

Cite this: *Mater. Adv.*, 2021,
2, 2021

Structure-controlled growth of vertically-aligned carbon nanotube forests using iron–nickel bimetallic catalysts†

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Although the progress towards catalyst engineering and growth parameters to tune the properties of carbon nanotubes have been widely studied, critical properties like chirality still need addressing in the case of bimetallic catalysts. Herein, we synthesized six different catalysts *via* solid-state reactions using a mechanical alloying (MA) method to investigate the role that these catalysts have in the growth of high-quality vertically aligned carbon nanotubes (VACNTs). The growth of the VACNTs was performed using a chemical vapor deposition method onto silicon wafers. Focusing on the impact of Ni addition into the Fe matrix, we explored the diameter distribution and chirality profiles of the VACNTs. The conductive band branches, the overall mapped morphology, chemical purity, structural quality, and physical alignment of the VACNTs were analysed in detail using Raman micro spectroscopy, thermogravimetric analysis, and scanning electron and transmission electron microscopies. The relative radial breathing modes were explored for the (12,3), (9,3), (9,6), and (10,4) chiralities, with different counts for overall samples. Furthermore, the (11,9) chirality index, which is the most encountered type by far, was observed in the synthesis of VACNTs on the Fe50Ni catalyst, with a relatively high count of 28%. The current Fe–Ni bimetallic catalyst study thus employs Raman spectroscopy to determine the role that Ni addition into the Fe matrix has in the possible difference in the conductive branches of the VACNTs.

Received 23rd October 2020,
Accepted 9th February 2021

DOI: 10.1039/d0ma00826e

rsc.li/materials-advances

1. Introduction

Vertically aligned carbon nanotubes (VACNTs) have received much attention in the few last decades due to their use in many applications, including polymer nanocomposites (PNCs),^{1,2} nanoporous membranes,^{3,4} supercapacitor electrodes^{5,6} *etc.* However, the use of carbon nanotubes (CNTs) in demanding applications such as electronics, supercapacitors and/or sensors is still limited due to their low quality in batch production with a narrow diameter and non-selective chirality growth.⁷ Even though VACNTs can be easily synthesized using monometallic transition metals such as Fe, Ni, Co, Mo, Cu *etc.*, which have

unique properties related to their 3d-orbitals, it is still a challenge to synthesize fully aligned CNTs with few defects at a high density that are homogeneous and show controlled chirality specifically for electronic applications.^{8–12} Recent studies have been reported that bimetallic catalysts have the potential for size control, quality enhancement and chirality distribution compared to monometallic transition metals since these mixed metal catalysts have a combined effect on the overall properties besides characterized by complex relations such as catalyst particle size and combination.^{13–16}

To date, controlling the diameter of CNTs with various methods has been proposed through changing the catalyst particle types (monometallic or bimetallic), modifying the carbon source (CH₄, C₂H₂, C₂H₄, *etc.*), varying the feedstock flux, support materials (*e.g.* SiO₂, Al₂O₃, MgO, *etc.*), surface treatments (*e.g.* chemical etching, ion bombardment, *etc.*) reaction time and growth temperature.^{17–19} However, with a large number of variables, it is still a hurdle to achieve control over the tube diameter, which has a direct influence on the chiral distribution of the material.²⁰ Recently, the most common and highest yielding growth method for VACNTs has been reported to be catalytic chemical vapor deposition (CCVD), which can be used to synthesize CNTs *via* a well-accepted

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0ma00826e



vapor-liquid-solid (VLS) growth mechanism.^{21,22} According to the VLS mechanism, the catalyst nanoparticles can be viewed as a seed for the nucleation and growth of CNTs, where the diameter, areal density and number of walls firmly depend on the size of the catalyst nanoparticles.^{23,24} However, Liu *et al.* reported that as a monometallic catalyst, Fe nanoparticles tend to diffuse into the substrate at a high reaction temperature, which limits the achievement of a uniform diameter distribution of the VACNTs despite its high catalytic efficiency.²⁵ Additionally, many researchers have also reported that bimetallic catalyst nanoparticles such as FeMo,²⁶ FeNi,^{27–29} and FeCu³⁰ improve the reducibility of Fe and inhibit the aggregation of Fe clusters, different from monometallic catalysts. Li *et al.* achieved controlled diameter distribution, wall number and thermal stability of VACNTs grown on a FePt catalyst compared to VACNTs grown from an Fe catalyst.³¹

Aside from diameter control, quality enhancement is another critical issue for CNTs for a wide range of applications.³² Apart from the mobility of the carbon source increasing with temperature, the CNT growth rate is also proportional to the reaction temperature owing to the high face-centered cubic phase of Fe (γ_{Fe}).³³ Lee *et al.* determined that both lateral alignment and crystallinity of CNTs are extremely dependent on the growth temperature using Raman spectroscopy.³⁴ However, when the optimal temperature is reached, the CNT growth rate decreases due to the accumulation of amorphous carbon on the catalyst surface and deactivates the catalyst particles which reduce the CNT quality in the presence of the amorphous carbon.^{35,36} Hence, Hata *et al.* developed a new technique for the super growth of CNTs and they significantly increased the catalytic activity by removing amorphous carbon through the addition of water vapor.³⁷ Moreover, Amama *et al.* reported that water vapor hinders the diffusion and Ostwald ripening of catalyst particles through the mitigation of oxygen and hydroxyl groups.³⁸ Consequently, several researchers focused on Co–Cu, Fe–Mo, Fe–Ni bimetallic catalyst particles to inhibit Ostwald ripening and reported that the synthesis of CNTs with low defect levels and enhanced crystallinity was achieved.^{26,32,39}

