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## Silver frameworks based on a tetraphenylethylene–imidazole ligand for electrocatalytic reduction of CO<sub>2</sub> to CO†

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Metal–organic frameworks (MOFs) can be used as electrocatalysts for the CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR) because of their well-dispersed metal centers. Silver is a common electrocatalyst for reduction of CO<sub>2</sub> to CO. In this study, two Ag-MOFs with different structures of [Ag<sub>8</sub>O<sub>2</sub>(TIPE)<sub>6</sub>](NO<sub>3</sub>)<sub>4</sub> (**Ag-MOF1**) and [Ag(TIPE)<sub>0.5</sub>CF<sub>3</sub>SO<sub>3</sub>] (**Ag-MOF2**) [TIPE = 1,1,2,2-tetrakis(4-(imidazol-1-yl)phenyl)ethene] were synthesized and used for CO<sub>2</sub> electroreduction. The results show that **Ag-MOF2** is superior to **Ag-MOF1** and exhibits high CO faradaic efficiency (FE) of 92.21% with partial current density of 29.51 mA cm<sup>-2</sup> at -0.98 V versus reversible hydrogen electrode (RHE). The FE<sub>CO</sub> is higher than 80% in the potential range of -0.78 to -1.18 V. The difference may be caused by different framework structures leading to different electrochemical active surface areas and charge transfer kinetics. This study provides a new strategy for designing and constructing CO<sub>2</sub> electroreduction catalysts and provides potential ways for solving environmental and energy problems caused by excessive CO<sub>2</sub> emission.

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## Introduction

The environmental problems caused by rapid consumption of fossil fuels such as coal, oil and natural gas disrupt the carbon balance of nature.<sup>1</sup> Researchers have designed and developed a series of methods for storing and converting carbon dioxide (CO<sub>2</sub>) to reduce its content in the atmosphere and thereby alleviate the greenhouse effect.<sup>2–4</sup> It has been reported that CO<sub>2</sub> can be converted into usable high-value energy products or chemicals, which can simultaneously solve the problem of energy depletion.<sup>5,6</sup> Due to the latest developments in renewable energy generation, such as solar and wind, the electrocatalytic CO<sub>2</sub> reduction reaction (CO<sub>2</sub>RR) has become one of the most attractive pathways for CO<sub>2</sub> conversion due to its simple installation, high efficiency and low energy consumption.<sup>7,8</sup> However, due to the chemical inertness of CO<sub>2</sub> and its high C=O chemical bond dissociation energy (~750 kJ mol<sup>-1</sup>), the transformation of CO<sub>2</sub> needs to overcome considerable kinetic obstacles arising from the thermo-

dynamic stability and low electron affinity of CO<sub>2</sub>.<sup>9</sup> In addition, the process of electrocatalytic reduction of CO<sub>2</sub> is accompanied by the production of the competitive by-product H<sub>2</sub>, which leads to a decrease in the faradaic efficiency of the catalyst.<sup>10</sup> Therefore, development of efficient, highly selective and stable CO<sub>2</sub> electrocatalysts is the focus of current research works.<sup>11</sup>

As a typical CO<sub>2</sub> reduction product, CO is of great significance for synthesizing carbon-based fuels and reducing CO<sub>2</sub> emissions. In addition, the reduction of CO<sub>2</sub> to CO has high selectivity and low separation cost. In the past few decades, electrocatalysts for CO<sub>2</sub>RRs such as metals, non-metallic two-dimensional (2D) materials and metal–organic frameworks (MOFs) have been widely studied.<sup>9,12–14</sup> Compared with traditional catalysts, MOFs composed of metal nodes and organic connectors have advantages, such as well-defined framework structures and well-dispersed metal centers,<sup>15–19</sup> which are conducive to establishing accurate structure–activity relationships for the CO<sub>2</sub>RR and exhibit excellent prospects in CO<sub>2</sub> electroreduction.<sup>20–22</sup> However, the conductivity of MOFs is poor, and they cannot provide sufficient electrons during the CO<sub>2</sub> electrochemical reduction process. In addition, the hydrogen evolution reaction (HER) is a competitive reaction in the CO<sub>2</sub> electroreduction process; thus, inhibiting the HER is key to improving selectivity. Therefore, simultaneously assembling electron donating groups and active components for CO<sub>2</sub> electroreduction into MOFs to improve CO<sub>2</sub> electroreduction efficiency and selectivity is a major challenge at present. It has been reported that metals such as Ag, Zn and In, located on

