



Cite this: *RSC Adv.*, 2019, 9, 31572

Theoretical investigation of the ORR on boron–silicon nanotubes (B–SiNTs) as acceptable catalysts in fuel cells

Razieh Razavi^{*a} and Meysam Najafi ^{*b}

Here, the potential of boron doped silicon nanotubes (7, 0) as ORR catalysts is examined. Acceptable paths for the ORR on studied catalysts are examined through DFT. The optimum mechanism of the ORR on the surface of B₂–SiNT (7, 0) is shown. The ORR on the surface of B₂–SiNTs (7, 0) can continue through LH and ER mechanisms. The calculated beginning voltage for the ORR on B₂–SiNTs (7, 0) is 0.37 V and it is smaller than the beginning voltage (0.45 V) for platinum-based catalysts. In the acidic solution the beginning voltage for the oxygen reduction process can be evaluated to be 0.97 V, which corresponds to 0.37 V as a minimum overvoltage for the ORR. The B₂–SiNTs (7, 0) are suggested as an ORR catalyst in acidic environments.

Received 3rd July 2019
 Accepted 29th September 2019

DOI: 10.1039/c9ra05031k

rsc.li/rsc-advances

1. Introduction

Fuel cells as energy machines are important due to their low contamination and great efficiency. The ORR rate in electrodes of cells is slow, therefore the ORR can be evaluated as a significant reason to increase the full cell efficiency.^{1–4} Platinum-compounds have been used as catalysis in the ORR but platinum-compounds have low ability to endure CO.^{5–9}

The potential of various compounds was investigated to find and propose effective catalysts for the ORR. Nanostructures and doped nanostructures with high ability for CO endurance can be used as suitable replacements for platinum-compounds.^{10–15} The B-nanostructures are acceptable catalysts for the ORR in alkaline conditions and mechanisms of action of B-doped nanostructures in acidic position are not clear.^{16–23}

The nanostructures due to their electrical conductivity and thermal conductivity can be used to product the transistors and non-volatile memory devices.^{24–28} The electrical conductivity of doped carbon/silicon nanotubes indicated that the adoption of carbon/silicon nanotubes (with various atoms such as B, N, O and some metals) increased their electrical conductivity, significantly. These findings improved the application of carbon/silicon nanotubes in nano-electronic devices and novel catalyst to ORR.^{29–38} Results demonstrated that the adoption of carbon/silicon nanotubes increased their electrical conductivity and enhanced the ORR efficiency.^{39–48}

Wang *et al.*⁴⁹ demonstrated that boron-doped graphene nano-ribbons are suitable catalyst to ORR catalyst. Xiao *et al.*⁵⁰ proved

that the layered silicon–carbon nano sheets represented the high activity in ORR without CO poisoning. Xia and Zhang *et al.*^{51,52} investigated the mechanisms of ORR of fuel cells in acidic environment on graphene cathodes. Stevenson *et al.*⁵³ proved in ORR the O₂ in a 2-electron path is reduced to form OOH on carbon nanotubes. Hu and Xiong *et al.*^{54,55} confirmed that nitrogen and boron-doped nanostructures as ORR catalysts have low price, great durability and excellent potential. Zhao and Wei *et al.*^{56,57} confirmed the doping of carbon nanotubes have vital roles on performance of ORR. Ferrighi *et al.*⁵⁸ demonstrated that boron atoms of nano-sheets increase the reactions of oxygen with graphene.

In current study, ORR on B-doped silicon nanotube (7, 0) as acceptable catalysts is examined to find possible mechanisms to ORR on B₂–SiNT (7, 0) and to suggest high activity nano-catalysts to ORR.

2. Computational details

In this study the silicon nanotube (length and diameter are 1 and 0.475 nm) is modeled and their open elements are saturated with hydrogen atoms to elude border effects. The geometries of nanotubes and studied molecules (such as OOH, OH, H₂O and CO) are optimized by M06-2X/6-311G+ (2d, 2p) in GAMESS package.^{59–72} The consistent field is investigated by 10^{–6} Hartree as convergence value. Vibrational frequencies of nanotubes and molecules by M06-2X/6-311G+ (2d, 2p) are calculated.

In the density functional, M06 functionals are extremely parameterized proximate exchange functionals theory and they are supported on generalized gradient approximation (meta-GGA). These functionals are used for traditional quantum chemistry, solid-state physics calculations and thermodynamic values of reactions.^{73–82} M06-2X as the most accurate functional

^aDepartment of Chemistry, Faculty of Science, University of Jiroft, Jiroft, Iran. E-mail: R.Razavi@ujiroft.ac.ir

^bMedical Biology Research Center, Health Technology Institute, Kermanshah University of Medical Sciences, Kermanshah, Iran. E-mail: iau.mnajafi@yahoo.com



of Minnesota functional is a Global hybrid functional with 54% HF exchange and it is the ascendancy constructor within the 06 functionals for thermochemistry, kinetics and various chemical interactions.^{73–82} The M06-2X functional as hybrid meta exchange–correlation functionals present 32 empirically improved factors within the exchange–correlation functional.^{83–87}

The energy and Gibbs free energy ($G = E_0 + ZPE + \Delta H + RT - TS$) values of nanotubes are calculated. The E_0 and ZPE are electronic energy and zero-point energy and T is 298.15 K.^{59–72} Adoption energy (E_{doped}) and Gibbs free energy adoption (G_{doped}) of B atoms in SiNT (7, 0) are calculated:

$$E_{\text{doped}} = E(\text{B-SiNT}) - E(\text{SiNT}) - E(\text{B}) \quad (1)$$

$$G_{\text{doped}} = G(\text{B-SiNT}) - G(\text{SiNT}) - G(\text{B}) \quad (2)$$

$$E_{\text{doped}} = E(\text{B}_2\text{-SiNT}) - E(\text{SiNT}) - 2E(\text{B}) \quad (3)$$

$$G_{\text{doped}} = G(\text{B}_2\text{-SiNT}) - G(\text{SiNT}) - 2G(\text{B}) \quad (4)$$

$E(\text{B-SiNT (7, 0)})$ and $E(\text{B}_2\text{-SiNT (7, 0)})$ are energies of B-SiNT (7, 0) and B₂-SiNT (7, 0).

Energy adsorption (ΔE_{ad}) and Gibbs free energy adsorption (ΔG_{ad}) of molecules (such as OOH, OH, H₂O and CO) on surfaces of studied nanotubes (SiNT, B-SiNT, B-B-SiNT and B₂-SiNT) are calculated:

$$\Delta E_{\text{ad}} = E(\text{molecule-nanotube}) - E(\text{nanotube}) - E(\text{molecule}) \quad (5)$$

$$\Delta G_{\text{ad}} = G(\text{molecule-nanotube}) - G(\text{nanotube}) - G(\text{molecule}) \quad (6)$$

$E(\text{molecule-nanotube})$ and $G(\text{molecule-nanotube})$ are E and G of molecules–nanotubes. $G(\text{nanotube})$ and $G(\text{molecule})$ are G of nanotubes and molecules. $E(\text{nanotube})$ and $G(\text{molecule})$ are energy and G of nanotubes and molecules. Natural bond orbital charges (q) and gap energy (E_{HLG}) complexes are calculated^{59–72} and transition state, reaction energy (ΔE_{r}) and activation barrier energy (ΔE_{a}) are examined by LST/QST method and M06-2X.⁶⁴

