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Nature of intramolecular $O-H\cdots\pi$ interactions as elucidated by QTAIM dual functional analysis with QC calculations†

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The intrinsic dynamic and static nature of intramolecular $OH^{-*}-\pi$ interactions is elucidated using a QTAIM dual functional analysis (QTAIM-DFA) after clarifying the structural features. Asterisks (*) are employed to emphasize the presence of bond critical points (BCPs) on the bond paths (BPs), which correspond to the interactions in question. Data from the fully optimized structures correspond to the static nature of the interactions. In our treatment, data from the perturbed structures, which are based around the fully optimized structure, are employed for the analysis in addition to those from the fully optimized structure, which represent the dynamic nature of the interaction. Seven intramolecular $OH^{-*}-C(\pi)$ interactions were detected in six-membered rings, with six BPs and BCPs for each, among the 72 conformers of the species examined here (1–15). The interactions are predicted to have a vdW or t-HB_{nc} (typical hydrogen bonds with no covalency) nature, which appeared in the pure closed shell region. They appear to be stronger than the corresponding intermolecular interactions. Nine BPs with BCPs were also detected for the intramolecular $O^{-*}-X$ interactions ($X = C(\pi)$ and $H(\pi)$, joined to $C(\pi)$) in the 5–7-membered rings. The E(2) values of the interactions, as obtained by NBO, are discussed in relation to the stabilities of the conformers and the BPs with BCPs.

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Introduction

Hydrogen bonds (HBs) are of ongoing interest in all fields of chemical and biological sciences.1-4 The conventional HBs in the shared proton interaction type (cv-HBs: B···H-X) are basic HBs. The B···H-X directions are controlled through the formation of HBs from X-H and B due to the contribution of the unsymmetric $\sigma(3c-4e)$ (three centre-four electron interactions of the σ -type).⁵⁻⁷ The energies involved in the formation of cv-HBs are typically 10-40 kJ mol⁻¹ for the neutral form.⁵ Another type of HB will form if π -orbitals are provided from ethyne, ethene, benzene and the derivatives to X-H. These X- $H \cdots \pi$ interactions, which are called π -HBs here, seem weaker than cv-HBs. The weaker proton-accepting ability of π -orbitals relative to the lone pair orbitals must primarily be responsible for the differences. We reported the behaviour of cv-HBs among the neutral and charged forms very recently by applying the QTAIM (quantum theory of atoms-in-molecules) approach. 4,8 The cv-HBs of the neutral form are predicted to have a vdW to

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CT-TBP (trigonal bipyramidal adduct formation through charge transfer) nature, while the cv-HBs of the charged form show a covalent bond nature (Cov).⁴ The natures of the XH··· π interactions were also reported recently for the π -systems of benzene,^{9,10} naphthalene,¹¹ anthracene¹² and/or coronene,¹³ where X = F, Cl, Br, I, HO, HS, HSe, MeO, H₂N, MeHN and/or Me₂N.

We also closely observed the intramolecular π -HBs since they play a very important role in the chemical and biological sciences. Hhat is the behaviour of the intramolecular π -HBs? What are the differences and similarities between the intramolecular and intermolecular π -HBs? How does steric hindrance affect the strength of the intermolecular π -HBs? It is challenging to clarify the nature of intramolecular π -HBs to understand the fundamental behaviour of π -HBs. Chart 1 illustrates species 1–15, which were examined in this study.

Chart 1 Candidates 1–15, to examine the intramolecular HB interactions.

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The bond critical point (BCP, *) is an important concept in the QTAIM approach that was introduced by Bader, 16,17 where the $\rho(r)$ (charge density) reaches its minimum along the interatomic (bond) path while reaching its maximum on the interatomic surface separating the atomic basins. The $\rho(\mathbf{r})$ at BCP is described by $\rho_b(\mathbf{r}_c)$, as are other QTAIM functions, such as the total electron energy densities $H_{\rm b}(r_{\rm c})$, potential energy densities $V_{\rm b}(\mathbf{r}_{\rm c})$ and kinetic energy densities $G_{\rm b}(\mathbf{r}_{\rm c})$ at BCPs. A chemical bond or interaction between atoms A and B is denoted by A-B, which corresponds to the bond path (BP) in QTAIM. We will use A-*-B for BP, in which the asterisk emphasizes the existence of a BCP in A-B.16,17 Eqn (1), (2) and (2') represent the relations between $G_b(\mathbf{r}_c)$, $V_b(\mathbf{r}_c)$, $H_b(\mathbf{r}_c)$ and $\nabla^2 \rho_b(\mathbf{r}_c)$.

$$H_{b}(\mathbf{r}_{c}) = G_{b}(\mathbf{r}_{c}) + V_{b}(\mathbf{r}_{c}) \tag{1}$$

$$(\hbar^2/8m)\nabla^2\rho_b(\mathbf{r}_c) = H_b(\mathbf{r}_c) - V_b(\mathbf{r}_c)/2$$
 (2)

$$(\hbar^2/8m)\nabla^2\rho_b(\mathbf{r}_c) = G_b(\mathbf{r}_c) + V_b(\mathbf{r}_c)/2$$
 (2')

A QTAIM dual functional analysis (QTAIM-DFA) was recently formulated based on the QTAIM approach. $H_{\rm b}(r_{\rm c})$ is plotted versus $H_b(\mathbf{r}_c) - V_b(\mathbf{r}_c)/2$ (= $(\hbar^2/8m)\nabla^2 \rho_b(\mathbf{r}_c)$) (cf.: eqn (2)) at BCPs in QTAIM-DFA. Data from the fully optimized structures are analysed using the polar coordinate (R, θ) representation, which corresponds to the static natures of the interactions. 18-21 In our treatment, data from the perturbed structures around the fully optimized ones are employed in addition to those from the fully optimized structures. Each interaction plot, which contains data from both the perturbed and fully optimized structures, includes a specific curve that provides important information about the interaction. This plot is expressed by (θ_p, κ_p) , where θ_p corresponds to the tangent line of the plot and κ_p is the curvature. The dynamic nature of the interactions has been proposed based on (θ_p, κ_p) . ^{18 *a*-21} We call (R, θ) and (θ_p, κ_p) the QTAIM-DFA parameters, which are illustrated in Fig. 4, as exemplified by the intramolecular OH-*-C(π) interaction in **12a** (C_1).

We very recently proposed a highly reliable method for generating the perturbed structures necessary for QTAIM-DFA.²² The method is called CIV, and it employs the coordinates derived from the compliance force constants C_{ii} , the diagonal elements in C_{ij} , for the internal vibrations. Eqn (3) defines C_{ij} as the partial second derivatives of the potential energy due to an external force, where i and j refer to internal coordinates, and the force constants f_i and f_i correspond to i and j, respectively. The dynamic nature of interactions based on perturbed structures with CIV is described as the "intrinsic dynamic nature of interactions," since the coordinates are invariant relative to the choice of the coordinate system.

$$C_{ij} = \partial^2 E / \partial f_i \partial f_j \tag{3}$$

QTAIM-DFA is applied to standard interactions, and rough criteria that distinguish the interaction in question from others are obtained. QTAIM-DFA has excellent potential for evaluating, classifying, characterizing and understanding weak to strong interactions according to a unified form. 18a-21 The QTAIM-DFA and the criteria are explained in the ESI using Schemes S1-S3, Fig. S1 and S2, Table S1 and eqn (S1)-(S7).† The basic concept of the OTAIM approach is also discussed.

