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Visible light-mediated hydrogen atom transfer and proton transfer for the conversion of (2-vinylaryl) methanol derivatives to aryl aldehydes or aryl ketones†

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In this paper, we report a photochemical strategy for the visible light-mediated efficient conversion of (2-vinylaryl)methanol derivatives to the corresponding aryl aldehydes or aryl ketones in moderate to excellent yields with broad substrate scope under mild conditions. This photochemical process takes place from the generation of the triplet state of olefins and involves 1,5-hydrogen atom transfer, enol tautomerization, and a subsequent proton transfer process. The plausible reaction mechanism has been verified by deuterium labeling and control experiments, kinetic and Stern–Volmer analyses, and DFT calculations.

Introduction

Visible light-mediated chemical transformations have become a powerful tool in synthetic chemistry[1] and materials science[2] for the efficient construction of novel chemical skeletons under mild conditions in the last decade. On the basis of visible light-mediated reaction modes of single electron transfer (SET) and energy transfer (EnT), various robust organic transformations can be achieved under mild conditions, such as C–H bond functionalizations,[3] C–C and C–X (X = heteroatom) bond formations as well as their selective cleavage and the subsequent downstream functionalizations and so on.[4] The SET pathway relies heavily on a redox process between a photosensitizer and a reaction substrate or a reaction intermediate,[5] so the excited-state photosensitizer often performs the role of an oxidant or a reductant. For example, in transition metal-catalyzed reductive coupling reactions,[6] the photoredox strategy is often used in place of a stoichiometric metal reductant, and hence the reaction can be performed under more environmentally-benign conditions. It is worth noting that the EnT reaction mode is also an effective way to obtain highly reactive intermediates in the excited state (or the ground state) under mild conditions.[7] These highly reactive excited intermediates often undergo pivotal chemical transformations that are not possible with regard to their corresponding ground states.

Olefins as unsaturated hydrocarbons are remarkably valuable synthetic scaffolds in organic synthetic chemistry.[8] Excited olefins allow for many chemical transformations such as atom transfer reactions,[9] [2 + 2] cyclizations,[10] and $E \rightarrow Z$ isomerizations[11] that cannot be realized in the ground state. Nevertheless, the direct excitation of most olefinic compounds *via* visible light irradiation is disfavored, basically due to the comparatively low absorption at longer wavelengths. In general, harsh UV light irradiation is required for the direct excitation of styryl organic molecules, which leads to negative impacts on the selectivity, functional group tolerance and general applicability of the reaction, as well as competitive and uncontrollable side reactions.[12,13] For example, the conversion of (2-(1-phenylvinyl)phenyl)methanol to the corresponding 2-(1-phenylethyl)benzaldehyde under UV irradiation was reported by Hornback's group in 1979 (Scheme 1a).[14] As shown in Scheme 1a, it must be emphasized here that due to the harsh UV irradiation, the photochemical reaction product was only obtained in 6.3% yield in the absence of a photosensitizer, and even with the addition of xanthone as a photosensitizer, the corresponding benzaldehyde was obtained only in 15.1% yield. The low yield is probably due to the Norrish-type reaction[15] of 2-(1-phenylethyl)benzaldehyde under UV light irradiation, which leads to the decomposition of the obtained product. Therefore, it is conceivable that the use of the recently

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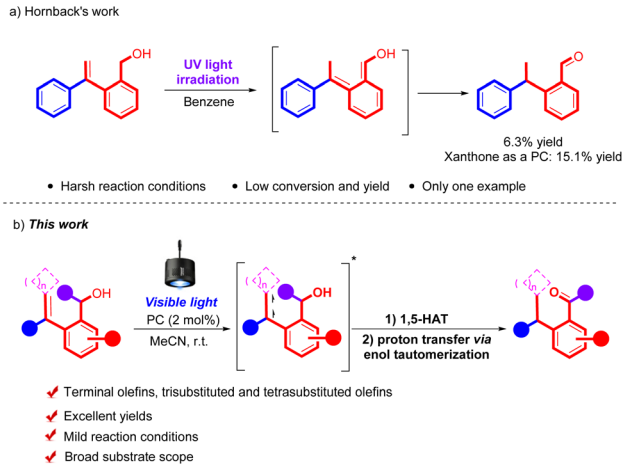


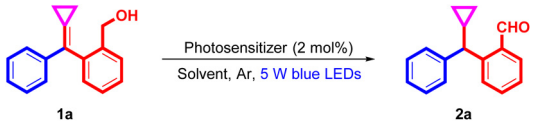
Scheme 1 Previous work and this work.

developed visible light-mediated strategy can avoid the further excitation of the formed aldehyde and realize an effective photochemical conversion of (2-vinylaryl)methanol derivatives to aryl aldehydes or aryl ketones.

Based on the above description and our previous work,[9c,d] we envisaged that the triplet state of (2-vinylaryl)methanol derivatives would be achieved through an EnT process upon visible light irradiation in corporation with a photocatalyst, in which the unprotected benzyloxy functional group could be used as a hydrogen atom transfer (HAT) partner and the subsequent enol interconversion isomerization as well as proton transfer would take place, resulting in an efficient conversion to aryl aldehydes or aryl ketones (Scheme 1b, this work). In this scenario, we anticipated that the desired product can be obtained in high yield along with broad substrate scope.

