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Synthesis of a fully protected long-chain polyamine subunit of aculeine B using the photoremovable NPEC group†

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Polyamines are ubiquitously found in nature. In this paper, we disclose our iterative coupling strategy for the synthesis of a structurally defined polymer of 1,3-propanediamine, and the polymer can be used for the synthesis of both the initially proposed structure and the revised structure of protoaculeine B isolated from a marine sponge. We first attempted the synthesis of polyamines using “the Ns strategy” but found that a polyamine with eleven Ns groups has solubility problems. We then examined the versatility of the photoremovable NPEC protecting group in polyamine synthesis. Finally, the synthesis of a suitably protected 12-mer polyamine was achieved employing the NPEC group for the temporary protection of a terminal amino group.

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Introduction

Aculeine B (**1a** in Fig. 1A, ACU-B) is a natural product isolated by Sakai *et al.* in 2011¹ from a marine sponge *Axinyssa aculeata* collected inshore around Iriomote island, Japan. The structure of ACU-B consists of a 44-amino acid residue peptide, designated AcuPep, bound to a 12-, 13-, 14-, or 15-mer of 1,3-propanediamine (long-chain polyamine, LCPA)² via a modified tryptophan (W' , shown in red in Fig. 1A). The Trp-AcuPep peptide is thought to be post-translationally modified by structurally heterogeneous LCPAs by an enzymatic Pictet–Spengler reaction, in the biosynthesis of ACU-B.³ The sponge contains analogs of ACU-B with various amino acids/peptides conjugated to 12- to 15-mer LCPAs via W' .¹

The biological functions of ACUs were investigated mainly on aculeine A (ACU-A, structure not shown),¹ which has an additional LCPA chain conjugated to ACU-B. ACU-A was found to be neurotoxic and hemolytically active in erythrocytes. ACU-A showed membrane permeabilizing activity when exposed to 1-palmitoyl-2-oleoyl-*sn*-glycero-3-phosphocholine (POPC) liposomes with 10% cholesterol.¹ Furthermore, ACU-A was recently shown to be a naturally occurring cell penetrating peptide (CPP) that can internalize in the cells in an energy dependent manner.⁴

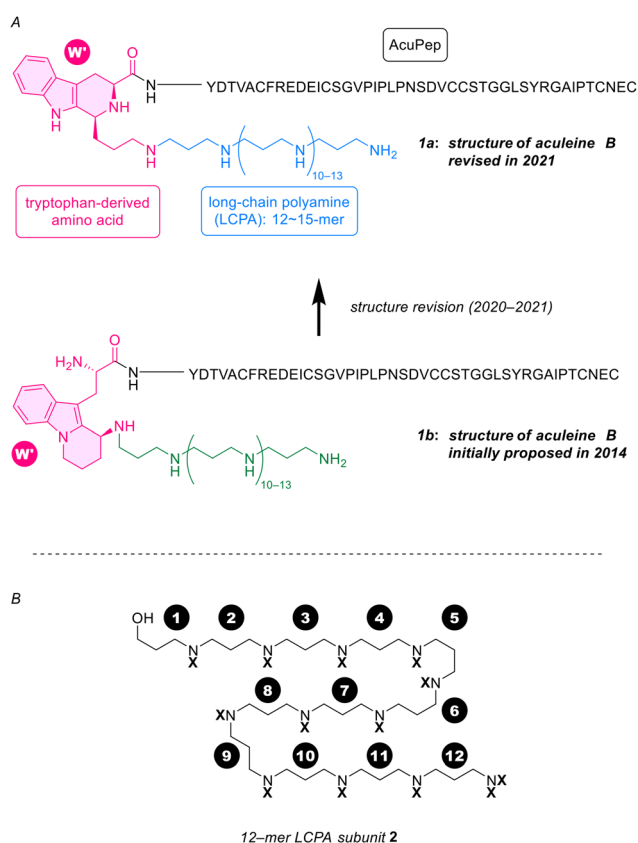


Fig. 1 (A) The structure of aculeine B (ACU-B) **1a**¹ isolated from *A. aculeata* and its initially proposed misassigned structure **1b**.⁹ (B) Suitably protected 12-mer LCPA subunit **2** for chemical syntheses of **1a** and **1b**.

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We then planned to study the biological properties of ACUs quantitatively using structurally less complex ACU-B. However, as noted above, the LCPA in the natural ACUs is a mixture of 12-mer to 15-mer of 1,3-propanediamine and hence is structurally heterogeneous. Furthermore, ACU-B is less abundant among the series of ACUs isolated from the sponge, and was obtained in only a limited amount. It was, therefore, necessary to establish a chemical synthesis method for obtaining large amounts of homogeneous ACUs (1) to quantitatively evaluate the biological properties and (2) to investigate the membrane permeabilizing and cell penetrating activities.

