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# Ultrasonic-template technology inducing and regulating cationic microblocks in CPAM: characterization, mechanism and sludge flocculation performance†

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In this study, the ultrasonic-template polymerization technique (UTPT) was used to generate and regulate the distribution of cationic microblocks in a polymer. The ultrasonic-template copolymer (TPAD-U) of acrylamide (AM) and methacryloxyethyl trimethyl ammonium chloride (DMC) with a novel cationic microblock structure was successfully synthesized through UTPT using sodium polymethacrylate (PMAA) as the template. Fourier transform infrared spectroscopy (FT-IR), <sup>1</sup>H (<sup>13</sup>C) nuclear magnetic resonance spectroscopy (<sup>1</sup>H (<sup>13</sup>C) NMR), scanning electron microscopy (SEM) and thermogravimetric analysis (TGA) were employed to characterize the properties of the polymers. The results showed that evident cationic microblocks formed in TPAD-U. Moreover, the template polymerization mechanism and reaction kinetics were analyzed, and the results showed that the I (ZIP) mechanism and free radical termination were assigned to template copolymerization. The I (ZIP) template mechanism convincingly indicated the formation of the cationic microblocks. The sludge dewatering results demonstrated that TPAD-U showed a better sludge flocculation performance than flocculants prepared by the non-template polymerization technique. During the sludge flocculation process, the cationic microblocks in TPAD-U greatly enhanced the effects of charge neutralization and bridging, which contributed much to a prominent flocculation performance.

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## 1. Introduction

As a result of the accelerating rate of urbanization and massive increase in human population size during the past few decades, a large amount of domestic sewage wastewater was produced, which posed a serious threat to the aquatic environment.<sup>1–4</sup> In order to meet the requirements of the current environmental legislation and more stringent disposal regulations, sewage wastewater was treated by the activated sludge process in which huge volumes of excess sludge were generated.<sup>5–8</sup> In China, the present annual sludge production has reached almost  $3 \times 10^7$  tons with a high moisture content of approximately 80%.<sup>9</sup>

Therefore, sludge dewatering is of paramount importance in the process of sludge treatment prior to disposal as it reduces the sludge volume and, consequently, affects the subsequent treatment and disposal operations.<sup>10</sup> However, high water content, presence of organic components which are mainly bacterial cells and extracellular polymeric substances (EPS), colloidal and supracolloidal range particles in the sewage sludge and biological gel structure properties of sludge lead to difficulties in dewatering even at high pressure.<sup>11,12</sup> Nevertheless, sludge dewatering remained the most expensive and most poorly understood wastewater treatment process.<sup>13</sup>

At present, cationic polyacrylamide (CPAM) has obtained more and more attention because the positive charged CPAM showed a high efficiency in the sludge colloids destabilization through charge neutralization, and the destabilized sludge colloids subsequently agglomerated to form large flocs under the effect of bridging, hence the large flocs could be more easily separated from the water.<sup>14,15</sup> Meanwhile, various methods have been used to enhance the efficiency of CPAM in sludge dewatering, including using different composite methods, changing different monomers and adding hydrophobic monomers.<sup>16,17</sup> However, the deficiency remains, for example, the disordered distribution of cationic microblocks in the polymer lead to the weakening of charge neutralization

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ability and the corresponding bridging ability.<sup>18</sup> The random distribution of cationic monomers can not maximize the whole function and efficiency in charge neutralization and bridging. By contrast, if the flocculant is characterized with the cationic microblocks, more cationic microblocks will be used effectively to increase the ability of charge neutralization and further enhance the ability of bridging.<sup>19</sup> As a result, the sludge dewatering performance is effectively improved. Thus, a new method should be proposed to control the distribution of cationic units.

Fortunately, the template polymerization technology can be employed to synthesize the flocculant with microblock structures in the study of CPAM.<sup>20</sup> The ability of charge neutralization can be improved by the novel cationic microblock structures because flocculants with these novel structures are more efficient in neutralizing counterions, which can promote flocculation and improve the flocculation efficiency.<sup>18</sup> Ultrasonic technology is no doubt recognized as one part of the emerging science and technology. Due to its easy operation, high efficiency, and environmental-friendly rapid processing, it is applied extensively in synthesis of materials,<sup>21,22</sup> disinfection,<sup>23</sup> and wastewater treatment.<sup>24</sup> High shear forces generated by acoustic cavitation can result in homogeneous distribution in aqueous solution,<sup>25</sup> so it may be applied to distribute and assemble cationic microblocks during the process of template polymerization. It was reported that ultrasonic-template technique has been employed to synthesize novel nanoparticles.<sup>26</sup> Therefore, ultrasonic-template polymerization technique shown in ESI Fig. S1† is proposed to synthesize CPAM with the novel cationic microblocks, and it is beneficial for overcoming the problems associated with traditional template technology. The polymer prepared by copolymerization of methacryloxyethyl trimethyl ammonium chloride (DMC) and acrylamide (AM) is named as CPAD which is regarded as one important and widely used flocculant in the family of CPAM. It will be more significant to design a CPAD with an evident cationic microblock structure through ultrasonic-template polymerization technique (UTPT), and this new method will give an insight into the domain of polymer science and the similar.

In this study, the ultrasonic-template technology was used to prepare TPAD-U with a cationic microblock structure and a rough surface morphology. In order to study the effect of ultrasonic-template polymerization technique on the structural characters of the products, many kinds of advanced analysis were used, including Fourier transform infrared spectroscopy (FT-IR), <sup>1</sup>H (<sup>13</sup>C) nuclear magnetic resonance spectroscopy (<sup>1</sup>H (<sup>13</sup>C) NMR), scanning electron microscopy (SEM) and thermogravimetric analysis (TGA). Besides, the ultrasonic-template polymerization mechanism was investigated though the reaction kinetics and association constant ( $K_M$ ). At last, the polymers were used to make a comparison to observe the impact of cationic microblocks on sludge flocculation, and the sludge flocculation performance was expressed by filter cake moisture content (FCMC), specific resistance to filtration (SRF), residual turbidity (RT), floc size and floc morphology.

## 2. Materials and methods

### 2.1. Materials

Reagents used in this experiment were of analytical grade except AM, DMC, and template sodium polymethacrylate (PMAA) which were of technical grade. The monomer AM was provided by Chongqing Lanjie Tap Water Company (Chongqing, China). DMC (78 wt% in water) and urea [CO(NH<sub>2</sub>)<sub>2</sub>] were purchased from Guangchuangjing Import and Export Co., Ltd. (Shanghai, China). The sodium polymethacrylate (PMAA) was obtained from Shandong Xintai Water Treatment (Zaozhuang, China) and its molecular weight is 5100. Initiator 2,2'-azobis [2-(2-imidazolin-2-yl)propane] dihydrochloride (VA-044) was purchased from Apotheker Chemical Reagent Co., Ltd. (Chendou, China). Deionized water was used throughout the experiment. Commercial flocculants CPAM-1 (copolymerization of AM and acryloyloxyethyl trimethyl ammonium chloride (DAC)) and CPAM-2 (copolymerization of AM and dimethyldiallyl ammonium chloride (DMDAAC)) were used for comparison in flocculation test, and their intrinsic viscosities were all of 9.25 dL g<sup>-1</sup>.

