


 Cite this: *RSC Adv.*, 2021, 11, 16388

Thermal decomposition of ammonium perchlorate-based molecular perovskite from TG-DSC-FTIR-MS and *ab initio* molecular dynamics†

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(H₂dabco)[NH₄(ClO₄)₃] (DAP, dabco = 1,4-diazabicyclo[2.2.2]octane) is a recently synthesized ammonium perchlorate-based molecular perovskite energetic material. The high-symmetry perovskite configuration assembles the oxidant ClO₄⁻ and fuel H₂dabco²⁺ into a compact cubic crystal, realizing a high energy-releasing efficiency. In this study, the thermal decomposition of DAP has been investigated by thermogravimetric analysis (TG) and differential scanning calorimetry (DSC) coupled with Fourier transform infrared (FTIR) spectroscopy and mass spectroscopy (MS). The TG-DSC profiles show that DAP has an intense one-stage heat release process with a weight loss of 94.7%. The evolved gas products are identified as H₂O, CO₂, CO, HCl, HCN, NH₃, HNCO by FTIR spectrum, in which the infrared characteristic peak at 2283 and 2250 cm⁻¹ is clarified not from N₂O and assigned to HNCO. The principal products are H₂O and CO₂ together with significant amounts of HCl, HCN, NH₃ in MS, while few nitrogen oxides and O₂ are detected. The experimental results show that organic components have been the prominent media for the degradation of ClO₄⁻. To refine the mechanism observed in experiment, *ab initio* molecular dynamics simulations are carried out to reveal the atomistic reaction mechanisms. The decomposition of DAP starts with proton transfer from NH₄⁺ and H₂dabco²⁺ to ClO₄⁻. The deprotonated carbon skeleton is preferable to NH₃ in capturing O atoms, realizing a faster consumption of O atoms. Amounts of H atoms enter the environment being active free radicals, realizing an efficient autocatalytic chain propagation of degradation of ClO₄⁻. The atomistic thermal decomposition reaction mechanism of DAP uncovers the role of organic components in promoting the degradation of ClO₄⁻, which will help improve the synthesis strategy of molecular perovskite energetic materials with improved performance.

Received 15th December 2020

Accepted 21st April 2021

DOI: 10.1039/d0ra10559g

rsc.li/rsc-advances

1. Introduction

Ammonium perchlorate (AP), as a powerful oxidant, can produce a self-sustaining deflagration reaction when mixed with fuel, and is the most commonly used energetic component of solid composite propellants.¹ The deflagration performance of a propellant is directly related to the complex thermal decomposition mechanism of AP. People have been committed to enhancing the combustion rate by reducing the high-temperature decomposition (HTD) temperature and increasing the apparent heat release of AP.^{1–3} One of the most commonly used methods is nano-additives, whose catalytic

activity can effectively reduce the HTD temperature of AP. The two-stage thermal decomposition reaction of AP can be merged into one-stage one with appropriate amount of addition, so the heat release of AP is more concentrated. Among various nano-additives, carbon materials such as carbon nanotubes and graphene that can combust exothermically like a fuel with oxidizing gas products generated from the thermal decomposition of AP, realize a much higher apparent heat release (>3000 J g⁻¹).^{2,3} As the amount of addition is further increased, the thermal decomposition temperature can be further reduced, but the heat release is also reduced, indicating that the main heat source is still from AP. Therefore, although the combustion heat of carbon materials contributes to the higher heat release, the additional oxygen-consuming reaction, which greatly promote the thermal decomposition of AP, should be the main reason.

Another method is crystal engineering, which assembles AP molecules with appropriate ligand molecules at the molecular level to obtain the desired improvement. Cocrystal is a new crystal formed by combining two or more known neutral

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0ra10559g



molecules at a certain stoichiometric ratio through non-bonding interaction.⁴ Three cocrystals of AP and crown ethers – AP/18-crown-6, AP/benzo-18-crown-6 and AP/dibenzo-18-crown-6, have been synthesized.⁵ The last two cocrystals have a single obvious exothermic process at a lower temperature, with heat release increased to 1304.2 J g⁻¹ and 1488.4 J g⁻¹. Looking into the spatial configuration, the large and stretched ring skeletons of crown ethers are packed disorderly and thus the space of crystal is filled with many free volumes. Consequently, the crystal density is decreased far below that of AP and so is the corresponding energy density. Therefore, although the heat release of the cocrystal is increased relative to AP, which can be attributed to the catalytic effect of organic component and its exothermic combustion, the low crystal density limits the heat release to a low level.

