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Ni-catalysed regioselective 1,2-diarylation of unactivated olefins by stabilizing Heck intermediates as pyridylsilyl-coordinated transient metallacycles†

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We report a Ni-catalysed diarylation of unactivated olefins in dimethylpyridylvinylsilane by intercepting Heck C(sp³)-NiX intermediates, derived from aryl halides, with arylzinc reagents. This approach utilizes a modifiable pyridylsilyl moiety as a coordinating group that plays a dual role of intercepting oxidative addition species to promote Heck carbometallation, and stabilizing the Heck C(sp³)-NiX intermediates as transient metallacycles to suppress β-hydride elimination, and facilitate transmetalation/reductive elimination. This method affords 1,2-diarylethylsilanes, which can be readily oxidized to 1,2-diarylethanol that occur as structural motifs in 3-aryl-3,4-dihydroisocoumarin and dihydrostilbenoid natural products.

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Introduction

Interception of Heck C(sp³)-[M] intermediates with carbon nucleophiles by cross-coupling is an attractive approach to construct simultaneously two carbon-carbon (C-C) bonds across an olefin in one synthetic step.¹ This olefin dicarbofunctionalization process integrates the Heck reaction and cross-coupling in one synthetic platform (Scheme 1, path A) to rapidly build molecular complexity, an endeavour that would otherwise require a multi-step process if pursued through traditional synthetic disconnection strategies. However, execution of such a process on an unactivated olefin remains formidably challenging especially with transition metals such as Pd. The difficulty arises due to the sheer requirement to overcome two of the most fundamental Pd-catalysed processes as side reactions – direct cross-coupling between an organohalide and an organometallic reagent prior to olefin insertion (Scheme 1, path B), and the Heck reaction by β-hydride (β-H) elimination from the C(sp³)-[M] intermediates formed after carbometallation to an olefin (Scheme 1, path C).

Prior reports have exploited two strategies to overcome these problems and enable the transition metal-catalysed three-component olefin dicarbofunctionalization reactions – (1) the use of dienes and styrenes as substrates that would stabilize the

Heck C(sp³)-[M] intermediates by intrinsic π-allyl and π-benzyl formation,² and (2) the generation and addition of alkyl radicals to olefins.³ In simple olefins that lack the means to stabilize the Heck C(sp³)-[M] intermediates,⁴ reactions generally follow the Heck carbometallation/β-H elimination/M-H reinsertion cascade to furnish 1,1-dicarbofunctionalized products.⁵

Recently, we⁶ developed a (cod)₂Ni-catalysed three-component dicarbofunctionalization of olefins in which we strategically employed imines as a readily removable coordinating group to facilitate the reaction.⁷ The idea was to intercept the organonickel species 3 by a bidentate coordination mode, promote intramolecular Heck carbonylation onto the bound olefin and stabilize the resultant Heck C(sp³)-[Ni] intermediate 6 as a transient metallacycle (Scheme 2). We envisioned that

Path A: Olefin diarylation by combined Heck reaction and cross-coupling



Path B: leads to undesired coupling products by bypassing olefin insertion



Path C: leads to undesired Heck products after β-H elimination



Scheme 1 Two well-known transformations working as side reactions during regioselective olefin diarylation.

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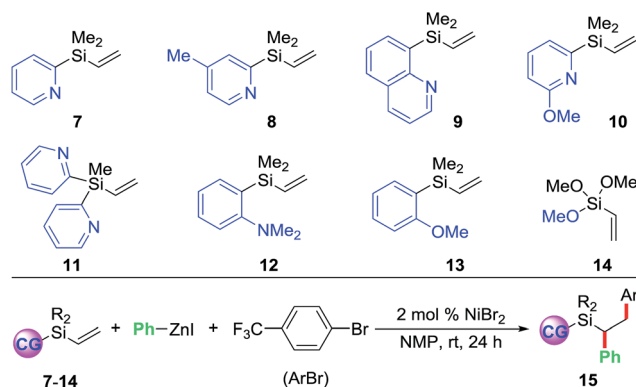
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this strategy would afford sufficient residence time for the Heck C(sp³)-[Ni] species **6** in order to promote the requisite transmetalation/reductive elimination steps, and furnish the expected product. We believe that this approach should operate regardless of the reaction proceeding *via* a Ni⁰/Ni^{II} or Ni^I/Ni^{III} catalytic cycle. Herein, we report diarylation of unactivated olefins in pyridylvinylsilanes that relies upon the same strategy