### 2.2. Preparation of copolymers

The ultrasonic-template polymerization technique was employed to synthesize the template copolymer TPAD-U. A certain amount of AM, DMC, urea, PMAA and deionized water was added in 100 mL quartz jar. The molar ratio of PMAA and DMC kept 1 : 1. Then, the pH of the reaction solution was adjusted to 4 by 0.5 mol L<sup>-1</sup> HCl or NaOH. An ultrasonicator with a frequency of 45 kHz (KQ 2200E, Kunshan ultrasonic instrument Co., LTD, China) was used as an ultrasonic source. In order to form a homogeneous solution, the mixture was

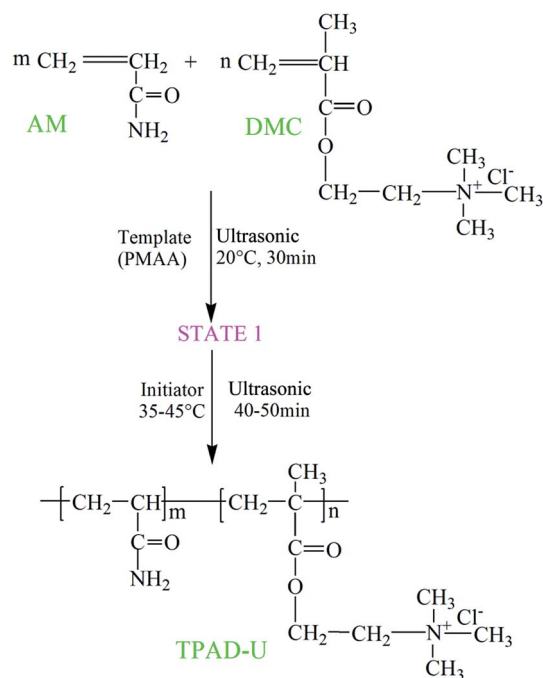


Fig. 1 The proposed reaction route for synthesis of TPAD-U.



sonicated at 45 kHz and 20 °C for 30 min. Subsequently, the purified nitrogen was bubbled into the aqueous solution at low temperature for about 30 min to remove the oxygen. The quartz jar was sealed immediately after addition of a predetermined dose of initiator VA0-44 to the mixture, and then continuously sonicated at 35–45 °C for 40–50 min. When the reaction was over, the copolymer was aged for 2 h at room temperature to increase the polymerization degree. The copolymer was dissolved in water, and the pH was adjusted to a value of less than 2 and then purified by acetone and ethanol. Finally, the flocculant was dried in a vacuum oven at 60 °C. The preparation of CPAD was similar to that of TPAD-U except that no ultrasonic and template was used. Besides, the polyacrylamide (PAM) prepared in laboratory was used for contrastive analysis of characterization. The proposed reaction route for TPAD-U was outlined in Fig. 1.

### 2.3. Characteristics of copolymers

The FT-IR spectra of the products were recorded using KBr pellets on a 550 Series II infrared spectrometer (Mettler Toledo Instruments Co., Ltd., Switzerland) in the range between 500 and 4000  $\text{cm}^{-1}$ . The  $^1\text{H}$  ( $^{13}\text{C}$ ) NMR of copolymers were obtained in  $\text{D}_2\text{O}$  with an Avance 500 nuclear magnetic resonance spectrometer (Bruker Company, Ettlingen, Germany). Meanwhile, the SEM analysis was conducted on MIRA 3 LMU SEM system (TES-CAN Company, Czech Republic). Thermogravimetric (TGA) and differential scanning calorimetry (DSC) were carried out at a heating rate of 10 °C  $\text{min}^{-1}$  under a nitrogen flow of 20  $\text{mL min}^{-1}$  from 20 to 600 °C on a DTG-60H synchronal thermal analyzer (Shimadzu, Kyoto, Japan).

### 2.4. Determination of $K_M$ and reaction kinetics

The ultrasonic-initiated template polymerization mechanism was discussed and determined through the investigation of association constant ( $K_M$ ) and reaction kinetics. The association constant ( $K_M$ ) between DMC and PMAA and the polymerization rate ( $R_p$ ) were examined through bag filter method and dilatometer method, respectively.<sup>27,28</sup> The details about the two methods were listed in ESI Text S1.†

### 2.5. Dewatering experiment

In this study, the polymers used for structure properties characterization and flocculation tests were displayed in Table 1. Raw sludge used in this experiment was collected from Jiguanshi Wastewater Treatment Plant (Chongqing, China) with a cyclic-activated sludge system. The sludge collected from this plant was characterized by a moisture content of 97.9%, a mass density of 0.997  $\text{kg L}^{-1}$ , a zeta potential of  $-22.3$  mV, a VSS/TSS of 0.75, and an initial sludge pH value of 7.2. The sludge dewatering experiments could be conducted on a buchner funnel or a Jar-test apparatus, and in this study, a program-controlled Jar-test apparatus (ZR4-6, Zhongrun Water Industry Technology Development Co. Ltd., China) was used to proceed the sludge dewatering experiments at room temperature. The sludge flocculation performance of the flocculants was expressed in terms of FCMC, RT and floc properties. Each measurement was

Table 1 The details of flocculants used in the characterization and flocculation tests

Flocculants <sup>a</sup>	Cationic degree	Intrinsic viscosity ( $\text{dL g}^{-1}$ )	Conversion rate (%)	Synthetic method
TPAD-U	30.0	9.19	99.7	UTPT
CPAD	30.0	9.23	99.5	N-UTPT
PAM	—	9.25	99.9	N-UTPT
CPAM-1	30.0	9.25	—	—
CPAM-2	30.0	9.25	—	—

<sup>a</sup> TPAD-U: copolymer of DMC and AM by ultrasonic-template technique (UTPT); CPAD: copolymer of DMC and AM by non ultrasonic-template technique (N-UTPT); PAM: homopolymer of AM by N-UTPT; CPAM-1: commercial copolymer of DAC and AM; CPAM-2: commercial copolymer of DMDAAC and AM.

performed in triplicate and the final results were the average of three runs. In sludge dewatering experiments, a certain amount of flocculant was added to 500 mL sludge in a glass beaker. The sludge was rapid stirred at 200 rpm for 30 s, followed by a slow stirring at 50 rpm for 10 min.<sup>29</sup> Flocs formed and settled for 30 min. The zeta potential of the supernatant and the turbidity of the supernatant collected from 2 cm under the surface were conducted on a Zetasizer Nano ZS90 (Malvern Instruments Ltd., Malvern, UK) and a 2100P turbidity meter (HACH, Loveland, CO), respectively. After a rapid agitation, the floc size was measured by a laser diffraction instrument (Mastersizer 2000, Malvern, U.K.). The filterability of the sludge was measured by FCMC and SRF which were illustrated in ESI Text S2.† Besides, a camera (Millet five, China) was used to observe the flocs morphology of conditioned sludge.

## 3. Results and discussion

### 3.1. The association constant ( $K_M$ ) and template polymerization mechanism

The  $K_M$  played an important role in the template copolymer system because the adhesion strength between cationic monomer and template could be intuitively reflected from  $K_M$ , and thus the type and mechanism of template copolymerization could be determined. In this study, a  $K_M$  value of 12.26 was obtained at  $n_{\text{DMC}} : n_{\text{PMAA}} = 1$ , which indicated that approximately three quarters of the DMC was pre-absorbed and anchored on the polymer molecular chain of template PMAA to form precursor of microblock structure under the electrostatic force, and the precursor was homopolymerized by ultrasonic-initiated radiation to breed a microblock segment. Thus, the template polymerization mechanism ultrasonic-initiated belonged to (I) ZIP mechanism rather than (II) pick-up exhibited in Fig. 2.<sup>30</sup> Besides, it provided a powerful support for the formation of the cationic microblock structure in TPAD-U.