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the left branch of the Trassati volcanic curve, are valid catalysts with high CO<sub>2</sub>RR selectivity.<sup>10,23,24</sup>

Inspired by the above-mentioned considerations, 1,1,2,2-tetrakis(4-(imidazol-1-yl)phenyl)ethene (TIPE) as an electron-rich organic ligand with the  $\pi$ -conjugated tetraphenylethylene moiety and different silver salts were used to synthesize Ag-MOFs. A cubic Ag-MOF based on Ag–oxygen clusters, named as [Ag<sub>8</sub>O<sub>2</sub>(TIPE)<sub>6</sub>](NO<sub>3</sub>)<sub>4</sub> (**Ag-MOF1**), was obtained by using silver nitrate (AgNO<sub>3</sub>). Meanwhile, silver trifluoromethanesulfonate (AgCF<sub>3</sub>SO<sub>3</sub>) and TIPE were utilized to synthesize an Ag-MOF, and [Ag(TIPE)<sub>0.5</sub>CF<sub>3</sub>SO<sub>3</sub>] (**Ag-MOF2**) was obtained. The results show that **Ag-MOF2** has superior electrocatalytic performance for CO<sub>2</sub> reduction with high faradaic efficiency (FE) as well as high partial current density for CO compared with that of **Ag-MOF1**, implying that the MOF structure has great impact on electrocatalytic reduction of CO<sub>2</sub>.

## Experimental

### Synthesis of [Ag<sub>8</sub>O<sub>2</sub>(TIPE)<sub>6</sub>](NO<sub>3</sub>)<sub>4</sub> (**Ag-MOF1**)

TIPE (0.05 mmol, 29.83 mg), 4,4'-oxybisbenzoic acid (H<sub>2</sub>OBA) (0.1 mmol, 25.80 mg) and AgNO<sub>3</sub> (0.12 mmol, 20.40 mg) in a mixed solvent of 5 mL CH<sub>3</sub>CN and 4 mL H<sub>2</sub>O were added to a Teflon-lined stainless autoclave, which was heated at 160 °C for 56 h. The resulting mixture was cooled to ambient temperature and orange transparent cubic crystals were obtained in 35% yield based on TIPE. Anal. calcd for C<sub>228</sub>H<sub>168</sub>Ag<sub>8</sub>N<sub>52</sub>O<sub>14</sub>: C, 57.98%; H, 3.59%; N, 15.42%. Found: C, 57.66%; H, 3.38%; N, 14.97%. IR (cm<sup>-1</sup>, Fig. S1, ESI<sup>†</sup>): 3120 (m), 3043 (s), 1682 (s), 1602 (s), 1520 (w), 1491 (m), 1423 (s), 1355 (w), 1304 (m), 1254 (m).

### Synthesis of [Ag(TIPE)<sub>0.5</sub>CF<sub>3</sub>SO<sub>3</sub>] (**Ag-MOF2**)

TIPE (0.05 mmol, 29.83 mg) and AgCF<sub>3</sub>SO<sub>3</sub> (0.12 mmol, 30.83 mg) in a mixed solvent of 8 mL CH<sub>3</sub>CN and 4 mL H<sub>2</sub>O were added to a Teflon-lined stainless autoclave, which was heated at 160 °C for 96 h. The resulting mixture was cooled to room temperature, and yellow transparent crystals were obtained in 81% yield based on TIPE. Anal. calcd for C<sub>20</sub>H<sub>14</sub>AgF<sub>3</sub>N<sub>4</sub>O<sub>3</sub>S: C, 43.26%; H, 2.54%; N, 10.09%. Found: C, 42.97%; H, 2.32%; N, 9.76%. IR (cm<sup>-1</sup>, Fig. S1, ESI<sup>†</sup>): 3118 (m), 1756 (s), 1602 (s), 1519 (w), 1420 (s), 1307 (s), 1263 (w), 1160 (m).

