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Synthesis and application of Bi₂WO₆ for the photocatalytic degradation of two typical fluoroquinolones under visible light irradiation†

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Bismuth tungstate (Bi₂WO₆) was successfully synthesized by a method combining ultrasonic solvothermal treatment and high-temperature calcination. The products were affirmed by X-ray diffraction, scanning electron microscopy, UV-vis diffuse reflection spectroscopy, X-ray photoelectron spectroscopy, and Fourier transform infrared spectroscopy. The characterization results indicated that calcination could improve the crystallinity and visible light utilization capacity of Bi₂WO₆. The photodegradation experiments showed that Bi₂WO₆ calcined at 450 °C for 3 h exhibited better photocatalytic activity for the degradation of norfloxacin and enrofloxacin under visible light irradiation than the catalyst prepared without calcination or calcined at other temperatures. Meanwhile, the effects of the amount of 450-Bi₂WO₆, the initial concentration of targets, and the pH of the solutions on the degradation were studied. Under the optimal conditions, the removal ratios reached to 92.95% (for norfloxacin) and 94.58% (for enrofloxacin) within 75 min. Furthermore, h⁺ and ·O₂⁻ were identified to affect the photodegradation process significantly, and the possible photocatalytic mechanism was proposed. The as-prepared sample was verified to possess good stability and reusability, suggesting its potential application prospect in the treatment of fluoroquinolone antibiotics.

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1 Introduction

Fluoroquinolone antibiotics (FQs), as common clinical antibiotics, have been widely used in the medical industry, animal husbandry and aquaculture.¹ They have been detected in the environment all over the world, such as Asia,^{2,3} European regions,⁴ the United States,⁵ and Australia⁶ in recent years. The ecological risk of FQs is of great concern mainly due to the potential long-term impacts on the environment even at low concentrations.⁷⁻⁹ They could increase the tolerance of bacteria,^{10,11} cause the generation of resistant genes,^{12,13} inhibit the photosynthesis of plants,¹⁴ and gradually threaten both environmental and human health.^{13,15-17} However, the stability of FQs makes their effective removal by conventional treatment processes challenging, such as biodegradation,^{18,19} bio-sorption,²⁰ sludge digestion,²¹ adsorption removal,^{22,23} and so on.

To date, photocatalytic degradation technology has been deemed as one of the most promising treatments for the refractory pollutants.²⁴⁻²⁶ Modern chemistry aims to identify high-efficiency and environmental-friendly photocatalytic materials which could utilize solar energy efficiently to mineralize organic pollutants thoroughly.²⁷⁻³⁰ In the earlier study, bismuth (Bi) was used to modify TiO₂.³¹ Gradually, excellent photocatalytic properties have been found in various Bi-based compounds, such as Bi₂O₃,³² BiOX (X = Cl, Br, I),³³⁻³⁶ Bi₂WO₆,³⁷⁻³⁹ BiVO₄,⁴⁰ Bi₂MoO₆,⁴¹ and so on.

Bismuth tungstate (Bi₂WO₆) is one of the simplest Aurivillius oxides with a layered structure and visible light response (the forbidden bandwidth is about 2.8 eV). As early as 1999, Kudo and Hiji obtained Bi₂WO₆ by a solid-state method and proved that the product had the photocatalytic activity for the first time.⁴² Tang *et al.*⁴³ confirmed its photocatalytic activity for mineralization of chloroform and acetaldehyde. Then the application market of Bi₂WO₆ in photocatalysis was opened. But there were still some limits for Bi₂WO₆ to perform photocatalysis due to the high recombination rate of the photo-generated electron-hole pairs and the boundedness of visible light response region.^{30,38,39} In order to enhance the photocatalytic activity and widen the visible light response region of Bi₂WO₆, some preparation methods have been developed and optimized to control the crystal form, crystallinity, particle size, specific surface area, and so on.^{30,44} Therein, hydrothermal and

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solvothermal synthesis have got into the researcher's good graces in recent years, because the corresponding reaction conditions were comparatively mild, and the obtained products had good dispersibility with a controlled particle size distribution. Factors that would affect the crystal structure of Bi_2WO_6 were usually adjusted to obtain the excellent photocatalytic activity, such as the pH and temperature of the reaction system. Meanwhile, a variety of auxiliary methods were also introduced into the preparation processes to improve the photocatalytic activity. Ultrasonic was introduced for its strong penetrability and excellent rapid mixing capacity,⁴⁵ and calcination was reported to increase the crystallinity and reduce the particle size.^{46,47} However, long reaction time (up to 20–48 h) was usually needed for the hydrothermal/solvothermal stage.^{37–39,48,49}

In this study, a combined method of ultrasonic solvothermal treatment and high-temperature calcination was attempted to synthesize Bi_2WO_6 with a shorter duration. The calcination temperatures were also investigated. The structure, morphology, chemical status and optical properties of the prepared samples were characterized by powder X-ray diffraction (XRD), scanning electron microscopy (SEM), field emission transmission electron microscope (FETEM), X-ray photoelectron spectroscopy (XPS), Fourier transformation infrared spectra (FT-IR) and UV-vis diffuse reflectance spectroscopy (DRS). The photocatalytic capacities were verified by degrading norfloxacin (NOR) and enrofloxacin (ENR) under visible light irradiation. The degradation kinetics were investigated, and the impacts of different reaction factors on the degradation were examined, such as the amount of photocatalyst, the initial concentration of drugs, and initial pH of the degradation system. Additionally, the dominant active species and probable photocatalytic degradation mechanism of NOR were also discussed.

2 Experimental

2.1 Materials

Bismuth(III) nitrate pentahydrate ($\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$, 99.0%), sodium tungstate dihydrate ($\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$, 99.5%), ethylene glycol ($(\text{CH}_2\text{OH})_2$, EG, 96.0%), sodium hydroxide (NaOH, 96.0%), ethanol ($\text{C}_2\text{H}_5\text{OH}$, 99.7%), isopropanol ($(\text{CH}_3)_2\text{CHOH}$, IPA, 99.7%), and triethanolamine ($(\text{CH}_2\text{CH}_2\text{OH})_3\text{N}$, TEOA, 99.0%) were purchased from Sinopharm Chemical Reagent Co., Ltd (Shanghai, China). Norfloxacin ($\text{C}_{16}\text{H}_{18}\text{FN}_3\text{O}_3$, NOR, 98.0%), enrofloxacin ($\text{C}_{19}\text{H}_{22}\text{FN}_3\text{O}_3$, ENR, 98.0%), and *p*-benzoquinone ($\text{C}_6\text{H}_4\text{O}_2$, *p*-BQ, 97.0%) were supplied by Aladdin Company (Shanghai, China). Degussa P25- TiO_2 nanoparticle was purchased from Degussa Corporation (Germany). Indium-tin oxide glass (ITO) was obtained from South China Science & Technology Co., Ltd (Guangzhou, China). All chemicals used in this research were analytical reagent grade without further purification. Milli-Q water (a minimum resistivity of 18.25 $\text{M}\Omega \text{ cm}$) was used as the solvent for all the solutions or dispersions.

2.2 Synthesis of Bi_2WO_6 samples

The solvothermal method was carried out to synthesize Bi_2WO_6 according to a previous study with modification.³⁷ Initially,

0.9701 g $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ was dissolved in 30 mL distilled water to form a homogeneous solution (solution A, 2 mM) and stirred vigorously for 10 min and sonicated for 20 min; 0.3298 g $\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$ was dissolved in 20 mL EG and stirred for 30 min to form solution B (1 mM). The solution B was added dropwise into solution A under ultrasonication for 60 min, and a white suspension was formed. After adjusting the pH of the mixture solution to 6, the solution was further stirred for 60 min. Then, the mixture solution was transferred into a 100 mL Teflon-lined stainless steel autoclave to undergo a solvothermal treatment at 180 °C for 12 h. After that, the autoclave was cooled down naturally to room temperature. The resulting precipitate was rinsed with ultrapure water and ethanol for three times. The obtained product was dried at 80 °C and denoted as bulk- Bi_2WO_6 . Subsequently, the bulk- Bi_2WO_6 samples were calcined at different temperatures ($T = 350$ °C, 450 °C, and 550 °C) for 3 h. The corresponding products were named 350- Bi_2WO_6 , 450- Bi_2WO_6 , and 550- Bi_2WO_6 , respectively.