### 3.2. Reaction kinetics

**3.2.1. Effect of pH.** Fig. 3(a) and (b) displayed the impact of pH on conversion and  $R_p$  at  $n_{\text{DMC}} : n_{\text{PMAA}} = 1 : 1$ ,  $n_{\text{DMC}} : n_{\text{AM}} = 1 : 3$ , total monomer concentration ( $C_M$ ) = 4.0  $\text{mol L}^{-1}$  and



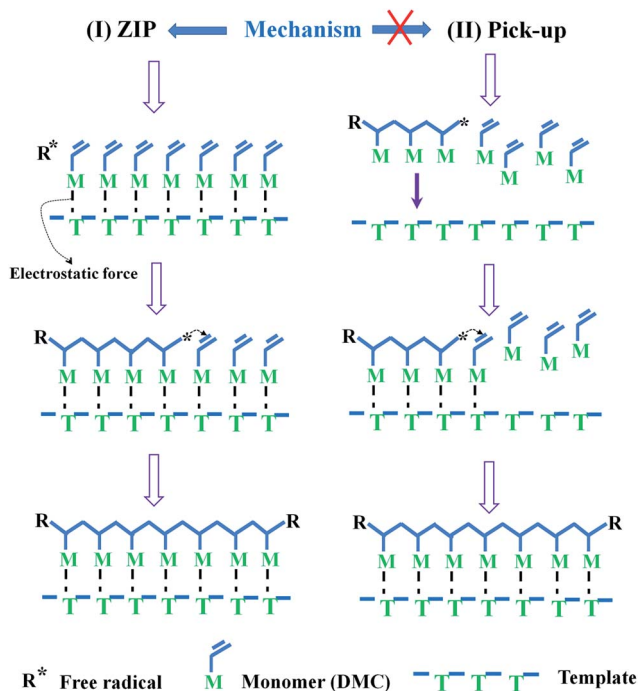


Fig. 2 The mechanism of template copolymerization.

initiator concentration ( $C_I$ ) =  $8.0 \times 10^{-4}$  mol L<sup>-1</sup>. It was clear that both conversion and  $R_p$  initially increased from pH = 2 to pH = 4 and reached their respective topic value at pH = 4, and then declined rapidly with the increasing of pH form 4 to 10. The low pH value resulted in imidization reaction of AM and a decrease of template dissociation degree, thus weakening the template effect. Meanwhile, when the pH value was too high, the hydroxyl methylation reaction occurred, which was adverse to template copolymerization. Therefore, the optimal range of pH was from 4 to 6.

**3.2.2. Effect of total monomer concentration ( $C_M$ ).** As shown in Fig. 3(c) and (d), the effect of  $C_M$  on conversion and  $R_p$  was investigated at  $n_{DMC} : n_{PMAA} = 1 : 1$ ,  $n_{DMC} : n_{AM} = 1 : 3$ , pH = 4 and initiator concentration ( $C_I$ ) =  $8.0 \times 10^{-4}$  mol L<sup>-1</sup>. Obviously, the conversion and  $R_p$  were positively associated with the  $C_M$ , and the higher the  $C_M$  was, the larger the value of conversion and  $R_p$  became. In Fig. 3(d), the slop of the fitting straight line was 1.698, which meant that a relationship of  $R_p \propto C_M^{1.698}$  was between  $C_M$  and  $R_p$ . The power value (1.698) of  $C_M$  was more than 1, this phenomenon was caused by the product precipitation in the copolymerization process, and it was common in free radical polymerization reaction.<sup>31</sup>

**3.2.3. Effect of mole ratio of PMAA to DMC.** In this part, the effect of mole ratio of PMAA to DMC on conversion and  $R_p$  was examined, and the results was illustrated in Fig. 3(e) and (f). The reaction conditions were set at  $n_{DMC} : n_{AM} = 1 : 3$ , pH = 4, total monomer concentration ( $C_M$ ) = 4.0 mol L<sup>-1</sup> and initiator concentration ( $C_I$ ) =  $8.0 \times 10^{-4}$  mol L<sup>-1</sup>. With the increase of mole ratio of PMAA to DMC from 0.25 : 1 to 1 : 1, the value of conversion and  $R_p$  all reached the maximum, and then displayed a decrease after  $n_{PMAA} : n_{DMC} > 1 : 1$ . The variation trend of conversion and  $R_p$  as a function of  $n_{PMAA} : n_{DMC}$  was accorded with

the typical I (ZIP) template polymerization mechanism.<sup>32</sup> Significantly, the template PMAA acted as a role of assembling and distributing cationic monomer. The template would be fully filled by DMC at the  $n_{PMAA} : n_{DMC} < 1$ , and therefore the conversion and  $R_p$  all displayed an increase with the increasing of  $n_{PMAA} : n_{DMC}$ . However, if the  $n_{PMAA} : n_{DMC}$  was more than 1 : 1, a part of template would be lack of DMC and the continuous distribution of DMC segments was destroyed, and thus leading to the termination of initiation reaction. As a result, the conversion and  $R_p$  decreased.

### 3.3. Characterization of flocculants

**3.3.1. FTIR spectral analysis.** As showed in Fig. 4, the FT-IR spectra of TPAD-U and CPAD were almost the same, but not quite identical. The FT-IR spectra of the two samples displayed the following similar peaks: -NH<sub>2</sub> stretching vibration was at 3433 cm<sup>-1</sup> in AM;<sup>33</sup> -CH<sub>3</sub> and -CH<sub>2</sub>- asymmetric stretching vibration was at 2934 and 2840 cm<sup>-1</sup>, respectively;<sup>34</sup> C=O stretching vibration was at 1662 cm<sup>-1</sup>; -CH<sub>2</sub>- bending vibration was at 1452 cm<sup>-1</sup> in -CH<sub>2</sub>-N<sup>+</sup>; C-O- stretching vibration was at 1161 cm<sup>-1</sup> in ester groups; and -N<sup>+</sup>-(CH<sub>3</sub>)<sub>3</sub> bending vibration was at 953 cm<sup>-1</sup> in DMC;<sup>19</sup> the characteristic chemical groups of AM and DMC were all observed in TPAD-U and CPAD, revealing that TPAD-U and CPAD were successfully synthesized through copolymerization of AM and DMC. Meanwhile, it also indicated that the cationic microblock structure almost had no impact on the chemical structure of the functional group in the prepared template copolymer except the slight shift in the peak area. For instance, the peak area of C=O at 1662 cm<sup>-1</sup> in TPAD-U was larger than that in CPAD.

**3.3.2. DSC/TGA analysis.** The thermal gravimetric curves of the polymers were shown in Fig. 5. It could be found that both CPAD and TPAD-U displayed three similar stages of the weight loss. In the first stage occurred in the range of 30–185 °C, mass loss was about 14.9 wt% for CPAD and 11.7 wt% for TPAD-U, which was associated with the evaporation of moisture in the polymers.<sup>35</sup> In the second stage occurred in the range of 185–350 °C, mass loss of 32.9 wt% for CPAD and 25.5 wt% for TPAD-U were observed, which was due to the imine reaction of the amide group and the thermal decomposition of methyl in the quaternary ammonium groups.<sup>36</sup> The final stage occurred beyond 350 °C, the mass loss of CPAD was about 38.1 wt% while that of TPAD-U was about 41.5 wt%, which was attributed to the carbonization of the copolymer.<sup>37</sup> Furthermore, two evident heat absorption peaks at 343.7 °C and 395.7 °C for TPAD-U were observed in the final stage, whereas just a single absorption peak at 372.9 °C appeared in CPAD. An explanation could be that the number of the absorption peak was related to the microstructure of the polymers. The random distribution of cationic monomer DMC in CPAD could not form microblock structure, thus resulting in a single absorption peak during the thermal gravimetric test. However, the template copolymer could generate evident microblock structure, namely, the DMC microblock structure and AM microblock structure, and thereby two sharper and larger heat absorption peaks were observed. The results were corresponding with the analysis shown in Section 3.1 and previous report.<sup>18,19</sup>