Activation energy ($\Delta E_{\text{a}} = E_{\text{TS}} - E_{\text{IS}}$) is difference of energy between transition (E_{TS}) and initial (E_{IS}) studied complexes. In this study, the reaction energy ($\Delta E_{\text{a}} = E_{\text{FS}} - E_{\text{IS}}$) is difference of energy between final (E_{FS}) and initial (E_{IS}) studied complexes.^{88–95} The ΔG of ORR on B₂-SiNT (7, 0) in according to standard hydrogen electrode was evaluated through $\Delta G = \Delta E + \Delta ZPE - T\Delta S + \Delta G_{\text{U}} + \Delta G_{\text{pH}}$ (obtained data were reported in Fig. 5).^{96–104} ΔG_{pH} ($\Delta G_{\text{pH}} = kT \ln 10 \times \text{pH}$) is modification of proton Gibbs free energy and $\Delta G_{\text{U}} = -neU$. The n , e and U are electrons, first charge and electrode potential. The U is requested potential and overvoltage is $\eta = U_0 - U$.^{49–58} Conductor like screening method is used to estimate water environment (dielectric constant is 78.54).^{59–72}

3. Results and discussion

3.1. Molecule adsorptions on nanotube

In this section, the B adoption of SiNT (7, 0) were investigated and then interactions of B-SiNT (7, 0) structures with O₂, OOH,

OH, H₂O and CO molecules were investigated. The one Si atom of the SiNT (7, 0) was replaced with one B atom and the B-SiNT (7, 0) was produced (Fig. 1). Also the two Si atoms of the SiNT (7, 0) in two difference positions were replaced with two B atoms and B-B-SiNT (7, 0) and B₂-SiNT (7, 0) structures were produced (Fig. 1). Structures of SiNT, B-SiNT, B-B-SiNT, B₂-SiNT, O₂, OH, H₂O, H₂O₂ and CO are presented in Fig. 1. The adoption energy (E_{doped}), adoption free Gibbs energy (G_{doped}) and bond lengths of B-Si of B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) were reported in Fig. 1. In the B-SiNT (7, 0) the B atom was connected with three neighboring silicon atoms and the E_{doped} and G_{doped} were -2.18 and -2.10 eV and average of bonds of B-Si in B-SiNT (7, 0) is 1.95 Å. In the B-B-SiNT (7, 0) the B atoms are connected with four neighboring silicon atoms and the E_{doped} and G_{doped} are -2.23 and -2.14 eV and average of bonds of B-Si in B-B-SiNT (7, 0) is 1.93 Å. In the B₂-SiNT (7, 0) the B atoms are connected with six neighboring silicon atoms and the E_{doped} and G_{doped} values are -2.28 and -2.17 eV and average of bonds of B-Si in B₂-SiNT (7, 0) are 1.92 Å.

The q and E_{HLG} of SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) are stated in Table 1. E_{HLG} of SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) are 1.84, 1.75, 1.68 and 1.64 eV. The q of B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) are 0.58, 0.69 and 0.73| e |. The $|E_{\text{doped}}|$, $|G_{\text{doped}}|$ and q values of B₂-SiNT (7, 0) are higher than B-SiNT (7, 0) and B-B-SiNT (7, 0). Results showed that the E_{HLG} value of B₂-SiNT (7, 0) is lower than corresponding values on surfaces of B-SiNT (7, 0) and B-B-SiNT (7, 0). Therefore, it can be concluded that the B₂-SiNT (7, 0) is the most stable than B-SiNT (7, 0) and B-B-SiNT (7, 0) from thermodynamic view point. The B atoms in B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) structures are advantageous to adsorption of O₂ molecule and these B atoms are active positions of B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) as catalyst for ORR. Therefore, B atoms are essential location to adsorption of O₂ molecule and B atoms can be considered as initiator of first ORR step.

Therefore, O₂ adsorption on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) were investigated. The possible positions of SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) to O₂ adsorption including top position B atom and bridge positions of B-Si, B-B and Si-Si bonds were examined in Fig. 2. The B-O, O-O and Si-O in SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) with O₂ are presented in Fig. 2 (2a–2f structures). The ΔE_{ad} and ΔG_{ad} of O₂ on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) are displayed in Table 2. $|\Delta G_{\text{ad}}|$ of O₂ on B-SiNT, B-B-SiNT and B₂-SiNT are greater than SiNT (7, 0). The $|\Delta E_{\text{ad}}|$ and $|\Delta G_{\text{ad}}|$ of O₂ on B₂-SiNT (7, 0) are greater than B-SiNT (7, 0), B-B-SiNT (7, 0). The bridge position of B-B in B₂-SiNT (7, 0) is stable than top position B in B₂-SiNT (7, 0) to O₂ adsorption.

Wang, Xiao and Xia *et al.*^{49–51} calculated the O₂ adsorption on surfaces of boron-doped graphene nanoribbon, silicon–carbon nano sheets and B and N doped-graphene by theoretical methods, respectively (results reported in Table 4). The ΔE_{ad} of O₂ on B-doped graphene, silicon–carbon nano-sheets and N-doped graphene are -0.62 , -0.53 and -0.60 eV. Therefore ΔE_{ad} value (-0.61 eV) of O₂ on B₂-SiNT (7, 0) in present study is



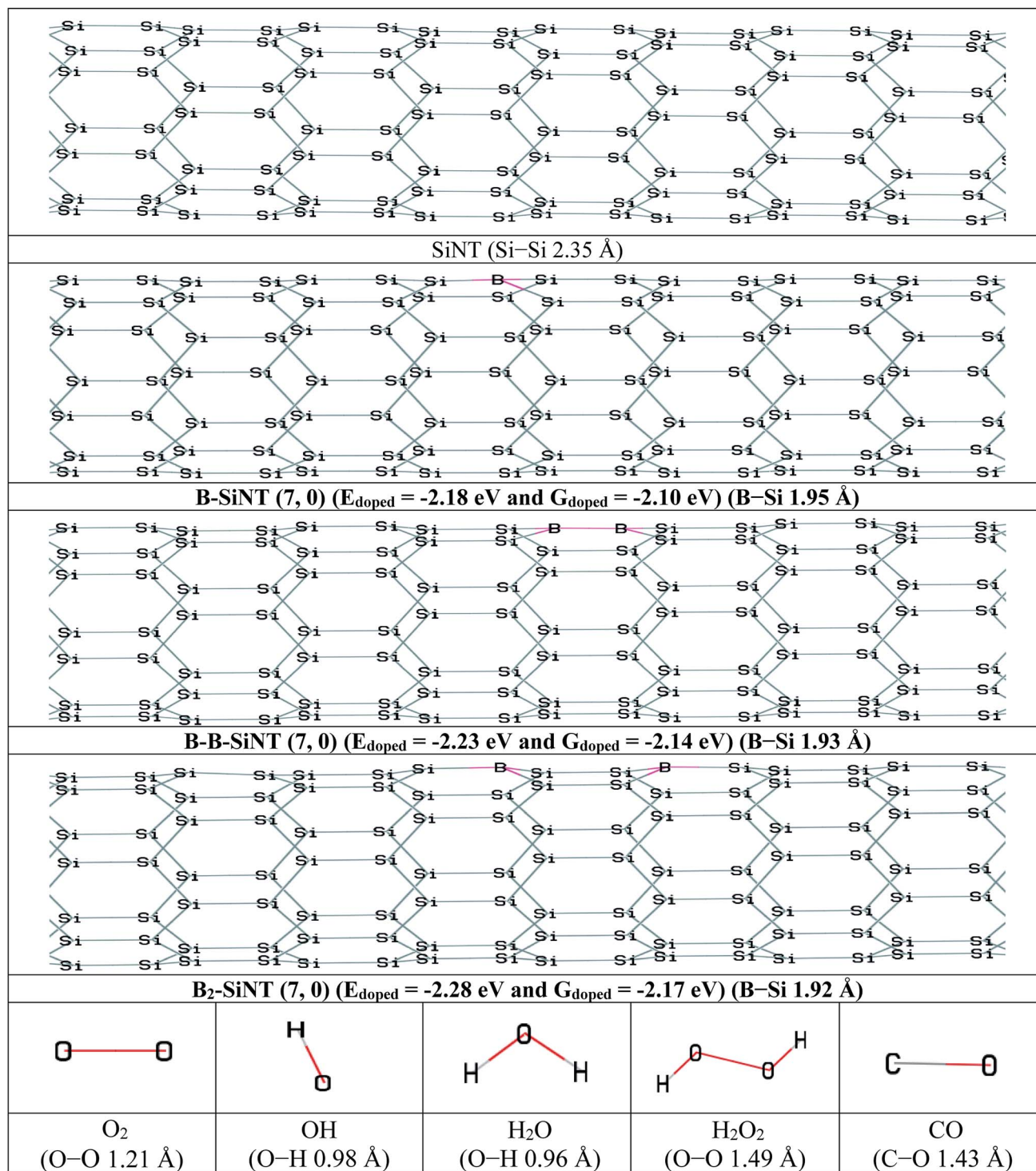


Fig. 1 The initial structures of SiNT, B-SiNT, B-B-SiNT and B₂-SiNT and O₂, OH, H₂O, H₂O₂ and CO molecules.

