} and antibacterial purposes,⁴¹ CO₂ reduction, and^{89,91,98–100} nitrogen fixation.^{92,94,101}

According to Scopus data, illustrated in Fig. 3(b), the predominant focus of photocatalytic applications involving SnS₂ is wastewater treatment, constituting 66% of the total research publications to date. Following closely, hydrogen production and CO₂ reduction make up 21% and 10% of the total publications on SnS₂ as a photocatalyst. The least explored applications include nitrogen fixation, antibacterial activity, and air purification. Wastewater treatment represents the most extensively studied application when investigating SnS₂ photocatalytic performance. Given the high demand for clean water in our daily lives, the removal of wastewater from water resources is of paramount importance. Various types of wastewater, containing pollutants such as organic dyes, heavy metals, pharmaceutical waste, and even uranium(vi), pose significant threats to our health. Until now (2024 January 10th), based on Scopus data there are 337 publications that have reported the use of SnS₂ photocatalyst for wastewater treat-



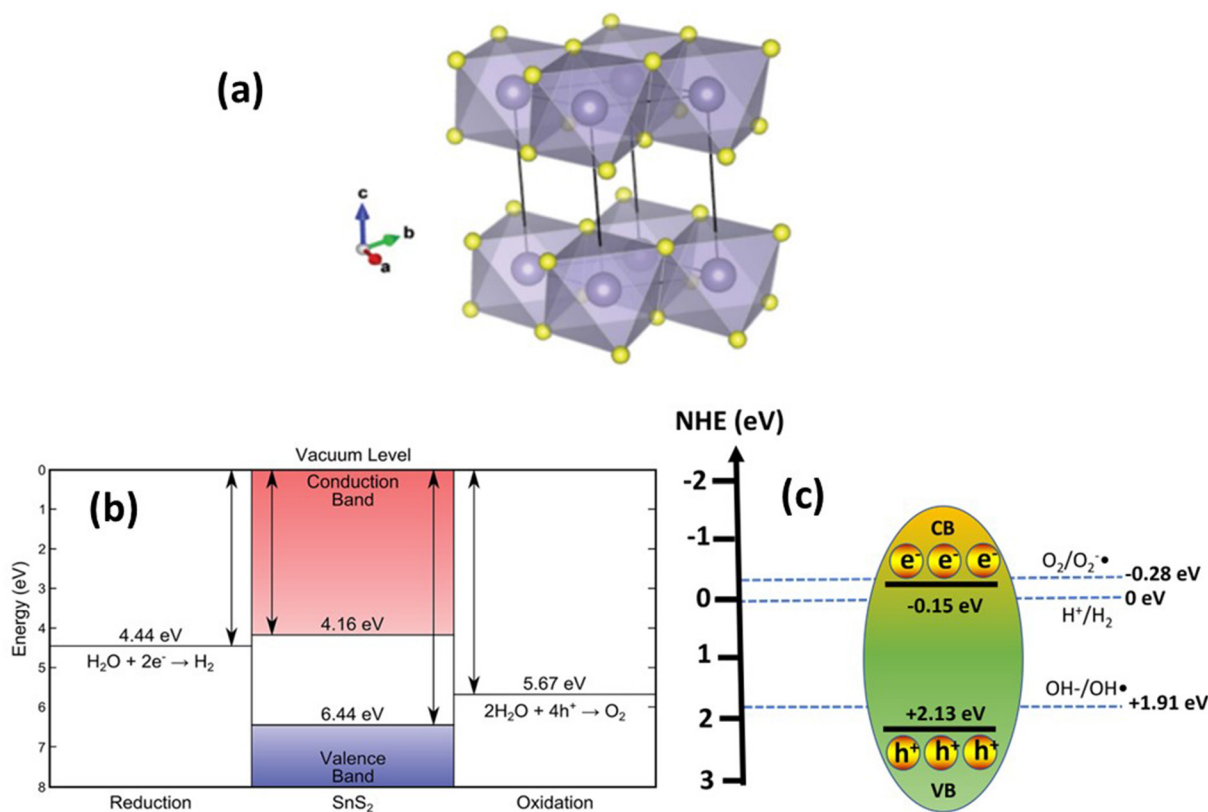


Fig. 2 (a) Crystal structure and (b) band position of SnS₂. (Reproduced from ref. 79 with permission from the Royal Society of Chemistry, copyright 2016.)⁷⁹ (c) Band position of SnS₂ for photocatalytic reactions.

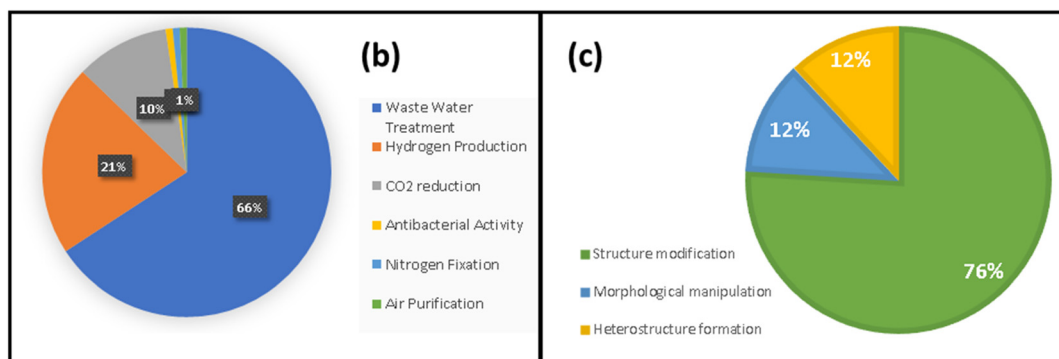
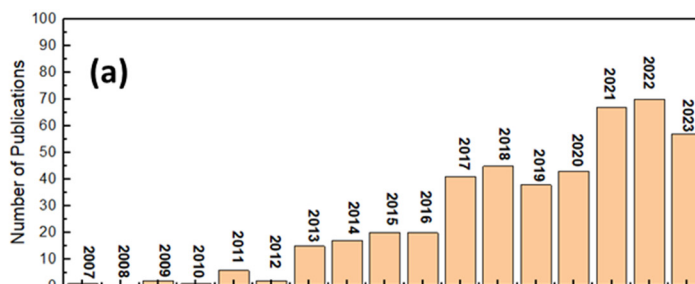


Fig. 3 (a) Number of SnS₂ photocatalytic publications; (b) summary of practical applications of SnS₂ photocatalytic. (c) Graphical illustration for photocatalytic improvement strategies for SnS₂.



ment under visible light irradiation and 192 of them are about organic dye removal such as methylene blue,^{102–104} methylene orange,¹⁰⁵ and rhodamine B (RhB).^{49,106} For heavy metals, hexavalent chromium (Cr(vi)) is the most studied waste water model for photocatalytic investigations of SnS₂, with 99 publications, which makes SnS₂ one of the most promising materials for Cr(vi) reduction under visible light irradiation. In the pharmaceutical waste field, tetracycline is the most studied wastewater, with 11 reports showing a good performance for tetracycline removal by a SnS₂ photocatalyst. Table 1 shows the summary of various types of application of SnS₂ photocatalysts. It can be seen that the SnS₂ photocatalyst exhibits great potential for various types of photocatalytic application.

2.2.1 CO₂ reduction. Recently, CO₂ conversion has attracted considerable attention for photocatalytic results due to its great potential for reducing CO₂ pollution alongside producing more valuable products.¹⁰⁷ Sun *et al.*¹⁰⁸ initiated efforts to reduce CO₂ through photocatalytic processes by SnS₂ under visible light irradiation. Their research successfully converted CO₂ into CO with a yield of approximately 80 ppm after 8 hours of xenon irradiation. Subsequently, Li *et al.*¹⁰⁹ explored the influence of SnS₂ morphology on CO₂ reduction *via* photocatalytic reactions. They examined two distinct morphologies—tablet-like and flower-like—and observed the conversion of CO₂ gas into CH₄ and CO gases. Notably, each morphology exhibited a unique behavior in CO₂ reduction. The flower-like morphology demonstrated a higher CH₄ yield at 97.5 μmol g⁻¹ compared with CO gas at 75.1 μmol g⁻¹, while the tablet morphology showed stronger CO production at 61.6 μmol g⁻¹ relative to CH₄ gas at 17.2 μmol g⁻¹. This underscores the impact of SnS₂ morphology on CO₂ reduction, which is directly linked to variations in charge separation processes, surface properties, and other factors (further elaborated in the subsequent discussion). Recent endeavors have focused on optimizing and enhancing CO₂ reduction through various techniques, including the formation of Z-scheme heterojunctions,¹¹⁰ type II heterojunctions, vacancy generation,¹⁰⁶ and doping processes.⁹⁷ Among these strategies, Z-scheme heterojunctions have shown significant promise in improving the CO₂ reduction process. Wang *et al.* reported the synthesis of a Z-scheme photocatalyst, SnO₂/SnS₂/Cu₂SnS₃, which achieved a CO₂ reduction to ethanol rate of 28.44 mmol g⁻¹ h⁻¹ as a byproduct. The fundamental process of CO₂ reduction can be described as follows: when subjected to irradiation, electrons and holes are generated. Electrons in the conduction band, possessing a band position more negative than the reduction potential of CO₂, participate in reactions that transform CO₂ into various chemicals. The specific transformation depends on the suitability of the conduction band energy for reducing CO₂ into alternative chemical compounds.

2.2.2 Nitrogen reduction. The exploration of SnS₂ photocatalytic applications in nitrogen fixation is still relatively limited, with only four studies investigating its potential in this area. Li *et al.*¹⁰¹ reported on the nitrogen fixation capability of SnS₂ using nanosheet and nanoplate morphologies. They found that the NH₄⁺ concentration reached 40.5 μmol g⁻¹

and 110.1 μmol g⁻¹ on SnS₂-plate and SnS₂-sheet samples, respectively, after 4 hours, with average reaction rates of 10.1 μmol g⁻¹ h⁻¹ and 27.5 μmol g⁻¹ h⁻¹, respectively. These findings highlight the sensitivity of SnS₂ photocatalytic performance to particle morphology. Three other reports discussed the nitrogen fixation application of SnS₂ in heterostructure forms. For instance, Nie *et al.*⁹² demonstrated the formation of 1D/2D heterojunctions of SnS₂@MoO₃ as an efficient catalyst for visible photoelectric nitrogen fixation. Compared with pure MoO₃ and SnS₂, the nitrogen fixation performance improved by 2.3 and 4.6 times, respectively, and the Faraday efficiency was enhanced by 2.0 and 3.5 times, respectively. Another study by Lan *et al.*⁹⁴ also explored the formation of a Z-scheme heterojunction Bi/Bi₂S₃/SnS₂ for nitrogen fixation applications. The results showed that the Z-scheme formation of Bi/Bi₂S₃/SnS₂ successfully enhanced NH₄⁺ production, yielding approximately 96.4 μmol g⁻¹ h⁻¹, which is 1.9 and 3.0 times higher than that of SnS₂ and Bi/Bi₂S₃, respectively. The basic mechanism of the nitrogen fixation applications can be explained as follows: upon irradiation, the electrons and holes are generated. After that the photogenerated holes reacted with water to produce H⁺ ions and in the final reaction N₂, H⁺, and electrons reacted to produce NH₃ or NH₄⁺ ions.¹⁰¹

2.2.3 H₂ production. Hydrogen production has become one of the popular applications for photocatalytic research since the growth of green sources of energy for switching from fossil fuels.¹¹¹ Initial investigations into the hydrogen production capabilities of SnS₂ were conducted in 2014. Yu *et al.*⁴⁵ prepared monodisperse SnS₂ nanosheets for hydrogen production under visible light irradiation, achieving a notable hydrogen production yield of approximately 0.566 mmol g⁻¹ h⁻¹, which is considerably high for a visible light photocatalyst. The fundamental mechanism of H₂ production in photocatalytic reactions involving SnS₂ can be elucidated as follows: upon irradiation, the SnS₂ semiconductor generates electron-hole pairs. The photoexcited electrons reside in the conduction band and subsequently react with H⁺ to produce H₂ gas. However, for the generation of H₂ gas, it is essential that the reduction potential of electrons in the conduction band exceeds the reduction potential of H⁺/H₂ (0 V). The conduction band position of SnS₂ is influenced by various factors such as facet orientation and band gap. Huang *et al.*⁶⁹ demonstrated that the band position of SnS₂ is influenced by its facet orientation. Two facet orientations, (100) and (101), exhibit distinct performance characteristics in the photocatalytic production of H₂ gas. The (101) facet orientation is more favorable as it possesses a higher conduction band level than the reduction potential of H⁺/H₂, whereas the dominance of the (100) facet results in a lower reduction potential in the conduction band compared with H⁺/H₂. These findings suggest that the (101) facet of SnS₂ may be more conducive to H₂ production. Recent studies have focused on the development of S-scheme and Z-scheme photocatalysts (see Table 1) due to their efficient charge separation and band modulation properties.^{112–117}



Table 1 Photocatalytic data of SnS₂ for various types of application

Sample	Sample	Light source	Photocatalytic result	Additional information	Ref.
CO ₂ reduction	Cu-doped SnS ₂	420 nm LED	CH ₃ OH yield 0.99 mmol g ⁻¹		97
	S-vacancy SnS ₂	400 W metal halide lamp	CO yield = 2.44 μmol g ⁻¹ h ⁻¹ , 1.48 times that of pure SnS ₂		106
	SnS ₂ flower-like	300 W xenon light	CO production rate 8.91 μmol g ⁻¹ h ⁻¹		135
	Z-scheme SnO ₂ /SnS ₂ /Cu ₂ SnS ₃	300 W xenon lamp	Ethanol yield 28.44 mmol g ⁻¹ h ⁻¹ 3 times higher than SnS ₂		110
	SnS ₂ /SnO ₂ nanoflower	300 W xenon lamp	CO yield 60.85 μmol g ⁻¹ h ⁻¹		136
	Partly oxidized SnS ₂	300 W xenon lamp	CO yield 12.28 μmol g ⁻¹ h ⁻¹		67
	Carbon-doped SnS ₂	300 W halogen lamp	CH ₃ CHO yield 125.66 μmol per 100 g per 3 h		67
	Carbon-doped SnS ₂ SnS ₂ QDs/g-C ₃ N ₄	150 W halogen lamp 300 W xenon lamp (420 nm cut-off)	CH ₄ yield 0.64 μmol g ⁻¹ h ⁻¹ CH ₃ OH yield 2.3 μmol g ⁻¹ h ⁻¹		67 137
Hydrogen production	SnS ₂ -nanosheets	400 W mercury lamp	H ₂ production rate: 0.566 mmol g ⁻¹ h ⁻¹		31
	2.5 mol% Ni-SnS ₂	400 W mercury lamp	1429 μmol g ⁻¹ h ⁻¹		138
	In ³⁺ -doped SnS ₂	300 W xenon lamp	H ₂ production rate: 470 mmol g ⁻¹ for 12 h		139
	S-scheme SnS ₂ /CdS	Simulated solar light (100 mW cm ⁻²)	360.75 μL h ⁻¹		116
	SnS ₂ /g-C ₃ N ₄	300 W xenon lamp (λ ≥ 420 nm)	1818.8 μM g ⁻¹ h ⁻¹		124
	SnS ₂ /CdS/TiO ₂	300 W xenon lamp (λ ≥ 420 nm)	97.14 μmol h ⁻¹ cm ⁻²		140
	SnS ₂ /BiVO ₄	100 mW cm ⁻² xenon lamp, 0.8 V	900 μmol cm ⁻²		117
SnS ₂ /Cd _{0.5} Zn _{0.5} S	300 W xenon lamp (λ ≥ 420 nm)	3.19 mmol g ⁻¹ h ⁻¹		141	
Air purification	Bi ₄ O ₅ Br ₂ -SnS ₂	Visible light	40% NO _x removal	NO _x abatement	44
	Ni-doped SnS ₂ /SnO ₂	8 W, 256 nm light	85%	Trichloroethylene	54
	SnS ₂ /TiO ₂	40 W fluorescence lamp	55% removal	Formaldehyde	55
Nitrogen fixation	SnS ₂ nanosheets	330 W xenon lamp	27.5 μmol g ⁻¹ h ⁻¹		101
	SnS ₂ /MoO ₃	300 W xenon lamp	30.04 μg h ⁻¹ mg ⁻¹		92
	Bi/Bi ₂ S ₃ /SnS ₂	300 W xenon lamp	96.4 μmol g ⁻¹ h ⁻¹		94
Antibacterial activity	SnS ₂ nanoparticles	500 W xenon lamp	75.28 ± 2.3%	<i>S. aureus</i> (ATCC 25923)	41
	SnS ₂ nanoparticles	500 W xenon lamp	78.95 ± 2.0%	<i>E. coli</i> (ATCC 13534)	41
	SnS ₂ nanoparticles	500 W xenon lamp	81.22 ± 1.1%	<i>E. coli</i> (ATCC 25922)	41
	SnS ₂	Visible light	50.39	<i>E. coli</i>	43
	Pr-doped SnS ₂	Visible light	71.03	<i>E. coli</i>	43
Waste water treatment	SnS ₂ nanoparticle	250 W xenon lamp (λ ≥ 420 nm)	k _{MO} : 0.0594 min ⁻¹	Methyl orange	105
	MoS ₂ /SnS ₂	Visible light	k _{MB} : 0.0594 min ⁻¹	Methylene blue	122
	SnS ₂ decorated biochar	Natural sunlight	k _{MB} : 0.013 min ⁻¹	Methylene blue	142
	S-Sn-S vacancy SnS ₂	300 W Xe lamp	k _{RHB} = 0.136 min ⁻¹	Rhodamine B	143
	CdZnS/SnS ₂ /SnO ₂	500 W xenon lamp	k _{RHB} = 0.0267 min ⁻¹	Rhodamine B	49
	Zr-doped SnS ₂	Larger than 420 nm	k _{Cr(VI)} = 0.059 min ⁻¹	Cr(VI)	144
	SnS ₂ nanoparticle	300 W xenon lamp	k _{Cr(VI)} = 0.00401 min ⁻¹	Cr(VI)	129
	SnS ₂ @SnO ₂	300 W xenon lamp	k _{Cr(VI)} = 0.0907 min ⁻¹	Cr(VI)	129
	SnS ₂ /SnS	Direct sunlight	k _{Cr(VI)} = 0.086 min ⁻¹	Cr(VI)	145
	SnS ₂ /CNTs	300 W xenon lamp (λ ≥ 420 nm)	k _{Cr(VI)} = 0.07750 min ⁻¹	Cr(VI)	146
	SnS ₂ nanoplates	Solar irradiance was assumed to be 221 W m ⁻²	k _{TC} = 0.001 min ⁻¹ , k _{CP} = 0.0005 min ⁻¹	Tetracycline and ciprofloxacin	134
	SnS ₂ /Bi ₂ WO ₆	Solar irradiance was assumed to be 221 W m ⁻²	k _{TC} = 0.027 min ⁻¹ , k _{CP} = 0.024 min ⁻¹	Tetracycline and ciprofloxacin	134
	SnS ₂ nanoplates	Natural sunlight	k _{CP} = 0.0024 min ⁻¹	Ciprofloxacin	132
	SnS ₂ /BiVO ₄	Natural sunlight	k _{CP} = 0.018424 min ⁻¹	Ciprofloxacin	132
	SnS ₂ nanoplates	LED light 500 W m ⁻²	k _{GM} = 0.1492 min ⁻¹	Gemifloxacin	37
	Platinum/polypyrrole-carbon black/SnS ₂	LED light 500 W m ⁻²	k _{GM} = 0.0429 min ⁻¹	Gemifloxacin	37
	SnS ₂ nanoparticle	Visible light (λ ≥ 420 nm)	k _{TC} = 0.00129 min ⁻¹	Tetracycline	38
	SnS ₂ /Bi ₂ MoO _{6-x}	Visible light (λ ≥ 420 nm)	k _{TC} = 0.01245 min ⁻¹	Tetracycline	38
	SnS ₂ nanoplates	Solar simulator	k _{NP} = 0.014 min ⁻¹	Norfloxacin	147



2.2.4 Antibacterial activity. The SnS₂ photocatalyst has also been explored for antibacterial and antifungal applications. Fakhri *et al.*⁴¹ demonstrated the removal of *S. aureus* (ATCC 25923), *E. coli* (ATCC 13534), and *E. coli* (ATCC 25922) with removal rates of 75.28 ± 2.3%, 78.95 ± 2.0%, and 81.22 ± 1.1%, respectively. The basic mechanism of antibacterial removal by SnS₂ nanoparticles can be elucidated as follows: upon irradiation, electrons and holes are generated, and the photo-generated holes and electrons react with OH⁻ and O₂ to produce ·OH and O₂^{·-}, respectively. These radical species are then converted into H₂O₂, serving as potent oxidative agents for killing bacterial cells. In another study, Fakhri *et al.*¹¹⁸ investigated the impact of heterojunction formation of SnS₂/SnO₂ on antibacterial and antifungal activity. Their results indicate that SnS₂/SnO₂ exhibits superior bacteria-killing ability compared with pure SnS₂ and SnO₂, possibly due to the inhibition of the charge carrier separation of electrons and holes. Furthermore, antifungal activity was assessed against pathogenic *C. albicans*. The largest zones of inhibition for the control, SnO₂ nanoparticles, SnS₂ nanoplates, and SnO₂/SnS₂ nanocomposite are 5.89 mm, 13.20 mm, 14.01 mm, and 21.11 mm, respectively. A recent study by Ech-Chergui *et al.*⁴³ has corroborated the antibacterial activity of Pr-doped SnS₂ for combating *E. coli*. The Pr-doped SnS₂ exhibits enhanced bacteria-killing activity compared with SnS₂ alone, attributed to the smaller crystallite size providing more active sites for photocatalytic reactions.

2.2.5 Air purification. One of the major environmental issues is the deterioration of air quality due to heavy industry. Various types of air pollutant, including NO_x gas, VOCs, and other gases, are released into our environment, leading to serious health problems. Photocatalytic technology has been acknowledged as one of the solutions to address this issue.¹¹⁹ In the case of SnS₂ photocatalysts, SnS₂ has been reported to effectively remove various types of air pollutant, including NO_x gas, formaldehyde, and trichloroethylene, as shown in Table 1. However, the number of reports on the use of SnS₂ for air purification is very low, with fewer than 10 publications available. Moreover, all these reports involve heterojunction formation. Therefore, for further development, it is necessary to investigate the use of SnS₂ alone for this application.