Results and discussion

Towards our long-term goal of olefin dicarbofunctionalization by cross-coupling,^{4a,k} we aspired to expand the scope of our three-component diarylation of olefins to silicon-based molecules. We believe that development of such a method could provide rapid access to complex differently substituted 1,2-diarylethylsilanes, which are difficult to synthesize.⁹ The 1,2-diarylethylsilanes could then be readily oxidized to 1,2-diarylethanol that occur as structural motifs in wide range of 3-aryl-3,4-dihydroisocoumarin and dihydrostilbenoid natural products.¹⁰ In this process, we examined the diarylation of pyridylvinylsilane **7** with 4-bromobenzotrifluoride and PhZnI (Table 1). Examination of various parameters revealed that the reaction furnished the expected product **15a** in best yield (79%) using 2 mol% NiBr₂ as a catalyst in NMP at room temperature (entry 1). We then investigated substituted pyridyl (**8** and **10**), 8-quinolinyl (**9**), dipyridyl (**11**), *N,N*-dimethylaniliny (12), *o*-

Table 1 Optimization of reaction conditions^a

Entry	Reaction condition deviation	Vinylsilanes	Product	Yields of 15 (%)
1	None	7	15a	79 (72)
2	None	8	15b	51
3	None	9	15c	56
4	None	10	15d	0
5	None	11	15e	0
6	None	12	15f	0
7	None	13	15g	0
8	None	14	15h	0
9	15 h	7	15a	62
10	(cod) ₂ Ni instead of NiBr ₂	7	15a	35
11	(Ph ₃ P) ₄ Ni instead of NiBr ₂	7	15a	30
12	DMF or DMA instead of NMP	7	15a	40–44
13	DMSO, dioxane or MeCN instead of NMP	7	15a	10–15
14	Benzene of THF instead of NMP	7	15a	<5
15	CuI or FeCl ₃ instead of NiBr ₂	7	15a	0
16	Pd(OAc) ₂ or Co(OAc) ₂ instead of NiBr ₂	7	15a	0 ^b

^a Yields were determined by ¹H NMR using pyrene as an internal standard. Value in parenthesis is the isolated yield from a 0.5 mmol scale reaction.

^b Heck product was formed in >90% NMR yield with Co(OAc)₂.

Table 2 Reaction scope with aryl halides^a

7

16-32

16, X = Br, 75%^b

17, X = I, 52%^b

18, X = I, 49%^b

19, X = Br, 70%^b

20, X = I, 42%^c

21, X = Br, 53%^c

22, X = Br, 57%^c

23, X = Br, 65%^b

24, X = Br, 55%^b

25, X = I, 61%^b

26, X = I, 44%^b

27, X = Br, 57%^b

28, X = Br, 55%^d

29, X = Br, 56%^b

30, X = Br, 50%^b

31, X = Br, 52%^c

32, X = Br, 54%^c

^a Values are isolated yields from 0.5 mmol scale reactions. ^b 50 °C. ^c Room temperature. ^d 80 °C.

anisoyl (13) and trimethoxysilyl (14) as coordinating groups where the nitrogen or oxygen atoms could coordinate to Heck C(sp³)-[Ni] intermediates to generate 4-, 5- or 6-membered transient metallacycles (entries 2–8). Among them, only the 4-methylpyridyl (8) and 8-quinolinyl (9) afforded the expected products **15b–c** in good yields (entries 2–3).¹¹ Running the reaction for a shorter time (15 h) formed the product in slightly lower yield (entry 9). Other Ni-catalysts such as Ni(cod)₂ and Ni(PPh₃)₄ also generated the product in lower yields (entries 10 and 11). The lower yields with Ni(cod)₂ and Ni(PPh₃)₄ could be rationalized based on difficulty to displace strongly coordinating ligands such as 1,5-cyclooctadiene (cod) and Ph₃P by the substrate imine or olefin. Using DMF or DMA as a solvent furnished product **15a** in moderate yields (entry 12). Other solvents such as DMSO, dioxane, MeCN, benzene and THF formed the product **15a** in <15% yield (entries 13 and 14). The use of a Ni-catalyst was found to be critical for this transformation as other



Scheme 3 Transformation of silyl group to alcohol.

catalysts based on Pd, Co, Fe and Cu did not catalyse the reaction (entries 15 and 16).

