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Modified top-down approach for synthesis of molybdenum oxide quantum dots: sonication induced chemical etching of thin films

 Dibya Jyoti Borah,^a Abu Talat Tahir Mostako,^b  ^{*,a} Angshuman Thunder Borgogoi,^a Prasanta Kumar Saikia^b and Ashim Malakar^c

A simple and modified top-down approach to synthesize molybdenum oxide (MoO_x; x = 2, 3) quantum dots (QDs) is proposed in this study. This modified approach involves the conversion of a bulk powder material into thin films followed by a sonication induced chemical etching process for synthesising QDs. X-Ray Diffraction (XRD) is used for crystal structural characterization of MoO_x thin films. The crystal structure properties of the MoO_x QDs are analysed by High Resolution Transmission Electron Microscopy (HRTEM) images and corresponding Selected Area Electron Diffraction (SAED) patterns. The optical band gap is estimated by Tauc's plot from UV-Vis-NIR absorption spectra. The excitation dependent photoluminescence (PL) emission of MoO_x QDs as a function of acid concentration is investigated. The growth mechanism of QDs in different crystalline phases as a function of acid concentration is also exemplified in this work. The micro-Raman and Fourier Transform of Infrared (FTIR) spectra are recorded to analyse the vibrational spectrum of the molybdenum–oxygen (Mo–O) bonds in the MoO_x QDs.

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Introduction

Transition Metal Oxide (TMO) nanostructures have shown great promise and are being investigated with great attention for basic scientific and industrial research owing to their unique physical and chemical properties.^{1–3} Among TMOs, molybdenum oxide (MoO_x) based nano-materials with different morphologies have attracted great attention in recent years due to their optical, electronic, and structural features.^{4–13} One dimensional (1D) and two dimensional (2D) nano-belts,^{4,5} nano-rods,^{3,6} nano-wires,^{7,8} nano-flowers,⁹ nano-ribbons,¹⁰ nano-sheets,^{11,12} and nano-films¹³ have stimulated great research interest in recent years. MoO_x nano-materials have been pioneering through their uses in solar cells,^{14,15} lithium ion batteries,^{16,17} field effect transistors,¹⁸ gas sensors,¹⁹ antiseptics and anticancer treatment.^{20,21} Zero dimensional MoO_x QDs exhibit enhanced optical and electronic properties due to the quantum confinement effect, which originates from their small size (~10 nm) compared to their 1D or 2D morphology.^{22–24} Low toxicity, superior photostability, biocompatibility and excellent chemical stability, enables MoO_x QDs a potential candidate for

metal ion sensing, small biomolecule detection, biological labelling and imaging.²⁵ Moreover, oxygen deficient MoO_{3–x} (2 < x < 3) QDs used for photothermal ablation of cancer cells due to strong near infrared (NIR) absorption and high photothermal conversion efficiency.^{26,27} MoO_x QDs containing in solution phase of hydrochloric acid (HCl) have a potential application in photothermal treatment and biological sensing due to low cost production and high sensitivity.^{26,27} This solution phase of QDs are also applied in rapid colorimetric detection of glucose.²⁶ MoO_x nano-materials can be well extracted from its solution phase. T. T. P. Pham *et al.* reported that the heat treatment at elevated temperature ~150 °C for 1 hour duration eliminates HCl acid and water from the solution containing MoO_x, and are able to extract MoO_x nano-materials in powder form.²⁸ As received powder of MoO_x nano-materials exhibited excellent catalytic properties in partial oxidation of methanol to formaldehyde.²⁸ As an advance branch of low dimensional materials, 2D QDs derived from atomically-thin 2D sheets *viz*; graphene, transition metal dichalcogenide, graphitic carbon nitride, hexagonal boron nitride, and phosphorene, are emerging extraordinary low dimensional materials.^{23,29} The 2D QDs have wide range of applications including bio-imaging, photo-induced therapy, chemical and metal ion sensors, and photo-voltaic devices.^{23,29}

Generally, the synthesis of MoO_x QDs can be classified in two approaches: top-down and bottom-up method.^{26,30} Top-down approach mainly relies on weakening the van der Waals interaction between adjacent layers and breaking the strong covalent bonding within each layer to exfoliate the structure or bulk

^aMaterial Science Laboratory, Department of Physics, Dibrugarh University, Dibrugarh-786004, Assam, India. E-mail: mostako@dibru.ac.in; abumostako@gmail.com

^bThin Film Laboratory, Department of Physics, Dibrugarh University, Dibrugarh-786004, Assam, India

^cCentral Instrumental Facility, Indian Institute of Technology Guwahati, Guwahati-39, India



materials.^{26,31} Alternatively, QDs of layered materials prepared by a top-down approach do not require an insulating ligand coating, which is deleterious to electrical transport of the materials, especially in the application of catalysis and electronics.^{22,31} Top-down approaches, such as laser ablation,³² e-beam lithography,³³ can be used to reduce the particle size. Unfortunately, such methods are weakly feasible for the high yield preparation of QDs in solution phase and are also restricted due to the need of extremely expensive instrumental facilities. QDs synthesis by bottom-up approaches, suffer from insulating ligand coating, which affect the electrical transport of properties of the material.³¹ In the meantime, Huang *et al.* reported that MoO_{3-x} nanotubes have been successfully synthesized *via* hydrothermal method in presence of thiol by means of bioligand and surfactant.³⁴ However, the poor hydrophilicity of the resulting products owing to the difficulties experienced for complete removal the surfactant or organic ligand molecules. A combination of sonication and solvothermal treatment of bulk MoS₂/WS₂ was used by Xu *et al.* for MoS₂/WS₂ QDs synthesis.³⁵ But, requirement of high temperature and frequent sonication resulted breaking of bulk crystals down to nano-sheets indicates further refinement.^{36,37} Difficulties persisted in the existing synthesis procedure of TMO QDs as per present literature survey depicts the need of alternative approach for the synthesis without any incongruities. Therefore, in this manuscript a simple and modified top-down method is proposed for synthesis of MoO_x QDs. This method is a combination of thermal evaporation and sonication induced chemical etching process. Every synthesis method needs a careful optimization of various parameters involved in that method. Sonication is an effective tool that can avoid the formation of large clusters and aggregation of MoO₃ nanoparticles. Sonication frequency, duration and temperature of the sonication bath has a vital role on the growth process of QDs. S. Bai *et al.* and Z. Wu *et al.* have reported that the increase in sonication duration produces more smoother and regular particles with better crystallinity.^{19,38} Keeping these in mind, at fixed sonication duration and sonication temperature, the effect of parametric variation of etchant concentration on the synthesis and their effect on the optical and structural properties of MoO_x QDs is investigated in the present work.

