Solar hydrogen production using epitaxial SrTiO\(_3\) on a GaAs photovoltaic

L. Kornblum,‡ab D. P. Fenning,¶cd J. Faucher,¶cd J. Hwang,‡ M. D. Morales-Acosta,‡ab Y. Zhu,‡ E. I. Altman,‡bi M. L. Lee,‡ek C. H. Ahn,‡ab F. J. Walker§*ab and Y. Shao-Horn§*cT

We demonstrate an oxide-stabilized III–V photoelectrode architecture for solar fuel production from water in neutral pH. For this tunable architecture we demonstrate 100% Faradaic efficiency for hydrogen evolution, and incident photon-to-current efficiencies (IPCE) exceeding 50%. High IPCE for hydrogen evolution is a consequence of the low-loss interface achieved via epitaxial growth of a thin oxide on a GaAs solar cell. Developing optimal energetic alignment across the interfaces of the photoelectrode using well-established III–V technology is key to obtaining high performance. This advance constitutes a critical milestone towards efficient, unassisted fuel production from solar energy.

Introduction

One of the grand challenges for creating a sustainable society is to develop practical materials and devices that produce fuels when exposed to sunlight. Solar fuel production, e.g. via photoelectrochemical (PEC) water splitting\(^1\)–\(^3\) or \(\text{CO}_2\) reduction,\(^4\) allows the storage of solar energy in chemical bonds for use on demand. Transition metal oxides\(^6\)–\(^7\) represent a flexible class of materials for promoting water splitting kinetics including hydrogen and oxygen evolution from water, but unfortunately their bandgaps are typically too wide for efficient solar energy collection. Meanwhile, the most efficient solar energy collectors, III–V semiconductor solar cells,\(^8\) are not chemically stable in the relevant environments for long term operation.\(^9\),\(^10\)

Recent advances in photoelectrochemical hydrogen\(^11\) and oxygen evolution\(^12\)–\(^14\) include the incorporation of III–V semiconductor solar cells,\(^15\),\(^16\) stabilized using surface protection oxides,\(^11\),\(^17\) phosphides and sulfides;\(^18\) often in combination with precious metal catalysts, e.g., Pt,\(^11\),\(^17\),\(^19\),\(^20\) to promote reaction kinetics. For example, ground breaking work using InP semiconductors\(^21\),\(^22\) achieved high efficiency using a protective oxide formed by chemical reaction of the semiconductor and electrolyte, followed by application of a precious metal catalyst. Further advances toward the development of more robust catalyst-oxide-semiconductor heterostructures can be made using an epitaxial oxide where each interface can be engineered to achieve catalytic activity without a precious metal and to provide robust protection from the electrolyte. In addition, incomplete control of the interfaces between the catalysts, surface protection layers, and the underlying semiconductor, has limited the ability to achieve further advances.\(^21\) Here, we demonstrate a stable, epitaxial oxide III–V photoelectrode architecture for solar fuel production from water in neutral pH. For this tunable architecture we demonstrate 100% Faradaic efficiency for hydrogen evolution, and incident photon-to-current efficiencies (IPCE) exceeding 50%. The high IPCE is a consequence of the low-loss interface achieved via epitaxial growth of a thin oxide on a GaAs solar cell with a band offset that promotes electron transfer to the hydrogen reduction couple. The SrTiO\(_3\) forms a single
crystal, epitaxial layer with an atomically abrupt interface with the GaAs so that the electron energetics can be traced in detail from electron–hole pair generation in the GaAs to the efficient delivery of the electron to a hydrogen ion at the SrTiO3–electrolyte interface.

We report an np-GaAs(001) photocathode that operates in neutral pH, stabilized by an epitaxial SrTiO3 surface layer to deliver 3.1 mA cm⁻² of hydrogen evolution current at 0.18 V above the thermodynamic potential. The photocathode consists of a 16 nm-thick single-crystal SrTiO3 layer that protects the GaAs photocathode and provides stability during 24+ hours of hydrogen production. IPCE measurements reveal little loss of photogenerated carriers at the atomically-sharp SrTiO3/np-GaAs(001) interface, permitting electron flow to the surface with a large thermodynamic driving force for hydrogen evolution reaction (HER, Fig. 1a). By achieving a quantitative description of the electronic band alignment at both the electrolyte–oxide interface and the buried oxide–semiconductor interface, we demonstrate the potential for this epitaxial oxide–semiconductor platform to leverage both oxide and III–V band-engineering toward highly efficient photoelectrochemical devices.

While photoelectrochemical water-splitting at extreme pH is currently preferred, neutral pH operation remains an important goal for environmentally-benign solar fuels. At the system level, the water splitting in extreme alkaline or acidic environments is important for limiting transport and concentration polarization. More fundamentally, extreme pH conditions enhance the activity of catalytic surfaces currently available. We also note that perovskite surfaces offer promising opportunities for CO2 reduction, which operates preferentially in near-neutral pH. Enhancing the activity of available surfaces, especially at neutral pH, is a fundamental challenge for the field and addressed in this work. The fundamental aspects of interfaces and band alignments presented here will enable the integration of oxide electronic materials and III–V semiconductors for efficient solar fuel production.