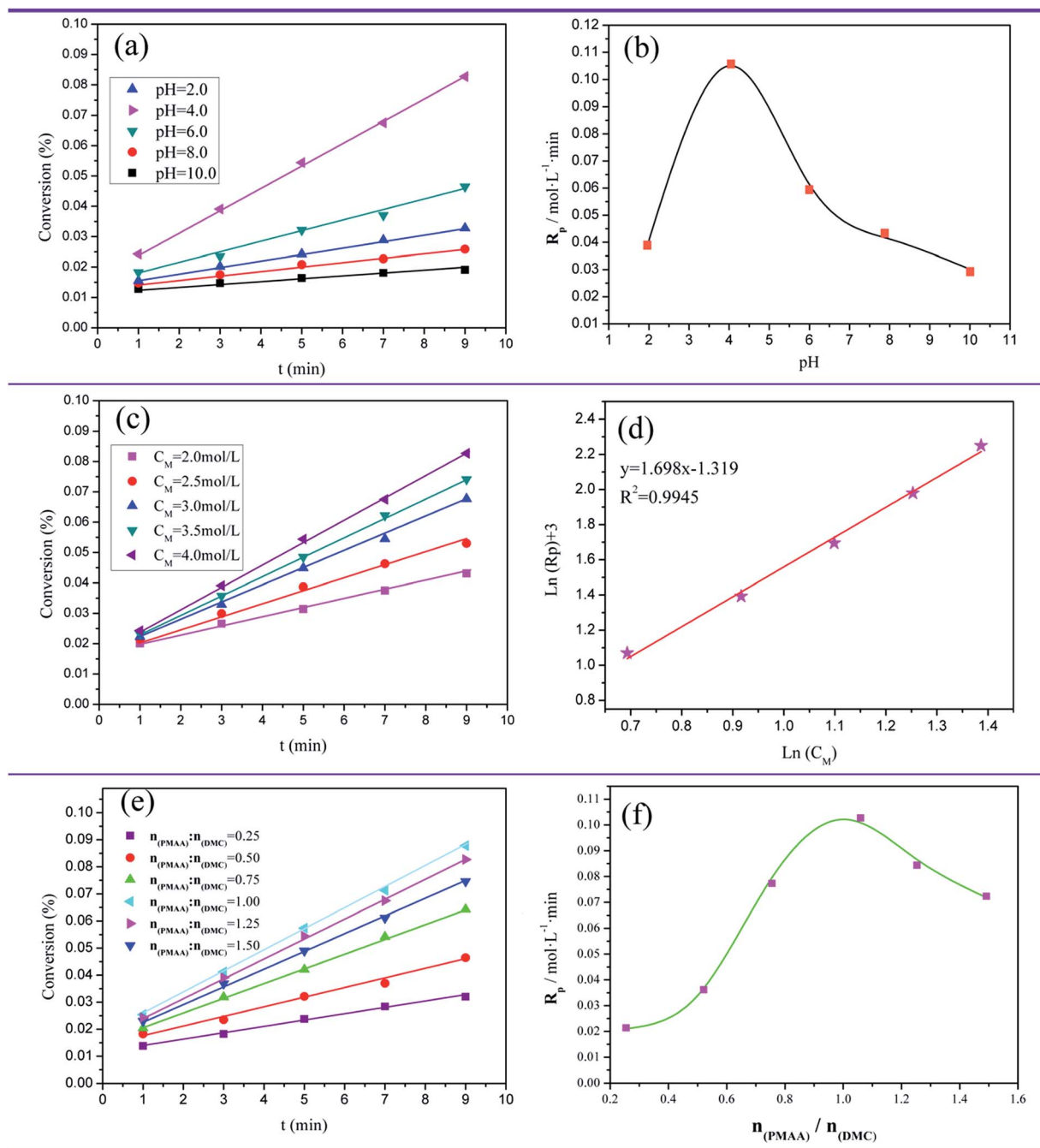


Fig. 3 Effect of pH on conversion (a) and  $R_p$  (b); effect of total monomer concentration ( $C_i$ ) on conversion (c) and  $R_p$  (d); and effect of  $n_{PMAA} : n_{DMC}$  on conversion (e) and  $R_p$  (f).

**3.3.3. <sup>1</sup>H NMR spectral analysis.** The <sup>1</sup>H NMR spectra shown in Fig. 6 further analyzed the microstructure of the PAM, CPAD and TPAD-U. Because the monomer DMC was grafted on the backbone of PAM, the proton chemical environment changed, and thus a slight proton peak shift of  $-CH_2-$  (a) at 1.629 ppm and  $-CH-$  (b) at 2.194 ppm were observed in both CPAD and TPAD-U compared with those of PAM. Moreover, CPAD and TPAD-U displayed many similarities, but several subtle differences were very significant to this study. CPAD and TPAD-U showed the following similar characteristic peaks: peak of

protons in methyl group  $-CH_3$  (c) was at  $\delta = 1.157$  ppm in DMC; peak of protons in the group of  $-O-CH_2-$  (d) was at  $\delta = 4.49$  ppm; peak of protons in the methylene group of  $(-CH_2-N^+)$  (e) was at  $\delta = 3.77$  ppm in DMC; the sharp peak of protons in the three equivalent methyl groups of  $-N^+(CH_3)_3$  (f) was at  $\delta = 3.209$  ppm in DMC; and the sharp peak of protons in solvent  $D_2O$  was at  $\delta = 4.79$  ppm. These results indicated that CPAD and TPAD-U were successfully synthesized by AM and DMC.<sup>38</sup> Compared with TPAD-U, four weak peaks were observed in the <sup>1</sup>H NMR spectra of CPAD. These weak peaks were at  $\delta = 3.19$  ppm,  $\delta = 3.44$  ppm,  $\delta =$



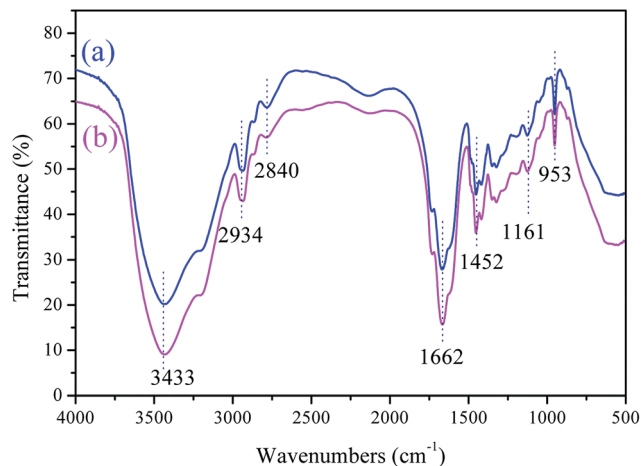


Fig. 4 FT-IR spectra of (a) CPAD and (b) TPAD-U.

3.65 ppm and  $\delta = 4.11$  ppm, respectively. However, those weak peaks did not show up in the  $^1\text{H}$  NMR spectra of TPAD-U. Although CPAD was the copolymerization of AM and DMC, the distribution of monomer DMC was disordered and uncontrollable. Due to the cationic microblock structure in TPAD-U, the steric conformation and the chemical environment for each group of  $-\text{CH}_2-\text{N}^+$  (e),  $-\text{O}-\text{CH}_2-$  (d), and  $-\text{N}^+(\text{CH}_3)_3$  (f) become identical, and the protons of each group were identical in principle.<sup>19,39</sup> Therefore, just only one resonance was observed for each group in TPAD-U rather than two or more extra weak peaks in CPAD. The result of  $^1\text{H}$  NMR spectra was demonstration that the microblock structures formed in TPAD-U.

