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One-pot synthesis of 2-amino-4*H*-chromene derivatives by MNPs@Cu as an effective and reusable magnetic nanocatalyst†

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In this research, MNPs@Cu as an effective and recyclable nanocatalyst was prepared and characterized using different methods including Fourier transform infrared spectroscopy (FT-IR), thermogravimetric analysis (TGA), vibrating sample magnetometry (VSM), scanning electron microscopy (SEM), energy dispersive X-ray spectroscopy (EDX), and X-ray diffraction (XRD). After the characterization of this new nanocatalyst, it was efficiently used for the promotion of the one-pot synthesis of 2-amino-4*H*-chromene derivatives *via* one-pot three-component reaction of the enolizable compound, malononitrile, and arylaldehydes under solvent-free conditions at 90 °C. The procedure gave the desired products in high-to-excellent yields in short reaction times. Also this catalyst, because of its magnetic nature, can be simply restored by a permanent magnetic field and comfortably reused several times without any significant loss of its catalytic activity.

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Introduction

Heterocyclic compounds have wide applications and are essential in various pharmaceutical compounds. One of the key reasons for the widespread use of heterocyclic compounds is constructing various structures to achieve the desired functions.^{1,2} Oxygen-bearing heterocyclic molecules such as 4*H*-pyrans and 4*H*-pyran-annulated heterocyclic scaffolds (4*H*-chromene moieties) are very important because of their ubiquitous presence in nature and provision of appropriate biological activity, which make them an interesting issue for many researchers in organic chemical science. These structural motifs are very beneficial derivatives that develop activities such as anti-HIV, anti-tumor, antimicrobial, anti-inflammatory, antifungal, and anti-allergenic with respect to the available substituents in the structure.^{3–12}

Moreover, it has been observed that pyran-containing heterocyclic compounds are widespread in nature and their uses in pharmaceuticals, biodegradable agrochemicals, pigments, and cosmetics industries are becoming more and

more important.^{13–16} Natural resources for preparing these compounds are limited and, in most cases, it is required to handle a difficult process for the isolation and purification of them. Therefore, to synthesize compounds from this group, many efforts have been made to investigate derivatives progress and their medicinal features.^{17–25} Many procedures have been developed for the preparation of functionalized 2-amino-4*H*-chromenes *via* one-pot Knoevenagel condensation of C–H-activated acids, aldehydes, and malononitrile using catalysts such as SDS,¹⁷ DAHP,²⁶ MNP@AVOPc,²⁷ [Sipim]HSO₄,²⁸ nano Al₂O₃,²⁹ POPINO,³⁰ CuO nanoparticles³¹ and TBBDA.³² Some of these procedures suffer from one or more limitations, such as low yields of the desired product, long reaction times, toxicity, poor recovery and reusability of the catalyst, and tedious work-up. Therefore, a great demand still exists for simple, versatile, and environmentally friendly processes for the synthesis 2-amino-4*H*-chromene derivatives under simple conditions.

In green chemistry, scientists are trying to develop, design, and implement ways to prevent the use or generation of hazardous substances in chemical processes. Although innovation processes with environmentally friendly solvents provide a perfect solution to overcome these problems, in most cases, the solvent-free conditions in chemical processes play an inevitable role. Nowadays, solvent-free procedures in the existence of magnetic nanoparticles as a catalyst have eliminated most of the problems associated to the use of organic solvents for avoiding tedious separation procedures through extraction or distillation. Metal-catalysis has played a fundamental role in the prosperity of the industry.³³ The utilization of transition-metal nanoparticles in catalysis is definitive as they assume metal

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surface activation and catalysis at the nanoscale and thereby afford efficiency and selectivity to heterogeneous catalysis.^{34–37} Magnetic nanoparticles and analysis of their manipulation in the manufacturing of fine heterogeneous catalysis synthesis have achieved a major area of research as an effective material because of their good stability, low toxicity, easy synthesis, and functionalization, and large surface area.^{38–42} Due to lose dispersibility and catalytic activity through anisotropic dipolar attraction of the Fe_3O_4 nanoparticles, these kinds of uncoated nanoparticles with metal oxides can be easily aggregate into large clusters.⁴³ The acidic environment also causes the loss of their magnetic properties. It is essential to note that relatively small Fe_3O_4 nanoparticles may tolerate rapid biodegradation when they are directly exposed to biological systems.⁴⁴ The suitable coating with a protective shell of silica to form a core-shell structure is essential to overcome such limitation.⁴⁵ Therefore, the use of magnetic coated nanoparticles as suitable alternative supports for the functionalization of organic compounds in the preparation of heterogeneous catalysis is of paramount importance.

Experimental

General

All the pure chemicals were prepared from Merck and Fluka chemical companies. Separation and purification of the products were performed with extraction and recrystallization techniques. Melting points were determined with an Electrothermal-9100 apparatus and no correction was made. FT-IR spectra were obtained by means of a PerkinElmer PXI spectrometer in KBr wafers. TGA analysis was conducted using a TGA thermoanalyzer (PerkinElmer) instrument. Samples were heated from 25 to 800 °C at ramp 10 °C min^{−1} under an N_2 atmosphere. The X-ray diffraction (XRD) measurements of catalyst were performed with a Siemens D-500 X-ray diffractometer (Munich, Germany). Scanning electron microscopy was

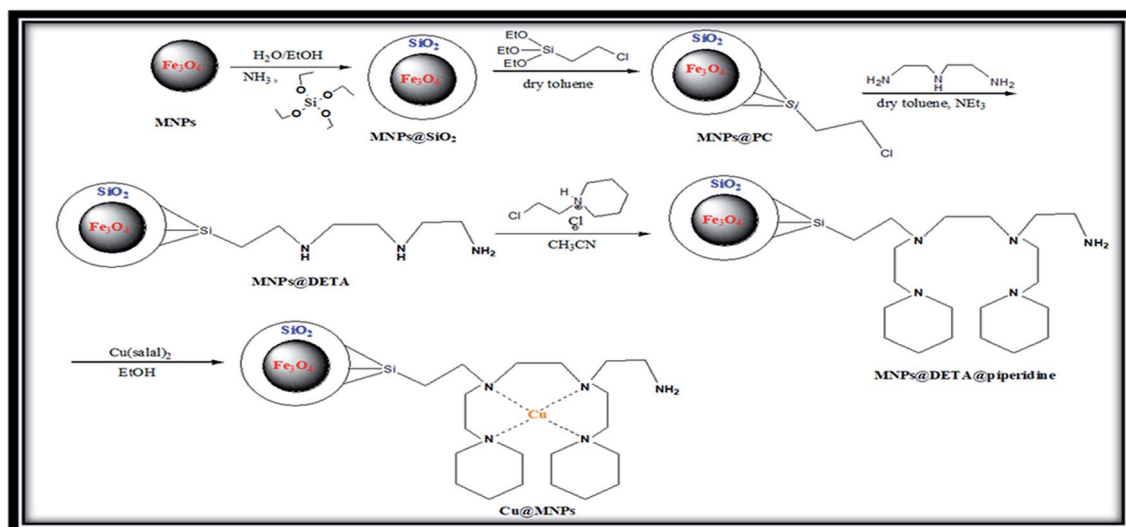
accomplished using an SEM-LEO 1430VP analyzer. Magnetic susceptibility measurements were accomplished using a vibrating sample magnetometer (VSM/AGFM, MDK Co, Ltd, Iran) in the magnetic field range of −8000 Oe to 8000 Oe at ambient temperature. Magnetic susceptibility measurements were performed by a VSM/AGFM (MDK Co, Ltd, Iran) in the magnetic field range of −8000 Oe to 8000 Oe at 25 °C. Elemental analyses were carried out on a Carlo-Erba EA1110CNNO-S analyzer and agreed (within 0.30) with the calculated values.