### Electrochemical measurements and CO<sub>2</sub> reduction experiments

The catalyst (2 mg) was uniformly ground to a powder and dispersed into a mixed solution of 10  $\mu$ L of Nafion (5 wt%) and 190  $\mu$ L of isopropanol, then sonicated for 10 min to form a homogeneous ink. 200  $\mu$ L of the well-dispersed ink was loaded onto carbon paper (CP) with an area of 1 cm<sup>2</sup> and dried at 120 °C on a heating table. Electrochemical measurements were performed on a CHI 730E electrochemical workstation (Shanghai, Chenhua, China) using a three-channel flow cell comprising the Ag-MOF/CP as a working electrode, Pt as a

counter electrode and Ag/AgCl (saturated KCl electrolyte) as a reference electrode. For CO<sub>2</sub> reduction experiments, the cathode and anode chambers were separated by an anion exchange membrane (Fumasep, FAB-PK-130) while the electrolyte was 1 M KOH. The electrolyte in the cathode and anode was circulated by a peristaltic pump (Kamoer, F01A-STP) with a flow rate of 10 mL min<sup>-1</sup>. High purity CO<sub>2</sub> (99.9995%) gas was continuously passed through the flow chamber with a flow rate of 20 mL min<sup>-1</sup> via a mass flow controller (Sevenstar, D07-7B). All potentials were measured versus an Ag/AgCl reference electrode (with saturated KCl) without *iR* compensation. The potentials were converted to the RHE scale using the following equation:  $E(\text{versus RHE}) = E(\text{versus Ag/AgCl}) + 0.197 + 0.059 \times \text{pH}$ . All experiments were carried out at room temperature.

## Results and discussion

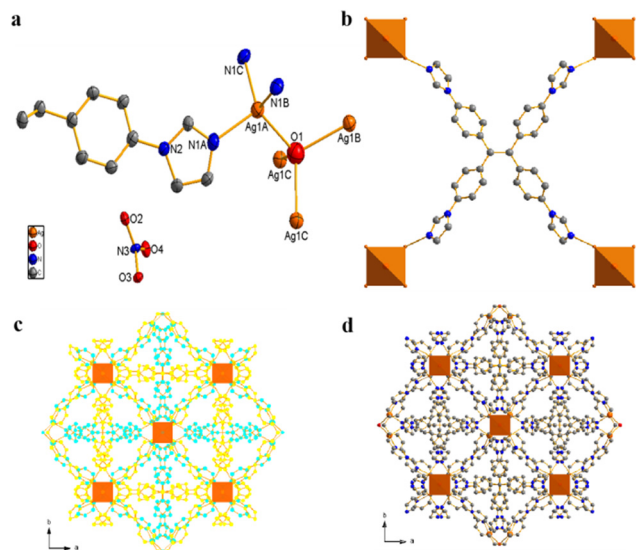
### Crystal structures and characterization of Ag-MOFs

**Ag-MOF1** was isolated by reaction of TIPE and AgNO<sub>3</sub> together with H<sub>2</sub>OBA. It is noteworthy that no **Ag-MOF1** was obtained without addition of H<sub>2</sub>OBA. This means that H<sub>2</sub>OBA may play a role in the formation of **Ag-MOF1** although there is no H<sub>2</sub>OBA in the formed **Ag-MOF1**. Single-crystal X-ray diffraction (SC-XRD) analysis revealed that **Ag-MOF1** belongs to cubic space group *P* $\bar{4}3n$  (Table 1). The asymmetric unit contains one twenty-fourth molecule of [Ag<sub>8</sub>O<sub>2</sub>(TIPE)<sub>6</sub>](NO<sub>3</sub>)<sub>4</sub> and there are Ag<sub>4</sub>O clusters, in which each O<sup>2-</sup> ion links four Ag(I) ions. Each Ag(I) is four-coordinated with distorted tetrahedral configuration and surrounded by one O<sup>2-</sup> and three imidazole N-atoms from three different TIPE ligands (Fig. 1a). The Ag(1)–O(1) and Ag(1)–N(1) distances are of 2.5898(14) Å and 2.314(9) Å, respectively, and the corresponding angles around the Ag(I)