**3.3.4.  $^{13}\text{C}$  NMR spectral analysis.** Parallel to the  $^1\text{H}$  NMR spectral, the  $^{13}\text{C}$  NMR spectra of TPAD-U, CPAD and PAM were represented in Fig. 7 to further characterize their properties. The peaks at  $\delta = 34.7$  ppm and  $\delta = 42.1$  ppm were attributed to the carbon resonances of methane ( $\text{CH}_2-$ ) (a) and methylene ( $-\text{CH}-$ ) (b) in TPAD-U, CPAD and PAM. Compared with the  $^{13}\text{C}$  NMR spectra of PAM, the characteristic peaks ( $^{13}\text{C}_c$ - $^{13}\text{C}_f$ ) of TPAD-U and CPAD were all observed at  $\delta = 17.9$  ppm,  $\delta = 45.1$  ppm,  $\delta = 179.7$  ppm,  $\delta = 176.0$  ppm,  $\delta = 58.9$  ppm,  $\delta =$

64.8 ppm,  $\delta = 53.5$  ppm, respectively. The above results further demonstrated that the copolymers of TPAD-U and CPAD successfully synthesized by AM and DMC. Moreover, the proportion of the monomer segment could be calculated by the area of the carbonyl carbon resonance peaks by using MestReNova software,<sup>40</sup> and the results were displayed in Table 2. The carbonyl carbon resonance peaks shown in the rectangle in Fig. 7 from left to right was from 1 to 6. If AM unit was abbreviated as A and DMC unit was abbreviated as D, the peaks of 1, 2 and 3 were represent AAA, AAD, DAD segments with A as the central, and 4, 5 and 6 were represent ADA, DDA, DDD segments with D as the central. Obviously, the microblock structures had an evident influence on the mole fraction of the triad sequence. Because of these microblocks in TPAD-U, the length of DDD and AAA became longer than those in TPAD without microblock structures. For example, the mole fractions of the AAA and DDD triad sequence in TPAD-U were 10.7% and 21.8%, respectively, whereas those in CPAD were 8.9% and 13.5%. Based on the above  $^{13}\text{C}$  NMR analytical results, it could be concluded that the evident DMC microblock structure was successfully prepared in TPAD-U rather than CPAD, which was in consistent with the analytical results of DSC/TGA and  $^1\text{H}$  NMR spectral.

**3.3.5. SEM of polymers.** The SEM images of the polymers (PAM, CPAD and TPAD-U) were displayed in Fig. 8 to investigate their amorphous morphology. It was more clear that there kind of polymers presented different surface morphology. As a homopolymer of AM, PAM showed a smooth and regular surface morphology.<sup>41</sup> However, with the grafting of DMC on AM, profound morphological change occurred. The introduction of DMC to AM chain destroyed the originally structure of AM, and thereby resulting in a porous structure with a rough surface area, which could be confirmed by fractal dimension. The linear correlation of the logarithm of projected area ( $A$ ) and the characteristic length ( $L$ ) was observed, and the average fractal dimensions could be calculated by using Image-Pro Plus 6.0 software.<sup>42</sup> The average fractal dimension of the polymers followed the order of PAM (1.218) < CPAD (1.426) < TPAD-U (1.631), which indicated that the surface of TPAD-U and CPAD was more rough and irregular than PAM. Besides, compared

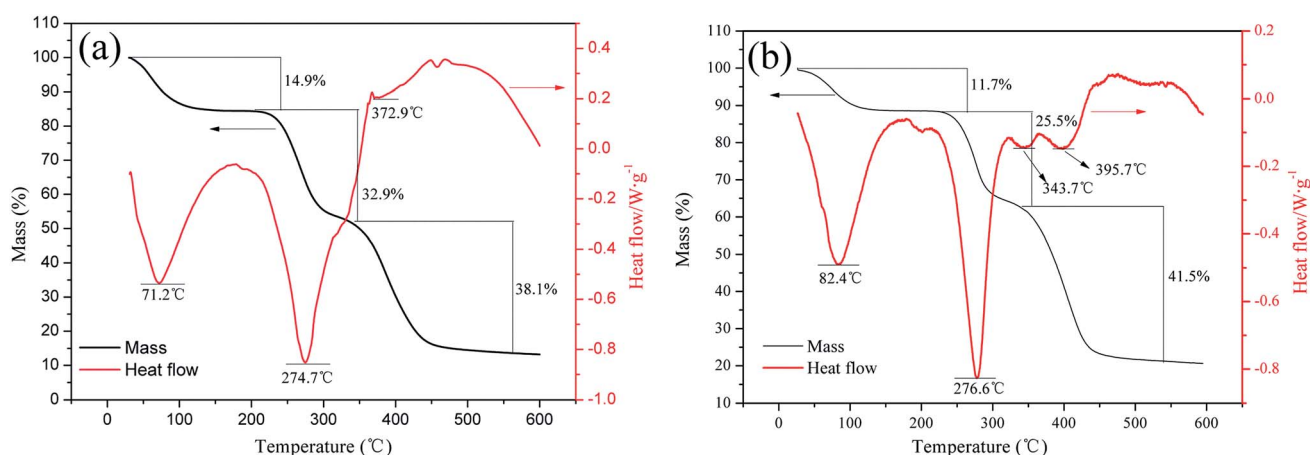


Fig. 5 Thermogravimetric curve of (a) CPAD and (b) TPAD-U.



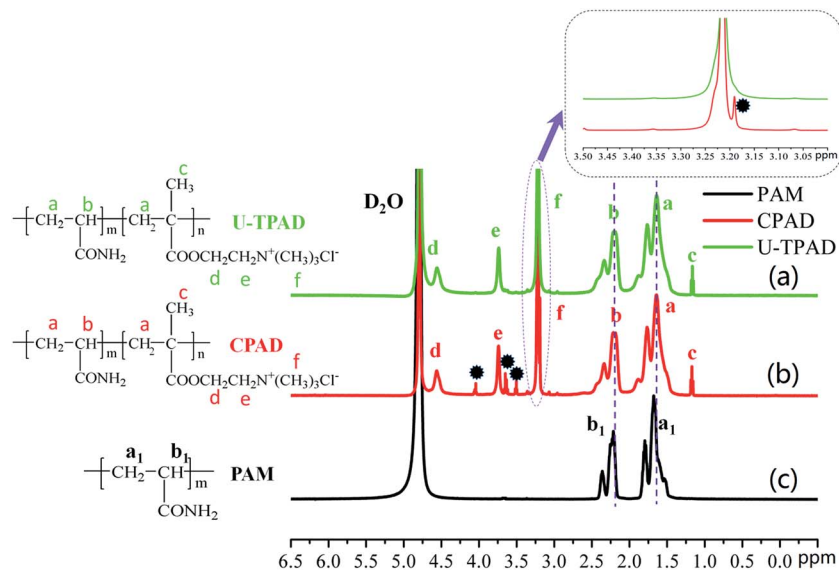


Fig. 6  $^1\text{H}$  NMR spectrum of (a) TPAD-U, (b) CPAD and (c) PAM.

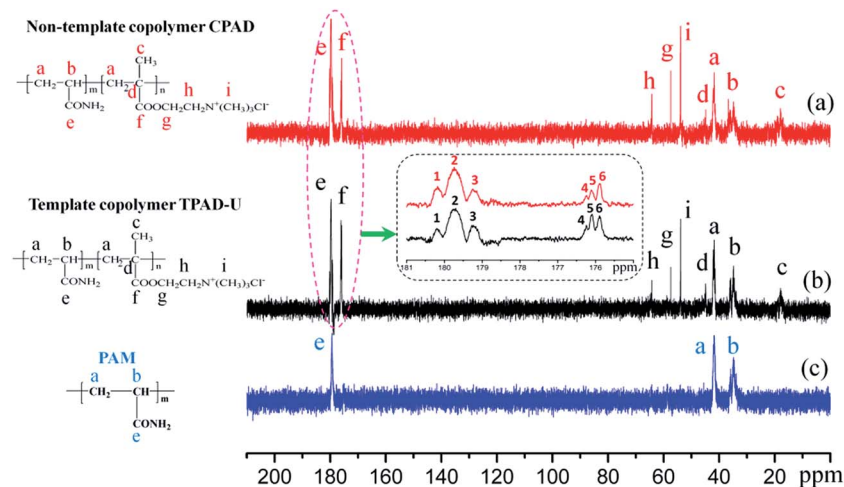


Fig. 7  $^{13}\text{C}$  NMR spectrum of (a) CPAD, (b) TPAD-U and (c) PAM.

with the CPAD, TPAD-U displayed a more rugged and rough surface morphology which was caused by the microblock structured, and this similar phenomenon has been observed in the template copolymer in previous study.<sup>18</sup>

Table 2 The mole fraction of the triads sequence structure using MestReNova software

Flocculants <sup>a</sup>	F(AAA) (%)	F(AAD) (%)	F(DAD) (%)	F(ADA) (%)	F(DDA) (%)	F(DDD) (%)
CPAD	8.9	41.3	15.7	3.2	17.4	13.5
TPAD-U	10.7	37.9	13.4	3.1	13.1	21.8

<sup>a</sup> TPAD-U: copolymer of DMC and AM by ultrasonic-template technique (UTPT); CPAD: copolymer of DMC and AM by non ultrasonic-template technique (N-UTPT); A: AM unit; D: DMC unit; F(AAA) to F(DDD): mole fraction of the AAA triad sequence to DDD triad sequence.