similar to corresponding values of O₂ on various nanostructures were calculated in previous theoretical works.⁴⁹⁻⁵¹

The charge transfer (q) and HOMO-LUMO band gap (E_{HLG}) of the complexes of SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) with O₂ molecule are displayed in Table 1. The E_{HLG} values of O₂ adsorption on B-SiNT (7, 0), B-B-SiNT (7,

0) and B₂-SiNT (7, 0) are lower than SiNT (7, 0). The E_{HLG} values of O₂ adsorption on B₂-SiNT (7, 0) are lower than B-SiNT (7, 0), B-B-SiNT (7, 0). The bridge position of B-B in B₂-SiNT (7, 0) has higher q and lower E_{HLG} than top position B in B₂-SiNT (7, 0) to O₂ adsorption. Complex of B₂-SiNT (7, 0) with O₂ molecule (**2d** structure) is the most stable than other complexes of B-SiNT (7,



Table 1 Charge transfer (q) (in $|e|$) and HOMO–LUMO band gap (E_{HLG}) (in eV) of studied complexes

| Complex | q | E_{HLG} | Complex | q | E_{HLG} | Complex | q | E_{HLG} | Complex | q | E_{HLG} |
|-------------|------|------------------|---------------|------|------------------|-----------------|------|------------------|---------------------------|------|------------------|
| <i>SiNT</i> | — | 1.84 | <i>B–SiNT</i> | 0.58 | 1.75 | <i>B–B–SiNT</i> | 0.69 | 1.68 | <i>B₂–SiNT</i> | 0.73 | 1.64 |
| <i>2a</i> | 0.36 | 1.69 | <i>2b</i> | 0.47 | 1.61 | <i>2c</i> | 0.59 | 1.52 | <i>2d</i> | 0.64 | 1.48 |
| <i>2e</i> | 0.29 | 1.75 | <i>2f</i> | 0.39 | 1.68 | <i>2g</i> | 0.51 | 1.59 | <i>2h</i> | 0.58 | 1.55 |
| <i>2i</i> | 0.32 | 1.72 | <i>2j</i> | 0.43 | 1.64 | <i>2k</i> | 0.54 | 1.56 | <i>2l</i> | 0.61 | 1.52 |
| <i>2m</i> | 0.49 | 1.35 | <i>2n</i> | 0.62 | 1.23 | <i>2o</i> | 0.73 | 1.17 | <i>2p</i> | 0.82 | 1.14 |
| <i>3a</i> | 0.41 | 1.54 | <i>3b</i> | 0.51 | 1.45 | <i>3c</i> | 0.65 | 1.38 | <i>3d</i> | 0.73 | 1.25 |
| <i>3e</i> | 0.59 | 1.14 | <i>3f</i> | 0.81 | 1.05 | <i>3g</i> | 0.84 | 0.99 | <i>3h</i> | 0.91 | 0.95 |
| <i>3m</i> | 0.11 | 1.80 | <i>3n</i> | 0.14 | 1.70 | <i>3o</i> | 0.19 | 1.62 | <i>3p</i> | 0.21 | 1.60 |
| <i>3q</i> | 0.08 | 1.82 | <i>3r</i> | 0.12 | 1.72 | <i>3s</i> | 0.15 | 1.64 | <i>3t</i> | 0.18 | 1.63 |

0) and B–B–SiNT (7, 0) with O₂ molecule from thermodynamic view point. It can be concluded that O₂ adsorbed on B₂–SiNT (7, 0) in figure 2*d* significantly and there are suitable interactions

between the O₂ molecule and B₂–SiNT (7, 0) and the adsorption of O₂ molecules on studied surfaces are chemical adsorption processes.

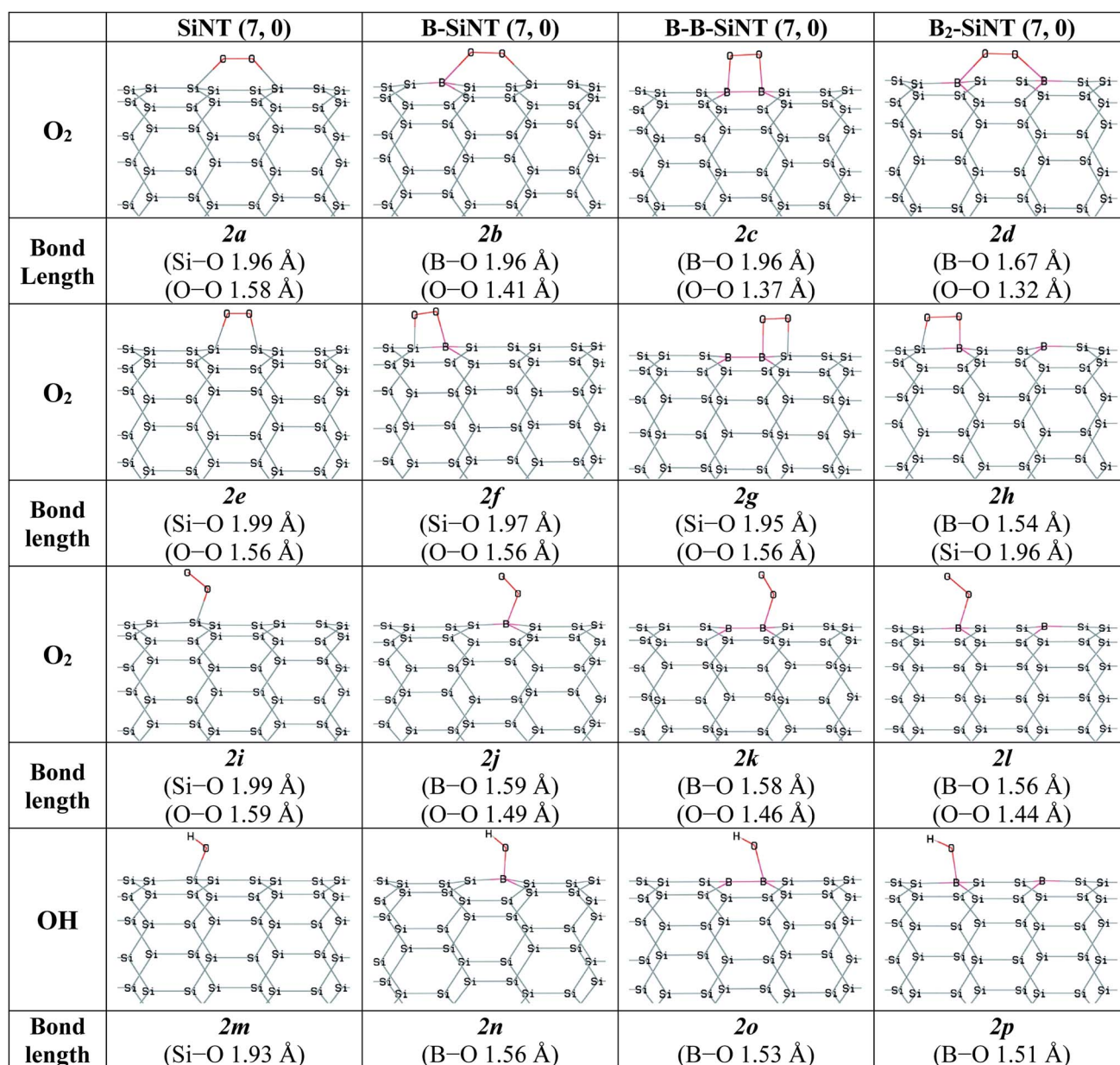
Fig. 2 Complexes of the SiNT (7, 0), B–SiNT (7, 0), B–B–SiNT (7, 0) and B₂–SiNT (7, 0) with O₂ and OH molecules.

