


 Cite this: *RSC Adv.*, 2020, 10, 37515

Systematic investigation of CO₂ : NH₃ ice mixtures using mid-IR and VUV spectroscopy – part 1: thermal processing†

 Rachel L. James,^a Sergio Ioppolo,^b Søren V. Hoffmann,^c Nykola C. Jones,^c Nigel J. Mason^d and Anita Dawes^a

The adjustment of experimental parameters in interstellar ice analogues can have profound effects on molecular synthesis within an ice system. We demonstrated this by systematically investigating the stoichiometric mixing ratios of CO₂ : NH₃ ices as a function of thermal processing using mid-IR and VUV spectroscopy. We observed that the type of CO₂ bonding environment was dependent on the different stoichiometric mixing ratios and that this pre-determined the NH₃ crystallite structure after phase change. The thermal reactivity of the ices was linked to the different chemical and physical properties of the stoichiometric ratios. Our results provide new details into the chemical and physical properties of the different stoichiometric CO₂ : NH₃ ices enhancing our understanding of the thermally induced molecular synthesis within this ice system.

 Received 3rd July 2020
 Accepted 25th September 2020

DOI: 10.1039/d0ra05826b

rsc.li/rsc-advances

1 Introduction

Despite over 200 molecules being detected in the interstellar medium (ISM), the formation pathways for most of these molecules remain elusive. However, processing of interstellar ices is believed to play an important role in the molecular synthesis of the majority of these astrochemical molecules. As a result, there exists extensive experimental investigations of condensed phase molecular films under ISM conditions.

Typical to all experiments, regardless of application, a range of experimental parameters can be controlled. For astrochemical experiments investigating interstellar ice analogues, these experimental parameters include, but are not limited to, the following: deposition temperature, deposition rate, ice composition (and ratio), processing type, processing time, processing energy and substrate type. Adjusting these experimental parameters will change the chemical and physical properties of the interstellar ice analogues, the properties of which govern the molecular synthesis within the ice and requires comprehensive analysis.

To demonstrate the impact that one discrete experimental parameter can have on the chemical and physical properties of

an interstellar ice system we investigated the stoichiometric mixing ratio of CO₂ : NH₃ ices as a function of thermal processing. We chose the CO₂ : NH₃ system as CO₂ and NH₃ are two of the most abundant molecules observed in interstellar ice and when combined contain the four most common elements to life. Furthermore, the two compounds do not share any of the elements between them, making identification of the origin of the constituents that make up the products easier, without the need for isotopic substitution. As such this is an attractive ice system to study and has been the subject of multiple previous studies which provide a benchmark to validate our systematic investigation of one discrete experimental parameter.^{1–9}

Table 1 summarises the previous thermal processing studies of CO₂ : NH₃ interstellar ice analogues. A consensus was that a thermally induced reaction within the CO₂ : NH₃ ice mixture resulted in the formation of carbamic acid and/or ammonium carbamate. Of the nine experimental studies shown in Table 1 only three studies^{1,6,7} investigated the stoichiometric mixing ratio of CO₂ : NH₃ ice, the discrete experimental parameter under focus in this paper. And, of these three studies, CO₂-rich mixtures were investigated by Frasco who makes no comment about the CO₂-rich mixtures¹ and Noble *et al.* who reported no thermal reaction in CO₂-rich mixtures.⁷ For equal parts or NH₃-rich mixtures Noble *et al.* observed that ammonium carbamate formed first and converted to carbamic acid at temperatures above 150 K.⁷ Rodríguez-Lazcano *et al.* reported that their CO₂ : NH₃ 1 : 2 ratio produced the highest yield of products compared to their 1 : 3 & 1 : 1 ratios, but offered no explanation as to why this preference existed.⁶ As such the effect of the stoichiometric mixing ratio parameter on the CO₂ : NH₃ ice system has potential for further investigation.

^aSchool of Physical Sciences, The Open University, Walton Hall, Milton Keynes, UK. E-mail: Rachel.James1@open.ac.uk; Fax: +44 (0)1908 654192; Tel: +44 (0)1908 332012

^bSchool of Electronic Engineering and Computer Science, Queen Mary University of London, Mile End Road, London, UK

^cISA, Department of Physics and Astronomy, Aarhus University, Ny Munkegade 120, DK-8000 Aarhus C, Denmark

^dSchool of Physical Sciences, University of Kent, Canterbury, Kent, UK

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0ra05826b



Table 1 Summary of the results of previous studies conducted on the thermal processing of CO₂ : NH₃ ice mixtures

Reference	CO ₂ : NH ₃ ratio	Deposition temperature (K)	Main products
Frasco ¹	Various (10 to 90% NH ₃) ^a	195	AC
Hisatsune ²	1 : 3 ^a	83	(NH ₃) ₂ CO ₂ & AC
Bossa <i>et al.</i> ³	1 : 1 ^b	10	AC & CA
Bertin <i>et al.</i> ⁴	1 : 1 ^a	10	AC & CA
Lv <i>et al.</i> ⁵	1 : 1 ^a	16	AC & CA
Rodríguez-Lazcano <i>et al.</i> ⁶	1 : 1, 1 : 2, 1 : 3 ^c	15	AC & CA
Noble <i>et al.</i> ⁷	Various (including 20 : 1, 13 : 1, 4 : 1, 1 : 1, 1 : 1.5) ^c	60	AC & CA
Potapov <i>et al.</i> ⁸	1 : 4 ^c	15	AC
Potapov <i>et al.</i> ⁹	1 : 4 ^c	15	AC

^a Ratios derived from partial pressures of the mixture in the gas line. ^b Method of determining ratio not specified. ^c Ratios derived from column density. AC = ammonium carbamate. CA = carbamic acid.

All of the studies shown in Table 1 employed IR spectroscopy, a powerful *in situ* technique commonly used for the investigation of the structure and composition of ice samples in the laboratory. Some of these studies also incorporated a second analytical technique, *e.g.* mass spectrometry^{3,7,8} or high resolution low energy electron loss spectroscopy,⁴ providing further characterisation of the ice. In addition to using mid-IR spectroscopy to investigate the thermal processing of CO₂ : NH₃ ice mixtures in this paper, we also present a complementary study using vacuum-UV (VUV) spectroscopy. This is the first time that CO₂ : NH₃ interstellar ice analogues have been investigated using VUV spectroscopy. Both mid-IR and VUV spectroscopy are used *in situ* and provide information on the end products allowing for the monitoring of the complex chemical and physical processes involved in forming these products.

This work is part of a wider, ongoing collection of experiments aimed at understanding how discrete experimental parameters impact both the chemical and physical properties of an interstellar ice analogue which govern molecular synthesis. In particular, this paper focusses on the influence that the stoichiometric mixing ratio has on the chemical and physical properties of CO₂ : NH₃ mixtures when subjected to thermal processing. A corresponding paper will investigate the influence that the stoichiometric mixing ratio has on the chemical and physical properties of CO₂ : NH₃ mixtures when subjected to both non-thermal and thermal processing.

In this paper, we present mid-IR spectra of the thermal processing of CO₂ : NH₃ mixtures (3 : 1, 2 : 1, 1 : 1, 1 : 3 & 1 : 10) deposited at 20 K. We complement the mid-IR spectra with the first VUV spectroscopic study of CO₂ : NH₃ mixtures (4 : 1, 1 : 1 & 1 : 3) which were deposited at 20 K and thermally processed. We characterise these stoichiometric mixing ratios in detail at 20 K and demonstrate that the stoichiometric mixing ratio has a significant impact on both the chemical and physical properties of the CO₂ : NH₃ ice system at deposition and throughout thermal processing.

2 Experimental

Both the mid-IR and VUV experiments were performed using The Open University Portable Astrochemistry Chamber (PAC). The PAC set-up is described in detail in Section S1.1 of the ESI.† All samples

were grown *via* physical vapour deposition onto a cooled substrate (mid-IR: ZnSe, Crystran; VUV: MgF₂, Crystran) at a base temperature of 20 K and a base pressure of low 10⁻⁹ mbar. CO₂ (99.8%, BOC) and NH₃ (99.96%, ARGO International Ltd) were premixed in the gas line prior to deposition. The ice samples were thermally processed at a heating rate of approximately 0.1 K s⁻¹ to a set temperature. The acquisition time for a mid-IR measurement was approximately 2 min. The acquisition time for a VUV measurement was dependent on the step size used, which for measurements taken at temperatures ≤ 80 K corresponded to ~30 min and at temperatures ≥ 90 K corresponded to ~10 min. As acquisition times were different for the mid-IR and VUV spectroscopic measurements, the samples were allowed to isothermally stabilise for 2 min before the spectroscopic measurement was taken at the set temperature.

The mid-IR experiments were performed at The Open University Molecular Astrophysics Laboratory, UK, using a FTIR spectrometer (Nicolet Nexus 670) with an external MCT detector. All mid-IR spectra were acquired in absorbance over the wavenumber range of 4000–800 cm⁻¹ at a resolution of 1 cm⁻¹. Background scans were averaged over 512 scans and sample scans were averaged over 128 scans. All mid-IR spectra presented in this paper were obtained at an oblique angle (45°) to the IR radiation unless otherwise stated. For the VUV experiments, the PAC was attached to the AU-UV beam line at the ASTRID2 storage ring, Aarhus University, Denmark. All VUV spectra were acquired in absorbance over the wavelength range of 120–340 nm with 0.05 to 1 nm wavelength step size depending on the width of the spectral features to be resolved. The average photon flux per point was 2 × 10¹⁰ photons s⁻¹ and collection time per point was 2.69 s. VUV processing (*e.g.* photoionisation) is wavelength dependent and given the low average photon flux and short collection time, any VUV processing during a measurement is negligible. All VUV spectra presented in this paper were obtained at an oblique angle (45°) to the UV radiation unless otherwise stated.

All mid-IR and VUV spectra are freely available on the Open Research Data Online (ORDO) Repository.¹⁰

2.1 Film thickness

Film thickness was determined from *in situ* laser interferometry measurements using a HeNe laser beam reflected off the



substrate during deposition (see Section S1.2 of the ESI† for more details). The same deposition conditions were maintained for both mid-IR and VUV experiments and the deposition rates were between 0.8–1.8 nm s⁻¹. On average, the film thickness of the mid-IR samples was 402 nm. Thinner films were required to prevent saturation of absorption peaks for the VUV samples compared to mid-IR samples and the average film thickness was 204 nm. Where comparisons were made between the mixtures, the spectra were normalised to specific film thickness, 400 nm for mid-IR spectra and 200 nm for VUV spectra, and is indicated in the figure captions. See Table S1 in the ESI† for the individual sample thickness and normalisation factors.

