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Cu₂CdGe(S,Se)₄†

Optimizing kesterite solar cells from Cu₂ZnSnS₄ to

Robert B. Wexler, Da Gopalakrishnan Sai Gautam D‡ and Emily A. Carter **D**ab

Kesterite solar cells, based on the prototypical absorber material Cu₂ZnSnS₄ (CZTS), are cheap, nontoxic, and chemically stable, thus rendering them promising, beyond-Si photovoltaic technologies. Their efficiencies, however, are limited by the formation of defects that decrease the short-circuit current by creating deep traps where nonradiative recombination of photoexcited charge carriers occurs via the Shockley-Read-Hall mechanism. To suppress the formation of these defects, specifically the most deleterious 2Cuzn + Snzn antisite cluster, we devised an ion substitution strategy involving complete Cdand Ge-substitution and partial selenization, ultimately arriving at the optimal composition, Cu₂CdGeS₃Se (CCdGSSe). Using density functional theory and ab initio thermodynamics, we predict that complete Cd- and Ge-substitution leads to a 125% increase in the formation energy of the deep-trapinducing $2Cu_{Cd} + Ge_{Cd}$. Additionally, 25% selenization optimizes the predicted band gap (1.43–1.47 eV, as calculated from a hybrid functional) with respect to the Shockley-Queisser limit. In addition to providing a practical and novel ion substitution strategy, we also elucidate the mechanisms of defect suppression and promotion by Ge and Se, highlighting the key role of the inert pair effect and metalchalcogen bond covalency, respectively. Due to its ideal thermodynamic and electronic characteristics, CCdGSSe should reinvigorate research on kesterite-based solar cells, optimizing the rich materials space afforded by ion substitution and post-quinary compositions.

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Introduction

The potential impact of solar energy is unmistakable as it far exceeds global energy needs while simultaneously combating global warming due to its sustainability and carbon-neutrality. To capitalize on this opportunity, the scientific community has spent many decades searching for materials that efficiently convert sunlight to electricity. A number of solar-cell

technologies have been commercialized, most notably those based on Si;2-5 thin-film technologies such as CdTe,6-8 Cu(In,Ga) Se2,9-11 GaAs,12-14 and hybrid organic-inorganic halide perovskites;15-17 conductive organic polymers;18-20 and molecular dyes.21-23 Although eco-friendly printing techniques have been used to fabricate inexpensive, nontoxic organic/polymer solar cells, 24,25 solar cells with higher efficiency made of promising inorganic materials are as yet unable to supplant fossil-fuel power stations for a number of reasons, including because they contain expensive (Te, In, and Ga) and/or toxic (Cd, Pb, and As) elements or, in the case of perovskite solar cells, have poor resistance to moisture.26-28 Consequently, solar-energy conversion efficiency, cost, toxicity, ease of production, and stability all must be considered concurrently to design an efficient, scalable, and environmentally friendly solar infrastructure. Over the last decade, there has been growing interest in the solar absorber material Cu₂ZnSnS₄ (CZTS), which contains inexpensive and nontoxic elements, possesses ideal sunlight absorption characteristics (namely, a band gap of $\approx 1.39-1.52$ eV (ref. 29)), and is stable, even in the presence of moisture.30-33 Unfortunately, under processing conditions (i.e., annealing at 600 K (ref. 34 and 35)), defects can form (e.g., antisites, vacancies, and their clusters), which leads to solar-cell inefficiencies. 36,37

Two important parameters for optimizing kesterite solar cells are the band gap ($E_{\rm g}$) and the concentration of Shockley-Read-Hall (SRH) recombination centers^{38,39} ($x_{\rm SRH}$), as illustrated

^aDepartment of Mechanical and Aerospace Engineering, Princeton University, Princeton, NJ 08544-5263, USA. E-mail: eac@princeton.edu

^bOffice of the Chancellor and Department of Chemical and Biomolecular Engineering, University of California, Los Angeles, CA 90095, USA. E-mail: eac@ucla.edu

[†] Electronic supplementary information (ESI) available: Further details about the effect of Mg- and Si-substitution on CZTS; stability of CCdGSSe against phase separation and decomposition; lowest energy ionic configurations for bulk structures of Ge-substituted CZTS and Cd- and Se-substituted CZGS; polymorph stability for quaternary chalcogenides containing Zn; defect generation protocol; ICSD bulk structures; stability of individual point defects vs. clusters; chemical potential limits for CZGSSe and CCdGSSe; lowest energy defect configurations of Ge-substituted CZTS and Cd- and Se-substituted CZGS; effect of Cd- and Se-substitution on the concentration of recombination centers in CZGS; influence of anion order in CCdGSSe on Cu_{Cd} + Cd_{Cu} formation energies; intrinsic stability/instability of CTZS/CZGS-based compounds; experimental band gaps of Cu₂CdGeS₄ and Cu₂CdGeSe₄ and predicted defect formation energies for Cu₂CdGeS₃Se and Cu₂CdGeSe₂; and formation energy of anion vacancies in CCdGSSe. See DOI: 10.1039/d0ta11603c

[‡] Current address: Department of Materials Engineering, Indian Institute of Science, Bengaluru, Karnataka 560012, India.

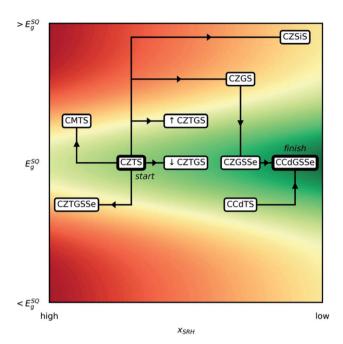


Fig. 1 Schematic road map of substitution strategies considered, from Cu₂ZnSnS₄ (CZTS) to Cu₂CdGeS₃Se (CCdGSSe), where C is Cu¹⁺, Z is Zn^{2+} , M is Mg^{2+} , T is Sn^{4+} , G is Ge^{4+} , \downarrow CZTGS is $Cu_2ZnSn_{0.875}Ge_{0.125}S_{4+}$ and \uparrow CZTGS is Cu₂ZnSn_{0.5}Ge_{0.5}S₄. The horizontal and vertical axes correspond, respectively, to the concentration of Shockley-Read-Hall recombination centers $(x_{SRH})^{38,39}$ and the band gap (E_q) , where E_q^{SQ} is the band gap that maximizes the Shockley-Queisser (SQ) limit.^{42,43} Red and green contours indicate regions of lower and higher solar cell efficiency, respectively.

in Fig. 1. The band gap of the absorber, which depends on the polymorph ($E_g^{\text{Stannite}} < E_g^{\text{Kesterite}}$, typically^{40,41}), is proportional to the open-circuit voltage (V_{oc}) of the solar cell whereas x_{SRH} is a good measure of the short-circuit current (I_{sc}) , with lower concentrations corresponding to higher currents. Given that efficiency is proportional to the product of V_{oc} and I_{sc} , it is clear that increasing I_{sc} by decreasing x_{SRH} increases solar cell efficiency (from high to low x_{SRH} , contours become greener in Fig. 1). The $V_{\rm oc}$, however, cannot be increased without bound as the Shockley–Queisser limit⁴² dictates an optimal E_g of 1.34 eV,⁴³ with quasi-exponential decline in efficiency upon deviations away from that value.

While in principle CZTS can host a wide range of neutral and charged defects including antisites, vacancies, and their clusters, the effects of the neutral defects listed in Table 1 on the physics and performance of kesterite solar cells are both significant and well-understood. We label defects using a simplified Kröger-Vink notation M_S (in Table 1 and thereafter), where M corresponds to the point defect species, which can be an atom (e.g., Cu) or a vacancy (V), and S indicates the lattice site that the species occupies. For Cu_X + X_{Cu}, Cu⁺ and X²⁺ cations swap sites. This leads to the formation of shallow donor (X_{Cu}) and acceptor (Cu_x) levels within the band gap, resulting in spatial fluctuations of the conduction and valence band edges. These fluctuations reduce the effective band gap, thus lowering the V_{oc} of the material. $^{36,44-49}$ For Cu vacancies (V_{Cu}), a Cu⁺ site is unoccupied, leaving the crystal Cu-deficient. Previous quantum mechanics simulations predicted that Cu vacancies reduce potential fluctuations along the valence band edge, which can mitigate any band gap reduction caused by Cu_x + X_{Cu} clusters. 36 Consequently, V_{Cu} tends to mitigate any reduction in V_{oc} , consistent with the Cu-poor synthesis conditions that typically are employed to achieve highest efficiency.34,50-52 Finally, for 2Cu_X + Y_X, three X²⁺ are replaced by two Cu⁺ and one Y⁴⁺, leaving the crystal X²⁺-deficient and Cu⁺/Y⁴⁺-rich. Previous quantum mechanics studies indicate that $2Cu_{Zn} + Sn_{Zn}$ generates localized trap states near the center of the band gap, which promote SRH recombination and, therefore, reduce the I_{sc} of CZTS. 41,53,54 We note that interface recombination also can reduce the $V_{\rm oc}$, 55,56 however, the goal of the present work was to suppress the formation of bulk defects that are well-known, via both experiment and theory, to be detrimental to kesterite solar cell efficiency.

