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### Introduction

 $CO<sub>2</sub>$  is a rich C1 resource and its conversion into chemicals has attracted more and more attention.<sup>1-3</sup> Particularly, reducing  $CO<sub>2</sub>$  into  $CO$ , which is a key starting material for the synthesis of important chemicals and commodities, is of great importance.<sup>4</sup> Among the transformation approaches, the electrochemical  $CO<sub>2</sub>$  reduction reaction  $(CO<sub>2</sub> RR)$  into CO using an electrocatalyst is a promising way.<sup>5-9</sup> Compared with precious metal electrocatalysts (Au, Ag and Pd), $10-12$  it is more desirable to develop alternative earth-abundant metal catalysts for the  $CO<sub>2</sub>$ RR to CO. However, non-noble metal electrocatalysts often suffer from low activities and selectivities. It has been proposed that the formation of metal alloy or bimetallic catalysts (especially with noble metals) can improve the efficiency of nonprecious metals for the  $CO<sub>2</sub>$  RR, through tuning their binding energy to reaction intermediates and regulating their inherent activity.<sup>13</sup>–<sup>22</sup> However, the performances of these electrocatalysts for the  $CO<sub>2</sub>$  RR still need improvement in terms of two issues. First, other carbon products (e.g. formic acid, methane, etc.) are inevitably produced in many catalytic systems. Second, the current density is usually low, especially at relatively low

## $Cu<sub>x</sub>Ni<sub>v</sub>$  alloy nanoparticles embedded in a nitrogen–carbon network for efficient conversion of carbon dioxide†

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The electrocatalytic conversion of  $CO<sub>2</sub>$  to CO using non-noble metal catalysts under mild conditions is of great importance. Achieving the combination of high activity, selectivity and current density by developing electrocatalysts with desirable compositions and structures is challenging. Here we prepared for the first time  $Cu<sub>x</sub>Ni<sub>y</sub>$  alloy nanoparticles embedded in a nitrogen–carbon network. Such an electrocatalyst not only well overcomes the disadvantages of single Cu and Ni catalysts but has a high  $CO<sub>2</sub>$  adsorption capacity. Outstandingly, the catalyst can effectively convert  $CO<sub>2</sub>$  into CO with a maximum faradaic efficiency of 94.5% and current density of 18.8 mA cm<sup>-2</sup> at a low applied potential of  $-0.60$  V (versus reversible hydrogen electrode, RHE). Moreover, the catalyst is very stable during long-term electrolysis owing to the stabilization of the nitrogen–carbon network. **EDGE ARTICLE**<br> **(a)** Check for updates<br>
Cu<sub>s</sub> Ni<sub>i</sub>, alloy nanoparticles embedded in<br>
Cu<sub>s</sub> Ni<sub>i</sub> and Custom in a **and introgen — carbon network for efficient conversional comparations of carbon dioxide in a mitrogen — c** 

overpotentials. It is challenging to develop non-precious metal electrocatalysts with combined high activity, selectivity and current density for the CO<sub>2</sub> RR into CO.

Here we demonstrate for the first time the formation of Cu/ Ni alloy nanoparticles embedded in a three-dimensional (3D) nitrogen–carbon network. Outstandingly, the as-synthesized catalyst can effectively convert  $CO<sub>2</sub>$  into CO with a maximum faradaic efficiency of 94.5% and current density of 18.8 mA  $cm^{-2}$  at a low applied potential of  $-0.60$  V. Such an electrocatalyst well overcomes the disadvantages of single Cu and Ni catalysts. For example, Cu catalysts usually produce multiple products due to their moderate binding energy to most reaction intermediates.<sup>13,23</sup> Ni tends to favor the competitive hydrogen evolution reaction (HER) in aqueous electrolyte.<sup>24</sup> Moreover, the Cu/Ni catalyst is very stable during long-term electrolysis owing to the stabilization of the nitrogen–carbon network.