2.2 Determining the CO₂ : NH₃ mixing ratios

It is well known that discrepancies exist between the ratio of the partial pressures in the gas line and the resultant mixing ratio of the deposited sample. For the mid-IR samples the CO₂ : NH₃ ratios were determined from the derived column densities of CO₂ (ν_3 absorption band) and NH₃ (ν_2 absorption band) when the substrate was positioned at normal incidence to the IR radiation. For specific details on the calculations used, integration range of the absorption peak and integrated band strengths see Section S1.3 of the ESI.† VUV spectroscopy does not have the same protocol for determining the ratios directly from the spectra as mid-IR spectroscopy. A method for determining the ratios from VUV spectra without needing to calculate photoabsorption cross sections is proposed in Section S1.4 of the ESI.† The VUV samples used the same mixing partial pressures as the mid-IR samples. While the partial pressures do not correlate with the deposited mixtures, a consistent mixing ratio was obtained from the corresponding partial pressures and this was used to check the validity of the method proposed.

3 Mid-IR results & discussion

3.1 Deposition at 20 K

Fig. 1 shows the mid-IR spectra of CO₂ : NH₃ mixtures (3 : 1, 2 : 1, 1 : 1, 1 : 3 & 1 : 10) deposited at 20 K compared with pure CO₂ (1 : 0) and pure NH₃ (0 : 1) also deposited at 20 K. The band assignments and positions are given in Table 2.

In addition to the vibrational modes associated with NH₃ and CO₂ two new vibrational modes were observed in the CO₂ : NH₃ mixtures in Fig. 1(a) and are marked with dashed lines. The first vibrational mode at 3417 cm⁻¹ was present for all ratios except the 1 : 3 and 1 : 10 ratios and the second vibrational mode at 3253 cm⁻¹ was observed for all ratios except the 1 : 10 ratio. Previous studies have assigned these vibrational modes to a CO₂ : NH₃ molecular complex in a ‘T-shape’ whereby the NH₃ molecule is complexed to the C atom in the CO₂ molecule *via* the lone pair of electrons on the N atom of NH₃.^{3,14,15}

Due to the oblique angle of the ice films to the incident IR radiation, longitudinal optical (LO)-transverse optical (TO) splitting of the ν_3 vibrational mode of CO₂ was observed for pure CO₂ ice and the 3 : 1, 2 : 1 & 1 : 1 mixtures as shown in Fig. 1b.^{16,17} LO-TO splitting arises due to long-range dipole

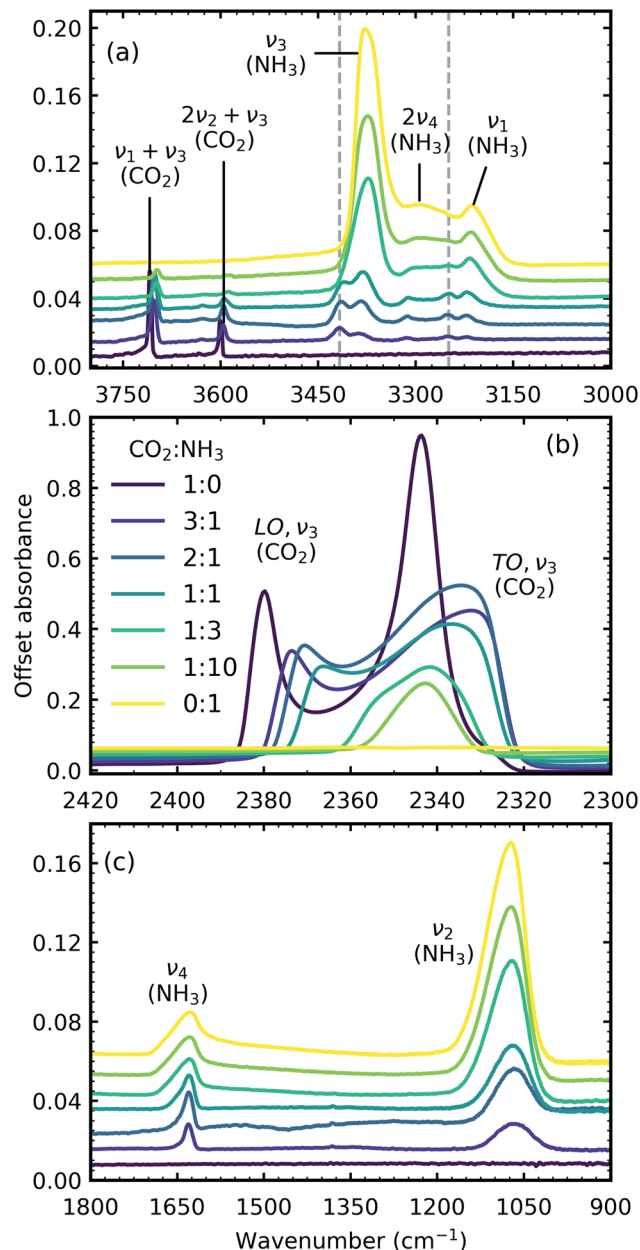


Fig. 1 Mid-IR spectra of CO₂ : NH₃ mixtures (3 : 1, 2 : 1, 1 : 1, 1 : 3 & 1 : 10) at 20 K compared with pure CO₂ (1 : 0) and pure NH₃ (0 : 1). (a) Dashed lines represent new features which formed upon deposition. (b) LO-TO splitting of the ν_3 absorption band of CO₂ (c) NH₃ ν_4 and ν_2 absorption bands. Band assignments are given in Table 2. Spectra are offset on the y-axis for clarity. All spectra are normalised to a thickness of 400 nm.

interactions in the CO₂ lattice and the position of the LO mode is sensitive to defects within the CO₂ lattice.¹⁷ The CO₂ in the NH₃-rich mixtures (1 : 3 & 1 : 10) can be thought of as defects within the NH₃ ice and hence the LO mode was not observed. The shape of the LO mode can also give insight into the mixing environment of CO₂ mixtures. For example, Cooke *et al.* noted that the LO mode would split when inhomogeneous mixing occurred.¹⁷ No splitting of the LO mode was observed for 3 : 1,



Table 2 Band assignments and positions of the vibrational modes of pure CO₂ ice (1 : 0), pure NH₃ ice (0 : 1) and CO₂ : NH₃ mixtures (3 : 1, 2 : 1, 1 : 1, 1 : 3 & 1 : 10) deposited at 20 K

Molecule	Vib. mode	Assignment	Ref.	Position (cm ⁻¹)						
				1 : 0	3 : 1	2 : 1	1 : 1	1 : 3	1 : 10	0 : 1
CO ₂	$\nu_1 + \nu_3$	Combination	11	3709	3703	3702	3701	3698	3697	
	$2\nu_2 + \nu_3$	Combination	11	3601	3596	3596	3595	3590	3588	
	ν_3	C=O <i>asym.</i> stretch (LO)	12	2380	2374	2371	2366			
	ν_3	C=O <i>asym.</i> stretch (TO)	11 and 12	2345	2332	2334	2337	2342	2343	
¹³ CO ₂	ν_3	<i>asym.</i> stretch	11	2283	2279	2279	2278	2277	2277	
CO ₂ : NH ₃ complex			3		3417	3415	3413			
			3		3250	3251	3248	3248		
NH ₃	ν_3	N-H <i>asym.</i> stretch	13		3384	3382	3379	3373	3373	3377
	$2\nu_4$	Overtone	13		3313	3312	3311	3301	3298	3294
	ν_1	N-H <i>sym.</i> stretch	13		3220	3220	3220	3215	3214	3212
	$\nu_4 + \nu_L$	Combination	13					1872	1872	1883
	ν_4	Deformation	13		1630	1630	1629	1629	1628	1627
	ν_2	Umbrella	13		1066	1067	1069	1071	1073	1073

2 : 1 & 1 : 1 CO₂ : NH₃ mixtures which suggested homogeneous mixing occurred.

3.1.1 Differing absorption band shapes and positions. Broadening or narrowing of the vibrational absorption bands in the mixtures compared to the respective pure ice vibrational absorption bands were observed as shown in Fig. 1. In solids, the vibrational absorption band shapes are strongly influenced by intermolecular interactions or its local environment. Each molecule may perceive a different local bonding environment and hence, the observed vibrational absorption band is a normal distribution of oscillators around an average band intensity. Changes in the vibrational absorption band shapes of the mixtures from the pure ices reflect the differing local environments. Broadening of the absorption band suggests a wider distribution of bond lengths within the different local environments such as in a disordered matrix. Whereas, narrowing of the absorption band typically suggests an increased order such as a crystalline matrix contributing to a narrower distribution of bond lengths within the local bonding environment.

Shifts in the position of the vibrational absorption bands in the mixtures compared to the respective pure ice vibrational absorption bands can also be seen in Fig. 1. A red shift is associated with an increase in bond length, and hence a weakening of the bond associated with the absorption band. The reverse is true for observed blue shifts in the spectra. Fig. 2 shows the difference between the position of the pure ice vibrational absorption band and the position of the ice mixture vibrational absorption band ($\Delta\tilde{\nu} = \tilde{\nu}_{\text{pure ice}} - \tilde{\nu}_{\text{mixed ice}}$) against the [CO₂]/[NH₃] ratio (R).

All the CO₂ vibrational absorption band positions in the mixtures (Fig. 2a) red shifted compared to the pure CO₂ vibrational absorption bands. For all vibrational absorption bands, except for the ν_3 TO mode, a progressive red shift was observed for increasing concentrations of NH₃ (lower R values). This suggested a progressive weakening of the C=O bonds with increasing NH₃ concentration. In contrast, a progressive red

shift with decreasing concentrations of CO₂ (higher R values) was observed for the ν_3 TO mode.