The main bottleneck in the development of CNT-based technologies such as electronics and optics is the precise control of the chirality of CNTs that have high-crystallinity, a low-defect density and low diameter.⁴⁰ As such, huge efforts have been dedicated towards chirality controlled synthesis in the last few years, including optimizing growth conditions such as the growth temperature, the carbon source type, support materials, surface treatments and the design of the catalyst nanoparticles.^{41–43} Many researchers have focused on understanding how the catalyst composition and the structure affect the chirality distribution of CNTs.⁴⁴ Despite several studies being reported, establishing a link between the catalyst properties and CNT structures still remains elusive owing to the complex relationship between them.^{45,46} Thus, density functional theory (DFT) calculations have led to the investigation of the epitaxial relationships between bimetallic catalyst nanoparticles such as Ni_xFe_{1–x}, FeGa, NiGa, and Pd–Co and nanotube cap formation.^{47–49} However, DFT calculations are carried out at zero

temperature, so they only provide a thermodynamic overview of nanotube nucleation and growth.⁴⁹ Rao *et al.* determined that kinetic aspects are important in assigning the chirality of CNTs using *in situ* Raman spectroscopy.⁵⁰ Thus, experimental studies are important for determining the selective-chirality of CNTs combined with molecular dynamics simulations.³⁹

In this study, through the use of six distinct catalyst types based on Fe–Ni combinations with different weight fractions, VACNT synthesis *via* CCVD was investigated using Raman spectroscopy, thermogravimetric analysis (TGA), scanning electron microscopy (SEM) and transmission electron microscopy (TEM). The main goal of this study is to determine the effects that catalyst composition have on several parameters, including the chirality index, diameter, quality, purity, and morphology of the VACNTs. Fe–Ni based catalysts with different weight fractions were synthesized using a mechanical alloying (MA) method. These catalysts were then used to coat silicon (Si) wafers using the E-beam method. The CCVD parameters were kept the same for all of the catalyst types used in this study to compare the effect of the catalysts. Subsequently, the synthesized VACNTs were characterized by Raman spectroscopy. A Lorentzian area function was utilized to fit the peaks of the radial breathing mode (RBM) for each sample and chirality indexes of (*n,m*) were assigned from Kataura plots. Furthermore, the morphological and oxidation behavior of the CNTs were evaluated using SEM, TEM and TGA. We revealed that the counts of the special chirality species such as (11,9), (9,3), (9,6) (12,3) and (10,4) significantly change upon the addition of Ni into the Fe matrix. In this way, this study paves the way for the selective growth of VACNTs depending on their chiral index in future studies.

2. Experimental study

2.1 Preparation of the bimetallic catalysts

MA was performed to prepare Fe–Ni powder alloy catalysts from high purity powders (Fe 99.9% purity, Ni 99.5% from the ABCR Corporation) through varying the Ni amounts in the catalysts to 5, 10, 15, 20 and 50 wt%. Powder blends were milled for 8 h using a Spex 8000 D ball mill at a rotational speed of 1200 rpm. A ball-to-powder weight ratio (BPR) of 10 : 1 using stainless steel balls (diameter 6 mm) in a stainless-steel box was employed for MA. Milling vials were placed into a glove box (PlasmaLab TM) under an Ar atmosphere (Linde TM 99.999% purity) to protect oxidation of the powders at high speed. All of the powder alloy catalysts were annealed at 600 °C for 2 h to remove the residual stress that arises from the MA process.

2.2 Synthesis of the VACNTs by CCVD

VACNTs were synthesized using the CCVD method at 750 °C on highly n-doped Si wafer substrates coated with Al₂O₃/catalyst using an E-beam. The Si substrate was first coated with a 300 nm oxide layer *via* plasma-enhanced chemical vapor deposition (Oxford Instrument Plasma Lab System 100) and then E-beam evaporation was performed to obtain a uniform Al₂O₃ thin film of 10 nm in thickness as a buffer layer at a rate



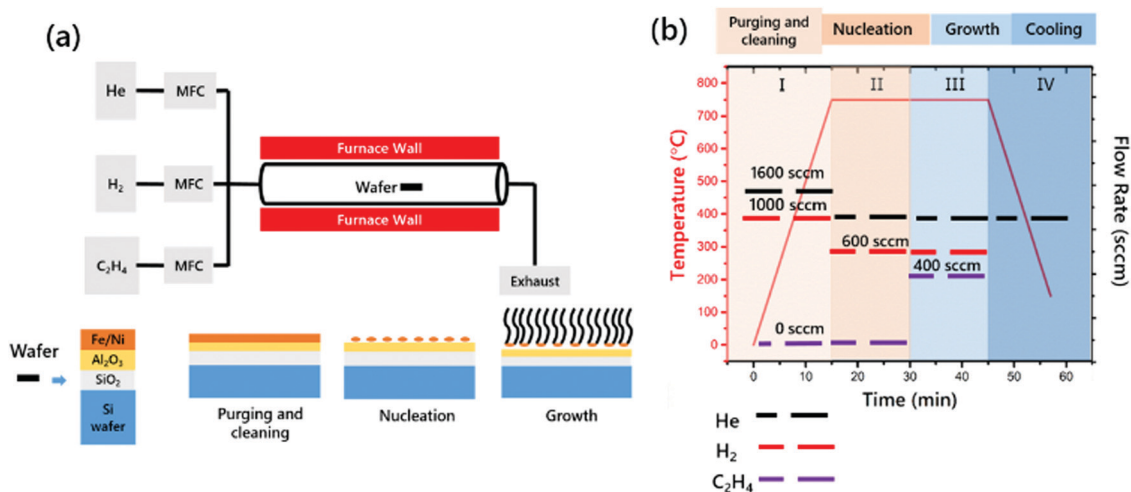


Fig. 1 (a) Schematic of the CCVD system and growth mechanism of VACNTs onto Si wafers and (b) the growth protocol for the preparation of VACNTs in a quartz tube furnace using He, H₂ and C₂H₄ as gas sources.