### Results and discussion

To get the target product, a Cu/Ni bimetallic complex was first prepared by a hydrothermal method using copper powder (Fig.  $S1\dagger$ ) and nickel( $\pi$ ) nitrate hexahydrate as precursors. From scanning electron microscopy (SEM) and transmission electron microscopy (TEM) images, the Cu/Ni bimetallic complex has a 3D network structure (Fig. 1a, d and  $S2\dagger$ ). After calcination of the Cu/Ni bimetallic complex at 900  $^{\circ}$ C for 1 h, the target product was obtained. It well maintains the network structure of the bimetallic complex, but it is embedded with nanoparticles in the range of 50–300 nm (Fig. 1b and c). The highresolution transmission electron microscopy (HRTEM) image (Fig. 1e) shows a lattice fringe corresponding to the crystal plane



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<sup>†</sup> Electronic supplementary information (ESI) available: SEM and TEM images, XRD patterns, EXAFS spectra, XPS survey spectra, faradaic efficiency, <sup>1</sup>H NMR spectra and partial current density details. See DOI: 10.1039/c9sc00174c



Fig. 1 SEM image (a) and TEM image of the Cu/Ni bimetallic complex (d), and SEM image (b), TEM image (c), HRTEM image (e), EDS line scans (f) and EDX mapping of  $Cu_{1.0}Ni_{1.0}/N-C$  (g). Scale bars, 300 nm in (a) (i) and EDA mapping of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N-C (g). Scale bars, 500 nm in (a) Fig. 2 High-resolution XPS spectrum of Cu 2p (a), Ni 2p (b), N 1s (c) and (d), and 2 nm in (e).

(111)  $(d = 0.21$  nm) of the face centered cubic Cu/Ni bimetallic alloy, $25,26$  which is consistent with the X-ray diffraction (XRD) pattern (Fig. S3†). Energy dispersive X-ray spectroscopy (EDS) line scans on nanoparticles reveal that Cu and Ni elements are distributed nearly at the same locations (Fig. 1f), further proving the formation of the Cu/Ni alloy. Energy dispersive X-ray elemental (EDX) mapping images show that C, N and O elements are evenly distributed in the whole sample, while Cu and Ni elements aggregate into granular forms, corresponding to the Cu–Ni alloy (Fig. 1g). The Cu and Ni contents were determined to be 37.8 and 33.5 wt% by inductively coupled plasma-mass spectrometry (ICP-MS), respectively. The contents of C and N are 23.2 and 4.2 wt%, respectively, as determined by elemental analysis. The product was denoted as  $Cu_{1,0}Ni_{1,0}/N-C$ .

The chemical compositions and elemental states of the  $Cu_{1.0}Ni_{1.0}/N-C$  catalyst were investigated by X-ray photoelectron spectroscopy (XPS). In the high-resolution Cu 2p spectrum (Fig. 2a), the binding energies at 934.8 eV and 954.9 eV correspond to  $Cu^{2+}$ , while those at 932.8 eV and 952.6 eV can be assigned to Cu<sup>0</sup> and possibly existing Cu<sup>+</sup>.<sup>27,28</sup> The Auger Cu LMM spectrum shows a peak at a binding energy of 918.5 eV (Fig.  $S4\dagger$ ), which excludes the presence of Cu<sup>+</sup> species.<sup>28</sup> In the high-resolution Ni 2p spectrum (Fig. 2b), the binding energies at 852.7 eV, 870.1 eV and 855.9 eV, 873.5 eV correspond to  $Ni<sup>0</sup>$ and  $Ni<sup>2+</sup>$ , respectively.<sup>29,30</sup> The four nitrogen species peaks in the high-resolution N 1s spectrum (Fig. 2c) correspond to pyridinic N (B.E. 398.9 eV), Cu/Ni–N (B.E. 399.5 eV), pyrrolic N (B.E. 400.9 eV) and graphitic N (B.E. 402.1 eV), respectively.<sup>9,31</sup> This indicates that the N species on the catalyst surface is mainly in the pyridinic and pyrrolic form, and there is a small amount of the Cu/Ni–N form. The presence of graphitic N species indicates that the 3D network is mainly composed of nitrogen-doped graphene.<sup>31</sup> The high-resolution C 1s spectrum (Fig. 2d) shows five peaks at 284.8, 286.1, 287.7, 289.2 and 291.3 eV, which correspond to C–C/C=C, C–N, C–OH, C=O and