The ν_3 TO mode absorption bands for the 3 : 1, 2 : 1 & 1 : 1 ratios also had a distinct broad, asymmetric profile which reduced in asymmetry for the NH₃-rich mixtures, but not equivalent to the profile of pure CO₂ ice. The asymmetry or inhomogeneous broadening of the TO mode absorption band suggested that the absorption band of the C=O asymmetric stretch contained contributions from CO₂ molecules in different bonding environments. In contrast, the relatively narrow profile of the TO mode absorption band for pure CO₂ was indicative of a contribution from a largely single bonding environment and previous studies suggested that pure CO₂ formed dimers upon deposition at 20 K.^{18,19} Where NH₃ was dominant in the 1 : 10 ratio the CO₂ molecules were akin to defects in the NH₃ ice and were essentially matrix-isolated. This reduced CO₂ dimerisation and thus presented a single bonding environment of the C=O stretch as that of an isolated CO₂ molecule. While NH₃ was still in excess in the 1 : 3 ratio, evidence of the formation of CO₂ : NH₃ molecular complexes were observed through the absorption band at 3253 cm⁻¹. The shoulder on the TO mode absorption band at 2354 cm⁻¹ also suggested that there were contributions from two bonding environments to the C=O asymmetric stretch, most likely that of isolated CO₂ and CO₂ : NH₃ molecular complex. This was further supported by the increased asymmetry of the TO mode absorption band for the other mixtures which was likely a combination of several bonding environments that influenced the C=O asymmetric stretch (*e.g.* CO₂ : NH₃ molecular complex, CO₂ dimers and isolated CO₂ molecules). Therefore the progressive red shift in the peak position with decreasing concentration of CO₂ reflected the change in asymmetry of the TO mode as a result of different local CO₂ bonding environment, which arose due to the change in relative CO₂ and NH₃ concentrations. The combination modes of CO₂ were considerably weaker and less sensitive to changes in the lattice compared to the ν_3 asymmetric stretch of CO₂. The progressive



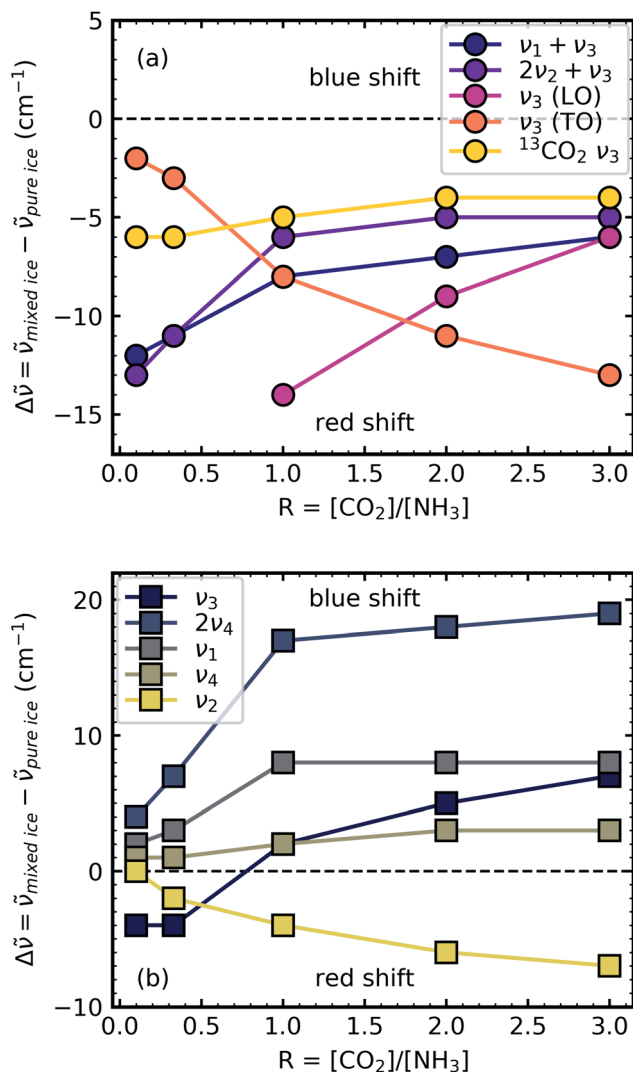


Fig. 2 Scatter plots of the difference between the pure ice vibrational absorption band and the ice mixture vibrational absorption band ($\Delta\tilde{\nu} = \tilde{\nu}_{\text{pure ice}} - \tilde{\nu}_{\text{mixture ice}}$) against the $[\text{CO}_2]/[\text{NH}_3]$ ratio (R) for (a) CO_2 vibrational absorption bands and (b) NH_3 vibrational absorption bands. Lines between the scatter points are to guide the eye only.

weakening of the CO_2 combination modes with increasing NH_3 concentration compared to pure CO_2 reflected the changes in the CO_2 bonding environments, as well as the relative decrease in the CO_2 concentration.

For the NH_3 vibrational modes in the mixtures (Fig. 2b) a progressive blue shift was observed as the CO_2 concentration increased (higher R values) for the $2\nu_4$, ν_1 & ν_4 absorption bands. This suggested a progressive strengthening of the N–H bonds with increasing concentrations of CO_2 . At lower concentrations of CO_2 (1 : 10 & 1 : 3 ratios) the ν_3 absorption band of NH_3 was red shifted ($\Delta\tilde{\nu} = -4 \text{ cm}^{-1}$). However, a progressive blue shift was observed for increasing concentrations of CO_2 for the other ratios. In addition, the NH_3 vibrational modes in the CO_2 : NH_3 mixtures also progressively narrowed with increasing concentration of CO_2 compared to pure NH_3 absorption bands. Pure NH_3 ice has extensive intermolecular H-bonding present

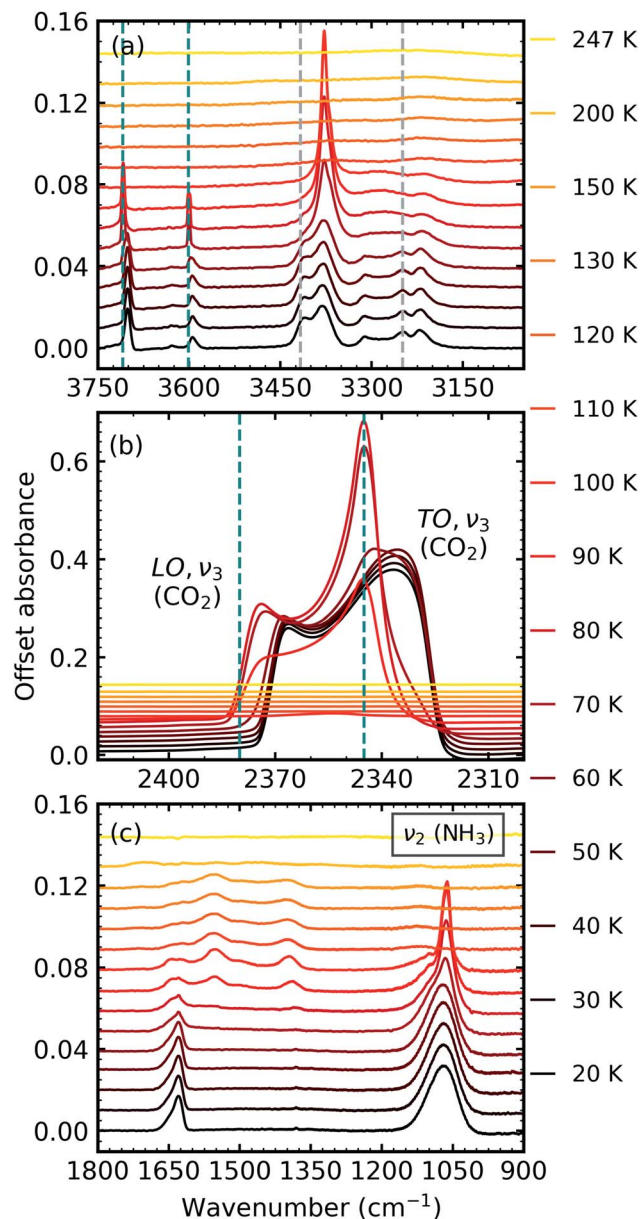


Fig. 3 Mid-IR spectra of a CO_2 : NH_3 mixture in a 1 : 1 ratio during thermal processing from 20 to 250 K. Spectra are offset on the y-axis for clarity. See Section S2 in the ESI† for the mid-IR spectra of the 3 : 1, 2 : 1, 1 : 3 & 1 : 10 ratios. (a) Segregation of the mixture was observed through the shift in the position of the CO_2 absorption bands towards the position of pure CO_2 absorption bands when deposited at 20 K which are indicated by blue dashed lines. Grey dashed lines indicate the CO_2 : NH_3 molecular complex absorption bands which disappeared between 70–80 K. (b) LO–TO splitting of the ν_3 absorption band of CO_2 . Segregation of the mixture was observed through the shift in the position of the LO and TO modes towards pure CO_2 positions which are indicated by blue dashed lines. (c) A phase change was observed for NH_3 through the splitting of the ν_2 absorption band between 70–80 K. New bands between 1800–1200 cm^{-1} at 80 K indicated thermal reaction.

between the NH_3 molecules.²⁰ For the 1 : 10 ratio where the NH_3 ice essentially has CO_2 defects, the intermolecular H-bonding between the NH_3 molecules was slightly perturbed. As the CO_2 concentration increased the intermolecular H-bonding was



progressively reduced and the formation of CO₂ : NH₃ molecular complexes also occurred.

A progressive red shift was observed for increasing concentrations of CO₂ for the ν_2 absorption band in contrast to the behaviour of the other NH₃ absorption bands which mainly blue shifted. At higher concentrations of CO₂, NH₃ molecules were likely to be bonded in a CO₂ : NH₃ complex. The 'T-shape' of the CO₂ : NH₃ complex resulted in a less restricted bending motion of the ν_2 mode due to reduced intermolecular H-bonding in a CO₂-rich environment.¹⁴

3.2 Thermal processing

After deposition at 20 K, the CO₂ : NH₃ mixtures were thermally processed and analysed at discrete temperatures until desorption. Fig. 3 shows the mid-IR spectra of a CO₂ : NH₃ mixture in a 1 : 1 ratio during thermal processing. Mid-IR spectra of the other ratios can be found in Section 2 of the ESI.† For reference, the thermal processing mid-IR spectra of pure CO₂ and pure NH₃ are shown in Fig. S8 and S9 of the ESI,† respectively.

Thermal processing induced several changes within the 1 : 1 CO₂ : NH₃ mixtures as shown in Fig. 3. Segregation of the homogeneously mixed CO₂ : NH₃ ice mixture was identified through a blue shift in the CO₂ vibrational modes between 60–70 K towards the position of pure CO₂ vibrational modes when deposited at 20 K (Fig. 3a and b). We note that the only IR absorption band of CO₂ to shift during thermal processing was the LO mode of the ν_3 asymmetric stretch which reached a maximum blue shift of ~ 40 cm⁻¹ at 80 K. Splitting of the ν_2 fundamental mode of NH₃ between 70–80 K signified a phase change in NH₃ (Fig. 3c). A thermally induced reaction was initiated between 70–80 K and was observed through the appearance of new vibrational modes in the wavenumber region of 1800–1200 cm⁻¹ (Fig. 3c). The vibrational modes of the CO₂ : NH₃ complex also disappeared between 70–80 K (Fig. 3a). Both CO₂ and NH₃ desorbed between 110–120 K leaving behind

a residue material which changed between 150–200 K before desorbing by 250 K.

The temperatures at which these changes occurred were dependent on the ratio and are listed in Table 3. Notably, no thermal reaction was observed for the 3 : 1 ratio and the mixture desorbed between 100–110 K.

3.2.1 Phase change of NH₃. Fig. 4 shows the ν_2 absorption band of NH₃ for all CO₂ : NH₃ mixtures and pure NH₃ deposited at 20 K and thermally processed to 90 K. Different splitting patterns were observed in the ν_2 umbrella absorption band of NH₃ for the CO₂ : NH₃ mixtures compared to pure NH₃ and also within the mixtures.