Density functional theory (DFT) and thermodynamic analyses have been employed to improve fundamental understanding of how to control defect formation and to identify promising doping schemes to limit the formation of detrimental defects and improve solar cell performance. For example, theory^{57,58} and experiment⁵⁹⁻⁶¹ both conclude that Agcontaining phases are more ordered in the 1+ and Zn2+ sublattices (i.e., lower concentration of $Cu_X + X_{Cu}$ defects) due to the anisotropic expansion of the unit cell upon replacement of Cu by Ag. 59 Specifically, Ag2ZnSnSe4 affords a higher predicted maximum photovoltaic efficiency than Cu₂ZnSnS₄,⁵⁸ consistent with experiments. 62-70 The alkali metals are another promising group of isovalent replacements for Cu,62-73 with theory indicating that <25% Na-doping in CZTS suppresses the formation of $Cu_{Zn} + Zn_{Cu}$. In addition to the Cu^+ site, several studies establish isovalent doping on the Zn2+ site as a promising strategy to improve the performance of CZTS-based solar cells as well. Among the 2+ cations considered, which includes the alkaline earth72,74,75 and transition metals,57,76-78 Cd has been identified, by both theory57 and experiment,77 as one of the most effective at reducing 1+/2+ and 2+/4+ cation disorder. Finally, on the 4+ site, Ge stands out as an exceptional candidate for

Table 1 Defects considered and their effects on the physics of kesterite solar cells and solar cell parameters. X is Zn, Cd, or Mg; Y is Sn, Ge, or Si; $E_{\rm q}$ is the band gap; $V_{\rm oc}$ is the open-circuit voltage; and $I_{\rm sc}$ is the short-circuit current

Defect	Effect on the physics of kesterite solar cells	Primary effect on solar cell parameters
$\begin{aligned} &Cu_X + X_{Cu} \\ &V_{Cu} \\ &2Cu_X + Y_X \end{aligned}$	Causes electrostatic potential fluctuations Mitigates potential fluctuations along valence band edge Causes Shockley-Read-Hall recombination	Decreases $V_{ m oc}$ Increases $V_{ m oc}$ Decreases $I_{ m sc}$

substitution, as numerous studies, mostly experimental, have shown that the combination of partial Ge- and Se-alloying, where the latter serves primarily to optimize the band gap, leads to significant increases in both the $V_{\rm oc}$ and $I_{\rm sc}$ and thus solar-cell efficiency. However, there remains scarce mechanistic understanding of these dopants' influence on defect formation and consequently uncertainty regarding promising directions for improved materials design.

To address this need, here we present predictions of bulk stability, band gap, and formation energies for the key defects in Table 1 and analyze the trends to develop not only a deeper understanding of the quantum mechanical effects that govern defect formation but also a practical strategy for overcoming the efficiency stagnation of kesterite solar-cell technologies over the last six years. 79 To provide a clear picture of the ion substitution strategies we adopted, we first outline the path from CZTS to our newly proposed quinary chalcogenide, Cu₂CdGeS₃Se (CCdGSSe) in Fig. 1. Since CZTS offers a nearly optimal band gap (\approx 1.5 eV), 80 one would like post-CZTS absorbers to horizontally traverse our schematic road map (Fig. 1). First, we find that Mgand Si-substitution for Zn and Sn, respectively, increases E_g and, for Mg-substitution, increases x_{SRH} as well, which will lead to inefficient solar cells. Therefore, we do not consider Mg and Si further in the main text (see Tables S1-S3 in Section S1 of the ESI†). Second, we find that complete Ge-substitution (CZGS) decreases x_{SRH} but increases E_{g} . While partial Ge-substitution (CZTGS) increases E_g by a lesser extent, we find that this approach is not as promising as partial selenization (CZGSSe). Third, inspired by recent studies on Cu₂CdSnS₄ (CCdTS),^{57,77} we find that complete Cd-substitution, along with Ge-substitution coupled with partial selenization, provides the optimal band gap and significantly decreases x_{SRH} , thus identifying CCdGSSe as a promising candidate for improving the efficiencies of kesterite-based solar cells. By discussing the steps in our roadmap, we reveal both the independent and concerted effects of Ge, Se, and Cd on defect thermodynamics and electronic structure that leads to the optimization of kesterite solar cells, resulting in CCdGSSe.

Computational methods

DFT calculations

We calculated polymorph relative energies, defect formation energies, and band gaps using spin-polarized DFT as implemented in the Vienna *Ab initio* Simulation Package (VASP), ⁸¹⁻⁸³ which employs the projector augmented-wave (PAW) method. ^{84,85} We used a 520 eV kinetic energy cutoff for the plane wave basis set and initialized atomic magnetic moments of 0.6 μ_B in a ferromagnetic configuration so as not to preclude open shell configurations induced by uncompensated neutral defects, *e.g.*, V_{Cu}, which generates a hole that oxidizes Cu⁺, Zn²⁺, and/or Sn⁴⁺. That being said, we find that all final configurations are nonmagnetic. For the calculation of bulk thermodynamic quantities such as polymorph relative energies and defect formation energies, we used both the strongly constrained and appropriately normed (SCAN) meta-generalized-gradient approximation (meta-GGA)⁸⁶ and the Heyd–Scuseria-

Ernzerhof (HSE) hybrid 87-90 exchange-correlation functionals, where the latter was used to evaluate our most promising candidate, CCdGSSe. For band gap calculations, we used both DFT-HSE and the Perdew-Burke-Ernzerhof (PBE)91 GGA augmented by a Hubbard U term (PBE+U) where U was derived from unrestricted Hartree-Fock (UHF) calculations 92,93 of binary transition metal (TM) oxide clusters.94 For PBE+U calculations, we adopted the rotationally invariant approach introduced by Dudarev et al. 95 where U = 3.6 eV, 4.5 eV, and 4.8 eV were applied on Cu 3d, Zn 3d, and Sn 4d/Ge 3d, respectively. 94 U = 4.8 eV was not derived specifically for Ge 3d but the charge state of Sn and Ge should be the same and therefore we used the value derived for Sn here. For Cd-containing compounds, we used DFT-HSE instead of PBE+U despite its significant computational expense since DFT-HSE yields better agreement with experimental band gaps,96-102 which was needed to evaluate accurately CCdGSSe.

We used the PBE PAW data sets recommended by the Materials Project¹⁰³ for Cu (4s¹ 5d¹⁰), 2+ cations [Zn (4s² 3d¹⁰), $Mg (2p^6 3s^2)$, and $Cd (5s^2 4d^{10})$], 4+ cations $[Sn (5s^2 4d^{10} 5p^2)$, Ge $(4s^2 3d^{10} 4p^2)$, and Si $(3s^2 3p^2)$], and 2- anions [S $(3s^2 3p^4)$ and Se (4s² 4p⁴)]. Additionally, we applied 0.05 eV of Gaussian electronic smearing to accelerate self-consistent field (SCF) convergence and sampled 32 Γ -point-centered k-points per \mathring{A}^{-1} to integrate over the Brillouin zone (symmetry was not used to reduce the number of k-points). Fast Fourier transform grids and real space projectors were set using the "accurate" VASP precision mode. All structures were optimized using the conjugate gradient algorithm. For bulk structures we relaxed the lattice constants, angles, and ion positions. For defect structures, however, we optimized the ion positions in a fixed cell corresponding to the predicted equilibrium lattice parameters of the defect-free bulk. For DFT-HSE calculations, due to their computational cost, we used the DFT-SCAN-relaxed geometries followed by a single SCF DFT-HSE calculation. Note that we did not use a SCAN+U framework for any energy evaluation or structure relaxation because we previously found good agreement between DFT-SCAN-predicted and experimental formation enthalpies among binary transition metal sulfides. ⁵⁷ Finally, we set convergence thresholds of 1×10^{-5} eV for SCF loops and 3×10^{-2} eV \mathring{A}^{-1} for structure optimization loops.

The ESI† in ref. 104 contains additional details regarding the convergence of bulk thermodynamic quantities with respect to the kinetic energy cutoff, *k*-point grid density, and the inclusion of semicore states in the PBE PAW data sets for Cu and Ge.

We performed charged defect calculations using the PBE+*U*+D functional instead of SCAN because the latter grossly underestimates the band gaps of CZTS⁵⁷ and CZGS.¹⁰⁴ Further details about the phonon calculations to evaluate the thermodynamic stability of CCdGSSe can be found in Fig. S1–S3 and Table S4 in Section S2 of the ESI.†

Structural models

Polymorph structures. Prior to defect formation energy calculations, we determined the lowest energy crystal structure

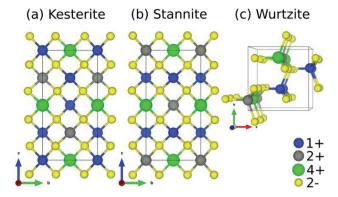


Fig. 2 Crystal structures of quaternary chalcogenides adopting the (a) kesterite, (b) stannite, and (c) wurtzite polymorphs. The prototypical composition for kesterite solar cells is Cu₂ZnSnS₄ where Cu¹⁺, Zn²⁺, Sn^{4+} , and S^{2-} occupy the 1+ (blue spheres), 2+ (grey), 4+ (green), and 2- (yellow) sites, respectively.

of the defect-free material. CZTS-like quaternary chalcogenides typically exist in one of three polymorphs: kesterite, stannite, and wurtzite-stannite (which, for simplicity, we will refer to as wurtzite),80 which are displayed in Fig. 2. All three polymorphs comprise corner-sharing tetrahedrons of four-fold anioncoordinated 1+, 2+, and 4+ cations. The kesterite and stannite polymorphs have similar space groups ($I\bar{4}$ and $I\bar{4}2m$, respectively) but differ in terms of their cation layer composition along c, where kesterite has alternating 1+/2+ and 1+/4+ layers and stannite has alternating 1+ and 2+/4+ layers. The main difference between wurtzite (space group: P63mc) and kesterite/ stannite is the underlying S-arrangement.