Here we report our practical synthesis of suitably protected, homogeneous polyamines.⁵ Although many attempts to chemically synthesize polyamines have been reported,⁶ synthesis of the LCPAs of ACUs is challenging because of their large molecular size. In this study, therefore, we *de novo* investigated the synthesis of a suitably protected 12-mer of 1,3-propanediamine. Our synthetic plan includes (1) employment of “the Ns strategy”⁷ and (2) the use of a 3-mer as the smallest structural unit and elaboration of a larger 6-mer and 12-mer by its iterative coupling.

As shown in Fig. 1A, the chemical structure of the modified tryptophan (W', red part) of ACUs underwent a revision in 2021.³ The 12-mer LCPA subunit 2 (Fig. 1B) targeted in this study was first designed as an alcohol to directly serve the synthesis of the misassigned structure 1b (see Fig. 1A, the green part)⁸ initially proposed in 2014.^{8,9} However, the alcohol 2 was expected to also be utilized in the chemical synthesis of the revised structure 1a, the part shown in blue. Indeed, we have recently succeeded in the total synthesis of the revised LCPA-W' part³ named protoaculeine B using 2 (see Scheme 11 for the structure).¹⁰ Furthermore, our preliminary studies using homogeneous synthetic LCPA prepared in this study have recently shown that the above activities characteristic of ACU-A require the LCPA moiety in addition to the AcuPep motif (unpublished). This paper reports our efforts to develop and synthesize the LCPA subunit 2 of ACUs in the following order.

(1) Our unfruitful attempt at the synthesis of the first-generation 12-mer LCPA subunit 2a (see Scheme 3) with eleven Ns groups by the Ns strategy.

(2) Revisiting synthesis strategies, and the evaluation of the suitability of the photoremovable NPEC protecting group in the modified Ns strategy.⁵

(3) Achievement of the synthesis of a suitably protected, second-generation 12-mer LCPA subunit 2b (see Scheme 10), with favourable physical properties owing to fewer Ns groups (four).^{9b}

Results and discussion

Synthesis of the first-generation 12-mer LCPA subunit 2a protected with eleven Ns groups

For the synthesis of the 12-mer LCPA subunit 2, we decided to adopt a strategy starting from the smallest 3-mer structural

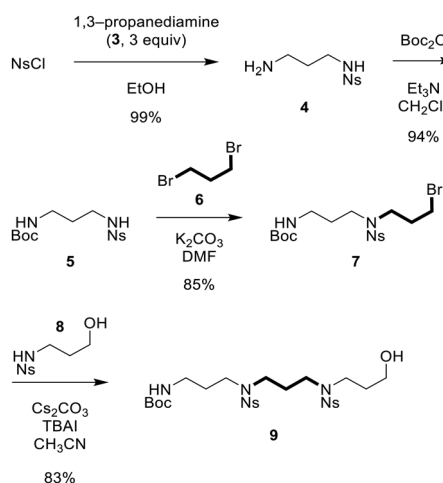
unit and first doubling the size to construct a 6-mer polyamine. A similar process was planned to further extend the 6-mer polyamine to a 12-mer polyamine. In such an iterative strategy, the choice of the protecting group seemed important, and we decided to use the Boc group and the Ns group as persistent protecting groups for amino groups.

We first attempted the synthesis of a suitably protected 12-mer polyamine 2a (see Scheme 3 for the structure) by a well-established method called the Ns strategy.⁷ Scheme 1 shows the synthesis of a 3-mer triamine 9 which is the smallest unit for iterative synthesis of the 12-mer LCPA 2a. The synthesis was performed according to the procedure reported by Fukuyama and Kan.^{7d} Selective mono-Ns-protection of 1,3-propanediamine (3) gave 4 quantitatively (99% yield). *N*-Boc protection of 4, followed by coupling with 1,3-dibromopropane (6), provided 2-mer diamine 7 in 80% in 2 steps. After treatment of the bromide 7 with *N*-Ns-3-aminopropanol (8) prepared from 3-aminopropanol by *N*-Ns protection,^{7d} the key 3-mer alcohol 9 was obtained in good yield (83%).

From the 3-mer triamine unit 9, a 6-mer polyamine 14 was synthesized as shown in Scheme 2. First, the hydroxy group was converted to iodide 11 *via* mesylate 10 in 93% yield over 2 steps. NMR data of 11 were consistent with those reported.^{7d} On the other hand, a two-step protecting group manipulation (SOCl₂, MeOH; NsCl, Et₃N) was performed on 9 to obtain 3-mer alcohol 13 with three Ns groups *via* free amine 12 in 94% yield.