To test our working hypothesis, we started our investigation by using (2-(cyclopropylidene(phenyl)methyl)phenyl)methanol **1a** (0.2 mmol, 1.0 equiv.) as a substrate, Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6 (2 mol%) as a photosensitizer, and degassed MeCN (2.0 mL) as the solvent under irradiation with 5 W blue LEDs for 12 hours under an argon atmosphere. To our delight, the desired aromatic aldehyde product **2a** was obtained in 94% isolated yield (Table 1, entry 1). However, when using photosensitizers such as [Ir(2′,4′-dF-5-CF_3-ppy)$_2$(bpy)]PF_6 having a slightly lower triplet state energy (Table 1, entry 2) and Ir(dFppy)$_2$pic having a higher triplet state energy (Table 1, entry 3) as the photocatalysts, **2a** was obtained in lower NMR yields of 31% and 25%, respectively. The use of *fac*-Ir(ppy)$_3$, [Ru(phen)$_3$](PF_6)$_2$, and [Ru(bpy)$_3$](PF_6)$_2$ as the photocatalysts, which have lower triplet state energies, did not give **2a** at all (Table 1, entries 4–6). When xanthone was used as a photosensitizer, the expected product **2a** was not detected (Table 1, entry 7). This is probably due to that xanthone (λ_{max} = 340 nm) is not absorbed in the visible light region.[16] Notably, this photosensitizer was used by Hornback's group in the photochemical synthesis of benzaldehyde. The solvent effects on the reaction outcome were then examined, including dichloromethane (Table 1, entry 8), tetrahydrofuran (Table 1, entry 9),

Table 1 Optimization of the reaction conditions[a,b]



Entry	PC	E_T (kcal mol^{-1})	Solvent	Conc. [M]	Yield[b] (%)
1	Ir(dF-CF_3-ppy)2(dtbbpy)PF_6	60.8	MeCN	0.1	94[c]
2	[Ir(2′,4′-dF-5-CF_3-ppy)$_2$(bpy)]PF_6	60.4	MeCN	0.1	31
3	Ir(dFppy)$_2$pic	61.1	MeCN	0.1	25
4	*fac*-Ir(ppy)$_3$	57.8	MeCN	0.1	N.R
5	[Ru(phen)$_3$](PF_6)$_2$	48.4	MeCN	0.1	N.R
6	[Ru(bpy)$_3$](PF_6)$_2$	46.5	MeCN	0.1	N.R
7	Xanthone	74.3	MeCN	0.1	N.R
8	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		DCM	0.1	24
9	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		THF	0.1	29
10	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		Toluene	0.1	65
11	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		MeCN	0.05	85
12	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		MeCN	0.2	68
13	None		MeCN	0.1	N.R
14[d]	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		MeCN	0.1	N.R
15[e]	Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6		MeCN	0.1	90

[a] Reaction conditions: **1a** (0.1 mmol) and PC (2 mol%) were added in degassed solvent (2.0 mL) under an Ar atmosphere for 12.0 h, in a sealed tube under 5 W blue LED light irradiation. [b] Yields were determined by ^{1}H NMR using 1,3,5-trimethoxybenzene as an internal standard. [c] Isolated yield. [d] No light. [e] Under 100 W blue LEDs for 6 h.

and toluene (Table 1, entry 10), and they afforded **2a** in reduced yields compared to MeCN. The substrate concentration did not have a significant impact on the yield of **2a** as shown in entries 11 and 12 of Table 1. Control experiments elucidated that the target product was not obtained in the absence of a photosensitizer (Table 1, entry 13) or under dark conditions (Table 1, entry 14). When a higher power lamp was used as the light source, it speeded up the rate of the reaction, but did not increase the yield (Table 1, entry 15).

Having determined the optimized conditions, the substrate scope of this newly developed photochemical reaction was investigated with a wide variety of substrates **1** and we found that most of them provided the corresponding arylaldehydes **2** smoothly in good to excellent yields (Table 2). For terminal olefins, substrates **1** with various substituted aryl groups as R^1 moieties, regardless of whether they had electron-donating or electron-withdrawing groups, were tolerated in the protocol, delivering the target products **2b–2g** in moderate to excellent yields ranging from 65% to 96%. Substrate **1h**, in which R^1 was a thienyl group, also afforded the corresponding arylaldehyde **2h** in 83% yield. We also examined the effects of substituents at the aryl alcohol moiety on the reaction outcome, revealing that substrates **1i–1n** having 4-Cl, 5-Cl, 4-Br, 5-Br, 4-Me, and 5-Me substituted aryl alcohols were all compatible in this transformation, providing the desired products **2i–2n** in moderate to good yields ranging from 57% to 92% yields. For Br and methyl substituents, changing the position from 5-substitution to 4-substitution induced over a 30% increase in yield (**2k** and **2l**, **2m** and **2n**). Therefore, we hypothesized that in