Table 2 The ΔE_{ad} (in eV) and ΔG_{ad} (in eV) values of studied complexes

| Complex | ΔE_{ad} | ΔG_{ad} | Complex | ΔE_{ad} | ΔG_{ad} | Complex | ΔE_{ad} | ΔG_{ad} | Complex | ΔE_{ad} | ΔG_{ad} |
|-----------|------------------------|------------------------|-----------|------------------------|------------------------|-----------|------------------------|------------------------|-----------|------------------------|------------------------|
| 2a | -0.41 | -0.36 | 2b | -0.61 | -0.54 | 2c | -0.84 | -0.79 | 2d | -0.88 | -0.81 |
| 2e | -0.39 | -0.33 | 2f | -0.42 | -0.38 | 2g | -0.55 | -0.51 | 2h | -0.61 | -0.55 |
| 2i | -0.27 | -0.24 | 2j | -0.64 | -0.61 | 2k | -0.69 | -0.64 | 2l | -0.73 | -0.68 |
| 2m | -2.07 | -1.97 | 2n | -2.23 | -2.15 | 2o | -2.31 | -2.21 | 2p | -2.39 | -2.28 |
| 3a | -0.87 | -0.81 | 3b | -1.06 | -0.99 | 3c | -1.11 | -1.03 | 3d | -1.17 | -1.12 |
| 3e | -3.11 | -2.97 | 3f | -3.59 | -3.45 | 3g | -3.67 | -3.56 | 3h | -3.79 | -3.67 |
| 3i | -2.24 | -2.13 | 3j | -2.25 | -2.15 | 3k | -2.27 | -2.17 | 3l | -2.31 | -2.21 |
| 3m | -0.15 | -0.10 | 3n | -0.18 | -0.13 | 3o | -0.20 | -0.15 | 3p | -0.24 | -0.19 |
| 3q | -0.09 | -0.05 | 3r | -0.10 | -0.07 | 3s | -0.11 | -0.08 | 3t | -0.14 | -0.12 |

In this study the interactions of important intermediates such as O, H, OOH, OH, H₂O and CO molecules with SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) surfaces in process of ORR were investigated (Fig. 2 and 3). The bonds of Si-O of SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) with molecules are stated. E_{HLG} , q , ΔE_{ad} , ΔG_{ad} of molecules on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) are reported in Tables 1 and 2. OOH and OH intermediates can be adsorb on B site of B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0). The O intermediate has tendency to adsorb on B-Si and B-B bridge positions of B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0). It can be concluded that the complexes of B₂-SiNT (7, 0) with O, H, OOH, OH and H₂O molecules are stable than SiNT (7, 0), B-SiNT (7, 0) and B-B-SiNT (7, 0).

Wang, Xiao and Xia *et al.*⁴⁹⁻⁵¹ calculated the O, OH and OOH adsorption on surfaces of boron-doped graphene nanoribbon, silicon-carbon nano sheets and B and N doped-graphene by theoretical methods, respectively (results reported in Table 4). ΔE_{ad} of O on B-doped graphene, silicon-carbon nano-sheets and N-doped graphene were -3.74, -4.11 and -3.55 eV. ΔE_{ad} of OH on boron-doped graphene, silicon-carbon nano-sheets and N-doped graphene were -2.38, -2.87 and -2.41 eV. ΔE_{ad} of OOH on B-doped graphene, silicon-carbon nano-sheets and N-doped graphene are -1.12, -1.18 and -1.06 eV. ΔE_{ad} values of O, OH and OOH (-1.17, -2.39 and -1.17 eV) on B₂-SiNT (7, 0) in present study are similar to corresponding values of O, OH and OOH on various nano-structures were calculated in previous theoretical works.⁴⁹⁻⁵¹

The H₂O molecule favored to adsorb on above ring position of SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0)

and the average of ΔE_{ad} and ΔG_{ad} values are -0.19 and -0.14 eV. The average of q and E_{HLG} values for adsorption of H₂O molecule on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) surface is 0.16|e| and 1.68 eV. H₂O molecule can be adsorbed on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) surfaces as physical adsorption processes.

In process of ORR the CO can occupy the positions of catalysts and the performance of ORR is reduced and efficiency of catalyst decreases sharply. Previous works showed that reactions between CO molecule and surface of platinum nano-catalyst was powerful (ΔE_{ad} is -1.90 eV) and CO poisoning was happen.^{20,66} The average of ΔE_{ad} and ΔG_{ad} of CO on SiNT, B-SiNT, B-B-SiNT and B₂-SiNT surfaces are -0.11 and -0.08 eV. The average of q and E_{HLG} values for adsorption of CO on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) surface is 0.13|e| and 1.70 eV. The CO molecule can be adsorbed on SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) surfaces as physical adsorption processes. It can be concluded that B₂-SiNT (7, 0) as acceptable catalyst can be endurance to CO poisoning and it can solve the major problem of platinum nano-catalysts.

Wang, Xiao and Xia *et al.*⁴⁹⁻⁵¹ calculated the H₂O and CO adsorption on surfaces of boron-doped graphene nanoribbon, silicon-carbon nano sheets and B and N doped-graphene. The ΔE_{ad} of H₂O on surfaces of B-doped graphene, silicon-carbon nano-sheets and N-doped graphene were -0.24, -0.18 and -0.08 eV. The ΔE_{ad} of CO on surfaces of B-doped graphene, silicon-carbon nano-sheets and N-doped graphene were -0.17, -0.07 and -0.12 eV. The ΔE_{ad} of H₂O and CO (-0.24 and -0.14 eV) on B₂-SiNT (7, 0) in present study are similar to

Table 3 The ΔE_{a} and ΔE_{r} for ORR on B₂-SiNT (7, 0)

| Path | Studied reaction steps | ΔE_{a} (eV) | ΔE_{r} (eV) |
|------|--|----------------------------|----------------------------|
| 1 | O ₂ + B ₂ -SiNT (7, 0) → B ₂ -SiNT (7, 0)-*O ₂ | — | -0.68 |
| 1 | B ₂ -SiNT (7, 0)-*O ₂ + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)-*OOH | 0.00 | -1.07 |
| 1 | B ₂ -SiNT (7, 0)-*OOH + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)-*O + H ₂ O | 0.18 | -2.75 |
| 1 | B ₂ -SiNT (7, 0)-*O + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)-*OH | 0.37 | -1.57 |
| 1 | B ₂ -SiNT (7, 0)-*OH + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)* + H ₂ O | 0.07 | -1.34 |
| 2 | O ₂ + B ₂ -SiNT (7, 0) → B ₂ -SiNT (7, 0)-*O ₂ | — | -0.68 |
| 2 | B ₂ -SiNT (7, 0)-*O ₂ + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)-*OOH | 0.00 | -1.07 |
| 2 | B ₂ -SiNT (7, 0)-*OOH + H ⁺ + e ⁻ → *OH-B ₂ -SiNT (7, 0)-*OH | 0.24 | -2.97 |
| 2 | *OH-B ₂ -SiNT (7, 0)-*OH + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)-*OH + H ₂ O | 0.35 | -1.20 |
| 2 | B ₂ -SiNT (7, 0)-*OH + H ⁺ + e ⁻ → B ₂ -SiNT (7, 0)* + H ₂ O | 0.07 | -1.34 |