Catalyst synthesis

Preparation of MNPs. This precursor was prepared according to the chemical co-precipitation method. To prepare MNPs, 5.2 g of $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ and 2 g of $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ were poured into a 250 mL flask reaction containing 40 mL deoxygenated water followed by adding 0.9 mL of concentrated HCl under a continuous flow of nitrogen gas. Next, 250 mL of NaOH solution (1.5 M) was poured dropwise into the reaction solution and stirred until the reaction was completed. After this period, the resulting black MNPs were magnetically separated and rinsed with deionized water and dried at 40 °C for 24 h.

Preparation of MNPs@ SiO_2 . About 1 g of MNPs was dispersed in a mixture containing water (5 mL), ethanol (45 mL), and concentrated aqueous ammonia (3 mL, 25 wt%) by ultrasonication. Next, 0.5 mL of tetraethylorthosilicate (TEOS) was poured and the mixture was mechanically stirred for 24 h at room temperature. After this period, the precipitates of core-shell MNPs@ SiO_2 nanoparticles were collected by an external magnetic and washed several times with ethanol. Finally, the MNPs@ SiO_2 precipitate was dried under vacuum oven at 70 °C for 5 h.

Preparation of MNPs bonded propyl chloride (MNPs@PC). In the following step, 1 g of core-shell Fe_3O_4 @ SiO_2 nanoparticles was dispersed in 20 mL dry toluene with the aid of ultrasonication for 20 min. Next, 2 mL of 3-chloropropyltriethoxysilane (CPTCSi) was poured in to the reaction solution



Scheme 1 All stages of the MNPs@Cu synthesis.



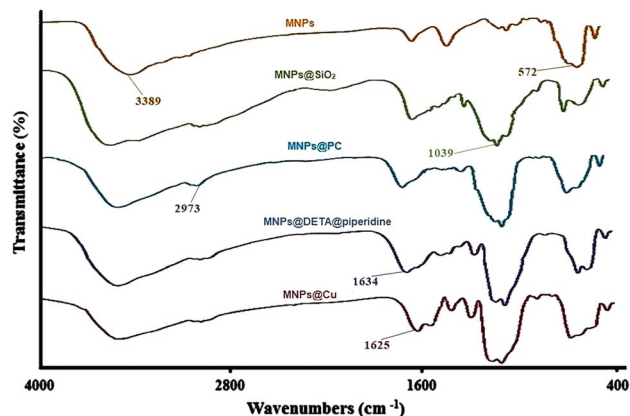


Fig. 1 FTIR spectra of MNPs, MNPs@SiO₂, MNPs@PC, MNPs@DETA@piperidine, and MNPs@Cu.

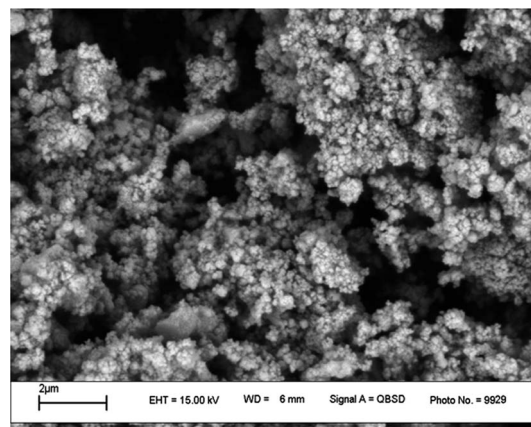


Fig. 4 SEM image of MNPs@Cu.

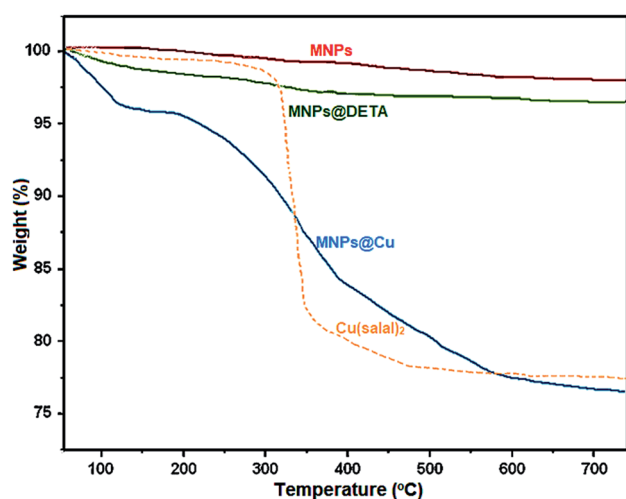


Fig. 2 TGA curves of MNPs, MNPs@DETA, Cu(salal)₂ and MNPs@Cu.

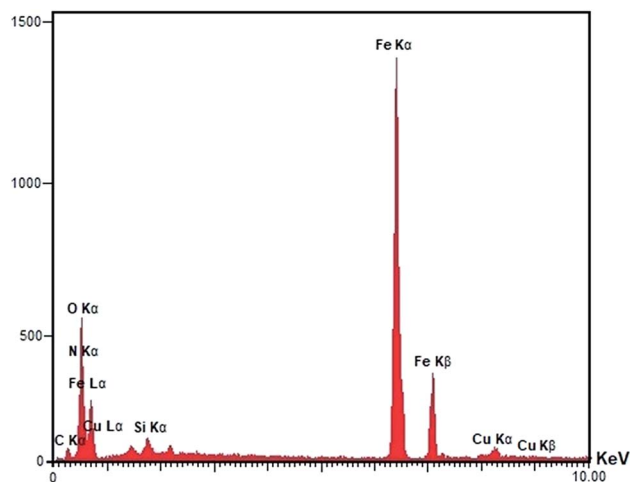


Fig. 5 EDX spectra of MNPs@Cu.

and the mixture was refluxed for 24 h under nitrogen atmosphere. At the end of the reaction, the solid phase was separated from the solvent by a powerful magnet and then washed several times with ethanol and distilled water. Eventually, the resulted product (MNPs@PC) was dried in a vacuum oven.

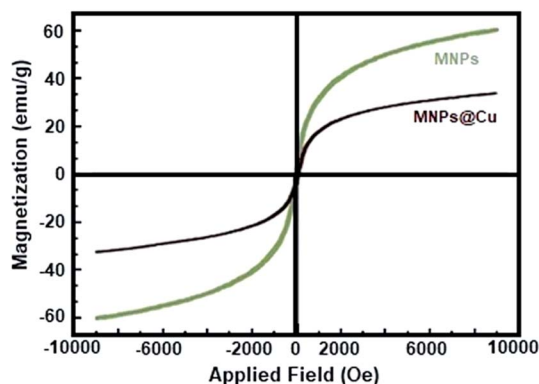


Fig. 3 VSM magnetization curves corresponding to the bare MNPs and MNPs@Cu.