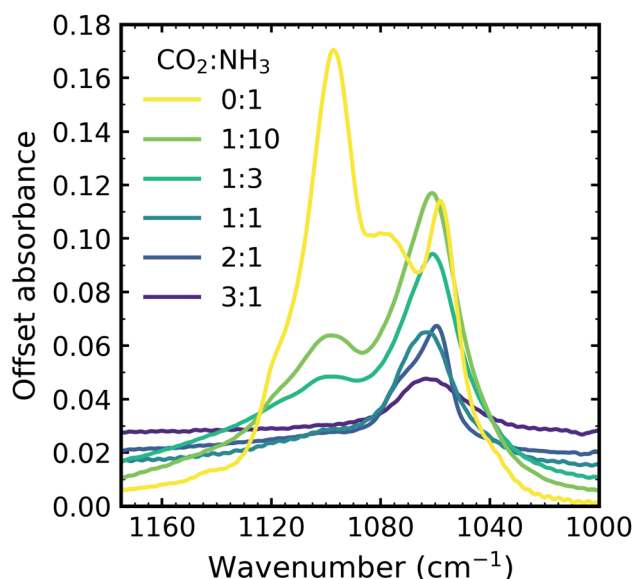


Fig. 4 Mid-IR spectra of NH₃ ν_2 umbrella absorption band in the CO₂ : NH₃ mixtures which were deposited at 20 K and thermally processed to 90 K compared with pure NH₃ ice deposited at 20 K and thermally processed to 90 K. Spectra are offset on the y-axis for clarity and normalised to a thickness of 400 nm.

Table 3 Summary of key observations during the thermal processing of CO₂ : NH₃ mixtures (3 : 1, 2 : 1, 1 : 1, 1 : 3 & 1 : 10)

Observation	Ratio				
	3 : 1	2 : 1	1 : 1	1 : 3	1 : 10
CO ₂ segregation (K)	By 60	50–60	60–70	70–80	70–80
NH ₃ phase change (K)	By 60	60–70	70–80	70–80	70–80
New bands (K)		>90	>80	>80	>80
New band positions (cm ⁻¹)				1735	1734
		1647	1647	1646 ^{sh}	1646
		1620	1623	1627	1624 ^{sh}
		1553	1554	1553	1554
			1501 ^{sh}	1501	1506
Disappearance of CO ₂ : NH ₃ complex (K)	80–90	1396	1396	1390	1389
		70–80	70–80	70–80	
CO ₂ desorption (K)	90–100	90–100	100–110	100–110	110–120
NH ₃ desorption (K)	100–110	100–110	100–110	100–110	110–120
Change in residue (K)		150–200	150–200	170–200	170–200
Residue desorption (K)		200–250	200–250	200–250	200–250

^{sh}shoulder



Pure NH_3 deposited at low temperatures (10–20 K) has been described as lacking long-range order or as ‘amorphous’.^{20,21} A more quantitative description has also been presented as containing nano-crystallites comprised of 64 NH_3 molecules with a lower crystallite size limit of 1.58 nm at 20 K. Depositions at higher temperatures were comprised of larger crystallites (*e.g.* NH_3 deposited at 60 K formed crystallites comprised of 756(\pm 20%) molecules with a crystallite size of 3.62 nm).²² A phase change for pure NH_3 deposited at 20 K and then thermally processed occurred at 57 K (ref. 21) and was signified by a complex splitting pattern similar to that presented in Fig. 4 for pure NH_3 (0 : 1). Different splitting patterns have also been observed for NH_3 deposited between 65–85 K and above 85 K.²⁰ The splitting pattern of the 1 : 10 & 1 : 3 ratios were similar to that of pure NH_3 deposited between 65–85 K and the splitting pattern of the other ratios was similar to that of pure NH_3 deposited above 85 K.

The bending umbrella motion of the NH_3 molecules described by the ν_2 vibrational mode was highly perturbed by the surrounding environment. The deposition temperature affected the crystallite size and hence the number of crystallite grain boundaries present. A crystallite grain boundary can be considered as a defect within the ordered H-bonded crystal structure. Broader absorption bands for the ν_2 vibrational mode were observed at lower deposition temperatures of NH_3 as there were more defects within an ordered H-bonded crystal structure due to the larger surface-area-to-volume ratio of the crystallites.

Narrower peaks were observed for NH_3 deposited at higher temperatures as the crystallite sizes were larger resulting in fewer defects present with the H-bonded crystal (*i.e.* lower crystallite surface-area-to-volume area) and more long-range order. However, thermal processing of the low deposition temperature ices was unable to overcome pre-existing H-bonding within the crystallites that were already formed during deposition, hence forming an ice sample with more crystallite grain boundaries than the equivalent ice deposited at the thermal processing temperature.²⁰ This was reflected in the profile of the ν_2 absorption band in Fig. 4 with an extensive splitting pattern resulting from the formation of crystallites of random shapes and sizes.

In Section 3.1 we refer to the presence of CO_2 within the mixtures disrupting the H-bonding network between NH_3 molecules. As a consequence, larger NH_3 crystallites were able to form (with fewer crystallite grain boundaries) in CO_2 : NH_3 mixtures with higher concentrations of CO_2 . The difference in the splitting patterns within ν_2 absorption band was a direct effect of the different concentrations of CO_2 present within the mixtures and we propose that the NH_3 crystallite formation is dependent on the CO_2 : NH_3 ratio.

3.3 Residue

A thermal reaction was observed for all CO_2 : NH_3 mixtures above \sim 80 K, except for the 3 : 1 ratio where no thermal reaction

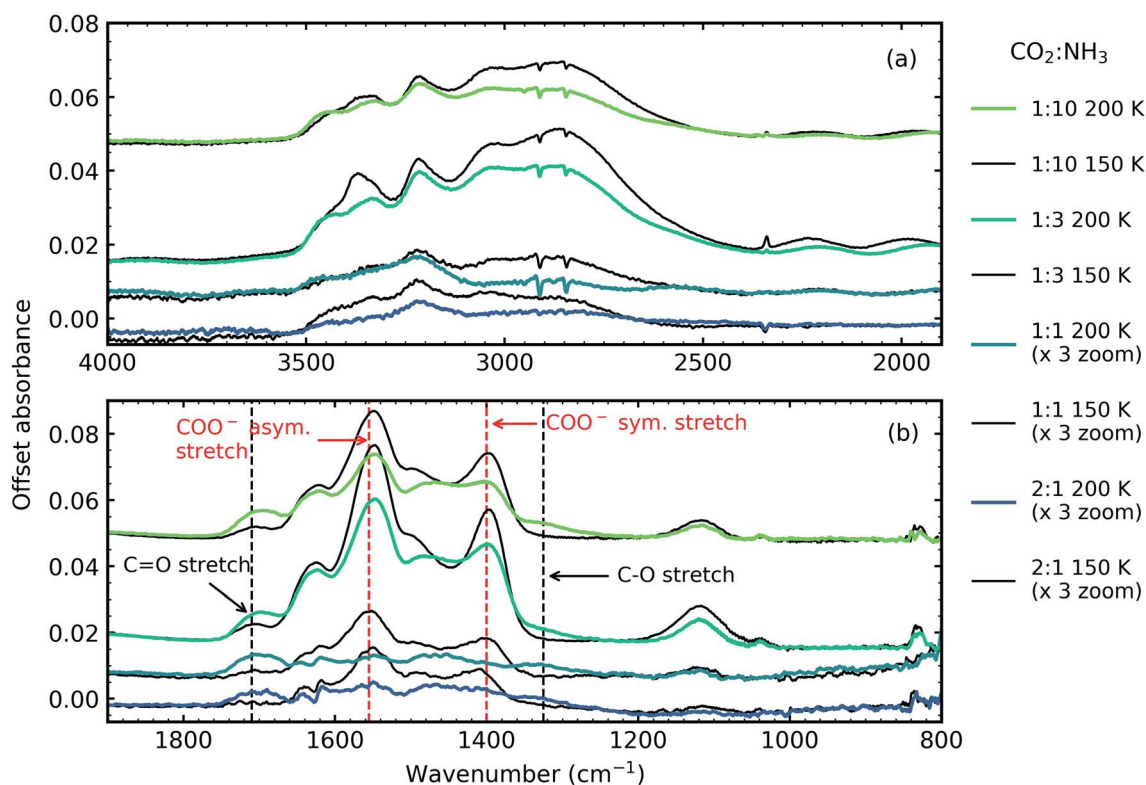


Fig. 5 Mid-IR spectra of CO_2 : NH_3 ices thermally processed to 150 K (black traces) and 200 K (coloured traces) (a) N–H and O–H stretching region (4000–1900 cm^{-1}) and (b) 1900–800 cm^{-1} with dashed lines indicating COO^- asymmetric and symmetric stretches and C=O and C–O stretches. The 2 : 1 and 1 : 1 ratios have been magnified by a factor of 3 for clarity. All spectra are normalised to a thickness of 400 nm and offset on the y-axis for clarity.



was observed. Previous thermal processing studies of CO₂ : NH₃ mixtures have identified ammonium carbamate and carbamic acid as thermal products above 150 K.³⁻⁷ However, discrepancies exist over the assignment of vibrational modes (see Table S2 in the ESI† for examples). The focus of this study was not to assign all the vibrational modes as this requires complementary theoretical calculations. Instead, differences between the functional groups which compose ammonium carbamate or carbamic acid were used *e.g.* ammonium carbamate is characterised by strong COO⁻ asymmetric and symmetric stretches, while carbamic acid is characterised by C=O and C-O stretches.

Fig. 5 shows the residue spectra at 150 and 200 K for the 2 : 1, 1 : 1, 1 : 3 and 1 : 10 ratios. The N-H and O-H stretching region is shown in Fig. 5a. Ammonium carbamate and carbamic acid were identified as products in the residue in Fig. 5b at 150 and 200 K and was in agreement with several previous studies.^{3,5,6} The intensity of the residues, and hence the amount of residue formed, was ratio-dependent with the 1 : 2 & 1 : 1 residues significantly weaker in intensity compared to the 1 : 3 & 1 : 10 ratios. The most intense residue was observed for the 1 : 3 ratio similar to a previous study which investigated CO₂ : NH₃ mixtures (1 : 1, 1 : 2, 1 : 3) and observed that a 1 : 2 ratio had the largest amount of residue material.⁶ While we also observed a reaction in the 2 : 1 ratio, this residue had the least amount of residue material and no reaction was observed in our 3 : 1 ratio, agreeing with Noble *et al.* that CO₂-rich mixtures hinder thermal reaction.⁷

Two NH₃ molecules per one CO₂ molecule are required to form ammonium carbamate, so it is not surprising that our 1 : 3 ratio had the most intense residue. Yet, the 1 : 10 ratio that was comprised of only 9% CO₂ produced a more intense residue than the 1 : 1 & 2 : 1 ratios indicating a complex interplay between the different chemical and physical properties of the mixtures and is discussed further in Section 5.

4 VUV results & discussion

4.1 Deposition at 20 K

Fig. 6 shows the VUV spectra of CO₂ : NH₃ mixtures (4 : 1, 2 : 1, 1 : 3) deposited at 20 K compared with pure CO₂ (1 : 0) and pure NH₃ (0 : 1) also deposited at 20 K.