When employing the perturbed structures generated with the CIV, we consider QTAIM-DFA to be well-suited for elucidating the nature of the intramolecular π -HB interactions in 1-15. Herein, we present the results of investigations on the intrinsic dynamic and static nature of the intramolecular π -HBs. BPs with BCPs corresponding to the intramolecular $OH^{-*}-\pi(C)$ interactions are detected in seven conformers, together with four intramolecular O^{-*} - $C(\pi)$ and five O^{-} *-H(π) interactions among 72 conformers examined in 1-15, where $H(\pi)$ indicates an H joined directly to a $C(\pi)$. The nature of the intramolecular interactions is clarified for those detected by the BPs with BCPs. The intramolecular interactions are classified and characterized by employing the criteria as a reference. An NBO analysis23 is applied to some selected conformers of 1-15. The nature of the intramolecular interactions will be discussed in relation to the results of the NBO analysis and the structural features.

Methodological details in calculations

The Gaussian 09 programs²⁴ were employed for the calculations. The calculations containing the NBO analysis²³ were performed with the 6-311++G(3df,3pd) basis set (BSS-A) at a second-order Møller-Plesset energy correlation (MP2) level (MP2/BSS-A).25 The optimized structures were confirmed by the frequency analysis. The results of the frequency analysis were used to obtain the C_{ii} values and the coordinates corresponding to C_{ii} . ^{26–29} The B3LYP³⁰/BSS-A and M06-2X³¹/BSS-A methods were also applied to the limited cases for the examination of the effect from the DFT level. The optimizations were not corrected with the BSSE method.

Eqn (4) explains the method used to generate the perturbed structures with CIV. A k-th perturbed structure in question (S_{iw}) is generated by the addition of the coordinates (Ci) corresponding to C_{ii} to the standard orientation of a fully optimized structure (S_0) in the matrix representation. The coefficient g_{kw} in eqn (4) controls the structural difference between S_{iw} and S_{o} : g_{kw} is determined to satisfy eqn (5) for r, the interaction distance in the perturbed structure. 32 The C_i values of five digits are used to predict the S_{iw} .

$$S_{iw} = S_0 + g_{iw}C_i \tag{4}$$

$$r = r_0 + wa_0$$
 ($w = (0), \pm 0.025$ and ± 0.05 ; $a_0 = 0.52918$ Å) (5)

$$y = c_0 + c_1 x + c_2 x^2 + c_3 x^3 (R_c^2)$$
: square of the correlation coefficient) (6)

The QTAIM functions were calculated using the same method as the optimizations at the MP2 level, unless otherwise noted. The calculated values were analysed with the $AIM2000^{33}$ and AIMAll³⁴ programs. $H_b(r_c)$ is plotted versus $H_b(r_c) - V_b(r_c)/2$ for the data from five points pertaining to w = 0, ± 0.05 and ± 0.1 in eqn (5) in QTAIM-DFA. Each plot is analysed using

a regression curve of the cubic function, as shown in eqn (6), where $(x, y) = (H_b(r_c) - V_b(r_c)/2, H_b(r_c)) (R_c^2 > 0.99999)$ in usual).³⁵

Results and discussion

Optimizations of species, 1-15

Species 1-15 were optimized with MP2/6-311++G(3df,3pd) (MP2/BSS-A). The most extended conformer was optimized for each of 1-15, first. The conformers were searched by the optimizations with the changing all torsional angles for each species. 72 conformers were optimized for 1-15. The conformers are explained first, as exemplified in 3. Fig. 1 illustrates the conformers of 3 with $\phi(C_{\beta}C_{\gamma}OH) \approx 180^{\circ}$ (t), which are the ttt, tgt, gtt, ggt and gg't conformers around C_{β} , C_{γ} and O. Three conformers around C_{α} in 3 are identical, as shown in Fig. 1. Therefore, the conformers will be distinguished by the conformations around C_{β} , C_{γ} and O. The gauche (g) and gauche' (g') notations are used for $\phi(C_BC_YOH) \approx 60^\circ$ and -60° , respectively, for example, in addition to the trans (t) notation for $\phi(C_{\beta}C_{\gamma}OH) \approx 180^{\circ}$. The g and g' conformers around C_{β} (from the ttt conformer) are the same in this study. The optimizations were further performed with the torsional angles changing compared with those of the optimized structures, as mentioned above. Fourteen conformers were optimized for 3, although the systematic conformation analysis is not applied. The optimized conformer of the shortest $OH \cdots C(\pi)$ distance $[r(H \cdots C(\pi))]$ in 3 is called 3a. The optimized conformers will be called 3b, 3c, ..., **3m** and **3n**, in the increasing order of the optimized $OH \cdots C(\pi)$ distances.

In the case of 5, the most extended structure of the $C_{\rm s}$ symmetry (5 $(C_{\rm s})$) has one imaginary frequency. The optimization converged a conformer of the $C_{\rm 1}$ symmetry (5A $(C_{\rm 1})$) with the torsional angle of $\phi(C_{\rm sp2}C_{\rm sp2}C_{\rm H2}H)\approx 8.5^{\circ}$ if started from

(a) (b) (c) (c) $g: \phi(C_{p}C_{p}OH) \approx 60^{\circ}$ $g: \phi(C_{p}C_{p}OH) \approx -60^{\circ}$ $f: \phi(C_{p}C_{p}OH) \approx -60^{\circ}$ $g: \phi(C_{p}C_{p}OH) \approx -60^{\circ}$ $f: \phi(C_{p}C_{p$

Fig. 1 Survey of the structural optimizations for $\bf 3$ and $\bf 5$, with MP2/BSS-A.

the C_1 structure, close to 5 (C_s). However, the similar conformer with $\phi(C_{sp2}C_{sp2}C_{H2}H) \approx 173.5^{\circ}$ (5A' (C_1)) was not optimized. Another type of conformer with C_s symmetry (5B (C_s)) was optimized, of which $\phi(C_{sp2}C_{sp2}C_{H2}C_{H2}) = 0$. Fig. 1 contains the process from 5 (C_s) to 5A (C_1) and 5B (C_s) with 5A' (C_1). The optimizations for 5 were performed by changing the torsional angles around the $-CH_2OH$ group in 5A (C_1) and 5B (C_s). As a result, fifteen different conformers were optimized for 5. Conformers 5 (C_s), 5A (C_1) and 5B (C_s) correspond to 5n, 5m and 5o, respectively, among the 15 conformers (see Fig. S4 of the ESI†).

The optimizations of 1–15 other than 3 and 5 were performed in a similar way. There were greater efforts to search for the conformers with the shorter $OH\cdots C(\pi)$ distances than the ones with the longer distances, which would prevent the trivial optimizations of the conformers with no intramolecular $OH\cdots C(\pi)$ interactions. Finally, 72 conformers were optimized for 1–15. The optimized conformers in this study are denoted as xa, xb,... (x=1-15), similar to 3. The selected structural parameters around the intramolecular $OH\cdots C(\pi)$ interactions in 1–15, r(O-H), $r(H\cdots C(\pi))$, $\angle OHC(\pi)$ and $\angle HC(\pi)C(\pi)$, are collected in Table S2 of the ESI.† The optimized structures of 1–15 are not shown in the figures, but they can be found in the molecular graphs drawn on the optimized structures (see Fig. S3–S5 of the ESI† and Fig. 3).