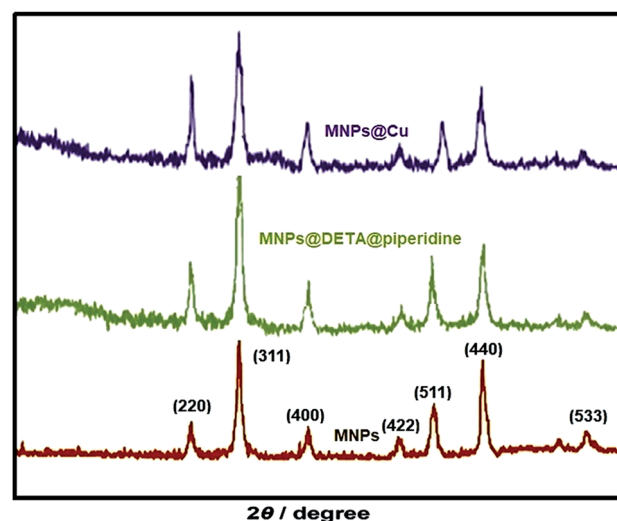
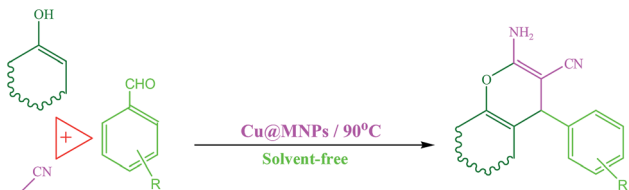


Fig. 6 The XRD patterns of MNPs, MNPs@DETA@piperidine, and MNPs@Cu.





Scheme 2 Synthesis of 2-amino-3-cyano-4H-pyran derivatives using MNPs@Cu.

Preparation of MNPs bonded *N*-propyl diethylenetriamine (MNPs@DETA). The prepared MNPs@PC (2 g) was dispersed in 100 mL of dry toluene for 30 min through sonication. Then, it was mixed with diethylenetriamine (5 mL) and triethylamine (1 mL) under reflux and nitrogen atmosphere and agitated for 12 h. The resultant solid (MNPs@DETA) was isolated using a permanent magnet and rinsed with ethanol several times to

eliminate the unreacted excess of the organic compounds and next dried in a vacuum oven.

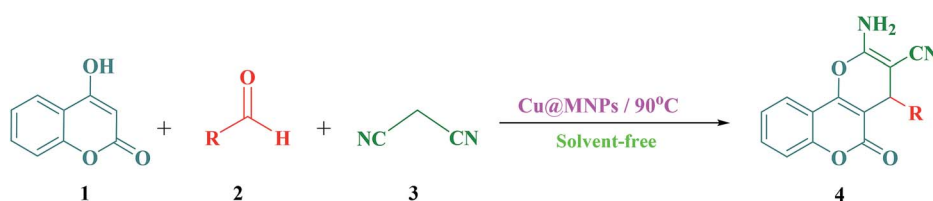
Preparation of MNPs@DETA@piperidine. The prepared MNPs@DETA (2 g) dispersed in 100 mL of acetonitrile was mixed with 1-(2-chloroethyl) piperidine (3 g) and the resultant mixture was agitated at 80 °C for 12 h. The resultant solid (MNPs@DETA@piperidine) was isolated using an external magnet and rinsed with water several times to eliminate the unreacted excess of the inorganic salts and next dried in a vacuum oven.

Coordination of Cu(II) with MNPs@DETA@piperidine. A mixture of MNPs@DETA@piperidine (1 g) and [Cu(salal)₂] (2 g) was dissolved in 20 mL of ethanol and then was stirred for 24 h under dry argon atmosphere. At the end of the reaction, the resultant precipitate MNPs@Cu formed was separated by filtration, washed twice with ethanol (15 mL) to remove unreacted metal precursors, and dried under vacuum oven to

Table 1 Optimization of the three-component reaction of 4-hydroxycoumarin (1), 4-chlorobenzaldehyde (2e), and malononitrile (3) under various conditions^a

Entry	Solvent	Catalyst (mg)	Temp.	Time (min)	Yield ^b (%)
1	CH ₂ Cl ₂	MNPs@Cu/10	Reflux	8	53
2	CH ₃ CN	MNPs@Cu/10	Reflux	8	59
3	toluene	MNPs@Cu/10	Reflux	8	59
4	DMF	MNPs@Cu/10	Reflux	8	63
5	H ₂ O	MNPs@Cu/10	Reflux	8	69
6	EtOH	MNPs@Cu/10	Reflux	8	77
7	Solvent-free	—	90 °C	8	Trace
8	Solvent-free	MNPs@Cu/10	60 °C	8	80
9	Solvent-free	MNPs@Cu/10	80 °C	8	87
10	Solvent-free	MNPs@Cu/10	90 °C	8	95
11	Solvent-free	MNPs@Cu/10	100 °C	8	95
12	Solvent-free	MNPs@Cu/10	110 °C	8	92
13	Solvent-free	MNPs@Cu/5	90 °C	8	64
14	Solvent-free	MNPs@Cu/10	25 °C	8	25
15	Solvent-free	MNPs@Cu/15	90 °C	8	91
16	Solvent-free	MNPs@DETA@piperidine/10	90 °C	20	85
17	Solvent-free	Cu(salal) ₂ /10	90 °C	25	78

^a Reaction conditions: 4-hydroxycoumarin (1 mmol), 4-chlorobenzaldehyde (1 mmol), malononitrile (1.1 mmol), and required amount of the catalysts. ^b The yields refer to the isolated product.



Scheme 3 One-pot three-component condensation of 4-hydroxycoumarin (1), various aldehydes (2), and malononitrile (3).



Table 2 MNPs@Cu-catalyzed synthesis of 2-amino-3-cyano-pyrano[3,2-c] chromen-5(4*H*)-one (4) derivatives^a