and C 1s (d) of  $Cu_{1.0}Ni_{1.0}/N-C$ 

 $Ph-\pi^*$  species, respectively.<sup>32,33</sup> Furthermore, synchrotron extended X-ray absorption fine structure (XAFS) analysis was performed to further investigate the alloy structure of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C. The Cu–Cu bond distance of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C shifts to 2.209 Å, which is lower than that of Cu foil  $(2.240 \text{ Å})$ . However, the Ni–Ni bond distance of  $Cu_{1.0}Ni_{1.0}/N-C$  is 2.175 Å, which is a higher bond distance than that of Ni foil  $(2.082 \text{ Å})$ . This phenomenon also proves the existence of the Cu/Ni alloy in  $Cu_{1.0}Ni_{1.0}/N-C$  (Fig. S5†).<sup>34</sup>

A series of  $\text{Cu}_x\text{Ni}_v/\text{N}-\text{C}$  catalysts with different Cu to Ni molar ratios were synthesized and characterized (Fig. S6–S9 and Table S1<sup>†</sup>), where  $x/y$  represents the molar ratio of Cu to Ni. All the catalysts present similar morphologies to  $Cu_{1.0}Ni_{1.0}/N-C$ , *i.e.*, the Cu/Ni alloy nanoparticles are embedded in the nitrogen– carbon network. The X-ray diffraction peaks of the  $\text{Cu}_x\text{Ni}_v/\text{N}-\text{C}$ catalysts shift toward a lower diffraction degree with increasing Cu content, implying the mergence of more Cu atoms into the Ni lattice (Fig. S3†).1,35,36 For comparison, Cu/N–C and Ni/N–C catalysts were synthesized using a method similar to that for  $Cu<sub>x</sub>Ni<sub>v</sub>/N-C$  synthesis in the absence of nickel( $\pi$ ) nitrate hexahydrate and copper powder, respectively (Fig. S10 and 11†).

The electrocatalytic activities of  $Cu_xNi_y/N-C$  for the  $CO_2$  RR were studied in a two-compartment cell. To assess the  $CO<sub>2</sub>$ electroreduction performance of  $Cu<sub>x</sub>Ni<sub>y</sub>/N-C$ , linear sweep voltammetry (LSV) measurements were performed in  $N_2$  and  $CO_2$ saturated 0.5 M KHCO<sub>3</sub>, respectively (Fig. 3a). For  $Cu_{1.0}Ni_{1.0}/N$ -C and  $\text{Cu}_{4.8}\text{Ni}_{1.0}\text{/N--C}$ , the increase of current density relates to the HER in the  $N_2$  saturated solution, and the combination of the  $CO<sub>2</sub>$  RR and HER in the  $CO<sub>2</sub>$  saturated solution.<sup>37</sup> The sharp increase of current density in the  $CO<sub>2</sub>$  saturated solution compared to that in the  $N_2$  saturated solution indicates that Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C and Cu<sub>4.8</sub>Ni<sub>1.0</sub>/N–C are more favorable for the  $CO<sub>2</sub>$  RR than the HER. However, the current density of Cu/N–C in CO<sub>2</sub> saturated 0.5 M KHCO<sub>3</sub> is close to that in the N<sub>2</sub> saturated solution. This implies that the performance of Cu/N–C for



Fig. 3 LSV curves acquired in  $N_2$ -saturated (dotted line) or  $CO_2$ saturated (solid line) 0.5 M KHCO<sub>3</sub> solution at a scan rate of 50 mV s<sup>-1</sup> (a). CO faradaic efficiency at various potentials (b). Total current density and partial current density of different catalysts at an applied potential of  $-0.60$  V vs. RHE (c). Stability of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C at a potential of -0.60 V vs. RHE during 38 h (d). Catalyst loading: 0.5 mg  $cm^{-2}$ .