The relative energies (ΔE) in **1–15** are calculated on the energy surface ($\Delta E_{\rm ES}$) and those with the corrections for the zero-point energies ($\Delta E_{\rm ZP}$). The energies of the conformers in **1–15**, with the smallest $r({\rm H}\cdots{\rm C}(\pi))$ values among the optimized ones (called **1a–15a**, respectively), were chosen as the standards. The ΔE values ($\Delta E_{\rm ES}$ and $\Delta E_{\rm ZP}$) for **1–15** are also presented in Table S2 of the ESI.†

Fig. 2 shows the plot of $\Delta E_{\rm ZP}$ versus $\Delta E_{\rm ES}$ for the optimized 14 conformers in 3. The plot showed an excellent correlation (y=

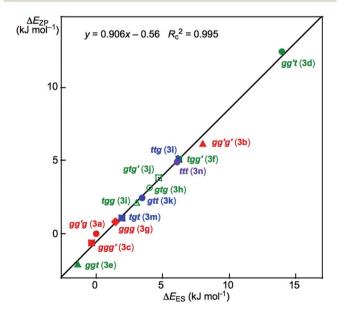


Fig. 2 Plot of $\Delta E_{\rm ZP}$ versus $\Delta E_{\rm ES}$ for the conformers in 3, as optimized with MP2/BSS-A.

0.906x - 0.56: $R_c^2 = 0.995$). The dependence of ΔE on the conformers in 3 seems well understood based on Fig. 2. The ΔE values for the conformers increase in the order shown in eqn (7) (less stable in the order), if $\phi(C_\beta C_\gamma OH)$ are limited to 180° (t). The ttt conformer is predicted to be the second-most unstable one in the stability sequence shown in eqn (7).

$$ggt$$
 (3e) $\leq tgt$ (3m) $\leq gtt$ (3k) $\leq ttt$ (3n) $\leq gg't$ (3d) (7)

The prediction seems unusual at first glance since the conformers in 3 are expected to be less stable as the steric crowding increases if no mechanisms other than the steric one are operating to stabilize the conformer. The repulsive energy from the steric hindrance is expected to be the lowest in ttt (3n). However, ggt (3e), tgt (3m) and gtt (3k) are predicted to be more stable than ttt (3n), although gg't (3d) is predicted to be most unstable relative to the others. The ΔE_{ES} values are calculated to be 14.0 and 8.0 kJ mol^{-1} for gg't (3d) and gg'g' (3b), respectively, which are the most and secondmost unstable conformers in 3. They seem to be the second and third-most sterically crowded ones in 3, respectively. The conformer 3a is expected to be the most sterically crowded one among those in 3, where the ΔE_{ES} value for gg'g (3a) is used as the standard (0.0 kJ mol⁻¹) for the conformers of 3. However, gg'g (3a) is close to the most stabilized conformer among those in 3. The intramolecular OH-*-C(π) interaction contributes to stabilizing gg'g (3a) by approximately 15 kJ mol⁻¹ in 3. Similar phenomena were observed among the optimized conformers in 1-15.

The $\Delta E_{\rm ZP}$ values are similarly plotted *versus* the $\Delta E_{\rm ES}$ for the 72 conformers. The plot is shown in Fig. S6 of the ESI.† The plot also gave a very good correlation (y = 0.897x - 0.20: $R_{\rm c}^2 = 0.990$). As a result, $\Delta E_{\rm ES}$ can be employed for the discussion of ΔE .

Before presenting a detailed discussion of the nature of the intramolecular $OH^{-*}-C(\pi)$ interactions, it is instructive to examine the molecular graphs with the contour plots.

Molecular graphs with contour plots for 1-15

A BCP on a BP corresponding to the intramolecular OH*-C(π) interaction was detected for **3a**, **6a**, **9a**, **11a**, **12a**, **13a** and **15a**, whereas one corresponding to the intramolecular O-*-C(π) interaction was recorded for **3b**, **6c**, **9b** and **14b**, and that corresponding to the intramolecular O-*-H(π) interaction was for **5e**, **5i**, **12b**, **15b** and **15c**, where the H(π) joined directly to the C(π). Fig. 3 illustrates the molecular graphs with the contour plots for the intramolecular OH*-C(π), O-*-C(π) and O-*-H(π) interactions for the conformers discussed above. All the expected BCPs are clearly detected, and they contain the components for the intramolecular interactions. These BCPs seem to be appropriately located at the (three-dimensional) saddle points of $\rho(r)$.

As shown in Fig. 3, each BP with BCP corresponding to the intramolecular OH–*– $C(\pi)$ interaction appears in the six-membered ring of the –COH–*–CCC type. However, each BP with BCP corresponding to the intramolecular O–*– $C(\pi)$ or

 O^{-*} -H(π) interaction appears in the five, six, or sevenmembered ring. A BCP on a BP corresponding to the intramolecular OH-*-C(π), O-*-C(π) or O-*-H(π) interaction is not detected for HOCC≡CH (1), HOCCC≡CH (2), HOCC= CH (4), HOCCC=CH (5) and HOCH₂Ph (7), HOCH₂CH₂Ph (8) and $HOC_6H_4C=CH-o$ (10). The cyclic interaction seems not to satisfy the conditions for the appearance of BP with BCP in each of the above species. BPs with BCPs corresponding to both intramolecular OH-*-C(π) and O-*-C(π) interactions are detected in 6, while BPs with BCPs corresponding to the intramolecular OH-*-C(π) and O-*-H(π) interactions are detected in 12 and 15. However, only BP with BCP corresponding to the intramolecular $OH^{-*}-C(\pi)$ interaction appears in 13, whereas only BP with BCP corresponding to the intramolecular $O^{-*}-C(\pi)$ interaction appears in 14. The intramolecular OH-*-C(π) interactions, appearing in the sixmembered rings, show clear contrast to the similar CH-*-C(π) interaction, expected to occur in the five-membered ring in the species similar to 5.36

Molecular graphs are given in Fig. S3–S5 of the ESI† for the conformers in **1–15** without a BCP on a BP corresponding to the OH–*-C(π) or O–*-X interaction being recorded for each.