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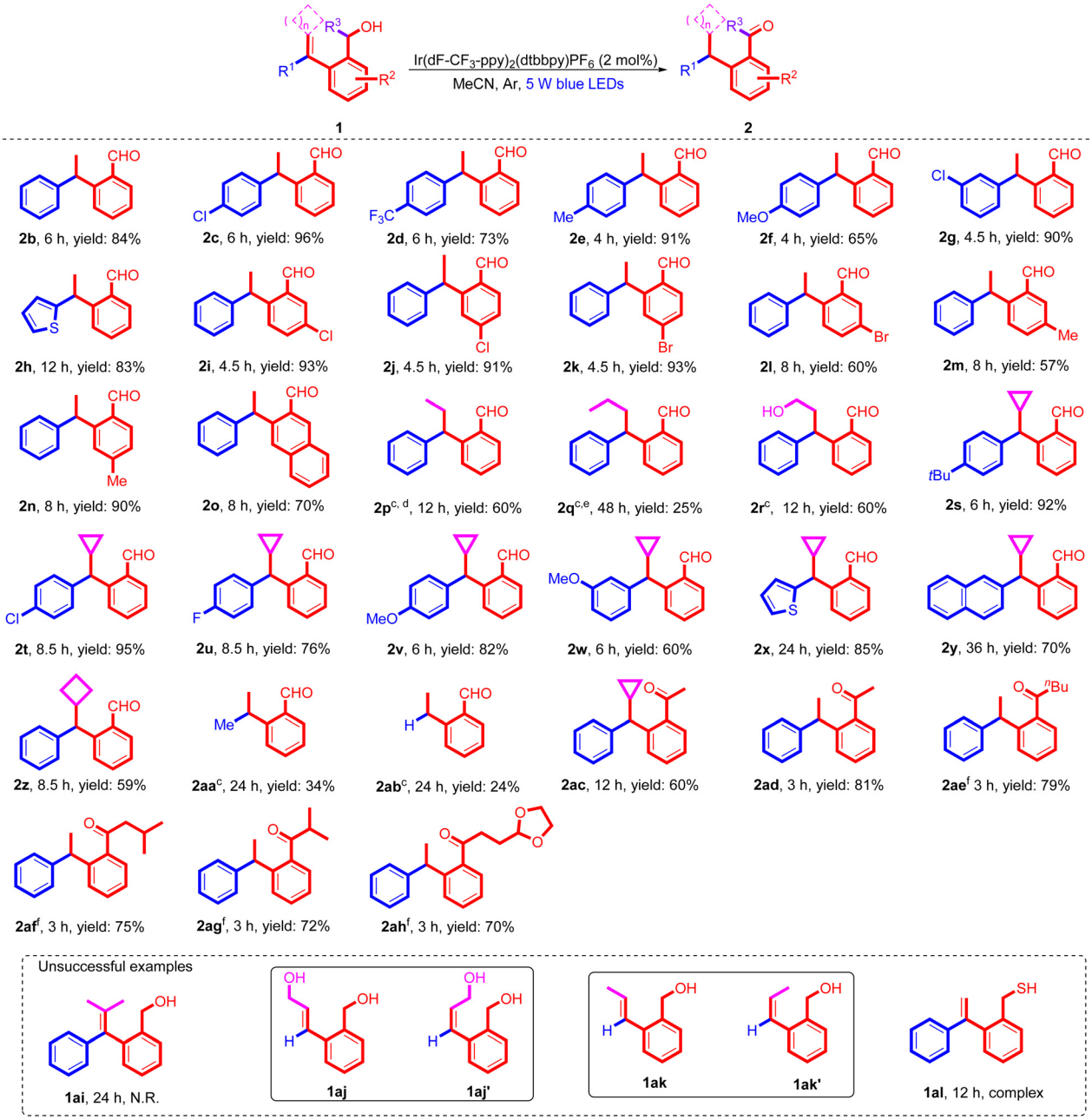
Table 2 Substrate scope of **1**[a,b]



[a] Reaction conditions: **1** (0.2 mmol) and **PC** (2 mol%) were added in solvent (2.0 mL) under an Ar atmosphere, in a sealed tube under 5 W blue LED light irradiation. [b] Yield of the isolated products. [c] Using 100 W blue light irradiation. [d] The *E*/*Z* ratio of substrate **1p** is 4 : 1. [e] The *E*/*Z* ratio of substrate **1q** is 6 : 4. [f] A 6 : 1 mixture of MeCN and H_2O was used as the solvent.

terms of electronic effects, the Br and methyl substituents at the 4-position probably make the corresponding intermediate **IV** (shown in Scheme 4) unstable, thus leading to a decrease in yields (**2l** and **2m**). Substrate **1o** bearing a naphthalenemethanol unit as a HAT partner was also smoothly converted to the corresponding aldehyde **2o** in 70% yield under the standard conditions. When using substrates **1p**, **1q** and **1r** containing a trisubstituted olefinic moiety in this photochemical reaction, the corresponding products **2p**–**2r** were also obtained in moderate yields under 100 W blue light irradiation. For substrates **1p**, **1q**, and **1r** containing trisubstituted olefinic molecules, the *E*/*Z* isomerization occurred as a side reaction; thus, 100 W blue LED light was used to ensure that the 1,5-HAT process as a critical step was able to proceed more efficiently to give the desired products in moderate yields. The desired product **2p** was obtained in 60% yield when using **1p** (*E*/*Z* = 4/1) as a substrate. However, it is worth noting that the *E*/*Z* ratio of substrate **1q** is 6/4, and the *Z*-trisubstituted olefin **1q′** cannot be