 $CO<sub>2</sub>$  reduction is lower than that of  $Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N-C$  and  $Cu_{4.8}Ni_{1.0}/N-C$  catalysts. Moreover, the current density of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C (40.3 mA  $cm^{-2}$ ) and Cu<sub>4.8</sub>Ni<sub>1.0</sub>/N–C (30.1 mA  $\text{cm}^{-2}$ ) at  $-1.0$  V is much higher than that of Cu/N–C (15.5 mA  $\text{cm}^{-2}$ ).

The gaseous products and liquid products formed in  $CO<sub>2</sub>$ reduction were analyzed by gas chromatography (GC) and  $^{1}\mathrm{H}$ nuclear magnetic resonance (<sup>1</sup>H NMR) spectroscopy, respectively. The faradaic efficiencies of CO formation during the  $CO<sub>2</sub>$ RR at different applied potentials are presented in Fig. 3b. CO and  $H_2$  were produced when catalyzed by  $Cu_xNi_y/N-C$  and there is no liquid product in the whole electrolysis process. In contrast,  $H_2$  is the main catalytic product with a small amount of CO and HCOOH when catalyzed by Cu/N–C, and  $H_2$  is the only catalytic product when catalyzed by Ni/N–C (Fig. S12–S14†). Outstandingly, the maximum faradaic efficiency FE(CO) of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C is 94.5% at a low potential of  $-0.60$  V, with a high current density of 18.8 mA  $\rm cm^{-2}.$  In sharp contrast, the FE(CO) of Cu/N–C is only 20.4% at the same potential, with a small current density of 6.5 mA  $\rm cm^{-2}.$  Such a high FE(CO) of  $Cu_{1.0}Ni_{1.0}/N-C$  at low potential provides a maximum energy conversion efficiency of 69.2%, which is about 4.6 times that of Cu/N–C (Fig. S15†).38,39 Compared with the reported Cu-based alloys and bimetallic catalysts, the catalytic performance of  $Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N-C$  is superior owing to the combined high FE(CO) and current density at relatively low potential (Table S2†).18,20,22,27,40–<sup>42</sup>

Fig. 3c shows the total current density and partial current density of different catalysts at  $-0.60$  V for CO formation. Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C has a  $j_{\text{CO}}$  of 17.7 mA cm<sup>-2</sup> at -0.60 V, which is about 13.4 times that of the Cu/N–C catalyst  $(1.32 \text{ mA cm}^{-2})$ . Moreover, the partial current density for CO formation increases first and then decreases with applied potential

(Fig. S16†), indicating that high applied potentials can accelerate the reaction rate for  $CO_2$  reduction.<sup>38,43</sup> Attractively, the catalytic performance of  $Cu_{1.0}Ni_{1.0}/N-C$  is very stable, because the current density and  $FE(CO)$  remain undiminished after 38 h of constant potential electrolysis (Fig. 3d). The reused  $Cu_{1.0}Ni_{1.0}/N-C$  catalyst was characterized by different techniques (Fig. S17 and 18†). It is observed that the  $Cu_{1,0}Ni_{1,0}/N-C$ catalyst is very stable against long-term electrolysis.