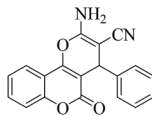
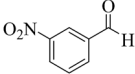
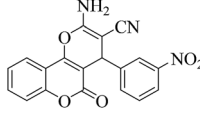
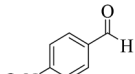
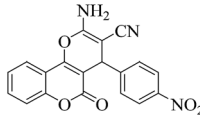
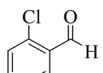
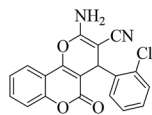
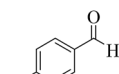
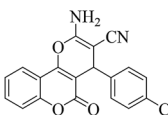
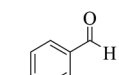
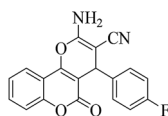
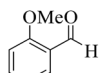
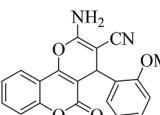
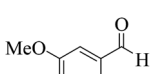
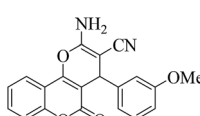
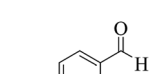
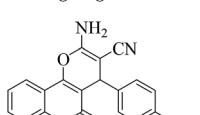
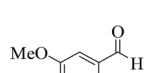
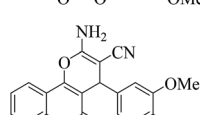
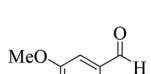
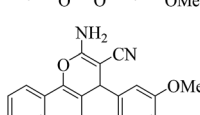
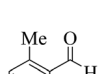
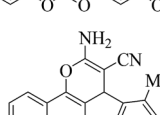
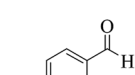
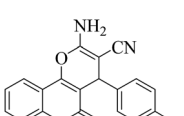
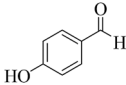
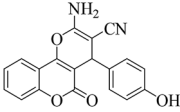
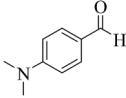
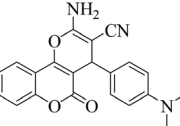
Entry	Product	RCHO (2)	Product	Time (min)	Yield (%) ^b	MP (obsd) (°C)	MP (lit.) (°C)
1	4a			12	91	258–260	256–258 ¹⁷
2	4b			12	92	260–263	260–262 ¹⁹
3	4c			8	94	249–251	250–252 ¹⁹
4	4d			8	94	243–245	244–246 ¹⁸
5	4e			8	95	262–264	263–265 ⁴⁸
6	4f			8	95	259–261	260–262 ¹⁷
7	4g			15	92	233–235	236–238 ⁵³
8	4h			20	90	240–243	242–244 ⁴⁶
9	4i			15	93	222–224	221–223 ¹⁸
10	4j			20	90	226–229	228–230 ¹⁷
11	4k			15	92	256–259	253–254 ⁴⁷
12	4l			15	90	261–264	260–262 ¹⁷
13	4m			15	90	250–252	252–254 ¹⁷



Table 2 (Contd.)

Entry	Product	RCHO (2)	Product	Time (min)	Yield (%) ^b	MP (obsd) (°C)	MP (lit.) (°C)
14	4n			15	95	258–260	259–261 ⁴²
15	4o			20	89	261–264	265–267 ¹⁷

^a Reaction conditions: 4-hydroxycoumarin (1 mmol), aldehyde (1 mmol), malononitrile (1.1 mmol), MNPs@Cu (10 mg).

obtain the pure product. All stages of the MNPs@Cu synthesis are shown in Scheme 1.

General procedure for the synthesis of 2-amino-4H-chromenes (4, 6, 9 and 10)

A mixture of, enolizable compound (1, 5, 7–8, and 1 mmol), malononitrile (1.1 mmol), aldehyde (1 mmol), and MNPs@Cu was combined with each other at 90 °C under solvent-free conditions. The development of the reaction was controlled with TLC analyses. At the end of the reaction, the mixture was rinsed with water and next, recrystallized from ethanol to achieve the pure product.

Results and discussion

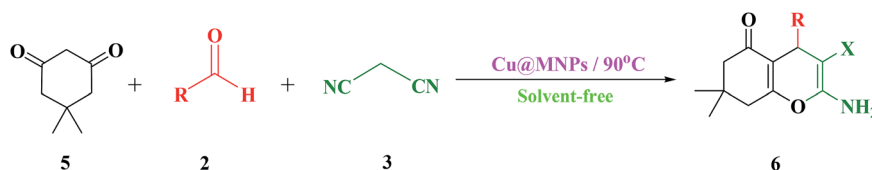
Catalyst characterization

FTIR analysis of MNPs@Cu. The corresponding FT-IR spectra of the MNPs, MNPs@SiO₂, MNPs@PC, MNPs@DETA@piperidine, and MNPs@Cu samples in the region of 400–4000 cm^{−1} are presented in Fig. 1. The bands at 572 and 3389 cm^{−1} in the Fe₃O₄ MNPs are related to the Fe–O stretching frequency and the OH groups attached to the iron, respectively. In the case of MNPs@SiO₂, the band's appeared at 1039 cm^{−1} can be ascribed to the stretching vibration of the Si–O groups. In about the MNPs@PC, the absorption bands at 2973 cm^{−1} are associated to the C–H stretching vibration mode. The obtained FT-IR spectra for MNPs@DETA@piperidine contained characteristic amine stretching vibrations at 1634 cm^{−1} related to diethylene triamine and piperidine groups. In the case of MNPs@Cu, the signals at 1634 cm^{−1} are shifted to 1625 cm^{−1}. The results clearly proved the effective coordination of metal groups with MNPs@DETA@piperidine.

Thermal analysis of MNPs@Cu. As observed in Fig. 2, the thermal constancy of the MNPs, MNPs@DETA, Cu(salal)₂, and MNPs@Cu were measured by thermogravimetric analysis (TGA) under a nitrogen atmosphere at 10 °C min^{−1} of heating rate. The TGA of Cu(salal)₂ has three stages: below 250 °C, 300–400 °C, and 400–500 °C which is related with the decomposition of salicylic chelate ligand. TGA data for another three samples show approximately 5% weight loss below 200 °C due to the dehydration of the surface hydroxy groups as well as desorption of physically adsorbed solvents. In the TGA graph of MNPs@DETA, a mass percentage loss of about 9% appeared that is related with the organic parts decomposition. The TGA of MNPs@Cu undergoes four stages of thermal degradation with a more mass loss at about 24.5%. In addition to the mentioned step in the range below 250 °C, the second stage can be seen in the range between 300–400 °C, which can be related to the elimination of salicylic chelate ligand. The second stage occurs in the range between 500–600 °C which is attributed to formation of a Fe–O–Cu bond and, finally, at 650 °C because of the decomposition of MNPs.

VSM analysis of MNPs@Cu. To investigate the magnetic properties of the MNPs and MNPs@Cu, magnetic measurements were performed by means of a vibrating sample magnetometer (VSM) with the field sweeping from −8500 to +8500 Oersted at ambient temperature (Fig. 3). As shown in VSM patterns, the saturation magnetization (*M_s*) values of MNPs and MNPs@Cu are 60.39 and 34.55 emu g^{−1}, respectively.

SEM analysis of MNPs@Cu. The morphology and size of the MNPs@Cu were surveyed to determine the variations in the surface of this nanocatalyst by the SEM analysis (Fig. 4). The SEM image of MNPs@Cu shows that the catalyst has a nearly spherical structure with nanometer-sized particles.



Scheme 4 One-pot three-component reaction of dimesedone (5), different aldehydes (2) and malononitrile (3).