Experimental section

Materials and reagents

Molybdenum(vi) oxide powder (Sigma-Aldrich, 99.97% purity) was used as source material. Hydrochloric acid (Emplura®, ≥ 35%), nitric acid (Emplura®, 68–70%) and acetone (Emplura®, ≥ 99.0%) were of analytical grade and used without further purification. Glass microscope slides (76.2 mm × 25.4 mm × 1.2 mm) were used as substrates for thin film fabrication. Throughout all the experiments, deionized water was used.

Synthesis of MoO_x QDs

MoO_x QDs were synthesized from molybdenum oxide powder in two step process. In the first step: an amount of 0.1 g

molybdenum(vi) oxide powder is evaporated from molybdenum boat by supplying a current ~3.0–3.4 A in a thermal evaporation setup (Model: HINDHIVAC 12" vacuum coating unit, 12A4D). The duration of deposition is ~10 min. MoO_x films are deposited on clean glass substrate (5.52 cm × 2.54 cm) by thermal evaporation technique at substrate temperature ~100 °C and base pressure ~2 × 10⁻⁵ torr. Prior to thin film deposition, the glass substrates were cleaned by usual cleaning protocols as described in our previous work.³⁹ The second step depicts the chemical etching of as-deposited MoO_x thin films. Prior to chemical etching, a 4 mL 25% diluted hydrochloric acid (HCl) was prepared in a glass breaker of 10 mL. This 4 mL solution was prepared by mixing 3 mL concentrated HCl (~35%) and 1 mL deionized water. During the process of chemical etching, the breaker containing the mixture solution and submerged film in it, was partially dip into the water bath of the sonicator and continuously sonicated with frequency ~50 Hz for 30 minutes. The temperature of the water bath was maintained at 45 °C. After that, a light yellow colour solution is obtained. In this work, two samples are prepared by changing the concentration of HCl acid from 25% to 30%. The samples are coded as S₁ and S₂, respectively.

Characterizations

The crystal structural properties of MoO_x film is characterized by XRD (Model: Rigaku Ultima (IV)), using Cu-K_α monochromatic radiation ($\lambda = 1.541 \text{ \AA}$) with a potential of 40 kV and current of 30 mA. The 2θ angle of the detector was scanned typically in between 10° and 70° with a step size of 0.15°. The size of MoO_x particles are investigated by field emission transmission electron microscope (Model: JEOL 2100F) image analysis, operated with an acceleration voltage of electron beam ~200 kV. HRTEM images were captured to observe the nanoparticle formation while the SAED pattern were recorded for analysing the modification in crystalline structure of the QDs as a function of HCl acid concentration. The average particle size and lattice spacing are estimated from fringing pattern in the HRTEM images by using Gatan Digital Micrograph software. For spectroscopic analysis, the absorption, fluorescence, Raman and FTIR spectra of the MoO_x QDs are recorded. Absorption spectra are recorded by using a UV-Vis-NIR Spectrophotometer (Model: Shimadzu UV-3101 PC) in wavelength range, 200–3000 nm with an interval of 1 nm. The optical band gap of MoO_x QDs is calculated from absorption spectra by using Tauc's plots. Fluorescence spectra of QDs are recorded by using a Spectrophotometer (Model: Horiba Scientific Fluoromax-4 Spectrofluorometer) at different excitation wavelengths 260, 270, 300, 320, 350 nm and 280, 300, 320, 350, 380 nm, for the sample S₁ and S₂, respectively. Raman spectra of the QDs were recorded at room temperature using laser micro-Raman setup (Model: Horiba Jobin Vyon, Lab-Ram HR 800) under 532 nm wavelength as an excitation source. Fourier Transform Infrared (FTIR) absorption spectra are recorded by FTIR Spectrometer (Model: PerkinElmer (BX)) to study the bonding nature of Mo–O in the MoO_x QDs. For FTIR studies, a small amount of sample solution was dried onto silicon substrate.



Results and discussion

Crystal structure and particle size analysis

The synthesis procedure of MoO_x QDs from bulk molybdenum oxide powder is shown schematically in Fig. 1. Thermally deposited MoO_x layers are etching slowly into small particles. These particles further fragmented to even more small particles by ultra-sonication, which lead to the formation of well dispersed MoO_x QDs. The MoO_x thin films are deposited on glass substrate at 6 cm source-substrate distance and room temperature. The XRD spectrum of bare glass and thermally deposited MoO_x thin film are shown in Fig. 2.

The broad XRD peak ranging from $2\theta \sim 20^\circ$ to 30° is observed in the MoO_x film due to the glass substrate only. No significant XRD peaks are observed in the MoO_x film. This may be attributed to the low mobility of evaporated atoms on the substrate surface and less interaction between the molybdenum and oxygen since this film is deposited at room temperature.³⁹ Similar thin film samples are prepared under same deposition parameters that are subjected to chemical etching process for synthesis of QDs.