## 4. Flocculation test

### 4.1. Effect of flocculant dosage on dewatering performance

Contrast tests were conducted among the four flocculants (TPAD-U, CPAD, CPAM-1 and CPAM-2) to evaluate the effects of cationic microblocks on sludge flocculation performance in terms of SRF, FCMC and RT. The raw sludge had a FCMC of 96.8% and a SRF of  $132.9 \times 10^{12} \text{ m kg}^{-1}$ . Meanwhile, the zeta potential regarded as a main and extensive index were investigated to analyze the flocculation mechanisms.<sup>43,44</sup> As shown in Fig. 9, the FCMC, SRF and RT of the four flocculants initially decreased at a dosage ranging from  $10 \text{ mg L}^{-1}$  to  $40 \text{ mg L}^{-1}$ , and then increased at a dosage ranging from  $40 \text{ mg L}^{-1}$  to  $90 \text{ mg L}^{-1}$ . At low dosage, the flocculants failed to completely neutralize the negative charged sludge particles, thus resulting in a small, loose and fragile flocs. As a result, a high FCMC, SRF



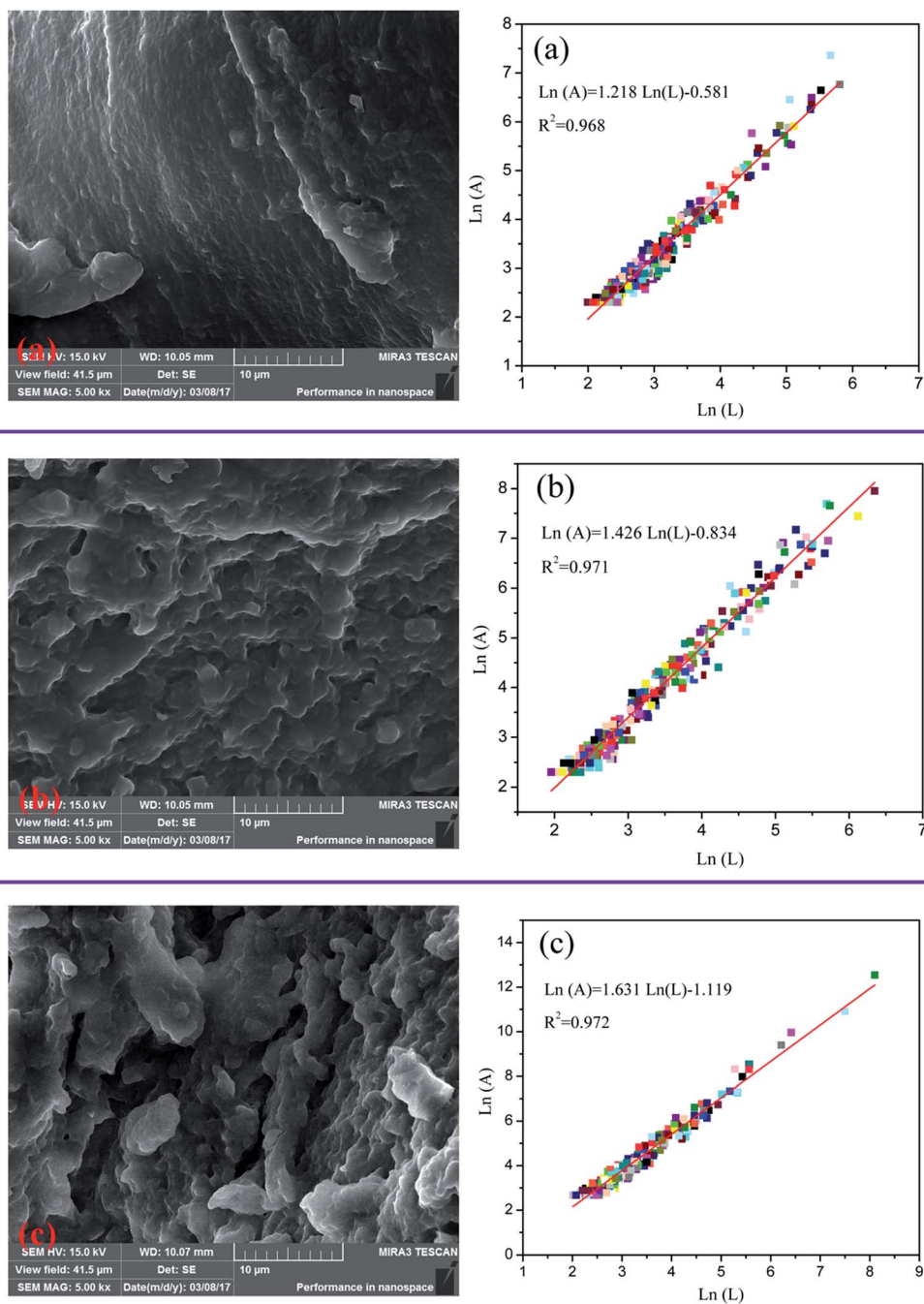


Fig. 8 SEM images of (a) PAM, (b) CPAD and (c) TPAD-U.

and RT was obtained.<sup>45</sup> However, the overdose would increase the positive charge of the colloidal system and the electronic repulsion would be increased, hence the destabilized particles re-stabilized and the flocculation efficiency was weakened.<sup>33,42</sup> It was worth noting that, in the full dosage range, TPAD-U showed the most prominent flocculation performance than the other three. When the dosage was at  $40 \text{ mg L}^{-1}$ , the minimum FCMC of 63.6%, SRF of  $4.9 \times 10^{12} \text{ m kg}^{-1}$  and RT of 5.2 NTU were acquired for TPAD-U. The phenomenon could be explained as follows. The cationic microblocks in TPAD-U with a high zeta

potential could extremely enhance the charge neutralization ability and improve the usage efficiency of DMC, and thus it would reach the isoelectric point at a relative low dosage. In other words, TPAD-U was more capable in charge neutralization and destabilization of sludge colloidal particles than CPAD, CPAM-1 and CPAM-2 at the same dosage. Besides, the optimal dosages for the four flocculants were not same with those at the isoelectric point, it indicated that mechanisms such as bridging or patching also worked with the charge neutralization mechanism in the flocculation process.