| | SiNT (7, 0) | B-SiNT (7, 0) | B-B-SiNT (7, 0) | B ₂ -SiNT (7, 0) |
|------------------|--|--|---|--|
| OOH | | | | |
| Bond length | 3a (Si-O 1.95 Å) (O-O 1.54 Å) | 3b (B-O 1.51 Å) (O-O 1.51 Å) | 3c (B-O 1.49 Å) (O-O 1.49 Å) | 3d (B-O 1.53 Å) (O-O 1.47 Å) |
| O | | | | |
| Bond length | 3e (Si-O 1.85 Å) | 3f (Si-O 1.83 Å) (B-O 1.47 Å) | 3g (B-O 1.45 Å) | 3h (Si-O 1.81 Å) (B-O 1.43 Å) |
| H | | | | |
| Bond length | 3i (Si-H 1.66 Å) | 3j (Si-H 1.64 Å) | 3k (Si-H 1.63 Å) | 3l (Si-H 1.64 Å) |
| H ₂ O | | | | |
| Distance | 3m (distance 3.21 Å) | 3n (distance 3.17 Å) | 3o (distance 3.14 Å) | 3p (distance 3.11 Å) |
| CO | | | | |
| Distance | 3q (distance 3.34 Å) | 3r (3.27 Å) | 3s (distance 3.25 Å) | 3t (distance 3.22 Å) |

Fig. 3 Complexes of the SiNT (7, 0), B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) with O and H atoms, OOH, H₂O and CO molecules.

corresponding values of H₂O and CO on various nanostructures were calculated in previous theoretical works.^{49–51}

3.2. B₂-SiNT (7, 0) as catalyst to ORR

Nano-catalysts processed the chemical reactions through the ER and LH paths. The paths for ORR *via* B₂-SiNT (7, 0) as

acceptable catalyst through the LH and ER mechanisms were investigated. As start, O₂ adsorption is investigated *via* O₂ dissociation or hydrogenation of O₂ to create B₂-SiNT (7, 0)-*OOH. Firstly, the O₂ dissociation process can be defined as B₂-SiNT (7, 0)-*O₂ → *O-B₂-SiNT (7, 0)-*O. The dissociated O atoms were elected to link on B-Si position and activation barrier energy is 0.96 eV (figures 2a (IS), 2b (TS) and 2c (FS)).



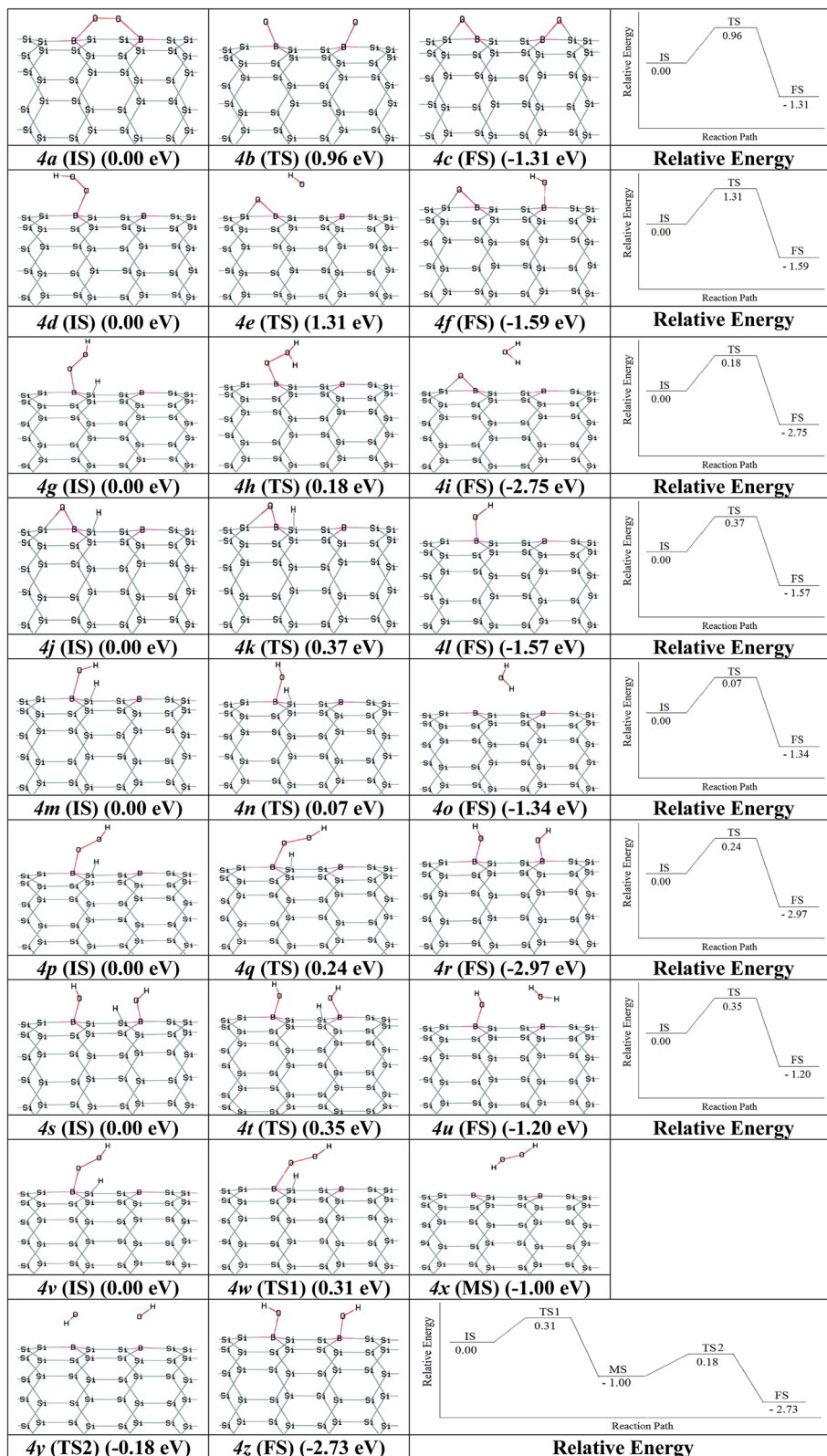


Fig. 4 The intermediates of ORR and relative energies: (1) $B_2-SiNT(7,0)-*O_2 \rightarrow *O-B_2-SiNT(7,0)-*O$; (2) $B_2-SiNT(7,0)-*OOH \rightarrow *O-B_2-SiNT(7,0)-*OH$; (3) $B_2-SiNT(7,0)-*OOH \rightarrow B_2-SiNT(7,0)-*O + H_2O$; (4) $B_2-SiNT(7,0)-*O \rightarrow B_2-SiNT(7,0)-*OH$; (5) $B_2-SiNT(7,0)-*OH \rightarrow B_2-SiNT(7,0)-* + H_2O$; (6) $B_2-SiNT(7,0)-*OOH \rightarrow *OH-B_2-SiNT(7,0)-*OH$; (7) $*OH-B_2-SiNT(7,0)-*OH \rightarrow *OH-B_2-SiNT(7,0) + H_2O$; (8) $B_2-SiNT(7,0)-*OOH \rightarrow B_2-SiNT(7,0) + H_2O_2 \rightarrow *OH-B_2-SiNT(7,0)-*OH$.



Table 4 The ΔE_{ad} (in eV) values of O_2 , O, OH and OOH on B-doped graphene, silicon-carbon nano-sheets and N-doped graphene^{65–67} and B-SiNT (7, 0), B-B-SiNT (7, 0) and B₂-SiNT (7, 0) in this study

| Catalysts/species | B-doped graphene ⁴⁹ | Si-C nano-sheet ⁵⁰ | N-doped graphene ⁵¹ | B-SiNT (7, 0) | B-B-SiNT (7, 0) | B ₂ -SiNT (7, 0) |
|-------------------|--------------------------------|-------------------------------|--------------------------------|---------------|-----------------|-----------------------------|
| O_2 | -0.62 | -0.53 | -0.60 | -0.69 | -0.84 | -0.88 |
| O | -3.74 | -4.11 | -3.55 | -3.59 | -3.67 | -3.79 |
| OH | -2.38 | -2.87 | -2.41 | -2.23 | -2.31 | -2.39 |
| OOH | -1.12 | -1.18 | -1.06 | -1.06 | -1.11 | -1.17 |

Secondary, adsorbed O_2 can interact *via* H atom to create B₂-SiNT (7, 0)-*OOH as follow: B₂-SiNT (7, 0)-*O₂ + H⁺ + e⁻ → B₂-SiNT (7, 0)-*OOH, this process has no any activation barrier energy.

