A. Szumna et al.
The templation effect as a driving force for the self-assembly of hydrogen-bonded peptidic capsules in competitive media
The templation effect as a driving force for the self-assembly of hydrogen-bonded peptidic capsules in competitive media†

M. Grajda, M. J. Lewińska and A. Szumna

Peptide-based cavitands (resorcin[4]arenes substituted with histidine and glutamine hydrazides) exist as monomeric species in polar solvents (DMSO and methanol). Upon complexation of fullerenes, the cavitands wrap around the hydrophobic guests forming dimeric capsular shells (as evidenced by DOSY). The self-assembly of the cavitands is based on the formation of beta-sheet-like binding motifs around the hydrophobic core. In a polar environment, these hydrogen bonded structures are kinetically stable and highly ordered as manifested by a 100-fold increase of intensity of circular dichroism bands, as well as a separate set of signals and substantial differences in chemical shifts in NMR spectra. This behavior resembles a protein folding process at the molten globule stage with non-specific hydrophobic interactions creating a protective and favourable local environment for the formation of secondary structures of proteins.

Our work takes inspiration from the native structures of proteins, which are exemplars of excellence in folding of ordered three-dimensional architectures in polar environments. At the initial stages of protein folding a pivotal role is attributed to hydrophobic interactions. Although these interactions are non-directional, they lead to the formation of a hydrophobic core, a molten globule that constitutes a perfect environment for further formation of directional hydrogen bonds by lowering the local dielectric constant. Thus, hydrogen bonds around a hydrophobic core are stronger and contribute more significantly to the stability of the native state than hydrogen bonds exposed to the aqueous environment.

This bioinspired approach based on a combination of hydrophobic core formation with an ordering power of hydrogen bonds on the peripheries is widely used for the self-assembly of infinite fibrous structures or nanoparticles. When the number of building blocks decreases, the strength of hydrophobic clustering is reduced and the above-mentioned strategy becomes less efficient. Therefore, the formation of discrete structures made of only a few building blocks is challenging. Gibb and co-workers reported on the self-assembly of discrete capsular dimers, tetramers and hexamers in water using hydrophobic interactions only. The group of Rebek developed a series of hydrocarbon-based templates for ordering of hydrogen bonded multicomponent capsules in non-polar solvents. Fullerenes have been used for the templation of self-assembled structures based on 4-pyrimidinone derivatives (isocytosine or 2-ureido-4[H]-pyrimidinone derivatives), amino acid functionalized naphthalenediimides, S-trityl-cysteine derivatives or through pyridinium–anion–pyridinium interactions. We have also presented resorcinarenes appended to hydrophobic peptides (Phe, Val) that form well-ordered hydrogen bonded dimers and complex fullerenes (Fig. 1).

In all these examples, hydrogen bonded shells are formed in non-polar media (chloroform, toluene, tetrachloroethane, carbon disulfide) and they are not able to survive in polar environments (even DMSO). In this work, we demonstrate that hydrophobic interactions together with beta-sheet-like binding motifs are effective (a) in inducing the self-assembly of peptidic cavitands that otherwise do not self-assemble and (b) in ordering hydrogen bonded shells around hydrophobic cores even in competitive, polar environments (Fig. 1).

In our quest towards self-assembly in polar environments we have designed resorcin[4]arene cavitands substituted with hydrophilic peptides, here containing glutamine and histidine residues (3a and 3b, Table 1). The side chains of these amino acids increase solubility in polar solvents and they additionally can take part in stabilizing or competitive (destabilizing) hydrogen bonding interactions due to their dual donor/acceptor character. The final thermodynamic outcome of such a complex system of interactions is highly context-dependent. For example, polyglutamine sequences in the cores of proteins
stabilize β-sheets and β-barrels through hydrogen bonding interactions between side chains.\textsuperscript{18–21} On the other hand, glutamine present in the external shells of other proteins, e.g. insulin does not form stabilizing interactions.