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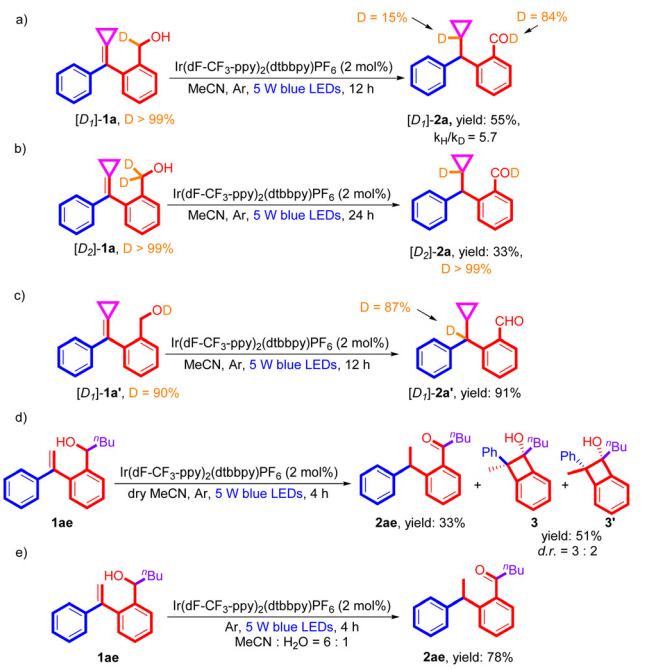
excited under the standard conditions due to its higher EnT energy,[11a] thus leading to a significant decrease in the yield of **2q** to 25% yield. Substrate **1r** with a free hydroxyl group was also tolerated in this protocol and the target product **2r** was obtained in 60% yield.

To further extend the scope of substrates **1**, a series of substrates bearing a methylenecyclopropane unit were used in this newly explored visible light-mediated photochemical reaction, and the corresponding products were obtained in good to excellent yields. For substrates **1s–1w**, in which R^1 was a substituted aryl group, the desired products **2s–2w** were also successfully obtained in good to high yields ranging from 60% to 95% regardless of whether the aryl group had an electron-withdrawing group or an electron-donating group. In addition, for substrates **1x** and **1y**, in which R^1 was a 2-thienyl group and a 2-naphthyl group, the reactions proceeded smoothly, delivering the desired products **2x** and **2y** in 85% and 70% yields, respectively, although a prolonged reaction time was required. This protocol also performed very well for substrate **1z** with a methylenecyclobutane unit, providing the desired aldehyde **2z** in 59% yield. Moreover, when substrates **1aa** (R^1 = Me) and **1ab** (R^1 = H) were used in the reaction, the corresponding aldehydes **2aa** and **2ab** were obtained in 34% and 24% yields, respectively, despite that the photochemical reactivity decreased dramatically. For the monoaryl substrates **1aa** and **1ab**, the excitation efficiency was lower than that of the diaryl substrates; thus, 100 W blue LED light is required.

Next, we turned our attention to examine the substrate scope of secondary alcohols. For substrates **1ac** and **1ad**, in which R^3 = Me, the reactions proceeded smoothly, furnishing the corresponding acetophenones **2ac** and **2ad** in moderate yields of 60% and 81%, respectively. When the methyl group was changed to a sterically hindered group such as an *n*-butyl, an isobutyl, or an isopropyl group or a 2-ethyl-1,3-dioxolane moiety, the desired products **2ae–2ah** were also obtained in moderate yields ranging from 70% to 79%. It is worth noting that in these cases, additional water was necessary to be added into the reaction system as a proton source to accelerate the proton transfer process efficiently.[17] Unfortunately, some limitations of this newly developed HAT and proton transfer reaction were observed for several substrates. For example, when non-cyclic tetrasubstituted olefin **1ai** was employed as a substrate, no reaction occurred under the standard conditions, probably due to the steric hindrance leading to the non-occurrence of the 1,5-HAT process. Furthermore, none of the desired products **2aj** and **2al** were detected for substrates **1aj** and **1ak** containing *E*-disubstituted olefins, and most of them isomerized to the corresponding *Z*-olefins **1aj′** and **1ak′** under the standard conditions.[11] For the monoaryl disubstituted olefins **1aj** and **1ak**, the desired reactions did not occur, probably due to that they were more difficult to excite, as well as the presence of *E*/*Z* isomerization. When the OH group was changed to the SH group (**1al**), the reaction system became complex.