FETEM image analysis is carried out to investigate the size of the particles and crystallographic orientation of prepared QDs. FETEM image of the sample S₁ is shown in Fig. 3(a). It confirmed the formation of nearly spherical shape of MoO_x QDs of sample S₁. The size distribution histogram gives the average size among thirty QDs is ~ 4.3 nm, shown in Fig. 3(b) for the sample S₁. The Fig. 3(c) shows the bright field image of corresponding SAED pattern. The SAED pattern revealed that QDs are polycrystalline in nature. There are two lattice planes each (218) and (420) corresponding to h-MoO₃ phase (JCPDS: 21-0569) and that of (200) and (060) corresponding to α -MoO₃ phase (JCPDS: 05-0508), respectively. The HRTEM images, Fig. 4(a)–(d) reveals that lattice spacing of QDs are ~ 0.20 nm correspond to (320) lattice plane of h-MoO₃ phase (JCPDS: 21-0569) and ~ 0.22 nm, 0.23 nm and 0.19 nm correspond to (150), (060) and (200) lattice planes of α -MoO₃ phase (JCPDS: 05-0508), respectively.

Moreover, FETEM image confirmed the nearly spherical shape of MoO_x QDs for the sample S₂, as shown in Fig. 5(a)–(d). The size distribution histogram gives the average size ~ 4.6 nm (Fig. 5(b)). The SAED pattern, Fig. 5(c), revealed that these QDs are also polycrystalline in nature. Highly ordered and parallel

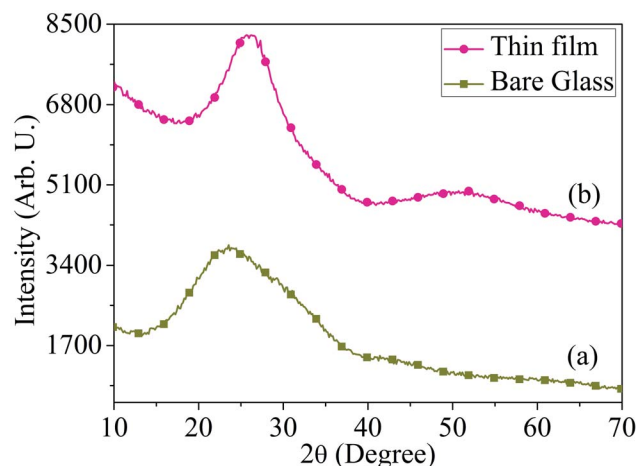


Fig. 2 XRD spectrum of (a) bare glass and (b) thermally deposited MoO_x thin film, respectively.

lattices in the HRTEM image, Fig. 5(d) of the QDs suggested its crystallization with lattice spacing ~ 0.26 nm, which is associated with (140) lattice plane of α -MoO₃ phase (JCPDS: 05-0508). For the samples: S₁ and S₂, the inter planner spacing along with crystal planes observed in HRTEM image and estimated from SAED pattern are listed in Table 1. These studies indicate that the small nano-crystals are QDs of MoO₃. The bulk MoO₃ powder can be transformed to the QDs with tuneable phase and size with the careful optimization of various parameters involved in this modified technique. In addition, by this modified technique amorphous MoO_x thin film is successfully transformed into the polycrystalline QDs with co-existence of α -MoO₃ and h-MoO₃ phase, as confirmed by XRD and SAED patterns analysis. There are several reports on the synthesis of MoO_x nano-materials in presence of surfactant.^{4,19,40} However, in the present work a well dispersed stable MoO_x QDs are successfully synthesized without any surfactant. The observed polydispersity of the particle size distribution can be minimized with the use of suitable capping or stabilizing agent during the chemical etching process.

The increase of size of the particles with the increase of solution concentration (more than two sample concentration) were reported by H. S. Dehsari *et al.* and J. H. Huang *et al.*^{41,42} In

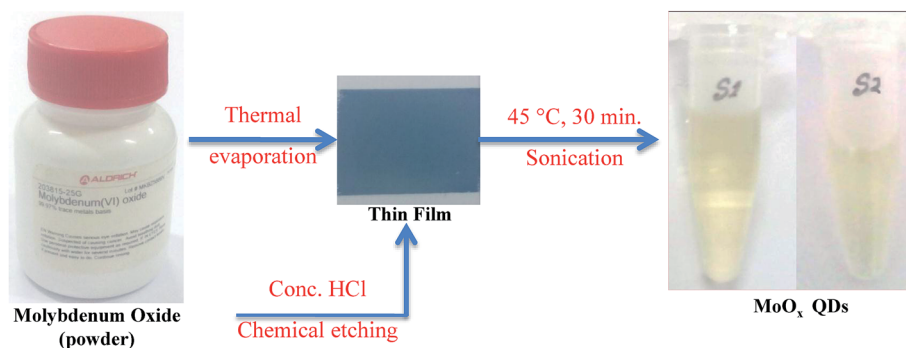


Fig. 1 Schematic representation of synthesis of MoO_x QDs from bulk molybdenum oxide powder.