The electrochemical surface area (ECSA), which can be evaluated by comparing the double-layer capacitance  $(C_{d_l})$  of catalysts, is an important factor affecting the catalytic performance.<sup>44</sup> From Fig. 4a, the  $C_{d1}$  of Cu<sub>1.0</sub>Ni<sub>1.0</sub>/N–C is about 2 times that of Cu/N–C. The large ECSA will provide abundant active sites, and thus can enhance the activity for the  $CO<sub>2</sub>$  RR.<sup>45</sup> The Tafel slopes of  $Cu_{1.0}Ni_{1.0}/N-C$  and  $Cu_{4.8}Ni_{1.0}/N-C$  are 113.4 and 104.2 mV dec $^{-1}$ , respectively (Fig. 4b), which are close to the theoretical value of 116 mV dec $^{-1}$ . It indicates that the single electron transfer to  $CO<sub>2</sub>$  to form the surface adsorbed  $CO_2$ <sup>--</sup> intermediate is the rate-determining step for  $Cu_x$ - $Ni<sub>v</sub>/N-C.<sup>9</sup>$  The values of Cu<sub>x</sub>Ni<sub>v</sub>/N–C are obviously lower than those of Cu/N–C, implying that the formation of CO by  $\text{Cu}_x\text{Ni}_v$ / N–C is more favorable kinetically. Electrochemical impedance spectroscopy (EIS) was used to further explain the reaction kinetics of the  $CO<sub>2</sub>$  RR. EIS was carried out in 0.5 M KHCO<sub>3</sub> solution at an open circuit potential (Fig. 4c), and the equivalent circuit (Fig.  $S19\dagger$ ) was then used to fit the impedance data. The Nyquist plots confirm that  $Cu_{1.0}Ni_{1.0}/N-C$  and  $Cu_{4.8}Ni_{1.0}/N-C$  possess smaller charge transfer resistance than Cu/N–C. This indicates a faster electron transfer to  $CO<sub>2</sub>$  for forming the reduced  $CO_2$ <sup> $\sim$ </sup> intermediate in the  $CO_2$  RR catalyzed by  $Cu_xNi_y/N-C$ , thus leading to greatly enhanced catalytic efficiency for  $CO<sub>2</sub>$  reduction.<sup>46</sup> Experimental on 15 March 2019. Downloaded on 15 March 2019. The common article is licensed under a creative Common and the Creative Common and the Creative Commons Article is licensed under a creative Common and the Creat

The above results reveal that the  $Cu_xNi_y/N-C$  catalyst has superior electrocatalytic performance for  $CO<sub>2</sub>$  reduction to  $CO<sub>2</sub>$ in particular compared with mono-metal catalysts. In



Fig. 4 Charging current density differences  $\Delta j$  plotted against scan rates (a). Tafel plots for CO production (b). The experimental (inset) and simulated EIS spectra of the samples (c).  $CO<sub>2</sub>$  adsorption isotherms of different catalysts at 298 K (d).

comparison with Cu/N–C, the enhanced electrocatalytic activity of  $\text{Cu}_x\text{Ni}_v/\text{N}-\text{C}$  can be attributed to the more stabilized COOH\* adsorption and decreased overpotential caused by the CuNi alloy.<sup>13,47</sup> Compared with Ni/N–C which produces  $H_2$  only, the release of CO from the CuNi alloy surface is more feasible than from the Ni surface, resulting from the reduced binding energy of CO\*. 48,49 Moreover, it is worth noting that the 3D porous nitrogen–carbon network would improve the  $CO<sub>2</sub>$  adsorption capacity, contributing to the enhancement of reactivity.<sup>50,51</sup> To testify this, the  $CO<sub>2</sub>$  adsorption isotherms of different catalysts were determined (Fig. 4d). Obviously, the  $Cu<sub>x</sub>Ni<sub>y</sub>/N-C$  catalysts present improved  $CO<sub>2</sub>$  adsorption capacity than Cu/N–C. For example,  $Cu_{1.0}Ni_{1.0}/N-C$  reaches a  $CO<sub>2</sub>$  adsorption capacity of 23.4 mg  $g^{-1}$  at 298 K and 1 atm, which is 3.3 times higher than that of Cu/N–C. Therefore, the large  $CO<sub>2</sub>$  adsorption capability facilitates the adsorption of a large amount of  $CO_2$  on the  $Cu_xNi_y/$ N–C catalyst surface, thus promoting its catalytic performance. Chemical Science<br>
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### **Conclusions**