Pure CO₂ deposited at 20 K has an absorption band centred around 126 nm due to the ¹Π_g ← ¹Σ_g⁺ electronic transition and exhibits extensive vibrational bands.²³ The average separation between the vibrational bands was 619 cm⁻¹ in keeping with previous studies²³⁻²⁵ and corresponded to the ν₂ bending mode of CO₂ suggesting a change in geometry of the molecule from linear to bent upon excitation.²⁵ A second, weaker absorption band was observed centred around 143 nm due to the ¹Δ_u ← ¹Σ_g⁺ electronic transition.²³ Pure NH₃ deposited at 20 K has an absorption band centred at 128 nm which was likely due to contributing D, E, F and G ← X Rydberg transitions.²² A second, weaker absorption band centred at 178 nm was observed due to the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ electronic transition.²⁰

The positions of the absorption bands of the ¹Π_g ← ¹Σ_g⁺ and ¹Δ_u ← ¹Σ_g⁺ electronic transitions of CO₂ and the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ electronic transition of NH₃ overlap between

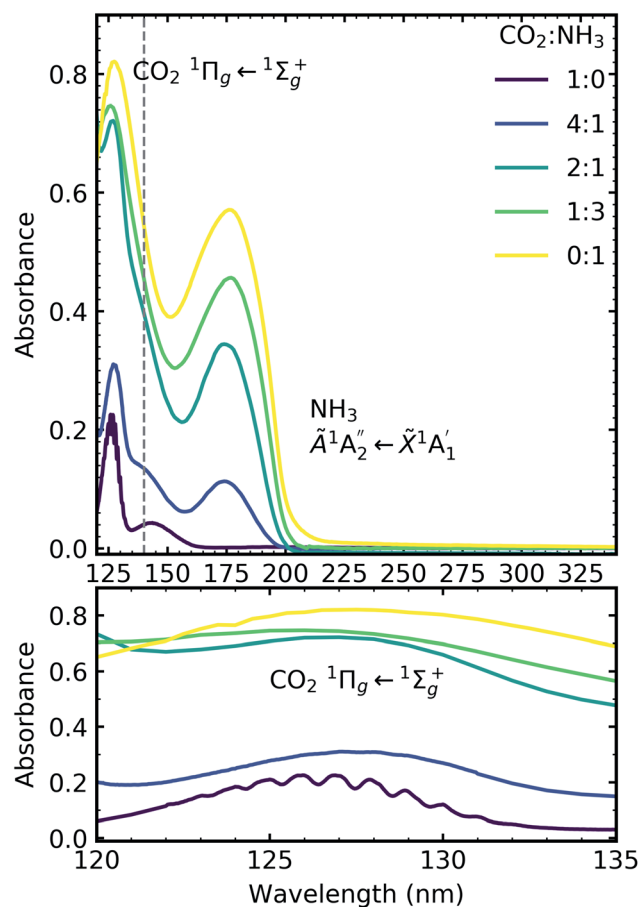


Fig. 6 VUV spectra of CO₂ : NH₃ mixtures (4 : 1, 2 : 1 & 1 : 3) compared with pure CO₂ (1 : 0) and pure NH₃ (0 : 1) between 120–340 nm (top plot). Dashed line highlights the ¹Δ_u ← ¹Σ_g⁺ electronic transitions of CO₂ at ~140 nm for the 4 : 1 and 2 : 1 ratios. Bottom plot shows the vibrational structure of the CO₂ ¹Π_g ← ¹Σ_g⁺ transition between 120–135 nm. Spectra are normalised to a thickness of 200 nm.

120–150 nm. The NH₃ electronic transitions have a higher cross section compared to that of the CO₂ electronic transitions and for the 1 : 3 ratio, $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ electronic transition of NH₃ largely obscures the CO₂ electronic transitions. For the 4 : 1 & 2 : 1 ratios, the ¹Π_g ← ¹Σ_g⁺ electronic transition of CO₂ was observed centred at 127 nm. Unlike pure CO₂, no vibrational bands were observed in the ¹Π_g ← ¹Σ_g⁺ transition for the 4 : 1 and 2 : 1 mixtures at 20 K. The ¹Δ_u ← ¹Σ_g⁺ electronic transition of CO₂ was observed as a shoulder on the absorption band of the contributing D, E, F and G ← X Rydberg transitions of NH₃ at 141 nm for the 4 : 1 & 2 : 1 ratios. The $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ transition of NH₃ was observed in all CO₂ : NH₃ mixtures and slightly blue shifted compared to the pure NH₃ band.

4.2 Thermal processing

Fig. 7, 8 and 9 show the thermal processing VUV spectra of CO₂ : NH₃ mixtures in ratios of 4 : 1, 2 : 1 & 1 : 3 respectively.

Thermal processing VUV spectra of pure CO₂ and pure NH₃ ices are shown in Fig. S9 and S10 of the ESI,† respectively and



the results are summarised briefly here. For pure CO₂, vibrational structure was observed on the ${}^1\Delta_u \leftarrow {}^1\Sigma_g^+$ transition of CO₂ at 70 and 80 K with an average space separation of 1484 cm⁻¹ at 70 K and 1444 cm⁻¹ at 80 K. This vibrational structure was not observed at deposition (20 K). For pure NH₃, factor-group (Davydov) splitting of the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ transition of NH₃ was observed at 70 K indicating a phase change.²⁰

For the CO₂ : NH₃ mixtures at 20 K no vibrational structure was observed for the ${}^1\Pi_g \leftarrow {}^1\Sigma_g^+$ transition of CO₂ for the 4 : 1 & 2 : 1 ratios (observed in the 1 : 3 ratio). However, vibrational structure was observed upon thermal processing to 70 K for the 4 : 1 ratio and 60 K for the 2 : 1 ratio (VUV spectra below 70 K were not obtained for the 4 : 1 ratio). In agreement with the mid-IR results (see Table 3), desorption of CO₂ occurred between 90–100 K and was observed through the disappearance of the ${}^1\Pi_g \leftarrow {}^1\Sigma_g^+$ transition for both the 4 : 1 & 2 : 1 ratios.

For all ratios, a blue shift in the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ transition of NH₃ upon thermal processing was observed. Factor-group splitting observed in the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ transition of pure NH₃ in Fig. S11† of the ESI was not observed in the mixtures. However, a shoulder at 194 nm was observed on the absorption band of the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ for the 1 : 3 ratio which was not observed in pure NH₃ ice, *vide infra*. The $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$

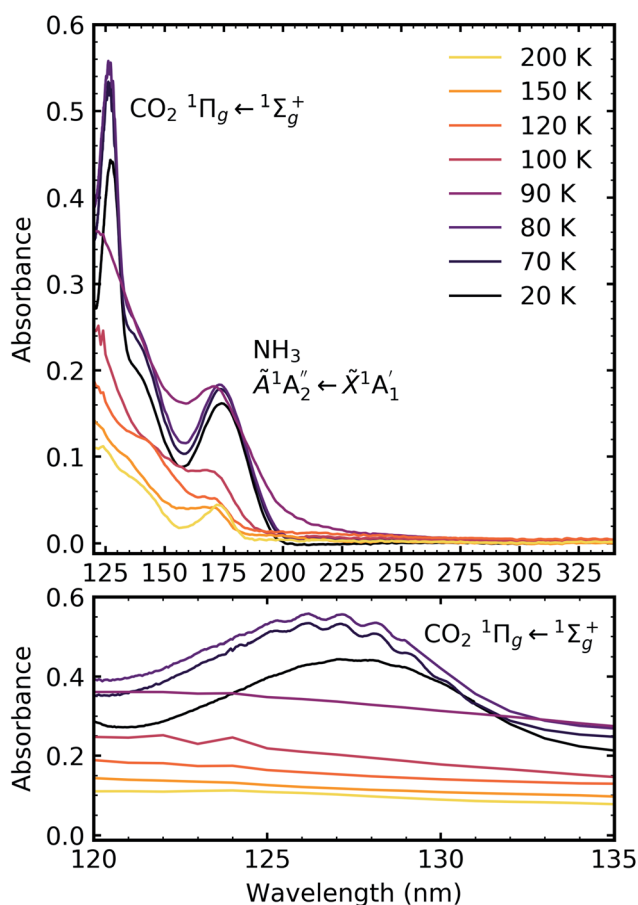


Fig. 7 VUV spectra of the thermal processing of a 4 : 1 CO₂ : NH₃ mixture from 20–200 K between 120–340 nm (top plot). Bottom plot shows the vibrational structure of the CO₂ ${}^1\Pi_g \leftarrow {}^1\Sigma_g^+$ transition between 120–135 nm.

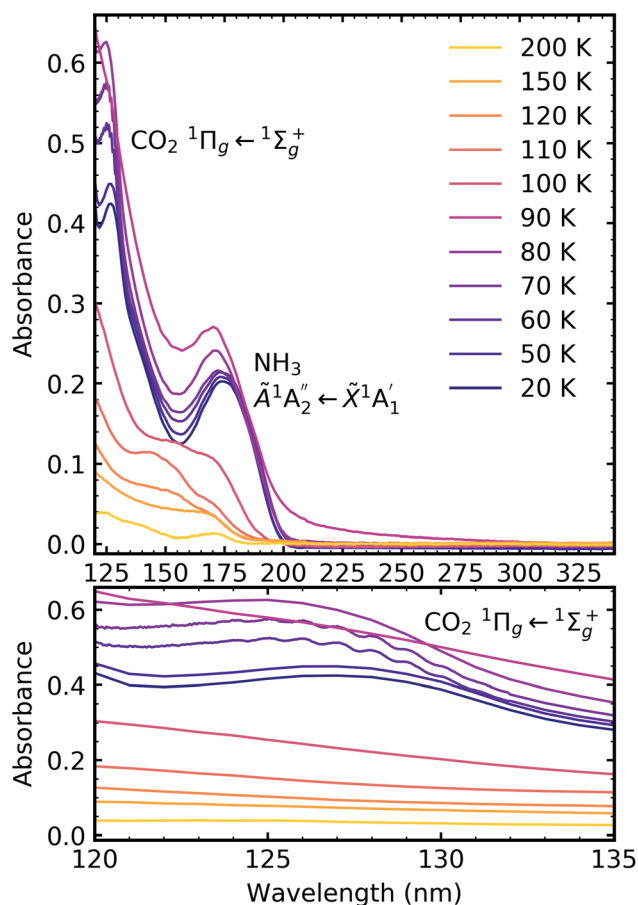


Fig. 8 VUV spectra of the thermal processing of a 2 : 1 CO₂ : NH₃ mixture from 20–250 K between 120–340 nm (top plot). Bottom plot shows the vibrational structure of the CO₂ ${}^1\Pi_g \leftarrow {}^1\Sigma_g^+$ transition between 120–135 nm.

transition of NH₃ disappeared by 100 K revealing new absorption bands due to thermal reaction for all ratios including the 4 : 1 ratio where the equivalent mid-IR 3 : 1 ratio did not have an observable residue.