2.3 Characterization

The crystal structures and phase identification analyses were characterized by XRD (Bruker D8 Advance, Germany) at the angle range of $2\theta = 10$ – 90° using $\text{Cu-K}\alpha$ irradiation ($\lambda = 0.15418$ nm). The morphologies and structures were examined by SEM (Zeiss Sigma HD, Germany) and FETEM (Tecnai G2 F20 S-TWIN TMP, USA). FT-IR was performed on Nicolet iS50 (Thermo Fisher Scientific, USA) spectrophotometer in the range of 400–4000 cm^{-1} . The chemical status and elemental compositions were analyzed by XPS (Thermo Fisher Scientific, USA) with monochromatic Al- $\text{K}\alpha$ source ($h\nu = 1486.6$ eV, 6 mA \times 12 kV), and the deconvolution of the spectrum was performed using the XPS PEAK 41 program with Gaussian functions after the subtraction of a Shirley background. The optical properties of as-prepared samples were investigated by UV-vis DRS (TU-1901, China) with the range of 200–800 nm, and BaSO_4 was used as a reflectance standard. The specific surface areas were recorded by the Brunauer–Emmett–Teller (BET) technique (Quadrasorb SI-3MP, Quantachrome, USA) with nitrogen adsorption–desorption isotherm.

2.4 Photocatalytic degradation of NOR and ENR

The photocatalytic activities of the as-prepared Bi_2WO_6 samples were investigated by the degradation of NOR and ENR at room temperature under visible light irradiation. A Xenon arc lamp (300 W) was used as the visible light source with a UV cutoff filter ($\lambda > 400$ nm). In each experiment, 50 mg of Bi_2WO_6 sample was suspended into 100 mL of NOR or ENR aqueous solution (10 mg L^{-1}). Prior to irradiation, the suspension was stirred in the dark for 60 min to ensure the adsorption–desorption equilibrium. After that, 2 mL of suspensions were withdrawn and centrifuged (10 000 rpm, 10 min) to produce the supernatant which was analyzed on a TU-1901 spectrophotometer, and the obtained result acted as the initial concentration, C_0 . Then, the irradiation was conducted with continuous magnetic stirring, and the sample solutions were withdrawn every 15 min and centrifuged. The concentrations of NOR or ENR (generally



labeled as C_t) were monitored with the spectrophotometer. In this study, 0.1 mol L^{-1} NaOH was used as a reference solution, and the detected λ_{max} of NOR and ENR in the solutions were 273 and 271 nm, respectively (Fig. S1 and S2†). The degradation efficiencies were calculated in the form of $(C_0 - C_t)/C_0$.

2.5 Detection of active species

In order to understand the active species and possible photocatalytic mechanism, the quenching experiments were carried out. The experimentation was similar to the photocatalytic test except that scavengers were incorporated into the reaction solutions prior to the addition of photocatalyst. In this work, 20 mg L^{-1} *p*-benzoquinone (*p*-BQ), 0.01 mol L^{-1} triethanolamine (TEOA), and 0.01 mol L^{-1} isopropanol (IPA) were used as the scavengers in the photodegradation systems to quench superoxide radicals ($\cdot\text{O}_2^-$), photo-holes (h^+), and hydroxyl radicals ($\cdot\text{OH}$), respectively.

2.6 Electrochemical characterization

The Mott-Schottky plots were measured in Na_2SO_4 electrolyte (0.1 mol L^{-1}) at room temperature by an electrochemical workstation (CHI604E, China) at a constant frequency of 1000 Hz using a 5 mV amplitude signal. And the potential ranged from 0.7 V to -0.8 V at a rate of 0.01 V s^{-1} .

The ITO ($10 \text{ mm} \times 10 \text{ mm} \times 1 \text{ mm}$) modified with photocatalyst was used as the working electrode. The Pt plate and saturated calomel electrode (SCE, 0.2415 V vs. NHE) were acted as the counter electrode and reference electrode, respectively.

2.7 Recycling experiments

Recycling experiments were programmed to investigate the reusability and stability of the prepared Bi_2WO_6 samples. The tests were conducted in four sequential cycles. Multiple sets of parallel experiments were performed at the same time, and finally the remaining solutions were merged to recovery the catalysts. After centrifugation and filtration, the recovered catalysts were washed and dried at 80°C . The obtained powders were used in the next cycle experiment. The amount of catalyst used in each experiment was guaranteed to be the same (50 mg).

3 Results and discussions

3.1 Characterizations of the as-prepared samples

Powder XRD patterns of the prepared Bi_2WO_6 samples were shown in Fig. 1. There were similar diffraction peaks for the four samples, indicating that they might have the same crystal form. The main characteristic diffraction peaks situated at $2\theta = 28.3^\circ$, 32.9° , 47.1° , 55.8° , 76.1° are related to the (131), (200), (202), (133) and (333) crystal planes of orthorhombic phase of Bi_2WO_6 (PDF #79-2381), respectively; which were consistent with those in other literatures,^{39,44,48} implying that the as-prepared samples were Bi_2WO_6 with high purity. Although some additional peaks were observed in bulk- Bi_2WO_6 , all of these peaks were disappeared after calcination. As for the bulk- Bi_2WO_6 , the shortened hydrothermal reaction time may affect the crystallinity of the products. However, the characteristic diffraction peaks

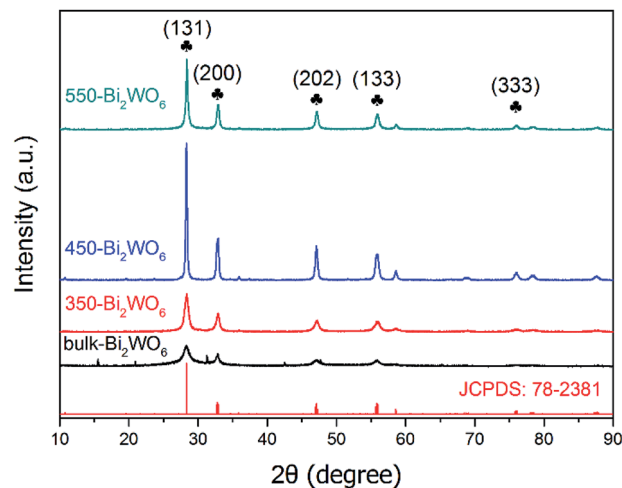


Fig. 1 XRD patterns of the as-prepared Bi_2WO_6 samples.

became more intensive and acute with the increase of calcination temperature, indicative of the growth of Bi_2WO_6 nanocrystals. Fumiaki, *et al.*⁴⁷ and Sheng, *et al.*⁵⁰ had also affirmed that calcination would affect the crystallinity of Bi_2WO_6 .

The morphology and structure were investigated by SEM (Fig. 2). Agglomerated small particles were observed for bulk- Bi_2WO_6 (Fig. 2(a)), while flower-like, lamellar-like and rod-like nanostructures were observed for 350- Bi_2WO_6 , 450- Bi_2WO_6 , and 550- Bi_2WO_6 (Fig. 2(b-d)), respectively. It suggested that calcination could affect the morphologies of Bi_2WO_6 samples efficiently. The surface areas of bulk- Bi_2WO_6 , 350- Bi_2WO_6 , 450- Bi_2WO_6 , and 550- Bi_2WO_6 gauged by BET analysis were 3.73, 9.53, 15.83, and $15.27 \text{ m}^2 \text{ g}^{-1}$, respectively. Among them, 450- Bi_2WO_6 had the largest surface area. FETEM was used to study the ultrastructure of 450- Bi_2WO_6 sample, and the typical image was presented in Fig. S3.† The lattice fringe spacing was about 0.315 nm, which was corresponding to the (131) crystal plane lattice fringes of orthorhombic Bi_2WO_6 .^{39,51}

UV-vis DRS was used to analyze the optical property of catalysts and deduce their forbidden bandwidth (E_g). Fig. 3 showed the UV-vis absorption spectra and Tauc plots ($(\alpha h\nu)^{1/2}$ vs. $h\nu$) of the samples. There were strong broad absorption bands in the range from 200 to about 450 nm for all the as-prepared samples (Fig. 3(a)), indicating their potential photocatalytic activity under the visible light. The optical absorption edge of bulk- Bi_2WO_6 was about 435 nm and its E_g value was estimated to be 2.85 eV, which was in agreement with the results from other literatures.^{48,51} Nevertheless, for 350- Bi_2WO_6 , 450- Bi_2WO_6 , and 550- Bi_2WO_6 , the optical absorption edges were 464, 482, and 471 nm, respectively. Undoubtedly, the absorption edges of the calcined samples were red-shifted compared with that of bulk- Bi_2WO_6 . The forbidden bandwidths (E_g) of 350- Bi_2WO_6 , 450- Bi_2WO_6 , and 550- Bi_2WO_6 were also calculated with the Kubelka-Munk equation (eqn (1)) based on the optical absorption edge from the UV-vis DRS spectrum. The values were 2.67, 2.57, and 2.63 eV, respectively (Fig. 3(b)). It was validated that the visible light utilization capacity of the Bi_2WO_6 samples could be improved by calcination. Among the four as-prepared