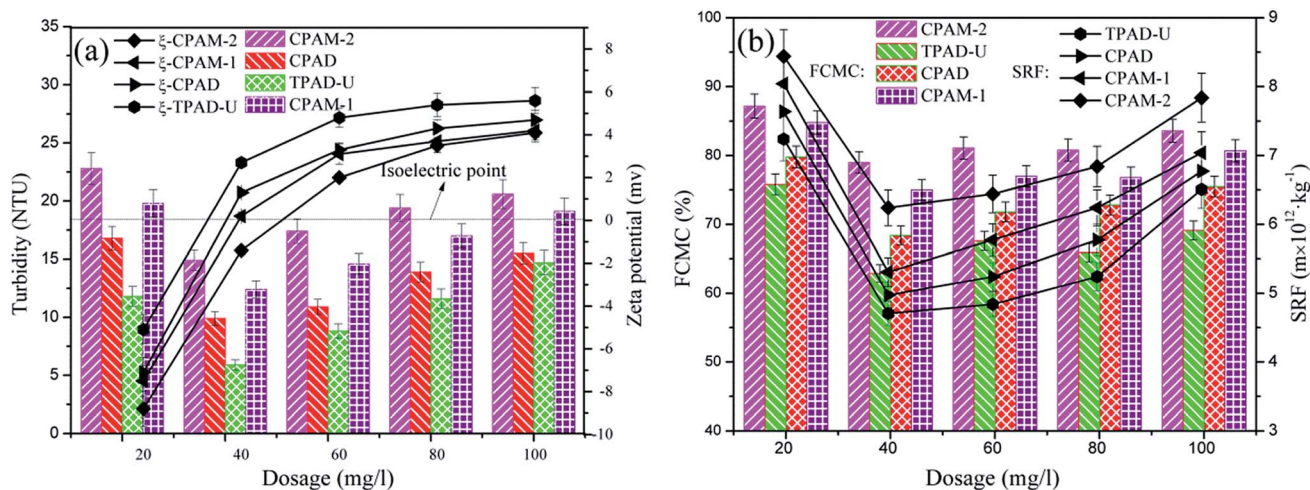


Fig. 9 Effect of dosage on (a) RT and zeta potential, (b) FCMC and SRF.

#### 4.2. Effect of pH on sludge dewatering performance

Previous study has demonstrated that the pH displayed an important impact on the sludge flocculation performance, and the flocculant surface properties could be changed under different pH value.<sup>46</sup> In this section, the influence of pH on sludge flocculation performance was investigated at the optimum flocculant dosage of  $40 \text{ mg L}^{-1}$ . Fig. 10 displayed the impact of pH on the RT, FCMC and SRF, where TPAD-U, CPAD, CPAM-1 and CPAM-2 showed the similar trend of decreasing initially at the pH range of 1 to 7 and then increasing at the pH range of 7 to 11. However, the zeta potential gradually declined at the pH range of 1 to 11. The strong acid (pH: 2.0–4.0) would lead to excessive hydrogen ions to enhance the positive charge repulsion of polymers, and the strong alkali (pH: 9.0–11.0) would result in sufficient hydroxyl ions instead of negative charged particles to neutralize the positive charge of polymers.<sup>47</sup> The strong acid and alkali would severely weaken the charge neutralization ability and more colloidal particles would not be neutralized thoroughly, and thus a desirable flocculation performance could not be obtained.

Moreover, TPAD-U showed the minimum value of RT (4.9 NTU), FCMC (62.4%) and SRF ( $4.8 \times 10^{12} \text{ mg}^{-1}$ ) and the strongest zeta potential among the four flocculants in the full pH range. It indicated that TPAD-U had the strongest charge neutralization ability generated by the cationic microblocks to possess the most excellent sludge flocculation performance in a broad scope of pH. Besides, the cationic microblocks had a strong electrostatic repulsion between the polymer chains of TPAD-U, which was favorable for the extension and spreading of the polymer chain to improve its bridging effect.<sup>19,48</sup>

#### 4.3. Sludge floc size

As reported previously, sludge floc size was regarded as an extensive and widely used index to evaluate the sludge flocculation performance.<sup>33,49</sup> The sludge floc with a large size usually performed well in separation and settlement, thus having an excellent flocculation performance. Therefore, the floc size distributions of the sludge conditioned by TPAD-U, CPAD, CPAM-1 and CPAM-2 at the optimum condition (dose:  $40 \text{ mg}$

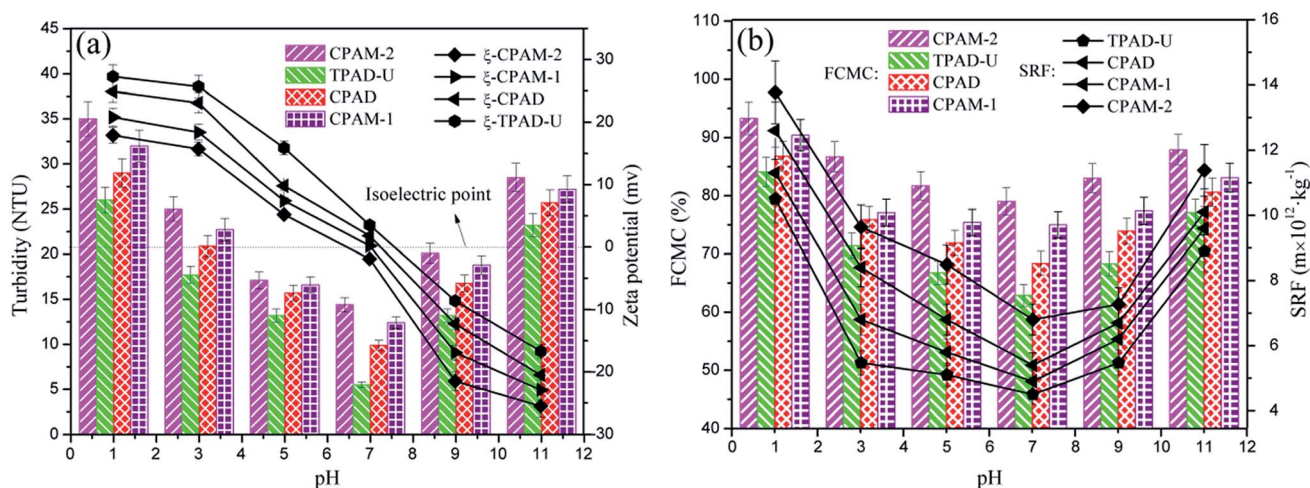


Fig. 10 Effect of the pH on (a) RT and zeta potential, (b) FCMC and SRF.



$L^{-1}$ , pH: 7.0) were recorded, and the results were shown in Fig. 11 and Table 3, where the  $d_{10}$  represented 10% in the range  $[0 \mu\text{m}-d_{10}]$ ,  $d_{50}$  represented the median equivalent volumetric diameter, and  $d_{90}$  represented 90% in the range  $[0 \mu\text{m}-d_{90}]$ . The sludge floc size of TPAD-U was demonstrated the largest in the flocculants, and those were 229.931  $\mu\text{m}$  of  $d_{10}$ , 562.428  $\mu\text{m}$  of  $d_{50}$ , 1130.207  $\mu\text{m}$  of  $d_{90}$ , respectively. Meanwhile, it was found that the morphology of the sludge flocs conditioned by TPAD-U was more compact in ESI Fig. S2,<sup>†</sup> compared with CPAD. The large and compact flocs was beneficial for the formation of drainage channel, and thereby a low SRF was obtained. By contrast, the small and loose floc would likely plug the channels and holes and result a big SRF, which was in consistent with the discussion in Section 4.1 and 4.2. These results all manifested that the cationic microblocks in TPAD-U contributed much to the excellent sludge flocculation behavior with the large and compact flocs *via* improving and strengthening the charge neutralization and bridging ability.

Table 3 The sludge flocs size distribution of  $d_{10}$ ,  $d_{50}$  and  $d_{90}$  for flocculants

Flocculant	$d_{10}^a$ ( $\mu\text{m}$ )	$d_{50}$ ( $\mu\text{m}$ )	$d_{90}$ ( $\mu\text{m}$ )
TPAD-U	229.931	562.428	1130.207
CPAD	195.396	470.622	858.349
CPAM-1	143.832	358.052	770.832
CPAM-2	123.004	313.347	609.930

<sup>a</sup>  $d_{10}$ : 10% in the range  $[0 \mu\text{m}-d_{10}]$ ;  $d_{50}$ : 50% in the range  $[0 \mu\text{m}-d_{50}]$ ; and  $d_{90}$ : 90% in the range  $[0 \mu\text{m}-d_{90}]$ .