The OOH adsorption on of B₂-SiNT (7, 0) has higher ΔE_{ad} than O_2 *ca.* 0.29 eV and also O_2 dissociation on surface of B₂-SiNT (7, 0) has high activation barrier energy. H atom is added into Si in B₂-SiNT (7, 0)-*OOH and H atom reacted *via* B₂-SiNT

(7, 0)-*OOH. Then the B₂-SiNT (7, 0)-*OOH dissociated to *O-B₂-SiNT (7, 0)-*OH (figures 2d (IS), 2e (TS) and 2f (FS)), due to great activation barrier energy (1.31 eV) this process is impossible. The creation of B₂-SiNT (7, 0)-*OOH in ORR on B₂-SiNT (7, 0) is suitable than dissociation of O_2 molecule.

The ORR is done through the B₂-SiNT-*OOH intermediate as follows:

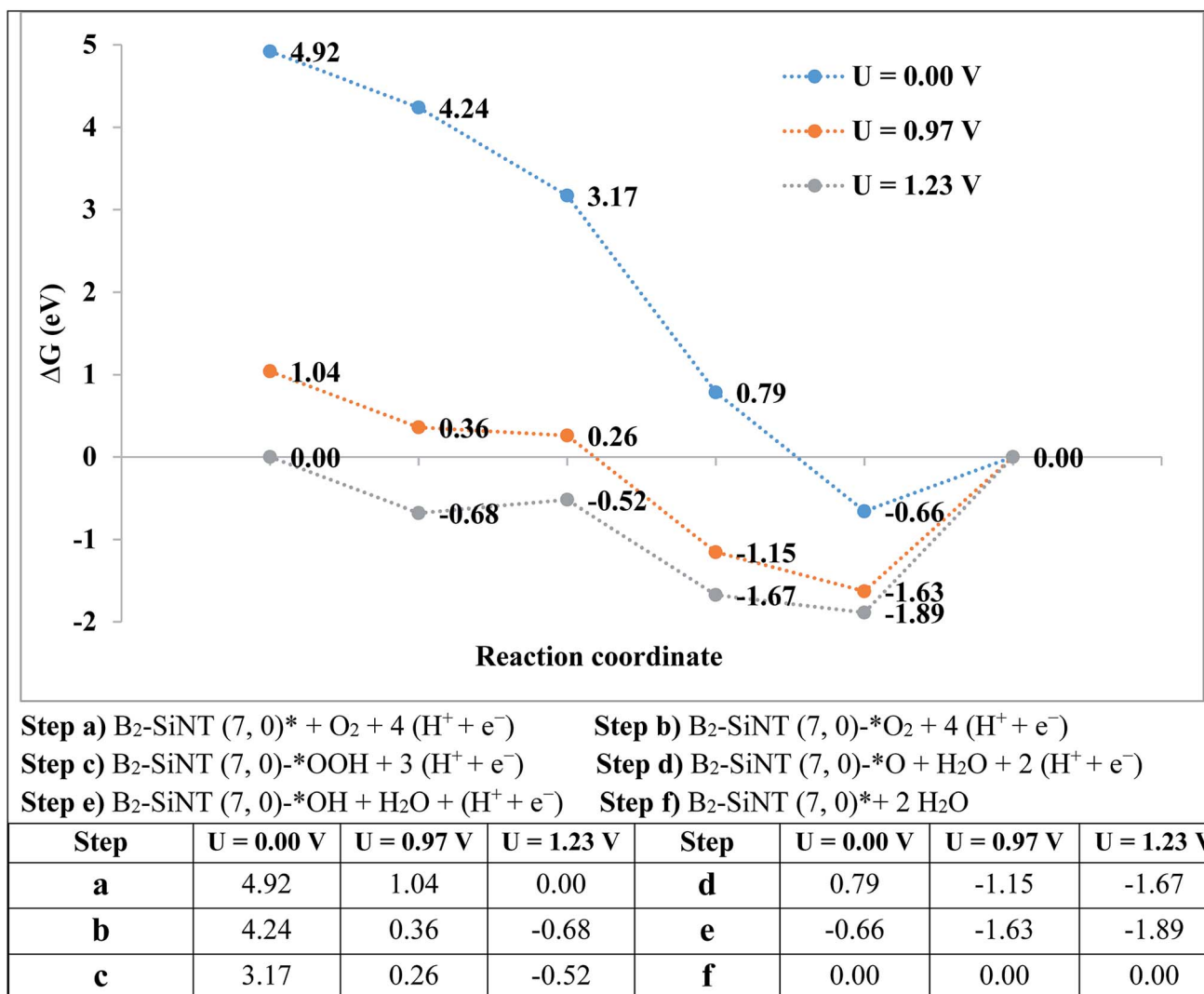
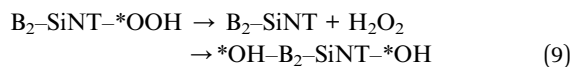
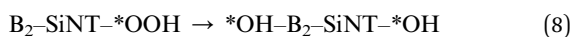


Fig. 5 The G values for ORR on B₂-SiNT (7, 0).



Table 5 The onset-potential (in eV) values for the ORR performed on several catalysts^{105–110}

| Catalyst | Onset potential | Catalyst | Onset potential | Catalyst | Onset potential |
|-------------------------------------|-----------------|------------------------|-----------------|----------------------|-----------------|
| Pd/CNT ¹⁰⁵ | 0.764 | Pd-Ni ¹⁰⁷ | 1.105 | PdNi ¹⁰⁹ | 1.040 |
| Pd/MWCNT ¹⁰⁵ | 1.014 | PtCo/C ¹⁰⁷ | 0.836 | Pd ¹⁰⁹ | 0.901 |
| Pd-Ni(3 : 1)/C ¹⁰⁶ | 1.005 | Pd-Fe/C ¹⁰⁸ | 0.865 | Pd-Cu ¹¹⁰ | 1.001 |
| Pd ₂ Co/C ¹⁰⁶ | 0.735 | Pd/C ¹⁰⁸ | 0.920 | Pt/C ¹¹⁰ | 0.900 |



In path 1, B₂-SiNT (7, 0)-*OOH intermediate was decreased to H₂O molecule and B₂-SiNT (7, 0)-*O (ΔE_a = 0.18 eV). In this process, O-O is fragmented and the first H₂O molecule is created (figures 2g (IS), 2h (TS) and 2i (FS)). Then two hydrogenation stages were done and B₂-SiNT (7, 0)-*OH (figures 2j (IS), 2k (TS) and 2l (FS)) and the second H₂O molecule was created (figures 2m (IS), 2n (TS) and 2o (FS)). The activation barrier energies of these two hydrogenation processes are 0.37 and 0.07 eV, respectively.

In path 2, H atom is linked to O and *OH-B₂-SiNT (7, 0)-*OH is created and activation barrier energy is 0.24 eV (figures 2p (IS), 2q (TS) and 2r (FS)). Then, *OH-B₂-SiNT (7, 0)-*OH linked to H atom and the first H₂O molecule is created and activation barrier energy of this stage is 0.35 eV (figures 2s (IS), 2t (TS) and 2u (FS)). In the end stage of path 2, the B₂-SiNT (7, 0)-*OH is hydrogenated and in this step the second H₂O molecule is separated.