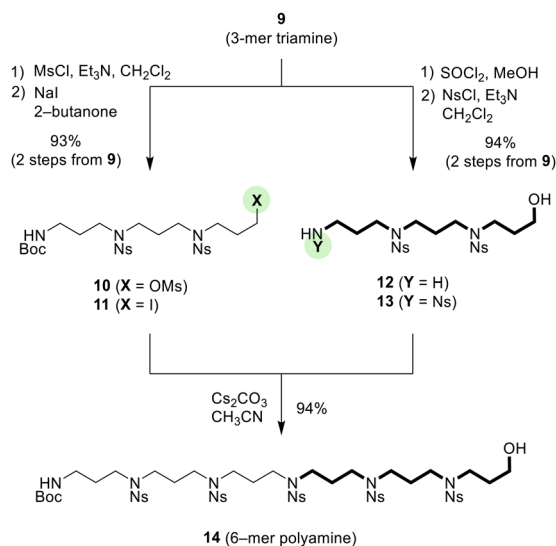
With two 3-mer triamines 11 and 13 in hand, we then attempted to construct a 6-mer polyamine bearing five Ns groups by coupling them. Gratifyingly, it was found that the coupling induced by Cs₂CO₃ proceeded smoothly in CH₃CN to give 14 in excellent yield (94%).

From the 6-mer polyamine 14, we further attempted to synthesize a 12-mer polyamine. Here, polyamine 2a (Scheme 3) of twice the size of 14 was synthesized following the procedure for synthesizing the 6-mer polyamine 14 from the 3-mer triamine 9 shown in Scheme 2. That is, as shown in Scheme 3, the hydroxy group of 14 was converted to iodide 16 in 83%

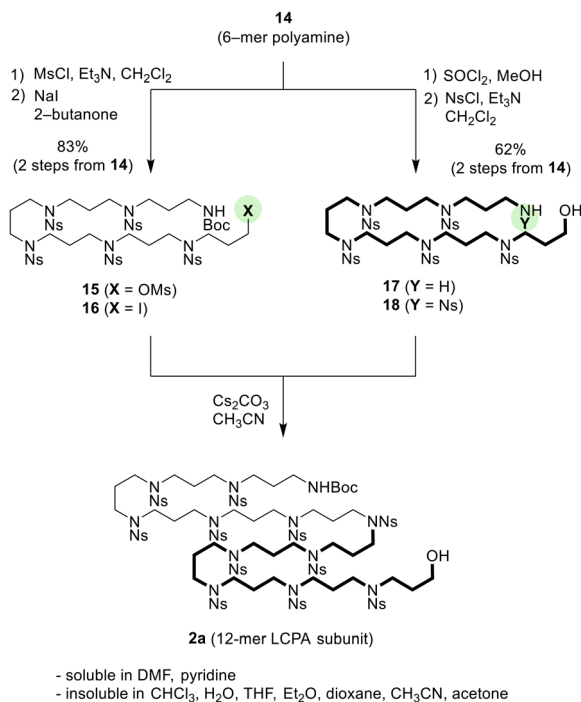


Scheme 1 Preparation of the smallest 3-mer triamine unit 9.^{7d}





Scheme 2 Synthesis of 6-mer polyamine **14** with five Ns protecting groups.



Scheme 3 Synthesis of the first-generation 12-mer LCPC subunit **2a** with eleven Ns protecting groups.

yield (2 steps) by mesylation followed by iodination. On the other hand, the *N*-Boc group was replaced with an *N*-Ns group over two-step manipulation to furnish hexa-*N*-Ns-polyamine **18** in moderate yield (62%, 2 steps). The two 6-mer polyamines (**16** and **18**) were then subjected to a coupling reaction induced by Cs₂CO₃ to give 12-mer polyamine **2a**. Unfortunately, **2a** was found to be insoluble in common organic solvents except DMF and pyridine and could not be purified. The poor solubility of

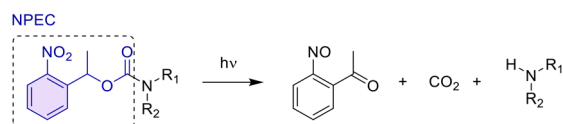
2a may be due to the many (eleven) Ns groups that cause hydrophobic and/or π - π stacking interactions. Therefore, we decided to consider another 12-mer LCPC subunit with a different pattern of *N*-protecting groups.

Revisiting synthesis strategies

Since the insolubility of the initially synthesized 12-mer LCPC subunit **2a** was thought to be due to the interactions between *N*-Ns groups, we decided to synthesize a 12-mer LCPC subunit with fewer Ns groups. To make this possible, the number of *N*-Boc and *N*-Ns groups in the 3-mer triamine **9** must be changed. We planned to increase the number of *N*-Boc groups to “two”, and reduce the number of *N*-Ns groups to “one” in the newly designed 3-mer triamine unit (see **36** in Scheme 8).