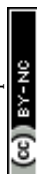


Table 3 MNPs@Cu-catalyzed synthesis of 2-amino-3-cyano-7,8-dihydro-4*H*-chromen-5(6*H*)-one (**6**) derivatives^a

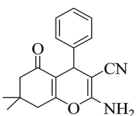
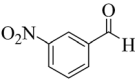
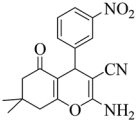
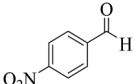
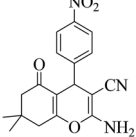
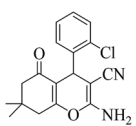
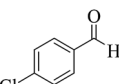
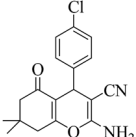
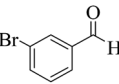
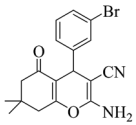
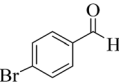
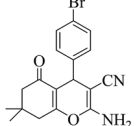
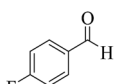
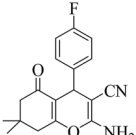
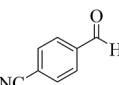
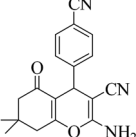
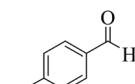
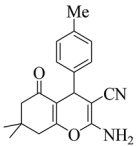
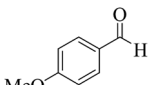
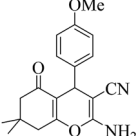
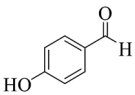
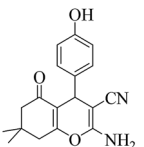
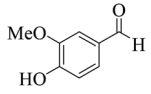
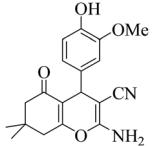
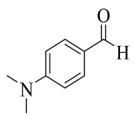
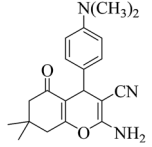
Entry	Product	RCHO (2)	Product	Time (min)	Yield (%) ^b	MP (obsd) (°C)	MP (lit.) (°C)
1	6a			17	91	232–234	234–235 ⁴⁹
2	6b			15	94	216–219	214–216 ²³
3	6c			15	92	180–182	179–180 ²⁵
4	6d			15	95	207–210	198–200 ²⁴
5	6e			15	93	212–214	215–217 ²³
6	6f			20	94	229–231	228–230 ³²
7	6g			15	92	208–211	207–209 ⁵²
8	6h			15	95	207–209	210–211 ⁴⁹
9	6i			15	92	225–229	227–230 ⁵²
10	6j			25	91	217–220	220–222 ⁵⁰
11	6k			25	90	203–205	201–202 ⁴⁹



Table 3 (Contd.)

Entry	Product	RCHO (2)	Product	Time (min)	Yield (%) ^b	MP (obsd) (°C)	MP (lit.) (°C)
12	6l			20	92	223–225	224–226 ⁵⁹
13	6m			25	94	237–239	238–240 ⁵¹
14	6n			20	92	209–211	210–212 ²⁶

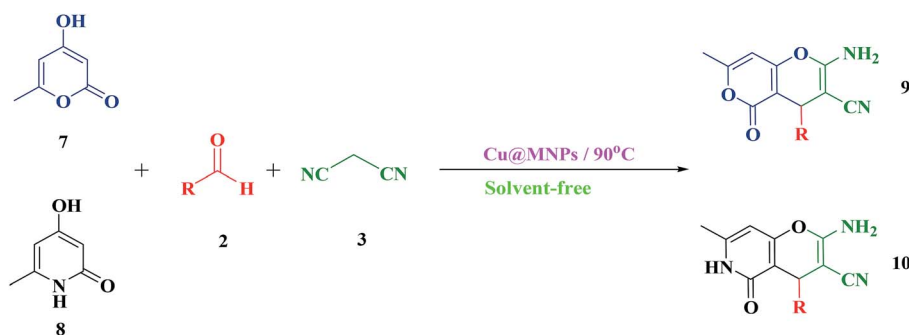
^a Reaction conditions: dimedone (1 mmol), aldehyde (1 mmol), malononitrile (1.1 mmol), MNPs@Cu (10 mg). ^b The yields refer to Isolated products.

EDX analysis of MNPs@Cu. Energy dispersive X-ray (EDX) mapping was carried out to verify the components of MNPs@Cu nanocatalyst (Fig. 5). The EDX analysis proved the presence of Fe, O, Si, C, N, and Cu characteristic peaks. Moreover, the appearance of copper bands in the nanocatalyst demonstrates the successful immobilization of copper(II) complex on the MNPs surface.

XRD analysis of MNPs@Cu. X-ray diffraction (XRD) is an effectual spectra method for the identification of the synthesis of magnetite crystal phase in the magnetic nano-catalyst. Fig. 6 exhibits the XRD patterns corresponding to the MNPs, MNPs@DETA@piperidine, and MNPs@Cu. In the XRD patterns of the all samples, the reflections at the values of 2θ – which are 30.27 (220), 35.34 (311), 43.73 (400), 54.45 (422), 57.51 (511), 63.19 (440), and 74.58 (533) – correspond to the spinel phase of magnetic iron oxide nanoparticles and are good agreement with (JCPDS card no. 79-0417). In the case of MNPs@DETA@piperidine, and MNPs@Cu, the broad reflection in $2\theta = 15\text{--}27^\circ$ is related to an amorphous silica shell. Therefore, it is concluded that the MNPs@Cu catalyst was synthesized successfully without damaging the crystalline structure of the MNPs core.

In current research, we reported our results for the efficient and rapid synthesis of 2-amino-3-cyano-4*H*-pyrans using MNPs@Cu as an effective and reusable heterogeneous nanocatalyst under solvent-free conditions (Scheme 2).

To monitor the reaction conditions for synthesizing 2-amino-4-(4-chlorophenyl)-3-cyano-pyrano[3,2-*c*] chromene-5(4*H*)-one (4e), the influence of the reaction temperature, the solvent, and the amounts of magnetic nanocatalyst were studied in the reaction of 4-hydroxycoumarin (1), 4-chlorobenzaldehyde (2e), and malononitrile (3) (molar ratio: 1 : 1 : 1.1) as a model reaction. The outcomes are presented in Table 1. To obtain the optimal reaction solvent, different solvents such as H₂O, CH₂Cl₂, CH₃CN, DMF, toluene, and EtOH in the existence of a certain amount of magnetic nanocatalyst were used. The trial reaction in the existence of EtOH, as a solvent, afforded the highest percentage of the target product 4e (Table 1, Entry 6) but the most favorable conditions in terms of rate and yield were obtained under solvent-free conditions for the reaction (Table 1, Entry 10). In the following phase of the survey, the influence of temperature on the completion of the



Scheme 5 One-pot three-component reaction of C–H-activated acids (7–8), different aldehydes (2) and malononitrile (3a).



Table 4 MNPs@Cu catalyzed synthesis of 2-amino-3-cyano-4H-pyran (9–10) derivatives^a