Survey of HB interactions in 1-15

The BPs corresponding to the intramolecular $OH^{-*}-C(\pi)$, O-*-C(π) and O-*-H(π) interactions shown in Fig. 3 appear somewhat curved, especially around the area close to the atoms at the ends of the BPs. To examine the linearity of the interactions further, the lengths of the BPs (r_{BP}) in question and the corresponding straight-line distances (R_{SL}) are calculated for those shown in Fig. 3. The values calculated with MP2/BSS-A are collected in Table S3 of the ESI,† together with the differences between them ($\Delta r_{\rm BP} = r_{\rm BP} - R_{\rm SL}$). The magnitudes of $\Delta r_{\rm BP}$ are 0.01-0.30 Å for the BPs. Consequently, the intramolecular OH-*-C(π) interactions shown in Fig. 3 should be recognized as the curved ones, more or less. The curved nature of the intramolecular interactions would have originated from the twisted interaction due to the steric constraints in the optimized conformers. The r_{BP} values are plotted *versus* the R_{SL} , which is displayed in Fig. S7 of the ESI.† The $\Delta r_{\rm BP}$ values seem to increase in the order $O^{-*}-H(\pi) < O^{-*}-C(\pi) < OH^{-*}-C(\pi)$ in the

The QTAIM functions of $\rho_{\rm b}(r_{\rm c})$, $H_{\rm b}(r_{\rm c})$ – $V_{\rm b}(r_{\rm c})/2$, $H_{\rm b}(r_{\rm c})$ and $k_{\rm b}(r_{\rm c})$ are calculated at the BCP on the BP corresponding to the intramolecular OH–*–C(π) interaction for 3a, 6a, 9a, 11a, 12a, 13a and 15a and at the factor corresponding to the intramolecular O–*–C(π) interaction for 3b, 6b, 9b and 14b, together with that corresponding to the intramolecular O–*–H(π) interaction for 5e, 5i, 12b, 15b and 15c. Table 1 shows the values, as evaluated with MP2/BSS-A. Fig. 4 shows the plots of $H_{\rm b}(r_{\rm c})$ versus $H_{\rm b}(r_{\rm c})$ – $V_{\rm b}(r_{\rm c})/2$ for the data in Table 1 and those from the perturbed structures generated with CIV. The (R, θ) and $(\theta_{\rm p:CIV}, \kappa_{\rm p:CIV})$ values were calculated by analysing the plots in Fig. 4 according to eqn (S1)–(S6) of the ESI.† The values are collected in Table 1, together with the $C_{\rm ii}$ values corresponding to CIV employed to generate the perturbed structures.

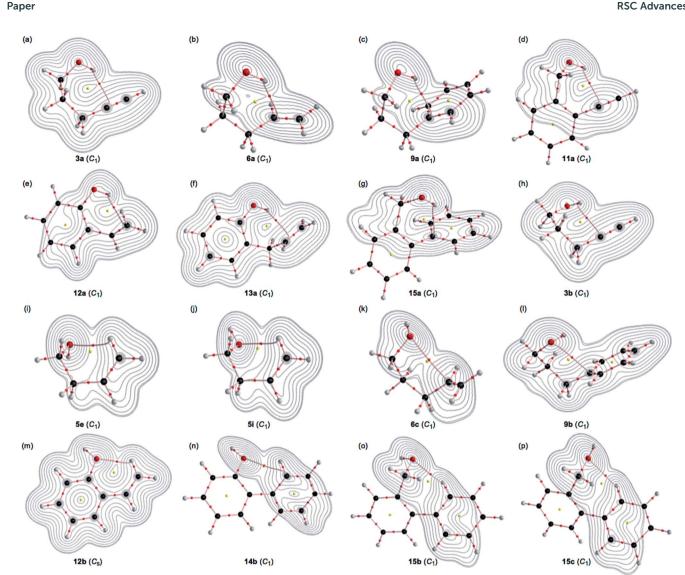


Fig. 3 Molecular graphs for the 16 conformers of 3a, 6a, 9a, 11a, 12a, 13a, 15a, 3b, 5e, 5i, 6c, 9b, 12b, 14b, 15b and 15c calculated with MP2/BSS-A (shown by (a)-(p), respectively, in the figure), where BPs with BCP corresponding to the intramolecular non-covalent interactions around the OH group are detected. The BCPs are denoted by red dots, RCPs (ring critical points) are indicated by yellow dots and BPs are indicated by pink lines. The carbon, hydrogen and oxygen atoms are shown in black, grey and red, respectively. Contour plots are drawn on the planes containing the intramolecular interaction for each. The contours (ea_0^{-3}) are at 2^l ($l = \pm 8, \pm 7 \dots$ and 0).

Nature of intramolecular OH-*-C(π), O-*-C(π) and O-*-H(π) interactions in the conformers of 1-15

The intramolecular OH-*-C(π), O-*-C(π) and O-*-H(π) interactions in the conformers of 1-15 given in Table 1 are classified and characterized based on the $(\theta, \theta_{p:CIV})$ values and evaluated with MP2/BSS-A. While θ classifies interactions, θ_p characterizes them. It is instructive to survey the criteria shown in Scheme S3 and Table S1 of the ESI† before engaging in detail discussion. The criteria indicate that $45^{\circ} < \theta < 180^{\circ} (0 < H_b(r_c) - V_b(r_c)/2)$ for the CS interactions and $180^{\circ} < \theta < 206.6^{\circ} (H_b(r_c) - V_b(r_c)/2 < 0)$ for the SS interactions. 18a,32 The CS interactions are sub-divided into $45^{\circ} < \theta < 90^{\circ} (H_b(r_c) > 0)$ for the pure CS interactions (p-CS) and $90^{\circ} < \theta < 180^{\circ} (H_{\rm b}(r_{\rm c}) < 0)$ for the regular CS interactions (r-CS). In the p-CS region of $45^{\circ} < \theta < 90^{\circ}$, the character of the interactions will be the vdW type for $45^{\circ} < \theta_{\rm p} < 90^{\circ} (45^{\circ} < \theta <$

75°), whereas it will be the typical hydrogen bonds type with no covalency (t-HB_{nc}) for $90^{\circ} < \theta_{\rm p} < 125^{\circ} (75^{\circ} < \theta < 90^{\circ})$, where $\theta =$ 75° and $\theta_{\rm p}=125^\circ$ are tentatively given for $\theta_{\rm p}=90^\circ$ and $\theta=90^\circ$, respectively. The CT interaction will appear in the r-CS region of $90^{\circ} < \theta < 180^{\circ}$. The t-HB interactions with the covalency (t-HB_{wc}) appear over the range $125^{\circ} < \theta_{p} < 150^{\circ} (90^{\circ} < \theta < 115^{\circ})$, where $(\theta_{p} < 150^{\circ})$ $\theta_{\rm p}$) = (115°, 150°) are tentatively given as the borderline between the nature of t-HB_{wc} and CT-MC (molecular complex formation through CT). The borderline of the interactions between CT-MC and CT-TBP (TBP adduct formation through CT) types is defined by $(\theta, \theta_p) = (150^\circ, 180^\circ)$, where $\theta = 150^\circ$ is tentatively given as corresponding to $\theta_p = 180^\circ$. As a result, the (θ, θ_p) values of (75°, 90°), (90°, 125°), (115°, 150°), (150°, 180°) and (180°, 190°) correspond to the borderlines between the nature of interactions for vdW/t-HB_{nc}, t-HB_{nc}/t-HB_{wc}, t-HB_{wc}/CT-MC, CT-

Table 1 QTAIM functions and QTAIM-DFA parameters for the intramolecular interactions around the O-H group, as elucidated with MP2/BSS-A and predicted using the predicted nature for the interactions^{a,b}