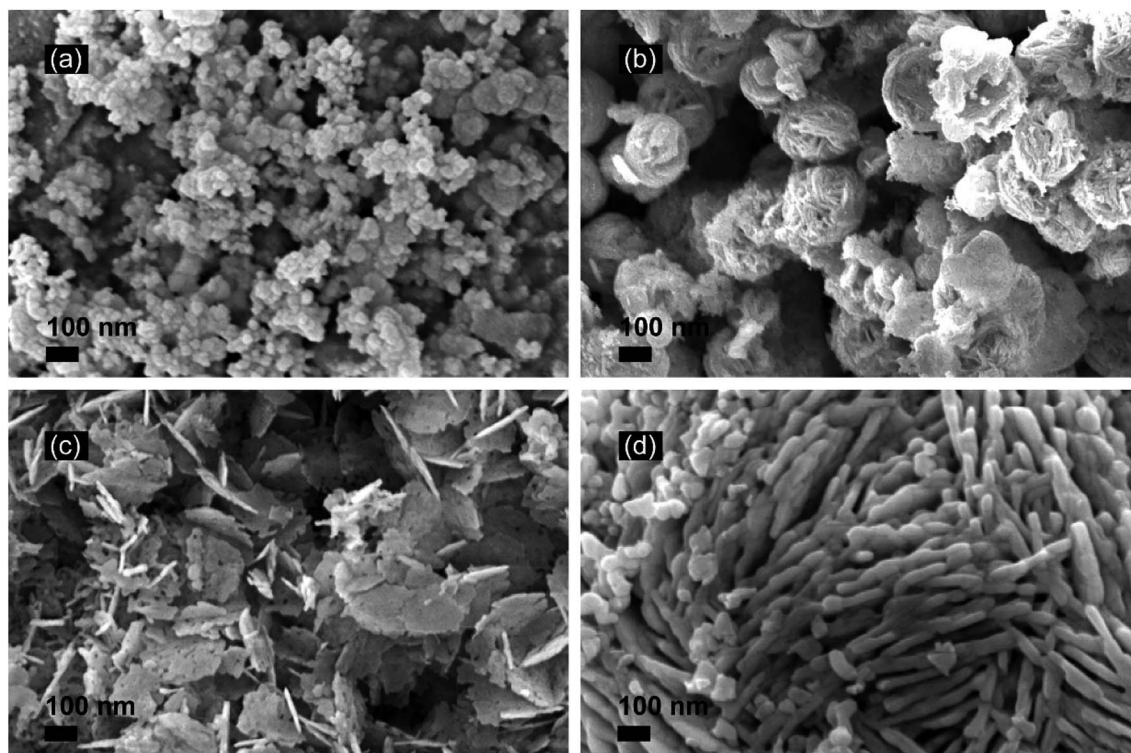


Fig. 2 SEM images of (a) bulk-Bi₂WO₆, (b) 350-Bi₂WO₆, (c) 450-Bi₂WO₆, (d) 550-Bi₂WO₆.

samples, 450-Bi₂WO₆ held the widest absorption region and the lowest band gap. Following is the Kubelka–Munk equation:

$$\alpha h\nu = A(h\nu - E_g)^{\frac{n}{2}} \quad (1)$$

where α is the absorption coefficient, h is Planck's constant, 6.626×10^{-34} J s, ν is the optical frequency (s^{-1}), A is a constant (generally 1), E_g is the forbidden bandwidth (eV), and n is decided by the optical transition type of typical semiconductors. Ordinarily, n is 1 or 4, and 1 was selected in this research.⁵²

Surface status and elemental composition of 450-Bi₂WO₆ were examined by XPS spectra (Fig. 4). As shown in Fig. 4(a), the presence of Bi, W, O, and C elements could be identified in the sample, and their mole ratio was estimated to be 1.93 : 1 : 6.04

(Bi : W : O) based on the XPS results, which were consistent with the results of other literature.³⁸ The C element might be derived from the pollutants in the XPS instrument or from the air. The binding energy of C 1s was used as the correction standard (284.8 eV). The two strong peaks located at 159.10 and 164.40 eV were attributed to Bi 4f_{7/2} and Bi 4f_{5/2} of Bi³⁺, respectively (Fig. 4(b)) (with the splitting energy $\Delta = 5.40 \pm 0.10$ eV).⁵¹ As shown in Fig. 4(c), the peaks at 35.37, 37.47, and 29.15 eV should be indexed to the W⁶⁺ in 450-Bi₂WO₆ sample.⁵³ The XPS spectrum in the O 1s region could be fitted into two peaks (Fig. 4(d)), located at 530.01 and 532.20 eV, which were corresponded to the [WO₄]²⁻ layers and crystal lattice oxygen [Bi₂O₂]²⁺ of 450-Bi₂WO₆.^{38,54}

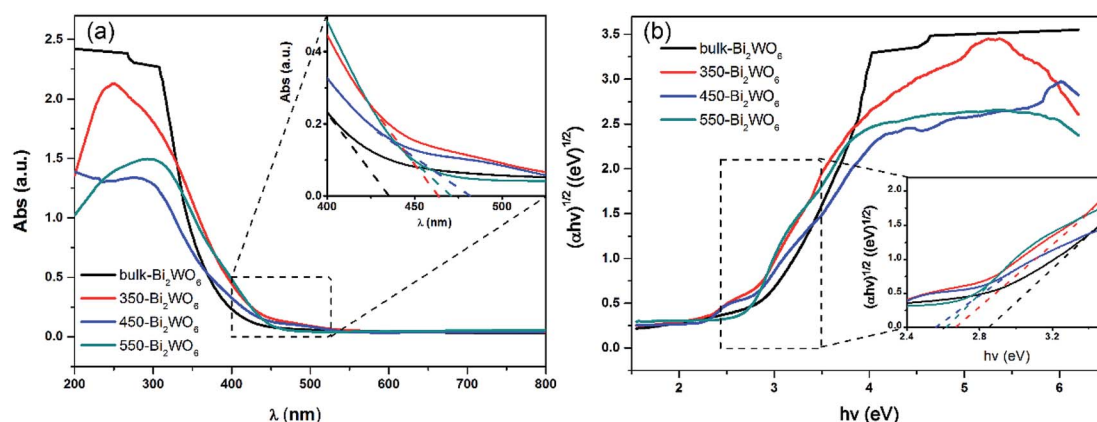


Fig. 3 (a) UV-vis DRS spectra and (b) forbidden bandwidths of Bi₂WO₆ samples.



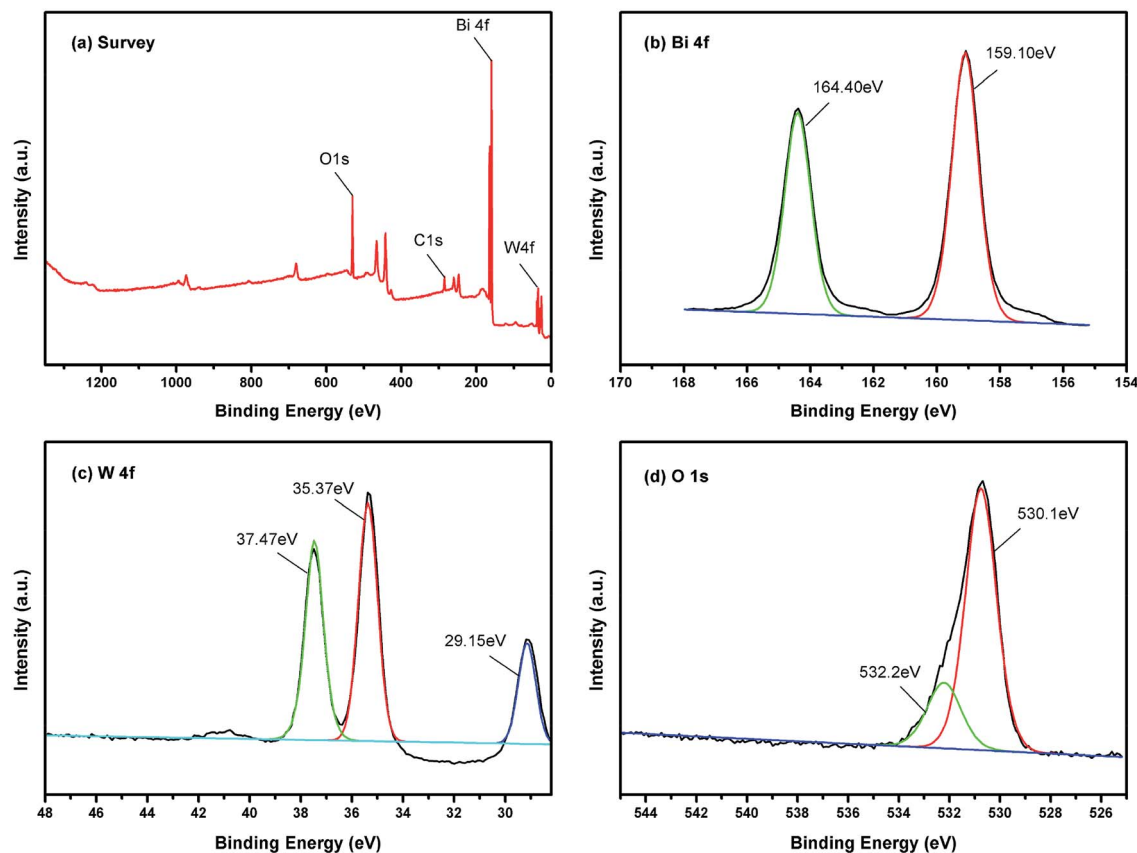


Fig. 4 High-resolution XPS of 450-Bi₂WO₆ (a) total survey, (b) Bi 4f, (c) W 4f, (d) O 1 s.