4.3 Crystallisation of NH₃ in the binary mixtures

It was established in Section 3.2.1 that the phase change that occurred between 60–80 K was dependent on the ratio of the CO₂ : NH₃ mixtures. Mixtures with higher concentrations of CO₂ formed larger crystallites and so had a more crystalline structure compared to mixtures with high concentrations of NH₃.

The subtle morphological changes observed in the mid-IR spectra were more difficult to observe in the VUV spectra. A shoulder observed at 194 nm on the $\tilde{A}^1A''_2 \leftarrow \tilde{X}^1A'_1$ transition of NH₃ in the 1 : 3 ratio was assigned as a Wannier–Mott exciton.²⁰ Wannier–Mott excitons were observed in NH₃ ices deposited above 65 K and are linked to the morphology of the ice with it being most prominent in ices with more crystallites.²⁰

Fig. 10 shows the VUV spectra of the Wannier–Mott exciton peak for the 1 : 3 CO₂ : NH₃ mixture compared to Dawes *et al.* data of pure NH₃ deposited at 75 K and 100 K (ref. 20) and the



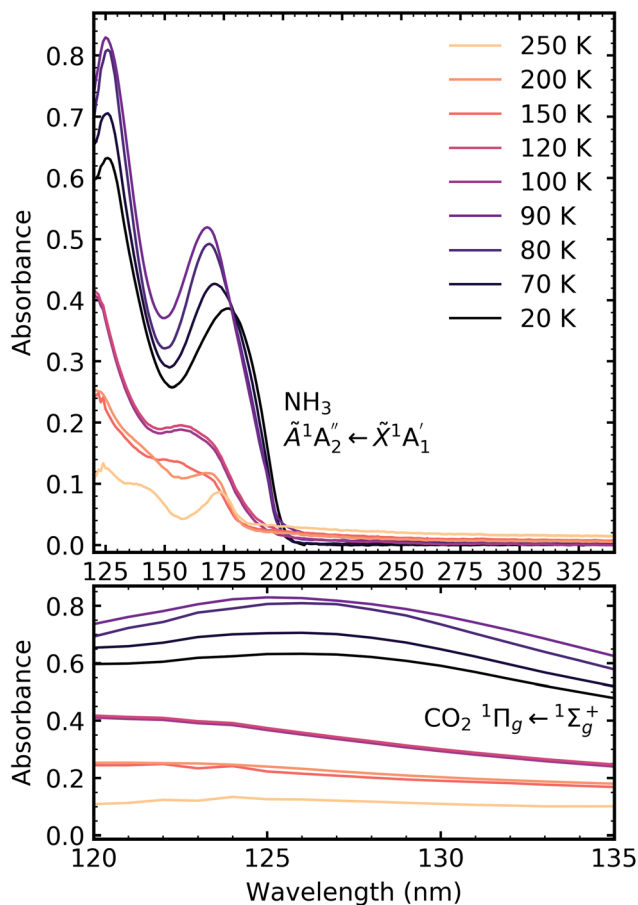


Fig. 9 VUV spectra of the thermal processing of a 1 : 3 CO₂ : NH₃ mixture from 20–200 K between 120–340 nm (top plot). Bottom plot shows the lack of vibrational structure of the CO₂ ¹Π_g ← ¹Σ_g⁺ transition between 120–135 nm.

2 : 1 & 1 : 3 CO₂ : NH₃ mixtures where no Wannier–Mott exciton peak was observed. The exciton peak was much more pronounced for pure NH₃ ice deposited at 75 and 100 K compared to the exciton peak observed in the 1 : 3 ratio as shown in Fig. 10. The intensity of the exciton peak has been linked to NH₃–NH₃ crystallite boundaries which is why we observed an exciton peak in the 1 : 3 ratio. However, the exciton peak in the 1 : 3 ratio was considerably less intense than the exciton peak in pure NH₃ ices shown in Fig. 10. As the mixture exists as regions of segregated NH₃ interspersed between regions of segregated CO₂ ice, there are fewer crystallite grain boundaries for the exciton to propagate similar to a thickness dependence observed in pure NH₃.²⁴

4.4 Residue

Ammonium carbamate and carbamic acid were identified at 150 K and 200 K for all mid-IR ratios apart from the 3 : 1 ratio (Fig. 5) with evidence of thermal conversion of ammonium carbamate to carbamic acid between 150–200 K. Fig. 11 shows the residue VUV spectra at 150 and 200 K for the 4 : 1, 2 : 1 & 1 : 3 ratios. At 150 K all ratios had a peak at ~150 nm which decreased upon thermal processing to 200 K. We have

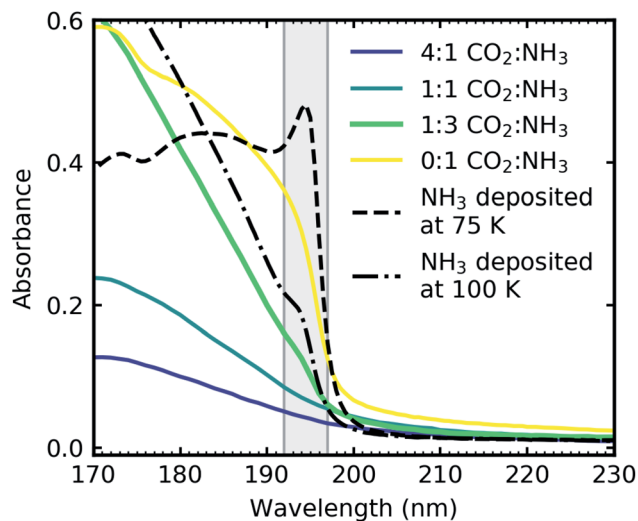


Fig. 10 VUV spectra CO₂ : NH₃ mixtures (4 : 1, 2 : 1 & 1 : 3) deposited at 20 K and thermally processed to 90 K and pure NH₃ deposited at 75 K (dash) and 100 K (dash-dot) from Dawes *et al.*²⁰ The grey shading indicates the area of the Wannier–Mott exciton which was observed in only the 1 : 3 CO₂ : NH₃ mixture (green) and pure NH₃ deposited at 75 K and 100 K. Spectra are normalised to a thickness of 200 nm and the pure NH₃ deposited at 75 and 100 K are further scaled by 0.3.

tentatively assigned this peak to an electronic transition of ammonium carbamate from the analysis of the mid-IR results which showed that vibrational absorption bands associated with ammonium carbamate decreased between 150–200 K. A dash-dot line in Fig. 11 indicates a shoulder at 170 nm which resolved into a more distinct peak upon thermal processing from 150 to 200 K. We tentatively assigned this to an electronic transition of carbamic acid as mid-IR results indicated that carbamic acid formation increased between 150–200 K.

A noticeable difference between the mid-IR and VUV residue spectra was the presence of a residue for the VUV 4 : 1 ratio. The almost equivalent mid-IR 3 : 1 ratio showed no observable residue. However, a residue material could have been present but in trace amounts below the sensitivity of mid-IR spectroscopy. Even the observed mid-IR 2 : 1 & 1 : 1 residues were significantly less intense than the mid-IR 1 : 3 & 1 : 10 ratios indicating less residue material.

4.5 Rayleigh scattering tails

Rayleigh scattering tails have been observed in VUV spectra of several astrophysical ices which did not fully wet the substrate and provided information on the morphology of the ice.^{26,27} Rayleigh scattering tails were observed in the VUV thermal processing spectra (Fig. 7–9). While the ices used in this work were of a thickness where coverage of the surface is expected, if the surface of the ice film was not uniformly smooth then a rough, clumpy surface can scatter the light causing a scattering tail to be observed.

Rayleigh scattering tails occur when particle sizes are less than $\lambda/10$ and the intensity of scattered light (I_s) is proportional λ^{-4} such that:^{28,29}



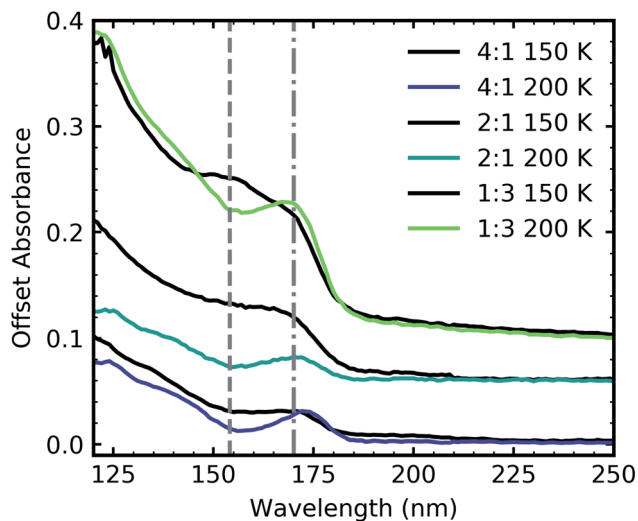


Fig. 11 VUV residue spectra of $\text{CO}_2 : \text{NH}_3$ mixtures (4 : 1, 2 : 1 & 1 : 3) at 150 and 200 K after thermal processing from deposition at 20 K. The dash-dot line highlights a feature at 152 nm and the dashed line highlights a feature at 170 nm. Spectra are normalised to a thickness of 200 nm and offset on the y-axis for clarity.

$$I_s = I_0 \times c \times \lambda^{-4} \quad (1)$$

where I_0 is the incident intensity and c is a constant of proportionality which is dependent on the particle size, the refractive index and the number density of scatterers present within the sample.

In the VUV absorption spectra at $\lambda > 215$ nm, where no absorption peaks were observed, the only contribution to the loss of intensity in transmitted intensity (I_t) was due to scattering. Therefore, by treating the I_t as equal to I_0 minus I_s and using the Beer-Lambert Law, the following simple Rayleigh model was fitted:

$$A = \log_{10} \left(\frac{I_0}{I_0 - I_s} \right) = \log_{10} \left(\frac{1}{1 - c\lambda^{-4}} \right) \quad (2)$$

A more meaningful way of representing the changes observed in the scattering from the ice samples is to calculate the fractional change in the constant of proportionality of the processed ice relative to the constant of proportionality at deposition (Δc):

$$\Delta c = \left(\frac{c_p - c_d}{c_d} \right) \quad (3)$$

where c_p is the constant of proportionality of the processed interstellar ice analogue and c_d is the constant of proportionality of interstellar ice analogue at 20 K.