A^{-*} –B(π), Compound	$\rho_{\rm b}(r_{\rm c})~(ea_{ m o}^{-3})$	$c\nabla^2 \rho_{\rm b}(\boldsymbol{r}_{\rm c})^c \ ({\rm au})$	$H_{\rm b}(\boldsymbol{r}_{\rm c})$ (au)	$k_{\rm b}(\pmb{r}_{\rm c})^d$	R^e (au)	θ^f (°)	C_{ii}^{g} (Å mdyn ⁻¹)	$\theta_{\mathrm{p:CIV}}^{ h}\left(^{\circ}\right)$	$\kappa_{\text{p:CIV}}^{i} \left(\text{au}^{-1} \right)$	Predicted nature
OH-*-C(π)										
OH-*- α C(π) in 3a	0.0161	0.0063	0.0015	-0.870	0.0065	77.0	7.81	97.7	335	p-CS/t-HB _{nc}
$OH^{-*}-^{\alpha}C(\pi)$ in 6a	0.0162	0.0059	0.0010	-0.909	0.0059	80.5	8.02	99.2	387	p-CS/t-HB _{nc}
OH-*- ${}^{i}C(\pi)$ in 9a	0.0155	0.0061	0.0011	-0.898	0.0062	79.5	7.46	101.8	288	p-CS/t-HB _{nc}
OH-*- $^{\alpha}$ C(π) in 11a	0.0117	0.0048	0.0011	-0.867	0.0050	76.8	14.93	79.3	56.2	p-CS/vdW
OH-*- $^{\beta}$ C(π) in 12a	0.0170	0.0072	0.0014	-0.888	0.0073	78.6	5.85	101.8	236	p-CS/t-HB _{nc}
OH-*- $^{\alpha}$ C(π) in 13a	0.0178	0.0064	0.0011	-0.907	0.0065	80.3	7.58	115.9	420	p-CS/t-HB _{nc}
OH-* $^{-i}$ C(π) in 15a	0.0135	0.0053	0.0009	-0.901	0.0054	79.8	12.45	94.0	193	p -CS/t-HB $_{nc}$
Ο-*-C(π)										
$O^{-*}-\alpha C(\pi)$ in $3b^{j}$	0.0125	0.0054	0.0011	-0.887	0.0055	78.5	7.31	49.7	5035	p-CS/vdW
$O^{-*}-\alpha C(\pi)$ in 6c	0.0098	0.0047	0.0011	-0.863	0.0049	76.4	7.48	107.2	7373	p-CS/t-HB _{nc}
O^{-*} -iC(π) in 9b	0.0108	0.0047	0.0009	-0.897	0.0048	79.4	9.14	90.8	85.0	p-CS/t-HB _{nc}
$O^{-*}-^{o}C(\pi)$ in 14b ^k	0.0108	0.0060	0.0019	-0.807	0.0063	72.0	4.79	86.1	1631	p-CS/vdW
O^{-*} - $H(\pi)^l$										
$O^{-*}-^{\beta}H(\pi)$ in 5e	0.0107	0.0051	0.0012	-0.867	0.0052	76.8	12.74	86.3	266	p-CS/vdW
$O^{-*-\beta}H(\pi)$ in 5i	0.0102	0.0049	0.0012	-0.862	0.0050	76.3	15.97	88.5	577	p-CS/vdW
$O^{-*-\beta}H(\pi)$ in 12b	0.0164	0.0086	0.0026	-0.822	0.0090	73.2	4.43	73.4	19.8	p-CS/vdW
$O^{-*}-^{o}H(\pi)$ in 15b	0.0129	0.0056	0.0012	-0.877	0.0057	77.7	13.09	77.9	0.4	p-CS/vdW
$O^{-*}-^{o}H(\pi)$ in 15c	0.0112	0.0047	0.0010	-0.883	0.0048	78.2	16.49	79.9	24.3	p-CS/vdW
ΟΗ-*-C(π)										
$OH^{-*}-\alpha C(\pi)$ in $3a^m$	0.0147	0.0061	0.0018	-0.832	0.0063	73.9	7.95	82.7	185	p-CS/vdW
OH-*- α C(π) in 3 \mathbf{a}^n	0.0132	0.0050	0.0016		0.0053		9.68	81.0	199	p-CS/vdW

^a Data are collected for the conformers, where the intramolecular non-covalent interactions around the OH group are detected. Fig. 3 illustrates the bata are conferred to the conformers, where the intramolecular interactions around the OH group are detected. Fig. 3 must are conformers with contour plots drawn on the optimized structures for the conformers shown in this table, while those other than the ones above are presented in Fig. S3–S5 of the ESI. b MP2/6-311+G(3df,3pd) for MP2/BSS-A. c $\nabla^2 \rho_b(\mathbf{r}_c) = H_b(\mathbf{r}_c) - V_b(\mathbf{r}_c)/2$, where $c = \hbar^2/8m$. d $k_b(\mathbf{r}_c) = V_b(\mathbf{r}_c)/G_b(\mathbf{r}_c)$, where $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, where $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, where $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, where $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, where $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, where $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, $(\mathbf{r}_c) = (\mathbf{r}_c) + (\mathbf{r}_c)/2$, where $(\mathbf{r}_c) = (\mathbf{r}_c)/2$, where 2.3277 Å versus 2.2797 Å (MP2)). Calculated with B3LYP/BSS-A $(r(H \cdots C(\pi)) = 2.3782$ Å versus 2.2797 Å (MP2)).

MC/CT-TBP and CT-TBP/Cov-w (weak covalent bonds), respectively. The parameters described in bold are superior to those tentatively given as parameters in the classification and/or characterization of interactions. The SS (180 $^{\circ}$ < θ) and r-CS $(90^{\circ} < \theta < 180^{\circ})$ interactions were not detected in each intramolecular HB interaction studied in this work. Therefore, R is not employed for the characterization in this work.

In the case of the intramolecular $OH^{-*}-C(\pi)$ interactions in Table 1, the $(\theta, \theta_{p:CIV})$ values are $(77.0-80.5^{\circ}, 94.0-115.9^{\circ})$ for all the intramolecular OH-*-C(π) interactions, except for 11a, of which $(\theta, \theta_{p:CIV}) = (76.8^{\circ}, 79.3^{\circ})$. Therefore, the intramolecular OH-*-C(π) interactions in 3a, 6a, 9a, 12a, 13a and 15a are predicted to have a t-HB_{nc} nature, as appeared in the p-CS region (p-CS/t-HB_{nc}), whereas the nature of the interaction in 11a has a p-CS/vdW nature. The intramolecular O^{-*} -C(π) interactions show a similar trend relative to the intramolecular OH-*-C(π) interactions. The (θ , $\theta_{p:CIV}$) values are (76.4–79.4°, 90.8–107.2°) for **6c** and **9b**, whereas the values are $(72.0-78.5^{\circ})$, 49.7-86.1°) for 3b and 14b. Consequently, the nature of the intramolecular O-*-C(π) interactions in **6c** and **9b** is predicted to have a p-CS/t-HB_{nc} nature, while the predicted nature is p-CS/ vdW for 3b and 14b. However, the $(\theta, \theta_{p:CIV})$ values are (73.2-78.2° and 73.4–88.5°) for the intramolecular O^{-*} – $H(\pi)$ interactions in 5e, 5i, 12b, 15b and 15c, which should have a p-CS/vdW nature. The strength of the three types of interactions is roughly predicted to be smaller in the order OH-*-C(π) > O-*- $C(\pi) > O$ -*- $H(\pi)$ (see Table 1).