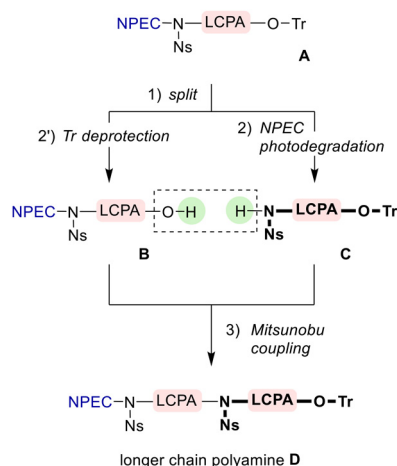
To achieve the synthesis, it was necessary to employ another temporary group orthogonal to the Boc and Ns groups, for protection of a terminal amino group. Although *N*-Alloc and *N*-Cbz groups had been used orthogonally,^{7c} in the present study we decided to examine the applicability of a photoremovable group in polyamine synthesis. Among the photoremovable protecting groups for amines, we focused on the 1-(2-nitrophenyl)ethoxy carbonyl (NPEC) group¹¹ (Scheme 4) used for photocaged compounds. The NPEC group belongs to the nitrophenylethyl (NPE)-type caging groups¹² and was developed by Walker *et al.* in 1989¹³ as a more reactive group than the popular *o*-nitrobenzyl group.¹⁴ The NPE-type caging groups had been used in biochemistry to control *in vitro* biological events.¹⁵ However, there were only a few instances in which the NPEC group was used in organic synthesis.¹⁶ One reason for this may be that the NPEC group contains a stereogenic center. We expected that the issue of the NPEC group will not be a major problem in the synthesis of achiral polyamines; rather, the photoremovable nature was expected to bring efficiency to polyamine synthesis.

Scheme 5 shows one cycle of the iterative process planned for the synthesis of polyamines, wherein the NPEC group is used to temporarily protect a terminal amino group. That is, starting with an oligomer of 1,3-propanediamine **A** bearing NPEC-NNs- and TrO- groups at both ends, the Tr group would be selectively removed under acidic conditions to give alcohol **B**. On the other hand, *N*-Ns amine **C** is expected to be synthesized by photoirradiating **A** to remove the NPEC group. A Mitsunobu reaction¹⁷ of **B** and **C** is expected to give polyamine **D** twice the size of **A**. Since the product **D** has NPEC-NNs- and TrO- groups at both ends, a polyamine twice the size of **D** would be further synthesized by repeating the same operation as **A** \rightarrow **D** for **D**. Thus, the number of Ns groups was expected to be reduced in this synthetic strategy.



Scheme 4 Mode of photodegradation of *N*-NPEC amine.¹²



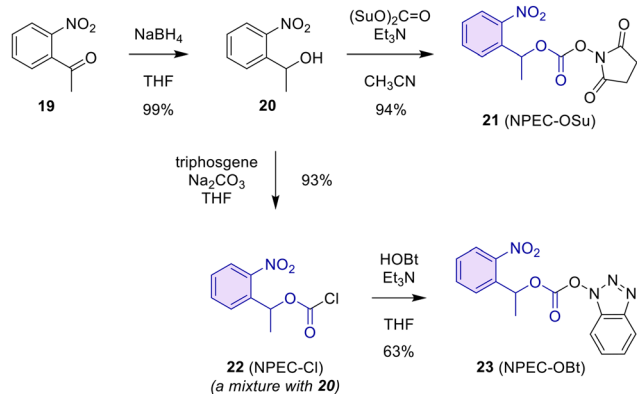


Scheme 5 Our plan for polyamine synthesis using the modified Ns strategy. "LCPA" denotes the protected oligomer of 1,3-propanediamine.

Evaluation of the photoremovable NPEC protecting group

Preparation of three NPEC reagents, NPEC-OSu **21**, NPEC-Cl **22**, and NPEC-OBt **23**, is shown in Scheme 6. The syntheses were carried out starting from commercially available 2'-nitroacetophenone (**19**).¹⁸ Reduction of **19** afforded alcohol **20** in 99% yield. Treatment of **20** with $(\text{SuO})_2\text{C}=\text{O}$ and Et_3N gave the chemically stable NPEC-OSu **21** in good yield (94%) after silica-gel column chromatography. Alcohol **20** was also converted to chemically rather unstable NPEC-Cl **22** upon exposure to triphosgene and Na_2CO_3 . From NPEC-Cl **22**, NPEC-OBt **23** (HOBt, Et_3N)¹⁹ was elaborated as another stable NPEC reagent in 63% yield.

With three NPEC reagents thus obtained, we synthesized NPEC-NHNs **24**. As shown in Table 1, succinimidyl ester **21** was treated with NsNH_2 , Et_3N , and DMAP to furnish NPEC-NHNs **24** in 69% yield (entry 1). When chloroformate **22** was used, the reaction gave a better yield (85%, run 2). With NPEC-OBt **23**, the reaction was sluggish and the yield for **24** was only 13% (entry 3). Therefore, NPEC-Cl **22** was decided to be employed for the preparation of NPEC-NHNs reagent **24**



Scheme 6 Preparation of three NPEC reagents NPEC-OSu **21**, NPEC-Cl **22**, and NPEC-OBt **23**.

Table 1 Preparation of NPEC-NHNs **24**

Entry	NPEC-X	Yield of 24
1	NPEC-OSu (21)	69%
2	NPEC-Cl (22)	85%
3	NPEC-OBt (23)	13%

that is used for the introduction of the NPEC-NNs- group (see below) in the present study.