To further investigate the reaction mechanism, a series of deuteration experiments and controlled experiments were carried out. First, the kinetic isotope effect of this photochemical reaction was determined using the mono-deuterated substrate [D_1]-**1a**, and $k_H/k_D \approx 5.7$ was obtained for the formation of [D_1]-**2a**, indicating that the C–H bond homolytic cleavage is involved in the rate-determining step (Scheme 2a) (see page S13 in the ESI†). Subsequently, the double-deuterated substrate [D_2]-**1a** was used in the reaction under the standard conditions, providing the target product [D_2]-**2a** in 33% yield along with >99% deuterium incorporation (Scheme 2b). These obtained results are consistent with our working assumption that the olefinic biradical species at the triplet state undergoes a 1,5-HAT process with the C(sp^3)–H bond at the benzylic alcohol position to give an open-shell biradical species. To clarify another benzylic proton source in the product, we carried out this reaction using substrate [D_1]-**1a′**, in which the alcoholic hydrogen atom was deuterated, giving the deuterated product in 91% yield along with 87% deuterium incorporation, suggesting that the benzylic proton of [D_1]-**2a′** indeed comes from the alcoholic hydrogen atom in the substrate (Scheme 2c). As shown in Scheme 2d, when the reaction was carried out using **1ae** as a substrate under the standard conditions in anhydrous MeCN, the reaction efficiency dropped down significantly, affording the desired product **2ae** in only 33% yield, and benzocyclobutane **3** was detected in 51% yield as a diastereomeric mixture, which was produced by [2 + 2] electrocyclization of the conjugated diene intermediate in the reaction.[9d] However, when the solvent was replaced with a mixture of MeCN and H_2O (6 : 1), the target product **1ae** was obtained in 78% yield as a sole product under otherwise identical conditions (Scheme 2e) (see Table 2 as well). This result suggests that the ambient H_2O in the reaction system is likely to accelerate the proton transfer process,[17] and as for sub-



Scheme 2 Deuterium labeling and control experiments.

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strates bearing sterically large groups, the reaction efficiency can be improved.

To further elucidate the interaction details of substrate **1** and the photocatalyst Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6, a fluorescence quenching experiment was performed using substrate **1a** with Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6 (Fig. 1a) (see page S16 in the ESI†). From Stern–Volmer analysis, it can be seen that the emission of Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6 was effectively quenched by substrate **1a**, suggesting that substrate **1a** may be photoexcited by the photosensitizer *via* an EnT pathway upon 5 W blue LED irradiation (Fig. 1b). The oxidation potential of substrate **1a** was determined by cyclic voltammetry as $E^{ox}_{p/2}$ = +1.75 V *vs.* SCE (Fig. 1c) (see page S17 in the ESI†), showing that it is impossible for substrate **1a** to undergo a SET process with the used photosensitizer (E_{red} (*Ir^{III}/Ir^{II}) = 1.21 V *vs.* SCE). For the terminal olefinic substrate **1l**, we also determined its oxidation potential as $E^{ox}_{p/2}$ = +1.35 V *vs.* SCE, suggesting that all these olefinic substrates undergo an EnT process under the standard conditions (see page S17 in the ESI†). Meanwhile, the light–dark interval experiments (Fig. 1d) and the quantum yield of 0.13 for this photochemical reaction (see page S18 in the ESI†) indicated that continuous visible light irradiation is an essential element in the product formation process and the radical chain process might not be involved in the turnover process.

DFT calculations were performed to gain insight into the reaction mechanism. All calculations were performed at the SMD(CH_3CN)/M06/def2tzvpp//M06/def2svp level with the Gaussian 16 program.[18] The solvation Gibbs free energy profile in acetonitrile for the suggested reaction pathway is shown in Scheme 3 (see page S174 in the ESI†). The first triplet-state of **1a** (**int-1a**) lies higher than the ground state of **1a** by 44.5 kcal mol^{-1}, and the Ir(dF-CF_3-ppy)$_2$(dtbbpy)PF_6 in the excited state (60.8 kcal mol^{-1}) can promote the formation of **int-1a** *via* an EnT pathway. Subsequently, **int-1a** undergoes 1,5-HAT *via* **ts-1a** to form the open-shell biradical intermediate **int-2a** (20.8 kcal mol^{-1}) with a reaction energy barrier of 11.8 kcal mol^{-1}. The intersystem crossing of **int-2a** gives the closed-shell conjugated diene intermediate **int-3a** (18.9 kcal mol^{-1}). Once **int-3a** is generated, the subsequent proton transfer takes place to generate the corresponding product **2a** (−17.2 kcal mol^{-1}) *via* **ts-2a** with a reaction barrier of 0.5 kcal mol^{-1}. We also examined the reaction barrier (**ts-3a**, $\Delta G^{\ddagger}$ = 10.4 kcal mol^{-1}) for **int-3a** to undergo electrocyclization to form benzocyclobutane **3a** (3.0 kcal mol^{-1}). The results show that the pathway to generate benzocyclobutane is unfavourable. For the reaction of the terminal olefin **1b**, we also performed DFT calculations; the calculation results are shown in Scheme 3b and they are similar to those of substrate **1a**, which undergoes the 1,5-HAT process and the subsequent proton transfer process. DFT calculations for the reaction of the substrate (**1ae**) having a secondary alcohol moiety were also performed (Scheme 3c). First, the energy transfer of the photosensitiser in the excited state occurs to obtain the