The calculated $\theta_{p:CIV}$ values are usually larger than or close to the θ values for the usual interactions. However, the $\theta_{\text{p:CIV}}$ in 3b (49.7°) is predicted to be much smaller than θ (78.5°) for the O-*-C(π) interaction. The reason is unclear when using the data in Table 1. It would have originated from the substantially distorted nature of the O-*-C(π) interaction in 3b ($\Delta r_{\rm BP} = 0.295 \, \text{Å}$). The BP for the intramolecular O^{-*} – $C(\pi)$ interaction in 3b seems to be very close to the O-H bond in 3b, which would also be a reason for this phenomena.

The effects from basis sets and levels on the optimized structures and the calculated natures of the interactions in question must be an important issue of QTAIM approach. The effects on the standard interactions, containing hydrogen bonds, are carefully examined, which will be discussed elsewhere.37 The effects from the DFT level of M06-2X (M06-2X/BSS-A//M06-2X/BSS-A: M06-2X/BSS-A) and B3LYP (B3LYP/BSS-A// B3LYP/BSS-A: B3LYP/BSS-A) on the nature of OH-*- α C(π) in 3a were examined, in this paper. Table 1 shows the results. The $r(H \cdots C(\pi))$ values were optimized as 2.3277 and 2.3782 Å at the Paper RSC Advances

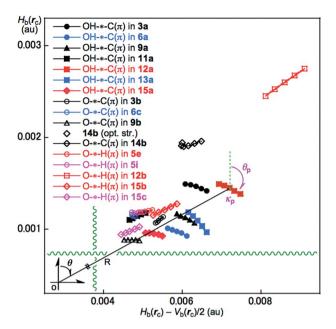


Fig. 4 Plots of $H_b(r_c)$ versus $H_b(r_c) - V_b(r_c)/2$ for the intramolecular interactions around the OH group in 3a, 3b, 5e, 5i, 6a, 6c, 9a, 9b, 11a, 12a, 12b, 13a, 14b, 15a, 15b and 15c, as evaluated with MP2/BSS-A. The perturbed structures are generated with a CIV.

M06-2X and B3LYP levels, respectively, which are 0.048 and 0.099 Å longer, relative to that optimized at the MP2 level (2.2797 Å). On the other hand, the $(\theta, \theta_{\text{p:CIV}})$ values of (73.9°, 82.7°) and (72.5°, 81.0°) were calculated for OH–*– $^{\alpha}$ C(π) in 3a at the M06-2X and B3LYP levels, respectively, irrespective of the calculated $r(\text{H····C}(\pi))$ values, while the (77.0°, 97.7°) values were calculated at the PM2 level. As a result, the p-CS/vdW nature was predicted for OH–*– $^{\alpha}$ C(π) in 3a at the M06-2X and B3LYP levels, whereas the p-CS/t-HB_{nc} nature was at the MP2 level. The effects from M06-2X/BSS-A and B3LYP/BSS-A seem not small, relative to the case of MP2/BSS-A.

The strength of the intramolecular interactions is discussed in relation to those of the NBO analysis in the next section.

NBO analysis for intramolecular interactions

The stabilization energy E(2) is calculated by NBO analysis²⁴ for each donor NBO (i) and acceptor NBO (j) based on the second-order perturbation theory according to eqn (8). The q_i values in eqn (8) are the donor orbital occupancy, E_i and E_j are diagonal elements (orbital energies) and F(i,j) is the off-diagonal NBO Fock matrix element. The treatments will evaluate the CT terms of the intramolecular interactions.

$$E(2) = q_{i}F(i,j)^{2}/(E_{i} - E_{i})$$
(8)

The NBO (version 3.0) was applied to the conformers, where BPs with BCPs corresponding to the intramolecular OH-*-C(π), O-*-C(π) and/or O-*-H(π) interactions were detected. The NBO (version 3.0) was also applied to the conformers for which the OH···C(π) distances are less than 2.9 Å. The E(2) values were successfully obtained under the

Table 2 Results of the NBO analysis with NBO 3.0 for the intramolecular interactions around the OH group, as evaluated with MP2/ BSS-A

Species	$E(2)^a$ (kJ mol ⁻¹)	$E(i) E(i)^b (cor)$	$F(i,j)^c$ (au)	$r(\mathbf{H}\cdots\mathbf{C}(\pi))$	
Species	E(2) (KJ IIIOI)	E(J)-E(I) (au)	r(i,j) (au)	(Å)	
CT term	of the π(C≡C/C=C	$C) \rightarrow \sigma^*(H-O)$ ty	pe		
3a	10.0	1.14	0.047	2.2797	
5a	4.8	1.09	0.032	2.4802	
5b	3.7	1.10	0.028	2.6111	
6a	13.8	1.11	0.054	2.3020	
8a	2.7	1.02	0.025	2.5218	
9a	7.2	1.05	0.042	2.3316	
10a	4.5	1.52	0.036	2.2584	
11a	2.1	1.13	0.021	2.4507	
12a	11.6	1.08	0.049	2.4139	
13a	20.1	1.09	0.065	2.2783	
14a	8.9	1.00	0.045	2.3601	
15a	4.4	1.05	0.032	2.3869	
CT term	of the $\sigma(H-O) \rightarrow \tau$	τ*(C≡C/C≕C) ty	pe		
10a	3.6	1.27	0.029	2.2584	
15b	2.1	1.47	0.027	4.1125	
CT term	of the $n_p(O) \rightarrow \pi^*(O)$	(C≡C/C=C) type	;		
1a	5.7	0.82	0.030	2.5155^{d}	
3b	2.3	0.79	0.018	2.4272^{e}	
4a	6.9	0.72	0.031	2.5383^{f}	
4c	5.4	0.73	0.027	2.6315^{g}	
6a	2.3	0.70	0.017	2.3020^{h}	
6b	3.7	0.73	0.023	3.4771^{i}	
6c	3.8	0.73	0.023	3.5116 ^j	
7a	3.7	0.67	0.024	2.5311^{k}	
CT term	of the $n_p(O) \rightarrow \sigma^*(O)$	(C-H) type			
15c	2.3	1.34	0.024	4.2452^{l}	
CT term	of the $n_s(O) \rightarrow \pi^*($	C-H) type			
10b	3.5	1.18	0.028	3.6982^{m}	
CT term	of the $n_s(O) \rightarrow \sigma^*($	C-H) type			
12b	3.3	1.59	0.032	3.8536^{n}	
15b	19.3	7.82	0.170	4.1125^{o}	