Reactivity of NPEC-NHNs **24** was then evaluated in the (A) coupling reaction with an alkyl bromide or (B) Mitsunobu reaction with an alcohol (Table 2).^{7c} The former coupling reaction (procedure A) was examined with two bromides using K_2CO_3 and Bu_4NI in DMF at 80 °C. As shown in entry 1, heptylamine **25** was readily synthesized from NPEC-NHNs **24** and heptyl bromide in 77% yield. With more reactive benzyl bromide, benzylamine **26** was obtained in higher yield (98%, entry 2).

Amine synthesis by the Mitsunobu reaction of NPEC-NHNs **24** with alcohols using DEAD and PPh_3 (procedure B in Table 2) was next examined in benzene at rt. From 1-octanol, *N*-NPEC-*N*-Ns-octylamine (**27**) was provided in good yield (92%) (entry 3). The product **27** was later used as a model compound for the photodegradation reaction (see Table 3 and the text). Naphthalenylethylamine **28** was smoothly provided in 95% yield by the Mitsunobu reaction of **24** with naphthalene-2-ethanol (entry 4). Even a secondary alcohol can react in the Mitsunobu reaction; alanine ester **29** was constructed from ethyl lactate in good yield (85%, entry 5).

Table 2 Synthesis of amines by *N*-alkylation of NPEC-NHNs **24**

Entry	RBr or ROH	Conditions	Yield
1	$\text{Br}-\text{C}_7\text{H}_{15}$	A ^a	77% (25)
2	$\text{Br}-\text{C}_6\text{H}_5$	A ^a	98% (26)
3	$\text{HO}-\text{C}_8\text{H}_{15}$	B ^b	92% (27)
4	$\text{HO}-\text{C}_8\text{H}_7$	B ^b	95% (28)
5	$\text{HO}-\text{C}_3\text{H}_5\text{CO}_2\text{Et}$	B ^b	85% (29) <i>dr</i> = 57 : 43

^a Conditions A: RBr (1.5 equiv.), K_2CO_3 (5 equiv.), Bu_4NI (0.2 equiv.), DMF, 80 °C. ^b Conditions B: ROH (2 equiv.), DEAD (2 equiv.), PPh_3 (2 equiv.), PhH, rt.



Table 3 Photodegradation of the NPEC group of *N*-NPEC-*N*-Ns-octylamine (**27**)

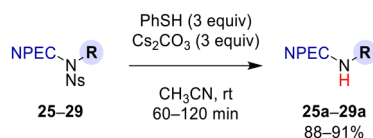
Entry ^a	Irradiation source	Wavelength	Solvent	Time	Yield ^b (NMR/isolation)
1	UV lamp (3 W)	365 nm	MeOH	1.5 h	36%/N.D.
2	UV lamp (3 W)	365 nm	MeOH	9 h	85%/N.D.
3	UV lamp (3 W)	365 nm	TFE	9 h	100%/92%
4	UV lamp (3 W)	365 nm	HFIP	9 h	100%/94%
5	LED lamp (3 W)	365 nm	MeOH	20 min	100%/95%
6	LED lamp (1.2 W)	395 nm	MeOH	9 h	3.7%/N.D.
7	LED lamp (1.2 W)	395 nm	TFE	9 h	6.1%/N.D.
8	LED lamp (1.2 W)	395 nm	HFIP	9 h	8.2%/N.D.
9	Blue LED (0.5 W)	ca. 460 nm	MeOH	9 h	6.5%/N.D.
10	High-pressure Hg lamp (435 W)	—	MeOH	20 min	100%/91%
11	High-pressure Hg lamp (435 W)	—	MeOH	60 min	100%/94%
12	High-pressure Hg lamp (435 W)	—	TFE	20 min	100%/N.D.
13	High-pressure Hg lamp (435 W)	—	HFIP	20 min	100%/N.D.

^a Milligram quantities of *N*-NPEC amine **27** were used for the photodegradation. ^b N.D.: not determined.

For efficient synthesis of primary and secondary amines, Boc-NHNs^{7c,20} and Cbz-NHNs^{7c,21} have been generally used in the Ns strategy.⁷ From the good reactivity of NPEC-NHNs **24** shown in Table 2, **24** is expected to serve as a new entry for this purpose.

With *N*-NPEC-*N*-Ns amines **25–29**, we studied the orthogonal reactivity of NPEC and Ns protecting groups under their deprotection conditions. Selective deprotection of the *N*-Ns group was first examined (Scheme 7, see the ESI for further details†). Treatment of **25–29** with PhSH and Cs₂CO₃ in CH₃CN, the standard reagents developed by Fukuyama and Kan,^{7c} afforded des-*N*-Ns amines **25a–29a** in good yields (88–91%). No side reaction was observed, and all reactions were completed at rt within 120 min.