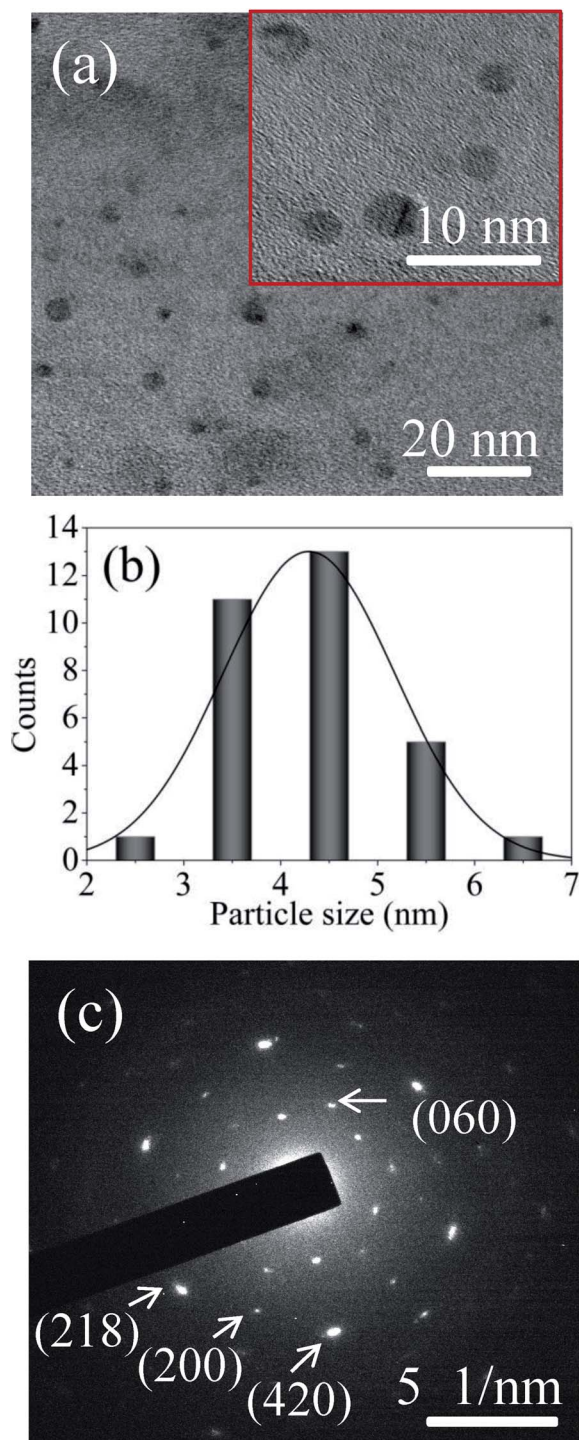


Fig. 3 (a) FETEM image of the sample S_1 (inset magnified area), (b) the corresponding particle size distribution histogram, and (c) SAED pattern of MoO_x QDs for the sample S_1 .

the present investigation, similar increase in the size of QDs with the increase in HCl acid concentration in the sample solution is also observed. It is because, only few particles are available at the lower concentration of the solution and they tend to stay apart due to strong repulsive force. This reduces the growth rate due to less mass transfer in the reaction

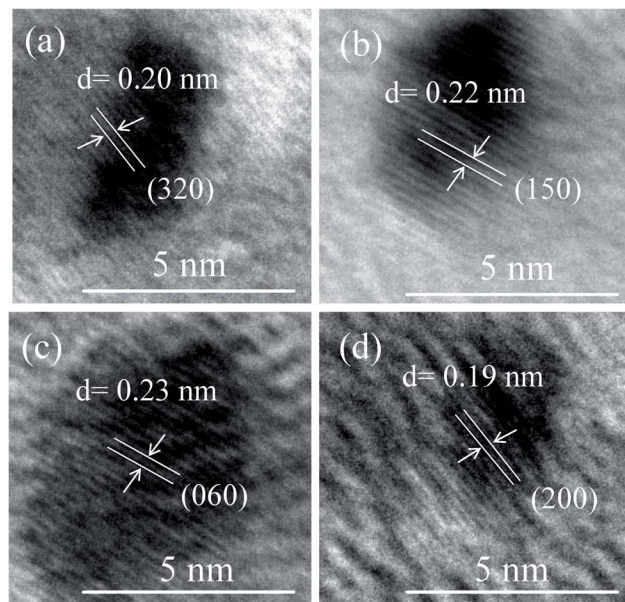


Fig. 4 (a–d) The HRTEM image of four different MoO_x QDs of the sample S_1 .

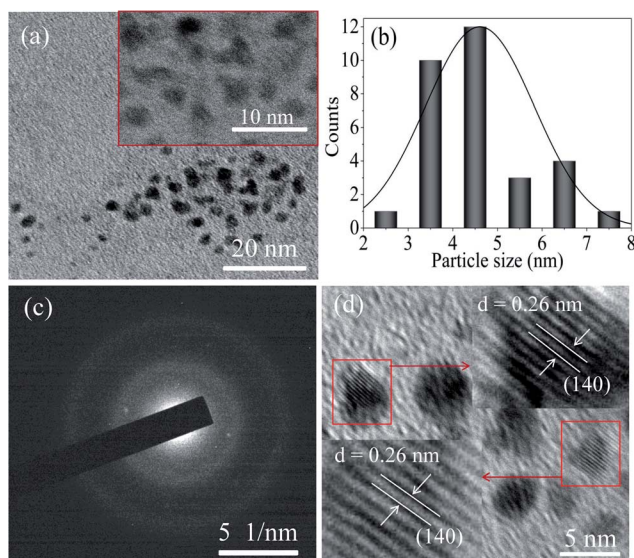


Fig. 5 (a) FETEM image of the sample S_2 (inset magnified area), (b) the corresponding particle size distribution histogram, (c) the SAED pattern of MoO_x QDs, and (d) the HRTEM image of MoO_x QDs for the sample S_2 .

medium.^{41,42} However, at higher concentration the number density of the particles is increased which in turn reduced the inter-particle distances. There is a high mass transfer at low inter-particle distance and thereby causing the higher growth rate of the QDs.^{41,42} The QDs are synthesised at lower HCl acid concentration showed the co-existence of $h\text{-MoO}_3$ and $\alpha\text{-MoO}_3$ phases. It is also observed that these mixed phase transformed to thermodynamically stable $\alpha\text{-MoO}_3$ phase with increase of HCl acid concentration from 25% to 30% in the solution. It is



Table 1 Estimated optical and structural parameters of MoO_x QDs along with the sample codes