of 1 \AA s^{-1} and a synthesized binary catalyst with a thickness of 2 nm at $2\text{--}4 \times 10^{-6}$ torr (Torr Instrument) at a rate of 0.1 \AA s^{-1} . A single zone quartz tube furnace was used for the CCVD process (Lindberg/Blue M) equipped with an electronic mass flow controller (Alicat Inc.) (Fig. 1(a)). CNT growth was performed under atmospheric pressure using the reactant gases C₂H₄, H₂ and He. The Si-wafer samples were placed into the hot zone of the furnace, as shown in Fig. 1(a). After purging and cleaning of the quartz tubes using He and H₂, the furnace was heated to 750 °C with 1600 sccm He and 1000 sccm H₂ for 15 min to nucleate the catalyst particles. Then, 1000 sccm He, 600 sccm H₂ and 400 sccm C₂H₄ were introduced into the quartz tubes for 15 min. The aforementioned four main steps are summarized in Fig. 1(b), showing the growth conditions such as temperature, time and gas flow rate.

2.3 Characterization of the VACNTs

VACNTs were characterized using SEM (FEI QUANTA) with an acceleration voltage in the range of 5–15 kV under high vacuum. Also, TEM analysis was performed using a JEOL ARM-200 CFEG microscope operated at 200 kV equipped with a Cs probe collector (JEOL, Tokyo) to measure the tube diameter and number of tube walls in the VACNTs. To prepare the TEM samples, VACNTs were immersed in isopropyl alcohol and ultrasonicated for 5 h. Then, a drop of sample was placed on a carbon coated 300 μm mesh size Cu grid. The grid was dried at room temperature and cleaned using plasma for 10 seconds before loading on the microscope. Raman spectroscopy was performed using a Renishaw inVia reflex microscope and spectrometer with an excitation energy of 2.33 eV over an acquisition range of $100\text{--}3200 \text{ cm}^{-1}$, coupled with an ultraviolet (UV) and near-infrared (NIR) enhanced deep depletion, Peltier cooled CCD array detector (1024×256 pixels) and a Leica microscope equipped with 5×, 20× and 50× magnification objectives. A computer controlled XYZ stage with a scan area of 40×40 mm (X and Y) and a step size of 1 mm allowed automated Raman

spectral mapping measurements (laser power $\sim 1 \text{ mW}$ in 1 \mu m^2) to be carried out. TGA was performed using an SDT Q600 TGA thermogravimetric analyzer at a heating rate of 10 °C min^{-1} up to 800 °C under a 100 ml min^{-1} airflow.

3. Results and discussion

3.1 Characterization of the Fe–Ni catalysts

The phase components, crystal structures and directions of the planes were detected by X-ray diffractometry (XRD) using a Bruker™ D8 advance diffractometer (CuK α radiation, $\lambda = 1.542 \text{ \AA}$). As shown in Fig. S1 (ESI[†]), the peaks marked with \blacklozenge , \clubsuit , \spadesuit at 44.8° , 66.1° , and 82.4° can be indexed to the (110), (200) and (211) planes of BCC Fe–Ni. A small peak marked with \bullet at 51.1° attributed to the (200) FCC plane of the Fe–Ni alloy was present in the XRD pattern of Fe₂₀Ni. Ni is a well-known austenite phase stabilizer used in the steel industry.⁵¹ A detailed investigation showed that the presence of Ni in an amount higher than 18.4 wt% stabilized the FCC-Fe content of the material into a BCC-Fe structure after the annealing process.⁵² Moreover, increasing the Ni content results in an increase in the stabilized FCC-Fe content. The XRD pattern of Fe₅₀Ni consists of diffraction peaks marked with \blacklozenge at 43.8° , \bullet at 51.1° and $*$ at 75.2° , which can be attributed to an Fe_{0.5}Ni_{0.5} alloy with a FCC crystalline structure.⁵³ The main diffraction peak of the FCC Fe–Ni alloy at the 2θ position of 43.8° , which cannot be detected in the XRD pattern of Fe₂₀Ni, may be caused by overlapping with the dominant peak of the BCC-Fe structure at 44.8° . Moreover, there was no trace of the peaks belonging to the BCC-Fe crystalline phase in the XRD pattern of Fe₅₀Ni, which can be attributed to the existence of the Fe₅₀Ni catalyst only in a FCC structure. In addition to that, when the Fe–C phase diagram was considered, the carbon solubility of the FCC structure was found to be higher than that of the BCC structure. Hence, by increasing the amount of Ni in Fe–Ni alloy, the carbon solubility in the FCC structure improves.¹²



3.2 The effects of Ni addition on the chirality and diameter of the VACNTs

Identifying the effects that the Ni addition in the catalyst alloy has on the chirality and diameter of the VACNTs was studied by Raman spectroscopy from the RBM peaks between 120 and 300 cm^{-1} . Using an empirical Kataura plot,⁵⁵ these RBM peaks were correlated with the chirality of the VACNTs. The diameter of the VACNTs was calculated as $\omega_{\text{RBM}} = 227/d_t$,³⁹ where ω is the Raman shift of the RBM peaks at 121 different points and the RBM spectra of the VACNTs were fitted using a Lorentzian area function. For each fitting, the R -squares adj. values were higher than 0.95. The positions of the RBM peaks are listed in Table 1 according to their RBM peak positions, assignment of the chirality, diameters, and type of conductivity, respectively (see more detailed information in Table S1 in the ESI,[†] which also gives more information about the equations).