Different from the direct “mixing” of cocrystal, molecular perovskites assemble molecules by introducing organic molecular unit into the perovskite architecture.⁶ (H₂dabco) [NH₄(ClO₄)₃] (DAP, dabco = 1,4-diazabicyclo[2.2.2]octane) introduces organic cation H₂dabco²⁺ as A component, with NH₄⁺ cation as B component and ClO₄⁻ anion as X component, to assemble a high symmetry ternary perovskite ABX₃ structure,^{7,8} which realize a homogeneous mixing of fuel and oxidant at the molecular level.⁹ The organic component H₂dabco²⁺ cation also has a ring skeleton but small and compact. And the electrostatic interaction between cation and anion is stronger than neutral molecules. Thus, the compact cubic crystal has a density close to AP, far above that of cocrystal. Furthermore, DAP is easier to be modified by replacing A, B or X component.^{10,11} Compared with AP, DAP has a lower thermal decomposition temperature and a single intense exothermic process, and its heat release has reached the level of nano-additives of carbon materials (>3000 J g⁻¹).^{12,13} These lay a good foundation for the further modification and application of DAP.^{14,15}

The catalytic performance of nano-additives depends on the effective contact area and the number of active sites with AP. Simply mixing their nanoparticles or making them porous nanocomposites, in essence, is physically mixed and subject to the strength of interaction and the uniformity of dispersion. The cocrystals of AP and crown ethers suffer from low crystal density, so more types of cocrystal need to be synthesized with more suitable ligand molecules. DAP seemingly integrate the advantages of all above methods. More importantly, it is evident that the organic components play a key role in the thermal decomposition of AP, which need to be further revealed by studying the thermal decomposition mechanism of DAP.

The structure changes of DAP have been studied by *in situ* FTIR in heating process.¹³ The cage skeleton constructed by NH₄⁺ and ClO₄⁻ ions decomposes before the organic moiety and determines the decomposition process. The combustion flames photographed with high speed photography and the gas release process recorded by TG-FTIR show a fierce one-stage energy-releasing effect of DAP.¹² However, the chemical reactions during the decomposition process of DAP still remain unclear. The infrared characteristic peaks of gas products in the FTIR spectra show significant changes compared with AP.¹² But the gas composition has not been identified, nor has the reason

been given for these changes. To examine the thermal decomposition reaction process of DAP, thermogravimetric analysis and differential scanning calorimetry coupled with Fourier transform infrared spectroscopy and mass spectroscopy (TG-DSC-FTIR-MS) experiments and *ab initio* molecular dynamics simulations are carried out to identify the evolved gas products. The concentration–time profile of gas products gives clues about the elementary reactions,¹⁶ while the *ab initio* molecular dynamics simulations trace the breaking and formation of bond on atomic scale.^{17–19} Theoretical calculations can well make up for the insufficiency of macroscopic experiments on the exploration of microscopic mechanism, while the corresponding results of macroscopic experiments verify the reliability of theoretical calculations.

2. Experimental and computational details

2.1 Experiments

2.1.1 Sample preparation. DAP samples are prepared according to the method proposed by Deng *et al.*¹² and purified to >99%.

2.1.2 Powder X-ray diffraction measurements. The Powder X-ray diffraction (PXRD) patterns of prepared samples are collected by a Philips X'Pert Pro X-ray diffractometer (PANalytical, Holland). The PXRD patterns of crystal model are simulated with the reflex module in Materials Studio 8.0.²⁰ Cu is selected as the target and the scanning range is 5–50° with a step-width of 0.02°.

2.1.3 TG-DSC-FTIR-MS measurements. Place ~2 mg of samples in an aluminum oxide crucible and load it on the Netzsch 449C TG-DSC synchronous thermal analyzer. Heat the samples from 30 to 500 °C at the heating rate of 10 °C min⁻¹. High-purity argon is used as the purge gas, and the evolved gas products are transferred into the coupled QMS403 quadrupole mass spectrometer and Nicolet 5700 infrared spectrometer at a flow rate of 25 ml min⁻¹. The FTIR spectral is recorded in the range of 4000–650 cm⁻¹ with a resolution of 4 cm⁻¹. The FTIR gas pool is maintained at 200 °C, while the tube connecting thermal analyzer, mass spectrometer and infrared spectrometer is maintained at 210 °C. The electron impact ionization energy of mass spectrometer is 70 eV. Operations of some characterization items can refer to the literature.^{21–25}