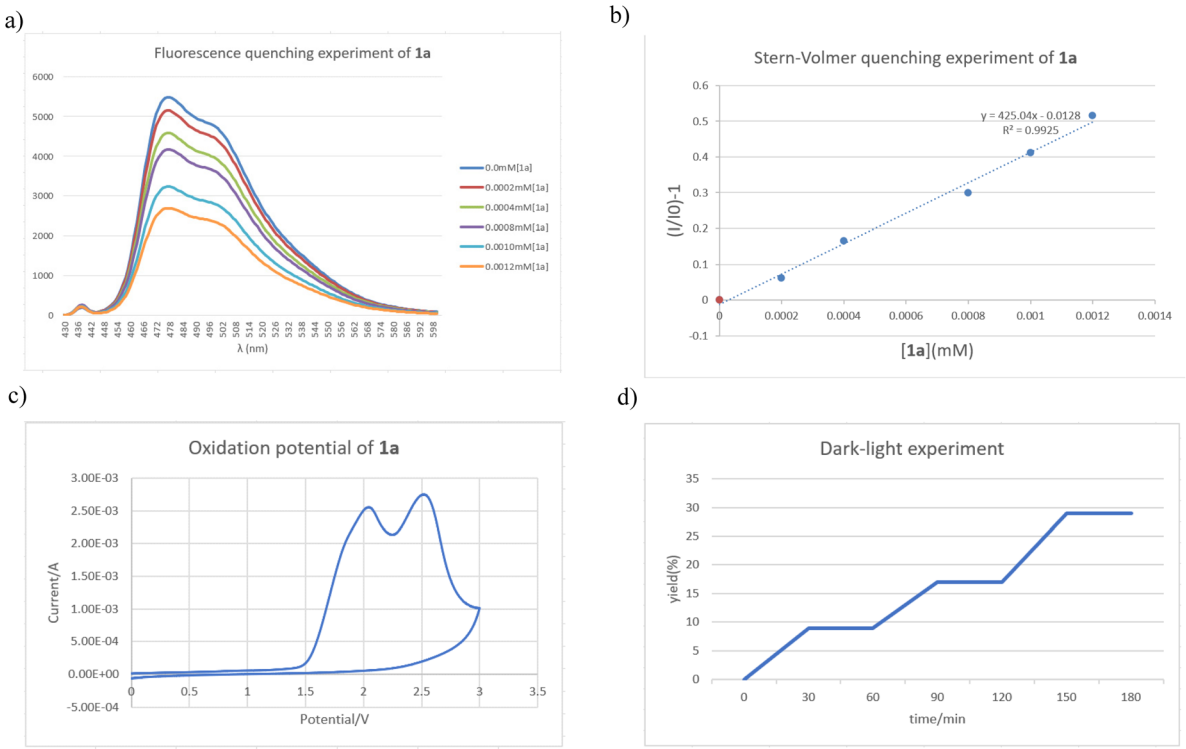

Fig. 1 Mechanistic studies. (a) An appropriate amount of **1a** was added to a 0.001 M photocatalyst solution in MeCN under an argon atmosphere and the emission from the sample was collected; (b) Stern–Volmer studies; (c) a solution of the substrate **1a** in MeCN (0.1 M) was tested with 0.1 M Bu_4NPF_6 as the supporting electrolyte. Scan rate = 0.1 V s^{-1}; (d) light on/off experiments.

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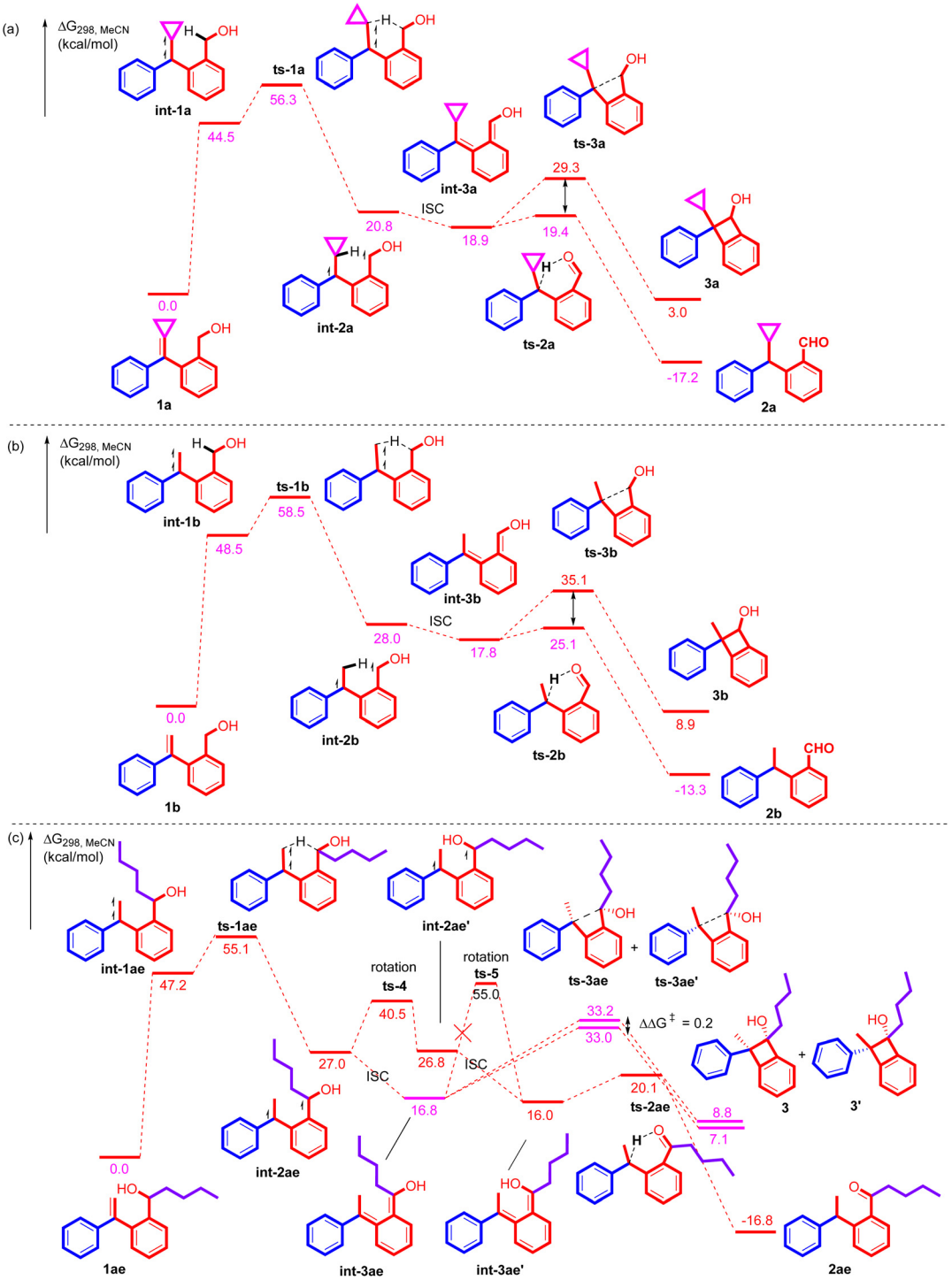