Fig. 2 presents the RBM region of the spectrum at 2.33 eV with metallic or semiconducting branches. For quantitative evaluation of each peak depending on the diameter and chirality, the integrated intensities of the RBMs were employed in the Raman spectra.^{58,59}

Table 1 RBM shifts of the synthesized VACNTs according to their Fe and Fe–Ni alloy combinations for green laser spots (2.33 eV, 532 nm) used in all the calculations (ω_{RBM} : mean value from 121 spectra)^{40,56,57}

Catalyst type	ω_{RBM} (cm^{-1})	(n,m)	d (nm)	S or M
Fe-Only	165	(11,9)	1.37	S
	187	(12,5)	1.20	S
	208	(12,3)	1.09	M
	220	(9,6)	1.03	M
	229	(10,4)	0.99	M
	267	(9,3)	0.85	M
Fe5Ni	131	(15,10)	1.73	S
	162	(16,3)	1.40	S
	182	(10,8)	1.24	S
	208	(12,3)	1.09	M
	220	(9,6)	1.03	M
	229	(10,4)	0.99	M
	267	(9,3)	0.85	M
	267	(9,3)	0.85	M
Fe10Ni	135	(16,8)	1.68	S
	164	(11,9)	1.37	S
	185	(15,1)	1.23	S
	208	(12,3)	1.09	M
	220	(9,6)	1.03	M
	230	(10,4)	0.99	M
	266	(9,3)	0.85	M
	267	(9,3)	0.85	M
	267	(9,3)	0.85	M
Fe15Ni	162	(16,3)	1.40	S
	187	(12,5)	1.20	S
	208	(12,3)	1.09	M
	220	(9,6)	1.03	M
	229	(10,4)	0.99	M
	267	(9,3)	0.85	M
	267	(9,3)	0.85	M
Fe20Ni	157	(11,10)	1.44	S
	188	(12,5)	1.20	S
	210	(12,3)	1.09	M
	221	(9,6)	1.03	M
	231	(10,4)	0.98	M
	267	(9,3)	0.85	M
	286	(10,0)	0.79	S
Fe50Ni	165	(11,9)	1.37	S
	220	(9,6)	1.03	M
	229	(10,4)	0.98	M
	267	(9,3)	0.85	M

As clearly presented, the diameter distribution is wide, ranging from 0.79 to 1.73 nm, and the most common species are at 1.20, 1.09, 1.03 and 0.85 nm for the six samples. The mean diameters of the CNTs are close and in between 1.01 and 1.20 nm. The RBM peaks were observed at ~ 131 and 135 cm^{-1} for the VACNTs synthesized using Fe5Ni and Fe10Ni, respectively. These peaks were presented from the S_{44} branch in the Kataura plot and they were not seen in other samples. However, as the branches S_{33} and S_{44} overlap at $\lambda_{\text{laser}} = 532 \text{ nm}$, utilizing a higher laser power is more convenient to detect larger tube diameters, as also indicated in the enlarged Kataura plot.⁶⁰ An RBM peak was observed at 165 cm^{-1} for VACNTs synthesized using Pure-Fe, Fe10Ni, and Fe50Ni. Fig. 2(a), (b) and (f) reveal that these VACNTs feature (11,9) as the most abundant chiral species, with concentrations of 14%, 14% and 28%, respectively.

This value was calculated as approximately 17% for the (11,9) chirality in previous reports.⁵⁸ Additionally, for the Pure-Fe, Fe15Ni, and Fe50Ni catalysts, the VACNTs presented a peak at 187 cm^{-1} . In the S_{33} branch, only this peak was detected for VACNTs synthesized using Fe50Ni. However, the most intense peaks were observed at 208, 220, 229 and 267 cm^{-1} for all samples in the M_{11} branch. Cheng *et al.* reported the same RBM peaks for bundled HiPco SWCNTs at an excitation energy of 2.33 eV.⁶¹ According to the RBM in the spectrum of the VACNTs synthesized using Fe10Ni, it was clearly seen that semiconducting VACNTs in 55% were grown compared to the other samples, meaning that they could have important potential use as high-performance field effect transistors after enrichment processes.^{62–64}

Exploring the structural quality of the large-scale VACNTs, Raman spectral mappings were recorded for all of the samples, on $40 \times 40 \mu\text{m}^2$ scan areas, with a step size of $5 \mu\text{m}$. Raman mappings of VACNTs synthesized on Pure-Fe, Fe5Ni, Fe10Ni, Fe15Ni, Fe20Ni, and Fe50Ni are presented in Fig. 3. The G-band, observed at 1600 cm^{-1} , is the first-order Raman scattering band, related to the degree of graphitization, while the defect-induced D-band provides information about the side walls of the CNTs.^{65,66} The quality of a sample is often identified from the intensity of the G- to D-band ratio.⁶⁷ The mappings of VACNTs synthesized on different catalyst particles are shown in Fig. 3, using different colors for each I_G/I_D ratio. The means of the I_G/I_D ratios of VACNTs synthesized on Pure-Fe, Fe5Ni, Fe10Ni, Fe15Ni, Fe20Ni, and Fe50Ni are approximately 2.04, 2.09, 2.68, 2.07, 1.75 and 1.73, respectively. The maximum I_G/I_D ratio was obtained for the VACNTs synthesized on Fe10Ni. Although this value is higher than the others, it does not give information on the types of the defects in the sample because perfect zigzag edges, charge impurities, intercalants, uniaxial and biaxial strain do not generate a D-band peak.⁶⁸