We obtained cavitands 3a and 3b in the reaction of respective amino acid hydrazides 2a or 2b and tetraformylresorcin[4]arene 1 (Table 1). Hydrazones 3a and 3b are soluble in polar solvents: DMSO and methanol but not in water. \textsuperscript{1}H NMR spectra in DMSO-d\textsubscript{6} reveal C\textsubscript{4}-symmetric structures for both products. Analysis of ROESY NMR spectra indicates the presence of only trivial contacts (originating from covalent proximity). The intramolecular f–j correlation signals that are hallmarks for dimerization via backbone interactions are not present. There are also no indications of specific interactions involving side-chains’ hydrogen bonding groups. These results suggest that hydrazones 3a and 3b do not self-assemble in DMSO and exist as monomeric semi-open cavitands (Fig. 1). This conclusion is also supported by DOSY measurements that reveal a diffusion coefficient of 0.95 × 10\textsuperscript{-10} m\textsuperscript{2} s\textsuperscript{-1} for 3a and 0.95 × 10\textsuperscript{-10} m\textsuperscript{2} s\textsuperscript{-1} for 3b (in DMSO-d\textsubscript{6} at 298 K), corresponding to an average diameter of molecule of ~2.2 nm (predicted value for the cavitand is 2.3 nm, the model is based on the previously reported X-ray structure of the valine derivative).\textsuperscript{17} This self-assembly behaviour and solubility characteristics of cavitands containing polar side chains are distinctly different from those of the previously reported hydrophobic derivatives that were soluble only in non-polar solvents and formed dimeric capsules in these media.\textsuperscript{17}

In subsequent experiments, we tested the possibility of templation of self-assembly by fullerene C\textsubscript{60}. For cavitands 3a and 3b a direct complexation experiment is not feasible due to the lack of solubility of fullerene in polar solvents. Therefore alternative ways of reaction activation were employed: (1) thermal activation of samples (heating of 3a + C\textsubscript{60} or 3b + C\textsubscript{60} in DMSO at 70 °C for 7 days with vigorous stirring, path A); (2) complexation during the chemical reaction of cavitand synthesis (path B); and (3) mechanochemical activation\textsuperscript{22} in the solid state (Table 1).

In all cases, the resulting products contained two types of species as revealed by \textsuperscript{1}H NMR spectra in DMSO-d\textsubscript{6} (Fig. 2b and e). One set of signals corresponds to the cavitands (3a or 3b) and the second one is attributed to the dimeric species with encapsulated fullerenes ([3a]\textsubscript{2} ⊂ C\textsubscript{60} or [3b]\textsubscript{2} ⊂ C\textsubscript{60}). This assignment is based on \textsuperscript{13}C NMR spectra that reveal the signal of encapsulated fullerene at 141 ppm and on DOSY

### Table 1 Synthesis of cavitands 3a and 3b, C\textsubscript{60} complexation strategies and condition optimization of ([3a]\textsubscript{2} ⊂ C\textsubscript{60} synthesis