Sample code	Concentration of HCl acid (%)	Band gap energy (eV)	Average particle size (nm)	Inter planner spacing, <i>d</i> (nm)	(<i>hkl</i>) planes	Phase assignment
S ₁	25	~5.01	~4.3	0.20	(320)	h-MoO ₃
				0.22	(150)	h-MoO ₃
				0.23	(060)	α-MoO ₃
				0.19	(200)	α-MoO ₃
S ₂	30	~4.98	~4.6	0.26	(140)	α-MoO ₃
				0.26	(140)	α-MoO ₃

because, at lower concentration of HCl acid, only few H⁺ ions were adsorbed on the negative polar plane of growth unit of the MoO₃ *i.e.*, MoO₆ octahedron. Thus, the corner to corner growth became the basis of the growth unit and it easily developed into a network shaped h-MoO₃ phase.⁴³ With the increase of HCl acid concentration in the sample solution, the concentration of H⁺ ion increased, which play a crucial role in MoO₃ phase transformation. The growth rate of the negative polar plane of MoO₆ octahedron is blocked, due to more and more H⁺ ions were adsorbed at the negative polar plane. Instead, the growth rate of the positive polar plane of MoO₆ octahedron was far greater than others. Therefore, the growth is changed to superposition by sharing the edges and by sharing the corners of the positive polar plane structure, which developed into an α-MoO₃ phase.⁴³ In addition, the lower concentration of HCl acid will favour the formation of h-MoO₃ phase.⁴⁴

Spectral analysis

Absorption characteristics. The UV-Vis absorption spectra of MoO_x QDs in the wavelength range, 200–400 nm for different concentration of HCl acid as shown in Fig. 6. The spectra confirm that the as-prepared MoO_x QDs shows characteristic UV-Vis absorption peak at ~219 nm and strong absorption between 200–300 nm. The strong absorption is due to the charge transfer of Mo–O band in the MoO₆⁶⁻ octahedron.^{40,45} The

broadening of the absorption peak is ~22 nm and ~23 nm for the sample S₁ and S₂, respectively. The change in full width at half-maximum (FWHM) broadening with the change in HCl acid concentration 25% to 30% is observed to be marginal. The narrow FWHM data of QDs reflects their smaller particle size distribution.^{46,47} This is consistent with the present findings. The light yellow colour of the solution of MoO_x QDs, as shown in Fig. 1, confirmed the formation of molybdenum trioxide (MoO₃).^{26,48,49} The absence of distinct precipitation of the yellow solution indicates the fine disparity and stability of MoO_x QDs. The absorbance is increased to ~1.2 for sample S₂ in comparison to that of S₁, ~0.9 along with bathchromic shift of ~8 nm at the wavelength 230 nm. Higher HCl acid concentration in sample S₂ produces more number of particles in the solution. This in turn led to the increase in absorbance. In addition, the red shift of absorption edge of sample S₂ in comparison to that of S₁ indicates the increase of particle size, which is in accordance with the result obtained from FETEM characterization. The optical band gaps (*E_g*) of QDs in the present work are estimated by using the eqn (1),^{50,51}

$$\alpha h\nu = B(h\nu - E_g)^m \quad (1)$$

where, '*B*' is the Tauc's slope, '*hν*' is the photon energy, '*m*' is the transition probability. The value of '*m*' is $\frac{1}{2}$ for allowed transition in direct band gap material.^{50,51} The $(\alpha h\nu)^2 - h\nu$ plot for both sample: S₁ and S₂, as shown in Fig. 7. The optical band gap is estimated by linear extrapolation of the $(\alpha h\nu)^2 - h\nu$ plot. The

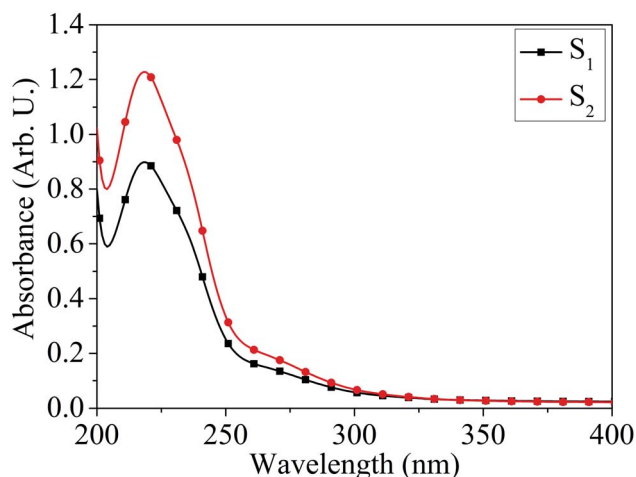


Fig. 6 UV-Vis absorption spectra of MoO_x QDs for the samples: S₁ and S₂.

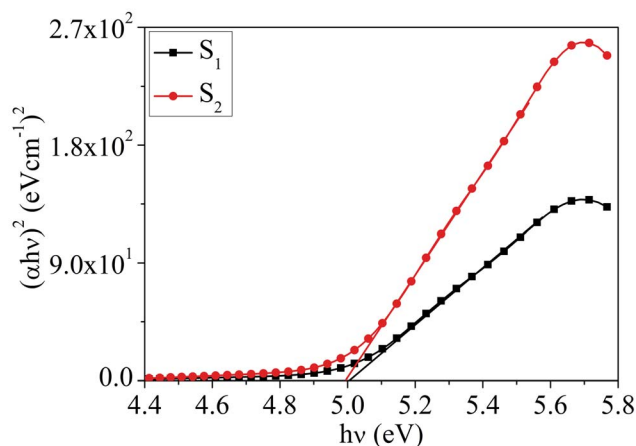


Fig. 7 The plot of $(\alpha h\nu)^2 - h\nu$ for the samples: S₁ and S₂.



calculated values of band gap are listed in Table 1. The optical band gap of the samples: S_1 and S_2 are ~ 5.01 eV and 4.98 eV, respectively. The change in optical band gap is marginal with the present change in acid concentration. The band gap energy estimation shows there is blue shift for the MoO_x QDs compared to that of bulk MoO_3 (~ 2.9 eV). This blue shift is due to the quantum confinement effect and hence, it reconfirmed the formation of QDs.^{52,53} The histogram of band gap and average particle size of MoO_x QDs for different acid concentration is shown in Fig. 8.