2.2.6 Waste-water

2.2.6.1 Organic dyes. Organic dyes are prevalent pollutants in wastewater within our immediate environment, often originating from direct discharge by the textile industry, presenting significant environmental concern. Utilizing photocatalytic wastewater treatment emerges as a promising solution to address this issue, offering notable advantages such as cost-effectiveness, simplicity, and the capability to convert the chemical composition of organic dyes into harmless byproducts such as CO₂.^{120,121} In the realm of employing SnS₂ as a photocatalyst for organic dye degradation, extensive research has been conducted, targeting various dye types including methylene blue,¹²² methyl orange,¹²³ rhodamine B,^{124–126} crystal violet,¹²⁷ and others. Table 1 depicts the efficacy of SnS₂ photocatalysts in removing organic dyes. Notably, the data

demonstrate promising outcomes across all the tested dyes. The underlying mechanism of organic dye degradation can be elucidated as follows: upon irradiation, electron–hole pairs are generated, which subsequently react with O₂ and OH⁻ to yield highly reactive O₂^{·-} and ·OH radicals. These radicals possess strong oxidative capabilities, facilitating the decomposition of diverse synthetic dyes into CO₂ and water.

2.2.6.2 Heavy metals. In addition to organic dyes, the reduction of heavy metals represents a significant area of study in the photocatalytic application of SnS₂. Among various heavy metals, hexavalent chromium (Cr(vi)) has been extensively researched. Several strategies have been developed to enhance Cr(vi) removal, including vacancy generation,⁴⁸ doping formation,¹²⁸ and heterostructure formation.¹²⁹ The basic mechanism of Cr(vi) reduction involves the generation of electrons and holes upon irradiation, where the photoexcited electrons can directly decompose Cr(vi) ions into Cr(III). The photocatalytic ability of SnS₂ depends on the extent to which electrons react with Cr(vi) molecules, which may be attributed to increasing active sites or improving charge carrier separation efficiency. Table 1 presents several studies confirming the degradation of Cr(vi) using SnS₂ photocatalysts with varying capabilities. A detailed analysis will be provided in the subsequent section.

2.2.6.3 Antibiotics. Pharmaceutical waste, including compounds like tetracycline and others, presents a significant challenge in wastewater management, particularly with many pharmaceutical industries directly discharging their waste into water bodies. Consequently, effective treatment methods are imperative to mitigate the presence of these hazardous chemicals in water. Photocatalysis has emerged as a promising technique for efficiently removing pharmaceutical waste from water. However, the utilization of SnS₂ photocatalysts for antibiotic removal remains limited. Only a small number of research papers, fewer than 30, have explored the use of SnS₂ photocatalysts for this purpose, with most focusing on the formation of heterojunctions with SnS₂, such as Z-scheme and S-scheme configurations.^{42,104,130,131} For instance, Kumar *et al.*¹³² demonstrated the effectiveness of Z-scheme SnS₂/BiVO₄ in removing the antibiotic ciprofloxacin from aqueous solutions. They found that the optimized batch with 0.15SnS₂/BiVO₄ exhibited a significantly higher apparent rate constant ($k = 0.0184 \text{ min}^{-1}$) compared with BiVO₄ ($k = 0.0049 \text{ min}^{-1}$) and SnS₂ ($k = 0.0024 \text{ min}^{-1}$) alone. This improved performance was attributed to enhanced charge separation efficiency, enabling electrons from SnS₂ to directly react with O₂ to form O₂^{·-}, ultimately degrading ciprofloxacin into CO₂ and water.¹³³ Other antibiotics have also been shown to be degradable by SnS₂ photocatalysts. For example, Kumar *et al.*¹³⁴ reported the use of SnS₂/Bi₂WO₆ for the removal of ciprofloxacin and tetracycline. While SnS₂ alone exhibited minimal removal capabilities for these antibiotics, the combination with Bi₂WO₆ significantly enhanced the degradation rates. These findings suggest that SnS₂ photocatalysts alone are insufficient for effectively decomposing various types of pharmaceutical waste. One of the key limitations is the rapid recombination of electron–hole



pairs, which impedes the reaction of electrons with O₂ to generate reactive oxygen species (ROS). Therefore, the development of hybrid photocatalytic systems, combining SnS₂ with other materials, holds promise for improving the efficiency of pharmaceutical waste removal from wastewater. Continued research in this area is crucial for advancing sustainable water treatment technologies.

Although there are many publications about photocatalytic performance with different applications, we can categorize their strategies into three different categories.

(1) Structural modification: This strategy involves altering the structure of SnS₂ to enhance its photocatalytic properties. This could include changes in the composition, arrangement of atoms, or other structural aspects.

(2) Morphological manipulation: Morphological manipulation focuses on changing the physical shape or form of SnS₂ to improve its photocatalytic performance. This might involve controlling the particle size, shape, or surface structure.

(3) Heterostructure formation: This approach involves creating composite materials by combining SnS₂ with other materials, possibly forming interfaces or layered structures. The high percentage suggests that researchers find this strategy particularly effective or promising for enhancing photocatalytic performance.

Based on these three classifications, the quantitative analysis has been made and plotted in Fig. 3(c) based on the Scopus data until now. The structural modification and morphological manipulation SnS₂ strategies are just 12% each of the total of publications of SnS₂ photocatalytic. On the other hands, the heterostructure formation exceeded more than 75%, which indicates, at the current stage, that the heterostructure formation of SnS₂ is the most popular strategy for improving strategies, which probably attributed to the unique characteristics of composite structures which imply to the better performance.

3 Photocatalytic improvement strategies

3.1 Structural modification

Structural modification is a valuable strategy for enhancing photocatalytic performance. By adjusting the structural properties of the photocatalyst, it is possible to improve its electronic and surface characteristics for a range of photocatalytic applications, including optimizing light utilization, enhancing charge transfer kinetics, and creating additional active sites.^{148,149} Several established techniques can be employed to modify the structure of the photocatalyst, including vacancy generation and doping. These methods can increase the photocatalyst's charge separation efficiency and surface reactivity by raising the adsorption energy.

3.1.1 Vacancy generation. Vacancies in 2D materials, such as graphene and transition metal dichalcogenides (TMDs), have garnered significant attention due to their easy formation in these materials. In addition to expediting the charge separ-

ation, the generation of vacancies in 2D materials like SnS₂ can increase the number of active sites, facilitating reactant adsorption and thereby implying photocatalytic improvements. In the case of SnS₂, sulfur vacancies have been well established as effective methods for modifying its structure in photocatalytic applications.

The synthesis of sulfur vacancy-generating SnS₂ was reported by Qiang *et al.*,⁴⁸ prepared *via* hydrothermal reaction for Cr(vi) reduction under visible light irradiation. They observed that adjusting the molar ratio between the Sn and S precursors during synthesis is crucial for controlling the defect concentration on SnS₂ nanoparticles. A higher Sn to S ratio indicates a higher defect content, confirmed by EPR analysis (Fig. 4(a)). An intense resonance signal from EPR analysis for the sample with a Sn to S ratio of five was observed, indicating a higher defect concentration compared with the sample with a lower Sn to S ratio due to sulfur vacancy. Photocatalytic Cr(vi) reduction also demonstrated that the sample with a high defect content exhibits the best photocatalytic ability (Fig. 4(b)). This phenomenon was attributed to the prolonged charge carrier lifetime of the sulfur vacancy sample, as demonstrated in the photocurrent analysis of all the prepared samples (Fig. 4(c)). SnS₂ nanosheets with an optimal amount of sulfur vacancies exhibited a superior photoreduction rate of Cr(vi) (100% in 20 min), roughly 18.09 times higher than that of pure SnS₂ (Table 2), indicating vast prospects for photocatalytic wastewater treatment. The presence of sulfur vacancies in the crystal structure of SnS₂ serves to create an electron trap between the valence band and conduction band, effectively slowing down the recombination of electrons and holes. In that case the trapped electrons are still able to generate chemical reactions with Cr(vi) ions to change into less hazardous ions (Cr(III)).

Sulfur vacancy generation is not only suitable for Cr(vi) reduction but also for other applications, including CO₂ reduction and phenol degradation. In a recent publication, Zhang *et al.*¹⁵⁰ employed a synthesis method using L-cystine as the sulfur precursor to generate S-vacancies in SnS₂ for CO₂ reduction (see Fig. 4(d)). Their investigation demonstrated a significant enhancement in the CO₂ reduction efficacy of SnS₂ with S-vacancies, showcasing a 14.3-fold improvement compared with commercially available SnS₂ products. Their objective was to elucidate the underlying mechanism of CO₂ reduction and discern the role of Sn⁰ atoms in CO₂ adsorption. Through DFT analysis, it was observed that the CO₂ molecule approached the surface of Sn/SnS₂, where the S_x²⁻ ions in proximity transferred photogenerated electrons to the oxygen atom of the CO₂, thereby activating the CO₂ molecule and facilitating adsorption. Sn⁰ particles functioned as electron transport bridges, enabling the further transfer of photogenerated electrons to the S_x²⁻ ions, consequently reducing the adsorption energy barrier and accelerating the reaction rate. In a parallel investigation, Li *et al.*¹⁰⁶ explored variations in the Sn precursor's valence state (Sn²⁺ and Sn⁴⁺) to induce S-vacancies in the crystal structure. Based on their investigation, the Sn²⁺ precursor yielded a higher S-vacancy content than the Sn⁴⁺ precursor.



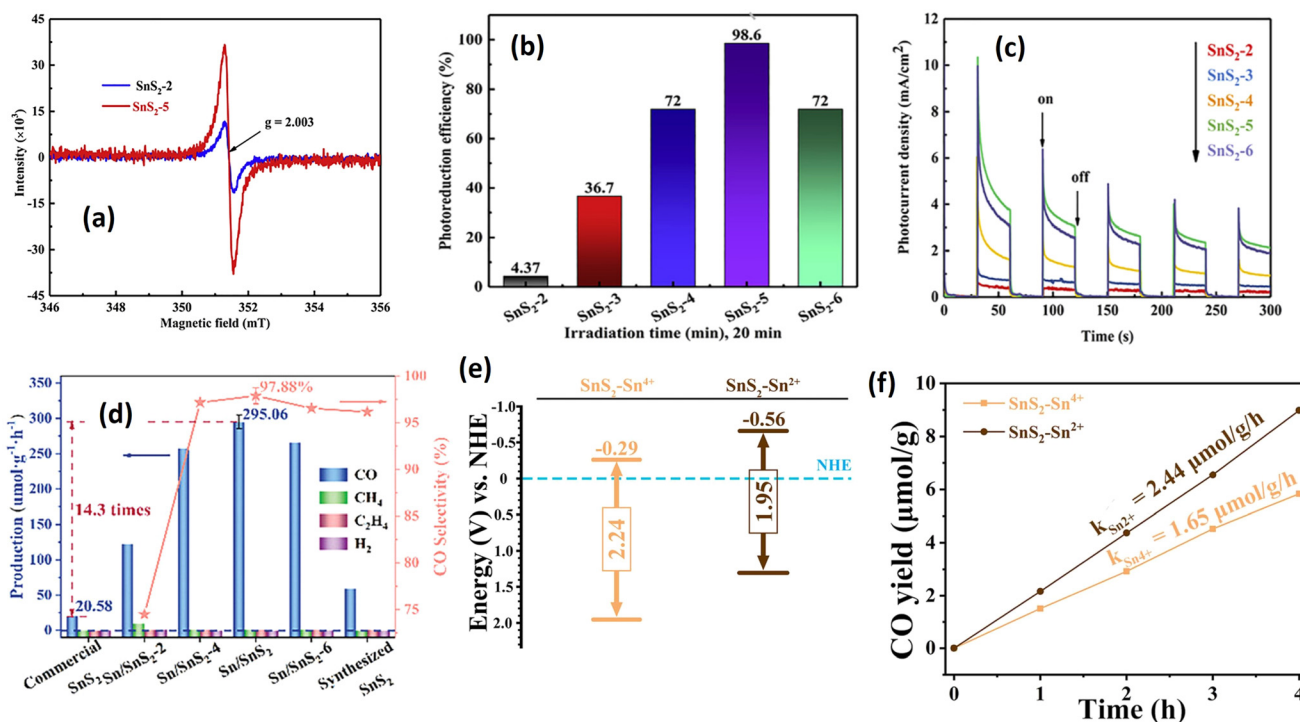


Fig. 4 (a) Room-temperature EPR spectra of SnS₂-2 and SnS₂-5. (b) Comparison of Cr(vi) photoreduction over SnS₂-*X* (*X* = 2, 3, 4, 5, 6) under visible light irradiation. (c) Photocurrent response of SnS₂-*X* (*X* = 2, 3, 4, 5, 6).⁴⁸ (Reproduced from ref. 48 with permission from Elsevier, copyright 2020.) (d) Photocatalytic CO₂ reduction performance of vacancy generated SnS₂ prepared at different L-cystine concentrations (−1, −2, −4, −6 mol).¹⁵⁰ (Reproduced from ref. 150 with permission from Wiley, copyright 2023.) (e) Energy band of SnS₂-Sn⁴⁺ and SnS₂-Sn²⁺. (f) CO₂ reduction of SnS₂-Sn⁴⁺ and SnS₂-Sn²⁺.¹⁰⁶ (Reproduced from ref. 106 with permission from Elsevier, copyright 2023.)

sor, with a determined vacancy ratio (*via* EPR analysis) of 4 : 1 for Sn²⁺ : Sn⁴⁺. This increase in S-vacancy quantity resulted in a band position shift, as illustrated in Fig. 4(e), leading to a more potent CO₂ reduction capability. Photocatalytic CO₂ reduction revealed a 1.47 times enhancement with the Sn²⁺ precursor compared with Sn⁴⁺, attributed to the heightened S-vacancy concentration (see Fig. 4(f)). The enhanced photocatalytic performance following an increase in S-vacancy concentration can be attributed to several factors, including improved light absorption and increased efficiency in the separation of photogenerated charge carriers. Beyond S-vacancies, other vacancy types, such as Sn-S vacancies and S-Sn-S vacancies, may exist, imparting distinct effects on the physical properties and photocatalytic performance of SnS₂. Guo *et al.*¹⁴³ addressed these defects by introducing a specific amount of 2,2'-bipyridine during the hydrothermal reaction. Their research has confirmed several phenomena directly correlated with the improvements in photocatalysis. The presence of S-vacancies and S-Sn-S vacancies could decrease the band gap energy, resulting in an upward shift of the valence and conduction bands. This shift facilitates a stronger reduction ability, enabling more efficient reactions with O₂ molecules. Furthermore, the S-Sn-S vacancy demonstrates the highest adsorption energy, followed by Sn-S and S vacancies, respectively. This trend is also reflected in the electron transfer ability, with the S-Sn-S vacancy exhibiting an electron transfer

ability 4.7 times higher than that of S vacancy. These findings suggest that the presence of S-Sn-S vacancies is more favorable for improving photocatalytic performance, a conclusion also supported by Guo *et al.*¹⁴³ for phenol decomposition (Table 2).

3.1.2 Doping SnS₂. Doping is widely recognized as an effective method for altering the structure of photocatalysts to improve their photocatalytic performance. In the case of SnS₂, the doping process involves the replacement of one or more SnS₂ atoms with other atoms, including cations or anions. Transition metal elements such as Cu,^{97,151} Ce,¹⁵² Zr,¹⁴⁴ and others have been successfully integrated into the SnS₂ structure. For instance, Liu *et al.*¹⁵³ demonstrated the doping of Cu onto the SnS₂ layer *via* a hydrothermal reaction for H₂ production applications under visible light irradiation. They varied the Cu concentration from 0 to 10% to determine the optimal concentration. The results indicated that a 5% Cu concentration represented the optimal condition for water-splitting photocatalytic reactions, exhibiting a sixfold increase compared with 0% Cu, as depicted in Fig. 5(a). Interestingly, the Cu doping process induces defects in the crystal structure of SnS₂, as investigated by EPR analysis, as depicted in Fig. 5(b). The electron paramagnetic resonance (EPR) signal for Cu-doping exhibits a stronger signal than undoped SnS₂, confirming the presence of a vacancy defect originating from the S-vacancy. This result indicates that a S-vacancy can be gener-



Table 2 Summary of photocatalytic performance through vacancy engineering and doping formation of SnS₂

Vacancy type	Light source	Photocatalytic testing	Photocatalytic performance	Ref.
S-vacancy SnS ₂	500 W Xe lamp	50 mg L ⁻¹ Cr(vi) reduction	Degradation efficiency (100% in 20 min), roughly 18.09 times that of pure SnS ₂	48
S-vacancy SnS ₂	400 W metal halide lamp	(1 × 10 ⁻⁵ M) rhodamine B (RhB) degradation	$k_{\text{RhB}} = 0.0102 \text{ min}^{-1}$ for S-vacancy SnS ₂ , ~2.1 times that of SnS ₂	106
S-vacancy SnS ₂	400 W metal halide lamp	CO ₂ reduction (1 atm)	CO yield = 2.44 μmol g ⁻¹ h ⁻¹ . 1.48 times that of pure SnS ₂	106
S-Sn-vacancy SnS ₂	300 W Xe lamp	RhB, and methylene blue (MB) degradation	$k_{\text{RhB}} = 0.054 \text{ min}^{-1}$ (3.17 times that of pure SnS ₂), $k_{\text{MB}} = 0.043 \text{ min}^{-1}$ (3.07 times that of pure SnS ₂)	143
S-Sn-S-vacancy SnS ₂	300 W Xe lamp	RhB and MB degradation	$k_{\text{RhB}} = 0.136 \text{ min}^{-1}$ (8.00 times that of pure SnS ₂), $k_{\text{MB}} = 0.116 \text{ min}^{-1}$ (8.28 times that of pure SnS ₂)	143
C-doped SnS ₂	Visible light	CO ₂ reduction (acetaldehyde formation)	Acetaldehyde formation: 1256.6 μmol g ⁻¹ (228.47 times that of SnS ₂)	67
Cu-doped SnS ₂	420 nm LED	CO ₂ reduction (CH ₃ OH production)	Cu-doped SnS ₂ : 0.99 mmol g ⁻¹ h ⁻¹ (2.06 times that of SnS ₂)	97
Cu-doped SnS ₂	300 W xenon lamp	Hydrogen evolution	Cu-doped SnS ₂ : 1.37 mmol g ⁻¹ h ⁻¹ far exceeding 6 times that of pure SnS ₂	153
Ce-doped SnS ₂	Solar irradiation	10 and 20 ppm of methyl orange (MO)	$k_{\text{MO}} = 0.096$ and $k_{\text{MO}} = 0.080$ for 10 ppm and 20 ppm. SnS ₂ : $k_{\text{MO}} = 0.045$ and $k_{\text{MO}} = 0.035$ for 10 and 20 ppm	87
Zr-doped SnS ₂	Larger than 420 nm	40 mg L ⁻¹ Cr(vi) reduction	$k_{\text{Cr(vi)}} = 0.059 \text{ min}^{-1}$ nearly 2.2 times that of SnS ₂	144
In ³⁺ -doped SnS ₂	White LED 48 W	20 ppm Cr(vi) reduction	$k_{\text{Cr(vi)}} = 0.123 \text{ min}^{-1}$. 39.4 times that of SnS ₂	128
In ³⁺ -doped SnS ₂	300 W xenon lamp	Hydrogen evolution	H ₂ production rate: 470 mmol g ⁻¹ for 12 h. 1.95 times higher than SnS ₂	139
Sr-doped SnS ₂	250 W xenon lamp	10 ppm RhB degradation	$k_{\text{RhB}} = 0.00721 \text{ min}^{-1}$ (1.55 times that of SnS ₂)	154

ated not only by controlling the S-source but also through the doping process. Moreover, the average electron lifetime of Cu-SnS₂ is longer than SnS₂, as shown in Fig. 5(c). The average lifetime of Cu-doped SnS₂ is 0.358 s, which is slower than SnS₂, with 0.341 s, indicating that defects caused by the S-vacancy act as electron acceptors, preventing electron and hole recombination. Upon exposure to visible light irradiation, electrons transition from the valence band to the conduction band. As defects are introduced, these electrons migrate to defect levels and promptly engage in reactions with adsorbed H₂O, rather than undergoing recombination with holes in the valence band. Additionally, compared with pristine SnS₂, Cu doping enables a broader range of visible light absorption, suggesting increased electron involvement in the redox reaction. Furthermore, due to the vacancies generated by Cu doping, the crystallinity decreases, and surface area increases. These factors are also beneficial for enhancing the photocatalytic reaction. Another study by Govindan *et al.*⁸⁷ conducted successful cation doping, introducing Ce atoms into the SnS₂ structure using a microwave-assisted method. In their research, they observed the formation of Sn-vacancies instead of S-vacancies. Additionally, alongside S-vacancy formation, they noted an increase in strain on the crystal structure of SnS₂ after Ce doping, measuring approximately 3.5×10^{-3} , which was higher than without Ce doping (2.3×10^{-3}). This behavior can be attributed to the higher atomic radius of Ce³⁺/Ce⁴⁺ (0.102 nm) compared with Sn⁴⁺ (0.07 nm). The modification of the SnS₂ structure with Ce doping was confirmed to enhance the photocatalytic performance, demonstrated by conducting methyl orange degradation under solar irradiation. The results indicate that Ce-doped SnS₂ improved the degradation of

methyl orange twice as fast as pristine SnS₂ (Table 2). The higher photocatalytic performance of Ce-doped SnS₂ could be attributed to the decreased crystallite size resulting from significant strain and a greater quantity of Sn vacancies. These factors facilitate a greater number of photon pathways within the SnS₂ host structure, consequently enhancing the photocatalytic activity for decomposing methyl orange molecule into CO₂ and water. In contrast to the above analysis, Park *et al.*¹²⁸ proposed a different mechanism to explain the doping behavior of In³⁺-doped SnS₂. They proposed that instead of the S-vacancy serving as an electron acceptor, the In³⁺ ions serve as electron acceptors, effectively improving the charge carrier separation efficiency as shown in Fig. 5(d). The photocurrent analysis of In³⁺-SnS₂ shows strong photoelectron excitation compared with SnS₂, which is directly correlated with the luminescence analysis in Fig. 5(e), indicating a slower recombination of electrons and holes after In³⁺ doping. The photocatalytic activity of In³⁺-doped SnS₂ tested for Cr(vi) reduction (Fig. 5(f)) shows a superior performance compared with SnS₂ and standard P-25 under visible light irradiation, providing clear evidence of the effect of In³⁺ in enhancing photocatalytic performance through extending the electron-hole lifetime. However, the effect of In³⁺ has been strengthened by Li *et al.*¹³⁹ who proposed that In³⁺ is responsible for S-vacancy generation, acting as an electron trap to improve hydrogen generation, with a hydrogen production rate of about 470 μmol g⁻¹ for 12 h as shown in Table 2.