Bulk structures. For (partial) ion-substituted CZTS, we also determined the lowest energy ionic configuration. In the case of Ge-substituted CZTS, we considered both the kesterite and stannite polymorphs and two concentrations of Ge: dilute (x_{Ge} = Ge/(Sn + Ge) = 0.0625) and high ($x_{\text{Ge}} = 0.5$). For dilute Ge, which we refer to as ↓CZTGS, there is only one possible symmetry-unique Sn/Ge configuration within a 2 \times 2 \times 2 supercell, i.e., Ge in the Cu/Sn layer of kesterite and the Zn/Sn layer of stannite (see Fig. S4 in Section S3.1.1 of the ESI†). Note that, in order to calculate defect formation energies at infinite Ge dilution, we fixed the cell to that of relaxed CZTS. For all other bulk structures considered, we relaxed the lattice vectors and ions. For high Ge-substitution, which we refer to as ↑CZTGS, we optimize the cell and, to reduce the number of Sn/ Ge configurations, we only used the one possible symmetryunique Sn/Ge configuration in the 16-atom conventional cell (see Fig. S5 in Section S3.1.2 of the ESI†).

For Se-substituted CZGS (i.e., CZGSSe), we considered all symmetry-unique S/Se configurations for $x_{Se} = Se/(S + Se) = 0.5$ in the 16-atom conventional cell of kesterite and stannite CZGS. We used pymatgen's structure matcher, which is powered by spglib, 106 to generate the symmetry-unique configurations of which there are 12 for kesterite and 10 for stannite (see Fig. S6 in Section S3.2 of the ESI†). Finally, for CCdGSSe, in the same way as for CZGSSe, we considered all symmetry-unique S/Se configurations for $x_{Se} = 0.25$ in kesterite (there are five),

stannite (five), and wurtzite (11) Cu₂CdGeS₄ (CCdGS; see Fig. S7 in Section S3.3 of the ESI†). We included the wurtzite polymorph for CCdGSSe because it is observed experimentally for CCdGS107 and CCdGSe108 and it is nearly isoenergetic with stannite. Wurtzite is not relevant, however, for the quaternary chalcogenides containing Zn (CZTS, CZGS, CZTSe, or CZGSe) because it is less stable than the kesterite (ground state) and stannite (metastable) phases (see Table S5 in Section S4 of the ESI†).

Defect structures. All defects were generated in $2 \times 2 \times 2$ supercells of their defect-free, bulk kesterite, stannite, or wurtzite structures (consistent with previous studies57,72,77) using the protocol described in Table S6 in Section S5 of the ESI.† The objective of this study was to identify strategies, specifically those involving ion substitution, for the suppression of Cu_X + X_{Cu} and 2Cu_X + Y_X and the promotion of V_{Cu}. Y in this paper refers strictly to either Ge or Sn and does not refer to the element Yttrium. To this end, we primarily focused on ion substitution at the 2+ and 4+ sites as these are directly involved in the defects we were trying to suppress and have not yet been studied exhaustively. As shown in Fig. 3, we considered Zn, Mg, and Cd at the 2+ site; Sn, Ge, and Si at the 4+ site; and S and Se at the 2- site. These selections were the result of discarding elements that are radioactive (green), rare (gold), redox-active (magenta), possess the wrong oxidation state (i.e., not 2+, 4+, or 2-), or too small/large in terms of their ionic radius (blue) (see Fig. 3 caption for ionic radius filtering criteria). We did not

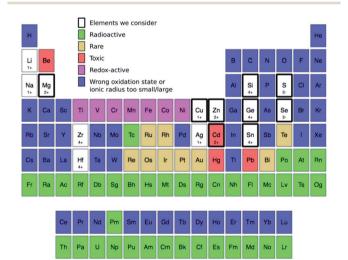


Fig. 3 Elements we consider in this study (thick black rectangles) to occupy the 1+, 2+, 4+, and 2- sites in Fig. 2: Cu (1+), Zn (2+), Cd (2+), Mg (2+), Sn (4+), Ge (4+), Si (4+), S (2-), and Se (2-). We filter these elements from the known 118 by removing, in the following order, those that are radioactive168 (green), rare (gold),132 redox-active (magenta), have either the wrong oxidation state or whose ionic radius is too different (blue). For ionic radius, we eliminate all elements that are too small (i.e., <30%) or too large (>170%) compared to the ionic radius of four-fold-coordinated Cu1+, Zn2+, Sn4+, and S2- corresponding to the occupation of 1+, 2+, 4+, and 2- sites, respectively. 149 While the use of toxic elements (red) should be restricted, 169 we retain them under consideration. For the remaining elements, we provide their most likely oxidation state in the prototype kesterite structure (Fig. 2a), e.g., Ge is 4+ and Se is 2-.

consider redox-active elements because they can promote defect formation, e.g., Mn can occupy both the 2+ and 4+ sites thereby promoting 2+/4+ disorder. We also did not consider Li⁺ and Zr⁴⁺ because Li₂ZnSnS₄ and Cu₂ZnZrSe₄ are wide-band-gap semiconductors (2.87 eV (ref. 109) and 1.95 eV, 110 respectively); Ag+, Na+, and Ti4+ because Ag- and Na-substituted CZTS, and Cu2ZnTi(S,Se)4 already have been studied computationally^{57,72,111} and experimentally;^{60,67,78,112} and Hf⁴⁺ because there is scarce experimental evidence of Hf forming quaternary chalcogenides (such as Cu2ZnHf(S/Se)4). We considered toxic elements (red) like Cd because toxicity can be mitigated through the practice of thoughtful device engineering.¹¹³ For example, previous efforts have successfully devised encapsulation schemes to ensure that CdTe solar cells, which contain more Cd by mass (47 wt%) than CCdGSSe (23 wt%), are significantly less toxic than they ought to be.114-116 Hence, depending on how sensitive to moisture and air CCdGSSe is in practice, similar encapsulation and water-proofing schemes will be useful for large-scale deployment. In contrast to toxicity, abundance, oxidation state, and ionic radius are, to a large extent, immutable characteristics and the effects of radioactivity are difficult to contain.

Ab initio thermodynamics

Convex hull construction. To determine the chemical potentials at which the quaternary chalcogenides (CZTS, CZGS, CZGSe, CCdGS, and CCdGSe) are in equilibrium with different combinations of secondary phases containing their constituent ions, we constructed 0 K phase diagrams using pymatgen, 105 which takes as inputs the DFT-SCAN (or DFT-HSE for CCdGSSe) total energies of all sub-quinary compounds and elements. Bulk structures for elements, binaries, ternaries, and quaternaries containing Cu, Zn, Mg, Cd, Sn, Ge, Si, S, and Se were taken from the inorganic crystal structure database (ICSD, 117 see Table S7 in Section S6 of the ESI†). We relaxed these structures with DFT-SCAN and used the same PAW potentials, kinetic energy cutoff, and k-point sampling density as above. Our previous work showed that DFT-SCAN systematically underestimates the formation energies of Ge-containing compounds by 0.27 eV/Ge.104 Therefore, we subtract 0.27 eV/Ge from all DFT-SCAN formation energies of Ge-containing compounds.