**Table 1** Crystal data and structure refinements for **Ag-MOF1** and **Ag-MOF2**

Compound	<b>Ag-MOF1</b>	<b>Ag-MOF2</b>
CCDC	2311795	2311796
Empirical formula	C <sub>19</sub> H <sub>14</sub> Ag <sub>0.67</sub> N <sub>4.33</sub> O <sub>1.17</sub>	C <sub>20</sub> H <sub>14</sub> AgF <sub>3</sub> N <sub>4</sub> O <sub>3</sub> S
Formula weight	393.59	555.28
Crystal system	Cubic	Monoclinic
Space group	<i>P</i> $\bar{4}3n$	<i>C2/m</i>
Temperature/K	193	193
<i>a</i> /Å	17.2542(3)	13.9254(4)
<i>b</i> /Å	17.2542(3)	14.9602(5)
<i>c</i> /Å	17.2542(3)	11.7909(4)
$\beta$ /°	90	119.6310(10)
Volume/Å <sup>3</sup>	5136.7(3)	2135.13(12)
<i>Z</i>	12	4
<i>D<sub>c</sub></i> /g cm <sup>-3</sup>	1.527	1.727
$\mu$ /mm <sup>-1</sup>	4.483	6.090
<i>F</i> (000)	2388	1104
Reflections collected	15 859	9034
Independent reflections	1885	2024
Goodness-of-fit	1.019	1.061
<i>R</i> <sub>1</sub> <sup>a</sup> [ <i>I</i> > 2 $\sigma$ ( <i>I</i> )]	0.0662	0.0567
<i>wR</i> <sub>2</sub> <sup>b</sup> [ <i>I</i> > 2 $\sigma$ ( <i>I</i> )]	0.1710	0.1616

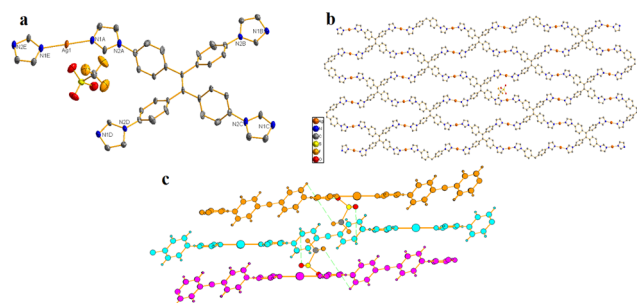
<sup>a</sup>  $R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ . <sup>b</sup>  $wR_2 = \sqrt{\sum w(|F_o|^2 - |F_c|^2)^2} / \sqrt{\sum w(F_o)^2}^{1/2}$ , where  $w = 1/[\sigma^2(F_o^2) + (aP)^2 + bP]$ .  $P = (F_o^2 + 2F_c^2)/3$ .



**Fig. 1** (a) Coordination environment of Ag(i) in **Ag-MOF1** with the ellipsoids drawn at the 30% probability level. Hydrogen atoms are omitted for clarity. (b) Coordination environment of TIPE. (c) and (d) Interpenetrated 3D structures of **Ag-MOF1** along the *c*-axis. Free nitrate anions are omitted for clarity.

atom vary in the range of  $106.8(2)$ – $130.5(7)^\circ$  (Table S1, ESI<sup>†</sup>), all of which coincide with those reported in the literature. Each TIPE ligand contains four imidazole N-atom donors and connects four  $\text{Ag}_4\text{O}$  clusters as a 4-connected node (Fig. 1b). Each  $\text{Ag}_4\text{O}$  cluster joins with 12 TIPE ligands, leading to 12-connected nodes. Thus, **Ag-MOF1** has a (4,12)-connected architecture. Moreover, owing to the action of the ligand and the bridge connection of metal ions, **Ag-MOF1** forms a two-fold interpenetrating three-dimensional (3D) framework (Fig. 1c and d). The void volume in **Ag-MOF1** is  $208.5 \text{ \AA}^3$  calculated by PLATON, which is 4.1% of the unit cell volume ( $5136.7 \text{ \AA}^3$ ).