Doping in SnS₂ structures, not only within the transition metal family but also through carbon infusion, has been demonstrated to markedly enhance photocatalytic efficacy. In a study by Shown *et al.*,⁶⁷ carbon-doped SnS₂ was investigated



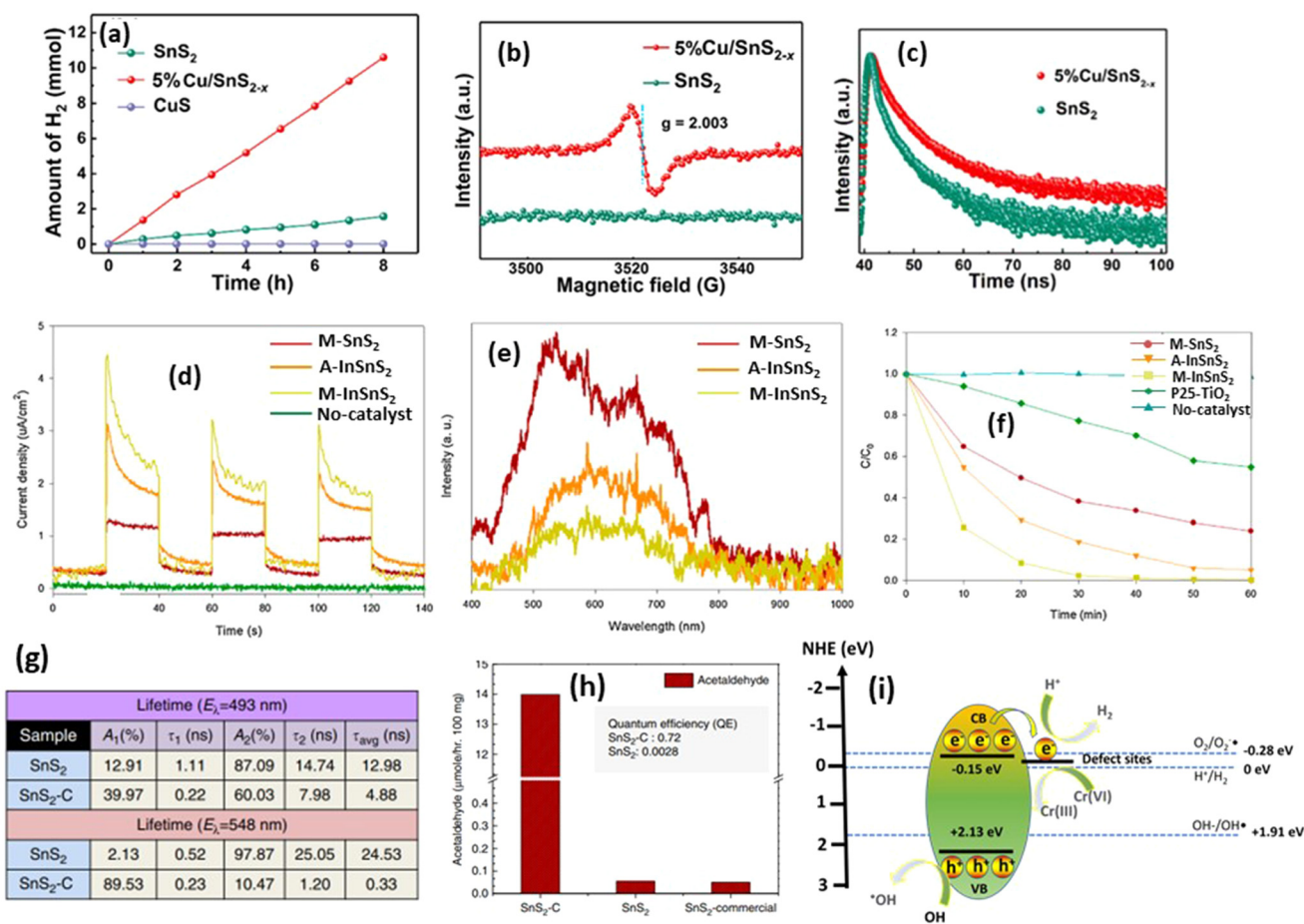


Fig. 5 (a) Photocatalytic water reduction capability of SnS₂, CuS, and Cu-doped SnS₂. (b) EPR spectra of SnS₂ and Cu-SnS₂, and (c) time-resolved photoluminescence spectra of SnS₂ and Cu-doped SnS₂.⁹⁷ (Reproduced from ref. 97 with permission from the Royal Society of Chemistry, copyright 2023.) (d) Photocurrent measurements of In³⁺-doped SnS₂. (e) Photoluminescence measurement of In³⁺-doped SnS₂. (f) Cr(VI) photodegradation by using In³⁺-doped SnS₂.¹²⁸ (Reproduced from ref. 128 with permission from Elsevier, copyright 2017.) (g) Charge life-time of C-doped SnS₂. (h) Photocatalytic CO₂ reduction into acetaldehyde by using carbon-doped SnS₂.⁶⁷ (Reproduced from ref. 67 with permission from Springer Nature, copyright 2018.) (i) Photocatalytic reaction mechanism of S-vacancy on the SnS₂ structure.

for its capacity to catalyze CO₂ reduction to acetaldehyde gas under 300 W halogen irradiation. The carbon doping regimen involved the use of L-cystine as both a sulfur and carbon source. The results revealed that carbon doping induced macrostrain within the SnS₂ lattice, leading to distinct photo-physical properties. Moreover, carbon doping facilitated the adsorption of CO₂ molecules on the surface, with a relatively low dissociation barrier observed. In Fig. 5(g), the observed average lifetimes for SnS₂-C are 4.88 and 0.33 ns, markedly shorter than the 12.98 and 24.53 ns for SnS₂ at 493 and 548 nm, respectively. This decrease in lifetime in SnS₂-C implies the presence of a nonradiative pathway, characterized by the delocalization of electrons from SnS₂ to C, thereby facilitating effective carrier separation. Additionally, the resulting carbon-doped SnS₂-C exhibited smaller nanosheets composed of only a few atomic layers, thereby reducing the charge diffusion times compared with pure SnS₂. These findings collectively underscore the exceptional quantum efficiency of carbon-doped SnS₂, which demonstrated an improvement of

over 257 times compared with pristine SnS₂. This enhanced quantum efficiency manifested as a remarkable increase in acetaldehyde production on carbon-doped SnS₂, outperforming pure SnS₂ by 228.47 times (Fig. 5(h)).

Based on the results, structural modification can enhance photocatalytic performance by several factors including S-vacancy generation, strain engineering, and ion trapping which successfully impeded the recombination of electrons and holes through an electron trap. However, since the majority of the publications state that the S-vacancy was the most responsible species for improving the photocatalytic reaction, therefore the photocatalytic mechanism of SnS₂'s structural modification can be generally explained as follows (Fig. 5(i)): upon exposure to light, electrons from the valence band become excited to the conduction band of SnS₂. Without a vacancy state, the electron would easily recombine with holes in the valence band. However, due to the presence of a vacancy state below the conduction band of SnS₂, electrons tend to be trapped, providing sufficient time to initiate photocatalytic



reactions. These reactions have potential applications in various fields, including hydrogen production, wastewater treatment, and many others.

All these structural modification strategies exhibit varying degrees of improvement in photocatalytic applications, as summarized in Table 2. Among these strategies, the carbon doping of SnS₂ emerges as one of the most promising approaches, showing an enhancement ability over 228.47 times greater than pristine SnS₂ in the reduction of CO to acetaldehyde.⁶⁷ Additionally, In³⁺ doping in the SnS₂ structure demonstrates a superior photocatalytic performance for wastewater treatment, exhibiting an improvement degree approximately 39.4 times higher than pristine SnS₂.¹²⁸ This suggests that In³⁺ doping enhances the charge separation efficiency more effectively than other cation doping methods and direct S-vacancy formation. However, the underlying mechanisms of this phenomenon remain unclear, as each doping process generates S-vacancies in the system, and the role of the dopant itself is not fully understood. Therefore, one plausible explanation for this phenomenon may lie in the variation in vacancy content and surface properties of the SnS₂ after doping, including an increased surface area and other factors.

3.2 Morphological manipulation

Beyond the structural factor, the photocatalytic activity of SnS₂ is also influenced by its surface properties. These properties are intricately linked to changes in morphology, impacting the surface area, the active facets, and the reaction process for photocatalytic reactions. Researchers have explored various SnS₂ morphologies, each demonstrating distinct photocatalytic performances. These morphologies encompass flower-like structures, nanosheets, nanoparticles, quantum dots, and more.

3.2.1 Flower-like. The flower-like morphology of SnS₂ is one of the most common shapes observed. It is formed due to the arrangement of several hexagonal shapes, creating a flower-like structure, as shown in Fig. 6(a).¹⁵⁵ This shape is typically prepared through a direct hydrothermal reaction without additional surfactants. The size of the flower can be adjusted by varying the reaction conditions, such as reaction time, temperature, and precursor amounts. One advantage of the flower-like morphology is the abundance of active sites on the crystal structure's edges.¹⁵⁶ Sengodu *et al.*¹⁵⁷ have successfully synthesized three distinct morphologies of SnS₂: nano-

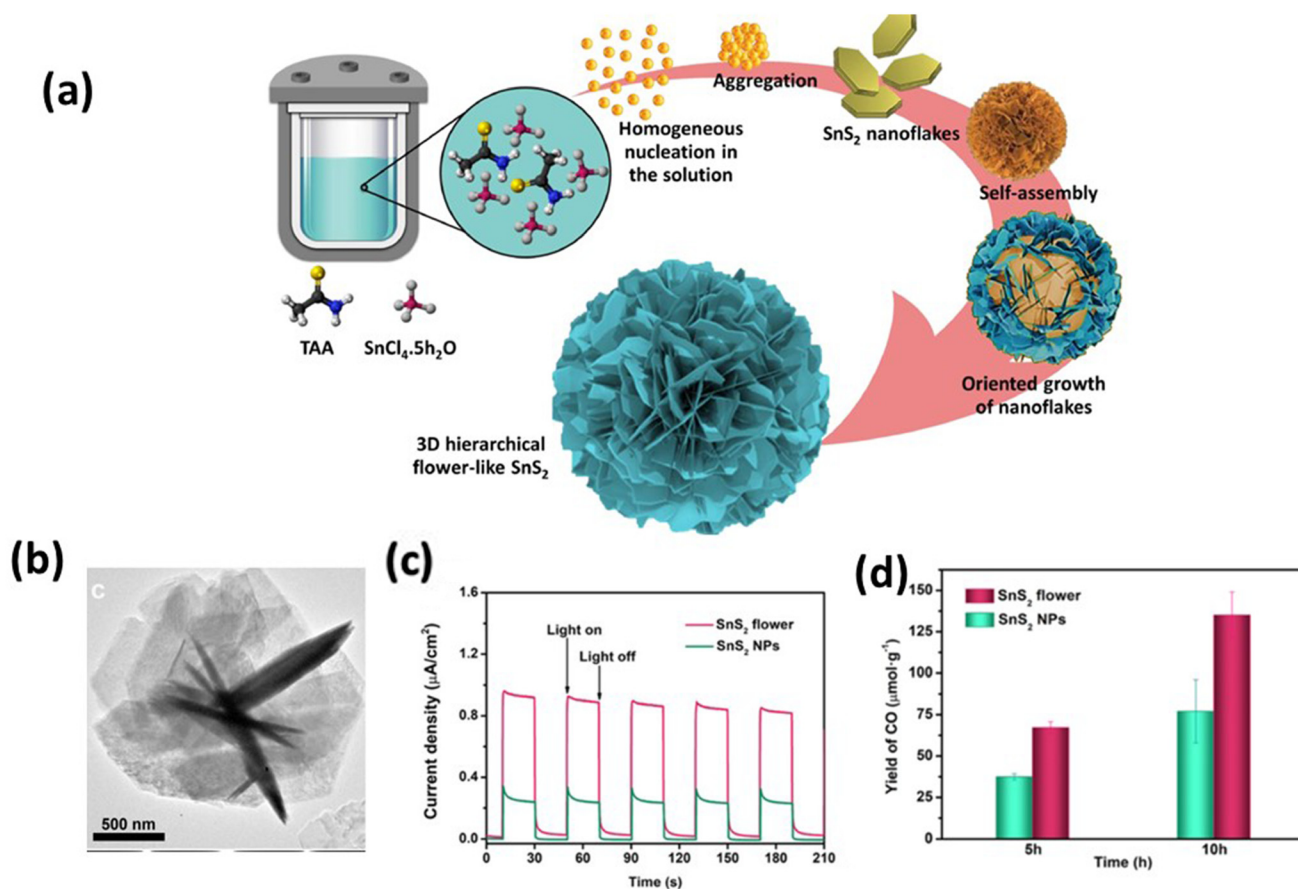


Fig. 6 (a) Formation of SnS₂ rose flower-like morphology.¹⁵⁵ (Reproduced from ref. 155 with permission from Elsevier, copyright 2021.) (b) TEM images of an individual SnS₂ flower. (c) Transient photocurrent and (d) comparative CO formation yield of the SnS₂ flower and SnS₂ NPs within the reaction times of 5 and 10 h, respectively.¹³⁵ (Reproduced from ref. 135 with permission from the American Chemical Society, copyright 2021.)



particles, nanoflowers (flower-like), and nanoplates, each exhibiting unique specific surface areas. Notably, the nanoflower morphology stands out with the highest specific surface area recorded, at $127.04 \text{ m}^2 \text{ g}^{-1}$, surpassing the values for nanoplates and nanoparticles, which are 96.40 and $69.4 \text{ m}^2 \text{ g}^{-1}$, respectively. Mondal *et al.*¹⁵⁸ corroborated the superior photocatalytic performance of the flower-like morphology, emphasizing its potential for environmental applications. Their study revealed that controlling the sulfur source precursor, particularly employing 0.06 g thioacetamide, led to the formation of nanoflowers, while 0.1 g thioacetamide resulted in a distinct nano-yarn morphology. Moreover, the nanoflower morphology of SnS_2 exhibited a markedly higher specific surface area than the nano-yarn, proving advantageous for enhanced photocatalytic $\text{Cr}(\text{vi})$ reduction under visible light irradiation. Upon photocatalytic reactions, the nanoflower with a high specific surface area can absorb more molecules onto the surface, increasing the likelihood of reaction occurrence. Moreover, the higher specific surface area enhanced light absorption leading to a more efficient generation of electron and holes for reducing $\text{C}(\text{vi})$ into $\text{Cr}(\text{iii})$. Another report from You *et al.*¹³⁵ has investigated the advantages of a flower-like morphology over nanoparticles. They observed that the flower-like structures prepared by the hydrothermal reaction have an average size of $2.39 \pm 0.26 \mu\text{m}$ and a thickness of around 20 nm , which is favorable for carrier transport to the catalyst surface (Fig. 6(b)). The flower-like morphology has a surface area around $20.11 \text{ m}^2 \text{ g}^{-1}$ which is higher than for the nanoparticle, at $12.99 \text{ m}^2 \text{ g}^{-1}$. Hence, SnS_2 flowers can supply more surface-active sites, leading to higher adsorption and photocatalysis. The flower-like structure also can absorb more light, as illustrated in Fig. 6(c), due to the scattering and reflection of light within the three-dimensional structure. Also, a thin sheet structure can shorten the transmission path, promoting more carriers to migrate to the surface of the photocatalyst.¹³⁵ The photocatalytic CO_2 reduction performance is undoubtedly directly correlated with the properties, as shown in Fig. 6(d). The CO_2 reduction performance of the flower-like morphology is $135.13 \mu\text{mol g}^{-1}$ which is 1.75 times higher than SnS_2 nanoparticles.

3.2.2 2D nanosheets. Two-dimensional (2D) layered morphologies can be defined as morphologies with a large two-dimensional ratio and a low vertical ratio. Due to their unique physicochemical properties, 2D material-based photocatalysts are expected to offer intriguing features such as porous structures, high specific surface areas, good crystallinity, a rich variety of host-guest species, better charge carrier separation, and abundant surface-active sites.¹⁵⁹ As mentioned earlier, the TMD family exhibits 2D properties, providing an advantage for photocatalytic applications. In the case of the SnS_2 photocatalyst, several reports have successfully demonstrated the preparation of 2D layered structures for photocatalytic applications.^{45,71,160–163} Yu *et al.*⁴⁵ have reported monodisperse SnS_2 nanosheets for hydrogen generation photocatalysis. Fig. 7 (a and b) shows the SnS_2 hexagonal nanosheets of $0.8\text{--}1 \mu\text{m}$ with a thickness of about 22 nm . The nanosheets are

assembled by *ca.* 35–38 hexagonal SnS_2 layers (the experimental value of single-layer SnS_2 is determined to be 0.61 nm). Yu *et al.*⁴⁵ prepared the SnS_2 nanosheets using a hydrothermal method with polyvinyl pyrrolidone (PVP) as a controlling agent. PVP can act in aggregation prevention during the hydrothermal reaction, preventing the SnS_2 layers from stacking into flower-like morphologies. Moreover, the monodisperse SnS_2 layers can also be achieved in an ethanol solution. The photocurrent measurement indicates that SnS_2 nanosheets have high electron mobility, which is beneficial for the H_2 -production reaction compared with the P-25 standard (Fig. 7(c)). The H_2 -production testing result shows various morphologies that have been compared with SnS_2 nanosheets, including nanoplates and sphere particles. The H_2 production yield for the nanosheet sample is much higher than for nanoplates, and spheres, as well as P-25 (Fig. 7(d)). The enhanced photocatalytic activity of SnS_2 nanosheets may be attributed to their low thickness, which provides more active sites. Additionally, their short and convenient approach allows photogenerated electrons and holes to migrate efficiently to the reaction sites on the surface, thereby decreasing the recombination probability and enhancing the photocatalytic efficiency.⁴⁵

Another report from Liu *et al.*¹⁶³ reported the formation of 2D nanosheets by the sonication of SnS_2/n -propylamine in HNO_3 solution to exfoliate the bulk SnS_2 into SnS_2 nanosheets, as shown in Fig. 7(e). The exfoliation of inorganic (SnS_2)-organic hybrids that are composed of few- or single-layered inorganic targeting slabs sandwiched by a single-layer organic molecule *via* coordinated bonds. The inserting organic molecule can be removed by the ion exchange method. In this research, they also confirmed the improvement of the charge carrier lifetime investigated by photocurrent measurement compared with bulk materials as shown in Fig. 7(f). Moreover, the charge carrier lifetime was observed to be around 0.78 ns which is almost 10 times higher than bulk SnS_2 . The $\text{Cr}(\text{vi})$ removal under visible light irradiation shows a significant improvement compared with bulk SnS_2 (Fig. 7(g)) with a degradation rate about 0.01436 min^{-1} which is 6 times higher than SnS_2 bulk materials. It indicates that the improvement of charge carrier separation from SnS_2 nanosheets has a direct correlation with the photocatalytic performance. Moreover, since SnS_2 nanosheets are a layered structure, the heterostructure of SnS_2 with other 2D materials was also interesting and has been explored by many researchers. For example, Zhang *et al.*¹⁶² have prepared a 2D heterojunction $\text{SnS}_2/\text{MoS}_2$ photocatalyst for methylene blue degradation. In this work, they prepared two-step hydrothermal reactions, as shown in Fig. 7(h). At the initial stage the nanolayer of SnS_2 was prepared, and the MoS_2 layers grew on the SnS_2 surface using a hydrothermal reaction. The formation of a 2D heterostructure of $\text{SnS}_2/\text{MoS}_2$ nanosheets can increase the charge separation efficiency, which leads to an increase in the photocurrent signal, as shown in Fig. 7(i). The photocurrent density of $\text{SnS}_2/\text{MoS}_2$ was 15.20 mA cm^{-2} in the stable second cycle, which exhibited an 8.8 times enhancement compared with that of SnS_2 ($\sim 1.55 \text{ mA cm}^{-2}$) and a 23.1 times improvement com-