Defect formation energies. In the screening part of this study, we considered only neutral defects, meaning that atoms are removed from or added to the structure in their neutral elemental form. Once we identified an optimal composition, we then characterized charged defects for that composition. Our reasons for this more efficient screening approach were four-fold. (1) Charged-defect calculations are used primarily to identify defect transition energy levels within the band gap and to indicate their type (*i.e.*, shallow or deep) and their influence on the majority charge carrier concentrations. (2) Previous studies already reported the defect transition energy levels for Cu_X , X_{Cu} , V_{Cu} , and Y_X (*i.e.*, the main antisites we considered), and showed that (a) charged Cu_X , X_{Cu} and V_{Cu} defects produce shallow gap states that do not significantly promote nonradiative recombination^{36,41,57} and (b) while charged Y_X defects do produce trap states, ^{41,53} their

depth does not depend strongly on the identity of Y.58 On the other hand, the depth of Yx charged defect transition levels does depend on the identity of X (vide infra). For X = Zn, ref. 58 shows that Y_{Zn} defects produce trap states within the band gap of CZTS. We therefore did examine X = Cd within our optimal composition case. Our results (right panel of Fig. 8) show that Ge_{Cd} exhibits deep donor levels, which lie within the valence band, resulting in Ge_{Cd} exhibiting a neutral charge state across the band gap. Thus, we do not expect the donor transition levels of GeCd to act as electron traps in CCdGSSe. (3) We consider antisite clusters that are locally charge-balanced (e.g., the electrondeficient Cux compensates the electron-excessive Xcu adjacent to it in $Cu_X + X_{Cu}$; non-charge-balanced defect clusters generally are less stable than charge-balanced ones (see Table S8 and Fig. S8 in Section S7 of the ESI†). For materials other than our newly proposed CCdGSSe, the only charge-imbalanced defect we consider is V_{Cu}, whose charged formation energies were studied previously. 41,50 Like Y_X in CZTSe, CZTS, and CZGSe, however, the depth of its transition energy level does not depend strongly on the identity of the 2+, 4+, and 2- ions; only upon complete Ag-substitution in CZTSe (which is not considered here) does the transition energy level change.58 Additionally, since its transition energy level is shallow, increasing the Fermi level stabilizes V_{Cu}, which increases the majority charge carrier (i.e., hole) concentration and $V_{\text{oc.}}^{36}$ (4) Neutral defect formation energies are better metrics for estimating intrinsic defect concentrations than those of charged defects because the latter depend strongly on the Fermi level, which can be influenced by external conditions. Indeed, a recent experimental study on the photoluminescence of CCdGS118 indicate that the material has minimal carrier traps that lie deep within the band gap, which is in qualitative agreement with our neutral defect formation energy calculations.

Neutral defect formation energies were calculated as

$$\Delta E_{\rm f}^{\rm d} = E_{\rm d}^{\rm SCAN} - E_{\rm b}^{\rm SCAN} + \sum_{i} n_{i} \mu_{i} \tag{1}$$

where d is the defect supercell, b is the bulk supercell, n is the number of neutral atoms removed from (n > 0) or added to (n < 0)0) the system to form defect d, i is an index that runs over the unique species in the compound, and μ is the corresponding chemical potential. When point defects are created, the atoms that contribute to forming such defects must be exchanged with an external reservoir. Experimentally, this external reservoir may be ambient atmosphere during annealing, the current collecting phases that are in contact, or a secondary phase, such as leftover binary phases (e.g., CdS) that were used for the actual synthesis of the chalcogenide absorber. While theoretically we can calculate the formation energies of various isolated defects and defect complexes over a range of thermodynamically possible chemical potentials, the specific experimental conditions will precisely define the formation energy of a given defect. Here, we adopt the Cu-poor experimental conditions,57 with the phases that coexist with the quaternary/quinary chalcogenide determined by the 0 K phase diagram (i.e., convex hull) of the overall quaternary/ quinary system. We selected Cu-poor chemical potential limits for CZGSSe and CCdGSSe because CZTS and CCdTS

synthesized under this condition achieve high efficiency.34,50-52,77 For example, under Cu-poor conditions, CZTS is in equilibrium with ZnS, SnS2, and S; CZGS is in equilibrium with ZnS, GeS2, and S; and CZGSe is in equilibrium with ZnSe, GeSe2, and Se. CZTGS, CZGSSe, and CCdGSSe are metastable compounds at 0 K with the equilibrium compounds at their compositions being CZTS, CZGS, ZnS, SnS₂, and S; CZGS, ZnS, GeS₂, S, and Se; and CCdGS, Cd₄GeS₆, GeS₂, S, and Se, respectively. For metastable compounds, we use the chemical potentials determined by the stable compounds (see Table S9 in Section S8 of the ESI†). We did not consider the effect of temperature on the variation of chemical potentials in our calculations but we expect this effect to be small between 0 K and 298 K. For additional details on the computational methods, please see the ESI† and ref. 104, from which the DFT calculations performed herein are based.

For completeness, we did consider the following charged defects and defect clusters to examine the possible role of defect transition levels in our newly proposed CCdGSSe: Cu_{Cd} (q = -1, $0, 1), Cd_{Cu}(q = -1, 0, 1), Ge_{Cd}(q = -1, 0, 1, 2), V_{Cu}(q = -1, 0, 1),$ $V_S(q = -1, 0, 1, 2), V_{Se}(q = -1, 0, 1, 2), V_S + Cu_{Cd}(q = -1, 0, 1),$ $V_{Se} + Cu_{Cd}$ (q = -1, 0, 1), and $Ge_{Cd} + Cu_{Cd}$ (q = -1, 0, 1). Charged defect (cd) formation energies were calculated as $\Delta E_{\rm f}^{\rm cd} = \Delta E_{\rm f}^{\rm d}$ + $qE_{\rm F} + E_{\rm c}$ where $E_{\rm F}$ is the Fermi energy of the pristine bulk structure and E_c is the electrostatic correction term, calculated using the correction scheme of Kumagai and Oba119 as implemented in the Spinney package. 120 For further details about the correction scheme, we refer the interested reader to ref. 57, as we used the same approach here. The PBE+U dielectric tensor of CCdGSSe used in our charge defect calculations is: $\varepsilon_{xx} = \varepsilon_{yy} =$ 10.5, $\varepsilon_{zz} = 9.97$, $\varepsilon_{xy} = \varepsilon_{yz} = -0.06$, and $\varepsilon_{zx} = 0.06$.

Results

Effect of Ge-substitution in CZTS

Given the deep-trap-inducing nature of 2Cu_{Zn} + Sn_{Zn} antisite clusters in CZTS53 and the primary role played by the Sn_{Zn} antisite, 41 we first explore the replacement of Sn⁴⁺ with varying amounts of Ge4+. Fig. 4 shows the effect of dilute (light blue), high (dark blue), and complete (purple) Ge-substitution in CZTS (green) on kesterite vs. stannite polymorph preference (panel a), neutral defect formation energies under experimentally relevant Cu-poor conditions (panel b), and the PBE+U band gap (panel c). First, our results show that Ge-substitution does not affect polymorph preference (panel a) with kesterite being the groundstate polymorph ($E_{\text{Stannite}} - E_{\text{Kesterite}} > 0$ eV per formula unit) for all x_{Ge} (see Table S5 in Section S4 of the ESI† for CZTS and CZGS wurtzite data). That Ge increases the relative stability of kesterite vs. stannite for all but dilute Ge-substitution is in general agreement with other theoretical work 59,121,122 and the experimental observation of enhanced grain growth in Ge-doped CZTSSe. 123-125 Recent scanning electron microscope images show that grain growth and crystallinity in CZTGS is not improved for $x_{Ge} > 0.2^{121}$ but this could be due to suboptimal annealing conditions for each Ge composition during fabrication.107,126,127 Accordingly, non-dilute Ge-substitution should favor kesterite formation and suppress $E_{\rm g}/V_{\rm oc}$ lowering due to the kesterite -> stannite phase transition.80

Next, we analyze trends in defect formation energies as a function of x_{Ge} (panel b). For the $Cu_{Zn} + Zn_{Cu}$ antisite clusters, we find that, while partial Ge-substitution does not influence their formation (0.22 eV for CZTS⁵⁷ and ~0.23 eV for both ↓CZTGS and ↑CZTGS), complete substitution has a promoting effect (0.15 eV for CZGS), which can be explained by the greater

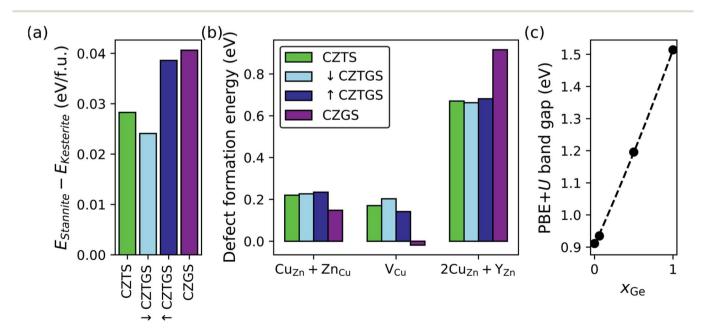


Fig. 4 Effect of Ge-substitution on (a) the DFT-SCAN relative stability of the stannite and kesterite polymorphs, (b) DFT-SCAN defect formation energies under Cu-poor conditions, and (c) the PBE+U band gap. In (b) Y is either Sn or Ge, whichever gives the lower defect formation energy; Y is Sn for CZTS, ↓CZTGS, and ↑CZTGS; and Ge for CZGS. (c) x_{Ge} is Ge/(Sn + Ge). Minimum energy defect configurations can be found in Fig. S9-S11 in Section S9.1 of the ESI.†

Paper

covalency of Ge–S bonds (Δ EN = 0.57, where EN is the Pauling electronegativity) compared to Sn–S bonds (Δ EN = 0.62). Ge–S bond covalency, which is greatest for complete Ge-substitution, reduces the charge density on S and promotes covalency across the Cu–S and Zn–S bonds. In turn, higher covalency reduces the effective 1+ and 2+ charges on the Cu and Zn, respectively, and consequently, the electrostatic energy penalty for Cu_{Zn} + Zn_{Cu} disorder. Therefore, CZGS should be more susceptible to Cu/Zn-disorder-induced potential fluctuations, as is seen in ref. 59, and potential $V_{\rm oc}$ deficit. Several studies found that partial Gesubstitution in CZTSSe increases the $V_{\rm oc}$, 124,128–131 however, we emphasize that these materials are selenized and, as we will explain in the next section, selenization suppresses stannite and promotes $V_{\rm Cu}$ in CZGS.