Cancado *et al.* calculated the defect density in a graphene structure using Raman spectroscopy, employing different laser powers.⁶⁸ They presented the relationships between I_D/I_G and the distance between the defects (L_D) depending on the excitation laser wavelength λ_L (nm), using the simple equation shown in eqn (1). They also reported that this empirical relationship can be



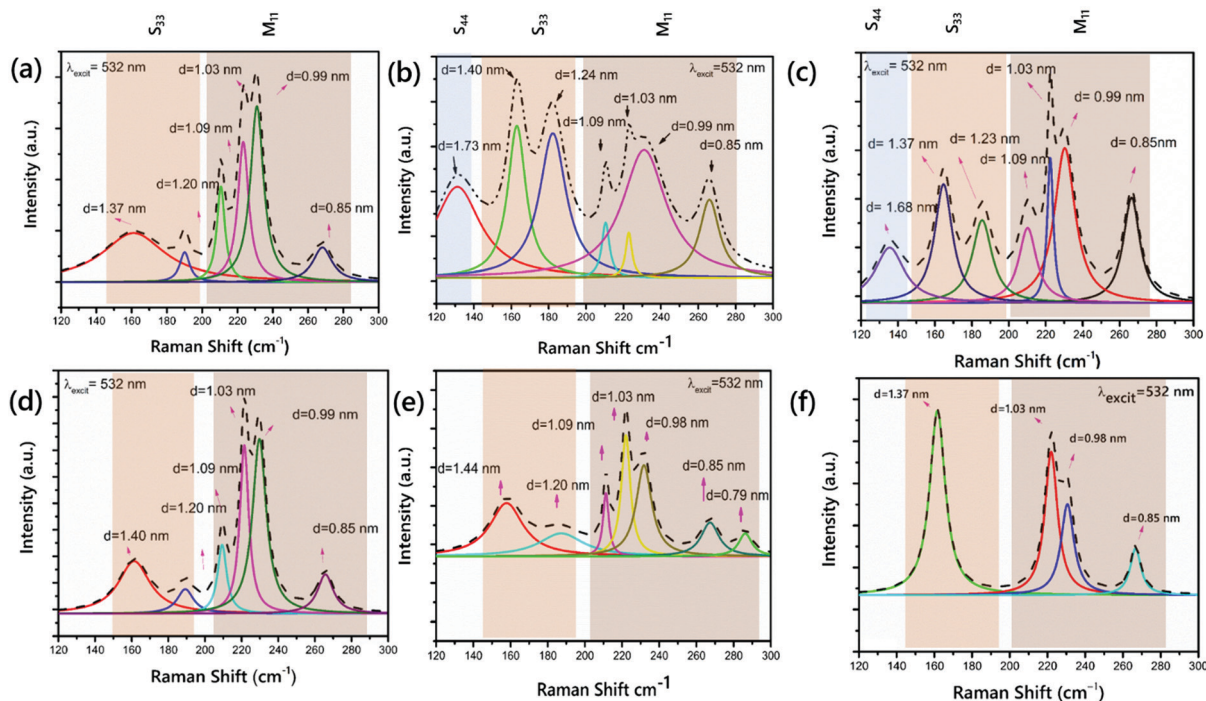


Fig. 2 RBM peaks fitted to a Lorentzian area function and calculation of the innermost diameters of the VACNTs synthesized on (a) Pure-Fe nanoparticles, (b) Fe5Ni, (c) Fe10Ni, (d) Fe15Ni, (e) Fe20Ni, and (f) Fe50Ni ($E_{\text{Laser}} = 2.33$ eV).

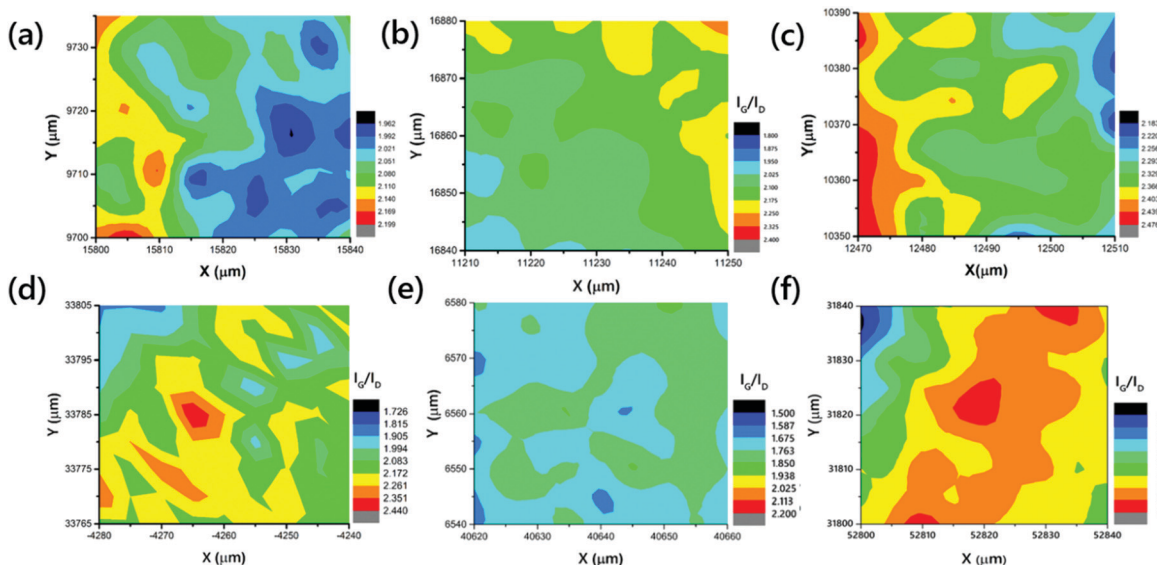


Fig. 3 Raman mappings of the VACNTs synthesized on (a) Pure-Fe, (b) Fe5Ni, (c) Fe10Ni, (d) Fe15Ni, (e) Fe20Ni, and (f) Fe50Ni.