We note that the reaction does not form products arising from the Hiyama coupling¹² despite the fact pyridylvinylsilanes are known to undergo such a reaction.¹³ The current reaction also proceeds with absolute regioselectivity with terminal arylation during Heck carbometallation, a result that is in sharp contrast to the undirected Ni-catalysed Heck reactions that proceed with excellent selectivity for internal arylation.¹⁴ The regioselectivity of the products was confirmed by oxidation of 1,2-diarylethylsilanes **19** and **23** (Table 2) to the corresponding known 1,2-diarylethanol **45** and **46** (*vide infra*, Scheme 3), respectively.¹⁵ We hypothesize that the remarkable terminal selectivity for Heck carbometallation in our reaction could be a function of both the coordinating effect of the pyridyl group that strives for forming stable and tight five-membered metal-lacycles after carbometallation on the vinyl group, and the β -cation stabilizing effect of Si due to hyperconjugation.¹⁶ Prior work on Pd- and Fe-catalysed, directed Heck reaction of dimethyl(2-pyridyl)vinylsilane has also shown that the reaction proceeds with high regioselectivity for terminal arylation.¹⁷

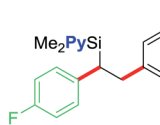
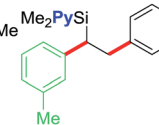
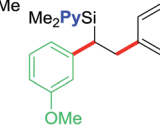
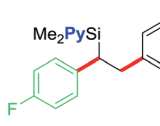
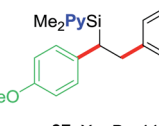


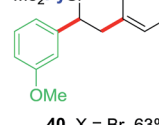
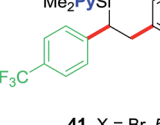
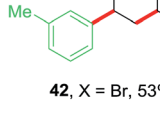
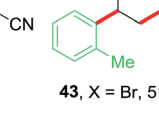
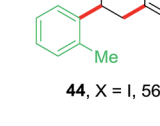
After optimizing the reaction conditions, we examined the scope of the current olefin diarylation reaction (Table 2). In general, reactions of electron-deficient aryl halides reacted at room temperature (**20–22**, **29** and **31**) and those with electron-rich substituents required 50 °C (**23–28** and **30**). In addition, reactions typically required 2–5 mol% NiBr₂. A wide variety of electron-rich, neutral and electron-deficient aryl iodides can be utilized as coupling partners along with the pyridylvinylsilane **7**, which furnishes variously substituted 1,2-diarylethylsilanes **16–28** in good yields.¹⁸ The reaction tolerates a variety of functional groups such as Me, OMe, OTBS, CF₃, F, Cl, COMe, CO₂Me and CN. The reaction also works well with aryl halides containing *ortho*-substituents (**16**, **24**, **28** and **32**) including sterically hindered isopropyl group (**25**).

We further examined the scope of the reaction with respect to arylzinc reagents (Table 3). Electron-rich, deficient and neutral arylzinc reagents containing Me, OMe, F and CF₃ can be used as coupling partners¹⁹ along with aryl halides bearing various functional groups like Me, OMe, CF₃, Cl, COMe, CO₂Me and CN, which afford the 1,2-diarylated products in good yields (**33–44**). The reaction also tolerates *ortho*-substituted arylzinc reagents (**43** and **44**).

The 1,2-diarylethylsilane products can be readily transformed into 1,2-diarylethanol by oxidation (Scheme 3). For example, we show that the products **19** and **23** can be efficiently oxidized to 2-(4-fluorophenyl)-1-phenylethanol (**45**) and 1-phenyl-2-(*p*-tolyl)ethanol (**46**), respectively, in excellent yields by reacting them with H₂O₂ in the presence of KF and KHCO₃.²⁰



Table 3 Reaction scope with arylzinc reagents^a

$\text{Py-Si(Me)}_2 + \text{Ar-ZnI} + \text{Ar'-X} \xrightarrow[\text{NMP, 50 } ^\circ\text{C, 24 h}]{2-5 \text{ mol \% NiBr}_2} \text{Py-Si(Me)}_2(\text{Ar})(\text{Ar'})$		
7		33-44
 33, X = Br, 55%	 34, X = Br, 50%	 35, X = Br, 60%
 36, X = Br, 52% ^b	 37, X = Br, 44% ^c	 38, X = Br, 52%
 39, X = I, 51%	 40, X = Br, 63%	 41, X = Br, 50%
 42, X = Br, 53% ^d	 43, X = Br, 51%	 44, X = I, 56%

^a Values are isolated yields from 0.5 mmol scale reactions. 2 mol% NiBr₂ for 38 and 42; 3 mol% NiBr₂ for 34, 35, 37, 39 and 40; 5 mol% NiBr₂ for 33, 36, 41, 43 and 44. ^b 40 °C. ^c 60 °C. ^d Room temperature, 48 h.