The FT-IR spectra of the Bi₂WO₆ samples were displayed in Fig. 5. The Bi–O–Bi asymmetric stretch located at 580 cm⁻¹, while W–O was stretching at 722 and 1065 cm⁻¹, and Bi–O was stretching at 810 cm⁻¹, which were matched well with the literatures.^{55,56} The absorption bands at 1631 and 3420 cm⁻¹ could be ascribed to the stretching and bending modes of O–H on the surface of the samples.⁴⁸

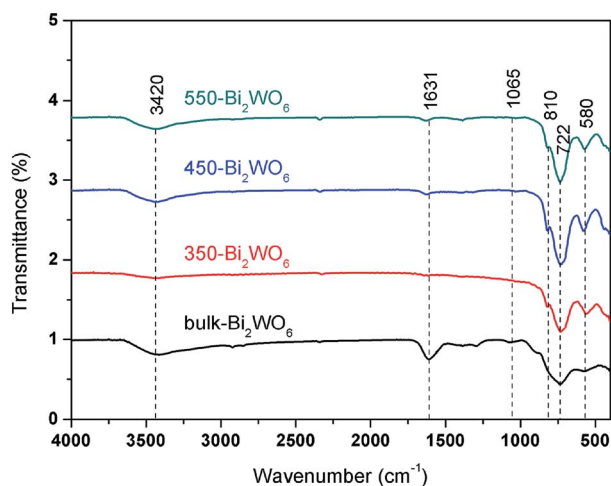


Fig. 5 FT-IR spectra of Bi₂WO₆ samples.

3.2 The kinetics of photocatalytic degradation

In order to investigate the kinetics, the as-prepared Bi₂WO₆ samples (50 mg) were put into 100 mL of 10 mg L⁻¹ NOR or ENR solution to undergo photocatalytic degradation, and P25-TiO₂ was used for comparison. Before irradiation, the adsorption-desorption experiment was firstly conducted in the dark, and the adsorption could reach equilibrium in 60 min (shown in Fig. S4†). The results also indicated that the adsorption ratio of the as-prepared catalyst was so small that the amount could be neglected. The Langmuir–Hinshelwood kinetic equation was usually used to calculate the rate constant of photocatalytic degradation, which was demonstrated in eqn (2):^{57,58}

$$r = -\frac{dC}{dt} = k_a k_b \frac{C_t}{1 + k_b C_0} \quad (2)$$

where r is the degradation rate of substrates (mol (L⁻¹ min⁻¹)), k_a is the physical constant of the reaction system (*i.e.*, the rate constant of solute molecules adsorbed on the surface of the catalyst (mol min⁻¹)), k_b is the photodegradation rate constant of reactive substrates (mol⁻¹), C_t and C_0 are the concentrations of the reactants at time t and “ $t = 0$ ” (mg L⁻¹).

When the initial concentration of substrates was sufficiently low, $k_b C_0 \ll 1$, the eqn (2) could be reduced into eqn (3):

$$r = -\frac{dC_t}{dt} = k_a k_b C_t = k_1 C_t \quad (3)$$



The eqn (4) was obtained from eqn (3) by integral processing:

$$\ln \frac{C_0}{C_t} = k_1 t \quad (4)$$

where $k_1 = k_a k_b$, which is the pseudo-first-order rate constant (min^{-1}). The slope of $\ln(C_0/C_t)$ vs. t is k_1 .

When there were secondary reactants or two first-order reactants, the reaction could be pseudo-second-order kinetics. The eqn (2) was simplified and integral processed to eqn (5) and (6):

$$r = -\frac{dC_t}{dt} = k_2 C_t^2 \quad (5)$$

$$\frac{1}{C_t} - \frac{1}{C_0} = k_2 t \quad (6)$$

where k_2 is the pseudo-second-order rate constant ($\text{L mol}^{-1} \text{min}^{-1}$). The fitted slope of $1/C_t$ vs. t is k_2 .

The degradation results were shown in Fig. 6 and Table 1. As shown in Fig. 6, all prepared Bi_2WO_6 samples exhibited higher activity than P25-TiO_2 . Moreover, the removal ratios of NOR and ENR with the presence of 450- Bi_2WO_6 could reach 85.36% and 88.01% within 75 min, respectively; suggesting an excellent

photocatalytic activity for the degradation of NOR and ENR. The coefficients (R^2) of pseudo-first-order kinetic model were ca. 0.99, which were much better than that of the pseudo-second-order kinetic model for the calcined Bi_2WO_6 samples. It implied that the pseudo-first-order kinetic model was more suitable for the prepared Bi_2WO_6 samples in this test. The sample of 450- Bi_2WO_6 had the highest rate constant (0.0261 min^{-1} of NOR and 0.0289 min^{-1} of ENR), and then it was chosen as the catalyst in the subsequent experiments.

3.3 Photocatalytic activity

3.3.1 Effect of dosage. The effect of the photocatalyst dosage (ranging from 0.3 to 1.2 g L^{-1}) on the NOR or ENR degradation was investigated with 10 mg L^{-1} NOR or ENR solution at the original pH. As shown in Fig. 7(a) and (b), both of the photocatalytic degradation efficiencies of NOR and ENR increased firstly and then decreased with the increase of photocatalyst dosage (the pseudo-first-order kinetic results were shown in Fig. S4(a) and (b)[†]). When the dosage was 0.5 g L^{-1} , the degradation efficiencies of NOR and ENR were the maximum with the values of 86.62% and 89.44% within 75 min, respectively. When the amount of catalyst was not enough, the

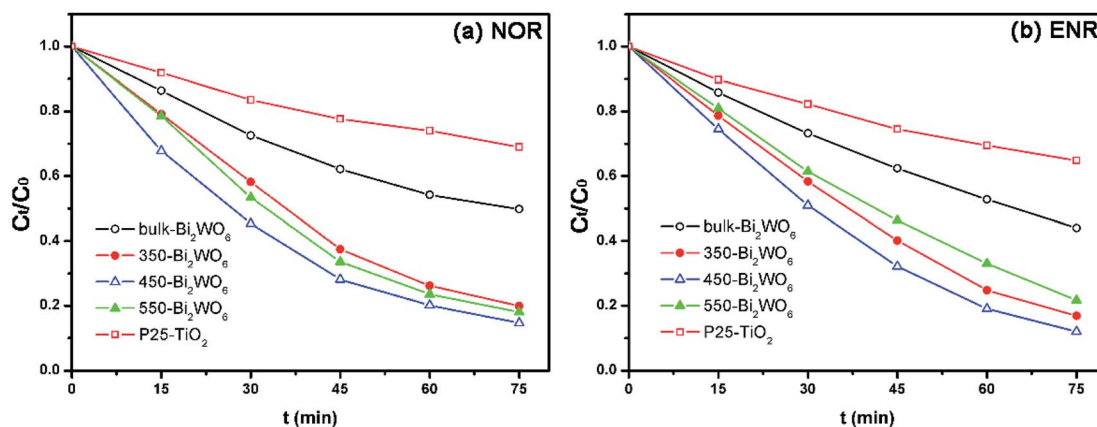


Fig. 6 Photocatalytic degradation of (a) NOR and (b) ENR with different photocatalysts (10 mg L^{-1} drug solution, 0.5 g L^{-1} photocatalyst, initial pH of 10.2 for NOR and of 10.7 for ENR).