<table>
<thead>
<tr>
<th>Entry</th>
<th>Path</th>
<th>Substrates</th>
<th>Equiv.</th>
<th>Conditions</th>
<th>Yield\textsuperscript{a} of:</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>A</td>
<td>3a + C\textsubscript{60}</td>
<td>2:1</td>
<td>DMSO (24 h, rt)</td>
<td>$\text{[3a]}<em>2 \subset \text{C}</em>{60}$ 3a</td>
</tr>
<tr>
<td>2</td>
<td>A</td>
<td>3a + C\textsubscript{60}</td>
<td>2:1</td>
<td>DMSO (24 h, 80 °C)</td>
<td>8% 92%</td>
</tr>
<tr>
<td>3</td>
<td>A</td>
<td>3a + C\textsubscript{60}</td>
<td>2:1</td>
<td>Ball-mill (2 h)</td>
<td>18% 82%</td>
</tr>
<tr>
<td>4</td>
<td>A</td>
<td>3a + C\textsubscript{60}</td>
<td>2:10</td>
<td>DMSO (24 h, rt)</td>
<td>12% 88%</td>
</tr>
<tr>
<td>5</td>
<td>A</td>
<td>3a + C\textsubscript{60}</td>
<td>2:10</td>
<td>DMSO (24 h, 80 °C)</td>
<td>15% 85%</td>
</tr>
<tr>
<td>6</td>
<td>A</td>
<td>3a + C\textsubscript{60}</td>
<td>2:10</td>
<td>Ball-mill (2 h)</td>
<td>19% 81%</td>
</tr>
<tr>
<td>7</td>
<td>B</td>
<td>1 + 2a + C\textsubscript{60}</td>
<td>2:8:1</td>
<td>DMSO (24 h, 80 °C)</td>
<td>24% 86%</td>
</tr>
<tr>
<td>8</td>
<td>B</td>
<td>1 + 2a + C\textsubscript{60}</td>
<td>2:8:1</td>
<td>Ball-mill (2 h)</td>
<td>21% 79%</td>
</tr>
<tr>
<td>9</td>
<td>B</td>
<td>1 + 2a + C\textsubscript{60}</td>
<td>2:8:10</td>
<td>DMSO (24 h, 80 °C)</td>
<td>42% 58%</td>
</tr>
<tr>
<td>10</td>
<td>B</td>
<td>1 + 2a + C\textsubscript{60}</td>
<td>2:8:10</td>
<td>Ball-mill (2 h)</td>
<td>22% 78%</td>
</tr>
</tbody>
</table>

\textsuperscript{a}Yields were determined by NMR.
experiments that clearly indicate a larger size of the newly formed species (Fig. 2c and f). Diffusion coefficients for \((3a)_2 \supset C_{60} (0.82 \times 10^{-10} \text{ m}^2 \text{ s}^{-1})\) and \((3b)_2 \supset C_{60} (0.80 \times 10^{-10} \text{ m}^2 \text{ s}^{-1})\) at 298 K in DMSO-d6 correspond to the diameter of 2.7 nm (predicted value for the dimeric capsule is \(\sim 2.7 \text{ nm}\)). Additionally, in the ROESY spectra the \(f-j\) correlation signals are present for \((3a)_2 \supset C_{60} \) and \((3b)_2 \supset C_{60}\) indicative of a close contact between hemispheres in a head-to-head arrangement (Fig. 3). This allows us to postulate the capsular structures of complexes \((3a)_2 \supset C_{60}\) and \((3b)_2 \supset C_{60}\) in DMSO-d6. The corresponding ESI-MS experiments for the products revealed the presence of \(3a\) and \((3a)_2 \supset C_{60}\) or \(3b\) and \((3b)_2 \supset C_{60}\), without any traces of “empty” capsules, \((3a)_2\) or \((3b)_2\) (ESI†), which proves that the presence of a hydrophobic templating agent is essential for the self-assembly of cavitands \(3a\) and \(3b\) in DMSO.

In order to check the stability of hydrogen bonded species in the presence of a competitive solvent that is a hydrogen bond donor and acceptor we performed experiments in methanol.

Table 2 Solubility table of \(3a\), \(3b\), \((3a)_2 \supset C_{60}\) and \((3b)_2 \supset C_{60}\) in MeOH and DMSO

<table>
<thead>
<tr>
<th>Solvent</th>
<th>(3a)</th>
<th>((3a)<em>2 \supset C</em>{60})</th>
<th>(3b)</th>
<th>((3b)<em>2 \supset C</em>{60})</th>
</tr>
</thead>
</table>

Sol. – well-soluble; insol. – insoluble.