2.2 *Ab initio* calculations

The lattice parameters of DAP (CCDC No. 1528108) determined by Chen *et al.*⁹ are used to construct the crystal model. The spatial configuration and molecular composition of DAP crystal are shown in Fig. 1. DAP molecule is composed of five elements – C, H, O, N, and Cl. They are described by first-principles pseudopotentials constructed by projection augmented waves.^{26,27} C 2s2p, H 1 s, N 2s2p, O 2s2p, Cl 3s3p are considered as valence state. The exchange correlation functional adopts the generalized gradient estimation (GGA) in the form of Perdew, Burke, and Ernzerhof formulation (PBE).²⁸ The cutoff energy of the plane wave basis set is set to 600 eV. Grimme's DFT-D2 (ref.



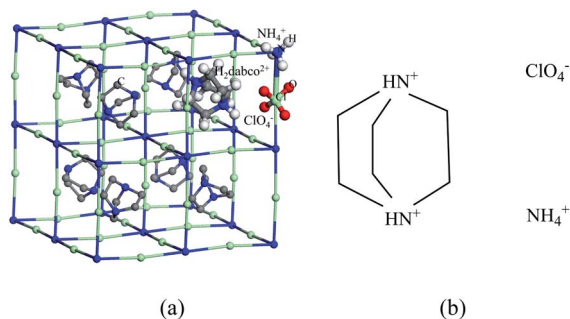


Fig. 1 Spatial configuration (a) and molecular composition (b) of DAP. C, H, N, O, Cl are represented by gray, white, blue, red, and green-yellow balls respectively. For clarity, only one ClO₄⁻ is shown with O atoms and one H₂dabco²⁺ and NH₄⁺ are shown with H atoms.

29) method is used to correct long-range dispersion interaction. Considering that the DAP unit cell contains 336 atoms, which is bigger than the supercell of common energetic material crystals used for *ab initio* molecular dynamics simulations,³⁰ the DAP unit cell is large enough to be used as the crystal model for *ab initio* molecular dynamics simulations with single Gamma point. In order to verify the applicability of the single Gamma point calculation to the DAP unit cell, the conjugate gradient method was used to perform complete structural relaxation (lattice and atom relaxed at the same time). The electronic iteration and ion relaxation convergence criteria are set to 1×10^{-7} and 1×10^{-6} eV. The optimized lattice parameters are in good agreement with the experimental values (see Table 1), which proves the reliability of the DAP calculation. Furthermore, the consistency of gas products obtained by experiment and calculation also verify the reliability of the calculation method. The Verlet algorithm is used for molecular dynamics integration with a time-step of 0.25 fs. The Nose-Hoover thermostat is used to realize a canonical ensemble (NVT). The optimized DAP unit cell is firstly relaxed at 300 K for 0.5 ps to eliminate the unreasonable internal stress. After that, the temperature is rapidly increased to 3000 K within 0.25 ps and the relaxed DAP unit cell is maintained at the temperature for 9 ps to be fully decomposed. In order to trace the evolution of product clusters, the cluster analysis method proposed by Lu *et al.*³⁰ is used. If the distance between two atoms is less than 1.33 times the normal covalent bond length³¹ and maintained for 25 fs (with hydrogen atoms bonded) or 50 fs (without

hydrogen atoms bonded),³² the two atoms are considered bonded.

For comparison, the decomposition of AP is also simulated. The lattice parameters of AP (CCDC No. 725655) determined by Kumar *et al.*³³ are used to construct the crystal model. Since the AP unit cell contains only 40 atoms, a $2 \times 1 \times 1$ supercell constructed from the optimized unit cell is used for the single Gamma point *ab initio* molecular dynamics simulation, while all other parameter settings are consistent with that of DAP.