used for $L_D > 10$ nm in Raman spectroscopy for any laser line in the visible range.⁶⁸

$$L_D^2 \text{ (nm}^2\text{)} = (1.8 \pm 0.5) \times 10^{-9} \lambda_L^4 \left(\frac{I_D}{I_G} \right)^{-1} \quad (1)$$

According to the calculations of L_D for all of the samples, the defect density of the VACNTs synthesized in this study using Fe10Ni is lower than other catalyst combinations as the distance between

the defects is greater, as presented in Fig. 4. In this calculation, 121 different I_G/I_D ratios were input into eqn (1), that correspond to Raman active defects.⁶⁸ It is critical to note here that silent defects such as perfect zigzags, intercalants, and impurities were not considered as they are correlated to G-band and 2D peaks.^{69,70}

3.3 Defining the chiral indexes of VACNTs with various catalysts

There are two main strategies for evaluating the chirality index (n, m) of CNTs quantitatively. First, the counts for each chiral



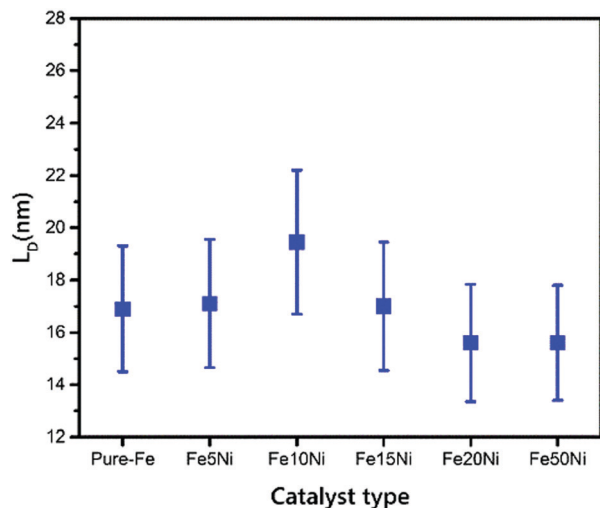


Fig. 4 Distance between defects with error bars for VACNTs synthesized on different catalyst particles.

index in a sample are assumed to be proportional to their integral RBM intensities.⁵⁸ This aspect is commonly used for bulk samples (e.g. thin films, arrays, forests *etc.*^{63,71,72}). The second strategy is to count the numbers of RBM peaks in the Raman spectral mapping.⁷³ In this study, the ratios of RBM intensities for all RBM peaks were compared to each other. The root of the chirality index can be calculated from the diameter (d_t), found using $d_t = \sqrt{3a_{c-c}/\pi\sqrt{n^2 + nm + m^2}}$, where a_{c-c} denotes the carbon-carbon bond length (as 0.144 nm). According to the results represented in Fig. S3–S8 (ESI[†]), in detail, all common tubes clearly fit the M₁₁ branch. The effect of the catalyst particles on the chirality of the VACNTs, in terms of the relative RBM intensities, was evaluated for the (9,6), (9,3), (12,3) and (10,4) chiralities, which are plotted against the catalyst type in Fig. 5. Cheng *et al.* and Kharlamova *et al.* identified these same chiral vectors from RBM frequencies for their samples.^{61,74} It should be

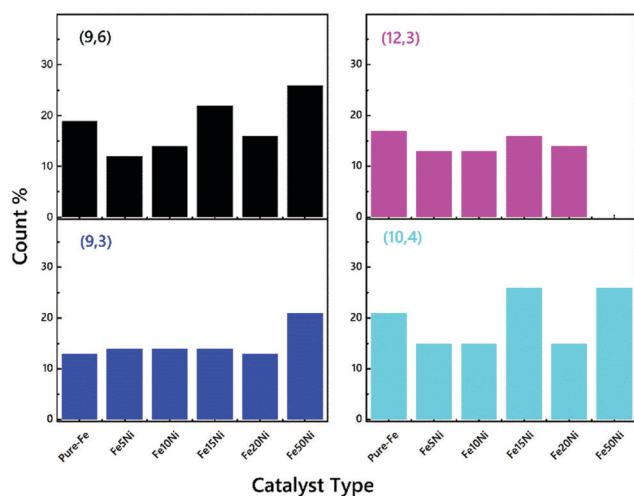


Fig. 5 The ratios of the CNTs with common (9,6), (9,3), (12,3) and (10,4) chiralities of the VACNTs synthesized using various catalyst particles.

noted that although the correlation between catalyst particles and chirality distribution is reported in this study, a more detailed investigation using sophisticated instrumentation such as microscopic and spectroscopic techniques is necessary for achieving the exact values for using the *in situ* CCVD system.