For Cu vacancies, with the exception of \downarrow CZTGS (0.20 eV), we predict that Ge-substitution decreases their formation energy (0.17 eV for CZTS, 57 0.14 eV for \uparrow CZTGS, and -0.02 eV for CZGS) to the extent that DFT-SCAN predicts a nonzero equilibrium concentration of V_{Cu} in CZGS at 0 K. In all likelihood, this negative value for V_{Cu} in CZGS is an artifact of DFT-SCAN's tendency to produce lower values of $\Delta E_{\rm f}^{\rm d}$ compared to DFT-PBE, PBE+U, SCAN+U, and DFT-HSE. ¹⁰⁴ For example, compared to DFT-HSE, DFT-SCAN predicts the $\Delta E_{\rm f}^{\rm d}$ for V_{Cu} to be lower by 0.53 eV in CZTS¹⁰⁴ and 0.58 eV in CZGS. Importantly, qualitative trends in $\Delta E_{\rm f}^{\rm d}$ are XC-functional-insensitive (see Fig. S5 in the ESI† for ref. 104) and we therefore simply interpret this negative value as "high to complete Ge-substitution promotes V_{Cu} formation".

The decrease in the $V_{\rm Cu}$ formation energy upon complete replacement of Sn with Ge can be rationalized as follows. The Ge–S 298 K neutral diatomic bond dissociation enthalpy (BDE = 5.54 ± 0.03 eV) is substantially larger than that of Sn–S (4.84 eV) 132 and, by replacing Sn–S bonds with stronger Ge–S bonds, Ge-substitution effectively weakens Cu–S bonds. As a result, less energy is required for Cu to break its bonds with S and form vacancies. In terms of solar cell parameters, the low $\Delta E_{\rm f}^{\rm d}$ for $V_{\rm Cu}$ in CZGS should mitigate the $V_{\rm oc}$ -lowering effects associated with complete Ge-substitution.

Finally, for 2Cu_{Zn} + Y_{Zn} antisite clusters, where Y is either Sn or Ge, it is clear that complete Ge-substitution strongly suppresses their formation (0.92 eV) compared to CZTS (0.67 eV), \downarrow CZTGS (0.66 eV), and \uparrow CZTGS (0.68 eV). The difference between CZTS and CZGS (0.25 eV) can be attributed to the inert pair effect, which stabilizes the 2+ oxidation state of Sn and, therefore, stabilizes Sn on the Zn²⁺ site by reducing it from 4+ to 2+. Ge, on the other hand, only exists stably in a 4+ oxidation state, as evidenced by the positive ΔG_{298} for the reduction of GeS₂ to GeS (0.80 eV), i.e., Ge(IV)S₂ \rightarrow Ge(II)S + S.¹³³ Furthermore, Ge-substitution suppresses 2Cu_{Zn} + Y_{Zn} only under complete substitution because Ge does not affect the 2Cu_{Zn} + Sn_{Zn} formation energies and these defect clusters will continue to form as long as Sn, which is susceptible to reduction via the inert pair effect, is present in the structure. Hence, unless the more redox active Sn is completely replaced by the less redox active Ge, the $\Delta E_{\rm f}^{\rm d}$ for $2Cu_{\rm Zn} + Y_{\rm Zn}$ should remain close to that for CZTS. With that being said, partial Gesubstitution still can suppress the formation of 2Cu_{Zn} + Y_{Zn}

by reducing $x_{\rm Sn}$, thus shedding light on the experimentally observed increase in minority charge carrier lifetimes for Gealloyed ($x_{\rm Ge}=0.3$) CZTSSe.¹³⁰

Via its substantial suppression of 2Cu_{Zn} + Y_{Zn}, as compared to other strategies in the literature, ⁷⁷ CZGS should offer a significant reduction in the rate of SRH recombination νs. CZTS. Unfortunately, despite the promising defect thermodynamics of CZGS, the band gap depends strongly on x_{Ge} (panel c) with a quasi-linear increase from CZTS (0.91 eV for PBE+U and ≈ 1.5 eV in experiments⁸⁰) to CZGS (1.51 eV for PBE+U and ≈ 2.1 eV in experiments¹³⁴). In summary, while only CZGS suppresses 2Cu_{Zn} + Y_{Zn} (rightward trajectory in Fig. 1), its E_g is well above E_g^{SQ} (upward trajectory), thus prompting our next ion substitution strategy.

Effect of selenization on CZGS

One way to reduce the band gap of CZGS is Se-substitution (also known as selenization);134-138 this has been used extensively to reduce/optimize the band gap of CZTS. 139,140 Band gap reduction in both CZTS and CZGS upon selenization can be explained by the slightly lower electronegativity of Se (2.55 on the Pauling scale141,142) compared to S (2.58), which pushes the valence band edge (comprising mostly Cu 3d and anion p states) up in energy but does not affect the conduction band edge (mostly Sn 5s).139 As a result, the band gap shrinks and the extent to which it does depends on $x_{Se} = Se/(S + Se)$. In Fig. 5, we consider both 50% (light green bars) and 100% (cyan bars) selenization (CZGSSe and CZGSe, respectively) in CZGS (yellow bars). Fig. 5a shows that the crystal structure of the lowest energy S/Se configuration of kesterite CZGSSe (see Section S3.2 in the ESI†) is composed of alternating S and Se (011) planes (red lines). As expected, our results show that selenization decreases E_g^{PBE+U} (panel b) from 1.51 eV (CZGS) to 0.59 eV (CZGSe), with the magnitude of the decrease in excellent agreement with the experimental 134,137,138 and theoretical literature. 135 Additionally, we find that the PBE+U band gap of CZGSSe (0.92 eV) is approximately equal to the optimal CZTS value (black dotted line at 0.91 eV). With respect to polymorph preference, the introduction of 50% Se (CZGSSe) marginally stabilizes the low- E_g stannite polymorph (green, panel c) relative to CZGS (yellow) but not with respect to CZTS (black dotted line), thus signaling that selenization should not exacerbate polymorphism-derived V_{oc} deficits.

In view of the favorable band gap and stability of kesterite CZGSSe, we investigate the dependence of defect formation energies on $x_{\rm Se}$ (panel d). First, we find, as do powder neutron diffraction measurements,⁵⁹ that selenization slightly suppresses the formation of $V_{\rm oc}$ -reducing ${\rm Cu_{Zn}} + {\rm Zn_{Cu}}$ antisite clusters relative to CZGS (0.15 eV for CZGS versus 0.20 and 0.16 eV for CZGSSe and CZGSe, respectively; CZTS is 0.22 eV). Furthermore, we predict that selenization promotes $E_{\rm g}$ -increasing/ $V_{\rm oc}$ -restoring Cu vacancies (-0.02 eV for CZGS versus -0.12 and -0.09 eV for CZGSSe and CZGSe, respectively), due to the decrease in bond strength from Cu–S (BDE = 2.85 ± 0.15 eV) to Cu–Se (2.64 ± 0.15 eV) 132 and the endoergicity of Cu₂S + Se \rightarrow Cu₂Se + S ($\Delta G_{298} = 0.12$ eV). In spite of these promising trends, selenization counteracts the Ge-induced suppression of $I_{\rm Sc}$ -

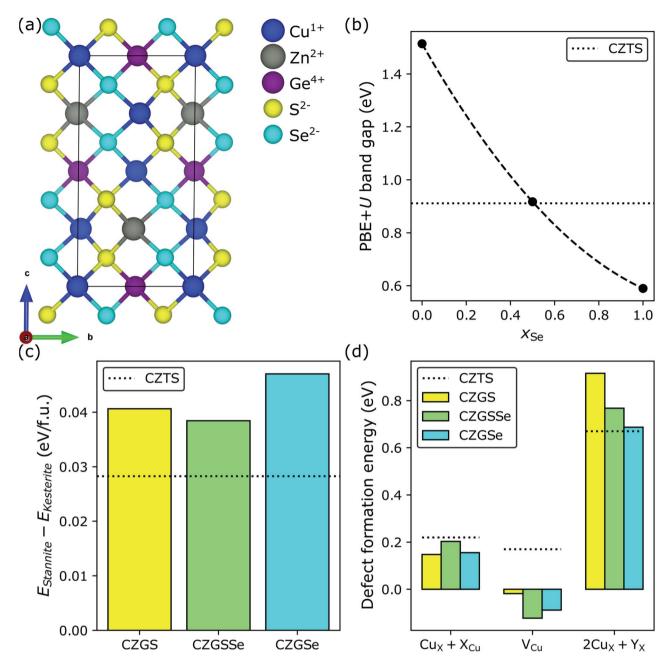


Fig. 5 Selenization of CZGS. (a) Crystal structure of the minimum-energy configuration of Cu₂ZnGeS₂Se₂ (CZGSSe), where the anion ordering is similar to how cations order in CuIn(S,Se)₂. Effect of selenization on (b) the PBE+U band gap, (c) the DFT-SCAN relative stability of the stannite and kesterite polymorphs, and (d) DFT-SCAN defect formation energies under Cu-poor conditions. Minimum-energy defect configurations can be found in Fig. S12 and S13 in Section S9.2 of the ESI.†

lowering $2Cu_{Zn} + Ge_{Zn}$ antisite clusters (0.92 eV for CZGS > 0.77 eV for CZGSSe > 0.69 eV for CZGSe \approx 0.67 eV for CZTS, black dotted line). We ascribe this phenomenon to the following mechanism. While 2Cu_{Zn} + Ge_{Zn} is charge-balanced, the balancing charges (two holes and two electrons from 2Cu_{Zn} and Ge_{Zn}, respectively) are slightly delocalized. Therefore, it is reasonable to expect that the stability of $2Cu_{Zn} + Ge_{Zn}$ depends, to some extent, on the anion's tolerance for delocalization. Since Se is less electronegative (and more polarizable) than S, it should be able to accommodate more delocalization and, therefore, stabilize 2Cu_{Zn} + Ge_{Zn}. Thus, selenization of CZGS should decrease I_{sc} , counteracting the beneficial suppression of Cu_{Zn} + Zn_{Cu} and promotion of V_{Cu}. A final ion substitution step therefore must be taken in order to minimize the need for E_g -remediating but I_{sc} -decreasing selenization in CZGS.