Photoluminescence (PL) characteristics. Time resolved PL was measured by a Life Spec-II, Edinburgh Instruments. The fluorescence decay curve of the sample S_2 is shown in Fig. 9(a). The residual counts of the fluorescence decay curve is shown in Fig. 9(b). A 308 nm LED from Pico Quant was employed as the excitation source with a Hamamatsu Micro Channel Plate (MCP) as the detector. Using the FAST software provided by Edinburgh Instruments Ltd, the fluorescence decays were analysed by the reconvolution method. The goodness of fit is determined by the reduced chi-square (χ^2) values and weighted residuals. The fluorescence decay of both the samples: S_1 and S_2 show lifetime of 1.40 and 1.42 ns, respectively and listed in Table 2. These results are quite indicative of the fact that the PL emissions of the QDs are fluorescent in nature. The fluorescence decay of the sample S_1 and S_2 show a single exponential decay as suggested by good χ^2 values.

The PL spectra of MoO_x QDs for the samples: S_1 and S_2 shown in Fig. 10(a) and (b), respectively. Similar to that of nanoparticles, MoO_x QDs also exhibit excitation dependent PL behaviour. For the sample S_1 , PL spectra of MoO_x QDs exhibit a strong emission peak at ~ 360 nm under an excitation wavelength 300 nm. This fluorescence emission peak at ~ 360 nm is due to the band to band transition and oxygen vacancies, Mo interstitial and surface defect of MoO_x QDs. In addition, the emission peak at ~ 363 nm is related to the indigo emission corresponding to the near band edge emission and it is close to the reported value of 365 nm.^{54,55} Similarly for the sample S_2 , the strong fluorescence emission peak appears at ~ 410 nm upon excitation with 350 nm wavelength. As per crystal field model,

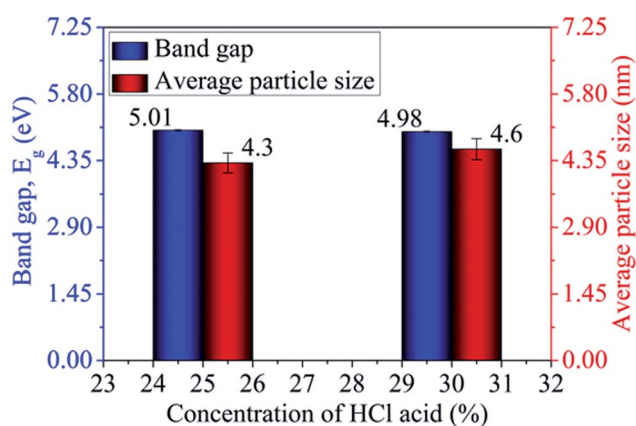


Fig. 8 Histogram of band gap and average particle size of MoO_x QDs for different concentration of HCl acid.

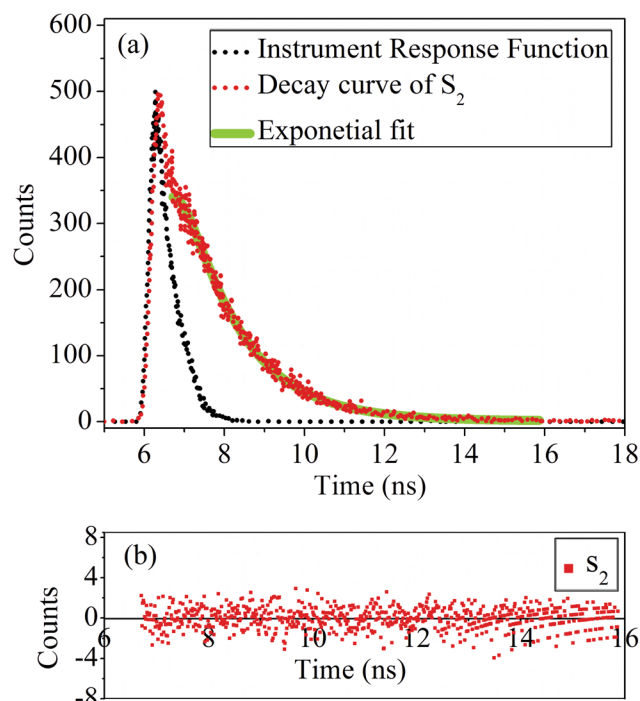


Fig. 9 (a) Fluorescence decay curve and (b) corresponding residual plot of the sample S_2 .

the fluorescence emission peak at ~ 410 nm may be attributed due to the $\text{Mo}^{5+} d_{yz^2} - d_{xy^2}$ band transition of a heavily distorted polyhedron (Mo-O) in an octahedral crystal field.⁵² The blue emission peaks at ~ 420 and 400 nm are mostly related to the band to band transition with high Stokes shift.⁵⁶⁻⁵⁸ Digging more into this experiment, PL spectra for both the samples were recorded at various excitation wavelengths. In these PL spectra, a very weak shoulders along with the strong fluorescence emission peaks are detected for both samples under different excitation wavelengths. The weak shoulders are appeared due to the electron-hole recombination between the conduction band and the sublevel of adsorbed oxygen acceptors; which are disappeared with the increase in excitation wavelengths.⁵⁹ The fluorescence emission peaks bathochromically shifted from ~ 363 to 360 nm for the sample S_1 with an increase in the excitation wavelength from 260 to 350 nm. The fluorescence intensity was observed to change with excitation wavelengths. Similar bathochromic shift from ~ 420 to 410 nm was also observed for the sample S_2 with the change in excitation wavelength from 280 to 380 nm. This phenomenon usually appears in QDs that derived from 2D materials.^{40,60-63} This is because of