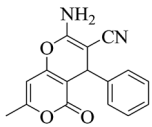
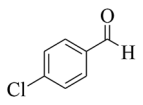
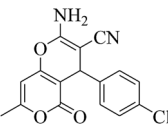
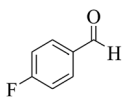
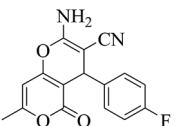
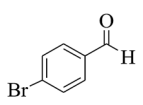
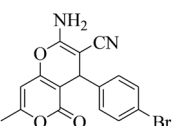
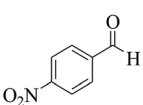
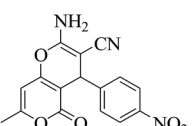
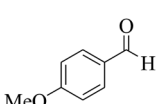
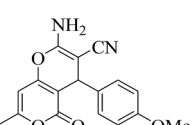
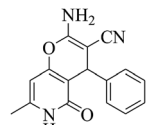
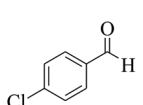
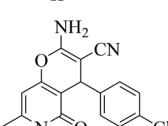
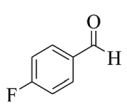
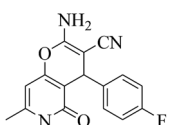
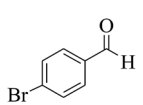
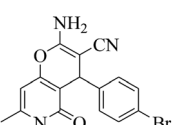
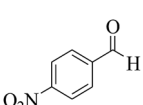
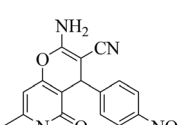
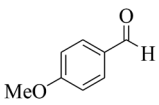
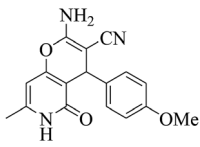
Entry	Product	RCHO (2)	Product	Time (min)	Yield ^b (%)	MP (obsd) (°C)	MP (lit.) (°C)
1	9a			20	91	237–239	236–238 ⁵³
2	9b			15	94	233–230	231–232 ⁵⁴
3	9c			15	95	224–221	223–225 ⁵⁵
4	9d			15	93	217–215	218–220 ⁵⁵
5	9e			15	94	223–219	220–222 ⁵⁶
6	9f			25	88	212–209	210–212 ⁵⁷
7	10a			25	91	276–280	279–282 ⁵⁸
8	10b			20	92	242–245	245–247 ⁵⁸
9	10c			20	93	258–261	258–259 ⁵⁵
10	10d			20	92	253–256	254–255 ⁵⁵
11	10e			20	94	276–279	278–279 ⁵⁵



Table 4 (Contd.)

Entry	Product	RCHO (2)	Product	Time (min)	Yield ^b (%)	MP (obsd) (°C)	MP (lit.) (°C)
12	10f			25	86	225–228	224–225 ⁵⁸

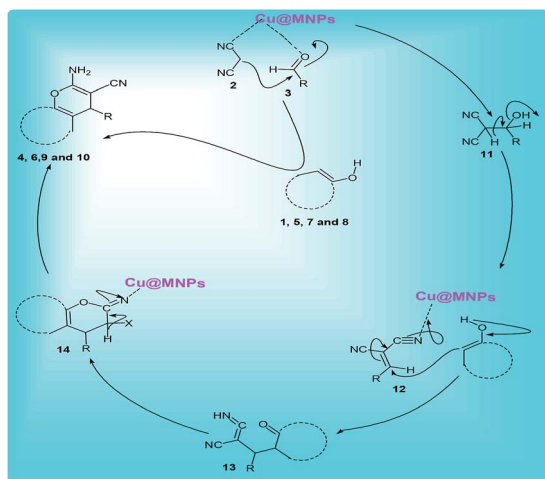
^a Reaction conditions: C–H-activated acid (**7** and **8**) (1 mmol), aldehyde (1 mmol), malononitrile (1.1 mmol), MNPs@Cu (10 mg). ^b The yields refer to isolated products.

reaction was investigated (Entries 8–12 and 14). Without heating, the reaction rate was very low and the yield was negligible (Table 1, Entry 14). The yield of the product increased up to 90 °C (Table 1, Entries 8–10), after which no enhancement was observed by increasing the temperature (Table 1, Entries 11 and 12). To obtain the optimum amount of magnetic nanocatalyst MNPs@Cu, diverse concentrations of nanocatalyst (5, 10, and 15 mg) were used to compare the reaction efficiency (Table 1, Entries 10, 13 and 15). After several screening tests with different concentrations of the nanocatalyst, it was found that the maximum yield (*i.e.*, 95%) at the short reaction time was achieved when 10 mg of MNPs@Cu was applied under solvent-free conditions (Entry 10). Enhancing the concentration of catalyst beyond 10–15 mg did not lead to any substantial progress in the product yield (Table 1, Entry 15) whereas the lower amount of catalyst needed for the reaction from 10 mg to 5 mg led to reducing the efficiency of the reaction (Table 1, Entry 13). Without using the catalyst, the product yields fell sharply and some stain was observed on thin-layer chromatography (TLC) (Table 1, Entry 7). Finally, when the model reaction was accomplished in the presence of 10 mg of MNPs@DETA@piperidine and Cu(salal)₂ under the optimized conditions, the yield of the product were 85 and 78%, respectively (Table 1, Entries 16 and 17). The favourable comparison of the

corresponding values for inputs 10, 16, and 17 accurately indicates that the catalyst activity increases when MNPs@DETA@piperidine is coordinate to the Cu(salal)₂ through the lone pair of nitrogen.

In order to generalize the optimum conditions (Table 1, Entry 10), we investigated the generality of this method with a range of appropriate aldehydes to prepare a series of 2-amino-3-cyano-pyrano[3,2-*c*] chromen-5(4*H*)-one derivatives (**4a–o**) under similar conditions (Scheme 3). The optimized results are represented in Table 2. It was observed that the aryl aldehydes reacted very well with both electron-withdrawing and electron-donating substituents and resulted the relating 2-amino-3-cyano-pyrano[3,2-*c*] chromen-5(4*H*)-one derivatives with high purity in good yields. The further precise analysis shows that aryl aldehydes having an electron-donating group (entries 7–15, Table 2) react slower compared to aryl aldehydes having electron withdrawing groups (entries 1–6, Table 2).

In the following, to expand the use of this nanomagnetic catalyst to the other reactions of these categories, a series of polyfunctionalized 2-amino-3-cyano-7,8-dihydro-4*H*-chromen-5(6*H*)-one derivatives (**6a–n**) were prepared from the reaction mixture of dimedone (**5**), different aldehydes (**2a–n**), and malononitrile (**3**) under the optimized reaction conditions mentioned above (Scheme 4). In this case, the use of dimedone, as an enolic component, improved the reaction time of the desired product **6** slightly (Table 3).



Scheme 6 A plausible mechanism for the one-pot three-component reaction of various enols (**1**, **5**, **7** and **8**), aldehydes (**2**), and malononitrile (**3**) catalyzed by MNPs@Cu under solvent-free conditions.

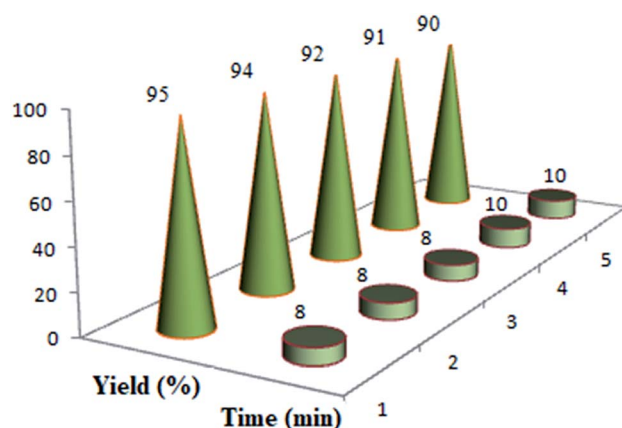


Fig. 7 The recycling of MNPs@Cu as a nanomagnetic catalyst under solvent-free conditions.