^a Second order perturbation energy given by eqn (8). ^b The diagonal elements (orbital energies). ^c The off-diagonal NBO Fock matrix element. ^d 2.4027 Å for $r(O \cdots C(\pi))$. ^e 2.9757 Å for $r(O \cdots C(\pi))$. ^f 2.4170 Å for $r(O \cdots C(\pi))$. ^g 2.4305 Å for $r(O \cdots C(\pi))$. ^h 3.0039 Å for $r(O \cdots C(\pi))$. ⁱ 3.0044 Å for $r(O \cdots C(\pi))$. ^j 2.9581 Å for $r(O \cdots C(\pi))$. ^k 3.0080 Å for $r(O \cdots C(\pi))$. ^l 2.4484 Å for $r(HO \cdots H)$. ^m 2.74345 Å for $r(O \cdots C(\pi))$. ⁿ 2.2155 Å for $r(HO \cdots H)$. ^o 2.3578 Å for $r(HO \cdots H)$.

threshold of 0.5 kcal mol $^{-1}$ (2.1 kJ mol $^{-1}$). Table 2 collects the results of the NBO analysis, as calculated with MP2/BSS-A. The CT terms of the $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ type contribute to E(2) in 3a, 5a, 5b, 6a, 8a, 9a, 10a, 11a, 12a, 13a, 14a and 15a, together with the inverse $\sigma(H-O) \rightarrow \pi^*(C \equiv C/C = C)$ type for 10a and 15b. The CT terms of the $n_p(O) \rightarrow \pi^*(C \equiv C)$ type contribute to E(2) in 1a, 3b, 4a, 4c, 6a, 6b, 6c and 7a. However, the CT term of the $n_p(O) \rightarrow \sigma^*(C-H)$ type was detected in 15c and that of the $n_s(O) \rightarrow \pi^*(C-H)$ type was in 10b, while the term of the $n_s(O) \rightarrow \sigma^*(C-H)$ type was in 12b and 15b.

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The E(2) values larger than 7.0 kJ mol⁻¹ were predicted for the CT terms of the $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ interactions in 3a, 6a, 9a, 13a and 14a, for which the $OH \cdots C(\pi)$ distances are less than 2.36 Å. A BP with a BCP corresponding to the intra-

molecular OH-*-C(π) interaction was detected for each case, except for 14a. The E(2) values of less than 4.8 kJ mol⁻¹ were similarly predicted for 5a, 5b, 8a, 11a and 15a, among which the $OH \cdots C(\pi)$ distances were longer than 2.38 Å. In this case, the BP with the BCP was detected for 11a, 12a and 15a, whereas it was not detected for 5a, 5b, 8a and 10a. The E(2) value of 20.1 kJ mol⁻¹ was evaluated for the intramolecular OH-*-C(π) interaction in 13a, which is larger than those in 3a, 6a, 9a, 11a, 12a and 15a $(2.1-13.8 \text{ kJ mol}^{-1})$. This must be the reason for the stronger intramolecular OH-*-C(π) interaction in 13a, relative to the cases in 3a, 6a, 9a, 11a, 12a and 15a evaluated with QTAIM-DFA. The large E(2) value of 20.1 kJ mol⁻¹ in 13a may come from the short $OH \cdots C(\pi)$ distance (2.28 Å), although the other advantageous structural parameters around the OH-*-C(π)-predicted CT interaction in **11a** are not of the π (C \equiv C) $\rightarrow \sigma^*(H-O)$ type but rather the inverse type of $\sigma(H-O) \rightarrow$ $\pi^*(C \equiv C)$. The E(2) value was evaluated to be 2.1 kJ mol⁻¹ for the intramolecular interaction of 11a.

Substantially large E(2) values are evaluated for the intramolecular $OH \cdots C(\pi)$ interaction by NBO, if a BP with a BCP corresponding to the intramolecular $OH^{-*}-C(\pi)$ interaction was detected for the conformer. In the case of 12b, a rather small E(2) value (3.3 kJ) was evaluated for the $n_p(O) \rightarrow \sigma^*(H-C)$ interaction.

However, a much larger E(2) value of 19.3 kJ was predicted for the $n_s(O) \rightarrow \sigma^*(C-H)$ interaction in 15b, which must be the reason for the predicted strong intramolecular O^{-*} -H(π) interaction for 15b using QTAIM-DFA. The CT terms were not printed out for the intramolecular interactions in 5e and 5i, although the BPs with the BCPs of the $O^{-*}_{-}^{\beta}H(\pi)$ type were detected. The results seem to be queries at first glance. They would be the results from the intramolecular vdW type interactions in 5e and 5i. The contributions of the CT terms must be (very) small for the vdW type interactions; therefore, the E(2)values should be evaluated to be (very) small, which would be buried in the threshold value of 2.1 kJ mol⁻¹ (0.5 kcal mol⁻¹).

The results of the NBO analysis are discussed in relation to the $\Delta E_{\rm ES}$ values in the next section.

Intramolecular $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ interactions as the factor to stabilize the conformers

Are the conformers effectively stabilized through the intramolecular CT interactions? The stability of the conformers are discussed in relation to the E(2) values calculated with the NBO, as exemplified by the energy differences between conformer a and b, $\Delta E_{\rm ES}(x\mathbf{b}/x\mathbf{a}) = \Delta E_{\rm ES}(x\mathbf{b}) - \Delta E_{\rm ES}(x\mathbf{a})$. The x values were limited to 3, 6, 9, 11-13 and 15, where the BP with BCP of the $OH^{-*}-C(\pi)$ type were detected in xa. The $OH\cdots C(\pi)$ distance must be the shortest in xa by definition; therefore, the steric hindrance is expected to reach its maximum in xa, although the xa will contain the attractive factor based on the intramolecular $OH\cdots C(\pi)$ interaction. The intramolecular $OH\cdots C(\pi)$ distance

in xb is the second shortest, by definition, and therefore the steric hindrance in xb would be somewhat released in most cases due to the change of $\phi(C_{\beta}C_{\gamma}OH)$ from xa. As a result, the $\Delta E_{\rm ES}(xb/xa)$ is expected to be a rough measure for the contribution from the intramolecular interaction in xa if the contribution from the intramolecular interaction is (almost) varnished in xb.

The nature of the interactions in question can be clarified based on the BPs with the BCPs, but the intramolecular interactions are carefully discussed based on BPs with BCPs. The theoretical treatment for the appearance and/or disappearance of BPs is very complex and very difficult.38 Namely, the theoretical treatment for the intramolecular interactions in detail is beyond the scope of this work. Therefore, the $\Delta E_{\rm ES}(x{\bf b}/x{\bf a})$ values are discussed here, where the BPs with BCPs are detected for some of the conformers from 1-15, whereas some are not. The $\Delta E_{\rm FS}(xb/xa)$ values are discussed in relation to the E(2) values for the intramolecular interactions evaluated with the NBO and the steric effect in the conformers.