We conducted further studies in order to probe the role of the pyridyl group (Scheme 4). We performed the reaction of phenylvinylsilane **47** with 4-bromobenzotrifluoride and PhZnI under the standard conditions. Despite Ni being a good catalyst for the Heck reaction,^{14,21} phenylvinylsilane **47** did not afford any Heck or the dicarbofunctionalized product. Only the direct cross-coupling product was formed in 81% yield. This result indicates that the pyridyl group in pyridylvinylsilane **7** is indeed required for both the Heck carbometallation of Ar[Ni] on the vinyl group, and stabilizing the Heck C(sp³)-[Ni] intermediates as five-membered metallacycles prior to the delivery of the desired 1,2-diarylated products *via* transmetalation/reductive elimination processes.

We also conducted a radical clock experiment in order to determine if aryl radicals were formed during the reaction of aryl halides with the Ni-catalyst (Scheme 5). For this purpose, we



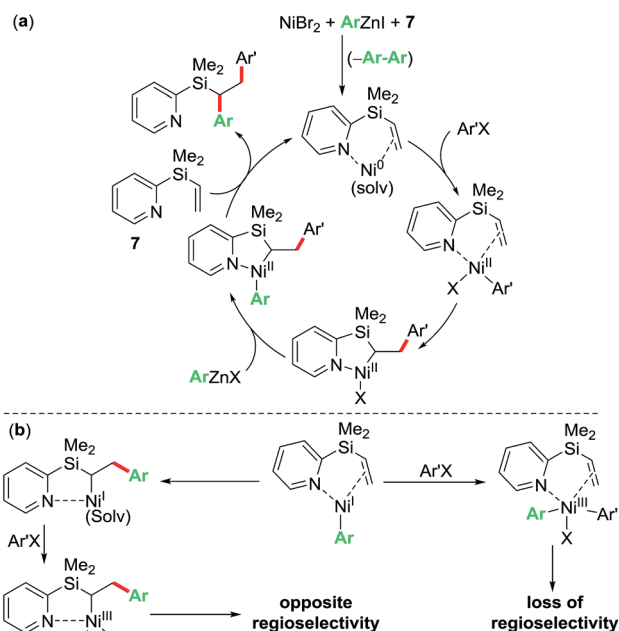
Scheme 4 Establishing the role of pyridylsilyl group.



Scheme 5 Radical clock experiment.

utilized 1-(but-3-enyl)-2-iodobenzene (**50**) as a radical probe. The aryl radical **52** generated from the radical probe **50** is known to undergo a fast radical cyclization with a k_{obs} of $5.0 \times 10^8 \text{ s}^{-1}$ to 1-methyldihydroindenyl radical **53**.²² When we reacted the radical clock **50** with the pyridylvinylsilane **7** and PhZnI using 5 mol% NiBr₂ as a catalyst, the diarylated product **51** was formed in 75% isolated yield. The cyclized products **54** and **55**, expected to arise if the aryl radical **52** was generated, were not observed. This experiment clearly indicates that oxidative addition of ArX to the Ni-catalyst proceeds without the generation of aryl free radicals.

Based on the control and radical clock experiments, and the observation of complete regioselectivity, we propose that the current diarylation reaction proceeds *via* a Ni(0)/Ni(II) catalytic cycle as outlined in Scheme 6a. Herein, the reaction proceeds *via* initial oxidative addition of ArX to Ni(0) bound to pyridylvinylsilane **7** followed by migratory insertion, transmetalation and reductive elimination steps to furnish the 1,2-diarylated products. An alternative non-radical Ni(I)/Ni(III) catalytic cycle as outlined in Scheme 6b proceeding *via* initial transmetalation of PhZnI²³ followed by either oxidative



Scheme 6 Proposed catalytic cycle.



addition/reductive elimination or migratory insertion/oxidative addition/reductive elimination sequence can be readily discounted because such a process would lead to products where the regioselectivity would be either lost or opposite to the one observed experimentally.

Conclusions

In summary, we have developed a Ni-catalysed diarylation of unactivated olefins in pyridylvinylsilanes by cross-coupling using aryl halides and arylzinc reagents as carbon sources. The reaction affords 1,2-diarylethylsilanes, which can be readily converted to 1,2-diarylethanol that are structural motifs in natural products and biologically important molecules. The reaction also tolerates various functional groups that are synthetically important. Preliminary studies indicate that the reaction proceeds *via* a Ni(0)/Ni(II) catalytic cycle.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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