Cooperative effect of Cd-substitution and selenization on **CZGS**

Recently, it was predicted and subsequently observed that replacing the Zn in CZTS with Cd, under Cu-poor conditions,

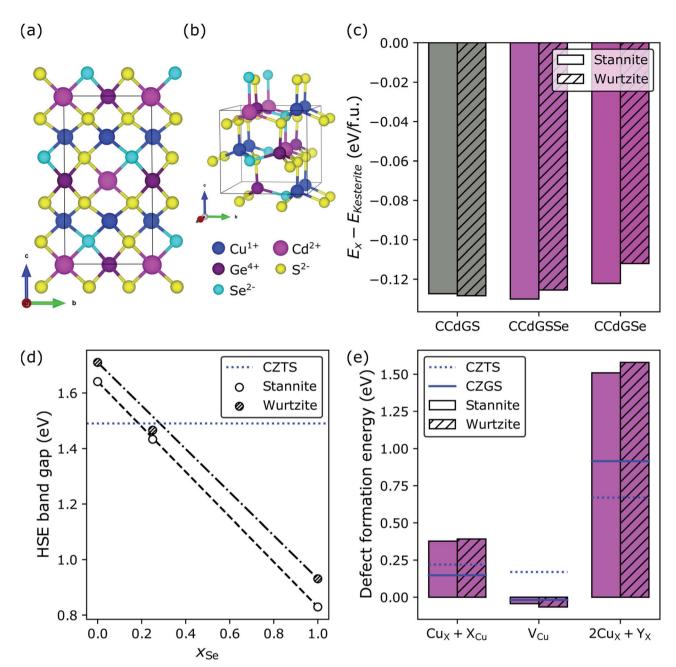


Fig. 6 Cd-substitution and selenization of CZGS. Crystal structures of the minimum-energy configurations of the (a) stannite and (b) wurtzite polymorphs of Cu_2CdGeS_3Se (CCdGSSe). Effect of Cd-substitution and selenization on (c) the DFT-HSE relative stabilities of the stannite, wurtzite, and kesterite polymorphs, (d) the DFT-HSE band gap (CZTS value is from ref. 143 while the other band gaps were calculated by us), and (e) DFT-SCAN defect formation energies under Cu-poor conditions for CCdGSSe. Minimum-energy defect configurations can be found in Fig. S14 and S15 in Section S9.3 of the ESI.†

decreases its band gap and suppresses the formation of the deep-trap-level-inducing $2\mathrm{Cu_X} + \mathrm{Sn_X}$, where X is either Zn or $\mathrm{Cd.^{57,77}}$ For these reasons, Cd-substitution in CZGS should limit the extent to which selenization is necessary and further suppress the formation of I_{sc} -lowering defects. To test this hypothesis, in Fig. 6, we consider complete Cd-substitution in CZGS, leading to CCdGS (grey bars), and with 25% (purple bars) and 100% selenization (magenta bars), where the former was chosen as an intermediate Se-substitution condition by

interpolating between the DFT-HSE band gaps of CCdGS and CCdGSe to find the value of $x_{\rm Se}$ for which $E_{\rm g}^{\rm HSE}[{\rm CCdGSSe}] \approx E_{\rm g}^{\rm HSE}[{\rm CZTS}] = 1.49~{\rm eV.}^{143,144}$ Fig. 6a and b show the ground state S/Se configuration in stannite and wurtzite CCdGSSe, respectively (see Section S3.3 in the ESI†). In what follows, we examine both of these polymorphs because they are nearly isoenergetic (panel c) and more stable than kesterite ($E_{\rm x} - E_{\rm Kesterite} < 0$). Fortunately, the DFT-HSE band gaps of stannite and wurtzite CCdGSSe are similar in terms of their dependence on $x_{\rm Se}$ (panel

d) and achieve optimality at 25% selenization (1.43 and 1.47 eV for stannite and wurtzite, respectively), which suggests that, despite the polymorphism that is likely present in real samples, the energy offset between bands at stannite-wurtzite interfaces and, therefore, interfacial recombination rates 145-148 should be small. Furthermore, optimality at $x_{Se} = 0.25$ indicates that complete Cd-substitution cuts the need for Se-substitution in half, compared to CZGS \rightarrow CZGSSe.

Perhaps the most dramatic effect of Cd-substitution is on the defect thermodynamics of CZG(S,Se). Fig. 6e reveals that Cd-substitution suppresses the formation of V_{oc} -lowering Cu_{Cd} + Cd_{Cu} antisite clusters in CCdGSSe (0.38 eV and 0.39 eV for stannite and wurtzite, respectively) compared to $Cu_{Zn} + Zn_{Cu}$ in CZTS (blue dotted line at 0.22 eV) and more than doubles the $Cu_{Zn} + Zn_{Cu} \Delta E_f^d$ in CZGS (blue solid line at 0.15 eV). The suppression of Cu_X + X_{Cu} can be attributed to the larger ionic radius of Cd²⁺ compared to Cu⁺ and Zn²⁺. Additionally, Cdsubstitution promotes the formation of V_{oc} -increasing Cu vacancies (-0.04 eV and -0.07 eV for stannite and wurtzite, respectively) compared to CZTS (0.17 eV), while the $\Delta E_{\rm f}^{\rm d}$ for $V_{\rm CH}$ in CCdGSSe is quite similar to CZGS (-0.02 eV). Thus, the stronger Ge-S bonds (versus Sn-S bonds) primarily govern the ease of formation of V_{Cu} in Ge-containing quaternary/quinary chalcogenides. Most importantly, Cd-substitution leads to a remarkably strong suppression of 2Cu_{Cd} + Ge_{Cd} in CCdGSSe (1.51 and 1.58 eV in stannite and wurtzite, respectively) relative to CZTS (0.67 eV) and CZGS (0.92 eV). The much higher $\Delta E_{\rm f}^{\rm d}$ corresponds to a roughly seven orders of magnitude decrease in x_{SRH} (see Section S10 in the ESI†) under CZTS annealing conditions (600 K),34,35 i.e., the maximum temperature to which CZTS is heated during synthesis, thus constituting an upper bound for crystalline defect concentrations. The significant suppression of 2Cu_{Cd} + Ge_{Cd} can be attributed to the large ionic radius difference between Cd²⁺ (0.78 Å in tetrahedral coordination) and Ge⁴⁺ (0.39 Å).¹⁴⁹ Note that the local anionic configuration of CCdGSSe does not affect significantly its defect formation energies ($\sigma_{\Delta E_{c}^{d}} = 0.06$ eV, see Fig. S16 and Table S10 in Section S11 of the ESI†).

To the best of our knowledge, there are neither reports on the synthesis of CCdGSSe (no entries in both the ICSD and Chemical Abstracts Service), experimental/computational characterizations of its solar cell parameters ($E_{\rm g}/V_{\rm oc}$, $I_{\rm sc}$, and FF), nor publications/patents detailing the fabrication of CCdGSSebased solar cells. Therefore, CCdGSSe may provide the innovation necessary to improve the efficiency of kesterite-inspired solar cells beyond their 12.6% efficiency record for CZTSSe⁷⁹ that has remained stationary since 2013, and propel them toward competitiveness¹⁵⁰ with Si (26.7%)³ and hybrid perovskites (25.2%).151

Discussion

Having discussed Ge-, Se-, and Cd-substitution individually (in order to reveal their independent effects on defect thermodynamics and electronic structure), we now, in Fig. 7, compare the performance of CZTS (green), CZGS (red), CZGSSe (blue), and CCdGSSe (magenta). The origin, *i.e.*, (0,0,0), is the center of the