We next screened selective deprotection conditions for the NPEC group using *N*-NPEC-*N*-Ns-octylamine (**27**, Table 3) as a representative substrate. The photodegradations were performed using various irradiation sources at rt in alcohol solvents. As shown in entry 1, irradiation of **27** in MeOH, with a UV lamp (365 nm), induced selective NPEC deprotection in 36% conversion (NMR yield) after 1.5 h. When the reaction

**Scheme 7** Selective deprotection of the Ns group in *N*-NPEC-*N*-Ns amines **25–29**. For structures of R groups, see Table 2. For further details, see the ESI†.

time was extended to 9 h, the conversion yield was 85% (entry 2). When the solvent for the reaction was changed to 2,2,2-trifluoroethanol (TFE, entry 3) or 1,1,1,3,3,3-hexafluoro-2-propanol (HFIP, entry 4), the reaction proceeded to completion and the isolated yield was satisfactory (92% for TFE, 94% for HFIP). Thus, TFE and HFIP were found to be the solvents that increase the conversion efficiency in the NPEC photodegradation. Interestingly, irradiation with an LED at the same wavelength (365 nm) completed the deprotection in a much shorter reaction time (20 min, entry 5). As shown in entries 6–8, however, irradiation with a longer wavelength LED (395 nm) was less effective in the three solvents (MeOH, TFE, and HFIP). A blue LED did not induce deprotection as well (entry 9).

Then irradiation with a high-pressure Hg lamp was examined. As shown in entry 10, the reaction smoothly proceeded in MeOH in 20 min and the product **27b** was obtained in 91% isolated yield. Even under prolonged irradiation conditions (60 min), the reaction was reproducible, and furthermore, no decomposition of the *N*-Ns amine product **27b** was observed (entry 11). With TFE (entry 12) or HFIP (entry 13) as a solvent, photoremoval of the NPEC group proceeded also smoothly in 20 min as expected from entries 3 and 4.

From these irradiation studies, we concluded that reactions under an LED (365 nm) or a high-pressure Hg lamp are sufficient to remove the NPEC group.

To examine the versatility of NPEC deprotection by LED (365 nm) irradiation, further reactions with other substrates (**25**, **26**, **28**, **29**) were also conducted (Table 4). Here, irradiation was carried out in MeOH at rt. Under these conditions, *N*-NPEC-*N*-Ns-heptylamine (**25**, entry 1) and two other substrates **26** (entry 2) and **29** (entry 4) wherein R = benzyl and R = ethyl lactate, respectively, were smoothly photodegraded in 20 min as in the case for **27** (Table 3, entry 5). The only exception was *N*-NPEC-*N*-Ns-naphthylamine (**28**, Table 4, entry 3), which took 40 min to complete the reaction. The substrate **28** has a light-absorbing naphthyl group, which is assumed to reduce irradiation efficiency.

Table 4 Photodegradation of the NPEC group in *N*-NPEC-*N*-Ns amines **25**, **26**, **28**, and **29** with an LED (365 nm, 3 W)^{a,b}

Entry	Substrate	Time	Yield (NMR)
1	25 (R = heptyl)	20 min	100% (25b)
2	26 (R = benzyl)	20 min	100% (26b)
3	28 (R = naphthylethyl)	2 × 20 min	63% (20 min), 100% (40 min) (28b)
4	29 (R = ethyl lactate, dr 57 : 43)	20 min	100% (29b)

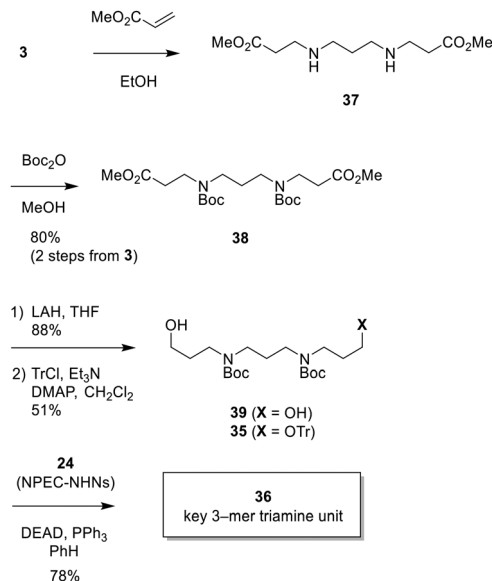
^a For structures of R groups, see Table 2. ^b Milligram quantities of *N*-NPEC amines **25**, **26**, **28**, and **29** were used for the photodegradation.



Achievement of the synthesis of a suitably protected, second-generation 12-mer LCPA subunit **2b**

With the *N*-NPEC method in hand, synthesis of the suitably protected 12-mer LCPA subunit **2** (see Fig. 1B) was retried. Our attempt to prepare the second-generation key 3-mer unit **36** is shown in Scheme 8. Di-*N*-Ns protection of 1,3-propanediamine (**3**) yielded **30** in 84% yield. Mono-*N*-alkylation of **30** with bromide **31**, prepared from commercially available 3-bromopropanol by tritylation,²² provided 2-mer **32** in good yield (94%) which, in turn, was subjected to the second *N*-alkylation with 3-bromopropanol (67% yield). Two Ns groups of **33** were then replaced with Boc groups, using the Fukuyama protocol (PhSH, Cs₂CO₃, CH₃CN)^{7c} followed by *N*-Boc protection (Boc₂O, Et₃N, CH₂Cl₂). However, the yields of these two steps were quite low (18%) due to the hydrophilic properties of intermediate **34**. Synthesis of **36** by this pathway was deemed impractical.