In conclusion, we prepared the  $Cu<sub>x</sub>Ni<sub>y</sub>/N-C$  catalyst with Cu/Ni alloy nanoparticles embedded in a highly porous nitrogen– carbon network, which has combined advantages for the electrocatalytic conversion of  $CO<sub>2</sub>$  to CO. For example, it well overcomes the disadvantages of single Cu and Ni electrocatalysts; the highly porous nitrogen–carbon network not only improves the  $CO<sub>2</sub>$  adsorption ability but stabilizes the Cu/Ni nanoparticles during long-term electrolysis. The  $Cu_xNi_y/N-C$  catalyst exhibits a maximum faradaic efficiency of 94.5% and current density of 18.8 mA cm<sup>-2</sup> at a relatively low potential of  $-0.60$  V vs. RHE. This study opens up new opportunities for improving the performance of metal electrocatalysts for  $CO<sub>2</sub>$  reduction under mild conditions, particularly for metals with low intrinsic activity.

### Experimental

#### Materials

 $CO<sub>2</sub>$  (>99.999% purity) was provided by Beijing Analysis Instrument Factory. Copper powder (>98% purity), Nafion N117 membrane and Nafion D-521 dispersion were purchased from Beijing InnoChem Science & Technology Co., Ltd. 2-Methylimidazole, cupric oxide and nickel oxide were purchased from J&K Scientific Co., Ltd. Absolute methanol was provided by Beijing Chemical Works. Nickel nitrate hexahydrate was provided by Sinopharm Chemical Reagent Beijing Co., Ltd.

#### Preparation of  $Cu<sub>x</sub>Ni<sub>y</sub>/N-C$

In a typical synthesis process, appropriate amounts of copper powder, nickel nitrate hexahydrate and 2-methylimidazole were dispersed in 10 mL of anhydrous methanol and stirred at room temperature for 10 min. After this, the solution was transferred into a Teflon lined stainless-steel autoclave, and the autoclave was sealed and heated at 120  $\degree$ C for 24 h under stirring. After the autoclave cooled to room temperature, the as-prepared sample was washed with methanol and dried in a vacuum oven at 80  $^{\circ}$ C

for 24 h. The solid was annealed in a tube furnace at 900  $^{\circ}$ C in a N<sub>2</sub> atmosphere for 1 h with a heating rate of 2  $^{\circ}$ C min<sup>-1</sup>.

#### Characterization

The morphologies of the catalysts were characterized by SEM (HITACHI S-4800). TEM images were taken on JEOL JEM-EM-1011 and JEOL JEM-2100F field-emission high-resolution transmission electron microscopes operated at 200 kV. XRD patterns were recorded on a Rigaku D/max 2400 diffractometer with Cu K $\alpha$  radiation ( $\lambda = 0.15418$  nm) with a scanning rate of  $4^\circ$  min<sup>-1</sup>. XPS was carried out with a multipurpose X-ray photoemission spectrometer (Thermo Scientific ESCALAB 250Xi). The elemental analysis of C, N and O was performed on a FLASH EA1112 elemental analysis instrument. The contents of Cu and Ni elements were determined by ICP-MS. The XAFS experiment was carried out at Beamline 1W1B at BSRF. Data of XAFS were processed using the Athena and Artemis programs of the IFEFFIT package based on FEFF 6. Data were processed with  $k^3$ -weighting and an Rbkg value of 1.0. Normalized XANES data were obtained directly from the Athena program of the IFEFFIT package.