Fig. 12 shows the thermal evolution of Δc for pure NH_3 and the $\text{CO}_2 : \text{NH}_3$ mixtures. The scattering tails for pure CO_2 ice were outside the Rayleigh regime. The Rayleigh scattering for pure NH_3 and the $\text{CO}_2 : \text{NH}_3$ indicated the presence of particles suggesting that both pure NH_3 and the mixtures did not cover the substrate as uniformly smooth films and instead formed rough ices with 'clumps' on the surface. This probably arose due

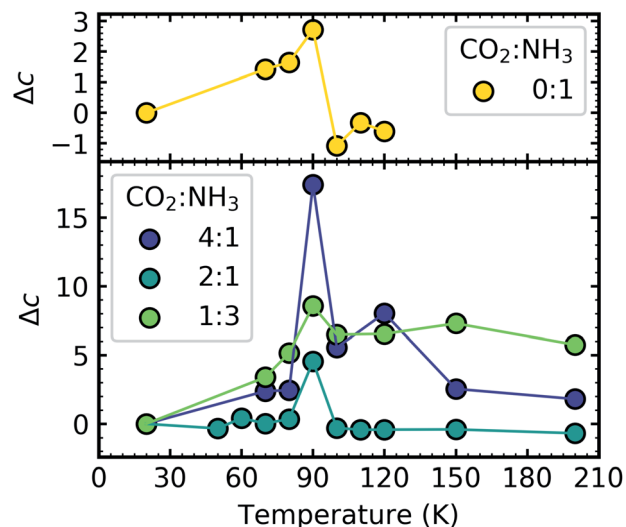


Fig. 12 Comparison of the fractional change (Δc) in the constant of proportionality of the processed ice (c_p) relative to the constant of proportionality at deposition (c_d) versus thermal processing temperature after deposition at 20 K for pure NH_3 ice (0 : 1) and $\text{CO}_2 : \text{NH}_3$ mixtures (3 : 1, 2 : 1, 1 : 1, 1 : 3, 1 : 10). Δc are normalised to a thickness of 200 nm.

to the non-wetting behaviour of NH_3 , investigated previously on Au and amorphous water substrates.³⁰ Single peaks from temperature-programmed desorption studies indicated that multilayers formed due to the preference of NH_3 to bind to neighbouring molecules rather than the substrate. While CO_2 ice is also known to exhibit non-wetting behaviour at low coverages, eventually CO_2 will cover the entire substrate in a uniform film,³¹ with a smooth surface or with particles that are outside of the Rayleigh regime and no Rayleigh scattering tails could be fitted using eqn (2) for CO_2 ice.

For pure NH_3 (0 : 1) and the $\text{CO}_2 : \text{NH}_3$ mixtures a spike in the Δc value at 90 K was observed. We suggest that the spike in the Δc value for NH_3 was due to molecular rearrangement and macroscopic change of morphology initiated by the phase change. However, the spike in the Δc value of the $\text{CO}_2 : \text{NH}_3$ mixtures at 90 K is unlikely to be due to an NH_3 phase change. Table 3 tells us that the observed phase change of NH_3 in the $\text{CO}_2 : \text{NH}_3$ mixtures in the mid-IR study, while ratio-dependent, occurred below 80 K. This was corroborated in Fig. 12 where a slight increase in the value of Δc was observed between 20–80 K. Segregation can also be ruled out as this, while also ratio-dependent, occurred below 80 K. This is further supported by the fact that the largest Δc spike was observed in the 4 : 1 ratio which had segregation and NH_3 phase temperatures below 60 K. We suggest that the spike in the Δc value was due to CO_2 desorption. The 4 : 1 and 2 : 1 ratios have the largest amount of CO_2 in the mixtures and the lowest CO_2 desorption temperatures (Table 3). Whereas, the CO_2 in the 1 : 3 ratio desorbed at the same temperature as NH_3 between 100–110 K likely due to the CO_2 being embedded within the NH_3 ice. Desorption of CO_2 will change the structure of the $\text{CO}_2 : \text{NH}_3$ mixture and possibly the refractive index of the ice which may cause the spike in Δc . The subsequent macroscopic smoothing of the surface could be



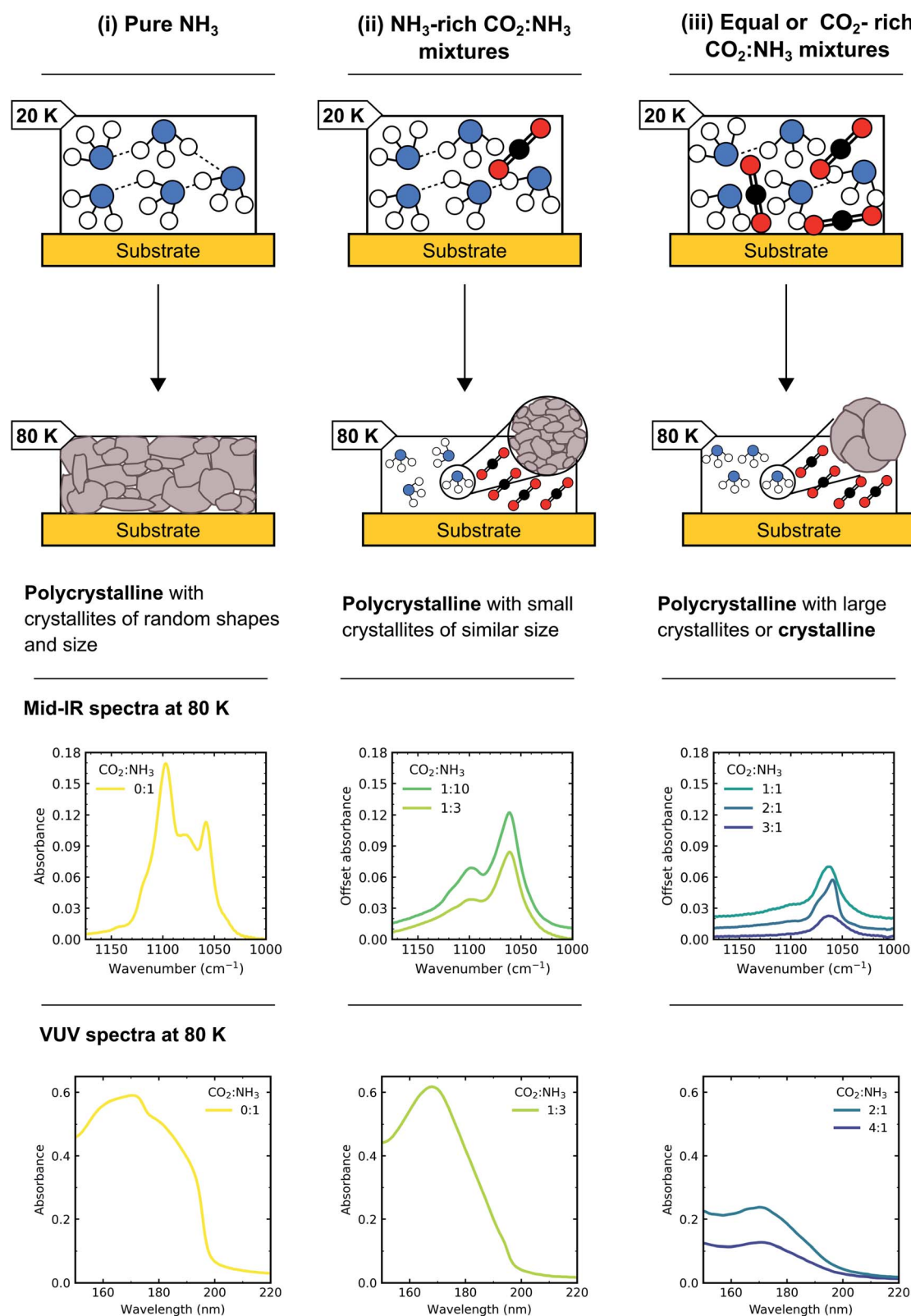


Fig. 13 Schematic diagram of how the morphology of NH₃ changed depending on the ratio of the CO₂ : NH₃ mixture. (i) Pure NH₃ (0 : 1) forms extensive H-bonds at 20 K some of which remain intact upon heating to 80 K so that a polycrystalline structure with crystallites of random shapes and sizes formed. (ii) CO₂ : NH₃ 1 : 10 and 1 : 3 ratios formed less extensive H-bonding than pure NH₃ at 20 K due to the presence of CO₂ and so upon heating to 80 K, a polycrystalline with small crystallites of similar sizes formed. (iii) CO₂ : NH₃ 1 : 1, 2 : 1 and 3 : 1 ratios form even less extensive H-bonding than 1 : 10, 1 : 3 and pure NH₃ at 20 K due to the presence of CO₂ and so upon heating to 80 K a polycrystalline with large crystallites or crystalline structure formed. Partly adapted from Dawes *et al.*²⁰



attributed to the thermal reaction which is initiated around 80 K causing rearrangement of the ice surface and hence a decrease in the Δc value after CO₂ desorption.

5 Discussion

We set out with the aim of demonstrating the impact that one discrete experimental parameter, the stoichiometric mixing ratio, had on the chemical and physical properties of the CO₂ : NH₃ ice system and the subsequent consequence that this may have on thermally induced molecular synthesis.

From our combined mid-IR and VUV spectroscopic results we were able to conclude that CO₂-rich, equal-parts and NH₃-rich mixtures had a thermally induced reaction at ~80 K. A thermally induced reaction at ~80 K was in agreement with several previous studies which deposited CO₂ : NH₃ mixtures at low temperatures (10–20 K).^{3,5,6} For other studies which deposited at higher temperatures^{1,2,7} and/or involved isothermal studies,^{8,9} direct comparison with our results is more difficult when concerning the temperature at which thermal reaction was initiated. For example, Potapov *et al.* reported a lower thermal reaction temperature of 65 K for a KBr substrate.⁹ It is uncertain what the cause of this discrepancy is, although it may be attributed to the different experimental conditions used.

Our residue material at 150–200 K was identified as a mixture of ammonium carbamate and carbamic acid from the literature^{3–7} which underwent a further reaction between 150–200 K as a conversion of ammonium carbamate to carbamic acid.^{6,7} We also identified a ratio-dependence on the amount of residue material present at 150 and 200 K with the largest amount observed in the 1 : 3 ratio and the least amount observed in the CO₂-richest ratios (*i.e.* mid-IR 3 : 1 and VUV 4 : 1). While it could simply be put that ammonium carbamate was the major product at 150 K and to form it requires two NH₃ molecules per every one CO₂ molecule, then stoichiometrically the 1 : 3 ratio was the most ideal. However, we observed differences within both the physical and chemical properties of the CO₂ : NH₃ mixtures which provided a comprehensive understanding of how thermally induced molecular synthesis occurred within this system.

We present the first mid-IR study of CO₂ : NH₃ mixtures with the substrate at an oblique angle with respect to the IR radiation. This allowed us to probe the LO and TO modes of the ν_3 absorption band of CO₂ and further characterise the bonding environment of CO₂ within the mixtures. The absence of a LO mode in the 1 : 10 & 1 : 3 ratio led us to suggest that the CO₂ molecules were largely matrix isolated in an NH₃ matrix. Additional vibrational modes associated with the CO₂ : NH₃ molecular complex in the 1 : 3 ratio also suggested the presence of CO₂ : NH₃ molecular complexes. LO-TO splitting and the asymmetry of the TO absorption bands suggested that the 3 : 1, 2 : 1 & 1 : 1 ratios had bonding environments which included CO₂ dimers, isolated CO₂ and CO₂ : NH₃ molecular complexes.