Scheme 3 (a) DFT calculations on the reaction of **1a**; (b) DFT calculations on the reaction of **1b**; (c) DFT calculations on the reaction of **1ae**.

excited-state intermediate **int-1ae**, and the energy of this state is 47.2 kcal mol^{-1} higher than that of **1a**. Subsequently, **int-1ae** undergoes 1,5-HAT across a low reaction energy barrier (**ts-1ae**, $\Delta G^{\ddagger}$ = 7.9 kcal mol^{-1}) to form the biradical intermediate **int-2ae** (27.0 kcal mol^{-1}). Considering that butyl is a sterically hindered group, we calculated the rotational energy barrier for the transformation of **int-2ae** into **int-2ae′**, which is 13.5 kcal mol^{-1} and indicates that both **int-2ae** and **int-2ae′** can exist in the reaction system. DFT calculations show that **int-3ae′** cannot be obtained through the direct rotation of **int-3ae** because the rotation barrier is 38.2 kcal mol^{-1} which is difficult to overcome under the standard reaction conditions. Finally, **int-2ae′** undergoes proton transfer *via* **ts-2ae** ($\Delta G^{\ddagger}$ = 4.1 kcal mol^{-1}) to produce the corresponding ketone **2ae**; the

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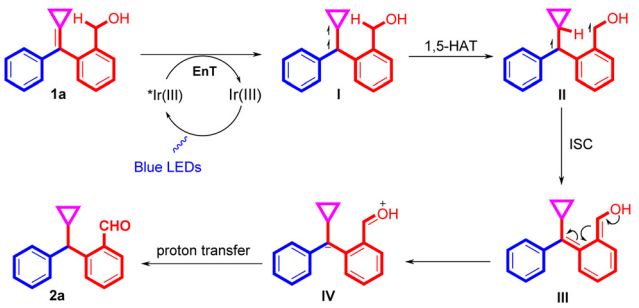
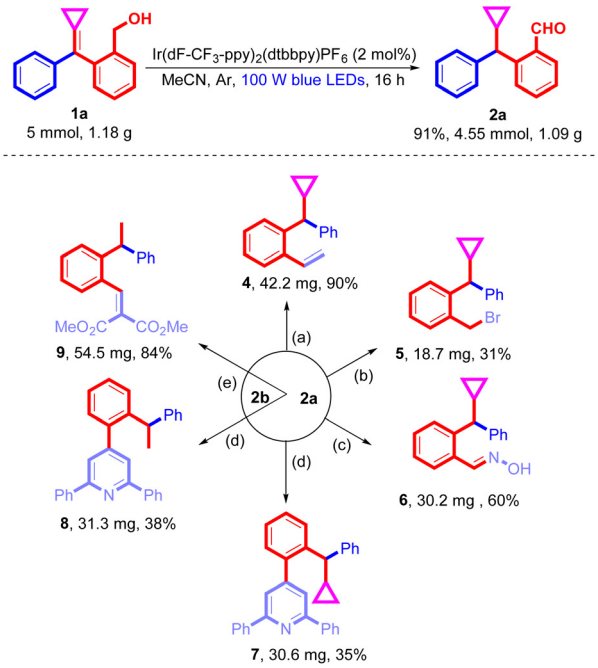
intermediate **int-2ae** undergoes electrocyclization *via* **ts-3ae** or **ts3–3ae′** to give a pair of diastereomers **3** and **3′**, ($\Delta\Delta G^{\ddagger}$ = 0.2 kcal mol^{-1}) which is basically in line with the dr value obtained in the experiment (dr = 3 : 2, see Scheme 2d).