Therefore, we decided to adopt the route of introducing the Boc groups in the beginning of the synthesis, not by replacing Ns groups. Ultimately, we were able to synthesize the 12-mer LCPA subunit **2b** by the new route (see below). Scheme 9 shows our successful synthesis of the fully protected key 3-mer structural unit **36**. First, 1,3-propanediamine (**3**) was reacted with 2.1 equiv. of methyl acrylate to give **37**,^{6a} which was then treated with Boc₂O to furnish bis-*N*-Boc amine **38** in 80% yield in 2 steps. After reduction of two ester groups with LAH (88%), the symmetric diol **39** was monoprotected as Tr ether to provide **35** (51%) and the unreacted **39** (38%) was recovered. Finally, NPEC-NHNs **24** was introduced into alcohol **35** by a

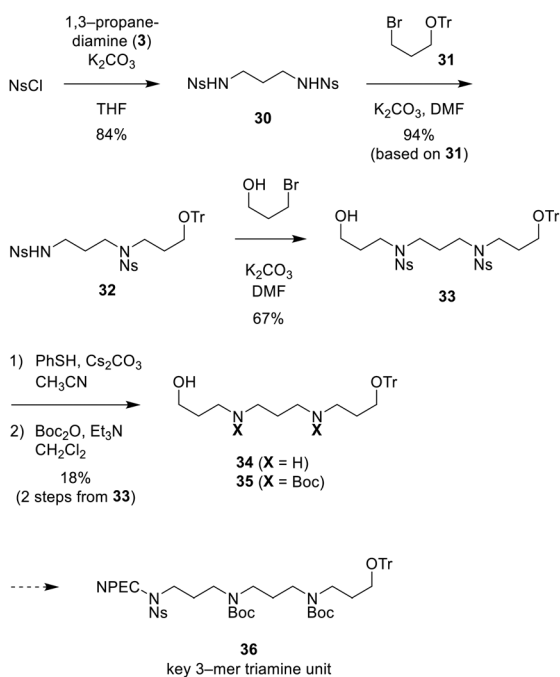


Scheme 9 Successful preparation of the second-generation 3-mer triamine unit **36** protected with the NPEC group.

Mitsunobu reaction (see Table 2) to obtain the suitably protected key 3-mer unit **36** in moderate yield (78%).

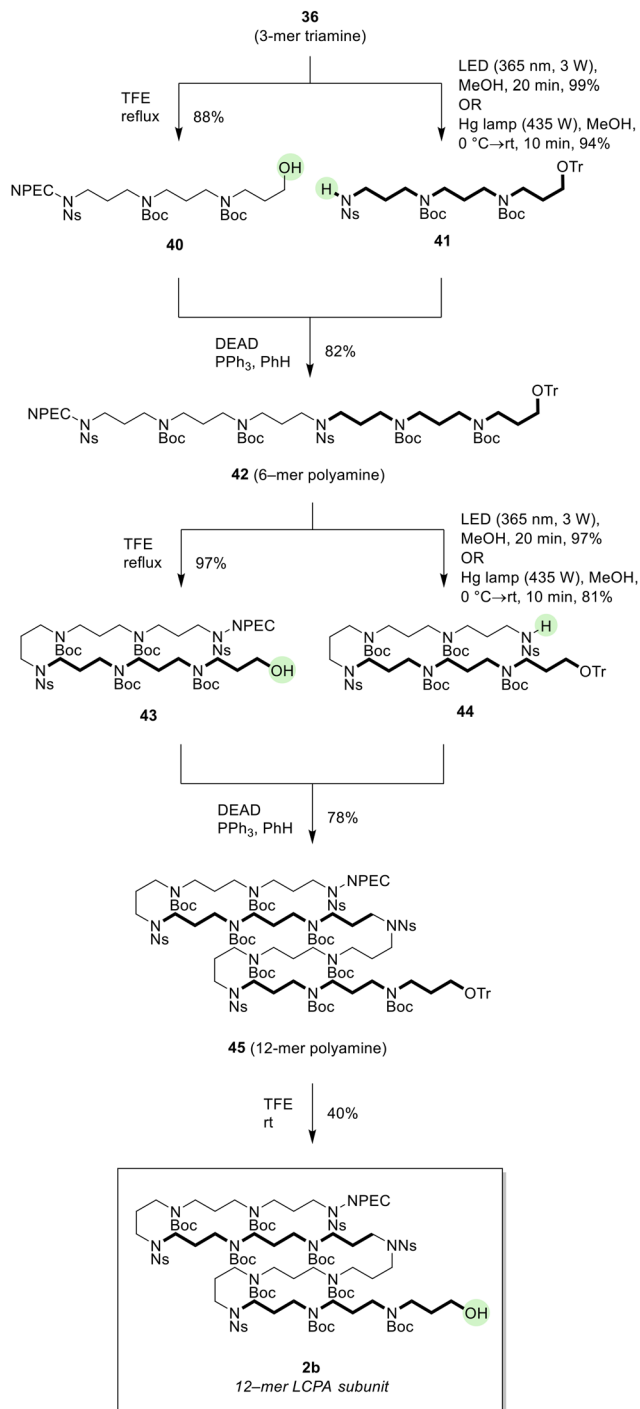
Repeated splitting and coupling operations converted the key 3-mer triamine unit **36** to the 12-mer LCPA subunit **2b** (Scheme 10). Removal of the Tr group of **36** by TFE under reflux conditions gave alcohol **40** (88%).^{9b} On the other hand, the NPEC group of **36** was removed by photoirradiation with LED (365 nm) quantitatively. It should be noted that the photodegradation can also be performed with a high-pressure Hg lamp in 94% yield. Mitsunobu coupling of alcohol **40** and the *N*-Ns amine **41** thus obtained proceeded smoothly, yielding 6-mer polyamine **42** in reasonable yield (82%). From **42**, the same deprotections provided alcohol **43** and *N*-Ns amine **44**, which were then coupled by a Mitsunobu reaction (DEAD, PPh₃, benzene) to furnish 12-mer polyamine **45** (78% yield). Finally, the Tr group of **45** was removed under mild conditions (TFE, rt) to accomplish the synthesis of the suitably protected 12-mer LCPA subunit alcohol **2b** in 40% yield. Unreacted Tr ether **45** (56%) was recovered intact. From 1,3-propanediamine (**3**), **2b** was thus synthesized in 6.9% total yield in 10 steps. The subunit **2b** contained only four Ns groups, fewer than **2a** (eleven Ns groups). As expected, 12-mer LCPA **2b** was found to be soluble in most common organic solvents.