#### Fabrication of electrodes

The working electrodes were prepared by loading the sample suspension onto carbon paper (CP). Briefly, the as-prepared catalysts and 10  $\mu$ L Nafion D-521 dispersion were dispersed in absolute ethanol and ultrasonicated for 30 min to form a uniform suspension, and the suspension was loaded on 1 cm  $\times$  1 cm CP. The electrode was dried in a vacuum oven at 80 °C for 6 h before electrochemical experiments.

A CHI660E electrochemical analyzer (Shanghai Chenhua Instrument co. LTD, China) was used in all the electrochemical experiments. A three-electrode test was carried out in a sealed H-cell which was separated by a Nafion N117 membrane. The reference and counter electrodes are Ag/AgCl and a Pt net  $(1 \times 1)$ cm), respectively. All experiments were carried out at atmospheric pressure and room temperature. Linear sweep voltammetry (LSV), amperometric  $i-t$  curve measurements and electrochemical impedance spectroscopy (EIS) were performed in 0.5 M KHCO<sub>3</sub> solution. Before experiments, the electrolyte solution in the working compartment was bubbled for 1 h with  $N_2$  (99.99%) and  $CO_2$  (99.999%) to form a saturated solution. The pH values of the electrolyte solutions with  $N_2$  and  $CO_2$  are 8.8 and 7.3, respectively. All potentials reported in this paper are with respect to the reversible hydrogen electrode (RHE), which were converted using eqn (1).

Potential in the RHE ¼ applied potential vs. Ag/AgCl + 0.21 V + 0.0592 pH (1)

The gaseous product in the gas bag was collected and analyzed using a gas chromatograph (GC, HP 4890D), which was equipped with TCD and FID detectors, using helium as the internal standard. The liquid mixture was analyzed by the  ${}^{1}$ H-NMR method, and the spectra were recorded on a Bruker Avance III 400 HD spectrometer in deuterium oxide-d<sub>2</sub> with TMS as an internal standard.

#### Electrochemical impedance spectroscopy (EIS)

The EIS measurement was carried out in  $0.5$  M KHCO<sub>3</sub> solution at an open circuit potential (OCP) with an amplitude of 5 mV in the frequency range of  $10^{-2}$  to  $10^5$  Hz.

#### Double-layer capacitance  $(C_{\rm dl})$

The  $C_{\text{dl}}$  was determined by measuring the capacitive current associated with double-layer charging from the scan-rate dependence of cyclic voltammetric stripping. The scan rates were 5 mV  $\rm s^{-1} ,$  10 mV  $\rm s^{-1} ,$  15 mV  $\rm s^{-1} ,$  20 mV  $\rm s^{-1} ,$  25 mV  $\rm s^{-1} ,$  $30$  mV  $\text{s}^{-1}$ ,  $40$  mV  $\text{s}^{-1}$  and  $50$  mV  $\text{s}^{-1}$ .

#### Energy efficiency

The energy efficiency of reduction of  $CO<sub>2</sub>$  to  $CO$  was obtained using eqn (2):

$$
\mathcal{O}_{\rm{CO}} = \frac{\rm{FE}(\rm{CO}) \times \Delta E_{\rm{CO}}^0}{\Delta E_{\rm{CO}}} \tag{2}
$$

In the equation,  $\mathcal{O}_{\text{CO}}$  stands for the cathodic energy conversion efficiency of the CO formation;  $\Delta E_{\rm CO}^0$  represents the difference between the standard half reaction potentials for water oxidation (1.23 V vs. RHE) and reduction of  $CO<sub>2</sub>$  to  $CO<sub>2</sub>$  $(-0.11 \text{ V} \text{ vs. RHE})$ ;  $\Delta E_{\text{CO}}$  stands for the difference between the standard water oxidation potential and the applied potential for CO<sub>2</sub> reduction. Edge Article<br>
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### Conflicts of interest

The authors of this manuscript have no conflicts of interest.

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