SC-XRD analysis shows that **Ag-MOF2** crystallizes in the monoclinic crystal system with  $C2/m$  space group (Table 1). The repeat unit of **Ag-MOF2** has a half molecule of  $[\text{Ag}(\text{TIPE})_{0.5}\text{CF}_3\text{SO}_3]$ . Each Ag(i) is coordinated by two nitrogen atoms from two distinct TIPE ligands (Fig. 2a). Each TIPE links



**Fig. 2** (a) Coordination environment of Ag(i) in **Ag-MOF2** with the ellipsoids drawn at the 30% probability level. Hydrogen atoms are omitted for clarity. (b) 2D structure of **Ag-MOF2**. (c) 3D supramolecular structure of **Ag-MOF2** with hydrogen bonds indicated by dashed lines.

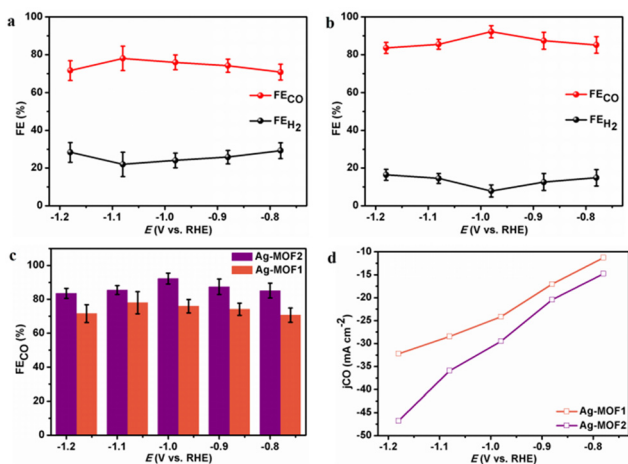
four Ag(i) atoms to form a two-dimensional (2D) layered structure of **Ag-MOF2** (Fig. 2b). The 2D layers are further connected by C–H...F and C–H...O hydrogen bonding interactions between TIPE and  $\text{CF}_3\text{SO}_3^-$  to form a 3D supramolecular structure of **Ag-MOF2** (Fig. 2c). The results indicate that anions have a significant impact on the formation of Ag-MOFs.

### Phase and morphology characterization of Ag-MOFs

The phase purity and chemical stability of Ag-MOFs were examined by powder X-ray diffraction (PXRD). As shown in Fig. S2 and S3, ESI<sup>†</sup> PXRD patterns of the synthesized samples are in good agreement with the simulated ones, indicating that the synthesized bulky samples are consistent with the structures determined by SC-XRD. It is worth noting that the PXRD patterns of samples do not change significantly after immersion in 1 M KOH solution for 24 h, implying that the samples have high chemical stability in harsh alkaline environments. Fourier transform infrared (FT-IR) spectroscopic data show that the bands around  $3000$ – $3100$  and  $1400$ – $1600 \text{ cm}^{-1}$  correspond to the stretching vibrations of aromatic C–H and C=C bonds, respectively (Fig. S1, ESI<sup>†</sup>). The characteristic bands at  $1254$  and  $1263 \text{ cm}^{-1}$  are caused by the stretching vibration of the C–N bond. Compared with the TIPE ligand, the shifts of the band at  $1254 \text{ cm}^{-1}$  in **Ag-MOF1** and the one at  $1263 \text{ cm}^{-1}$  in **Ag-MOF2** caused by C–N bond vibration ensure the coordination between TIPE and Ag. The characteristic band at  $1602 \text{ cm}^{-1}$  is caused by the stretching vibration of the C=N bond.<sup>25</sup> In addition, the thermal stability of Ag-MOFs was estimated by thermogravimetric analysis (TGA). As shown in Fig. S4, ESI<sup>†</sup> the stability of **Ag-MOF1** can be maintained up to  $320 \text{ }^\circ\text{C}$  while **Ag-MOF2** can be stable up to  $450 \text{ }^\circ\text{C}$  under an  $\text{N}_2$  atmosphere.