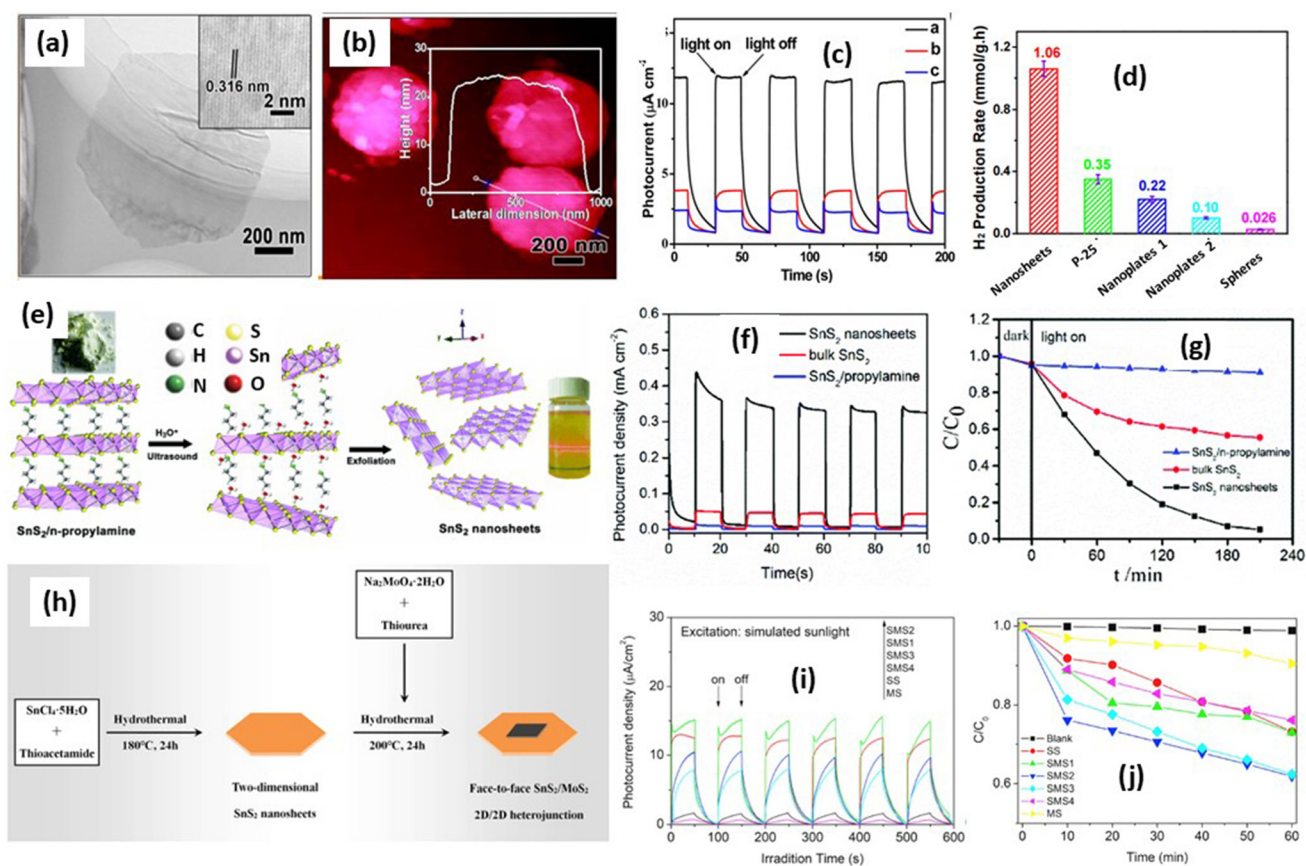


Fig. 7 (a) TEM image of an individual nanosheet. Inset is an HRTEM image. (b) High-magnification AFM image and corresponding height profile.⁴⁵ (c) The photoelectrochemical response of the samples at 0.8 V versus SCE electrode under 300 W Xe lamp illumination ($\lambda > 420$ nm). (Reproduced from ref. 45 with permission from the American Chemical Society, copyright 2014.) (d) Comparison of the photocatalytic hydrogen production rate of the samples. (e) Schematic illustration of the exfoliation of bulk SnS_2/n -propylamine into few-layer SnS_2 nanosheets. (f) Chronoamperometry at 0.6 V vs. $\text{Hg}/\text{Hg}_2\text{SO}_4$. (g) Photocatalytic reduction of 100 mg L^{-1} $\text{Cr}(\text{vi})$ with (1) the bulk SnS_2/n -propylamine hybrid, (2) bulk SnS_2 , and (3) SnS_2 nanosheets.¹⁶³ (Reproduced from ref. 163 with permission from the Royal Society Chemistry, copyright 2019.) (h) A schematic drawing of the synthetic process of $\text{SnS}_2/\text{MoS}_2$. (i) Photocurrent and (j) photocatalytic activities of SS (SnS_2), SMS1 ($\text{SnS}_2/\text{MoS}_2$ (Mo : Sn = 1%)), SMS2 ($\text{SnS}_2/\text{MoS}_2$ (Mo : Sn = 2.5%)), SMS3 ($\text{SnS}_2/\text{MoS}_2$ (Mo : Sn = 5%)), SMS4 ($\text{SnS}_2/\text{MoS}_2$ (Mo : Sn = 75%)) and MS (MoS_2) for methylene blue decomposition under UV-vis irradiation.¹⁶² (Reproduced from ref. 162 with permission from Elsevier, copyright 2019.)

pared with that of MoS_2 ($\sim 0.63 \text{ mA cm}^{-2}$). This phenomenon occurred due to the improvement of the charge separation efficiency due to the charge transfer between MoS_2 and SnS_2 , which can prolong the charge carrier lifetimes. The photocatalytic activity is also directly correlated with the photocurrent result as shown in Fig. 7(j). The best sample $\text{SnS}_2/\text{MoS}_2$ demonstrated an enhanced photocatalytic rate of 81% compared with SnS_2 .

3.2.3 Nanoparticles. Nanoparticles can be defined as particles with a diameter of less than 100 nm in size, offering numerous advantages over bulk-size materials, including an increased specific surface area, higher defect density, and improved dispersibility. Several reports have highlighted the photocatalytic performance of SnS_2 nanoparticles.^{16,25,41,105,164–167} Fakhri *et al.*⁴¹ reported the formation of SnS_2 nanoparticles by a hydrothermal reaction with the addition of 5 vol% acetic acid and 0.005 mol $\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$ and 0.02 mol thioacetamide. The produced nanoparticle was

about 20–30 nm as shown in Fig. 8(a) and (b). They conducted a photocatalytic investigation for ENRO degradation under 500 W xenon irradiation, revealing a degradation rate of approximately 0.0296 min^{-1} using SnS_2 nanoparticles as catalyst (Fig. 8(c)).

Zhang *et al.*¹⁰⁵ conducted a more detailed investigation into SnS_2 nanoparticle formation and its photocatalytic ability. They controlled the particle diameter and specific surface area of the SnS_2 particle by adjusting parameters such as sulfur source concentration and reaction temperature during the hydrothermal process. Increasing the reaction temperature resulted in larger particle sizes, as depicted in Fig. 8(d) and (e). For instance, particles prepared at 180 °C had an average size of around 142 nm, while those from the 120 °C hydrothermal reaction were about 23 nm in size. Conversely, higher sulfur source concentrations led to smaller particle sizes, illustrated in Fig. 8(e) and (f). For instance, the particle size decreased to 46 nm when the sulfur source concentration was



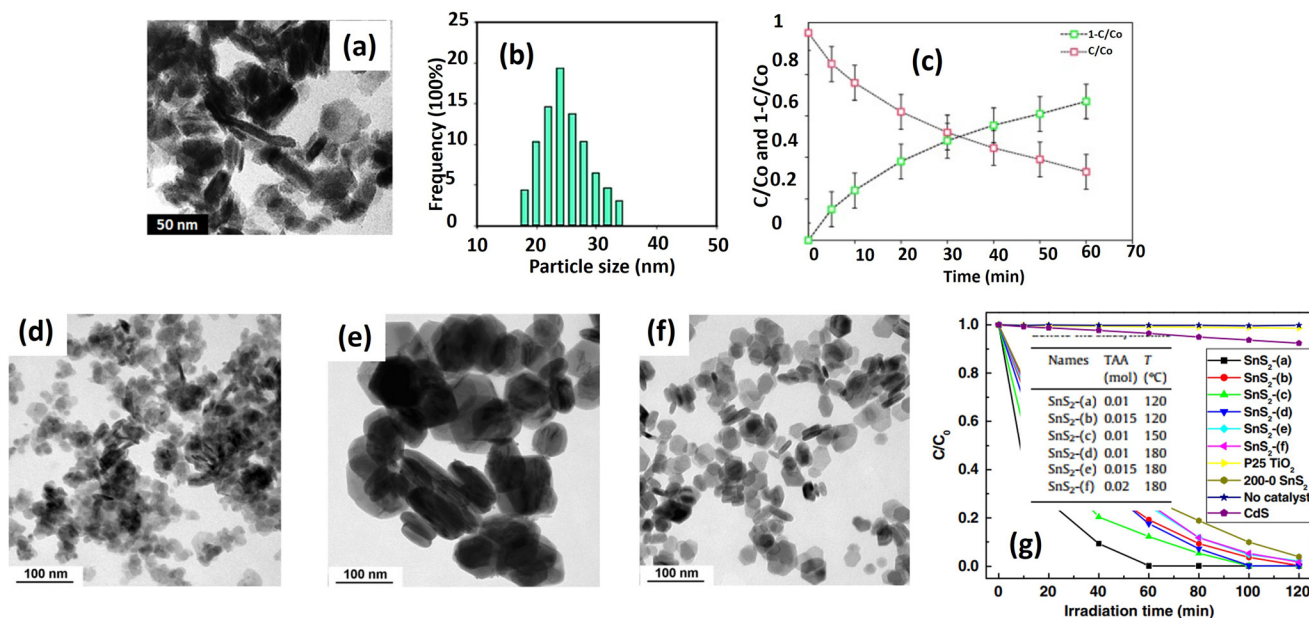


Fig. 8 (a) TEM image and (b) particle size distribution of SnS₂ NPs; (c) the photodegradation of ENRO by SnS₂ NPs under visible light.⁴¹ (Reproduced from ref. 41 with permission from Elsevier, copyright 2015.) TEM images of SnS₂ prepared with (d) 0.01 mol of TAA at 120 °C; (e) 0.01 mol of TAA at 180 °C; and (f) 0.02 mol of TAA at 180 °C; (g) photocatalytic activities of SnS₂, flower-like CdS and P25 TiO₂ in degrading MO in distilled water under visible light (420 nm) irradiation.¹⁰⁵ (Reproduced from ref. 105 with permission from Elsevier, copyright 2011.)

increased from 0.01 mol to 0.02 mol. Higher reaction temperatures promote lower viscosity and enhance the diffusion of constituent Sn⁴⁺ and S²⁻ ions. Moreover, they accelerate Ostwald ripening rates, resulting in larger particle sizes.¹⁰⁵

On the other hand, increasing the initial amount of thioacetamide from 0.01 to 0.02 mol releases more S²⁻ ions, saturating the reaction solution further. Consequently, a greater number of SnS₂ crystal nuclei form in the initial stages, reducing the available Sn⁴⁺ for subsequent crystal growth during the hydrothermal reaction and resulting in smaller particle sizes. Photocatalytic degradation of methyl orange reveals that SnS₂ synthesized at 120 °C exhibits superior performance compared with the sample prepared at 180 °C (Fig. 8(g)). The variation in particle size accounts for the observed differences, significantly influencing the available surface area and creating additional active sites to facilitate photocatalytic reactions. Furthermore, smaller particle sizes commonly lead to defect formation, serving as electron traps that prolong the lifetime of electrons and holes, thus promoting the photocatalytic degradation of methyl orange.

3.2.4 Quantum dots. Among other morphologies, quantum dots are one of the most unique morphologies regarding their physical properties. Quantum dots can be defined as particles of less than 10 nm. The uniqueness of these quantum dots is their properties, which are significantly affected by size, called the quantum size effect. Quantum dots are also recognized as having a high surface-to-volume ratio, which is comparable with larger two-dimensional nanosheets.¹⁶⁸ In the case of SnS₂, several reports have successfully synthesized SnS₂ quantum dots using various

methods.^{47,127,168–172} Zhao *et al.*¹⁶⁹ have reported the size-control of SnS₂ quantum dots using the sonication method with PEG400 as a solvent. They controlled the quantum dot size by varying the heating times after the sonication process. Fig. 9(a–d) shows the TEM images of SnS₂ quantum dots prepared by different heating times of 4, 8, 12, and 16 h, respectively. The particle size distribution of each sample plotted in the inset of each figure indicates the increase in the size with the increase of heating time from 5 to 10 nm.

Interestingly, the UV-Vis absorbance of SnS₂ quantum dots largely shifted to the larger wavelength, indicating the narrower band gap value after increasing the particle size (Fig. 9(e)). This behavior is a typical characteristic of quantum dot particles where the band gap energy is greatly affected by size. Quantum-sized particles have a higher surface area-to-volume ratio, which increases the probability of photon absorption; moreover, quantum-sized particles often exhibit higher charge carrier mobility due to their unique electronic structure and reduced dimensionality. Enhanced mobility facilitates the efficient separation and migration of photogenerated electron-hole pairs, which are essential for driving redox reactions in photocatalysis. The luminescence analysis (Fig. 9(f)) has also confirmed the emission shift because of electron excitation. The significant shift in the band gap with a slight increase in size is one of the unique features of quantum dots due to the quantum confinement effect, which is beneficial for luminescence applications. On the other hand, Wang *et al.*¹⁷³ reported the fabrication of a Si-photoanode incorporating SnS₂ quantum dots for the hydrogen evolution reaction. In Fig. 9(g), the TEM image displays the par-



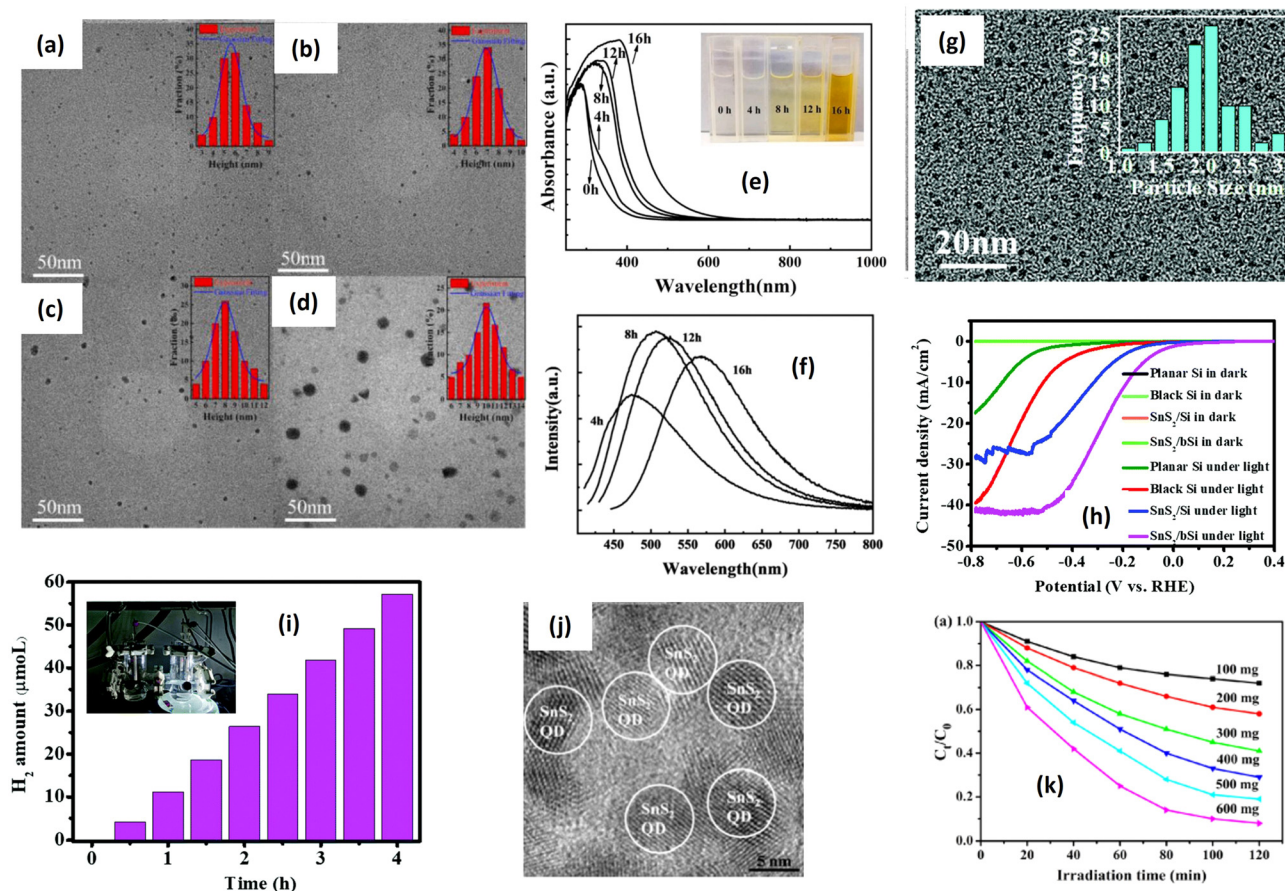


Fig. 9 TEM (inset: size distribution) images of SnS₂ QDs after heating for (a) 4 h; (b) 8 h; (c) 12 h; and (d) 16 h. (e) Absorption spectra of SnS₂ QDs with different heating times. (f) Luminescence analysis of SnS₂ QDs with different heating times.¹⁶⁹ (Reproduced from ref. 169 with permission from Elsevier, copyright 2023.) (g) TEM images of as-synthesized SnS₂ QDs.¹⁷³ (reproduced from ref. 173 with permission from the Royal Society of Chemistry, copyright 2019). (i) The calculated photocurrent density under various wavelengths; (j) HR-TEM images of SnS₂ QDs; (k) effect of the dosage of the SnS₂ QD photocatalyst on the photocatalytic reduction of 100 mg L⁻¹ Cr(vi) under visible-light irradiation ($\lambda = 420$ nm).⁴⁷ (Reproduced from ref. 47 with permission from Elsevier, copyright 2016.)

ticle distribution of SnS₂ quantum dots, with an average size of approximately 2 nm. Fig. 9(h) illustrates the polarization curves of planar Si, SnS₂/pSi, black Si, and SnS₂/bSi samples. All the Si electrodes exhibit noticeable current responses under illumination, contrasting with dark conditions. Additionally, the SnS₂/bSi photocathode demonstrates a maximum photocurrent of ~ 41 mA cm⁻² at approximately -0.51 V, attributed to the light-harvesting nanostructure of black Si and the active sites of the SnS₂ catalyst. The hydrogen evolution reaction (Fig. 9(i)) indicates a hydrogen amount of about 55 $\mu\text{mol L}^{-1}$ after irradiation. The photocatalytic performance of SnS₂ quantum dots has been also reported by Tu *et al.*⁴⁷ for Cr(vi) reduction under visible light irradiation. The quantum size SnS₂ has been successfully prepared for hydrothermal reaction with L-cystine as a particle-controlling agent. SnS₂ quantum dots with a size of 6.32 nm have been obtained, as shown in Fig. 9(j). The photocatalytic performance of SnS₂ quantum dots also shows good potential for wastewater treatment (Fig. 9(k)). Comparing various morphologies, it is anticipated that the most suitable morphology for photocatalytic

applications can be identified. Different morphologies can be achieved by adjusting several parameters during the synthesis process. As shown in Table 3, diverse morphologies, including flower-like structures, nanolayers, nanoparticles, and quantum dots, have been synthesized using different methods. Upon evaluating the apparent rate constants derived from the catalytic activities, it becomes evident that the flower-like structure exhibits the most effective photocatalytic reaction, with a degradation rate of approximately 0.0692 min⁻¹, followed by the nanoparticle morphology with the highest observed degradation rate of about 0.0594 min⁻¹. The superior photocatalytic performance of the flower-like morphology is attributed to its large specific surface area, excellent interfacial contact, and rapid electron transfer.¹⁵⁷ The enhanced photocatalytic properties of the flower-like structure can be elucidated as follows: the abundant petal-like structures provide more surface area for photochemical reactions to occur upon irradiation and interaction with chemicals, while the arrangement of flower-like structures traps light, thereby increasing electron-hole generation.¹⁷⁴ Furthermore, the open and porous nature of



Table 3 Photocatalytic activities of morphologically modified SnS₂

Morphology type	Light source	Synthesis method	Photocatalytic testing	Photocatalytic result	Ref.
SnS ₂ -flower-like	300 W xenon lamp	Hydrothermal (at 180 °C, 9 h) 80 mL 0.01 M SnCl ₄ ·5H ₂ O and L-cysteine	100 mg L ⁻¹ rhodamine B degradation	Degradation rate: 0.0251 min ⁻¹	50
SnS ₂ flower-like	300 W xenon lamp	Solvothermal in ethanol, (180 °C, 12 h) 2 mmol SnCl ₄ ·5H ₂ O and 5 mmol CH ₃ C(SNH ₂) (thioacetamide)	10 ppm rhodamine B degradation	Degradation rate: 0.0692 min ⁻¹	36
SnS ₂ flower-like	400 W metal halide lamp	Hydrothermal (200 °C, 24 h) 10 mmol SnCl ₄ ·5H ₂ O, and 3.04 g (40 mmol) thiourea	10 ppm methylene blue degradation	Degradation rate: 0.0179 min ⁻¹	176
SnS ₂ flower-like	300 W xenon lamp	Hydrothermal (180 °C, 3 days) 100.0 mg SnO ₂ hollow multi-shelled microspheres and 1.0 g thioacetamide	CO ₂ reduction	CO production rate 8.91 μmol g ⁻¹ h ⁻¹	135
SnS ₂ -nanosheets	300 W xenon lamp	Solvothermal in ethanol with acetic acid (180 °C, 12 h) 0.25 mmol SnCl ₄ ·5H ₂ O, 0.625 mmol of thioacetamide	Nitrogen fixation	NH ₄ ⁺ yield: 27.5 μmol g ⁻¹ h ⁻¹	101
SnS ₂ -nanosheets	300 W xenon lamp	Sonication in an ice-water bath for 1 h	50 mg L ⁻¹ Cr(vi) reduction	k _{Cr(vi)} = 0.01436 min ⁻¹	163
SnS ₂ -nanosheets	300 W xenon lamp 320 nm filter	Solvothermal in triethylene glycol (220 °C, 12 h). 1 mmol SnCl ₂ ·2H ₂ O, 2 mmol TAA and 0.5 g PVP	Hydrogen generation	H ₂ production rate: 1.06 mmol g ⁻¹ h ⁻¹	45
SnS ₂ -nanosheets	400 W mercury lamp	Hydrothermal with ethylenediamine (200 °C, 24 h). 3.5 g (10 mmol) of stannic(IV) chloride and 3.04 g (40 mmol) of thiourea	Hydrogen generation	H ₂ production rate: 0.566 mmol g ⁻¹ h ⁻¹	31
SnS ₂ -nanosheets	400 W mercury lamp	Hydrothermal with ethylenediamine (200 °C, 24 h). 3.5 g (10 mmol) of stannic(IV) chloride and 3.04 g (40 mmol) of thiourea	Methylene blue degradation	k _{MB} : 0.0178 min ⁻¹	31
SnS ₂ -nanoparticle	250 W xenon lamp 420 nm filter	Hydrothermal (130–170 °C). 5 mmol SnCl ₄ ·5H ₂ O, 10 mmol thioacetamide	50 mg L ⁻¹ Cr(vi) reduction	k _{Cr(vi)} : 0.0394 min ⁻¹	164
SnS ₂ -nanoparticle	250 W xenon lamp 420 nm filter	Hydrothermal (140–180 °C). 5 mmol SnCl ₄ ·5H ₂ O, 10 mmol thioacetamide	20 mg L ⁻¹ methyl orange degradation	k _{MO} : 0.0594 min ⁻¹	105
SnS ₂ -nanoparticle	500 W xenon lamp	Hydrothermal + acetic acid (130–170 °C, 12 h). 0.005 mol SnCl ₄ ·5H ₂ O and 0.01–0.02 mol thioacetamide	10 mL of log phase culture's antibacterial reaction	The bacteria killing ability of 0.11 g mL ⁻¹ NPs: 75.28 ± 2.3% for <i>S. aureus</i> (ATCC 25923), 78.95 ± 2.0% for <i>E. coli</i> (ATCC 13534), 81.22 ± 1.1% for <i>E. coli</i> (ATCC 25922)	41
SnS ₂ quantum dots	300 W Xe lamp equipped with a cut-off filter (λ > 420 nm)	Hydrothermal at (160 °C, 16 h) Sn source: SnCl ₄ ·5H ₂ O S source: L-cysteine modifier	100 mg L ⁻¹ of Cr(vi) reduction	Degradation ability: 92% for 2 h	47

flower-like structures facilitates a better mass transfer of reactants and products,¹⁷⁵ enhancing accessibility to active sites and promoting more effective contact between the photocatalyst and target pollutants in the surrounding environment.