3. Results and discussion

3.1 Experimental results

PXRD are often used for structural analysis. Pure substances have unique spatial configurations and crystal plane arrangements, which are depicted by certain diffraction peaks in PXRD figures like fingerprints. Fig. 2 shows the PXRD peaks of the prepared DAP samples and the DAP crystal model. The distribution of diffraction peaks of samples is consistent with that of crystal model, indicating that DAP has been successfully synthesized. The TG-DSC curves in Fig. 3 depicts the weight loss and energy release characteristics of DAP in the thermal decomposition process. The mass change of DAP samples starts at 292.7 °C and end at 402.6 °C with a weight loss of 94.7%. The weight loss curve has only one steep downward stage and so there is a single sharp exothermic peak on the DSC curve. The exothermic peak temperature is 381.8 °C. The thermal decomposition of DAP shows a rapid and concentrated one-stage heat release process different from that of AP.³⁴⁻³⁶

The infrared characteristic peaks in the FTIR spectrum correspond to certain characteristic vibrations of gas molecules which are fingerprints to identify the chemical species of gas products. By comparing the infrared characteristic peaks of gas products of DAP at 381.8 °C (Fig. S1(b)[†]) with the standard gas infrared data provided in HITRAN2016 molecular spectroscopic database (Fig. S1(a)[†]),³⁷ the gas components can be clearly identified to include CO₂, CO, H₂O, HCl, HCN, NH₃. It should be noted that high temperature will broaden the infrared characteristic peaks of some gases, but the central locations are almost unchanged (Fig. S2[†]). Since the temperature of FTIR gas cell is only maintained at 210 °C, for the infrared characteristic peaks not overlapped, the temperature-induced uncertainty can be ignored. More difficult is the identification of HNCO and N₂O, which are the key to understanding the thermal decomposition mechanism of DAP. N₂O is a symbolic gas product for

Table 1 Lattice parameters of DAP and AP

Lattice parameters	DAP			AP		
	Exp. ^a	Comp.	Relative error/%	Exp. ^b	Comp.	Relative error/%
<i>a</i> /Å	14.426	14.249	-1.23	9.176	8.779	-4.33
<i>b</i> /Å	—	—	—	5.797	5.976	3.09
<i>c</i> /Å	—	—	—	7.420	7.224	-2.64
ρ /(g cm ⁻³)	1.905	1.977	3.78	1.977	2.059	4.15

^a Cited from ref. 7. ^b Cited from ref. 24.



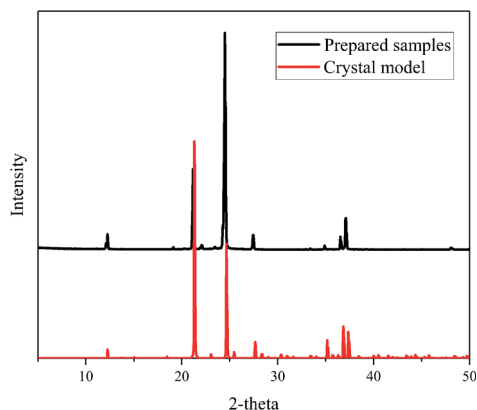


Fig. 2 PXRD of crystal model and prepared samples for DAP (red line for crystal model, black line for prepared samples).

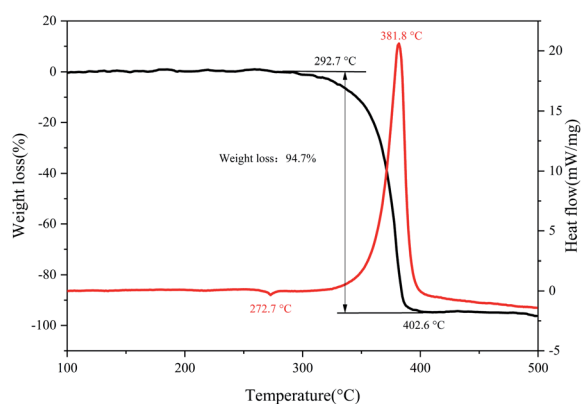


Fig. 3 Weight loss and heat release for DAP decomposed at the heating rate of $10\text{ }^{\circ}\text{C min}^{-1}$ (black line for TG, red line for DSC).