Table 5 Comparison of the current method with other reported strategies for synthesizing 2-amino-4-(4-chlorophenyl)-3-cyano-pyrano[3,2-c] chromene-5(4*H*)-one

Entry	Catalyst	Catalyst loading	Solvent	Temp (°C)	Time (min)	Yield (%)	Ref.
1	SDS	20 mol%	H ₂ O	60	150	88	17
2	CuO nanoparticles	15 mol%	H ₂ O	100	6	93	31
3	MNP@AVOPc	20 mg	—	25	15	94	27
4	Urea	10 mol%	H ₂ O/EtOH	25	420	93	59
5	POPINO	5 mol%	H ₂ O	Reflux	10	96	30
6	MNPs@Cu	10 mg	—	90	8	95	This work
7	DAHP	10 mol%	H ₂ O/EtOH	25	240	85	26
8	[Sipim]HSO ₄	0.08 mmol	—	100	30	90	28
9	Nano Al ₂ O ₃	25 mol%	EtOH	25	300	89	29
10	TBBDA	0.18 mmol	H ₂ O/EtOH	Reflux	170	91	32
11	DBU	10 mol%	H ₂ O	100	5	94	60
12	t-ZrO ₂ NPs	10 mol%	H ₂ O	80	38	89	61
13	ZnO NPs	10 mol%	EtOH	Reflux	10	80	62 ^a
14	Sodium alginate	10 mol%	EtOH	Reflux	50	93	63 ^a
15	KF/Al ₂ O ₃	250 mg	DMF	25	60–180	81	64 ^a

^a Reaction conditions: dimedone (1 mmol), 4-chlorobenzaldehyde (1 mmol), malononitrile (1 mmol).

After the successful synthesis of 2-amino-3-cyano-7,8-dihydro-4*H*-chromen-5(6*H*)-one derivatives in good to excellent yields, it was attempted to synthesize 2-amino-3-cyano-pyrano [4,3-*b*] pyran-5(4*H*)-ones (**9a–9f**) and 2-amino-3-cyano-pyrano [3,2-*c*] pyridine-6(5*H*)-ones (**10a–10f**) under similar conditions. For this purpose, were placed the 4-hydroxy-6-methylpyrone **9** and 4-hydroxy-6-methylpyridone **10** compounds with dimedone in the same conditions (Scheme 5). In this case, the use of C–H-activated acids, as cyclic 1,3-dicarbonyls, required longer reaction times compared to 4-hydroxycoumarin (**1**) under similar reaction conditions (Table 4).

A plausible mechanism for the preparation of 2-amino-4*H*-chromene derivatives catalyzed by MNPs@Cu is shown in Scheme 6. Initially, malononitrile **2** and aldehyde **3** as the reactant components react with each other *via* a Knoevenagel condensation reaction. MNPs@Cu coordinate with the C≡N and C=O functional groups of aldehyde and malononitrile, respectively, thus increasing the activity of these groups. After passing the transition state **11** and a dewatering step, it creates alkylidene malononitrile **12**. Then, enolate of the C–H-activated acid (**1**, **5**, **7**, and **8**) by Michael addition is added to this activated intermediate by MNPs@Cu and the polar transition state **13** is created. This compound is unstable and is converted by tautomerization to the intermediate **14**, which generates products **4**, **6**, **9**, and **10** though cyclization reaction.

To evaluate the recycled MNPs@Cu performance, after five periods of recycling, this nanocatalyst was reused in the three-component reaction of 4-hydroxycoumarin, 4-chlorobenzaldehyde, and malononitrile as a model reaction (Fig. 7). Negligible changes in the reaction efficiency represent the nanocatalyst performance and expression of the cost-effectiveness of our method.

As clearly shown Table 5, although all the mentioned homogeneous and heterogeneous catalysts are suitable for synthesizing 2-amino-4*H*-chromene derivatives, most of them suffer from one or more disadvantages, like the usage of toxic organic solvents, long reaction times, high catalyst loading, and

tedious work-up procedures. It is clear that a suitable methodology in terms of the compatibility with the environment, product yield, using green solvent, and reaction time in the existence of a low catalyst loading of MNPs@Cu comparing with several other catalysts has been developed.

Conclusion

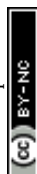
In general, we described an easy, effective, and ecofriendly method for a wide range of biologically and pharmacologically interesting functionalized 2-amino-4*H*-chromenes one in the existence of an environmental friendly and reusable nanocatalyst (MNPs@Cu) *via* one-pot Knoevenagel condensation of C–H-activated acids, aldehydes, and malononitrile under solvent-free conditions. This method suggests various privileges including the usage of a green catalyst, lower loading of the catalyst, no organic solvent, easy work-up, and good to high yields.

Conflicts of interest

There are no conflicts to declare.

References

- 1 F. Al-Assar, K. N. Zelenin, E. E. Lesiovskaya, I. P. Bezhan and B. A. Chakchir, *Pharm. Chem. J.*, 2002, **36**, 598.
- 2 R. P. Jain and J. C. Vederas, *Bioorg. Med. Chem. Lett.*, 2004, **14**, 3655.
- 3 S. J. Mohr, M. A. Chirigos, F. S. Fuhrman and J. W. Pryor, *Cancer Res.*, 1975, **35**, 3750.
- 4 M. Rueping, E. Sugiono and E. Merino, *Chem.–Eur. J.*, 2008, **14**, 6329.
- 5 L. Hanna and A. Calanolide, *Bulletin of Experimental Treatments for AIDS (BETA)*, 1998, vol. 12, p. 8.
- 6 M. T. Flavin, J. D. Rizzo, A. Khilevich, A. Kucherenko, A. K. Sheinkman, V. Vilaychack, L. Lin, W. Chen,