3.3 Heterostructure formation

Heterostructure photocatalysts have garnered widespread acceptance as a highly promising technique for the development of high-performance photocatalysts. The heterostructures can be defined as the combination of photocatalysts with other materials (other photocatalysts or co-catalysts). The key advantage of heterostructure formation lies in the synergistic effects between the different materials, wherein each component mutually reinforces the other, thereby enhancing the overall photocatalytic performance.^{177,178} As illustrated in Fig. 3(b), the current literature demonstrates that approxi-

mately 75% of published works on SnS₂-based photocatalysts involve heterostructure formation. Diverse materials, including semiconductors, metals, and 2D materials such as graphene and its derivatives, have been combined with SnS₂ and exhibit a significant impact on the photocatalytic properties, such as the improvement of specific surface area, heightened light sensitivity, and enhanced reactivity. Furthermore, a crucial aspect is its ability to mitigate the recombination of electrons and holes through the formation of heterojunctions. A heterojunction is an electrical transfer between two or more materials resulting from differing electronic states and band positions. This phenomenon extends the lifetimes of electrons and holes, providing them with more time to promote redox reactions for photocatalytic reactions. To date, various types of heterojunction formation exhibit distinct characteristics and effects on photocatalytic reactions, including type I, type II, p-



n junction, Schottky junction, and Z-scheme, as illustrated in Fig. 10.¹⁷⁹

3.3.1 Schottky junction and surface plasmon resonance. A Schottky junction typically forms during the interaction between a metal and a semiconductor. Upon contact, electrons from the semiconductor tend to flow towards the metal until a Fermi equilibrium is reached. Subsequently, a barrier potential arises, hindering the further flow of electrons into the metal unless they overcome this potential barrier. The Schottky barrier effectively acts as an electron trap, preventing the electrons from flowing back into the semiconductor. This characteristic turns the metal into an electron sink for the photoinduced electrons. The Schottky barrier serves to impede the recombination of electron-hole pairs, thereby extending the lifespan of electrons for the photoreaction.¹⁸⁰

The exploration of Schottky junction processes in SnS₂ photocatalysts remains limited in the existing literature, with only a few reports delving into this phenomenon with different types of noble metal including Pt,³⁷ Au,¹⁸¹ and Ag.¹⁸² Among other noble metals, gold (Au) has emerged as one of the extensively studied supports for SnS₂ photocatalysts. Feng *et al.*¹⁸³ reported the deposition of Au on SnS₂ nanoflowers prepared *via* hydrothermal synthesis for methylene blue degradation under visible light irradiation. The transmission electron

microscopy (TEM) images in Fig. 11(a and b) reveal a 20 nm size of Au deposition on 3 μm size SnS₂ nanoflowers, indicating intimate contact between the Au and SnS₂, which enhances the charge transfer efficiency. Fig. 11(c) demonstrates the methylene blue degradation ability of Au-deposited SnS₂ with varying Au concentrations, showcasing a significant enhancement in degradation after Au deposition. Specifically, the degradation rate of Au-SnS₂ was approximately 3.91 times higher than that of SnS₂ nanoflowers alone (Table 4). The pronounced improvement in methylene blue degradation post-Au deposition can be attributed to the enhanced charge carrier efficiency facilitated by the Schottky junction formation between Au and SnS₂ (Fig. 11(d)). As the Fermi level of Au is lower than that of SnS₂, electron transfer from SnS₂ to Au occurs, effectively preventing electron-hole recombination. This allows the generation of radicals necessary for methylene blue molecule dissociation. Charge separation efficiency, typically reflected in the number of mobile electrons on the surface, was assessed through photocurrent measurement (Fig. 11(e)). Mondal *et al.*¹⁸⁴ endeavored to decompose benzylamine using Au-SnS₂ nanosheets under solar irradiation. Synthesized *via* a hydrothermal reaction followed by gold photoreduction using a 400 W xenon lamp, Au-deposited SnS₂ nanosheets achieved a conversion rate of approximately 98%

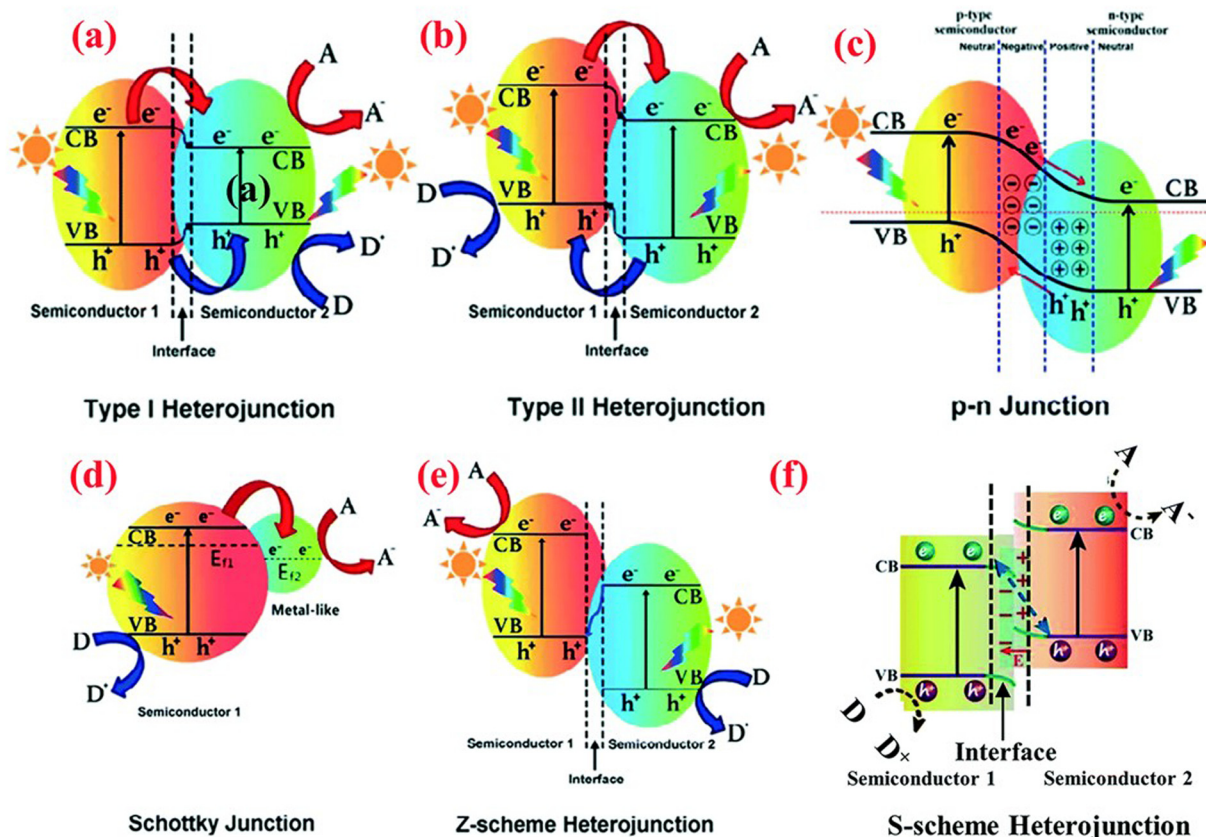


Fig. 10 Band structure of various types of heterojunction in a photocatalytic hybrid nanocomposite: (a) type I heterojunction; (b) type II heterojunction; (c) p-n junction; (d) Schottky junction; (e) Z-scheme heterojunction; and (f) S-scheme heterojunction.¹⁷⁹ (Reproduced from ref. 179 with permission from the Royal Society of Chemistry, copyright 2021.)



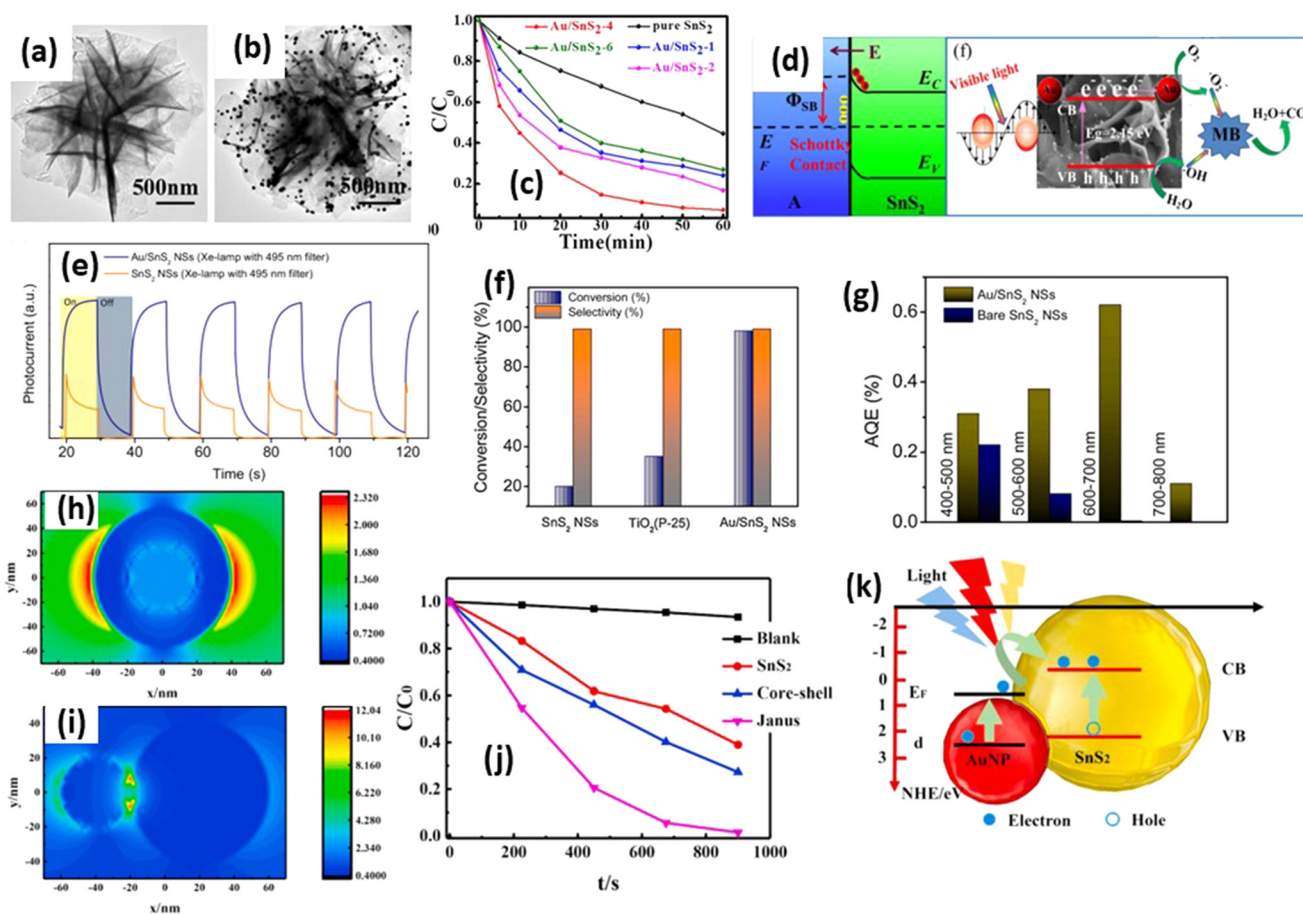


Fig. 11 TEM images of (a) SnS₂ and (b) Au/SnS₂. (c) Methylene blue degradation by Au/SnS₂. (d) Illustration of the band alignment formed at the interface between the SnS₂ and Au NPs.¹⁸³ (Reproduced from ref. 183 with permission from Elsevier, copyright 2019.) (e) Transient photocurrent recorded on bare SnS₂ and Au/SnS₂ NSs using a 400 W Xe lamp (with an attached 495 nm cut-off filter). (f) Comparison plots of benzylamine oxidation efficiency with different catalysts under natural sunlight and ambient conditions for 2 h in open air. (g) Plot of AQE at different wavelength ranges under LED illumination in air at room temperature for bare SnS₂ and Au/SnS₂.¹⁸⁴ (Reproduced from ref. 184 with permission from Elsevier, copyright 2021.) (h) The near-field distribution of the core-shell structure and the (i) Janus structure with Au nanoparticles, respectively. (j) Degradation curves of methyl orange under visible light. (k) The schematic diagram of the LSPR-enhanced visible light catalysis of SnS₂ and the heterostructures with Au nanoparticles.¹⁸⁶ (Reproduced from ref. 186 with permission from Elsevier, copyright 2021.)

under solar light (Fig. 11(f)), surpassing SnS₂ and standard P-25 conversion rates of about 20% and 35%, respectively (Table 4).

To investigate the photocatalytic enhancement in Au-SnS₂ nanosheets, the apparent quantum efficiency (AQE) was examined at different wavelengths (400–500 nm, 500–600 nm, 600–700 nm, 700–800 nm) (Fig. 11(g)). Compared with SnS₂ nanosheets, Au-SnS₂ exhibited a 1.2 to 3 times higher AQE at 400–500 nm and 500–600 nm wavelengths, attributed to improved electron-hole separation mediated by Au nanoparticles (Fig. 11(g)). Notably, at 600–700 nm, where SnS₂ excitation is minimal, Au-deposited SnS₂ exhibited a strong AQE, indicating an efficient ‘hot-electron’ injection mechanism from Au nanoparticles to SnS₂ nanosheets.¹⁸⁵ Furthermore, the formation of Au-SnS₂ is influenced by Au positioning relative to the SnS₂ photocatalyst. Fu *et al.*¹⁸⁶ investigated different composite structures of Au-SnS₂ for methyl orange degra-

ation under visible light. Two structures were examined: core-shell and Janus structures. Near-field analysis (Fig. 11(h and i)) revealed a low light enhancement effect in the core-shell structures, leading to diminished photocatalytic ability. Conversely, Janus structures exhibited a significant near-field enhancement, resulting in superior methyl orange degradation (Fig. 11(j)). Among the images, the Janus structures exhibit the highest reaction rate ($4.0 \times 10^{-3} \text{ s}^{-1}$), followed by the core-shell structures ($1.2 \times 10^{-3} \text{ s}^{-1}$). A significant increase in photocatalytic activity, approximately four times greater than that of sole SnS₂ nanoparticles ($1.0 \times 10^{-3} \text{ s}^{-1}$), is observed in the Janus structures. This enhancement is attributed to the improved extinction effect. Photoluminescence lifetime analysis revealed longer electron lifetimes in nanoparticle structures (1.74 ns) than Janus structures (1.53 ns) compared with core-shell structures (1.16 ns), indicating an enhanced photocatalytic efficiency. Fig. 11(k) illustrates the proposed



Table 4 Photocatalytic performance of heterojunction formation of SnS₂

Materials	Heterojunction type	Light source	Photocatalytic testing	Photocatalytic performance	Ref.
Au-SnS ₂	Schottky junction	330 W xenon lamp	100 mg L ⁻¹ rhodamine B degradation	k_{RhB} : 0.0514 min ⁻¹ . 3.91 times that of SnS ₂	183
Au-SnS ₂	Schottky junction	400 W xenon lamp	Benzylamine oxidation	Conversion: 98%, 4.9 times that of SnS ₂	184
Au-SnS ₂	Schottky junction	250 W xenon lamp	50 mg L ⁻¹ Cr(vi) reduction	$k_{\text{Cr(vi)}}$: 0.0234 min ⁻¹ . 2.07 times that of SnS ₂	73
BiFeO ₃ /SnS ₂	p-n junction	250 W mercury lamp	15 mg L ⁻¹ methylene blue degradation	k_{MB} : 0.0428 min ⁻¹ . 7.50 times that of SnS ₂	188
BiOI/SnS ₂	p-n junction	UV-vis (385–740 nm)	10 mg L ⁻¹ rhodamine B degradation	k_{RhB} 100% within 30 min 3 times that of SnS ₂	187
TiO ₂ /SnS ₂	Type I heterojunction	Xe lamp, 100 mW cm ⁻²	Hydrogen evolution reaction (HER)	H ₂ yield: 652.4 μmol g ⁻¹ h ⁻¹ 7.51 times that of SnS ₂	190
ZnIn ₂ S ₄ /SnS ₂	Type II heterojunction	300 W xenon lamp	50 mg L ⁻¹ Cr(vi) reduction	$k_{\text{Cr(vi)}}$: 0.01273 min ⁻¹ 2.2 times that of SnS ₂	196
ZnWO ₄ /SnS ₂	Type II heterojunction	Solar hotspot 5 kW h m ⁻² day ⁻¹	50 mg L ⁻¹ Cr(vi) reduction	$k_{\text{Cr(vi)}}$ = 0.061 min ⁻¹ 3.58 times that of SnS ₂	104
ZnWO ₄ /SnS ₂	Type II heterojunction	Solar hotspot 5 kW h m ⁻² day ⁻¹	20 mg L ⁻¹ tetracycline	k_{TC} = degradation rate: 0.023 min ⁻¹ 3.28 times that of SnS ₂	104
In ₂ S ₃ /SnS ₂	Type II heterojunction	300 W xenon lamp	10 mg L ⁻¹ rhodamine B degradation	k_{RhB} : 0.11716 min ⁻¹ . 15.3 times that of SnS ₂	192
Ag ₃ PO ₄ /SnS ₂	Z-scheme	A 500 W xenon-arc lamp	10 mg L ⁻¹ methylene blue	k_{MB} = 0.063 min ⁻¹ 7.08 times that of SnS ₂	201
Bi ₂ WO ₄ /SnS ₂	Z-scheme	Sunlight	20 mg L ⁻¹ tetracycline degradation	k_{TC} = 0.027 min ⁻¹ 27 times that of SnS ₂	134
BiOBr/SnS ₂	Z-scheme	400 W xenon lamp	10 mg L ⁻¹ rhodamine B degradation	k_{RhB} = 0.1203 min ⁻¹ 75 times that of SnS ₂	197
SnO ₂ /SnS ₂	S-scheme	500 W xenon lamp	10 mg L ⁻¹ methyl orange degradation	k_{MO} : 0.01182 min ⁻¹ . 2.19 times that of SnS ₂	199
ZnIn ₂ S ₄ /SnS ₂	S-scheme	300 W xenon lamp	H ₂ production	H ₂ yield: 1.113 mmol g ⁻¹ h ⁻¹ 16.14 times that of ZnIn ₂ S ₄	200
CdS/SnS ₂	S-scheme	300 W xenon lamp	H ₂ production	H ₂ yield: 5.18 mmol g ⁻¹ h ⁻¹ 5.95 times that of CdS	198

mechanism of surface resonance analysis of Janus Au-SnS₂; due to the near-field enhancement, electron excitation becomes more pronounced in SnS₂ nanoparticles, significantly enhancing photocatalysis. Besides the enhanced optical fields, the localized surface plasmon resonance (LSPR) induces electron injection from AuNPs to the conduction band edge of SnS₂, further contributing to the photocatalytic process. Among these mechanisms, the excited electrons in the conduction band edge of the SnS₂ nanoparticles facilitate the generation of reactive oxygen radicals, including superoxide, hydroxyl, and singlet oxygen, through a series of oxidation-reduction processes in the solution. These oxygen radicals exhibit high activity and play a crucial role in degrading methyl orange molecules.¹⁸⁶

3.3.2 p-n junction. The p-n junction arises from electronic interactions driven by differences in the majority carriers of semiconductors. In n-type semiconductors, electrons predominate as majority carriers, whereas in p-type semiconductors, holes are prevalent. Electrons flow from n-type to p-type, while holes migrate oppositely, impeding electron-hole recombination. SnS₂, an n-type semiconductor, is frequently combined with various p-type semiconductors like as BiOI,¹⁸⁷ BiFeO₃,¹⁸⁸ NiS,¹⁸⁹ and SnS nanoparticles¹⁶ to create p-n junction photocatalysts.

Wang *et al.*¹⁸⁷ confirmed the p-n junction formation between BiOI and SnS₂ for rhodamine B degradation under

visible light irradiation, as illustrated in Fig. 12(a). Notably, SnS₂ has a layered morphology, while BiOI comprises the nanoparticles that disperse on the top of the SnS₂ layer. The conduction band of SnS₂ is situated at -0.96 V, with a valence band at 1.5 V. Conversely, BiOI exhibits a valence band at +0.53 V and a conduction band at -1.29 V. Upon electrical contact, band alignment ensues, facilitating electron and hole transfer between the BiOI and SnS₂ until equilibrium is attained, with the Fermi level equalized. Upon irradiation, electron-hole pairs form in the conduction and valence bands of SnS₂ and BiOI. Due to the p-n junction, electrons readily transfer from BiOI to SnS₂, while holes undergo the reverse process. This reaction process enhances electron mobility and correlates directly with photocatalytic efficiency. Photocurrent analysis (Fig. 12(b)) demonstrates increased electron densities of BiOI/SnS₂ heterostructures compared with BiOI and SnS₂. The active electrons in the conduction band react with O₂ to form O₂^{•-}, while the holes react with OH⁻ to produce OH[•], facilitating rhodamine B degradation. Rhodamine B degradation analysis (Fig. 12(c)) indicates a significant improvement after p-n junction formation, three times higher than that observed with SnS₂ alone (Table 4), a finding consistent with the photocurrent analysis.