Table 1 Kinetics parameters for degradation of NOR and ENR with photocatalysts^a

Drugs	Materials	Pseudo first-kinetic		Pseudo second-kinetic	
		k_1 (min^{-1})	R^2	k_2 ($\text{L mol}^{-1} \text{min}^{-1}$)	R^2
NOR	Bulk- Bi_2WO_6	0.0096	0.9885	1.4287×10^{-5}	0.9953
	350- Bi_2WO_6	0.0225	0.9910	4.6973×10^{-5}	0.9312
	450- Bi_2WO_6	0.0261	0.9954	6.9244×10^{-5}	0.9438
	550- Bi_2WO_6	0.0241	0.9905	5.5157×10^{-5}	0.9410
	P25- TiO_2	0.0049	0.9861	0.5437×10^{-5}	0.9946
ENR	Bulk- Bi_2WO_6	0.0109	0.9890	1.7459×10^{-5}	0.9727
	350- Bi_2WO_6	0.0243	0.9840	6.4410×10^{-5}	0.8671
	450- Bi_2WO_6	0.0289	0.9894	6.9452×10^{-5}	0.8494
	550- Bi_2WO_6	0.0202	0.9840	4.4378×10^{-5}	0.8685
	P25- TiO_2	0.0058	0.9913	0.6309×10^{-5}	0.9986

^a The operating conditions were given in the caption of Fig. 6.



given light energy cannot be converted into chemical energy entirely so that the light energy failed to be fully utilized. With the increase of the catalyst, there were more active centers, which could effectively utilize the light energy to produce more charge carriers and active species. However, when the amount of catalyst exceeded 0.5 g L^{-1} , the degradation efficiencies of NOR and ENR were slightly reduced. This might be attributed to the light shielding effect since more Bi_2WO_6 powders suspended in the solution and scattered the light. Correspondingly, the efficiency of the light and the degradation ratios were decreased.^{59,60} In this test, 0.5 g L^{-1} 450- Bi_2WO_6 was selected as the optimum dosage.

3.3.2 Effect of initial concentration. The effect of initial NOR or ENR concentration (ranging from 10 mg L^{-1} to 30 mg L^{-1}) on the degradation efficiency was studied in the solutions with 0.5 g L^{-1} 450- Bi_2WO_6 at original pH, and the results were shown in Fig. 7(c) and (d) (the pseudo-first-order kinetic results were shown in Fig. S4(c) and (d)†). With the initial concentration increasing from 10 mg L^{-1} to 30 mg L^{-1} , the degradation efficiencies decreased from 86.62% to 72.92% for NOR, and from 89.44% to 74.07% for ENR within 75 min, respectively. This may be due to that the continuous increase of the substrates would cause the increase of light scattering. Moreover, with the increase of the initial concentration, more specific active centers would be occupied. Then, the adsorption of H_2O , O_2 , OH^- , and other molecules on active centers would reduce. Accordingly, the photocatalytic degradation efficiencies would also decrease.³⁷ Generally, the photocatalytic oxidation technology was more suitable for the treatment of pollutants with lower concentrations.

3.3.3 Effect of pH. The initial pH of the substrate solutions could have a significant influence on the charge properties of the catalyst surface, the adsorption of the solute molecules on the catalyst surface, as well as the dispersion and dissociation of the solvent molecules. Thus, the pH would play a key role in affecting the photocatalytic degradation.⁶¹ As shown in Fig. 7(e) and (f), with the presence of 0.5 g L^{-1} 450- Bi_2WO_6 and 10 mg L^{-1} NOR or ENR, the degradation efficiencies of NOR (ENR) were 92.95% (94.58%), 88.84% (89.90%), 84.32% (85.73%), 78.74% (80.83%) and 86.62% (89.44%) within 75 min at pH 3.0, 5.0, 7.0, 9.0, 10.2/10.7 (the original pH values of NOR/ENR solutions), respectively. The pseudo-first-order kinetic results were shown in Fig. S4(e) and (f)†. The degradation efficiencies in acidic conditions were significantly higher than that in alkaline conditions for both targeted substances. The highest degradation efficiencies of NOR and ENR were obtained at pH 3.0, and the rate constants were 0.0354 and 0.0389 min^{-1} , respectively. Firstly, H^+ could be trapped by O_2 on the surface ($\text{O}_{2\text{sur}}$) of 450- Bi_2WO_6 in the acidic condition to form $\cdot\text{OH}$ through a series of oxidation–reduction reactions. The $\cdot\text{OH}$ had a strong oxidation activity to react with the organics in water. Secondly, the pH_{pzc} of 450- Bi_2WO_6 was 3.5–4.5, which was obtained from the mass titration method and other literatures.^{62,63} The surface of 450- Bi_2WO_6 was positively charged at pH 3.0. It facilitated the formation of $\cdot\text{O}_2^-$ ($\text{e}^- + \text{O}_{2\text{sur}} \rightarrow \cdot\text{O}_2^-$) from the reactions of photo-electrons (e^-) and $\text{O}_{2\text{sur}}$ of Bi_2WO_6 , which had a strong oxidizing property for degrading the organics. The

reactions were also benefited for the separation of e^- and h^+ , which could improve photocatalytic activity. Therefore, pH 3.0 was selected as the optimal pH during the degradation experiments.

3.3.4 Comparison of the as-prepared Bi_2WO_6 to the already reported ones. As we known, Bi_2WO_6 is a kind of photocatalyst with good photocatalytic performance. Since its first synthesis, the preparation methods were adjusted by different researchers to improve its photocatalytic activity or utilizability. Table 2 summarized some applications of Bi_2WO_6 on the photo-degradation of antibiotics. As can be seen in the table, though the solvothermal reaction time was shortened in this study, the photocatalytic capacity of the prepared catalysts was comparable with that of the reported Bi_2WO_6 . It may be indicated that calcination could remedy the deficiency caused by the shortened hydrothermal reaction time. It had been ever confirmed that calcination at an appropriate temperature could promote the further growth of Bi_2WO_6 and improve the photo-degradation activity.⁵⁰

3.4 Detection of active species

To evaluate the role of active species ($\cdot\text{O}_2^-$, h^+ , and $\cdot\text{OH}$) during the photocatalytic degradation, the quench experiments were carried out for NOR by adding different scavengers. As can be seen in Fig. 8, the maximum removal ratio of NOR was achieved in the absence of any scavenger species. When *p*-BQ, TEOA, and IPA were added into the reaction systems, the removal ratios of NOR decreased from 90.62% (without scavenger) to 41.94%, 29.13%, and 73.78%, respectively. The pseudo-first-order rate constants were also decreased to 0.0072, 0.0045 and 0.0178 min^{-1} with the presence of *p*-BQ, TEOA, or IPA from 0.0354 min^{-1} , respectively. Thus, it could be reasonably deduced that the h^+ , $\cdot\text{O}_2^-$, and $\cdot\text{OH}$ were generated during the visible light irradiation with the presence of 450- Bi_2WO_6 , and participated in the photocatalytic reaction. Compared with $\cdot\text{OH}$, h^+ and $\cdot\text{O}_2^-$ played a significant role in the photo-degradation process.

3.5 Possible mechanism of photodegradation

The band edge positions of the valence band (VB) and conduction band (CB) of the semiconductor could be estimated by the following empirical equations:

$$E_{\text{VB}} = \chi - E^{\text{c}} + 0.5E_{\text{g}} \quad (7)$$

$$E_{\text{CB}} = \chi - E^{\text{c}} - 0.5E_{\text{g}} \quad (8)$$

where χ is the absolute electronegativity of compounds, E^{c} is the relative potential energy of free electron vs. NHE. E_{VB} and E_{CB} are the energy of VB and CB of semiconductor, respectively.

According to the principle of electronegativity equalization, the χ of 450- Bi_2WO_6 was calculated to be 6.39 eV, and the E^{c} was about 4.50 eV according to the previous research.³⁷ Fig. 3(b) showed the E_{g} of 450- Bi_2WO_6 sample was about 2.57 eV. Thus, the E_{VB} and E_{CB} of 450- Bi_2WO_6 sample were calculated to be 3.17 eV and 0.60 eV according to eqn (7) and (8), respectively,