In the path 3, the B₂-SiNT (7, 0)-*OOH is hydrogenated and the H₂O₂ molecule and B₂-SiNT (7, 0) catalyst are created and activation barrier is 0.31 eV (figures 2v (IS), 2w (TS1), 2x (MS), 2y (TS2) and 2z (FS)). The H₂O₂ molecule creation is a mediated state (MS) on surface of B₂-SiNT (7, 0) and it cannot effect on potential of the B₂-SiNT (7, 0), significantly. In next stage of path 3, separated H₂O₂ dissociated into *OH-B₂-SiNT (7, 0)-*OH structure and therefore H₂O₂ dissociation has activation barrier energy about 0.82 eV. The ORR *via* path 3 continued *via* two hydrogenation stage as presented in path 1 in Fig. 4 (figures 4s (IS), 4t (TS) and 4u (FS)) and path 2 in Fig. 4 (figures 4m (IS), 4n (TS) and 4o (FS)).

The parameters of two acceptable paths about reduction of B₂-SiNT (7, 0)-*OOH structure are stated in Table 3. In path 1, rate-determining stage (ΔE_a = 0.37 eV) on B₂-SiNT (7, 0) surface is creation of B₂-SiNT (7, 0)-*OH. In path 2, creation of B₂-SiNT (7, 0)-*OH structure and H₂O molecule is rate-determining stage (ΔE_a = 0.35 eV). In path 2, creation of *OH-B₂-SiNT (7, 0)-*OH has higher ΔE_a than formation of B₂-SiNT (7, 0)-*O + H₂O in path 1 *ca.* 0.06 eV and so path 1 can be considered as optimal pathway to ORR.

The over-potential of ORR on Pt-based compounds and graphene are 0.44 and 0.45 V.^{49–59} The experimental researchers

investigated the onset-potential for the ORR performed on several catalysts^{105–110} and results are stated in Table 5. The *G* of ORR steps are stated in Fig. 5. The level of the final produce (B₂-SiNT (7, 0)-* + 2H₂O) is considered as reference step and ORR steps in *U* = 0 V is downhill. Reaction steps become downward that *U* is decreased to 0.97 V and beginning voltage for ORR is 0.97 V. The B₂-SiNT (7, 0) is suggested as suitable ORR catalyst.

4. Conclusions

Performances of boron-silicon nanotube (7, 0) as novel catalyst to ORR are investigated. The ORR on surface of B₂-SiNT can be continued through LH and ER mechanisms. The rate-determining stage (ΔE_a = 0.35 eV) for ORR on B₂-SiNT (7, 0) surface is creation of B₂-SiNT (7, 0)-*OH structure. The calculated beginning voltage to ORR on surface of the B₂-SiNT (7, 0) is 0.37 V. In the acidic solution the beginning voltage to oxygen reduction process can be evaluated to 0.97 V. Results indicated that the B₂-SiNT (7, 0) is suggested as catalyst to ORR with suitable efficiency.

Conflicts of interest

There are no conflicts to declare.

References

- H. A. Gasteiger and N. M. Marković, *Science*, 2009, **324**, 48.
- J. B. Lee and Y. K. Park, *J. Power Sources*, 2006, **158**, 1251.
- N. M. Marković, T. Schmidt, J. Stamenkovic and V. Ross, *Fuel Cells*, 2001, **1**, 105.
- Z. W. Liu, F. Peng, H. J. Wang and H. Yu, *Catal. Commun.*, 2012, **29**, 11.
- N. S. Lewis and D. G. Nocera, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 15729.
- P. H. Matter and L. Zhang, *J. Catal.*, 2006, **239**, 83.
- M. Winter, *Chem. Rev.*, 2004, **104**, 4245.
- V. I. Zaikovskii, *J. Phys. Chem. B*, 2006, **110**, 6881.
- W. Song and E. J. M. Hensen, *ACS Catal.*, 2014, **4**, 1885.
- J. Liu, *ACS Catal.*, 2016, **7**, 34.
- Y. Nie, L. Li and Z. D. Wei, *Chem. Soc. Rev.*, 2015, **44**, 2168.
- M. E. Scofield, H. Q. Liu and S. S. Wong, *Chem. Soc. Rev.*, 2015, **44**, 5836.
- M. Zhou, H. L. Wang and S. Guo, *Chem. Soc. Rev.*, 2016, **45**, 1273.
- S. B. Yang, L. J. Zhi, K. Tang, X. L. Feng and J. Maier, *Adv. Funct. Mater.*, 2012, **22**, 3634.