3.3.3 Type I and II heterojunctions. Fig. 13(a) illustrates a type I heterojunction within a photocatalytic reaction. Type I heterojunctions manifest when a semiconductor with a rela-



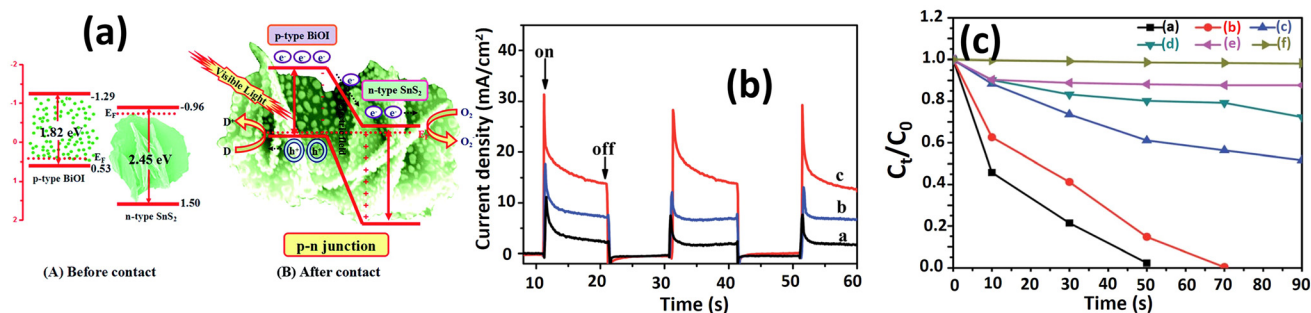


Fig. 12 (a) Schematic diagrams illustrating the energy bands of p-BiOI and n-SnS₂ before contact, the formation of a p-n junction, and the equilibrium energy band diagram. Additionally, the transfer of photoinduced electrons from p-BiOI to n-SnS₂ under visible-light irradiation is depicted; (b) comparison of the transient photocurrent response of SnS₂ (black), BiOI (blue), and BiOI/SnS₂ (red) in 0.1 M Na₂SO₄ aqueous solution under visible-light irradiation; (c) visible-light photocatalytic activities: (black) BiOI/SnS₂, (red) BiOI, (blue) SnS₂, (green) P25, (violet) adsorption in the dark, and (yellow) degradation of 10 mg rhodamine B without photocatalyst under visible light ($\lambda > 420$ nm) irradiation.¹⁸⁷ (Reproduced from ref. 187 with permission from the Royal Society of Chemistry, copyright 2015.)

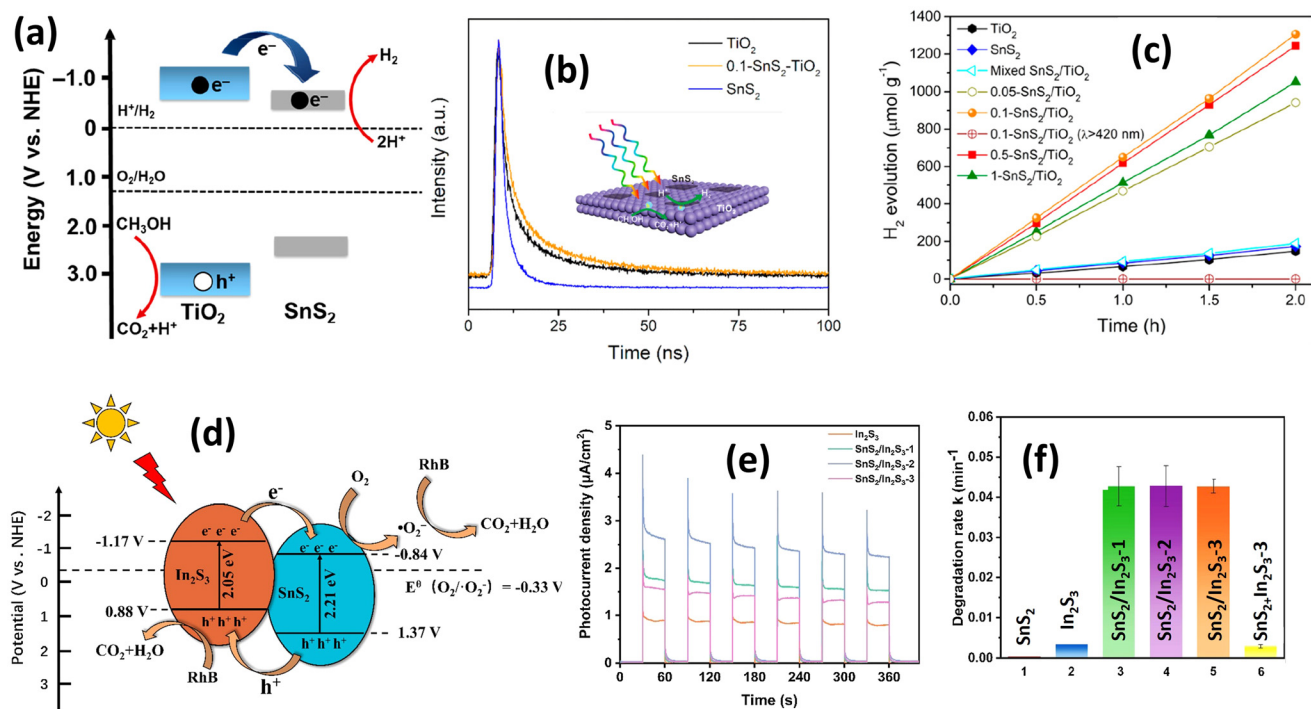


Fig. 13 (a) Corresponding band energy diagrams and charge transfer path; (b) time-resolved photoluminescence decay spectra of TiO₂, SnS₂, and 0.1-SnS₂/TiO₂; (c) photocatalytic H₂ evolution curves with different catalysts.¹⁹⁰ (Reproduced from ref. 190 with permission from Elsevier, copyright 2019.) (d) Schematic diagram of the possible photocatalytic degradation mechanism of RhB by the SnS₂/In₂S₃ heterostructure under visible light irradiation; (e) transient photocurrent responses of In₂S₃ and SnS₂/In₂S₃ heterostructures; (f) the apparent rate constants of SnS₂/In₂S₃.¹⁹² (Reproduced from ref. 192 with permission from Elsevier, copyright 2022.)

tively large band gap energy, such as TiO₂, interfaces with another semiconductor possessing a smaller band gap energy. The valence band and conduction band positions are intermediary between those of materials with larger band gaps. In this context, following photoexcitation, electrons and holes migrate from the large band gap photocatalyst to the smaller band gap photocatalyst, thereby facilitating reduction-oxidation reactions crucial for photocatalytic processes. Sun *et al.*¹⁹⁰ provide an exemplar of a type I heterojunction with a

SnS₂ photocatalyst, detailing the formation of SnS₂/TiO₂ heterostructures. Fig. 13(a) elucidates the catalytic reaction process of type I heterojunction formation between SnS₂ and TiO₂. Under these reaction conditions, they propose that photoelectrons from the conduction band of TiO₂ migrate into the conduction band, reducing H⁺ ions into hydrogen gas. Meanwhile, holes react with methanol to yield hydrogen gas. This heterojunction reaction prolongs the decay time of electrons, estimated at around 4.4 ns compared with TiO₂ (2.8 ns)



and SnS₂ (1.8 ns) as shown in Fig. 13(b). The hydrogen evolution reaction, as depicted in Fig. 13(c), showcased a remarkable enhancement facilitated by the 0.1-SnS₂/TiO₂ compared with SnS₂ and TiO₂ individually. The 0.1-SnS₂/TiO₂ nanohybrid exhibited an impressive hydrogen production rate, reaching 652.4 μmol h⁻¹ g⁻¹, surpassing the rates observed with pure TiO₂ and SnS₂ by approximately 8 times. However, a further increase in the amount of SnS₂ in the hybrid resulted in a slight decline in H₂ production activity, with values of 621.5 and 526.7 μmol h⁻¹ g⁻¹ for 0.5-SnS₂/TiO₂ and 1-SnS₂/TiO₂, respectively. This phenomenon can be attributed to the excess SnS₂ hindering the absorption of incident light and the generation of excited carriers by TiO₂. Additionally, an abundance of SnS₂ may act as charge recombination centers, leading to a reduced photocatalytic performance. Another type I heterojunction formation has been reported by Liu *et al.*¹⁹¹ investigating the formation of SnS/SnS₂ for the electrochemical water splitting process. They found that the energy gap of SnS₂ is 2.15 eV, while SnS has an energy gap of 1.1 eV. The conduction band (CB) and valence band (VB) positions of SnS₂ are -0.08 eV and 2.08 eV, respectively, whereas those of SnS are 0.12 eV and 1.22 eV, respectively. Given that the VB of SnS₂ is more positive while the CB is more negative, electrons and holes tend to migrate to SnS after electron excitation, effectively retarding electron and hole recombination. The photocurrent analysis revealed that the current density under irradiation for the SnS₂/SnS heterojunction is five times higher than that of SnS and ten times higher than that of SnS₂, confirming the charge carrier efficiency.

Differing from type I heterojunctions, type II heterojunctions involve semiconductor photocatalysts where two types of semiconductor exhibit distinct band structure levels. One semiconductor possesses a more positive valence band, while the other semiconductor has a more negative conduction band. In this arrangement, electrons flow from the semiconductor with the more negative conduction band position, while holes flow from the semiconductor with the more positive valence band. This electrical transfer scheme helps prevent electron and hole recombination. An example of a type II heterojunction formation of SnS₂ was reported by Zhang *et al.*¹⁹² investigating the formation of SnS₂ with In₂S₃. Fig. 13(d) illustrates the catalytic reaction mechanism of the SnS₂/In₂S₃ heterostructure for rhodamine B degradation under visible light irradiation. Upon irradiation, electron transfer occurs from the conduction band of In₂S₃ to the conduction band of SnS₂ due to the more negative level of the conduction band position of In₂S₃, while hole transfer operates oppositely due to the greater positivity of the valence band level of SnS₂. Similar to a heterojunction type I, this charge transfer process between In₂S₃ and SnS₂ significantly enhances charge mobility efficiency, as demonstrated in Fig. 13(e); the photocurrent of the In₂S₃/SnS₂ heterostructure is 3–4 times higher than In₂S₃ only, which shows a marked improvement in current density after heterojunction formation. RhB degradation testing indicates a significant enhancement in RhB degradation by conducting In₂S₃/SnS₂ heterojunction formation (Fig. 13(f)),

where the degradation rate was 99 times and 15 times higher than with SnS₂ and In₂S₃ alone, respectively. The theoretical explanation of the In₂S₃/SnS₂ heterojunction as type II heterojunction formation was confirmed by Wang *et al.*¹⁹³ through EPR analysis for detecting superoxide radicals. The lower intensity of the EPR signal of In₂S₃/SnS₂ compared with In₂S₃ indicates electron transfer from In₂S₃ to SnS₂, confirming type II heterojunction formation. Other materials have also formed type II heterojunctions combined with SnS₂, such as g-C₃N₄,¹⁹⁴ ZnWO₄,¹⁰⁴ SnO₂,¹⁹⁵ and ZnIn₂S₄.¹⁹⁶

3.3.4 Z-scheme and S-scheme photocatalyst. However, based on above existing heterostructure systems there is still an issue regarding the reducing of the oxidation and reduction ability of the photocatalyst. To address this issue, scientists have pioneered the development of Z-scheme and S-scheme photocatalysts. These novel configurations offer superior reduction and oxidation capabilities compared with conventional type I and type II heterojunction formations. To date, Z-scheme and S-scheme heterojunctions have emerged as the most investigated heterostructures for SnS₂-based photocatalysts. Notably, the direct Z-scheme photocatalyst for SnS₂ holds considerable promise for future advancements, as it obviates the need for a metal mediator to facilitate Z-scheme reactions.

Several studies have confirmed the formation of direct Z-scheme photocatalysts. For instance, Luo *et al.*¹⁵ initiated Z-scheme formation by integrating SnS₂ with Ag₃PO₄ for methyl orange degradation under visible light irradiation. The schematic representation in Fig. 14(a) illustrates the Z-scheme formation of the SnS₂/Ag₃PO₄ heterostructure. The realization of a direct Z-scheme photocatalyst depends on the band positions of each constituent material. In the case of the SnS₂/Ag₃O₄ heterostructure, the Z-scheme configuration is inferred from the disparate band positions of the SnS₂ and Ag₃PO₄.

The formation of a direct Z-scheme photocatalyst relies on the band positions of each material. In the case of the SnS₂/Ag₃PO₄ heterostructure, the existence of the Z-scheme can be inferred from the disparate band positions of each material. The valence band (VB) maximum and conduction band (CB) minimum of SnS₂ are about 1.21 eV and -0.89 eV, respectively, while the VB and CB potentials of Ag₃PO₄ are about 2.62 eV and 0.30 eV, respectively. The CB position of Ag₃PO₄ is more positive than the standard reduction potential of O₂/O₂^{•-} (0.13 eV). The electron from Ag₃PO₄ is unable to reduce O₂ into O₂^{•-} and instead reacts with the hole of the SnS₂. Consequently, recombination is prevented, leading to a more efficient charge separation process, as demonstrated by photocurrent analysis. This correlation directly corresponds to the observed improvements in photocatalytic activity (see Fig. 14(b and c)). As mentioned earlier, the Z-scheme can be generated if one of the conduction bands (CBs) of the sample position is more positive than the reduction potential of O₂/O₂^{•-} (0.13 eV), thereby preventing the reduction reaction. Other reports also confirm the same understanding about the Z-scheme formation, such as Qiu *et al.*,¹⁹⁷ Li *et al.*,⁴⁹ and Kumar *et al.*¹³⁴

Slightly different from the Z-scheme, the S-scheme can be generated through the inner electric field phenomenon. This



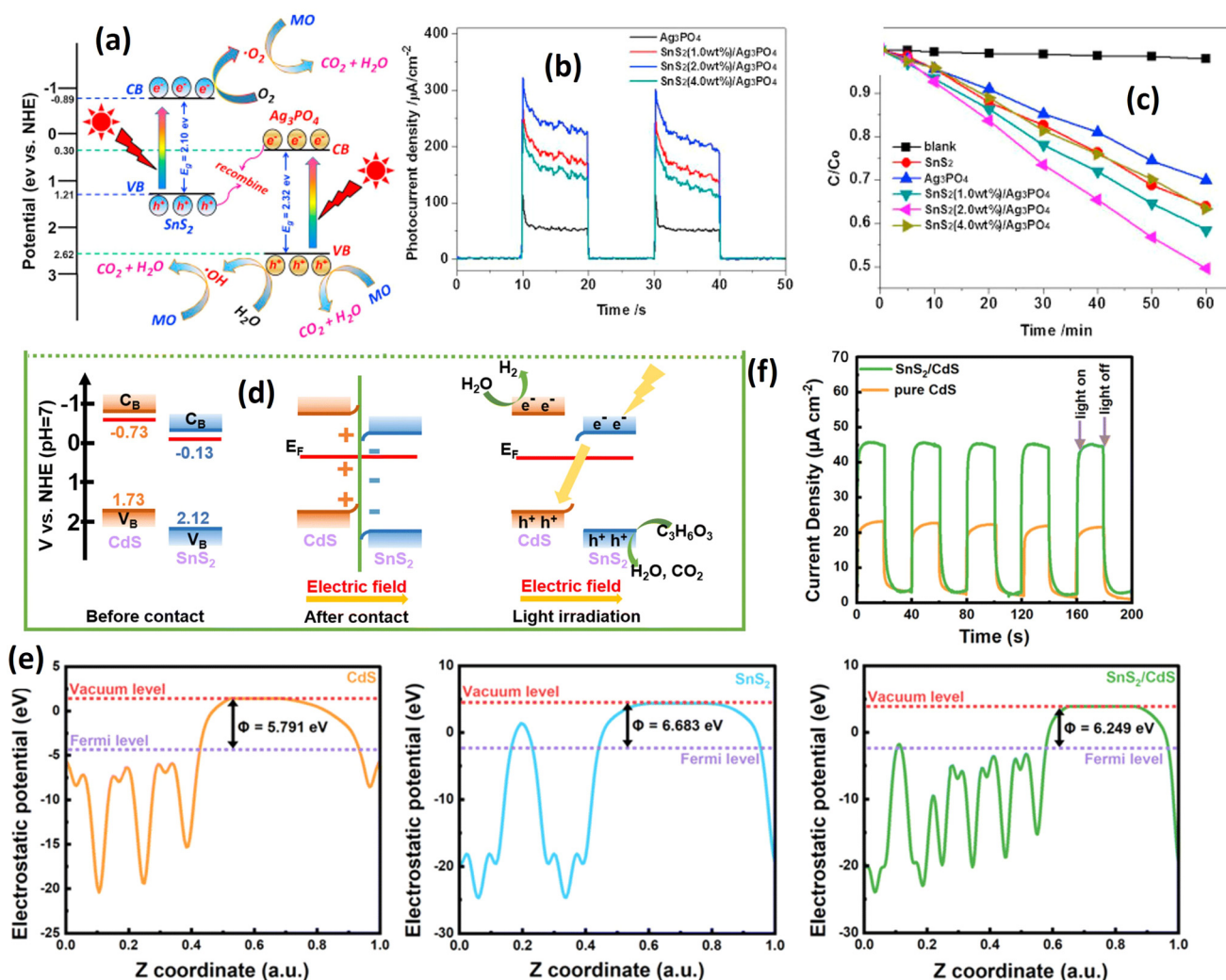


Fig. 14 (a) Schematic diagram of the possible photoinduced electron–hole pair separation process and direct Z-scheme photodegradation mechanism of methyl orange over the $\text{SnS}_2/\text{Ag}_3\text{PO}_4$ heterojunction photocatalyst under visible light irradiation. (b) Photocurrent response over the as-prepared samples under visible light irradiation. (c) Photocatalytic degradation curves for the photodegradation of MO.¹⁵ (Reproduced from ref. 15 with permission from Elsevier Ltd.) (d) S-scheme illustration of the possible charge transfer processes in the SnS_2/CdS heterostructures; (e) work function analysis of CdS, SnS_2 and SnS_2/CdS . (f) Photocurrent analysis of CdS and SnS_2/CdS .¹⁹⁸ (Reproduced from ref. 198 with permission from the Royal Society of Chemistry.)

electric field emerges when a charge moves between semiconductors, creating a field within each semiconductor due to differing Fermi energies. Fig. 14(d) illustrates an instance of S-scheme formation between SnS_2 and CdS nanoparticles.¹⁹⁸ Before contact, SnS_2 and CdS exhibit different Fermi energy levels; CdS registers at 5.791 eV (Fig. 14(e)), while SnS_2 is at 6.683 eV. Upon interaction, Fermi energy alignment occurs until an equilibrium is established. In this state, the electron density at the SnS_2 interface surpasses that of CdS, generating an inner electric field. Consequently, the valence band of CdS bends, and the conduction band of SnS_2 shows observable curvature. Upon light irradiation, both SnS_2 and CdS generate electron–hole pairs, and the inner electric field facilitates electron transfer from SnS_2 to CdS, enhancing the charge separation efficiency. This improvement is supported by photo-

current analysis (Fig. 14(f)). The photocatalytic activity of the S-scheme SnS_2/CdS photocatalyst for H_2 production, summarized in Table 4, yields $5.18 \text{ mmol g}^{-1} \text{ h}^{-1}$ of hydrogen, 5.95 times higher than that of CdS alone. Similar S-scheme SnS_2 -based photocatalysts have been reported for other materials, including BiOBr ,¹⁹⁷ SnO_2 ,¹⁹⁹ ZnIn_2S_4 .²⁰⁰

Table 4 presents various reports on the photocatalytic performance of SnS_2 heterojunction formations. Different strategies yield varying degrees of improvement compared with SnS_2 alone. According to current data, Au/SnS_2 demonstrates a 3.91-fold enhancement over SnS_2 . Among the p–n junction formations, $\text{BiFeO}_3/\text{SnS}_2$ exhibits the most significant improvement, with a degradation rate approximately 7.50 times higher than SnS_2 . Similarly, heterojunctions of type I and type II show comparable enhancement, boasting 7.51 times improvements



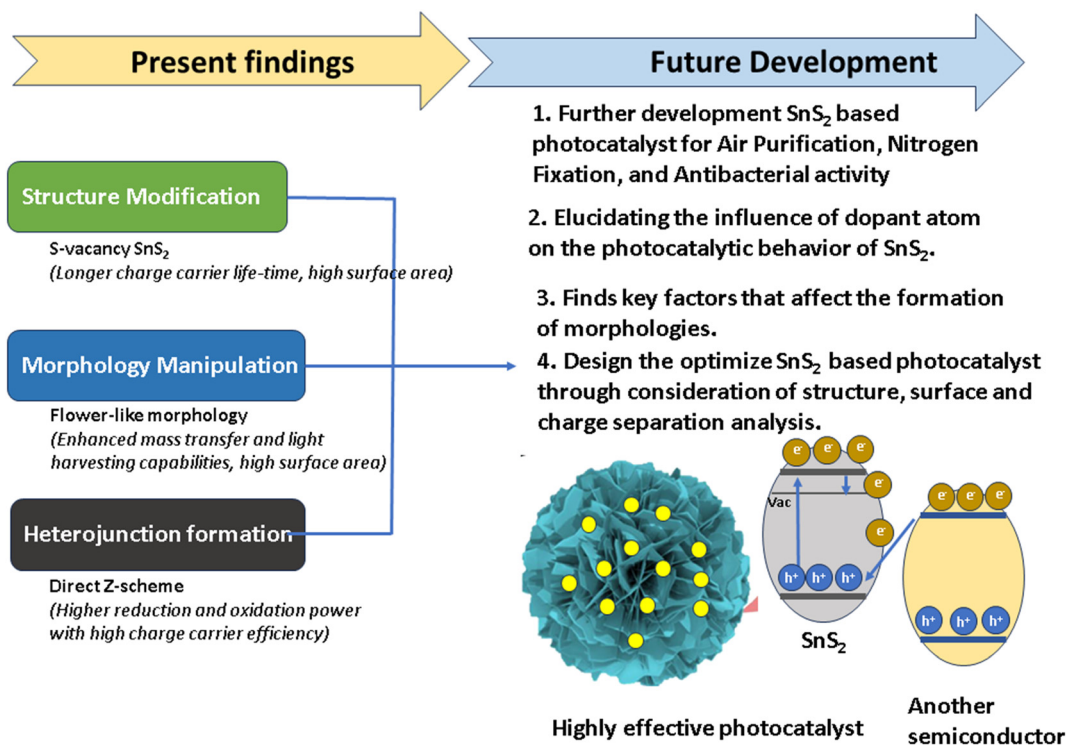


Fig. 15 Current findings and future research direction for SnS₂ photocatalytic research.

over SnS₂. Notably, Z-scheme configurations outperform other heterojunction types, with BiOBr/SnS₂ achieving the highest improvement, nearly 75 times higher than SnS₂ alone. This superior enhancement in Z-scheme heterojunctions is attributed to enhanced charge separation efficiency and stronger reduction and oxidation capabilities, resulting in accelerated degradation rates, and increased photochemical reaction yields.