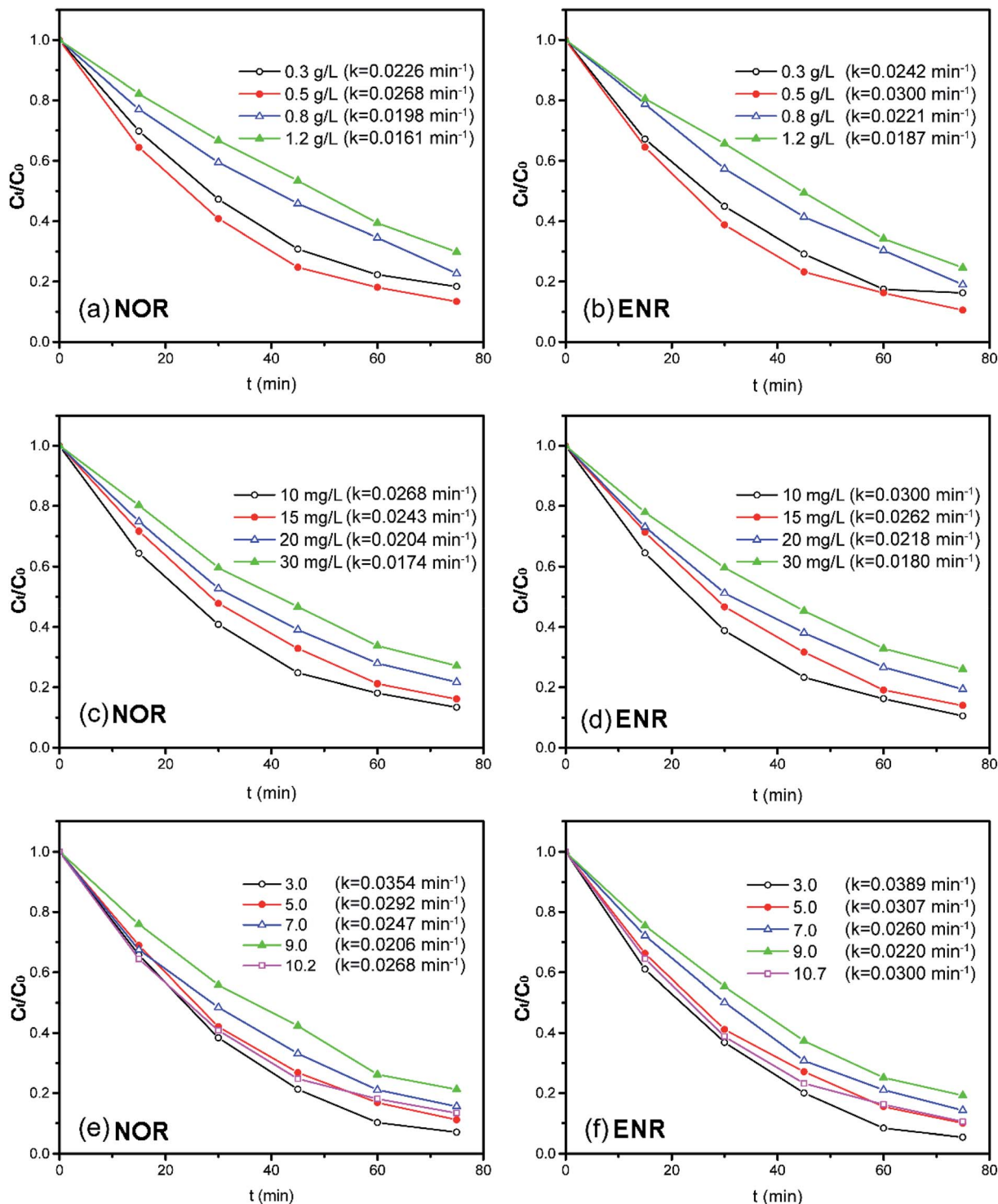


Fig. 7 Effects of the operation parameters on the degradation of NOR or ENR (different amounts of 450-Bi₂WO₆ for the solutions with the initial concentration of 10 mg L⁻¹ at original pH; solutions with different initial concentrations at 450-Bi₂WO₆ dosage of 0.5 g L⁻¹ and original pH; solutions with the initial concentration of 10 mg L⁻¹ at 450-Bi₂WO₆ dosage of 0.5 g L⁻¹ and different initial pH).

which were in accordance with the results from the literatures.^{48,64}

In general, when the visible light irradiated on the 450-Bi₂WO₆ particles suspended in aqueous solution, the photocatalyst absorbed the light energy ($h\nu$) and the electrons in the

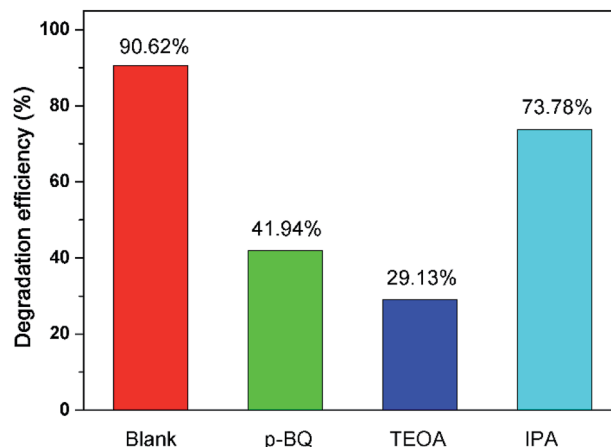
VB were excited to the CB, resulting in e^- in the CB and h^+ in the VB. The e^- and h^+ migrated to the surface of the catalyst and contributed to a series of reactions to produce active species.

From the view of thermodynamics, the amount of $\cdot\text{O}_2^-$ generated in the system should be greatly lower than that of

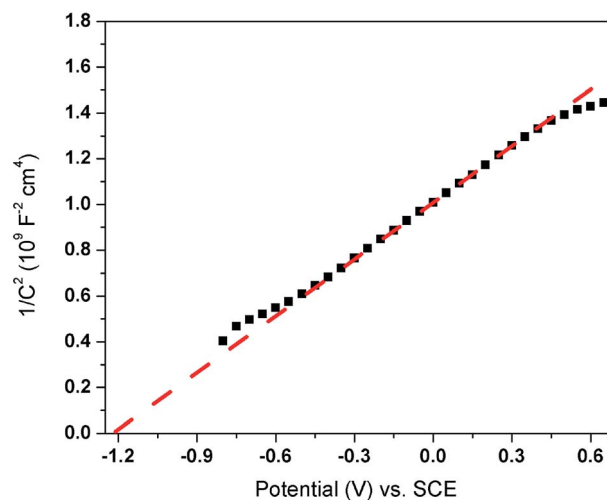


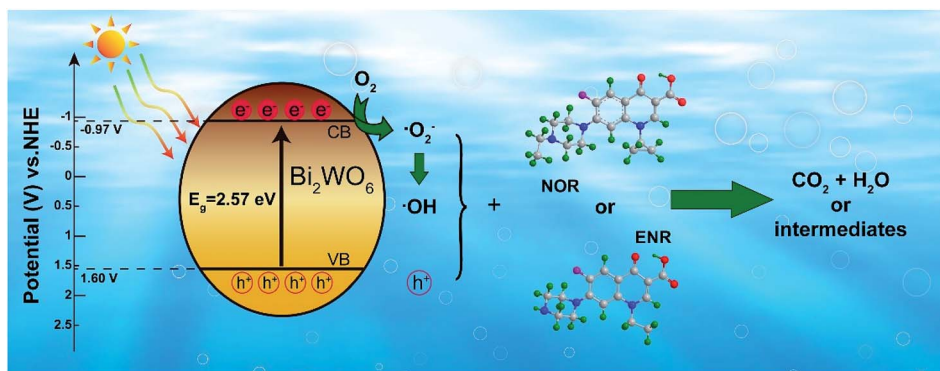
Table 2 Comparison of the as-prepared Bi₂WO₆ to that in the already reported literatures

Material	Preparation method	Solvothermal/hydrothermal time	Morphology	Substrate solution	Solid-to-liquid ratio	Degradation efficiency	Reference
Bi ₂ WO ₆	Hydrothermal method with the reaction solution at pH 4	20 h	Ultrathin nanoflakes	100 mL of 10 mg L ⁻¹ norfloxacin at pH 9	1 g L ⁻¹	90% (60 min)	1
Sponge-loaded Bi ₂ WO ₆	Solvothermal method	20 h	— ^a	70 mg L ⁻¹ tetracycline hydrochloride aqueous solution	8 g L ⁻¹	94% (90 min)	32
Bi ₂ WO ₆	Ultrasonic assisted hydrothermal method	20 h	Nanocuboids	100 mL of 10 mg L ⁻¹ levofloxacin at pH 7.14	0.75 g L ⁻¹	80% (150 min)	47
Bi ₂ WO ₆	Hydrothermal process	20 h	Nanoflowers	100 mL 10 mg mL ⁻¹ ceftriaxone sodium aqueous solution	1 g L ⁻¹	70.18% (240 min)	62
Bi ₂ WO ₆	Ultrasonic assisted hydrothermal method combined with calcination	12 h	Lamellar-like nanostructures	100 mL of 10 mg L ⁻¹ norfloxacin or enrofloxacin at pH 3	0.5 g L ⁻¹	92.95% for norfloxacin (75 min) and 94.58% for enrofloxacin (75 min)	The present study

^a Not mentioned.Fig. 8 Effect of different scavengers on the degradation efficiency of NOR with 450-Bi₂WO₆.

·OH, since E_{CB} of 450-Bi₂WO₆ was about 0.60 eV, which was positive than -0.33 eV (vs. NHE).⁵⁵ And the E_{VB} of 450-Bi₂WO₆ was calculated to be 3.17 eV, more positive than the potential of OH⁻/·OH (2.38 V vs. NHE) and H₂O/·OH (2.72 V vs. NHE). It seemed that the ·OH could be formed from h⁺ and OH⁻ or H₂O.^{36,65,66} This reasoning contradicted with the results of the quench experiments, which showed that ·O₂⁻ exhibited a strong degradation effect than ·OH, and h⁺ was the most important active species in this study. The possible reason might be the lack of rigor in calculation of the E_{CB} and E_{VB} of the catalyst, since the formula was based on the condition of pH = 1,⁶⁷ while pH of the investigated system was 10.2. In view of the effects of pH on surface charge and even the position of energy band, the E_{CB} and E_{VB} of 450-Bi₂WO₆ were likely to undergo a negative shift compared with the calculated values.⁶⁷ For proving the conjecture, the Mott-Schottky test (Fig. 9) was performed, and it was found that the Fermi energy level of the 450-Bi₂WO₆ was -1.21 V vs. SCE (*i.e.* -0.97 V vs. NHE).^{68,69} Correspondingly, the E_{CB} value of 450-Bi₂WO₆ should be -0.97 eV, suggesting that the reaction of producing ·O₂⁻ from

Fig. 9 Mott-Schottky plots of 450-Bi₂WO₆.