3.4 The effects of Ni addition on the oxidation behavior of the VACNTs

To understand the decomposition and purity of VACNTs, TGA was carried out at 800 °C under a 100 ml min⁻¹ airflow for all of the samples. According to the TGA results shown in Fig. S10 (ESI[†]), the highest temperature of the maximum rate of the oxidation resulted in the growth of the VACNTs on the Fe5Ni catalyst alloy. According to the TGA results, the mass of the residual metal catalyst nanoparticles obtained was ~2–3 wt% for each sample. The results of comparing the temperature of the maximum rate of oxidation depending on the catalyst type are shown in Table 2. Furthermore, these results can be correlated with the structural stability of the CNTs. In addition, the XRD pattern of the VACNTs is presented in Fig. S2 (ESI[†]). The three characteristic diffraction peaks of the VACNTs with 2θ angles of around 25°, 43°, and 77° are associated with the C (002), C (100) and C (110) diffractions of graphite, respectively.⁵⁴

3.5 Physical properties of the VACNTs

Fig. 6 shows micrographs of the VACNTs synthesized using neat Fe and Fe–Ni alloy catalysts from 5 to 50 wt% of Ni addition through imaging using SEM at an acceleration voltage of 5–15 kV at 50 000× magnification. The growth mechanism was determined as base-mode. This is due to the strong interaction between the catalyst and surface (the top-view SEM image of the VACNTs is shown in Fig. S9, ESI[†]). All of the samples show vertically aligned CNTs, without any significant difference in terms of their alignment. However, as depicted in Fig. 6(b), when the Fe5Ni alloy catalyst is used in the synthesis, an evident change was observed, with enhanced waviness and entanglements in between neighboring VACNTs. The carbon atoms precipitated during the catalyst cluster formation may present such entanglement and increased waviness similar to that observed in Fig. 6(b). According to ImageJ processing, the inter-wiring CNT–CNT spacing is approximately 75–80 nm for the synthesized VACNTs on both Pure-Fe and Fe5Ni. This is proof that the increase in waviness is due to the catalysis dynamics. This situation may have attributed to the high surface energy for the precipitation of carbon atoms. Li *et al.* conducted finite element

Table 2 TGA results of VACNTs synthesized using different catalyst nanoparticles

Type of catalyst	Temperature of the maximum rate of oxidation (°C)
Pure-Fe	696
Fe5Ni	704
Fe10Ni	700
Fe15Ni	651
Fe20Ni	668
Fe50Ni	703



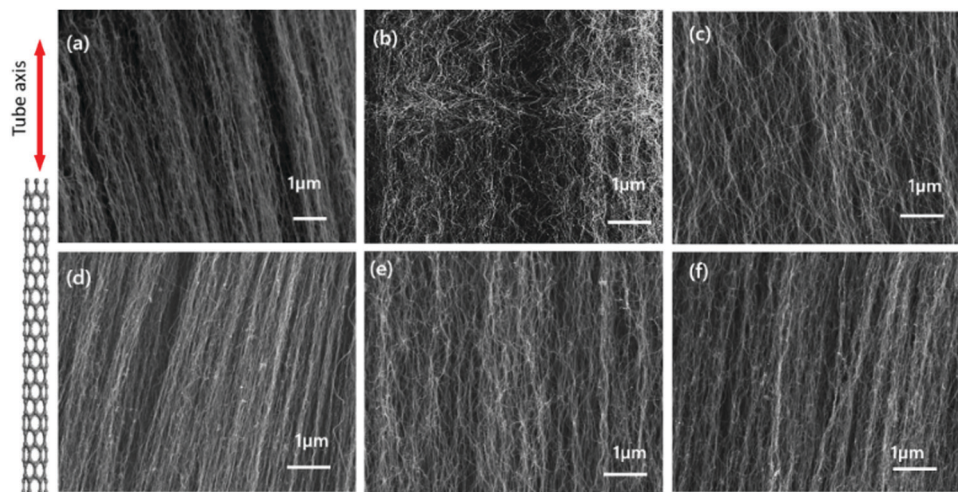


Fig. 6 FEG-SEM images of VACNTs synthesized on (a) Pure-Fe, (b) Fe5Ni, (c) Fe10Ni, (d) Fe15Ni, (e) Fe20Ni, and (f) Fe50Ni (scale bars: 1 μm).

simulations for CNT/polymer composites when CNT waviness was considered. The improved micromechanics model showed that

CNT waviness reduces the elastic modulus but enhances the ultimate strain of the composite.⁷⁵ From morphological analysis,