Based on the previous literature[9] and our own experimental results above, a possible mechanism for this photochemical conversion is depicted in Scheme 4. First, the photosensitizer enters the excited state under visible light irradiation, which is quenched by substrate **1a** through an EnT pathway to revert to the ground state, and **1a** is converted to its triplet state as a biradical species **I**. Subsequently, the species **I** abstracts the hydrogen atom at the benzylic alcohol position *via* a 1,5-HAT process to produce the open-shell biradical intermediate **II**, followed by intersystem crossing to give the closed-shell conjugated diene intermediate **III**. Finally, intermediate **IV** is formed by an enol intercalative isomerization, and it undergoes a proton transfer process to give the corresponding product **2a**.

To demonstrate the synthetic utility of this newly explored photochemical reaction, a gram-scale reaction of **1a** was executed and this reaction worked well, affording **2a** in 91% yield (Scheme 5). Subsequently, **2a** and **2b** were subjected to a series of transformations under different reaction conditions. First, **2a** underwent a Witting reaction in THF with the *in situ* generated phosphonium ylide to produce the corresponding substituted styrene **4** in 90% yield. Considering that product **2a** contains an unactivated cyclopropyl group, an attempt was made to achieve ring-opening 1,3-hydroboration of cyclopropane in DCM at room temperature in the presence of phenylsilane and boron tribromide according to the previous literature;[19] however, the corresponding benzyl bromide **5** was produced in 31% yield without the formation of the cyclopropane ring-opened product under the standard conditions. Next, substituted benzaldoxime **6** was obtained in 60% yield upon treating **2a** with 1.25 equiv. of $NH_2OH{\cdot}HCl$ and 2.50 equiv. of CH_3COONa in $HCOOH/H_2O$ (6/4) at 80 °C. Moreover, according to the previous literature,[20] products **2a** and **2b** were smoothly transformed into trisubstituted pyridines **7** and **8** under the treatment with acetophenone and an ammonium salt at 130 °C in PhCl in the presence of 3 equiv. of DMSO and molecular oxygen (1.0 atm). In addition, the Knoevenagel reaction of product **2b** with dimethyl malonate provided the corresponding condensation product **9** in 84% yield (Scheme 5).



Scheme 4 A possible mechanism.



Scheme 5 Gram scale and product transformations. (a) **2a** (0.2 mmol), *t*-BuOK (0.40 mmol, 2.0 equiv., 1.0 mol L^{-1}), methyltriphenylphosphonium bromide (0.40 mmol, 2.0 equiv.), DCM, r.t.; (b) **2a** (0.20 mmol), phenylsilane (0.24 mmol, 1.2 equiv.), boron tribromide (0.24 mmol, 1.2 equiv., 0.24 mL of 1.0 M solution in DCM), DCM, r.t.; (c) **2a** (0.20 mmol), $NH_2OH{\cdot}HCl$ (0.24 mmol, 1.2 equiv.), CH_3COONa (0.24 mmol, 1.2 equiv.), $HCOOH/H_2O$ (2.0 mL, 3 : 2), 80 °C; (d) **2a** (0.20 mmol, 1.0 equiv.) or **2b** (0.20 mmol), phenylacetophenone (0.60 mmol, 3.0 equiv.), NH_4I (0.40 mmol, 2 equiv.), DMSO (0.60 mmol, 3.0 equiv.), molecular oxygen (1.0 atm), PhCl, 130 °C; (e) **2b** (0.20 mmol), dimethyl malonate (0.40 mmol, 2 equiv.) AcOH (10 mol%), piperazine (10 mol%), toluene, 100 °C.

Conclusions

In conclusion, we have explored a novel approach for the transformation of (2-vinylaryl)methanol derivatives into the corresponding aryl aldehydes or aryl ketones in moderate to excellent yields under mild conditions *via* visible light-mediated 1,5-HAT, enol interconversion and proton transfer processes with wide substrate scope. Deuterium labeling and kinetic analysis experiments verify that the HAT process takes place at the excited triplet state of olefins and the homolytic cleavage of the C–H bond is the rate-limiting step in this photochemical reaction. Meanwhile, another enol tautomerization-involved proton transfer process is also involved in this photochemical reaction. In addition, the control experiments have demonstrated that the presence of water in the reaction system can accelerate the proton transfer process and the possibility that the photoreaction may involve a radical chain process has also been ruled out by on and off light experiments and the determination of the quantum yield. This novel photochemical strategy further broadens the reaction patterns of visible light-mediated reactions of excited-state olefins. Further investigations based on the visible light-mediated conversion of

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triplet-state compounds and their applications in materials science and synthetic chemistry such as in natural product total synthesis and medicinal chemistry are currently underway.

Author contributions

Yan, J. and Wei, H. Z. contributed to the investigation. Yu, Z. Q., Wei, Y., and Yan, J. contributed to the calculations. Yan, J., Wei, Y., and Shi, M. contributed to conceptualization and writing – original draft.

Conflicts of interest

There are no conflicts to declare.

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