AP and has the strongest infrared characteristic peak at 2239 and 2204 cm^{-1} (Fig. S3[†]). For DAP, the infrared characteristic peak at 2283 and 2250 cm^{-1} (Fig. S1(b)[†]) is also the strongest and has a similar profile with that of N_2O . Considering that DAP contains the same component with AP, it is easy to jump to conclusions and identify it as N_2O . However, the latter has an obvious shift to high wavenumber region, which is more consistent with that of HNCO. At the same time, the second characteristic peak of N_2O at 1303 and 1269 cm^{-1} is absent, while that of HNCO appears at 3525 cm^{-1} . Furthermore, the infrared characteristic peaks of CO_2 and HNCO are partially overlapped in the range of $2200\text{--}2400\text{ cm}^{-1}$, which is consistent with the FTIR spectrum in.^{16,38,39} The infrared characteristic peaks of CO_2 and N_2O are clearly not overlapped even under high temperature (Fig. S4[†]), which is consistent with the FTIR spectrum in.^{16,39} Therefore, the infrared characteristic peak at 2283 and 2250 cm^{-1} is assigned to HNCO. The gas products in the thermal decomposition process of DAP include CO_2 , CO , H_2O , HCl , HCN , NH_3 , HNCO (Fig. 4).

The concentration–time profiles obtained from mass spectrum show the evolution of product fragments. In Fig. 5, the product fragments all concentrate at $381.8\text{ }^{\circ}\text{C}$, indicating

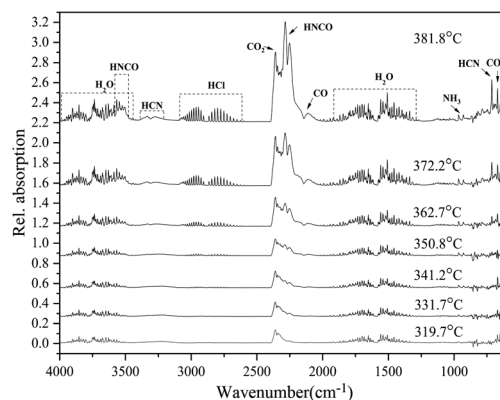


Fig. 4 Evolution of gas products from FTIR spectrum for DAP decomposed at the heating rate of $10\text{ }^{\circ}\text{C min}^{-1}$.

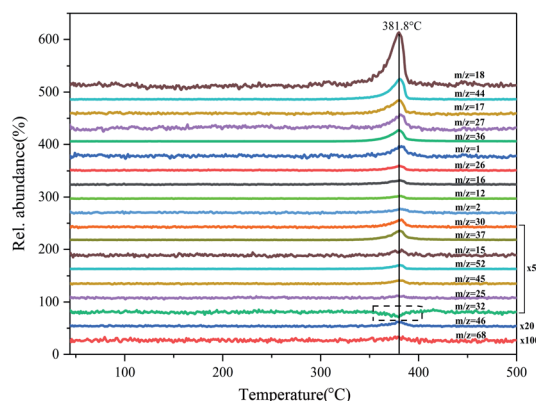


Fig. 5 Evolution of product fragments from MS for DAP decomposed at the heating rate of $10\text{ }^{\circ}\text{C min}^{-1}$.

a fierce thermal decomposition of DAP. It should be noted that the relative abundance of O_2 ($m/z = 32$) shows a negative increase at $381.8\text{ }^{\circ}\text{C}$. O_2 is a common gas product in the thermal decomposition process of AP.⁴⁰ The rupture of Cl–O bond will produce many oxygen free radicals like O, O_2 and O_3 . Their strong oxidation activity can oxidize NH_3 into various nitrogen oxides, which is a representation of oxidizing ability of AP. The negative increase of O_2 here indicates that there is no O_2 newly generated, but the O_2 from background gases is consumed by certain gas products. Considering that there is no fragments peak of CO ($m/z = 28$) in the mass spectrum, but the infrared characteristic peaks of CO do appear in the FTIR spectrum, it can be speculated that CO has been oxidized by O_2 in mass spectrometer. On the other hand, the absence of O_2 indicates that there may be few higher-order oxygen radicals generated in the thermal decomposition process of DAP. The single O radicals generated first from the rupture of Cl–O bond are quickly consumed, so there is not enough time for them to gather to higher-order oxygen radicals.