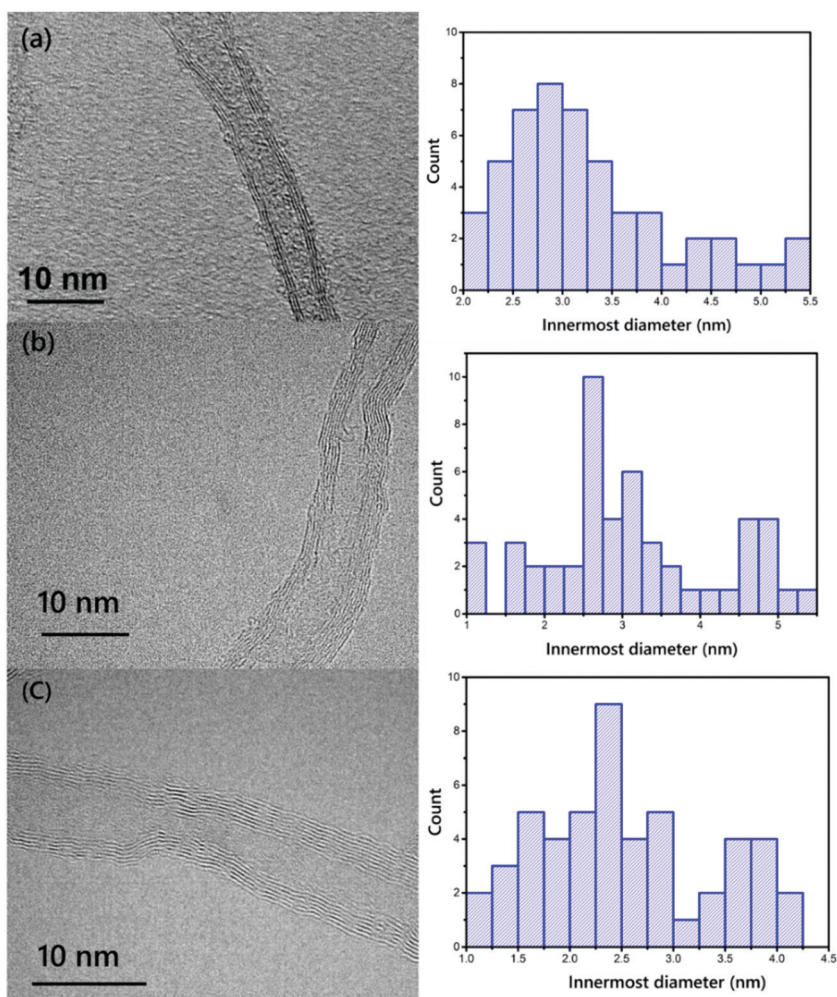


Fig. 7 TEM images and histograms of the distributions of the innermost diameters of the VACNTs synthesized on (a) Pure-Fe, (b) Fe5Ni, and (c) Fe50Ni (scale bars: 10 nm).



it can also clearly be concluded that the VACNTs synthesized using Fe5Ni have better potential as a reinforcing agent for use in PNC applications. However, the catalyst particles remain on the substrate, which is named as 'base-growth'. The strong interaction between the substrate and catalyst leads to a root-growth mechanism.

To further examine the morphology at the nanoscale, TEM images were captured of VACNT addition into the Fe catalyst, which (i) decreased the inner diameter of the CNTs, and (ii) increased the number of walls. 50 different lines were selected for calculation of the inner diameter of the CNTs using the ImageJ software. VACNTs grown from Pure-Fe show a wide distribution in terms of diameter of around 3.27 nm; the Fe15Ni catalyst synthesized VACNTs show a mean diameter of 3.11 nm; and the Fe50Ni catalyst yields mainly small tubes of around 2.52 nm. The TEM images are in clear agreement with the Raman spectroscopy results. The controlled CVD process using C₂H₄ as a carbon precursor also helps to achieve smaller diameter CNTs due to its low dissociation probability at high growth temperatures compared to other precursors. However, these TEM results prove that the number of walls of the VACNTs increases in line with the nickel in the catalyst particles, as can be clearly seen in Fig. 7. The number of walls of the VACNTs from Pure-Fe is 2–3 walls, whereas the numbers of walls of the VACNTs synthesized on Fe15Ni and Fe50Ni are in the range of 3–8.

4. Conclusion

Herein, the effect that the addition of Ni into an Fe catalyst has on VACNT growth and their many properties, including chirality, diameter, purity, quality, structural morphology and oxidation behavior, was investigated. Through using MA, the addition of Ni into the Fe catalysts was performed for VACNT growth using CVD. The synthesized VACNTs on different catalysts *via* CCVD were characterized using Raman spectroscopy, TGA, SEM and TEM. On the basis of the Raman spectroscopy at 532 nm laser wavelength, the following conclusions were drawn for all of the samples in terms of their peak positions, innermost diameters with metallic or semi conduction branches, Raman mappings, defect length of the VACNTs, and the counts of the common chiralities. To sum up, CNTs with metallic characteristics such as (9,3), (9,6), (10,4) and (12,3) were detected by Raman spectroscopy for all of the samples. The metallic behaviour of the synthesized CNTs with different Fe–Ni catalysts was seen to increase in line with the addition of nickel into the iron. Thus, this method improves the synthesis of VACNTs with metallic behaviour without the use of any etching and treatment methods. Furthermore, the oxidation behavior of the VACNTs was investigated using TGA. We reviewed the control of VACNTs with perfect alignment, as shown in the SEM images of all of the samples. According to the SEM images of VACNTs synthesized on the different catalyst particles, no significant change was observed in their morphologies. However, the VACNTs synthesized on Fe5Ni have a wavier structure compared to

others. TEM was used to determine the number of walls and inner diameters of the VACNTs. Further studies on the morphologies of the VACNTs with all of the Fe–Ni catalyst combinations will be carried out. With an increase in the Ni content in the catalysts, the innermost tube diameter of the VACNTs decreased from 3.27 to 2.52 nm for the Pure-Fe and Fe50Ni catalysts, respectively. These results are consistent with the Raman spectroscopy results. To the best of our knowledge, this is the first study to methodically investigate the relationship between Fe–Ni catalyst combinations and the selective growth of VACNTs. The mechanism was explained in terms of how the catalyst combinations affect the chirality distribution of few-walled VACNTs. In a future study, the synthesized different catalysts will be combined with a water-assisted chemical vapor deposition system to enable the growth of VACNTs with a narrow chirality distribution for use in electronic applications.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This study was financially supported by 38441 Scientific Research Project funded by ITU-Rectorate and TUBITAK 1001 projects (Project No. 112M807 and Project No. 117M290) for the materials and CVD system support. The authors would like to thank Dr H. Enis Karahan, and Zeki Semih Pehlivan for their feedback on the experimental design and discussions. The authors are also grateful to Prof. Haluk Küçük, Assist. Prof. Ali Fuat Ergenç, Assist. Prof. Didem Ovalı, Dr Alptekin Yıldız, Dr Kaan Yıldız, Dr Bihter Zeytuncu, Kıvanç Özışık, Fırat Turgut, Elçin Akar, and Fulden Kaygınoğlu for their support.

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