4 Future perspectives

The current strategies for modifying SnS₂ have been thoroughly examined, encompassing structural modification, morphological manipulation, and heterojunction formation. Despite the comprehensive elucidation of each photocatalytic advancement mechanism and synthesis technique in this review, challenges persist, hindering the further advancement of SnS₂ photocatalysts.

In terms of structural modification, it is established that S-vacancies play a pivotal role in enhancing the photocatalytic performance, primarily by improving electron-hole separation efficiency. While the authors attribute this enhancement to electron traps between the valence and conduction bands, empirical evidence remains elusive. Future research endeavors should focus on elucidating the mechanism of the S-vacancy effect on mid-gap formation to substantiate this claim. Additionally, the role of dopants in photocatalytic reactions remains unclear due to their concurrent generation with

S-vacancy formation during doping processes. To advance research in this area, techniques minimizing the impact of S-vacancies in doping must be developed to gain a comprehensive understanding of doping phenomena for photocatalytic enhancement.

In morphology engineering, the flower-like morphology has emerged as the most suitable structure for various photocatalytic applications, including wastewater treatment and hydrogen production. Enhanced mass transfer and light harvesting capabilities are key attributes contributing to the photocatalytic performance of flower-like structures. However, challenges in material preparation persist, particularly in hydrothermal synthesis methods, which remain dominant in SnS₂ flower-like morphology production. Although hydrothermal synthesis is known for producing flower-like morphologies, it is important to note that other morphologies can also be generated using the same method and precursor. This variability may hinder further development, as precise control over flower-like morphology parameters such as size, petal-like dimensions, and facet arrangements are essential for further research development of the morphological effect on the photocatalytic performance of SnS₂. Therefore, for future development, it is necessary to find the key factors which control the growth of morphology through hydrothermal reactions encompassing the precursor, solvent, temperature control, and other factors.

While the proliferation of heterojunction formation is widespread, it does not guarantee a comprehensive understanding of photocatalytic mechanisms. Heterostructure formation, con-



sidered one of the most extensively studied systems for SnS₂ photocatalysts, relies on the synergistic effects of its constituent materials. Various types of heterostructure have been reported, each with unique characteristics. Among these, the Z-scheme exhibits significant promise for enhancing photocatalytic performance due to its capacity to augment oxidation and reduction potentials for photocatalytic reactions. However, several considerations must be addressed for the optimal utilization of Z-scheme heterojunction photocatalysts. Beyond the reliance on electronic properties for photocatalytic enhancements, factors such as particle interaction and surface optimization are also pivotal. Particle interaction hinges on the specific facets of each particle, as different facets exhibit distinct electronic and adsorption behaviors. Theoretical investigations, particularly those utilizing density functional theory (DFT) calculations, are essential to fully comprehend these interactions. Additionally, optimizing the surface properties of Z-scheme particles for specific applications, including increasing the surface area and modifying the charge, is imperative for optimizing the adsorption process and enhancing active sites for photocatalytic reactions. To address these issues, for future research directions, it is interesting to employ the comprehensive investigation of photocatalyst design while considering the structural design alongside surface optimization and Z-scheme heterojunction formation.

We have thoroughly discussed the current results and outlined future perspectives for three distinct strategies. Building on this understanding, future research endeavors will focus on integrating these strategies into a unified photocatalyst system. The S-vacancy SnS₂, characterized by a prolonged charge carrier lifetime, can be elaborately engineered into a flower-like morphology with enhanced mass transfer, light harvesting capabilities, and a large surface area. Concurrently, nanoparticles with suitable band gaps, such as CdS and Ag₃VO₄, can be selectively deposited atop the flower-like SnS₂ system for generating a Z-scheme photocatalyst, as illustrated in Fig. 14. The successful amalgamation of these components is anticipated to yield a photocatalyst system of exceptional performance, aligning with industrial standards.

Furthermore, it is imperative to address a critical aspect pertaining to the chemical stability of SnS₂ itself, given its susceptibility to sulfur atom replacement by oxygen upon exposure to air. Factors influencing stability, including the synthesis process and the photocatalyst environment, necessitate meticulous examination. Such analysis assumes paramount importance for practical applications, wherein long-term stability is imperative for industrial viability.

5 Summary

For summarizing this review, Fig. 15 was used for understanding the current findings and future directions. In this review, we have summarized recent strategies for improving the photocatalytic performance of SnS₂, focusing on three fundamental approaches: crystal structure engineering, morphology engineering,

and heterojunction formation. SnS₂-based photocatalysts have been utilized in various applications, including CO₂ reduction, H₂ production, nitrogen fixation, antibacterial activity, air purification, and wastewater treatment technologies.

At the current stage, the structural engineering of SnS₂ has been carried out through two methods: vacancy generation and doping processes. These approaches have successfully increased the photocatalytic performance of SnS₂. The fundamental phenomenon observed in S-vacancy generation and doping processes is the ability to decrease the recombination of electrons and holes, thereby facilitating redox reactions. Additionally, altering crystal structures results in a decrease in the crystallite size of SnS₂, which is beneficial for increasing the surface area and the number of active sites for chemical reactions.

Various morphologies of SnS₂, such as flower-like structures, nanoparticles, nanosheets, and quantum dots, exhibit distinct effects on the photocatalytic performance. Flower-like structures, with their high specific surface area and unique morphology, demonstrate superior light harvesting ability and increased active sites compared with nanosheet and nanoparticle morphologies. Although nanosheets have a smaller surface area than flower-like structures, they facilitate the efficient migration of photogenerated electrons and holes to reaction sites, thereby enhancing the photocatalytic efficiency. Conversely, nanoparticle morphology is influenced by the size of the nanoparticles; smaller nanoparticles also contribute to an improved specific surface area, thereby enhancing photocatalytic activity.

Heterojunction formation is by far the most effective strategy for improving photocatalytic reactions, as it enables the generation of synergistic effects between SnS₂ and other materials, significantly enhancing the photocatalytic efficiency. Among various types of heterojunction, Z-scheme photocatalytic reactions are the most promising strategy for improving the photocatalytic performance of SnS₂. This is due to their stronger oxidation and reduction capabilities, coupled with highly efficient charge carrier separation efficiency.

For further development, several aspects are still limited in current literature. Firstly, there is a lack of research on the application of SnS₂ for antibacterial activity, N₂ fixation, and air purification. Therefore, future research should focus more on investigating these types of photocatalytic application. Secondly, there is uncertainty regarding the effect of dopants on vacancy generation. Since the doping process directly leads to vacancy formation, the individual impact of each dopant is not well understood. To address this issue, both synthesis processes and theoretical approaches should be conducted simultaneously to comprehensively understand the doping effect on improving charge carrier separation in SnS₂. Moreover, the growth of morphology is not well understood. Therefore, for future development, it is necessary to find the key factor which controls the growth of morphology through hydrothermal reactions encompassing precursor, solvent, temperature control, and other factors.



However, one of the most crucial areas of research is the design of the ideal photocatalyst based on the aforementioned findings. Developing a Z-scheme SnS₂-based photocatalyst with a high specific surface area and efficient charge carrier separation is essential. One promising technique is to combine SnS₂ flower-like structures with very small nanoparticles, allowing the particles to attach to the flower-like structures, thereby improving the light harvesting process and charge carrier separation efficiency. Lastly, the stability of SnS₂ is a critical consideration. Since SnS₂ is prone to oxidation, stability studies must be conducted to meet industrial standards.

Conflicts of interest

There are no conflicts to declare.

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References

- R. Yang, Y. Fan, Y. Zhang, L. Mei, R. Zhu, J. Qin, J. Hu, Z. Chen, Y. Hau Ng, D. Voiry, S. Li, Q. Lu, Q. Wang, J. C. Yu and Z. Zeng, *Angew. Chem., Int. Ed.*, 2023, **62**, 1–29.
- W. Peng, Y. Li, F. Zhang, G. Zhang and X. Fan, *Ind. Eng. Chem. Res.*, 2017, **56**, 4611–4626.
- V. Dhiman and N. Kondal, *Phys. B*, 2022, **628**, 413569.
- W. Yu, X. Chen, W. Mei, C. Chen and Y. Tsang, *Appl. Surf. Sci.*, 2017, **400**, 129–138.
- R. A. Geioushy, I. M. Hegazy, S. M. El-Sheikh and O. A. Fouad, *J. Environ. Chem. Eng.*, 2022, **10**, 107337.
- Y. J. Huang, L. M. Lyu, C. Y. Lin, G. C. Lee, K. Y. Hsiao and M. Y. Lu, *ACS Omega*, 2022, **7**, 2217–2223.
- B. Chen, Y. Meng, J. Sha, C. Zhong, W. Hu and N. Zhao, *Nanoscale*, 2018, **10**, 34–68.
- Y. Lin, P. Ren and C. Wei, *CrystEngComm*, 2019, **21**, 3439–3450.
- E. C. Cho, C. W. Chang-Jian, J. H. Zheng, J. H. Huang, K. C. Lee, B. C. Ho and Y. S. Hsiao, *J. Taiwan Inst. Chem. Eng.*, 2018, **91**, 489–498.
- W. Ho, J. C. Yu, J. Lin, J. Yu and P. Li, *Langmuir*, 2004, **20**, 5865–5869.
- X. Zhang, K. Zhu, C. Xie and P. Yang, *Carbon*, 2024, **220**, 118884.
- Y. Chen, H. Sun and W. Peng, *Nanomaterials*, 2017, **7**, 1–11.
- R. Tong, K. W. Ng, X. Wang, S. Wang, X. Wang and H. Pan, *J. Mater. Chem. A*, 2020, **8**, 23202–23230.
- L. Han, Y. L. Zhong, K. Lei, D. Mao, Y. Z. Dong, G. Hong, Y. T. Zhou and D. Fang, *J. Phys. Chem. C*, 2019, **123**, 2398–2409.
- J. Luo, X. Zhou, L. Ma, L. Xu, X. Xu, Z. Du and J. Zhang, *Mater. Res. Bull.*, 2016, **81**, 16–26.
- X. Hu, G. Song, W. Li, Y. Peng, L. Jiang, Y. Xue, Q. Liu, Z. Chen and J. Hu, *Mater. Res. Bull.*, 2013, **48**, 2325–2332.
- Y. Shan, Y. Li and H. Pang, *Adv. Funct. Mater.*, 2020, **30**, 200198.
- H. Li, Q. Su, J. Kang, H. Feng, P. Huang, M. Feng, M. Huang and G. Du, *Mater. Res. Bull.*, 2018, **108**, 106–112.
- W. J. Yan, D. Y. Chen, H. R. Fuh, Y. L. Li, D. Zhang, H. Liu, G. Wu, L. Zhang, X. Ren, J. Cho, M. Choi, B. S. Chun, C. Coileáin, H. J. Xu, Z. Wang, Z. Jiang, C. R. Chang and H. C. Wu, *RSC Adv.*, 2019, **9**, 626–635.
- J. Sun, W. Xiong, J. Zhang, Y. Zhang and B. Xie, *Mater. Lett.*, 2021, **308**, 131214.
- C. S. Diko, M. Abitonze, Y. Liu, Y. Zhu and Y. Yang, *Nanomaterials*, 2022, **12**, 4497.
- D. Ma, W. Zhang, Q. Tang, R. Zhang, W. Yu and Y. Qian, *J. Nanosci. Nanotechnol.*, 2005, **5**, 806–809.
- H. Zhu, X. Ji and D. Yang, *J. Mater. Sci.*, 2006, **41**, 3489–3492.
- R. K. Mishra, G. W. Baek, K. Kim, H. I. Kwon and S. H. Jin, *Appl. Surf. Sci.*, 2017, **425**, 923–931.
- B. Hai, K. Tang, C. Wang, C. An, Q. Yang, G. Shen and Y. Qian, *J. Cryst. Growth*, 2001, **225**, 92–95.
- A. Sánchez-Juárez, A. Tiburcio-Silver and A. Ortiz, *Thin Solid Films*, 2005, **480–481**, 452–456.
- G. Su, V. G. Hadjiev, P. E. Loya, J. Zhang, S. Lei, S. Maharjan, P. Dong, P. M. Ajayan, J. Lou and H. Peng, *Nano Lett.*, 2015, **15**, 506–513.
- J. Johnny, S. S. Guzman, B. Krishnan, J. A. A. Martinez, D. Avellaneda Avellaneda and S. Shaji, *Appl. Surf. Sci.*, 2019, **470**, 276–288.
- J. Johnny, S. Sepulveda-Guzman, B. Krishnan, D. Avellaneda and S. Shaji, *Appl. Surf. Sci.*, 2018, **435**, 1285–1295.
- W. Wang, T. Zhang, A. Seliverstov, H. Zhang, Y. Wang, F. Wang, X. Peng, Q. Lu, C. Qin, X. Pan, Y. J. Zeng, C. Van Haesendonck and Z. Ye, *Adv. Electron. Mater.*, 2019, **6**, 1901020.
- S. R. Damkale, S. S. Arbuji, G. G. Umarji, R. P. Panmand, S. K. Khore, R. S. Sonawane, S. B. Rane and B. B. Kale, *Sustainable Energy Fuels*, 2019, **3**, 3406–3414.
- N. Jawale, S. Arbuji, G. Umarji, M. Shinde, B. Kale and S. Rane, *RSC Adv.*, 2023, **13**, 2418–2426.
- A. C. Ok and C. Sarioğlu, *Int. J. Hydrogen Energy*, 2024, **52**, 561–568.
- B. Balan, M. M. Xavier and S. Mathew, *ACS Omega*, 2023, **8**, 25649–25673.
- T. Qiang, L. Chen, Y. Xia and X. Qin, *J. Cleaner Prod.*, 2021, **291**, 125213.
- X. Bai, Y. Du, W. Xue, X. Hu, J. Fan, J. Li and E. Liu, *Nanoscale Adv.*, 2020, **2**, 4220–4228.



- 37 M. Faisal, J. Ahmed, J. S. Algethami, A. S. Alkorbi and F. A. Harraz, *J. Saudi Chem. Soc.*, 2024, **28**, 101806.
- 38 X. Zou, B. Sun, L. Wang, H. Bai, X. Meng, C. Li and Z. Li, *Chem. Eng. J.*, 2024, **482**, 148818.
- 39 P. Huang, F. Chen, Y. Tang, W. Sun, Y. Song and Y. Sun, *Mater. Sci. Semicond. Process.*, 2024, **173**, 108182.
- 40 A. Fakhri, S. Behrouz and M. Pourmand, *J. Photochem. Photobiol., B*, 2015, **149**, 45–50.
- 41 A. Fakhri and S. Behrouz, *Sol. Energy*, 2015, **117**, 187–191.
- 42 V. Gadore, S. R. Mishra and M. Ahmaruzzaman, *J. Hazard. Mater.*, 2023, **461**, 132458.
- 43 A. N. Ech-Chergui, F. Bennabi, M. Isik, Y. Khane, F. J. G. García, A. S. Kadari, M. Guezoul, A. Rahman, M. M. Khan, A. Mehdi, K. Driss-Khodja and B. Amrani, *Colloids Surf., A*, 2024, **686**, 133362.
- 44 F. Chang, Z. Wei, Z. Zhao, Y. Qi and D. Liu, *J. Ind. Eng. Chem.*, 2023, **117**, 265–272.
- 45 J. Yu, C. Y. Xu, F. X. Ma, S. P. Hu, Y. W. Zhang and L. Zhen, *ACS Appl. Mater. Interfaces*, 2014, **6**, 22370–22377.
- 46 H. P. Nogueira, S. H. Toma, A. T. Silveira, A. A. C. Carvalho, A. M. Fioroto and K. Araki, *Microchem. J.*, 2019, **149**, 104025.
- 47 J. R. Tu, X. F. Shi, H. W. Lu, N. X. Yang and Y. J. Yuan, *Mater. Lett.*, 2016, **185**, 303–306.
- 48 T. Qiang and Y. Xia, *J. Alloys Compd.*, 2020, **845**, 156155.
- 49 L. Li, Y. Du, H. Sun, H. Zhang and Z. Zhong, *ChemCatChem*, 2022, **15**, 1436.
- 50 J. Zai, X. Wei, M. Sun, H. Tian, X. Liu, R. Qi and X. Qian, *J. Photochem. Photobiol., A*, 2021, **415**, 113320.
- 51 B. Xia, F. Deng, S. Zhang, L. Hua, X. Luo and M. Ao, *J. Hazard. Mater.*, 2020, **392**, 122345.
- 52 S. Singla, S. Basu and P. Devi, *J. Ind. Eng. Chem.*, 2023, **118**, 119–131.
- 53 M. Kovacic, J. Papac, H. Kusic, P. Karamanis and A. Loncaric Bozic, *Chem. Eng. J.*, 2020, **381**, 1222826.
- 54 N. Afzali, M. Torka Beydokhti, A. A. Khodadadi and Y. Mortazavi, *J. Environ. Chem. Eng.*, 2022, **10**, 107793.
- 55 G. Lu, X. Xie, X. Wang, G. Shi, Q. Zeng, D. Segets and J. Sun, *J. Photochem. Photobiol., A*, 2018, **364**, 725–731.
- 56 A. K. Ganguli, G. B. Kunde, W. Raza, S. Kumar and P. Yadav, *Molecules*, 2022, **27**(22), 7778.
- 57 Y. Liu, D. A. Cullen and T. Lian, *J. Am. Chem. Soc.*, 2021, **143**, 20264–20273.
- 58 W. Gao, J. Lu, S. Zhang, X. Zhang, Z. Wang, W. Qin, J. Wang, W. Zhou, H. Liu and Y. Sang, *Adv. Sci.*, 2019, **6**, 1244.
- 59 B. Ohtani, *Catalysts*, 2013, **3**, 942–953.
- 60 F. Amano, K. Nogami, M. Tanaka and B. Ohtani, *Langmuir*, 2010, **26**, 7174–7180.
- 61 H. Cheng, J. Wang, Y. Zhao and X. Han, *RSC Adv.*, 2014, **4**, 47031–47038.
- 62 N. Roy, Y. Park, Y. Sohn, K. T. Leung and D. Pradhan, *ACS Appl. Mater. Interfaces*, 2014, **6**, 16498–16507.
- 63 Z. Zheng, Z. Wang, L. Xie, Z. Fang, W. Feng, M. Huang and P. Liu, *Appl. Surf. Sci.*, 2015, **353**, 714–722.
- 64 X. Zhang, C. Bo, S. Cao, Z. Cheng, Z. Xiao, X. Liu, T. Tan and L. Piao, *J. Mater. Chem. A*, 2022, **10**, 24381–24387.
- 65 R. A. Carcel, L. Andronic and A. Duta, *Mater. Charact.*, 2012, **70**, 68–73.
- 66 L. Jiang, Z. Y. Jiang, Y. M. Lin and J. M. Zheng, *Phys. Status Solidi RRL*, 2021, **15**, 340.
- 67 I. Shown, S. Samireddi, Y. C. Chang, R. Putikam, P. H. Chang, A. Sabbah, F. Y. Fu, W. F. Chen, C. I. Wu, T. Y. Yu, P. W. Chung, M. C. Lin, L. C. Chen and K. H. Chen, *Nat. Commun.*, 2018, **9**(1), 169.
- 68 X. Shuai, Y. Wang, J. Zhang, R. Zhao, T. Guo, J. Du and J. Li, *ChemistrySelect*, 2022, **7**(33), 1068.
- 69 J. H. Liu, G. F. Huang, W. Q. Huang, H. Miao and B. X. Zhou, *Mater. Lett.*, 2015, **161**, 480–483.
- 70 Y. Zhao, D. Sun, K. Hu, W. Zhao and F. Huang, *Inorg. Chem. Commun.*, 2020, **114**, 107849.
- 71 S. Mondal, S. Das and U. K. Gautam, *J. Colloid Interface Sci.*, 2021, **603**, 110–119.
- 72 S. Paul, D. Barman, C. Chowdhury, P. K. Giri and S. K. De, *CrystEngComm*, 2021, **23**, 2276–2288.
- 73 F. Zhang, L. Shen, J. Li, Y. Zhang, G. Wang and A. Zhu, *Powder Technol.*, 2021, **383**, 371–380.
- 74 Y. Liu, D. Pan, M. Xiong, Y. Tao, X. Chen, D. Zhang, Y. Huang and G. Li, *Chin. J. Catal.*, 2020, **41**, 1554–1563.
- 75 A. Bumajdad and M. Madkour, *Phys. Chem. Chem. Phys.*, 2014, **16**, 7146–7158.
- 76 M. Lin, H. Chen, Z. Zhang and X. Wang, *Phys. Chem. Chem. Phys.*, 2023, **25**, 4388–4407.
- 77 K. Sharma, S. Patial, P. Singh, A. A. P. Khan, V. Saini, A. K. Nadda, C. M. Hussain, V. H. Nguyen, C. C. Nguyen, T. B. Hac Nguyen, S. Y. Kim, Q. Van Le and P. Raizada, *Sol. Energy*, 2022, **231**, 546–565.
- 78 X. Guo, F. Zhang, Y. Zhang and J. Hu, *J. Mater. Chem. A*, 2023, **11**, 7331–7343.
- 79 L. A. Burton, T. J. Whittles, D. Hesp, W. M. Linhart, J. M. Skelton, B. Hou, R. F. Webster, G. O'Dowd, C. Reece, D. Cherns, D. J. Fermin, T. D. Veal, V. R. Dhanak and A. Walsh, *J. Mater. Chem. A*, 2016, **4**, 1312–1318.
- 80 S. Deng, Y. Chen, Q. Li, J. Sun, Z. Lei, P. Hu, Z. H. Liu, X. He and R. Ma, *Nanoscale*, 2022, **14**, 14097–14105.
- 81 P. Shinde and C. S. Rout, *Mater. Chem. Front.*, 2021, **5**, 516–556.
- 82 W. Zhao, Z. Wei, L. Ma, J. Liang and X. Zhang, *Materials*, 2019, **12**(4), 582.
- 83 H. Y. He, J. Lu, L. Y. Cao and M. Li, *Res. Chem. Intermed.*, 2012, **38**, 537–547.
- 84 L. Dashairya, M. Sharma, S. Basu and P. Saha, *J. Alloys Compd.*, 2019, **774**, 625–636.
- 85 R. R. Srivastava, P. Kumar Vishwakarma, U. Yadav, S. Rai, S. Umrao, R. Giri, P. S. Saxena and A. Srivastava, *Front. Nanotechnol.*, 2021, **3**, 711368.
- 86 Z. Li, R. Fan, Z. Hu, W. Li, H. Zhou, S. Kang, Y. Zhang, H. Zhang and G. Wang, *J. Hazard. Mater.*, 2020, **394**, 122525.
- 87 V. Govindan, H. Imran, V. Dharuman and K. Sankaranarayanan, *J. Mater. Sci.: Mater. Electron.*, 2018, **29**, 17670–17680.
- 88 Z. Liu, C. Liu, S. Mao and X. Huang, *ACS Appl. Mater. Interfaces*, 2023, **15**, 7529–7537.