The chemical species of product fragments at $381.8\text{ }^{\circ}\text{C}$ are deduced from the mass-to-charge ratio (Table 2). Combined with FTIR spectrum, CO_2 , H_2O , HCl , HCN , NH_3 are confirmed to sole principal species, but there are no mass-to-charge ratios



Table 2 Assignment of product fragments for DAP decomposed at 381.8 °C

m/z	Ion current/ $\times 10^{-10}$ A	Species	m/z	Ion current/ $\times 10^{-10}$ A	Species
1	1.011	H	30	0.143	NO, CH ₂ O
2	0.337	H ₂	36	1.202	HCl
12	0.256	C	37	0.210	³⁷ Cl
15	0.070	NH, CH ₃	44	2.091	CO ₂ , N ₂ O
16	0.408	O, NH ₂ , CH ₄	45	0.07	HCO ₂ , HN ₂ O
17	1.359	OH, NH ₃	46	0.021	NO ₂
18	5.493	H ₂ O, NH ₄	52	0.083	HClO
25	0.036	C ₂ H	68	0.003	HClO ₂
26	0.492	CN			
27	1.206	HCN			

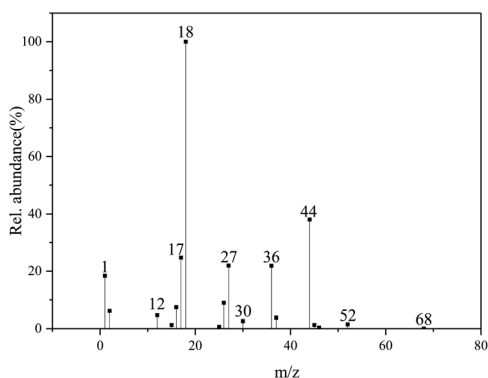


Fig. 6 Fragments of gas products from MS for DAP decomposed at 381.8 °C.

correspond to CO and HNCO. The absence of CO has been explained above. One possible reason for the absence of HNCO is that HNCO has been ionized to H₂NCO⁺ ($m/z = 44$) by adding H⁺, who will be covered by the fragments peak of CO₂, since there are many fragments of H⁺ in the mass spectrometer. In the molecular dynamics simulation, HNCO eventually exists in the form of HNCO \rightleftharpoons H₂NCO, indicating that it is easy for HNCO to capture another H. But high-precision mass spectrometry experiments are needed to confirm this. The relative abundance of primary product fragments is shown in Fig. 6. H₂O has the

largest abundance more than twice the second largest one of CO₂, indicating a large amount of O atoms decomposed from ClO₄⁻ have been captured by H atoms. In the decomposition of AP, H₂O is also a principal product.⁴⁰ The decomposition of AP starts with the proton transfer from NH₄⁺ to ClO₄⁻ and HClO₄ is degraded from high-order oxychloride or oxyacid to low-order ones, in which H atoms play an important role in the autocatalytic chain propagation.⁴¹ Since the cage skeleton constructed by NH₄⁺ and ClO₄⁻ ions determines the decomposition process of DAP,¹³ it can be speculated that the decomposition also start with the proton transfer from NH₄⁺ to ClO₄⁻. The presence of HClO and HClO₂ indicates that the degradation of HClO₄ is also an autocatalytic chain propagation from high-order oxychloride or oxyacid to low-order ones. More abundant H₂O origins from the organic components and higher proportion of ClO₄⁻. N₂O has been identified not included in the gas products by FTIR spectrum. Although there are fragments peaks with $m/z = 30$ and 46, which possibly belong to NO and NO₂ respectively, their abundance is too small. Therefore, nitrogen oxides, which are the main oxygen-consuming products in the decomposition of AP,^{34,35} have been replaced by carbon oxides. NH₃, HCN and HCl have close and significant relative abundance. NH₃ comes from the deprotonated NH₄⁺. HCl is a symbolic product for the degradation of ClO₄⁻. And HCN comes from the polyazine unit in the carbon-nitrogen heterocycle of H₂dabco²⁺, which is a common decomposition product of carbon-nitrogen

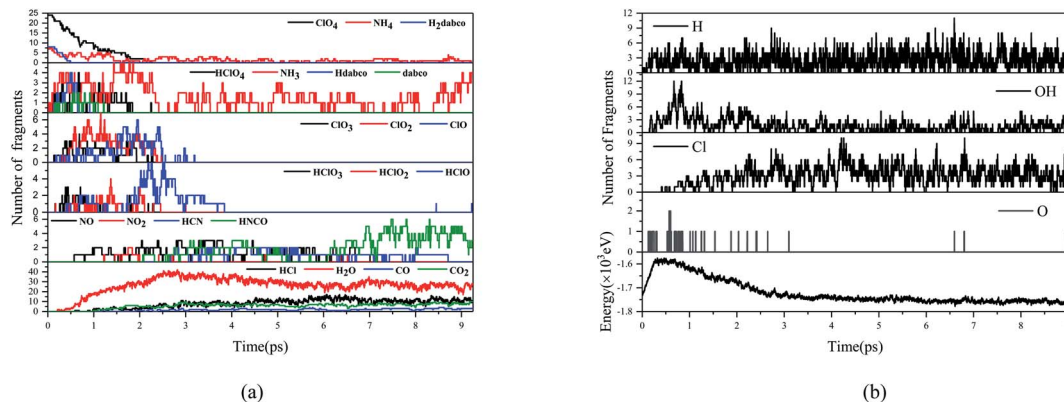


Fig. 7 Evolution of key chemical species (a), free radicals and potential energy (b) for DAP decomposed at 3000 K.