The 12-mer LCPA subunit **2b** was later converted to the structure **46b** initially proposed as protoaculeine B (pACU-B, LCPA-W') corresponding also to a terminal substructure of ACU-B (Scheme 11).^{8,9b} However, a comparison of synthetic **46b** with the natural product in terms of chemical reactivity and spectroscopic profiles unfortunately revealed that the initially proposed structure of pACU-B needed revision.⁸ Then **2b** was further used in the synthetic study of the revised structure **46a** of pACU-B. Recently, we have successfully completed the synthesis of **46a**, and determined the



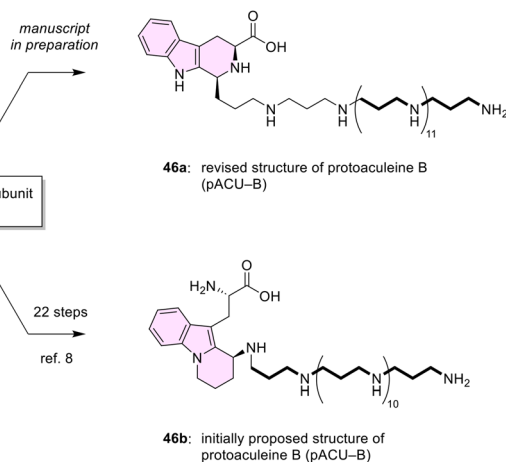
Scheme 8 An attempt to prepare the second-generation 3-mer triamine unit **36** protected with the NPEC group.





Scheme 10 Iterative synthesis of the second-generation 12-mer LCPA subunit **2b**.

true structure of pACU-B as **46a** (manuscript in preparation).¹⁰ These studies were made possible by the establishment of the synthesis of the protected polyamines reported here. In the future, we plan to study the structure–activity relationships of ACUs with structurally homogeneous LCPAs. Such studies are expected to reveal how ACUs, with their unique polyamine–peptide conjugate structures, show



Scheme 11 Syntheses of the revised structure **46a**¹⁰ and the initially proposed structure **46b**⁸ of protoaculeine B (pACU-B) from the LCPA subunit **2b**. The work established the true structure of pACU-B from *Axinyssa aculeata* as **46a**.¹⁰ Bold lines denote the part derived from the 12-mer LCPA subunit **2b**.

diverse activities such as membrane permeabilizing¹ and cell penetrating activities.⁴

Conclusions

In this study, we investigated the use of Boc and Ns groups for persistent protection of amino groups to efficiently synthesize suitably protected and structurally defined polymers of 1,3-propanediamine. To achieve this goal, we studied the use of an additional photoremovable NPEC group as a temporary protecting group, whose reactivity is orthogonal to that of Boc and Ns groups. Finally, a method was successfully developed to synthesize 6-mer LCPA **42** by dimerization and 12-mer LCPA **45** by tetramerization, using the 3-mer triamine **36** as the smallest structural unit. Polyamines are important structural components used in a variety of applications.²³ We believe that the method developed in this study will be useful for supplying structurally defined, diverse, and homogeneous polyamines for a variety of biochemical studies.

Author contributions

M. O. conceptualized the research. M. M. and R. W. performed the experiments and collected the data. R. I. and M. O. analysed the data and wrote the manuscript.

Conflicts of interest

There are no conflicts to declare.



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