### Electrochemical $\text{CO}_2\text{RR}$ performance

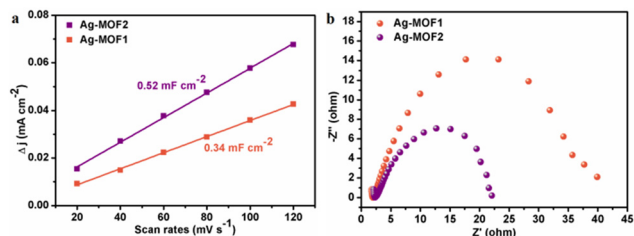
The  $\text{CO}_2\text{RR}$  behavior of the prepared samples was tested by using a standard three-electrode flow-cell system at room temperature in 1 M KOH. The experimental results presented in this work for all potential testing processes relative to the Ag/AgCl electrode are converted into reversible hydrogen electrode (RHE). As the voltage becomes more negative, the current gradually increases, which is mainly caused by two factors: one is the current generated by the hydrogen evolution reaction (HER), while the other is the current generated by electrocatalytic reduction of  $\text{CO}_2$ . In order to further test  $\text{CO}_2\text{RR}$  selectivity of the samples, we conducted FE testing. Gas chromatography analysis shows that CO and  $\text{H}_2$  are the gas products, while no liquid products were detected in the  $^1\text{H}$  NMR spectrum (Fig. S5, ESI<sup>†</sup>). From Fig. 3b, it can be seen that **Ag-MOF2** can reach the highest  $\text{FE}_{\text{CO}}$  of 92.21% at  $-0.98 \text{ V}$ , and it is worth noting that the selectivity of CO for **Ag-MOF2** is over 80% within a wide potential range, indicating that the reduction effect of electrocatalytic  $\text{CO}_2$  is superior to that of the HER. **Ag-MOF1** gives the maximum selectivity with 78.02%  $\text{FE}_{\text{CO}}$  at  $-1.08 \text{ V}$  (Fig. 3a). In order to more clearly compare the selectivity of CO for Ag-MOFs, the  $\text{FE}_{\text{CO}}$  of Ag-MOFs under different potentials was displayed. As illustrated in Fig. 3c, **Ag-**



**Fig. 3** (a) FEs of the HER and CO at different potentials for **Ag-MOF1**. (b) FEs of the HER and CO at different potentials for **Ag-MOF2**. (c) FEs of CO at different potentials for Ag-MOFs. (d) The partial current densities of CO for Ag-MOFs.

**MOF2** achieves better selectivity for CO than **Ag-MOF1** when covering the tested potential range. The partial current density for the products was studied to further illustrate the CO<sub>2</sub>RR performance of Ag-MOFs. Partial current density corresponding to CO ( $j_{CO}$ ) for **Ag-MOF2** reaches 29.51 mA cm<sup>-2</sup> at -0.98 V (Fig. 3d). As exhibited in Fig. 3d, the absolute value of  $j_{CO}$  for **Ag-MOF2** is larger than that of **Ag-MOF1** at all the measured potentials. The results indicate that the structures of Ag-MOFs have great impact on their electrocatalytic properties for CO<sub>2</sub> reduction and that the 2D-layered structure of **Ag-MOF2** is more beneficial for the CO<sub>2</sub>RR compared with the interpenetrated 3D framework of **Ag-MOF1**. Table S2, ESI† summarizes the performance comparison of the prepared Ag-MOFs with other catalysts used in the literature for CO<sub>2</sub> electroreduction to CO in terms of FE and current density. It can be seen that the FE<sub>CO</sub> of Ag-MOFs is higher than that of catalysts such as [Al<sub>2</sub>(OH)<sub>2</sub>(Co(tcpp))], PCN-222-Fe and {Ag<sub>49</sub>Mo<sub>16</sub>}.<sup>26–28</sup>

To uncover the origin of differences in the CO<sub>2</sub>RR performance between **Ag-MOF1** and **Ag-MOF2**, CO<sub>2</sub> adsorption and further electrochemical measurements were conducted. The CO<sub>2</sub> adsorption isotherms reveal that **Ag-MOF2** has higher CO<sub>2</sub> adsorption capacity than **Ag-MOF1** (Fig. S6, ESI†). The electrochemical double-layer capacitance ( $C_{dl}$ ) was determined by collecting cyclic voltammetry (CV) profiles in the non-faradaic region to investigate the electrochemical surface area (ECSA) of the catalysts. CV curves of each catalyst were recorded at different sweep speeds of 20, 40, 60, 80, 100 and 120 mV s<sup>-1</sup> in the non-pseudocapacitance region of -0.10–0 V vs. Hg/HgO (Fig. S7, ESI†). The current density difference,  $\Delta j$ , at a potential of -0.05 V vs. Hg/HgO was plotted with the corresponding sweep speed, and the  $k$  value is  $C_{dl}$ . The results show that  $C_{dl}$  increases in the order of **Ag-MOF1** < **Ag-MOF2** (Fig. 4a), which corresponds to the results of the electrochemical CO<sub>2</sub>RR, indicating that **Ag-MOF2** with its large ECSA possesses more