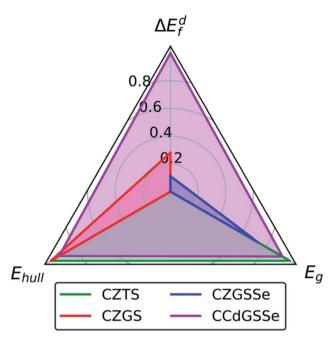


Fig. 7 Variation in performance among CZTS, CZGS, CZGSSe, and CCdGSSe where ΔE_f^d is the normalized $2Cu_X + Y_X$ formation energy under Cu-poor conditions, E_{hull} is the energy/atom above the hull, and E_q is the normalized absolute deviation of the band gap from that of CZTS. Normalization, i.e., $(x - x_{min})/(x_{max} - x_{min})$, is used to bring all values of the performance indicator x into the range [0,1]. For ΔE_f^d , values of 0 and 1 correspond to the $2Cu_{Zn} + Sn_{Zn}$ and polymorphaveraged $2Cu_{Cd} + Ge_{Cd}$ formation energies under Cu-poor conditions for CZTS (0.67 eV) and CCdGSSe (1.54 eV), respectively. For E_{hull} , values of 0 and 1 correspond to 0.073 eV per atom (CZGSSe) and 0 eV per atom above the hull, respectively (see Table S11 in Section S12 of the ESI†). For E_{g} , values of 0 and 1 correspond to 0.60 eV and 0 eV deviations from the DFT-HSE band gap of CZTS.143

black triangle and the axes, i.e., the closed line segments bounded by the origin and the vertices of the black triangle, correspond to the intrinsic stability/instability (energy above the convex hull at 0 K, E_{hull}), the band gap (E_g) , and the formation energy of deep-trap inducing 2Cu_X + Y_X defect clusters $(\Delta E_{\rm f}^{\rm d})$, which directly relate to the synthesizability, $V_{\rm oc}$, and $I_{\rm sc}$, respectively. Each axis is normalized as: (1) $\Delta E_{\rm f}^{\rm d} = 0 \equiv 0.67 \text{ eV}$ (taken from CZTS under Cu-poor conditions) and $= 1 \equiv 1.54 \text{ eV}$ (average of stannite- and wurtzite-CCdGSSe); (2) $E_{\text{hull}} =$ $0 \equiv 0.073$ eV per atom (from CZGSSe) and $1 \equiv 0$ eV per atom; and (3) E_g of 0 and 1 correspond to 0.60 eV and 0 eV deviations from the $E_{
m g}^{
m HSE}$ of CZTS, where we plot the DFT-HSE band gaps for CZTS,143 CZGS,152 and CCdGSSe (this work), and an experimental gap for CZGSSe.134 Fig. 7 shows that the performance of CZTS (roughly the area of the green triangle) comes from its stability ($E_{\text{hull}} = 0$ eV per atom) and optimal band gap ($E_{\text{g}}^{\text{HSE}} =$ 1.49 eV (ref. 143) $\approx E_{\rm g}^{\rm SQ}$). However, CZTS is limited by the ease with which I_{sc} -reducing $2Cu_{Zn} + Sn_{Zn}$ defect clusters form $(\Delta E_{\rm f}^{\rm d} = 0.67 \text{ eV})$. Complete Ge-substitution (CZGS, red), on the other hand, suppresses the formation of these detrimental defects ($\Delta E_{\rm f}^{\rm d}=0.92~{\rm eV}$), leading to a performance increase along the $\Delta E_{\mathrm{f}}^{\mathrm{d}}$ axis, but widens the band gap ($E_{\mathrm{g}}^{\mathrm{HSE}} = 2.09$ eV (ref. 152)) too far beyond that of the nearly optimal CZTS value. Partial

selenization (CZGSSe, blue) can be used to improve the band gap ($E_{\rm g}^{\rm exp}=1.65$ eV (ref. 134)) but it also reduces both the Ge-induced suppression of deep defects ($\Delta E_{\rm f}^{\rm d}=0.77$ eV) and the intrinsic stability of the material ($E_{\rm hull}=0.073$ eV per atom, $\approx 2.8 \times k_{\rm B}T$ at 298 K). Remarkably, complete Cd-substitution (CCdGSSe, magenta) practically eliminates $I_{\rm sc}$ -decreasing 2+/4+ disorder ($\Delta E_{\rm f}^{\rm d}=1.54$ eV), optimizes the band gap ($E_{\rm g}=1.45$ eV), and effectively stabilizes the material ($E_{\rm hull}=0.005$ eV per atom < $k_{\rm B}T$ at 298 K, which is likely thermally accessible).

The stability of CCdGSSe is supported by reports of the synthesis and characterization of the closely related compounds $\mathrm{Cu_2CdGeSe_4}$ (ref. 107) and $\mathrm{Cu_2CdGeSe_4}$, $^{108,153-155}$ which differ from CCdGSSe only in S–Se ratio. For both compounds, we predict $E_{\mathrm{hull}}=0$ eV per atom, which is consistent with their synthesizability (see Table S11 in Section S12 of the ESI†). Additionally, we predict that CCdGS and CCdGSe prefer the wurtzite and stannite polymorphs, respectively [see Fig. 6c], which agrees with X-ray diffraction measurements. $^{107,108,153-155}$ The excellent agreement between theory and experiment for these very closely related compounds CCdGS and CCdGSe suggests that our prediction of the bulk stability of CCdGSSe should be reliable.

We also calculated the thermodynamics of CCdGSSe phase separation, i.e., Cu₂CdGeS₃Se \rightarrow 3/4Cu₂CdGeS₄ + 1/4Cu₂CdGeSe₄, and Fig. S1 in Section S2 of the ESI† shows that the

configurational entropy of ideal S/Se mixing on the anion sublattice stabilizes CCdGSSe at temperatures above 133 K. Note that the 0 K convex hull at the composition of Cu₂CdGeS₃Se actually consists of Cu₂GeS₃ + CdSe and not the mixture of wurtzite-Cu₂-CdGeS₄ and stannite-Cu₂CdGeSe₄. We find that the S/Se mixing entropy and vibrational entropic contributions stabilize CCdGSSe with respect to the decomposition into Cu₂GeS₃ + CdSe at temperatures above 879 K (see Fig. S2 and S3 and Table S4 in Section S2 of the ESI†), which is comparable with typical sulfurization temperatures of CZTS and CCdTS, *i.e.*, 580–600 °C or 853– 873 K.⁷⁷

The calculated band gap of CCdGSSe (1.45 eV) is bounded by the measured band gaps of CCdGS (1.85–2.05 eV)^{118,156–158} and CCdGSe (1.14–1.27 eV)^{138,159–161} (see Table S12 in Section S13 of the ESI†), which suggests that our prediction of the optimality of the band gap of CCdGSSe is accurate. Since the measured band gaps of quinary metal sulfide-selenides are proportional to $x_{\rm Se}$, ^{140,162} that of CCdGSSe should be approximately 1.67–1.86 eV, as found by linear interpolation. While this is slightly greater than the nearly ideal band gap of CZTS (1.49 eV (ref. 143)), further selenization to Cu₂CdGeS₂Se₂ ($x_{\rm Se}=0.5$) optimizes the interpolated band gap (1.50–1.66 eV) and does not influence our conclusion that complete Cd- and Ge-substitution strongly suppresses the formation of detrimental defects (see Table S13 in Section S13 of the ESI†).

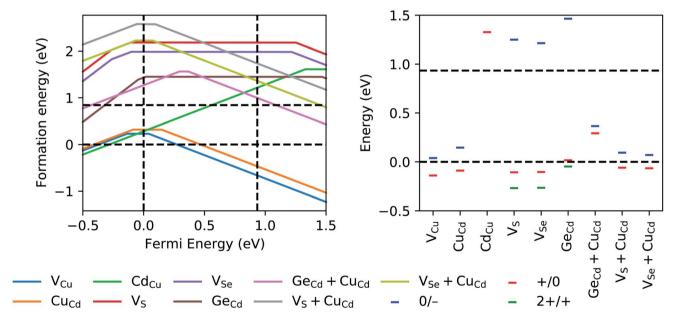


Fig. 8 (Left) GGA+U+D formation energies under Cu-poor conditions (*i.e.*, when CCdGS, Cd₄GeS₆, GeS₂, S, and Se are in equilibrium) of various charged vacancies, antisites, vacancy-antisite clusters, and antisite clusters considered within stannite CCdGSSe, plotted as a function of the Fermi energy. Vertical, dashed, black lines at 0 eV and 0.93 eV indicate the GGA+U+D VBM and CBM, respectively, where the zero of the Fermi energy is arbitrarily set to the VBM. For V_{Cu}, Cu_{Cd}, Cd_{Cu}, Ge_{Cd} + Cu_{Cd}, V_S + Cu_{Cd}, and V_{Se} + Cu_{Cd}, three ionized states are considered, namely, q = -1 (negatively charged, slope of -1), q = 0 (neutral, zero slope), and q = 1 (positively charged, slope of 1). For V_S, V_{Se}, and Ge_{Cd}, we also consider q = 2 (doubly positively charged, slope of 2). The solid lines indicate the minimum formation energy at each Fermi energy, indicative of the most stable charged state of the defect. Horizontal, dashed, black lines at 0 eV and 0.84 eV respectively indicate the threshold for exothermic defect formation and the GGA+U+D formation energy of the neutral 2Cu_{Zn} + Sn_{Zn} antisite cluster, ¹⁰⁴ which causes Shockley–Read–Hall recombination in CZTS. ^{41,53,54} The latter constitutes an upper bound on the formation energy of defects that affect solar cell parameters. (Right) Transition levels of the defects in left panel. Blue, red, and green dashes respectively signify acceptor $(q, 0 \rightarrow -1)$, donor $(q, 1 \rightarrow 0)$, and donor $(q, 2 \rightarrow 1)$ transition levels.