- 89 Y. Kawabe, Y. Ito, Y. Hori, S. Kukunuri, F. Shiokawa, T. Nishiuchi, S. Jeong, K. Katagiri, Z. Xi, Z. Li, Y. Shigeta and Y. Takahashi, *ACS Nano*, 2023, **17**, 11318–11326.
- 90 H. Wang, Z. Liu, L. Wang, Q. Shou, M. Gao, H. Wang, A. Nazir and P. Huo, *J. Mater. Sci.: Mater. Electron.*, 2023, **34**, 350.
- 91 C. Jin, W. Feng and X. Zhao, *IOP Conference Series: Earth and Environmental Science*, IOP Publishing Ltd, 2021, vol. 766.
- 92 Y. Nie, J. Liu, N. Li, Y. Wang, Q. Cheng, S. He, Q. Guo, R. Zhao and F. Pan, *Fuel Process. Technol.*, 2023, **250**, 107871.
- 93 L. Xiang, S. Liu, L. Zhao, S. Yuan, X. Li and N. Li, *J. Alloys Compd.*, 2023, **945**, 169201.
- 94 M. Lan, X. Dong, N. Zheng, X. Zhang, Y. Wang and X. Zhang, *J. Mater. Sci. Technol.*, 2023, **167**, 237–247.
- 95 S. R. Kadam, S. Ghosh, R. Bar-Ziv and M. Bar-Sadan, *Chem. – Eur. J.*, 2020, **26**, 6679–6685.
- 96 J. W. Shi, Y. Zou, D. Ma, Z. Fan, L. Cheng, D. Sun, Z. Wang, C. Niu and L. Wang, *Nanoscale*, 2018, **10**, 9292–9303.
- 97 T. Di, T. Cao, H. Liu, S. Wang and J. Zhang, *Phys. Chem. Chem. Phys.*, 2023, **25**, 5196–5202.
- 98 A. R. Woldu, P. Talebi, A. G. Yohannes, J. Xu, X. D. Wu, S. Siahrostami, L. Hu and X. C. Huang, *Angew. Chem., Int. Ed.*, 2023, **62**, e2023016.
- 99 T. Billo, I. Shown, A. Anbalagan, T. A. Effendi, A. Sabbah, F. Y. Fu, C. M. Chu, W. Y. Woon, R. S. Chen, C. H. Lee, K. H. Chen and L. C. Chen, *Nano Energy*, 2020, **72**, 104717.
- 100 X. Jiao, X. Li, X. Jin, Y. Sun, J. Xu, L. Liang, H. Ju, J. Zhu, Y. Pan, W. Yan, Y. Lin and Y. Xie, *J. Am. Chem. Soc.*, 2017, **139**, 18044–18051.
- 101 G. Li, W. Liu, J. Wang, B. Li, S. Yang, W. Wang, W. Sun and Y. Sun, *Chem. Phys. Lett.*, 2022, **807**, 140063.
- 102 G. Ma, Z. Pan, Y. Liu, Y. Lu and Y. Tao, *Materials*, 2023, **16**, 4436.
- 103 G. Kumar, J. Kumar, M. Bag and R. Kumar Dutta, *Sep. Purif. Technol.*, 2022, **292**, 121040.
- 104 G. Kumar and R. K. Dutta, *Environ. Sci. Pollut. Res.*, 2022, **29**, 57758–57772.
- 105 Y. C. Zhang, Z. N. Du, K. W. Li and M. Zhang, *Sep. Purif. Technol.*, 2011, **81**, 101–107.
- 106 L. Li, Z. Chai, W. Jin, H. Sun, J. He, G. Wu and W. Xia, *J. Alloys Compd.*, 2023, **932**, 167658.
- 107 X. Zhang and P. Yang, *Carbon*, 2024, **216**, 118584.
- 108 Y. Sun, G. Li, J. Xu and Z. Sun, *Mater. Lett.*, 2016, **174**, 238–241.
- 109 G. Li, Y. Sun, S. Sun, W. Chen, J. Zheng, F. Chen, Z. Sun and W. Sun, *Adv. Powder Technol.*, 2020, **31**, 2505–2512.
- 110 F. Wang, S. Zhang, W. Jing, H. Qiu, Y. Liu and L. Guo, *J. Mater. Sci. Technol.*, 2024, **189**, 146–154.
- 111 T. Song, X. Zhang, C. Xie and P. Yang, *Carbon*, 2023, **210**, 118052.
- 112 K. Perović, M. Kovačić, M. Kraljić Roković, H. Kušić, B. Genorio, U. Lavrenčić Štangar and A. Lončarić Božić, *Mater. Res. Bull.*, 2023, **167**, 112418.
- 113 J. Wang, J. Xuan, X. Wei, Y. Zhang, J. Fan, L. Ni, Y. Yang, J. Liu, Y. Tian and L. Duan, *Int. J. Hydrogen Energy*, 2024, **54**, 979–989.
- 114 Y. Li, Z. Liu, S. Wu, M. Zhu and Y. Zhang, *Chem. Phys. Lett.*, 2022, **812**, 140248.
- 115 M. Awais, S. Aslam, M. N. Ashiq, M. Mirza and M. Safdar, *New J. Chem.*, 2024, **48**, 3247–3257.
- 116 P. A. K. Reddy, H. Han, K. C. Kim and S. Bae, *ACS Sustainable Chem. Eng.*, 2024, **12**, 4979–4992.
- 117 N. Ma, C. Lu, Y. Liu, T. Han, W. Dong, D. Wu and X. Xu, *Small*, 2024, **20**, 2304839.
- 118 A. Fakhri, S. Behrouz and M. Pourmand, *J. Photochem. Photobiol., B*, 2015, **149**, 45–50.
- 119 X. Zhang, P. Yang, H. S. Chen and S. P. Jiang, *Chem. Eng. J.*, 2023, **479**, 147609.
- 120 D. Wang, P. Zhao, J. Yang, G. Xu, H. Yang, Z. Shi, Q. Hu, B. Dong and Z. Guo, *Colloids Surf., A*, 2020, **603**, 125147.
- 121 S. Thirumalairajan, K. Girija, V. R. Mastelaro and N. Ponpandian, *New J. Chem.*, 2014, **38**, 5480–5490.
- 122 K. Tamilarasu, R. Ranjith, P. Maadeswaran, R. Ramesh, R. Thammasak, G. Periyasami, P. Karthikeyan and C. Umarani, *J. Mater. Sci.: Mater. Electron.*, 2024, **35**, 607.
- 123 Y. Ben Smida, O. Oyewo, S. Ramaila, L. Mavuru, R. Marzouki, D. C. Onwudiwe and A. H. Hamzaoui, *J. Inorg. Organomet. Polym. Mater.*, 2022, **32**, 4679–4693.
- 124 Y. Cheng, J. He and P. Yang, *Colloids Surf., A*, 2023, **680**, 132678.
- 125 K. M. Alnahdi, *Inorg. Chem. Commun.*, 2023, **160**, 111870.
- 126 X. Zheng, M. Xu, C. Cai, Y. Yuan, F. Lin, W. Chen and F. Yang, *J. Alloys Compd.*, 2024, **980**, 173630.
- 127 V. Gadore, S. R. Mishra and M. Ahmaruzzaman, *J. Hazard. Mater.*, 2023, **444**, 130301.
- 128 S. Park, R. Selvaraj, M. A. Meetani and Y. Kim, *J. Ind. Eng. Chem.*, 2017, **45**, 206–214.
- 129 X. Wu, H. Hu, L. Cheng, Y. Zhang, Q. Jiang, P. Wang, L. Xu, P. Lin and C. Cui, *J. Mater. Sci.: Mater. Electron.*, 2024, **35**, 545.
- 130 P. Cao, Y. Zhang, D. Gao, H. Chen, M. Zhou, Y. He, P. Song and R. Wang, *J. Alloys Compd.*, 2022, **904**, 164061.
- 131 T. T. Salunkhe, V. Kumar, A. N. Kadam, M. Mali and M. Misra, *Ceram. Int.*, 2024, **50**, 1826–1835.
- 132 G. Kumar, *J. Inorg. Organomet. Polym. Mater.*, 2023, **33**, 2710–2720.
- 133 M. Rajamani, A. Rajan and B. Neppolian, *J. Environ. Chem. Eng.*, 2022, **11**, 109129.
- 134 G. Kumar and R. K. Dutta, *J. Phys. Chem. Solids*, 2022, **164**, 110639.
- 135 F. You, X. Hou, P. Wei and J. Qi, *Inorg. Chem.*, 2021, **60**(24), 18598–18602.
- 136 F. You, Y. Zhou, D. Li, H. Zhang, D. Gao, X. Ma, R. Hao and J. Liu, *J. Colloid Interface Sci.*, 2023, **629**, 871–877.
- 137 T. Di, B. Zhu, B. Cheng, J. Yu and J. Xu, *J. Catal.*, 2017, **352**, 532–541.
- 138 N. Jawale, S. Arbuji, G. Umarji, M. Shinde, B. Kale and S. Rane, *RSC Adv.*, 2023, **13**, 2418–2426.



- 139 G. Li, Y. Sun, X. Hong, W. Lu, W. Chen, Y. Deng, Z. Sun and W. Sun, *J. Mater. Sci.*, 2021, **56**, 10847–10858.
- 140 S. Zhu, G. Wu, Z. Liu, S. Zhao, D. Cao, C. Li and G. Liu, *Ceram. Int.*, 2023, **49**, 5893–5904.
- 141 H. Zhao, B. Zhao, H. Liu and X. Li, *Dalton Trans.*, 2023, **53**, 591–600.
- 142 F. Ganaie, Z. Haq, A. Bashir, A. Qureashi, I. Nazir, K. Fatima, A. H. Pandith and M. A. Bhat, *New J. Chem.*, 2024, **48**, 7111–7124.
- 143 S. Q. Guo, B. Yang, Z. Hu, M. Zhen, B. Gu and B. Shen, *Nano Res.*, 2023, **16**, 2102–2110.
- 144 S. Liu, *Mater. Lett.*, 2023, **330**, 133256.
- 145 X. He, T. He, R. Xia, Y. Qi, Y. Zhou, B. Song, F. Liao, Z. Kang and L. Jiang, *Sep. Purif. Technol.*, 2023, **324**, 124515.
- 146 Z. Wan, Q. Mao, J. Xiang, D. Ma and H. Tang, *J. Mater. Sci. Technol.*, 2023, **161**, 233–244.
- 147 S. Kar, T. Pal and S. Ghosh, *ChemistrySelect*, 2023, **8**, e202300878.
- 148 S. Bai, N. Zhang, C. Gao and Y. Xiong, *Nano Energy*, 2018, **53**, 296–336.
- 149 J. Liu, Z. Wei and W. Shangguan, *ChemCatChem*, 2019, **11**, 6177–6189.
- 150 H. Zhang, L. Mao, J. Wang, Y. Nie, Z. Geng, D. Zhong, X. Tan, J. Ye and T. Yu, *Small*, 2023, **20**, 5727.
- 151 M. Chen, S. Wan, L. Zhong, D. Liu, H. Yang, C. Li, Z. Huang, C. Liu, J. Chen, H. Pan, D. S. Li, S. Li, Q. Yan and B. Liu, *Angew. Chem., Int. Ed.*, 2021, **60**, 26233–26237.
- 152 G. Kiruthigaa, C. Manoharan, M. Bououdina, S. Ramalingam and C. Raju, *Solid State Sci.*, 2015, **44**, 32–38.
- 153 Y. Liu, Y. Zhou, X. Zhou, X. Jin, B. Li, J. Liu and G. Chen, *Chem. Eng. J.*, 2021, **407**, 127180.
- 154 A. Sharma, A. Makhija, S. Dahiya, A. Ohlan, R. Punia and A. S. Maan, *Mater. Res. Bull.*, 2023, **168**, 112464.
- 155 M. Setayeshmehr, M. Haghghi and K. Mirabbaszadeh, *Electrochim. Acta*, 2021, **376**, 137987.
- 156 P. Bharathi, S. Harish, M. Shimomura, M. Krishna Mohan, J. Archana and M. Navaneethan, *Mater. Lett.*, 2023, **335**, 133691.
- 157 P. Sengodu, C. H. Li, C. F. Wei, R. Bendi and C. C. Chen, *ChemistrySelect*, 2016, **1**, 3328–3334.
- 158 C. Mondal, M. Ganguly, J. Pal, A. Roy, J. Jana and T. Pal, *Langmuir*, 2014, **30**, 4157–4164.
- 159 X. Zhang, A. Chen, L. Chen and Z. Zhou, *Adv. Energy Mater.*, 2022, **12**, 2003841.
- 160 Y. Hu, X. Chen, X. Ren, Z. Huang, X. Qi and J. Zhong, *J. Mater. Sci.: Mater. Electron.*, 2018, **29**, 19614–19619.
- 161 B. Luo, G. Liu and L. Wang, *Nanoscale*, 2016, **8**, 6904–6920.
- 162 J. Zhang, G. Huang, J. Zeng, X. Jiang, Y. Shi, S. Lin, X. Chen, H. Wang, Z. Kong, J. Xi and Z. Ji, *J. Alloys Compd.*, 2019, **775**, 726–735.
- 163 Y. Liu, X. Mi, J. Wang, M. Li, D. Fan, H. Lu and X. Chen, *Inorg. Chem. Front.*, 2019, **6**, 948–954.
- 164 Y. C. Zhang, J. Li, M. Zhang and D. D. Dionysiou, *Environ. Sci. Technol.*, 2011, **45**, 9324–9331.
- 165 J. Gajendiran and V. Rajendran, *Adv. Nat. Sci.: Nanosci. Nanotechnol.*, 2021, **2**(1), 15001.
- 166 S. H. Chaki, M. P. Deshpande, D. P. Trivedi, J. P. Tailor, M. D. Chaudhary and K. Mahato, *Appl. Nanosci.*, 2013, **3**, 189–195.
- 167 X. L. Gou, J. Chen and P. W. Shen, *Mater. Chem. Phys.*, 2005, **93**, 557–566.
- 168 Y. Ren, H. An, W. Zhang, S. Wei, C. Xing and Z. Peng, *Nanophotonics*, 2022, **11**, 4781–4792.
- 169 H. Zhao, X. Li, Y. Chen, Z. Zhao, M. Zhang, H. Su, H. Liang, W. Zhang, K. Tao and L. Li, *J. Lumin.*, 2023, **235**, 118068.
- 170 Y. Huang, W. Jiao, Z. Chu, X. Nie, R. Wang and X. He, *ACS Appl. Mater. Interfaces*, 2020, **12**, 25178–25188.
- 171 Y. J. Yuan, D. Q. Chen, X. F. Shi, J. R. Tu, B. Hu, L. X. Yang, Z. T. Yu and Z. G. Zou, *Chem. Eng. J.*, 2017, **313**, 1438–1446.
- 172 T. Ali, X. Wang, K. Tang, Q. Li, S. Sajjad, S. Khan, S. A. Farooqi and C. Yan, *Electrochim. Acta*, 2019, **300**, 45–52.
- 173 B. Wang, M. Chen, J. Lv, G. Xu, X. Shu and Y. C. Wu, *Dalton Trans.*, 2021, **50**, 13329–13336.
- 174 D. Ursu, M. Vajda and M. Miclau, *Micro Nano Lett.*, 2019, **14**, 872–876.
- 175 Y. Liang, N. Guo, L. Li, R. Li, G. Ji and S. Gan, *Appl. Surf. Sci.*, 2015, **332**, 32–39.
- 176 J. M. Mali, S. S. Arbuj, J. D. Ambekar, S. B. Rane, U. P. Mulik and D. P. Amalnerkar, *Sci. Adv. Mater.*, 2013, **5**, 1994–1998.
- 177 Z. Liu, K. Xu, H. Yu and Z. Sun, *Int. J. Energy Res.*, 2021, **45**, 6850–6862.
- 178 L. L. Zulfa, R. Ediati, A. R. P. Hidayat, R. Subagyo, N. Faaizatunnisa, Y. Kusumawati, D. Hartanto, N. Widiastuti, W. P. Utomo and M. Santoso, *RSC Adv.*, 2023, **13**, 3818–3834.
- 179 Y. Zhong, C. Peng, Z. He, D. Chen, H. Jia, J. Zhang, H. Ding and X. Wu, *Catal. Sci. Technol.*, 2021, **11**, 27–42.
- 180 G. Li and K. A. Gray, *Chem. Phys.*, 2007, **339**, 173–187.
- 181 A. Mohanty and K. Kamali, *ACS Appl. Nano Mater.*, 2024, **7**, 3326–3338.
- 182 N. R. Barveen, J. L. Xu and Y. W. Cheng, *J. Environ. Chem. Eng.*, 2024, **12**, 112200.
- 183 J. Feng, Y. Sun, J. Mu, L. Chen, T. Han, H. Miao, E. Liu, J. Fan and X. Hu, *Mater. Lett.*, 2019, **236**, 534–537.
- 184 S. Mondal, L. Sahoo, C. P. Vinod and U. K. Gautam, *Appl. Catal., B*, 2021, **2**, 119927.
- 185 S. I. Naya, K. Kimura and H. Tada, *ACS Catal.*, 2013, **3**, 10–13.
- 186 R. Fu, L. Li, X. Li, B. Li, C. Shao, Z. Liu and A. Shen, *Mater. Chem. Phys.*, 2021, **267**, 124702.
- 187 T. Wang, H. Meng, X. Yu, Y. Liu, H. Chen, Y. Zhu, J. Tang, Y. Tong and Y. Zhang, *RSC Adv.*, 2015, **5**, 15469–15478.
- 188 M. Bagherzadeh and R. Kaveh, *J. Photochem. Photobiol., A*, 2018, **359**, 11–22.
- 189 H. Derikvandi and A. Nezamzadeh-Ejehieh, *J. Colloid Interface Sci.*, 2017, **490**, 628–641.



- 190 L. Sun, Z. Zhao, S. Li, Y. Su, L. Huang, N. Shao, F. Liu, Y. Bu, H. Zhang and Z. Zhang, *ACS Appl. Nano Mater.*, 2019, **2**, 2144–2151.
- 191 Q. Liu, S. Liu, A. Wu, H. Huang and L. Zhou, *J. Alloys Compd.*, 2020, **834**, 155174.
- 192 L. Zhang, X. Dong, Y. Wang, N. Zheng, H. Ma and X. Zhang, *Appl. Surf. Sci.*, 2022, **579**, 152088.
- 193 L. Wang, S. K. Karuturi and L. Zan, *Appl. Surf. Sci.*, 2023, **537**, 148063.
- 194 Y. Liu, C. Lv, J. Sun, X. Zhou, Y. Zhou and G. Chen, *Adv. Mater. Interfaces*, 2022, **9**, 2200153.
- 195 Sunaina, K. K. Yadav, Ankush, S. K. Guchhait, K. Sood, S. K. Mehta, A. K. Ganguli and M. Jha, *Sep. Purif. Technol.*, 2020, **242**, 116835.
- 196 J. Pan, Z. Guan, J. Yang and Q. Li, *Chin. J. Catal.*, 2020, **41**, 200–208.
- 197 F. Qiu, W. Li, F. Wang, H. Li, X. Liu and J. Sun, *J. Colloid Interface Sci.*, 2017, **493**, 1–9.
- 198 X. Chen, Z. Han, Z. Lu, T. Qu, C. Liang, Y. Wang, B. Zhang, X. Han and P. Xu, *Sustainable Energy Fuels*, 2023, **7**, 1311–1321.
- 199 W. Ren, J. Yang, W. Chen, J. Zhang, Y. Sun, Y. Zheng, H. Zhao and B. Liang, *Mater. Res. Bull.*, 2022, **153**, 111884.
- 200 C. Zhang, J. Ma, H. Zhu, H. Ding, H. Wu, K. Zhang, X. L. Zhao, X. Wang and C. Cheng, *J. Alloys Compd.*, 2023, **960**, 170932.
- 201 Q. Li, S. He, L. Wang, J. Song, J. Wang, C. Shao, Z. Tian and Y. Liu, *CrystEngComm*, 2023, **25**, 2882–2891.