This detailed characterisation of the CO₂ : NH₃ mixtures at deposition aided our interpretation of the differing NH₃ crystallite structures observed during thermal processing between the CO₂ : NH₃ mixtures. A schematic diagram of the different

NH₃ crystallite structures for pure NH₃ and the CO₂ : NH₃ mixtures is shown in Fig. 13. The NH₃-rich ratios were identified as having smaller crystallites of similar shapes and sizes (*i.e.* more crystallite grain boundaries) and equal-parts and CO₂-rich mixtures were identified as having larger crystallites or crystalline structure (*i.e.* less crystallite grain boundaries).

Further physical change in the CO₂ : NH₃ mixtures were observed in the Rayleigh scattering tails which tells us that the VUV 4 : 1 ratio underwent a large change in the Δc value at the desorption temperature of CO₂ at 90 K. This was observed to a lesser extent in the 2 : 1 ratio and only slightly in the 1 : 3 ratio.

If we return to the differing amounts of residue material in the CO₂ : NH₃ we now have a deeper understanding of the molecular synthesis occurring within the CO₂ : NH₃ mixtures presented in this paper. For example, smaller amounts of residue material were observed in the CO₂-rich and equal part mixtures. These mixtures mainly consisted of CO₂ bonded as CO₂-dimers and to a lesser extent CO₂ : NH₃ molecular complexes and isolated CO₂. Noble *et al.* observed no reaction in their CO₂-rich mixtures and attributed it to high reaction barriers which caused the CO₂ to desorb too quickly before a reaction could take place.⁷ While a comparison between our results and Noble *et al.* must be done with caution as they deposited their mixtures at a much higher deposition temperature (60 K) we do also see a large amount of CO₂ desorption within our CO₂-rich mixtures. We suggest here that this large desorption of CO₂ was that of CO₂-dimers leaving trace amounts of CO₂ : NH₃ molecular complexes and isolated CO₂ to undergo thermal reaction. CO₂-dimers are less prevalent in the mid-IR 1 : 1 ratio compared to the 2 : 1 (& 3 : 1) ratio which may explain the slightly larger amounts of residue material observed for this ratio. Looking at the NH₃-rich mixtures, the 1 : 10 ratio consists mainly of isolated CO₂ and the 1 : 3 ratio consists of isolated CO₂ and CO₂ : NH₃ molecular complexes. This may suggest that the presence of CO₂ : NH₃ molecular complexes within the 1 : 3 ratio enhanced the reactivity of the ice.

However, again it was probably not as simple as that. Larger amounts of residue material from the mid-IR study were observed in ratios with more NH₃ crystallite grain boundaries (1 : 3 & 1 : 10) compared to mixtures with less NH₃ crystallite grain boundaries (1 : 1 & 2 : 1). In non-astrophysical solids, it is well known that diffusion along crystallite grain boundaries is generally enhanced compared to the crystal or mineral equivalent.^{32,33} Experimental data³⁴ and astrochemical models^{35,36} have suggested that structural diffusion within the bulk ice may actually enhance reactivity. This would suggest that CO₂ diffusion along the crystallite grain boundaries was higher in the NH₃-rich mixtures which may also have increased reactivity.

6 Conclusions

We systematically investigated the stoichiometric mixing ratio in CO₂ : NH₃ ices as a function of thermal processing using mid-IR and VUV spectroscopy. This was the first time that CO₂ : NH₃ ice mixtures were studied using VUV spectroscopy which revealed a better sensitivity to the residue material. We



showed that the CO₂ bonding environment within the CO₂:NH₃ mixtures were highly dependent on the stoichiometric mixing ratio and that this bonding environment predetermined the NH₃ crystallite structures within the CO₂:NH₃ mixtures. By understanding the different chemical and physical properties within the CO₂:NH₃ we were able to extend our understanding of the thermally induced reactions in CO₂:NH₃ mixtures. There was a complex interplay between the different physical and chemical properties of the CO₂:NH₃ mixtures that drove the thermally induced molecular synthesis observed in the CO₂:NH₃ mixtures.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The research leading to this result has been supported by the project CALIPSOplus under the Grant Agreement 730872 from the EU Framework Programme for Research and Innovation HORIZON 2020. R. L. J. acknowledges the STFC for her PhD Studentship under grant no. ST/N50421X/1. S. I. acknowledges the Royal Society for financial support. All mid-IR and VUV spectra are available on the ORDO Repository (<https://doi.org/10.21954/ou.rd.12950255>).

Notes and references

- 1 D. Frasco, *J. Chem. Phys.*, 1964, **41**, 2134–2140.
- 2 I. C. Hisatsune, *Can. J. Chem.*, 1983, **62**, 945–948.
- 3 J. B. Bossa, F. Duvernay, P. Theulé, F. Borget and T. Chiavassa, *Chem. Phys.*, 2008, **354**, 211–217.
- 4 M. Bertin, I. Martin, F. Duvernay, P. Theule, J. B. Bossa, F. Borget, E. Illenberger, A. Lafosse, T. Chiavassa and R. Azria, *Phys. Chem. Chem. Phys.*, 2009, **11**, 1383–1845.
- 5 X. Y. Lv, P. Boduch, J. J. Ding, A. Domaracka, T. Langlinay, M. E. Palumbo, H. Rothard and G. Strazzulla, *Phys. Chem. Chem. Phys.*, 2012, **16**, 3433–3441.
- 6 Y. Rodríguez-Lazcano, B. Maté, V. J. Herrero, R. Escrivano and Ó. Gálvez, *Phys. Chem. Chem. Phys.*, 2014, **16**, 3371–3380.
- 7 J. A. Noble, P. Theule, F. Duvernay, G. Danger, T. Chiavassa, P. Ghesquiere, T. Mineva and D. Talbi, *Phys. Chem. Chem. Phys.*, 2014, **16**, 23604–23615.
- 8 A. Potapov, P. Theulé, C. Jäger and T. Henning, *Astrophys. J.*, 2019, **878**, L20.
- 9 A. Potapov, C. Jäger and T. Henning, *Astrophys. J.*, 2020, **894**, 110.
- 10 R. L. James, S. Ioppolo, S. V. Hoffman, N. C. Jones, N. J. Mason and A. Dawes, *Systematic investigation of CO₂:NH₃ ice mixtures using mid-IR and VUV spectroscopy – Part 1: thermal processing dataset*, <https://doi.org/10.21954/ou.rd.12950255>, 2020.
- 11 P. A. Gerakines, W. A. Schutte, J. M. Greenberg and E. F. van Dishoeck, *Astron. Astrophys.*, 1995, **296**, 810.
- 12 G. A. Baratta and M. E. Palumbo, *J. Opt. Soc. Am. A*, 1998, **15**, 3076–3085.
- 13 F. P. Reding and D. F. Hornig, *J. Chem. Phys.*, 1951, **19**, 594–601.
- 14 B. Jonsson and B. Nelander, *Chem. Phys.*, 1977, **25**, 263–269.
- 15 G. T. Fraser, K. R. Leopold and W. Klemperer, *J. Chem. Phys.*, 1984, **81**, 2577–2584.
- 16 R. M. Escrivano, G. M. Muñoz Caro, G. A. Cruz-Díaz, Y. Rodríguez-Lazcano and B. Maté, *Proc. Natl. Acad. Sci. U. S. A.*, 2013, **110**, 12899–12904.
- 17 I. R. Cooke, E. C. Fayolle and K. I. Öberg, *Astrophys. J.*, 2016, **832**, 5.
- 18 B. Sivaraman, B. N. Raja Sekhar, D. Fulvio, A. Hunniford, B. McCullough, M. E. Palumbo and N. Mason, *J. Chem. Phys.*, 2013, **139**, 074706.
- 19 G. A. Baratta and M. E. Palumbo, *Astron. Astrophys.*, 2017, **608**, A81.
- 20 A. Dawes, R. J. Mukerji, M. P. Davis, P. D. Holtom, S. M. Webb, B. Sivaraman, S. V. Hoffmann, D. A. Shaw and N. J. Mason, *J. Chem. Phys.*, 2007, **126**, 244711.
- 21 W. Zheng and R. I. Kaiser, *Chem. Phys. Lett.*, 2007, **440**, 229–234.
- 22 A. Cassidy, R. L. James, A. Dawes and D. Field, *ChemistryOpen*, 2020, **9**, 983–990.
- 23 K. M. Monahan and W. C. Walker, *J. Chem. Phys.*, 1974, **61**, 3886–3889.
- 24 N. J. Mason, A. Dawes, P. D. Holtom, R. J. Mukerji, M. P. Davis, B. Sivaraman, R. I. Kaiser, S. V. Hoffmann and D. A. Shaw, *Faraday Discuss.*, 2006, **133**, 311–329.
- 25 H.-C. Lu, H.-K. Chen, B.-M. Cheng and J. Ogilvie, *Spectrochim. Acta, Part A*, 2008, **71**, 1485–1491.
- 26 A. Dawes, N. Pascual, S. V. Hoffmann, N. C. Jones and N. J. Mason, *Phys. Chem. Chem. Phys.*, 2017, **19**, 27544–27555.
- 27 A. Dawes, N. Pascual, N. J. Mason, S. Gärtner, S. V. Hoffmann and N. C. Jones, *Phys. Chem. Chem. Phys.*, 2018, **20**, 15273–15287.
- 28 J. D. Jackson, *Classical Electrodynamics*, Wiley, New York, 2nd edn, 1975.
- 29 D. J. Lockwood, in *Rayleigh and Mie Scattering*, ed. M. R. Luo, Springer New York, New York, NY, 2016, pp. 1097–1107.
- 30 M. P. Collings, M. A. Anderson, R. Chen, J. W. Dever, S. Viti, D. A. Williams and M. R. S. McMoustra, *Mon. Not. R. Astron. Soc.*, 2004, **354**, 1133–1140.
- 31 J. A. Noble, E. Congiu, F. Dulieu and H. J. Fraser, *Mon. Not. R. Astron. Soc.*, 2012, **421**, 768–779.
- 32 R. Dohmen and R. Milke, *Rev. Mineral. Geochem.*, 2010, **72**, 921–970.
- 33 G. Stechauner and E. Kozeschnik, *J. Mater. Eng. Perform.*, 2014, **23**, 1576–1579.
- 34 P. Ghesquiere, A. Ivlev, J. A. Noble and P. Theulé, *Astron. Astrophys.*, 2018, **614**, A107.
- 35 C. N. Shingledecker, A. Vasyunin, E. Herbst and P. Caselli, *Astrophys. J.*, 2019, **876**, 140.
- 36 C. N. Shingledecker, T. Lamberts, J. C. Laas, A. Vasyunin, E. Herbst, J. Kästner and P. Caselli, *Astrophys. J.*, 2020, **888**, 52.