#### 4.4. Flocculation mechanism

According to the analytical results of flocculation tests, the flocculation mechanism involved in the flocculation process was discussed and shown in Fig. 12. The cationic microblock structure in TPAD-U played a vital role in improving the sludge flocculation performance. Specifically, the flocculation mechanism

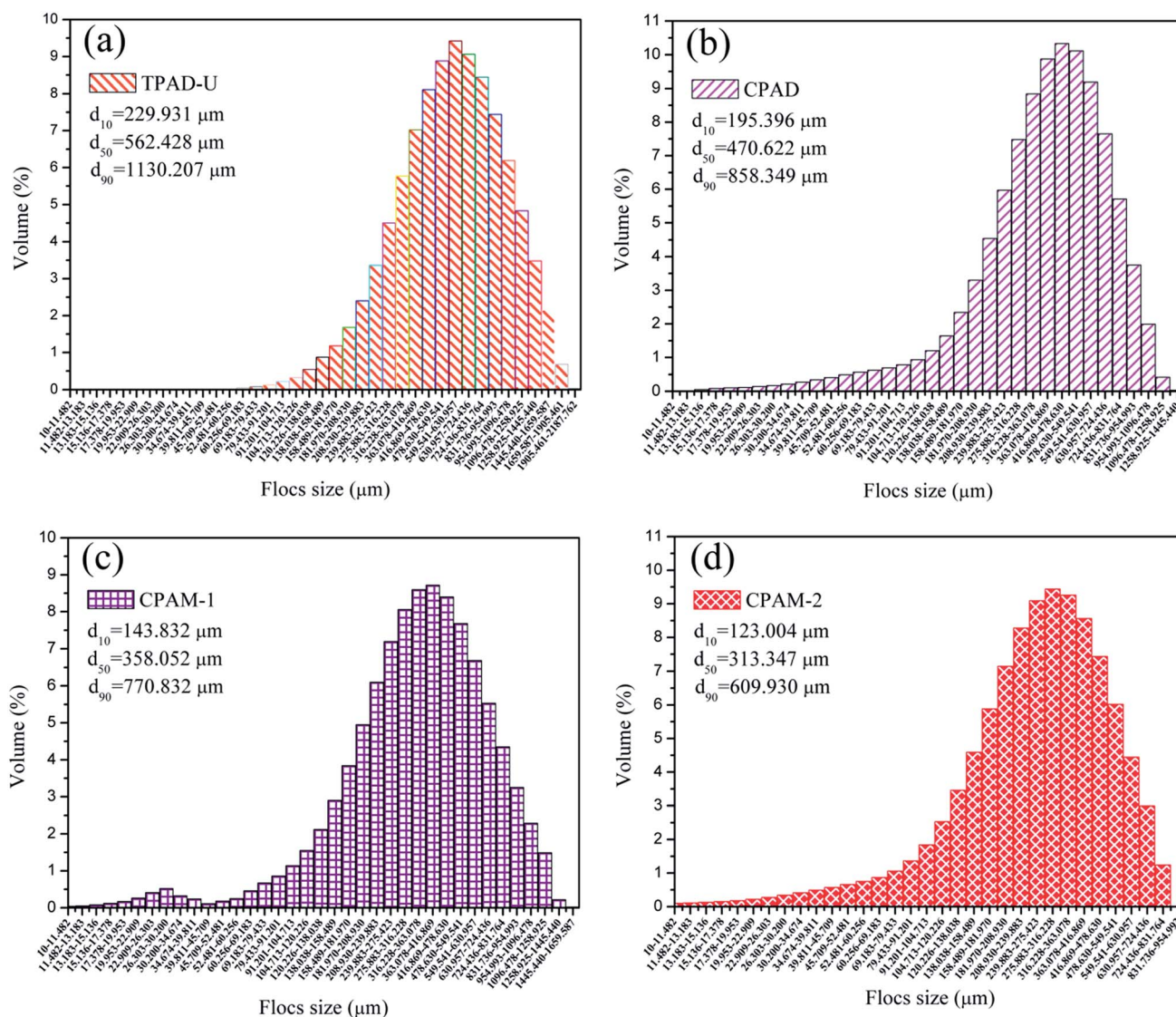


Fig. 11 Sludge flocs size distribution for (a) TPAD-U, (b) CPAD, (c) CPAM-1 and (d) CPAM-2.





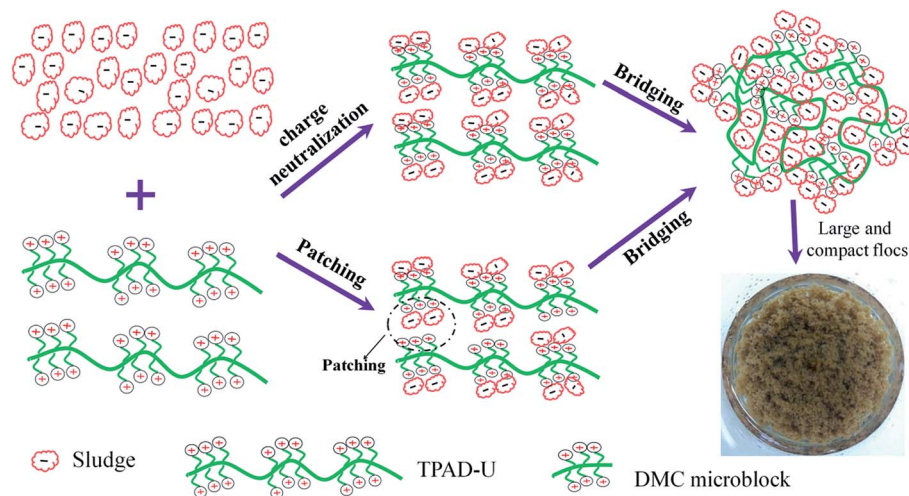


Fig. 12 Flocculation mechanism for TPAD-U.

related with the cationic microblock structure was summarized as follows: (1) the cationic microblocks in TPAD-U could greatly enhance the charge neutralization ability of the flocculant, and more negative charged sludge particles were neutralized completely and destabilized; (2) the ability of patching was also enhanced by the novel cationic microblock structure in TPAD-U, and which was helpful for the neutralization of negative charged sludge particles;<sup>18,19,37</sup> and (3) the cationic microblocks had a strong electrostatic repulsion between the polymer chain, which was helpful for the polymer chain extension to improve bridging. Finally, the destabilized sludge particles were captured and anchored on the polymer chain to aggregate and form large and compact flocs.

## 5. Conclusion

A new flocculant (TPAD-U) with the cationic microblock structure was successfully prepared through ultrasonic-template polymerization technique. The microblock structure was confirmed by characterizations such as FT-IR, <sup>1</sup>H (<sup>13</sup>C) NMR, SEM and TGA. Meanwhile, the (I) ZIP mechanism of the ultrasonic-template copolymerization was demonstrated by a high association constant ( $K_M$ ) value of 12.26, which also indicated that the formation of the DMC microblocks. The reaction kinetics manifested that (1) the conversion and  $R_p$  were extremely effected by the pH, and the optimal range of pH was from 4 to 6; (2) the free radical polymerization reaction was reflected by the power value (1.698) of  $C_M$  in the relationship of  $R_p \propto C_M^{1.698}$ , which was caused by the product precipitation in the copolymerization process; and (3) the variation trend of conversion and  $R_p$  as a function of  $n_{PMAA} : n_{DMC}$  was in line with the typical I (ZIP) template polymerization mechanism. Finally, the cationic microblocks in the TPAD-U significantly enhanced and improved the charge neutralization ability and further bridging, and thus improving the flocculation performance. The sludge flocculation results manifested that this enhanced charge neutralization and bridging effect of microblocky flocculant contributed much to lower value of SRF, FCMC and RT,

larger size, and denser floc structure, compared with the other three flocculants (CPAD, CPAM-1 and CPAM-2).

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