The solubilities in methanol are also different, but we were not able to separate the mixture, due to co-crystallization of the species. However, from a supersaturated methanolic solution we obtained enriched crystals containing up to 60% of a \((3a)_2 \supset C_{60}\) (ESI†). After dissolution of the crystals in DMSO intact complex \((3a)_2 \supset C_{60}\) was present indicating that also glutamine-based self-assembled species are stable in an environment of methanol.

It is important to note that fullerene-templated self-assembly results in the formation of a well-defined hydrogen bonding system with a \(\beta\)-sheet-like binding motif and an ordered inherently chiral conformation. In the \(^1\)H NMR spectrum, all signals for hydrogen capable of hydrogen bonding are considerably downfield shifted in \((3a)_2 \supset C_{60}\) as compared to the same species in MeOH (see Table 2).

**Fig. 2** \(^1\)H NMR (600 MHz, DMSO-d6) spectrum of \(3a\) (a) and a mixture of \(3a\) and \((3a)_2 \supset C_{60}\) (b); DOSY spectrum of \(3a\) and \((3a)_2 \supset C_{60}\) (c); \(^1\)H NMR spectrum of \(3b\) (d) and a mixture of \(3b\) and \((3b)_2 \supset C_{60}\) (e); DOSY spectrum of \(3b\) and \((3b)_2 \supset C_{60}\) (f).

**Fig. 3** Dimerization and ordering of the structures in DMSO: ROESY effects indicative of \(\beta\)-sheet-type interactions (a) for \((3a)_2 \supset C_{60}\); (b) for \((3b)_2 \supset C_{60}\) and (c) predicted binding motif (the 3D structure was obtained by Val → Gln mutation of an X-ray structure of a valine derivative†).
involving an inherently chiral chromophore. These results involve conformational inherent chirality and symmetry of these conformations (Fig. 4). These coupling constants are directly related to the backbone angles as described by the Karplus relationship for peptides. Higher values are hallmarks of extended β-sheet-like binding motif is also supported by the increase of three-bond coupling constants \( J(NH-CH_3) \) from 7.2 Hz for 3a to 8.5 Hz for \( 3a_2 \subset C_{60} \) and from 8.2 Hz for 3b to 8.9 Hz (DMSO) or 9.6 Hz (MeOH) for 3b \( \supset C_{60} \). These coupling constants are directly related to the backbone dihedral \( \phi \) angles as described by the Karplus relationship for peptides. Higher values are hallmarks of extended β-strand conformations (\( \phi \approx -139° \)) over a helix (\( \phi \approx -57° \)) or a random coil. An ordered conformation is also reflected in circular dichroism spectra that exhibit almost a 100-fold enhancement in the intensity of the lowest energy bands for \( 3b_2 \subset C_{60} \) as compared with 3b (Fig. 4d). These bands are the most sensitive to conformational inherent chirality and symmetry of these molecules, as they come from the HOMO–LUMO transition involving an inherently chiral chromophore. These results indicate that the capsular dimers indeed retain a hydrogen-bonded ordered structure in competitive medium. It is also important to note that there are no traces of dynamic exchange between species at the NMR timescale, indicating that the complexes are also kinetically persistent in this solvent, which is quite surprising considering that semi-open cavitands are also present in the solution.

In summary, we have demonstrated that wrapping of hydrogen-bonded shells around a hydrophobic template is effective in a highly disintegrative environment. DMSO is a strong competitor as a hydrogen bond acceptor and due to its hydrophobic character, while methanol is a strong hydrogen bond donor and acceptor. Therefore, both environments are competitive for aggregates based on hydrogen bonding and hydrophobic interactions. In fact, DMSO or methanol is commonly used for unfolding of natural proteins or disintegration of hydrogen bonded artificial aggregates. Here we demonstrated that the optimal combination of both interactions enables effective self-assembly, although none of the interactions is effective alone. The current system is not water-soluble, but we predict that this strategy can also be extended to aqueous solutions due to even stronger hydrophobic interactions.

Conflicts of interest

There are no conflicts of interest to declare.

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Notes and references