- E. M. Greenwood, T. Pengsuparp, J. M. Pezzuto, S. H. Hughes, T. M. Flavin, M. Cibulski, W. A. Boulanger, R. L. Shone and Z. Q. Xu, *J. Med. Chem.*, 1996, **39**, 1303.
- 7 D. O. Moon, K. C. Kim, C. Y. Jin, M. H. Han, C. Park, K. J. Lee, Y. M. Park, Y. H. Choi and G. Y. Kim, *Int. Immunopharmacol.*, 2007, **7**, 222.
- 8 T. Raj, R. K. Bhatia, A. Kapur, M. Sharma, A. K. Saxena and M. P. S. Ishar, *Eur. J. Med. Chem.*, 2010, **45**, 790.
- 9 L. Bonsignore, G. Loy, D. Secci and A. Calignano, *Eur. J. Med. Chem.*, 1993, **28**, 517.
- 10 A. G. Martinez and L. J. Marco, *Bioorg. Med. Chem. Lett.*, 1997, **7**, 3165.
- 11 L. L. Andreani and E. Lapi, *Bull. Chim. Farm.*, 1960, **99**, 583.
- 12 Y. L. Zhang, B. Z. Chen, K. Q. Zheng, M. L. Xu, X. H. Lei and X. B. Yaoxue, *Chem. Abstr.*, 1982, **96**, 135383e.
- 13 E. A. A. Hafez, M. H. Elnagdi, A. G. A. Elagamey and F. M. A. A. Eltaweel, *Heterocycles*, 1987, **26**, 903.
- 14 A. Shaabani, R. Ghadari, A. sarvary and A. H. Rezayan, *J. Org. Chem.*, 2009, **74**, 4372.
- 15 M. N. Elinson, A. S. Dorofeev, F. M. Miloserdov, A. I. Ilovaisky, S. K. Feducovich, P. A. Belyakov and G. I. Nikishina, *Adv. Synth. Catal.*, 2008, **350**, 591.
- 16 H. J. Bestmann, *Angew. Chem., Int. Ed. Engl.*, 1967, **6**(1), 104.
- 17 H. Mehrabi and H. Abusaidi, *J. Iran. Chem. Soc.*, 2010, **7**, 890.
- 18 A. Shaabani, S. Samadi, Z. Badri and A. Rahmati, *Catal. Lett.*, 2005, **104**, 39.
- 19 M. Khoobi, L. Ma'mani, F. Rezazadeh, Z. Zareie, A. Foroumadi, A. Ramazani and A. Shafiee, *J. Mol. Catal. A: Chem.*, 2012, **359**, 74.
- 20 L. Min, L. Ma and L. Hu, *Tetrahedron Lett.*, 2011, **52**, 2597.
- 21 X. Wang, Z. Zeng, D. Shi, X. Wei and Z. Zong, *Synth. Commun.*, 2004, **34**, 3021.
- 22 S. Jadhava, R. patila, D. Kumbhara, A. Patravaleb, D. Chandamb and M. Deshmukha, *Int. J. Pharm. Sci. Rev. Res.*, 2015, **35**, 75.
- 23 M. Kazemzad, A. A. Yuzbashi, S. Balalaie and M. Bararjanian, *Synth. React. Inorg., Met.-Org., Nano-Met. Chem.*, 2011, **41**, 1182.
- 24 J. K. Rajput and G. Kaur, *Catal. Sci. Technol.*, 2014, **4**, 142.
- 25 S. Balalaie, M. Bararjanian, A. M. Amani and B. Movassagh, *Synlett*, 2006, **2**, 263.
- 26 S. Gowravaram, K. Arundhathi, K. B. S. Sudhakar and J. S. Yadav, *Synth. Commun.*, 2009, **39**, 433.
- 27 M. Safaiee, M. A. Zolfigol, F. Afsharnadery and S. Bagheri, *RSC Adv.*, 2015, **5**, 102340.
- 28 K. Niknam and A. Piran, *Green Sustainable Chem.*, 2013, **3**, 1.
- 29 A. Montaghani and N. Montazeri, *Orient. J. Chem.*, 2014, **30**, 1361.
- 30 M. G. Dekamin, M. Eslami and A. Maleki, *Tetrahedron*, 2013, **69**, 1074.
- 31 H. Mehrabi and M. Kazemi-Mireki, *Chin. Chem. Lett.*, 2011, **22**, 1419.
- 32 S. Banerjee, A. Horn, H. Khatri and G. A. Sereda, *Tetrahedron Lett.*, 2011, **52**, 1878.
- 33 J. H. Clark, Catalysis for green chemistry, *Pure Appl. Chem.*, 2001, **73**, 103.
- 34 V. J. Mohanraj and Y. Chen, *Trop. J. Pharm. Res.*, 2006, **5**, 561.
- 35 D. Astruc, *Inorg. Chem.*, 2007, **46**, 1884.
- 36 L.-S. Zhong, J.-S. Hu, Z.-M. Cui, L.-J. Wan and W.-G. Song, *Chem. Mater.*, 2007, **19**, 4557.
- 37 M. Moreno-Manas and R. Pleixats, *Acc. Chem. Res.*, 2003, **36**, 638.
- 38 A. Yahyazadeh, E. Abbaspour-Gilandeh and M. Aghaei-Hashjin, *Catal. Lett.*, 2018, **148**, 1254.
- 39 F. Kamali and F. Shirini, *RSC Adv.*, 2017, **41**, 11778.
- 40 M. Sheykhan, A. Yahyazadeh and Z. Rahemizadeh, *RSC Adv.*, 2016, **6**, 34553.
- 41 X. Zheng, S. Luo, L. Zhang and J. P. Cheng, *Green Chem.*, 2009, **11**, 455.
- 42 W. Xiang-Shan, Z. Zhao-Sen, S. Da-Qing, W. Xian-Yong and Z. Zhi-Min, *Chin. J. Org. Chem.*, 2005, **25**, 1138.
- 43 D. K. Yi, S. S. Lee and J. Y. Ying, *Chem. Mater.*, 2006, **18**, 2459–2461.
- 44 J. Ramos, A. Millan and F. Palacio, *Polymer*, 2000, **41**, 8461.
- 45 Z. Ma, Y. Guan and H. Liu, *J. Magn. Magn. Mater.*, 2006, **301**, 469.
- 46 M. J. Khurana, B. Nand and P. Saluja, *Tetrahedron*, 2010, **66**, 5637.
- 47 W. Hong-juan, Z. jie and Z. Zhan-Hui, *Monatsh. Chem.*, 2010, **141**, 1107.
- 48 G. M. Ziarani, A. Badiei, M. Azizi and P. Zarabadi, *J. Chem. Eng.*, 2011, **30**, 59.
- 49 S. Gao, C. H. Tsai, C. Tseng and C. F. Yao, *Tetrahedron*, 2008, **64**, 9143.
- 50 R. Liangce, L. Xiaoyue, W. Haiying, S. Daqing, T. Shujiang and Z. Oiya, *Synth. Commun.*, 2006, **36**, 2363.
- 51 A. Patra and T. Mahapatra, *J. Chem. Res., Synop.*, 2010, **34**, 689.
- 52 S. Balalaie, M. Bararjanian, M. Sheikh-Ahmadi, S. Hekmat and P. Salehi, *Synth. Commun.*, 2007, **37**, 1097.
- 53 M. Poliakov and P. Licence, *Nature*, 2007, **450**, 810.
- 54 P. Tundo and P. T. Anastas, *Green Chemistry: Challenging Perspectives*, Oxford University Press, Oxford, UK, 2000, p. 79.
- 55 X. Fan, D. Feng, Y. Qu, X. Zhang, J. Wang, P. M. Loiseau, G. Andrei, R. Snoeck and E. Clercq, *Bioorg. Med. Chem. Lett.*, 2010, **20**, 809.
- 56 E. V. Stoyanov, I. C. Ivanov and D. Heber, *Molecules*, 2000, **5**, 19.
- 57 M. Z. Piao and K. Imafuku, *Tetrahedron Lett.*, 1997, **38**, 5301.
- 58 S. M. Baghbanian, *RSC Adv.*, 2014, **4**, 59397.
- 59 G. Brahmachari and B. Banerjee, *ACS Sustainable Chem. Eng.*, 2014, **2**, 411.
- 60 M. Hong and C. Cai, *J. Chem. Res., Synop.*, 2010, **34**, 568.
- 61 A. Saha, S. Payra and Su. Banerjee, *RSC Adv.*, 2015, **5**, 101664.
- 62 S. Zavar, *Arabian J. Chem.*, 2017, **10**, S67.
- 63 M. G. Dekamin, S. Z. Peyman, Z. Karimi, S. Javanshir, M. R. Naimi-Jamal and M. Barikani, *Int. J. Biol. Macromol.*, 2016, **87**, 172.
- 64 X.-S. Wang, D.-Q. Shi, S.-J. Tu and C.-S. Yao, *Synth. Commun.*, 2003, **33**, 119.