Table 2 Fluorescence lifetimes of MoO_x QDs along with the sample codes

Fluorophore	Monitoring λ_{max}	Lifetime (ns)	χ^2
S_1	360 nm	1.40	1.06
S_2	360 nm	1.42	1.05



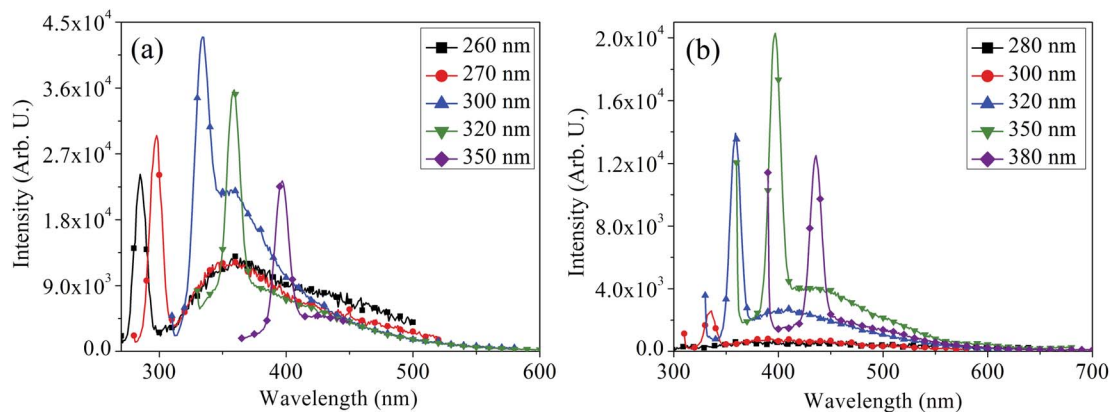


Fig. 10 Excitation dependent PL spectra of MoO_x QDs for the samples (a) S₁ and (b) S₂, respectively.

the inhomogeneity of chemical components and polydispersity of the lateral sizes,⁶⁴ which is again in consistent with FETEM observation. This phenomenon mostly likely due to the distribution of emissive trap sites on the surface of MoO_x QDs and quantum confinement effect.⁶⁰

Micro-Raman characteristics. In order to further validate the existence of MoO_x material, the Raman spectroscopy is used. The micro-Raman spectra recorded in the wavenumber range, 50–4000 cm⁻¹ for different HCl acid concentration is shown in Fig. 11. The micro-Raman peaks for α-MoO₃ phase of the sample S₁ and S₂ are located at ~89, 327, 952 cm⁻¹, and 89, 370, 473, 952 cm⁻¹, respectively.^{57,65,66} In addition, two Raman peaks of h-MoO₃ phase are appeared at ~398 and 492 cm⁻¹ for the sample S₁.⁶⁷ The magnified area of the micro-Raman spectra in wavenumber range 50–600 cm⁻¹ corresponding to the samples: S₁ and S₂ are shown in the inset of Fig. 11. Vibrational analysis by micro-Raman spectra reveals that the broad peaks around 1000–600 cm⁻¹, 600–200 cm⁻¹ and below 200 cm⁻¹ are associated with Mo–O stretching, bending and lattice vibration mode, respectively.^{66,68} The sharp intense peak located ~89 cm⁻¹ is assigned to the rigid transitional vibration of MoO₄

tetrahedral chain.⁶⁵ The peaks at ~327 and 370 cm⁻¹ are assigned to O–Mo–O bending and scissoring vibrations.⁵⁷ The characteristics Raman peaks of h-MoO₃ phase at ~398 and 492 cm⁻¹ are assigned to O–Mo–O vibrational mode.⁶⁷ The peak at ~473 cm⁻¹ corresponds to the Mo–O stretching and bending vibration of MoO₆ octahedra.⁶⁹ The presence of Mo⁶⁺ = O stretching vibration of terminal oxygen atoms is confirmed by the peak at ~952 cm⁻¹. These terminal oxygen bonds are created by breaking doubly coordinated oxygen (Mo₂–O) bonds at the corner-shared oxygen, which are common to two octahedra in the α-MoO₃ phase.⁶⁶ The QDs with mixed phase: α-MoO₃ and h-MoO₃ transformed to α-MoO₃ phase with increase in HCl acid concentration from 25% to 30%. The micro-Raman peaks at ~1645, 3250 and 3425 cm⁻¹ are due to H₂O scissor vibration and stretching vibration of the hydrogen-bonded OH groups.⁷⁰

FTIR characteristics. The IR absorption spectra of the samples: S₁ and S₂ shown in Fig. 12, in the wavenumber range 400–4000 cm⁻¹ are recorded for further study of vibrational properties of MoO_x QDs. The samples show the absorption peak

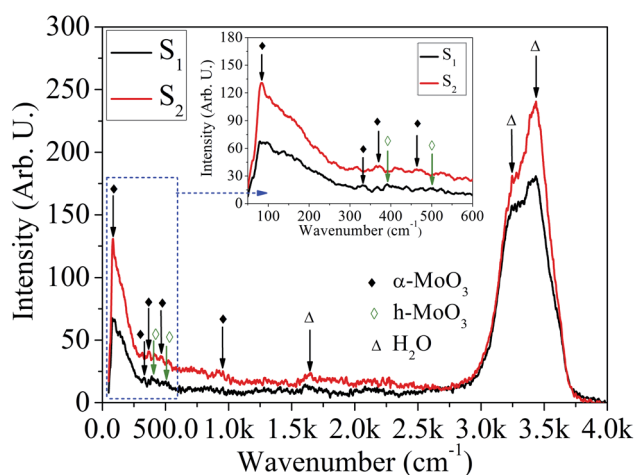


Fig. 11 Micro-Raman spectra of MoO_x QDs for the sample S₁ and S₂, respectively.