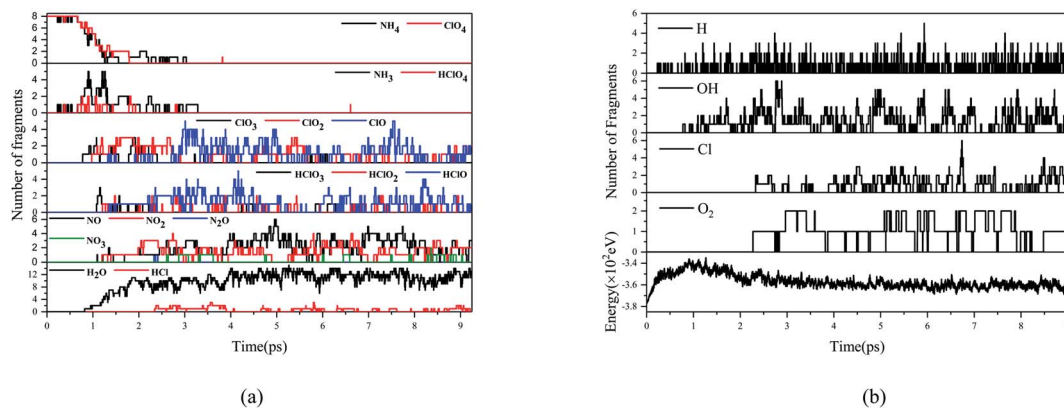


Fig. 8 Evolution of key chemical species (a), free radicals and potential energy (b) for AP decomposed at 3000 K.

heterocyclic molecules like RDX, HMX and CL-20.^{16,42} HNCO often accompanies with HCN, which is derived from the amide unit after the carbon skeleton is oxidized by O atoms. Another important gas product generated from the decomposition of carbon-nitrogen heterocycle is CH₂O. But it cannot be identified in the FTIR spectrum. The relative abundance of fragments with $m/z = 30$ is also much smaller than that of CO₂. This indicates that H₂dabco²⁺ may lose H atoms first before the ring skeleton breaks or is oxidized by O atoms. So H₂dabco²⁺ is more stable than its precursor of dabco under the confinement of cage skeleton.

In the thermal decomposition process of ClO₄⁻, the highly reactive transient radicals ClO_x ($x = 1-3$) can only survive for a very short time and few of them can escape in the form of gas. Therefore, they can hardly be detected by FTIR and quadrupole electron impact MS (the MS used in the paper).

In the paper,³⁶ the author summarized the detection results of various FTIR and MS methods on the thermal decomposition products of AP. Only the state-of-the-art single vacuum ultraviolet (VUV) photoionization coupled with reflectron time-of-flight mass spectrometry (PI-ReTOF-MS) can detect all ClO_x ($x = 1-3$) radicals, though the intensity of signal is much weaker than that of primary gas products. The FTIR and quadrupole electron impact MS can hardly detect the ClO_x ($x = 1-3$) radicals.

In the following simulations, the reason why ClO_x ($x = 1-3$) seem to have good half-lives is that they are transformed continuously from ClO₄⁻. When ClO₄⁻ is degraded completely, ClO_x ($x = 1-3$) disappear almost at the same time. Therefore, a single ClO_x can only survive for a very short time. Only HClO exists for a certain time after all ClO_x ($x = 1-4$) and HClO_x ($x = 2-4$) disappear, so it has a relatively good half-life. This is why for the only two HClO_x detected by MS, the intensity of signal of HClO is much stronger than that of HClO₂, but is still much weaker than that of primary gas products. This is consistent with detection results of AP, in which only HClO can be detected even with PI-ReTOF-MS.