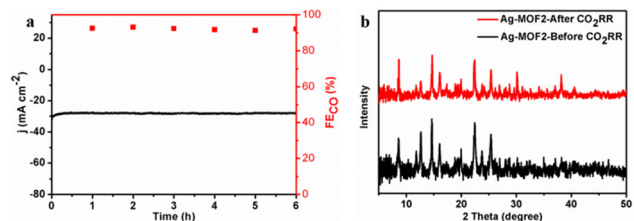


**Fig. 4** (a) Electrochemical double-layer capacitance plots for Ag-MOFs. (b) Nyquist plots of Ag-MOFs.

exposed active sites, benefiting its electrochemical performance for the CO<sub>2</sub>RR.<sup>29</sup> Electrochemical impedance spectroscopy (EIS) was used to estimate the rate of charge transfer and explore the kinetics of the CO<sub>2</sub> electroreduction reaction. The Nyquist plots show shapes of depressed semicircles corresponding to charge transfer reactance (Fig. 4b). The plot for **Ag-MOF2** exhibits a small curvature radius, proving that it has a fast charge transfer rate during the process of electrocatalysis. The low resistance of **Ag-MOF2** can speed up the transfer of electrons from the catalyst surface to the reaction intermediate, thus improving its catalytic activity and selectivity.<sup>30</sup>

#### Evaluation of the electrochemical and structural stability of Ag-MOFs

The stability and tolerance of catalysts are important indicators that reflect their suitability for practical utilization. Therefore, we conducted current–time experimental tests on **Ag-MOF2** at a potential of -0.98 V vs. RHE. As shown in Fig. 5a, the current density of **Ag-MOF2** exhibits almost no change after running the electrocatalytic test for 6 h, indicating that **Ag-MOF2** has good electrochemical stability in 1 M KOH. We conducted PXRD tests on the catalyst after the electrocatalytic test, as shown in Fig. 5b and Fig. S8, ESI†. The PXRD patterns of Ag-MOFs after the CO<sub>2</sub>RR test are basically consistent with the simulated ones, further proving that samples we prepared have electrochemical stability and tolerance, and can be used as stable and efficient CO<sub>2</sub> electroreduction catalysts. FT-IR spectra after the CO<sub>2</sub>RR also support the conclusion that Ag-MOFs have good stability (Fig. S9 and S10, ESI†), although there are tiny Ag nanoclusters after electrolysis



**Fig. 5** (a) Total current density and FE of CO over 6 h of electrolysis at -0.98 V on the **Ag-MOF2** catalyst. (b) PXRD patterns of **Ag-MOF2** before and after the CO<sub>2</sub>RR test.



reflected by transmission electron microscopy (TEM) measurements (Fig. S11, ESI†).

## Conclusions

In summary, a stable Ag<sub>4</sub>O cluster-based cubic **Ag-MOF1** and a 2D-layered **Ag-MOF2** have been synthesized and used for CO<sub>2</sub> electrocatalytic reduction. The selectivity to CO and H<sub>2</sub> of these Ag-MOFs is distinct. The results show that **Ag-MOF2** with high selectivity has a large electrochemically active surface area and fast electron transport rate, which may originate from its 2D-layered structure. In addition, the results of this work show that the N-enriched ligand with its  $\pi$ -conjugated system is unique due to its strong coordination ability with Ag(I) leading to Ag-MOFs with suitable stability for the CO<sub>2</sub>RR, which is important for designing new Ag-MOFs for the CO<sub>2</sub>RR.

## Conflicts of interest

There are no conflicts to declare.

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