Scheme 1 Possible photocatalytic mechanism of the 450-Bi₂WO₆ sample.

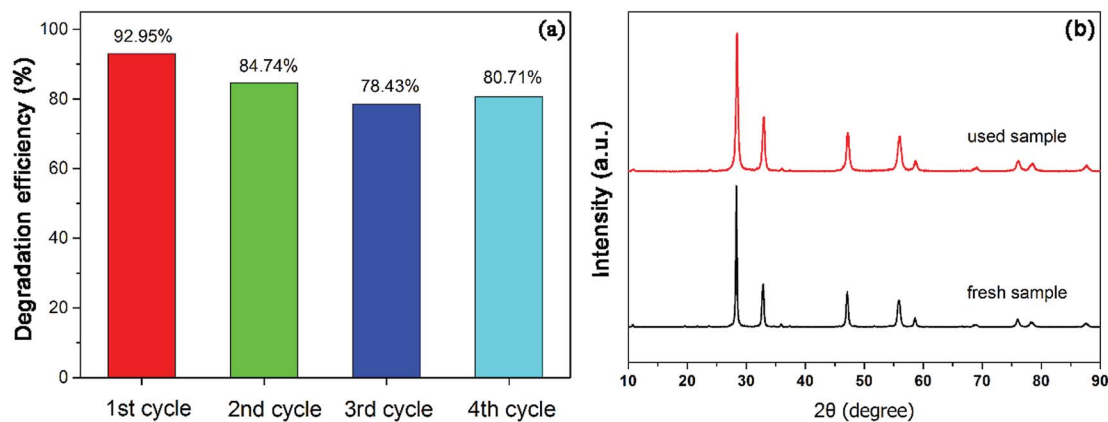
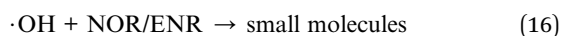
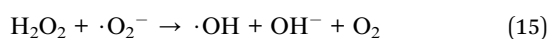
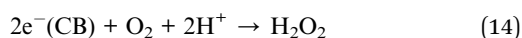
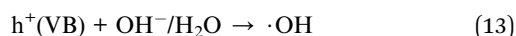
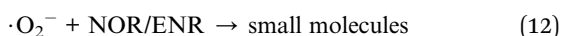
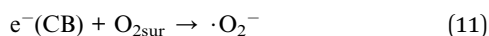
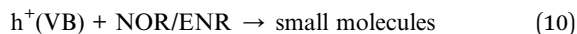
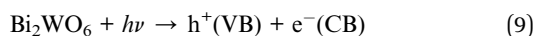


Fig. 10 (a) Results of recycling tests and (b) XRD of 450-Bi₂WO₆ after photocatalytic degradation of NOR for four cycles.

e^- and O_{2sur} could occur and cause a significant reduction in the production of $\cdot OH$. This was agree with the actual experiments. Based on the experimental results and the theory analyses above, the mechanism of photocatalytic degradation of NOR/ENR with 450-Bi₂WO₆ photocatalyst was presented in Scheme 1.

Therefore, the photocatalytic degradation of NOR or ENR with the as-prepared 450-Bi₂WO₆ catalyst was likely to occur through the reactions as followed:



3.6 Reusability and mineralization ability of 450-Bi₂WO₆

In order to study the reusability and stability of 450-Bi₂WO₆ photocatalyst, the stability test was examined. Four repetitive cycles of photodegradation experiments were carried out under the optimum conditions. The degradation efficiencies of NOR for the four successive cycles were 92.95%, 84.74%, 78.43% and 80.71%, respectively (Fig. 10(a)). The XRD pattern of the sample after been used for four cycles (Fig. 10(b)) was similar to that of the primary. It meant that the prepared 450-Bi₂WO₆ photocatalyst was stable under several reactions and was suitable for practical application.

The mineralization ability of the 450-Bi₂WO₆ was also investigated by measuring the decrease of TOC during the photocatalytic processes. As shown in Fig. S6,† under the optimal degradation conditions, the TOC removal efficiencies of ENR and NOR solutions could reach 42.17% and 53.24% after 75 min respectively, indicating that the 450-Bi₂WO₆ photocatalyst could mineralize the ENR and NOR molecules.

4 Conclusion

The Bi₂WO₆ photocatalysts were successfully synthesized by the combined method of ultrasonic solvothermal treatment and high-temperature calcination, and the calcination played a crucial role in enhancing the crystallinity and photocatalytic



activity of Bi₂WO₆. The 450-Bi₂WO₆ photocatalyst had a favorable photocatalytic performance for the degradation of NOR and ENR. The rate constant for the degradation of NOR with 450-Bi₂WO₆ catalyst was 2.72 and 5.33 times higher than those of bulk-Bi₂WO₆ and P25-TiO₂, respectively (for ENR, the rate constant was 2.65 and 4.98 times higher than those of bulk-Bi₂WO₆ and P25-TiO₂, respectively). The degradation efficiencies of NOR and ENR could reach to 92.95% and 94.58%, respectively (with the initial concentration of 10 mg L⁻¹ in presence of 0.5 g L⁻¹ 450-Bi₂WO₆ at initial pH 3.0). The quench experiments demonstrated that h⁺ and ·O₂⁻ were the dominant active species for the photocatalytic degradation of NOR and a possible mechanism was proposed. In addition, 450-Bi₂WO₆ showed enough stability and maintained the activity for the photocatalytic degradation of NOR after four cycles. This work could provide practical methods for synthesis of excellent photocatalyst and removal of fluoroquinolones antibiotics in water.

Conflicts of interest

The authors declare no competing financial interest.