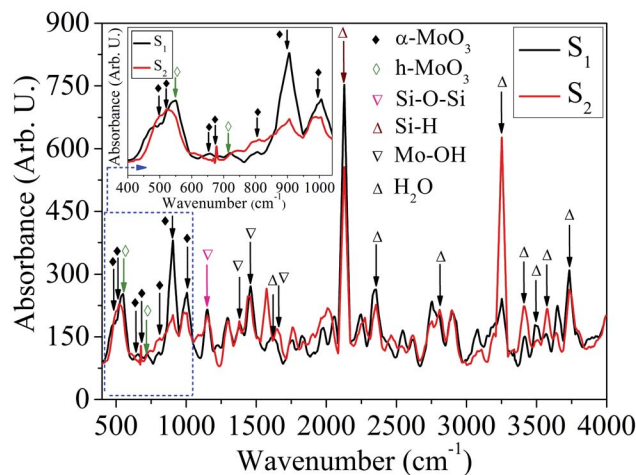


Fig. 12 FTIR absorption spectra of MoO_x QDs for the sample S₁ and S₂, respectively.



in the spectral range 400–1000 cm^{-1} corresponding to stretching and bending vibrations of metal–oxygen characteristics bonds.⁷¹ The characteristic absorption peaks for $\alpha\text{-MoO}_3$ phase of the sample S_1 and S_2 are located at ~ 502 , 648, 811, 901, 1002 cm^{-1} and ~ 502 , 512, 674, 811, 901, 1002 cm^{-1} , respectively.^{11,24,40,55,72–74} In addition, two absorption peaks of h-MoO₃ phase are appeared at ~ 548 and 711 cm^{-1} for the sample S_1 .⁶⁷ The magnified area of the FTIR absorption spectra in wavenumber range, 400–1040 cm^{-1} corresponding to the samples: S_1 and S_2 are shown in the inset of Fig. 12. The absorption peak located at ~ 502 and 512 cm^{-1} are attributed to the bending mode of vibration of the Mo–O–Mo entity, while the O^{2-} ion is shared by three Mo^{6+} ions.^{55,72} The absorption peaks of h-MoO₃ phase are ~ 548 and 711 cm^{-1} , associated with the vibration of Mo–O and Mo–O–Mo bond, respectively.⁶⁷ The peak at ~ 648 cm^{-1} could be assigned to the stretching vibration of Mo–O–Mo.¹¹ The absorption peak at ~ 674 cm^{-1} is due to the triply coordinated oxygen ($\text{Mo}_3\text{-O}$) stretching vibrational mode.⁴⁰ A broad peak observed at ~ 811 nm is due to the bridging vibration of Mo=O and indicated that the existence of Mo^{6+} oxidation state is related to the $\alpha\text{-MoO}_3$ phase.⁷³ Two absorption peak located at ~ 901 and 1002 cm^{-1} , associated with the stretching vibration of Mo=O, which is the specification of terminal double bonds. This confirms the basic characterises of a layered structure of $\alpha\text{-MoO}_3$.^{24,74} The phase transformation of MoO_x ($x = 2, 3$) QDs with the concentration of HCl acid is also confirmed by FTIR absorption spectra. The peak at ~ 1157 cm^{-1} is due to Si–O–Si asymmetric vibration in longitudinal optical mode.⁷⁵ In addition, three absorption peaks are observed at ~ 1384 , 1460 and 1643 cm^{-1} , where first peak is attributed to vibrational mode of the Mo–OH bond and other two correspond to bending mode of Mo–OH, which assigned the presence of adsorbed free water molecules.^{59,76} Moreover, the absorption peaks at ~ 1619 , and 3574 cm^{-1} are assigned to the stretching and bending vibration of O–H bonds of adsorbed water molecules, respectively.⁷⁷ The peak at ~ 2110 cm^{-1} corresponds to the stretching vibration of Si–H in dehydrate groups at the silicon surface.^{78,79} The absorption peaks located at ~ 2359 , 3245 and 3492 cm^{-1} assigned to the H–O–H bending mode, symmetric and anti-symmetric O–H mode of water.^{67,80} The absorption peaks at ~ 2809 and 3400 cm^{-1} are correspond to the typical O–H vibrations due to the adsorbed water molecule on the surface of MoO_x QDs.⁴⁰ The peak at ~ 3737 cm^{-1} associated with the stretching vibration of O–H in silanol groups (Si–OH) isolated by the hydrogen bond.⁷⁹ These FTIR results are consistent with the results obtained by HRTEM images and micro-Raman analysis.

Conclusions

In summary, this work developed a simple and modified top-down approach for synthesis of MoO_x QDs. This process is a combination of thermal evaporation and sonication induced chemical etching to get the desired particles in the solution. The XRD analysis reveals the thermally deposited MoO_x film is amorphous in nature. FETEM image analysis confirms the formation of spherical shaped MoO_x QDs with an average size

~ 4.3 and 4.6 nm. HRTEM, SAED, micro-Raman, and FTIR analysis reveals that as-prepared QDs exhibit polycrystalline in nature with co-existence of $\alpha\text{-MoO}_3$ and h-MoO₃ phase. MoO_x QDs exhibited excitation dependent fluorescence emission. The amorphous MoO_x thin film is able to transformed into the polycrystalline nature by this modified top-down approach. The bulk MoO_3 powder is successfully transformed to well-dispersed QDs *via* this modified top-down sonication induced chemical etching approach. Further, optimization of various parameters in this approach will lead to the development of tuneable crystal phase and size of the QDs. This work also confirms the successful synthesis of well-dispersed and stable QDs without any surfactants.

Conflicts of interest

The authors declare no competing financial interest.

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