3.2 *Ab initio* calculation results

Fig. 7 and 8 show the evolution of prominent atomic clusters and potential energy respectively for DAP and AP decomposed

at 3000 K. The chemical species for AP include NH₃, HClO_x ($x = 1,2,3,4$), ClO_x ($x = 1,2,3$), NO_x ($x = 1,2,3$), N₂O, H₂O, HCl, while for DAP including NH₃, hdabco, dabco, HClO_x ($x = 1,2,3,4$), ClO_x ($x = 1,2,3$), NO_x ($x = 1,2$), HCN, HNCO, H₂O, HCl, CO_x ($x = 1,2$). The products composition and content are consistent with the thermal decomposition experiments of AP and DAP, which verifies the reliability of *ab initio* molecular dynamics simulations. The presence of CO and HNCO verifies the experimental inference. When the temperature is heated to 3000 K, DAP decomposes quickly, while AP does not until 0.662 ps. AP has been heated to its HTD reaction, whose onset temperature is higher than that of DAP. The first appearance of NH₃ and HClO₄ indicates that the thermal decomposition of AP and DAP both begin with the proton transfer from NH₄⁺ to ClO₄⁻. For DAP, the cage skeleton collapses soon after the proton transfer and loses its confinement on h₂dabco²⁺. h₂dabco²⁺ is deprotonated to Hdabco⁺ and dabco almost simultaneously with the proton transfer, so H₂dabco²⁺ is also an important proton donor. It is worth mentioning that H₂dabco²⁺ seemingly prefers to lose the C-terminal H atoms instead of N-terminal H atoms, since there are more HNCO than HCN in the products. In fact, it's easy for H atoms to escape short-lived from X-H bond under thermal stimulation due to its minor mass. When losing H atoms, the exposed C atoms are easier to be oxidized by O atoms than N atoms, which produces -CO- groups instead of -NO- groups on the ring skeleton. Therefore, the C atoms have superiority over N atoms in the competition for O atoms. Although the decomposition starts from between NH₄⁺ and ClO₄⁻, only few nitrogen oxides generated and a significant amount of NH₃ exit till the end of reaction, which is opposite to that of AP. The degradation of HClO₄, which gradually changes from higher-order oxychloride or oxyacid to low-order ones, is a process of losing O atoms. When the cage skeleton collapsed, H₂dabco²⁺ loses H atoms first. Part of H atoms transfer to ClO₄⁻, while most of them enter the environment. The H atoms in the environment is maintained at a high level, which realize a more efficient autocatalytic chain propagation process, that is, a faster consumption of O atoms. This can be verified by the evolution of oxygen radicals. There are mainly single O radicals for DAP but O₂ radicals for AP in the environment, indicating that the O atoms are consumed much faster in the thermal



decomposition process of DAP, so there is not enough time for them to gather to form higher-order oxygen radicals. The first appearance of H₂O means that O atoms are first captured by H atoms. The number of OH radicals rise rapidly until the appearance of Cl atoms, which will be captured by H atoms to produce HCl. From the perspective of energy release, the inflection point of potential energy of AP lags behind and there is a second-stage energy release process when Cl atoms appear. On the one hand, NH₃ is relatively inert and it requires more energy to break N–H bond; on the other hand, breaking the last Cl–O bond to produce Cl atoms also requires more energy. However, it is much easier for C–H bond to break in DAP, and the rapid oxygen-consuming process accumulates enough heat to make the last Cl–O bond break earlier. Therefore, the appearance of Cl atoms is almost at the inflection point and the potential energy release is more concentrated and smoothed. The main energy release of DAP is finished when oxychloride or oxyacid are degraded completely (approach 3 ps), but the energy release of AP is insufficient due to the incomplete degradation of HClO₄.

4. Conclusions

In summary, the thermal decomposition of DAP starts with proton transfer from NH₄⁺ and H₂dabco²⁺ to ClO₄⁻, however, the degradation pathway of ClO₄⁻ has been changed by organic components compared with that of AP. Firstly, the deprotonated carbon skeleton is preferable to NH₃ in capturing O atoms, which realizes a faster consumption of O atoms. Secondly, amounts of H atoms released from H₂dabco²⁺ enter the environment being active free radicals, realizing an efficient autocatalytic chain propagation of degradation of ClO₄⁻. Therefore, the generation of gas products is fast and concentrated, and so is the energy release. Molecular perovskites have provided a novel assembly strategy for energetic molecules that can achieve high energy release.

Conflicts of interest

There are no conflicts of interest to declare.

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