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References

- J. J. Wang, L. Tang, G. M. Zeng, Y. Y. Zhou, Y. C. Deng, C. Z. Fan, J. L. Gong and Y. N. Liu, *Trans. Nonferrous Met. Soc. China*, 2017, **27**, 1794–1803.
- B. J. Dang, D. Q. Mao, Y. Xu and Y. Luo, *Water Res.*, 2017, **111**, 81–91.
- V. N. Binh, N. Dang, N. T. K. Anh, L. X. Ky and P. K. Thai, *Chemosphere*, 2018, **197**, 438–450.
- I. T. Carvalho and L. Santos, *Environ. Int.*, 2016, **94**, 736–757.
- H. Y. Magee, M. M. Maurer, A. Cobos, B. F. G. Pycke, A. K. Venkatesan, D. Magee, M. Scotch and R. U. Halden, *Sci. Total Environ.*, 2018, **643**, 460–467.
- K. Y. L. Chua, A. Bustamante, P. Jelfs, S. C. A. Chen and V. Sintchenko, *Pathology*, 2015, **47**, 678–682.
- K. H. Wammer, A. R. Korte, R. A. Lundeen, J. E. Sundberg, K. McNeill and W. A. Arnold, *Water Res.*, 2013, **47**, 439–448.
- K. Kümmerer, *Chemosphere*, 2009, **75**, 417–434.
- K. Kümmerer, *Chemosphere*, 2009, **75**, 435–441.
- R. T. Pei, J. M. Cha, K. H. Carlson and A. Pruden, *Environ. Sci. Technol.*, 2007, **41**, 5108–5113.
- T. Zhang, M. Zhang, X. X. Zhang and H. H. Fang, *Environ. Sci. Technol.*, 2009, **43**, 3455–3460.
- J. J. Macauley, Z. M. Qiang, C. D. Adams, R. Surampalli and M. R. Mormile, *Water Res.*, 2006, **40**, 2017–2026.
- X. H. Liu, G. D. Zhang, Y. Liu, S. Y. Lu, P. Qin, X. C. Guo, B. Bi, L. Wang, B. D. Xi, F. C. Wu, W. L. Wang and T. T. Zhang, *Environ. Pollut.*, 2019, **246**, 163–173.
- L. Aristilde, A. Melis and G. Sposito, *Environ. Sci. Technol.*, 2010, **44**, 1444–1450.
- S. J. Jiao, S. R. Zheng, D. Q. Yin, L. H. Wang and L. Y. Chen, *Chemosphere*, 2008, **73**, 377–382.
- K. Kümmerer, *Chemosphere*, 2001, **45**, 957–969.
- S. Zhang, H. L. Song, X. L. Yang, H. Li and Y. W. Wang, *Bioresour. Technol.*, 2018, **256**, 224–231.
- F. Santos, A. P. Mucha, D. A. M. Alexandrino, C. M. R. Almeida and M. F. Carvalho, *J. Environ. Manage.*, 2019, **231**, 1145–1153.
- L. Wang, Z. M. Qiang, Y. G. Li and W. W. Ben, *J. Environ. Sci.*, 2017, **56**, 263–271.
- E. M. Golet, I. Xifra, H. Siegrist, A. C. Alder and W. Giger, *Environ. Sci. Technol.*, 2003, **37**, 3243–3249.
- R. H. Lindberg, U. Olofsson, P. Rendahl, M. I. Johansson, M. Tysklind and B. A. V. Andersson, *Environ. Sci. Technol.*, 2006, **40**, 1042–1048.
- Y. F. Wang, J. X. Zhu, H. O. Huang and H. H. Cho, *J. Membr. Sci.*, 2015, **479**, 165–174.
- F. Tan, D. M. Sun, J. S. Gao, Q. Zhao, X. C. Wang, F. Teng, X. Quan and J. W. Chen, *J. Hazard. Mater.*, 2013, **244–245**, 750–757.
- Y. Li, J. F. Niu and W. L. Wang, *Chemosphere*, 2011, **85**, 892–897.
- M. Sturini, A. Speltini, F. Maraschi, A. Profumo, L. Pretali, E. Fasani and A. Albini, *Environ. Sci. Technol.*, 2010, **44**, 4564–4569.
- S. Y. Dong, J. L. Feng, M. H. Fan, Y. Q. Pi, L. M. Hu, X. Han, M. L. Liu, J. Y. Sun and J. H. Sun, *RSC Adv.*, 2015, **5**, 14610–14630.
- M. J. Chen and W. Chu, *Appl. Catal., B*, 2015, **168–169**, 175–182.
- D. H. Lan, H. T. Wang, L. Chen, C. T. Au and S. F. Yin, *Carbon*, 2016, **100**, 81–89.
- D. Liu, J. Wang, X. J. Bai, R. L. Zong and Y. F. Zhu, *Adv. Mater.*, 2016, **28**, 7284–7290.
- R. a. He, S. W. Cao, P. Zhou and J. G. Yu, *Chin. J. Catal.*, 2014, **35**, 989–1007.
- S. Rengaraj, X. Z. Li, P. A. Tanner, Z. F. Pan and G. K. H. Pang, *J. Mol. Catal. A: Chem.*, 2006, **247**, 36–43.
- S. Gong, Q. F. Han, J. W. Zhu, X. Wang and L. D. Lu, *Mater. Res. Bull.*, 2016, **76**, 222–228.
- J. L. Hu, W. J. Fan, W. Q. Ye, C. J. Huang and X. Qiu, *Appl. Catal., B*, 2014, **158–159**, 182–189.
- M. X. Du, Y. Du, Y. B. Feng, K. Yang, X. J. Lv, N. Jiang and Y. Liu, *Carbohydr. Polym.*, 2018, **195**, 393–400.
- H. B. Li, Z. J. Yang, J. N. Zhang, Y. C. Huang, H. B. Ji and Y. X. Tong, *Appl. Surf. Sci.*, 2017, **423**, 1188–1197.
- C. Y. Wang, X. Zhang, X. N. Song, W. K. Wang and H. Q. Yu, *ACS Appl. Mater. Interfaces*, 2016, **8**, 5320–5326.
- A. Kaur and S. K. Kansal, *Chem. Eng. J.*, 2016, **302**, 194–203.
- Y. K. Huang, S. F. Kang, Y. Yang, H. F. Qin, Z. J. Ni, S. J. Yang and X. Li, *Appl. Catal., B*, 2016, **196**, 89–99.



- 39 J. J. Wang, L. Tang, G. M. Zeng, Y. C. Deng, Y. N. Liu, L. L. Wang, Y. Y. Zhou, Z. Guo, J. J. Wang and C. Zhang, *Appl. Catal., B*, 2017, **209**, 285–294.
- 40 D. Cao, Y. B. Wang, M. Qiao and X. Zhao, *J. Catal.*, 2018, **360**, 240–249.
- 41 S. Jonjana, A. Phuruangrat, T. Thongtem and S. Thongtem, *Mater. Lett.*, 2016, **172**, 11–14.
- 42 A. Kudo and S. Hijii, *Chem. Lett.*, 1999, **1999**, 1103–1104.
- 43 J. W. Tang, Z. G. Zou and J. H. Ye, *Catal. Lett.*, 2004, **92**, 53–56.
- 44 L. Y. Liang, Y. Tursun, A. Nulahong, T. Dilinuer, A. Tunishaguli, G. Gao, A. Abulikemu and K. Okitsu, *Ultrason. Sonochem.*, 2017, **39**, 93–100.
- 45 S. W. Zhu, C. Y. Yang, F. Li, T. H. Li, M. Zhang and W. Cao, *Mol. Catal.*, 2017, **435**, 33–48.
- 46 G. Zhao, S. W. Liu, Q. F. Lu and L. J. Song, *Ind. Eng. Chem. Res.*, 2012, **51**, 10307–10312.
- 47 F. Amano, K. Nogami and B. Ohtani, *J. Phys. Chem. C*, 2009, **113**, 1536–1542.
- 48 Y. Y. Zhao, Y. B. Wang, E. Z. Liu, J. Fan and X. Y. Hu, *Appl. Surf. Sci.*, 2018, **436**, 854–864.
- 49 M. Shang, W. Z. Wang, S. M. Sun, L. Zhou and L. Zhang, *J. Phys. Chem. C*, 2008, **112**, 10407–10411.
- 50 J. Y. Sheng, X. J. Li and Y. M. Xu, *Acta Phys.-Chim. Sin.*, 2014, **30**, 508–512.
- 51 X. L. Song, H. D. Wang, Y. Y. Li, S. Y. Ye and D. D. Dionysiou, *Appl. Surf. Sci.*, 2018, **439**, 815–822.
- 52 Z. J. Zhang, W. Z. Wang, L. Wang and S. M. Sun, *ACS Appl. Mater. Interfaces*, 2012, **4**, 593–597.
- 53 Y. Tian, L. D. Zhang and J. X. Zhang, *J. Alloys Compd.*, 2012, **537**, 24–28.
- 54 Y. Y. Zhao, X. H. Liang, Y. B. Wang, H. X. Shi, E. Z. Liu, J. Fan and X. Y. Hu, *J. Colloid Interface Sci.*, 2018, **523**, 7–17.
- 55 F. Chen, D. Li, B. F. Luo, M. Chen and W. D. Shi, *J. Alloys Compd.*, 2017, **694**, 193–200.
- 56 J. G. Yu, J. F. Xiong, B. Cheng, Y. Yu and J. B. Wang, *J. Solid State Chem.*, 2005, **178**, 1968–1972.
- 57 R. W. Matthews, M. Abdullah and K. C. Low, *Anal. Chim. Acta*, 1990, **233**, 171–179.
- 58 J. Gamage McEvoy, W. Cui and Z. Zhang, *Appl. Catal., B*, 2014, **144**, 702–712.
- 59 W. F. Yao, H. Wang, X. H. Xu, J. T. Zhou, X. N. Yang, Y. Zhang and S. X. Shang, *Appl. Catal., A*, 2004, **259**, 29–33.
- 60 J. Grzechulska and A. W. Morawski, *Appl. Catal., B*, 2002, **36**, 45–51.
- 61 H. J. Zhang, P. Zhang, Y. Ji, J. Tian and Z. X. Du, *Chem. Eng. J.*, 2015, **262**, 1108–1115.
- 62 O. Moumeni, O. Hamdaoui and C. Pétrier, *Chem. Eng. Process.*, 2012, **62**, 47–53.
- 63 K. Bourikas, C. Kordulis and A. Lycourghiotis, *Environ. Sci. Technol.*, 2005, **39**, 4100–4108.
- 64 H. H. Wang, J. Lu, F. Q. Wang, W. H. Wei, Y. Chang and S. J. Dong, *Ceram. Int.*, 2014, **40**, 9077–9086.
- 65 G. J. Leigh, *J. Organomet. Chem.*, 1983, **251**, c37.
- 66 A. Akhundi and A. Habibi-Yangjeh, *Mater. Chem. Phys.*, 2016, **174**, 59–69.
- 67 B. Wang, H. Yang, T. Xian, L. J. Di, R. S. Li and X. X. Wang, *J. Nanomater.*, 2015, **16**, 194–200.
- 68 R. M. Fernández-Domene, R. Sánchez-Tovar, B. Lucasgranados, M. J. Muñoz-Portero and J. García-Antón, *Chem. Eng. J.*, 2018, **350**, 1114–1124.
- 69 D. Ma, J. Wu, M. C. Gao, Y. J. Xin, Y. Y. Sun and T. J. Ma, *Chem. Eng. J.*, 2017, **313**, 1567–1576.