Fig. 5 shows the plot of E(2) and $\Delta E_{ES}(x\mathbf{b}/x\mathbf{a})$ for x=3,6,9, 11-13 and 15 in red and blue, respectively. The $\Delta E_{\rm ES}(x\mathbf{b}/x\mathbf{a})$ values are evaluated over a range of 3.8 $\leq \Delta E_{\rm ES} \leq$ $13.8 \text{ kJ mol}^{-1} \text{ for } 3\text{b}/3\text{a}, 6\text{b}/6\text{a}, 9\text{b}/9\text{a} \text{ and } 11\text{b}/11\text{a} \text{ to } 13\text{b}/13\text{a}$ with $\Delta E_{\rm ES} = -0.2$ kJ mol⁻¹ for 15b/15a. However, the E(2)values are calculated at a range of $2.1 \le \Delta E_{\rm ES} \le 13.8 \text{ kJ mol}^{-1}$ for xa (x = 3, 6, 9, 11, 12 and 15) with 20.1 kJ mol⁻¹ for 13a, as shown in Fig. 5 and Table 2. The $\Delta E_{\rm ES}$ values of 3b/3a $(8.0 \text{ kJ mol}^{-1})$, 6b/6a $(13.8 \text{ kJ mol}^{-1})$ and 9b/9a $(10.0 \text{ kJ mol}^{-1})$ are close to the E(2) values of 3a (10.0 kJ mol⁻¹), 6a (13.8 kJ mol^{-1}) and 9a (7.2 kJ mol^{-1}). The results can be reasonably explained by assuming that the intramolecular $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ interactions can effectively stabilize the conformers of the ethenyl and ethynyl derivatives of the aliphatic alcohols. In the case of the phenol and benzyl alcohol derivatives, the $\Delta E_{\rm ES}$ values of 12b/12a (3.8 kJ mol⁻¹), 13b/13a (7.8 kJ mol⁻¹) and 15b/15a (-0.2 kJ mol⁻¹) are substantially smaller than the E(2) values of 12a $(11.6 \text{ kJ mol}^{-1})$, 13a $(20.1 \text{ kJ mol}^{-1})$ and 15a $(4.4 \text{ kJ mol}^{-1})$, respectively. Other factors seem to waste the contributions from the attractive intramolecular $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ interactions. A repulsive steric effect would greatly waste the attractive interactions in 12a and 13a. The intramolecular interactions operate more effectively to stabilize 12a and 13a relative to 12b and 13b, respectively, which would come from the steric hindrance in 12b and 13b larger than 12a and 13a, respectively. The $\Delta E_{\rm ES}(15b/15a)$ value is predicted to be -0.2 kJ mol^{-1} . The intramolecular $n_p(O) \rightarrow \sigma^*(C-H)$ interaction stabilizes 15b, very effectively, as shown by the NBO analysis (see, Table 2). In fact, the $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ interaction acts to stabilize 15a (4.4 kJ mol⁻¹), but the $\sigma(H-O)$ $\rightarrow \pi^*(C \equiv C/C = C)$, $n_p(O) \rightarrow \sigma^*(C-H)$ and $n_s(O) \rightarrow \sigma^*(C-H)$ interactions also operate to stabilize 15b (2.1 kJ mol⁻¹), 15c $(2.3 \text{ kJ mol}^{-1})$ and 15b $(19.3 \text{ kJ mol}^{-1})$, respectively. The contributions from the intramolecular interactions to stabilize 15a and 15b must be the primary factor in the negative value of $\Delta E_{\rm ES}(15b/15a)$, although the mechanism, similar to the case of 11b/11a and 11b/11a, must also be working.

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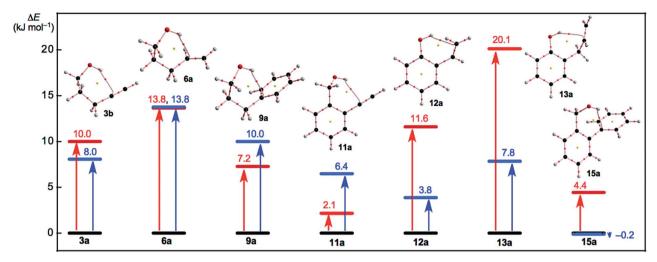


Fig. 5 Plots of E(2) and $\Delta E_{ES}(xb/xa)$ for x = 3, 6, 9, 11–13 and 15 in red and blue, respectively. Molecular graphs for xa are shown, where BCP with BPs corresponding to the intramolecular OH-*- $C(\pi)$, O-*- $C(\pi)$, or O-*- $H(\pi)$ interactions are given for each. Molecular graphs other than xaare drawn in Fig. S3-S5 of the ESI.†

The $\Delta E_{\rm ES}({\bf 1b/1a}: 6.6 \text{ kJ mol}^{-1})$ and $\Delta E_{\rm ES}({\bf 5b/5a}: 6.8 \text{ kJ mol}^{-1})$ values seem close to the $\Delta E_{\rm ES}$ (3b/3a: 8.0 kJ mol⁻¹) and smaller than the $\Delta E_{\rm ES}$ (6b/6a: 13.8 kJ mol⁻¹), for example. The intramolecular $n_p(O) \rightarrow \pi^*(C \equiv C/C = C)$ and $\pi(C \equiv C/C = C) \rightarrow$ $\sigma^*(H-O)$ interactions operate to stabilize **1a** (5.7 kJ mol⁻¹) and $5a (4.8 \text{ kJ mol}^{-1})$, respectively, while $3a (10.0 \text{ kJ mol}^{-1})$ and 6a(13.8 kJ mol⁻¹) are stabilized by the $\pi(C \equiv C/C = C) \rightarrow \sigma^*(H-O)$ interaction. In this case, a BP with a BCP is detected for 3a and 6a, whereas a BP with a BCP is not detected in 1a and 5a. The conditions for the appearance of the BP with BCP would not be satisfied for the corresponding interaction in 1a and 5a.

Conclusions

Intramolecular cv-HBs are extremely important in the all fields of the chemical and biological sciences as are intramolecular π -HBs. The intrinsic dynamic and static nature of intramolecular π -HBs is elucidated here using QTAIM-DFA. The perturbed structures necessary for QTAIM-DFA are generated using coordinates derived from the compliance constants (CIV). Over 70 conformers were optimized for 1-15. BPs with BCPs corresponding to the intramolecular $OH^{-*}-C(\pi)$ interactions were detected in seven conformers that appeared in the sixmembered rings of the -COH-*-CCC type. Those corresponding to the intramolecular O^{-*} - $C(\pi)$ interactions are also detected in four conformers together with those for the intramolecular O^{-*} -H(π) interactions in five conformers. The intramolecular O^{-*} – $C(\pi)$ and O^{-*} – $H(\pi)$ interactions appeared in the 5-7-membered rings. The BPs are somewhat curved. The intramolecular OH-*-C(π) interactions are predicted to have a p-CS/vdW to p-CS/t-HB_{nc} nature. The strength of the intramolecular interactions appears to be generally weaker in the order OH-*-C(π) > O-*-C(π) > O-*-H(π) (see Table 1). The contributions from the intramolecular $OH^{-*}-C(\pi)$, $O^{-*}-C(\pi)$ and O^{-*} -H(π) interactions towards stabilizing the conformers are also confirmed by the NBO analysis (see Table 2). The

contributions of the intramolecular interactions used to stabilize the conformers are considered, even for the conformers with no appearance of BPs, which corresponds to the intramolecular interactions. The intramolecular interactions between the OH and C≡C/C=C groups, such as the OH-*-C(π), O-*-C(π), or O-*-H(π) types, may help to stabilize the conformers even if this interaction is not detected as a BP with BCP. The intramolecular OH-*- π interactions appear to be evaluated as somewhat stronger than the intermolecular interactions.

Conflicts of interest

The authors declare no conflict of interest.

Acknowledgements

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