- 15 R. Li, Z. D. Wei, X. L. Gou and W. Xu, *RSC Adv.*, 2013, **3**, 9978.
- 16 K. Gong, P. Du, F. Xia and Z. H. Durstock, *Science*, 2009, **323**, 760.
- 17 M. Li, L. Zhang, Q. Xu, J. Niu and Z. J. Xia, *J. Catal.*, 2014, **314**, 66.
- 18 A. Wang, B. Lin and H. Zhang, *Catal. Sci. Technol.*, 2017, **7**, 2362.
- 19 I. A. Pasti, N. M. Gavrilov and A. S. Dobrota, *Electrocatalysis*, 2015, **6**, 498.
- 20 X. J. Bo and L. P. Guo, *Phys. Chem. Chem. Phys.*, 2013, **15**, 2459.
- 21 B. Delley, *J. Chem. Phys.*, 2000, **113**, 7756.
- 22 J. P. Perdew and K. Burke, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 23 A. A. Verma, S. A. Bates, T. Anggara and C. Paolucci, *J. Catal.*, 2014, **312**, 179.
- 24 S. Ata, H. Yoon and C. Subramaniam, *Polymer*, 2014, **55**, 5276.
- 25 D. Janas and A. P. Herman, *Carbon*, 2014, **73**, 225.
- 26 C. C. Lawlor and C. M. Schauerman, *J. Mater. Chem. C*, 2015, **3**, 10256.
- 27 J. Alvarenga, P. R. Jarosz and C. M. Schauerman, *Appl. Phys. Lett.*, 2010, **97**, 19.
- 28 I. Rubinstein and E. Gileadi, *J. Electroanal. Chem.*, 2017, **108**, 191.
- 29 D. Banham and S. Ye, *J. Power Sources*, 2015, **285**, 334.
- 30 G. Wu, K. I. More and P. Xu, *Chem. Commun.*, 2013, **49**, 329.
- 31 S. Gong and Z. H. Zhu, *J. Appl. Phys.*, 2013, **114**, 74303.
- 32 K. Kordek and L. Jiang, *Adv. Energy Mater.*, 2018, **2018**, 1802936.
- 33 J. Liang, Y. Jiao, M. Jaroniec and S. Z. Qiao, *Angew. Chem., Int. Ed.*, 2012, **51**, 11496.
- 34 B. C. Han, *Phys. Rev. B: Condens. Matter Mater. Phys*, 2008, **77**, 075410.
- 35 K. Shinozaki, J. W. Zack, S. Pylypenko, B. S. Pivovar and S. S. Kocha, *Kinet. Catal.*, 2015, **54**, 255.
- 36 A. N. Valisi, *Electrocatalysis*, 2013, **3**, 108.
- 37 N. V. Smirnova, *J. Electrochem. Soc.*, 2013, **162**, 1.
- 38 H. A. Gasteiger, *Appl. Catal., B*, 2005, **56**, 9.
- 39 K. Qu, Y. Zheng, S. Dai and S. Z. Qiao, *Nanoscale*, 2015, **7**, 12598.
- 40 V. Datsyuk, M. Kalyva and K. Papagelis, *Carbon*, 2008, **46**, 833.
- 41 J. C. Li and Z. Q. Yang, *NPG Asia Mater.*, 2018, **10**, 461.
- 42 C. Zhu, S. Fu, J. Song, Q. Shi and D. Su, *Small*, 2017, **13**, 1603407.
- 43 C. Zhu, S. Fu, Q. Shi and D. Du, *Angew. Chem., Int. Ed.*, 2017, **56**, 13944.
- 44 B. Bayatsarmadi, Y. Zheng and A. Vasileff, *Small*, 2017, **13**, 1702002.
- 45 J. C. Li, S. Y. Zhao and P. X. Hou, *Nanoscale*, 2015, **7**, 19201.
- 46 N. S. Parimi, Y. Umasankar and P. Atanassov, *ACS Catal.*, 2012, **2**, 38.
- 47 F. Jaouen, *Electrochim. Acta*, 2007, **52**, 5975.
- 48 Y. Chu, L. Gu, X. Ju and H. Du, *Catalysts*, 2018, **8**, 245.
- 49 L. Wang and H. Dong, *J. Phys. Chem. C*, 2016, **1203**, 117427.
- 50 P. Zhang and B. Xiao, *Sci. Rep.*, 2014, **4**, 3821.
- 51 L. Zhang and Z. Xia, *J. Phys. Chem. C*, 2011, **115**, 11170.
- 52 P. Zhang and J. S. Lian, *Phys. Chem. Chem. Phys.*, 2012, **14**, 11715.
- 53 J. D. Wiggins-Camacho and K. J. Stevenson, *J. Phys. Chem. C*, 2011, **115**, 20002.
- 54 D. Xiong and X. Li, *Catalysts*, 2018, **8**, 301.
- 55 X. Hu and C. Liu, *New J. Chem.*, 2011, **35**, 2601.
- 56 Q. Wei and X. Tong, *Catalysts*, 2015, **5**, 1574.
- 57 Y. Zhao and L. Yang, *J. Am. Chem. Soc.*, 2013, **135**, 1201.
- 58 L. Ferrighi and M. Datteo, *J. Phys. Chem. C*, 2014, **118**, 223.
- 59 A. Bruix, K. M. Neyman and F. Illas, *J. Phys. Chem. C*, 2010, **114**, 14202.
- 60 J. Cai, *Nature*, 2010, **466**, 470.
- 61 T. Wassmann, A. P. Seitsonen and A. M. Saitta, *Phys. Rev. Lett.*, 2008, **101**, 096402.
- 62 Y. H. Lu, R. Q. Wu, L. Shen, M. Yang and Z. D. Sha, *Appl. Phys. Lett.*, 2009, **94**, 122111.
- 63 R. S. Mulliken, *J. Chem. Phys.*, 1955, **23**, 18330.
- 64 N. Govind, *Comput. Mater. Sci.*, 2003, **28**, 250.
- 65 C. T. Campbell and J. R. V. Sellers, *Faraday Discuss.*, 2013, **162**, 9.
- 66 J. K. Nørskov and J. Rossmeisl, *J. Phys. Chem. B*, 2004, **108**, 17886.
- 67 S. Royer and D. Duprez, *ChemCatChem*, 2011, **3**, 24.
- 68 M. S. Chen, Y. Cai and Z. Yan, *Surf. Sci.*, 2007, **601**, 5326.
- 69 M. A. Yu, Y. Feng, L. Gao and S. Lin, *Phys. Chem. Chem. Phys.*, 2018, **20**, 20661.
- 70 J. A. Keith and T. Jacob, *Angew. Chem., Int. Ed.*, 2010, **49**, 9521.
- 71 Y. Sha, T. H. Yu, Y. Liu and B. V. Merinov, *J. Phys. Chem. Lett.*, 2010, **1**, 856.
- 72 X. W. Yu and S. Y. Ye, *J. Power Sources*, 2007, **172**, 145.
- 73 A. J. Cohen, P. Mori-Sánchez and W. Yang, *Chem. Rev.*, 2012, **112**, 289.
- 74 E. Skulason, *Phys. Chem. Chem. Phys.*, 2007, **9**, 3241.
- 75 Y. X. Chen, *Angew. Chem., Int. Ed.*, 2012, **51**, 8500.
- 76 W. Y. Zhang, *Electrochim. Acta*, 2016, **200**, 131.
- 77 H. Jeon, *J. Power Sources*, 2010, **195**, 5929.
- 78 J. Shui, *Sci. Adv.*, 2015, **1**, 1.
- 79 C. Bianchini, *Chem. Rev.*, 2009, **109**, 4183.
- 80 E. DeCarlos and J. G. Ángyán, *J. Chem. Phys.*, 2016, **145**, 124105.
- 81 P. Kasper, *J. Phys. Chem. A*, 2017, **121**, 2022.
- 82 G. Michael and I. S. Bushmarinov, *Science*, 2017, **355**, 49.
- 83 M. Nielsen, *Surf. Sci.*, 2015, **631**, 2.
- 84 M. H. Shao, *J. Power Sources*, 2011, **196**, 2433.
- 85 D. H. Guo, *Science*, 2016, **351**, 361.
- 86 S. E. Wheeler, A. Moran, S. N. Pieniazek and K. N. Houk, *J. Phys. Chem. A*, 2009, **113**, 10376.
- 87 J. Hohenstein, *Chem. Phys.*, 2006, **15**, 128.
- 88 G. Fazio, L. Ferrighi and C. Di Valentin, *J. Catal.*, 2014, **318**, 203.
- 89 V. Tripkovic and E. Skulason, *Electrochim. Acta*, 2010, **55**, 7975.
- 90 M. J. Janik, C. D. Taylor and M. Neurock, *J. Electrochem. Soc.*, 2009, **156**, 126.



- 91 J. X. Zhao, C. R. Cabrera, Z. H. Xia and Z. F. Chen, *Carbon*, 2016, **104**, 56.
- 92 L. H. Gan and J. Q. Zhao, *Phys. E*, 2009, **41**, 1249.
- 93 S. F. Boys and F. Bernardi, *Mol. Phys.*, 1970, **19**, 553.
- 94 L. Ma, J. M. Zhang, K. W. Xu and V. Ji, *Appl. Surf. Sci.*, 2015, **343**, 121.
- 95 B. Luo, *Small*, 2012, **8**, 630.
- 96 J. B. Zhu, *J. Mater. Chem. A*, 2016, **4**, 7422.
- 97 G. Kresse and M. Hafner, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1993, **47**, 558.
- 98 X. K. Kong, *Chem. Soc. Rev.*, 2014, **43**, 2841.
- 99 J. C. Liu, Y. G. Wang and J. Li, *J. Am. Chem. Soc.*, 2017, **139**, 6190.
- 100 Z. Qi, *J. Power Sources*, 2011, **196**, 5823.
- 101 M. Branda, N. C. Hernández, J. F. Sanz and F. Illas, *J. Phys. Chem. C*, 2010, **114**, 1934.
- 102 G. Henkelman, B. P. Uberuaga and H. Jónsson, *J. Chem. Phys.*, 2000, **113**, 9901.
- 103 A. J. Binder, T. J. Toops and R. R. Unocic, *Angew. Chem., Int. Ed.*, 2015, **54**, 13263.
- 104 E. Y. Ko, E. D. Park, H. C. Lee and D. Lee, *Angew. Chem., Int. Ed.*, 2007, **46**, 734.
- 105 S. Zhao, H. Zhang, S. D. House, R. Jin and R. Jin, *ChemElectroChem*, 2016, **3**, 1225.
- 106 G. Ramos-Sánchez and H. Yee-Madeira, *Int. J. Hydrogen Energy*, 2008, **33**, 3596.
- 107 M. Neergat, V. Gunasekar and R. Rahul, *J. Electroanal. Chem.*, 2011, **658**, 25.
- 108 Y. Dai, P. Yu, Q. Huang and K. Sun, *Fuel Cells*, 2016, **16**, 165.
- 109 L. Xiong, Y. X. Huang, X. W. Liu and W. W. Li, *Electrochim. Acta*, 2013, **89**, 24.
- 110 W. E. Mustain, K. Kepler and M. Prakash, *Electrochem. Commun.*, 2006, **8**, 406.