It also is crucial to know, for a given Fermi level, the most stable charge state and deep/shallow character of the defect and its potential trapping ability. We therefore calculated the formation energies of the following charged defects and defect clusters in stannite CCdGSSe, which are consistent with those considered in ref. 58: V_{Cu}, Cu_{Cd}, Cd_{Cu}, V_S, V_{Se}, Ge_{Cd}, V_S + Cu_{Cd}, and V_{Se} + Cu_{Cd} . The left panel of Fig. 8 shows that only V_{Cu} (q =-1, 0), Cu_{Cd} (q = -1, 0), and Cd_{Cu} (q = 1) should form in appreciable quantities for Fermi energies between the valenceband maximum (VBM) and conduction-band minimum (CBM), and under Cu-poor conditions (i.e., when CCdGS, Cd₄GeS₆, GeS₂, S, and Se are in equilibrium). The defect transition levels in the right panel of Fig. 8 agree qualitatively with those for CZTSe, CZTS, CZGSe, and Ag₂ZnSnSe₄ in ref. 58, with the exception of Ge_{Cd}, which exhibits deep donor levels within the valence band. The formation energy of Ge_{Cd} (1.44-1.45 eV), however, is well above that of the neutral $2Cu_{Zn} + Sn_{Zn}$ antisite cluster in CZTS (0.84 eV), 104 which constitutes an upper bound on the formation energy of defects that affect solar cell parameters. Ge_{Cd} + Cu_{Cd} also has deep donor and acceptor levels, however, we discard this defect as well because of its high formation energy. Therefore, CCdGSSe should exhibit lower non-radiative recombination than CZTS.

The acceptor transition levels of V_{Cu} and Cu_{Cd} at $E_F - E_{VBM} =$ 39 meV and 146 meV, respectively, agree quantitatively with those in ref. 118, i.e., 27-34 meV and 157 meV, respectively. Therefore, V_{Cu} and Cu_{Cd} likely are present in detectable quantities for CCdGSSe but we expect the absolute Cux concentrations to be lower than in CZTS ($\Delta E_{\rm f}^{\rm d} = -0.20$ eV for Cu_{Zn} in CZTS versus 0.02 eV for Cu_{Cd} in CCdGSSe - see Table S8 in Section S7 in the ESI†). The right panel of Fig. 8 also shows that only Cd_{Cu} has a donor transition level that is close to the conduction band but it is within the band itself. Therefore, we cannot assign the defects and/or defect complexes we considered to the experimentally observed deep donor transition levels at \sim 50-120 meV below the CBM.118

In addition to developing a feasible and original ion substitution scheme for optimizing kesterite solar cells, we also clarify the mechanisms of defect suppression and promotion by Ge and Se, emphasizing the crucial role of the inert pair effect and metal-chalcogen bond covalency, respectively. These insights reveal three important materials design considerations for tuning defect formation in kesterite-type absorbers via ion substitution: (1) bond order (BO) conservation, (2) cation redox inactivity, and (3) anion polarizability. The BO conservation principle states that the sum of the strengths of the bonds to an anion from its adjacent cations is nearly equal to its valence.163 In other words, if some of the anion-cation bonds become stronger, e.g., upon cation substitution, then the others must become weaker in order to conserve the BO of the anion. Such bond strength reciprocity can be used, for example, to tune the formation energy of V_{oc} -increasing Cu vacancies. By replacing Sn with an element that forms stronger bonds with S, such as Ge (BDE = 5.54 ± 0.03 eV for Ge-S vs. 4.84 eV for Sn-S), ¹³² BO conservation dictates that the Cu-S bonds must become weaker; therefore, Cu vacancies should form more easily in CZGS than CZTS, as is shown in Fig. 4b ($\Delta E_{\rm f}^{\rm d} = -0.02$ eV for

CZGS vs. 0.17 eV for CZTS). This same principle is illustrated upon examination of how the formation energies of neutral chalcogen vacancies and their neutral clusters with Cux antisites, which also have been predicted theoretically to be strong electron traps, 164 change with X and Y. Table S14 in Section S14 of the ESI† shows that generally Cd (Ge) decreases (increases) the formation energies of these defects compared to Zn (Sn). The effect of compositional changes on the formation energies of these defects again can be rationalized on the basis of 298 K neutral diatomic bond dissociation enthalpies (BDEs): Cd-S bonds (2.16 \pm 0.22 eV) are weaker than Zn-S bonds (2.33 \pm 0.13 eV) and Ge-S bonds (5.53 eV) are stronger than Sn-S (4.84 eV). Note that uncertainties in the Ge-S and Sn-S BDEs are unavailable.132

To illustrate the effect of cation redox activity on defect formation, consider the 2Cu_{Zn} + Sn_{Zn} antisite cluster. While Zn exhibits only one normal oxidation state (2+), Cu and Sn exhibit two (1+ and 2+ for Cu and 2+ and 4+ for Sn), where Sn²⁺ is stabilized by the inert pair effect. The redox activities of Cu and Sn promote the formation of Cuzn and Snzn antisites, respectively, because they both can adopt the 2+ oxidation state of Zn, thereby reducing the electrostatic energy penalty associated with the formation of these antisites and, consequently, their clusters. Replacing Sn (2+ and 4+) with an element that is redox inactive, such as Ge (4+), should suppress the formation of Isc-reducing Zn/4+ disorder, as is shown in Fig. 4b $(\Delta E_{\rm f}^{\rm d} = 0.92 \text{ eV for CZGS } \nu s. 0.67 \text{ eV for CZTS}).$

A final point to consider when designing kesterite solar cell materials is the polarizability of the anion. To explain its significance, we again consider the recombination-inducing $2Cu_{Zn} + Y_{Zn}$ antisite cluster, now focusing on the case where Y = Ge. Despite the fact that this defect is neutral and globally charge compensated (i.e., 2Cu_{Zn} and Ge_{Zn} produce two holes and two electrons, respectively, which cancel each other out), the generated charge carriers are, to some extent, delocalized. 104 Since Ge4+ is resistant to reduction, with reduction being a likely consequence of localized electrons, the stability of this particular defect depends on the anion's tolerance for delocalization, which can be measured in terms of its polarizability. Given that Se $(3.89 \times 10^{-24} \text{ cm}^3)$ is more polarizable than S $(2.87 \times 10^{-24} \text{ cm}^3)$, 132 it should be able to stabilize this detrimental defect, which is precisely what we find in Fig. 5d, where the formation energy of 2Cu_{Zn} + Ge_{Zn} decreases almost linearly with increasing x_{Se} .

Based on these considerations, there are a few ion substitution strategies worth exploring further. Note that we, by no means, intend to suggest that the following strategies are exhaustive or the only ones worth pursuing; they are simply those that were inspired by the considerations above. First, while several studies involve Ge- and Se-co-substitution, x_{Ge} and x_{Se} concurrently; x_{Ge} and x_{Se} concurrently; for that reason, there may be room to further improve CZGSSe via concurrent and comprehensive variation of x_{Ge} and x_{Se} . Additionally, since we consider only three possible compositions of CCdGSSe and a few, albeit important, defects, opportunities exist for theory to provide a better understanding of its composition-dependent properties and experiments to validate and augment its performance, *e.g.*, *via* cell fabrication. Another strategy that has proven effective is to replace $\mathrm{Cu^+}$ with redox inactive 1+ cations such as the alkali metal cations⁶²⁻⁷² and $\mathrm{Ag^+}$.^{57-61,167} While quaternary and quinary chalcogenides containing these cations have been studied extensively, there remains a great need to explore their post-quinary combinations in conjunction with Ge-, Se-, and Cd-substitution. For example, we suggest Na-doping in either CCdGSSe or CZGSSe (to avoid Cd) as a promising pathway to further mitigate the formation of disorder-causing and V_{oc} -reducing antisites and improve the performance of kesterite-based solar cells.

Conclusions

This work introduces the new photovoltaic material CCdGSSe, which we predict to be nearly thermodynamically stable, have an ideal band gap, and offer substantial suppression of the defects that plague the efficiency of other Cu2ZnSnS4-based materials, based on extensive DFT-based calculations. Additionally, we ascribe mechanisms to the dependence of key, performance-affecting defect formation energies on Ge and Se content, revealing how: (1) bond order conservation can be used to tune vacancy formation energies (e.g., decrease the formation energy of V_{oc} -increasing Cu vacancies); (2) the inert pair effect of Sn leads to its redox flexibility and thus promotes I_{sc} -reducing Zn/Sn disorder; and (3) the polarizability of Se stabilizes the local charge imbalances created by cation disorder, thereby reducing the thermodynamic barrier for defect cluster formation. These insights and the materials design principles gleaned in this work should provide the community of CZTS researchers with a heightened intuition for how to optimize kesterite solar cell materials, hopefully enabling this promising photovoltaic technology to reach its full potential.

Conflicts of interest

There are no conflicts to declare.

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