Results and discussion

A thin SrTiO3 metal oxide layer of 40 unit cells (~16 nm-thick) is epitaxially grown on GaAs(001) solar cells by molecular beam epitaxy (MBE). A schematic of the 16 nm-thick SrTiO3/np-GaAs(001) photocathode (STOPC) operating under illumination at 0 V vs. the reversible hydrogen electrode (RHE), or 0 V vs. RHE is shown in Fig. 1a. The np-GaAs(001) solar cells (Table S1 and Fig. S1, ESI†) consist of an np junction with a p-Al0.4Ga0.6As back surface field with a measured open-circuit voltage of 0.94 V (Fig. S2, ESI†). The np junction separates photoexcited electrons and holes and the electrons flow through the conduction band to the SrTiO3 surface where they reduce protons to hydrogen; in addition, the wide SrTiO3 bandgap blocks holes...
from reaching the surface. The MBE growth of the oxide results in an atomically sharp interface between SrTiO₃ and GaAs(001) as confirmed by scanning transmission electron microscopy (Fig. 1b). Streaks in the reflection high-energy electron diffraction (RHEED) indicate a smooth SrTiO₃ surface with high crystallinity (Fig. 1b, inset). High-quality crystalline [Fig. S3 and S4, ESI†] interfaces provide a platform for understanding and band engineering of oxide III–V photocathodes.

The proposed band alignments at the SrTiO₃/GaAs(001) and water–SrTiO₃ interfaces in the dark at equilibrium with the H⁺/H₂ couple (0 V RHE and −4.0 eV on the absolute energy scale at pH 7) are shown in Fig. 1c. X-ray photoelectron spectroscopy (XPS) analysis of the valence band (Fig. 1d) and As 3d and Ti 3p core levels (Fig. S5, ESI†) places the SrTiO₃ conduction band 13 0.7 ± 0.2 eV below that of GaAs. When the electron energy equilibrates with the H⁺/H₂ couple, this alignment also causes the GaAs bands at the interface with SrTiO₃ to bend upwards, forming a barrier for electron transport to the SrTiO₃ conduction band (Fig. 1c). Under illumination (Fig. 1a), a portion of the photovoltage of the solar cell is needed to reduce the magnitude of this barrier and produce HER current. In addition, the valence band of SrTiO₃ is positioned 2.6 ± 0.2 eV below that of n-GaAs, providing a large hole-blocking barrier that prevents hole transport from the n-GaAs into SrTiO₃ and GaAs corrosion.⁹

The SrTiO₃ conduction band edge is 0.37 eV higher than energy of the hydrogen on the absolute scale (Fig. 1c), as determined from the Mott–Schottky analysis of the flatband potential of a conductive, 1 (at)% Nb-doped SrTiO₃ single crystal (Fig. S6, ESI†). The doping induces negligible changes to the optical gap from that of undoped SrTiO₃ and is a good approximation to the epitaxial SrTiO₃. Fast HER kinetics can be expected given the large 0.37 eV thermodynamic driving force.¹⁵,³⁶

Electrochemical HER measurements were performed on 5 × 5 mm² pieces of this structure, which were contacted on the backside and insulated from the sides to form the photocathode. Photoelectrochemical HER kinetics were measured in 0.1 M potassium phosphate electrolyte solution (pH 7), selected for its non-corrosive, non-toxic buffering.³⁷ While similar SrTiO₃ films have exhibited excellent stability under acidic conditions,¹³ a neutral pH chemistry was chosen to provide an environmentally-benign, non-hazardous route for HER.

Because of the band alignments in the STOPC, large solar-to-hydrogen currents are realized under 1 Sun, and the catalyst-free STOPC provides a ~0.55 V voltage gain with respect to a Pt catalyst after extended chronoamperometry (at 1 mA cm⁻², Fig. 2a). Following 24 h solar hydrogen production at 0 V RHE under simulated 1 Sun illumination, cyclic voltammetry of the STOPC shows a HER onset potential of ~0.3 V RHE, and a large reduction current of 6 mA cm⁻² at 0 V RHE (Fig. 2a). Bubble formation and mass transport limitations are observed as noise in the current measurement below 0.1 V RHE (Fig. 2a). Gas chromatography measurements show that hydrogen evolution accounts for all the measured reduction current, yielding ~100% Faradaic efficiency (Fig. 2b). Comparing this onset potential to the V₆₇ of 0.94 V highlights the prospects for further improvements accessible in this material system.

In contrast to the illuminated photocathode, the reduction current is negligible on the SrTiO₃/np-GaAs in dark even at −0.4 V RHE. An illuminated Nb:SrTiO₃(001) photoelectrode that can only collect UV light³⁴,³⁸ produces negligible currents at 0 V RHE, and dark reduction currents become measurable at potentials below −0.8 V RHE. The large reduction current and gain in potential measured under illumination for the SrTiO₃–GaAs result from electrons generated by the III–V solar cell that are injected into the SrTiO₃ conduction band and reduce protons to hydrogen at the oxide surface.

HER currents from the STOPC remained stable under 1 Sun illumination during 24 h potentiostatic testing at 0 V RHE, as shown in Fig. 2b. A second round of CV and IPCE measurements, performed at reaction conditions for several hours following chronoamperometry (CA) demonstrates that the STOPC stability extends beyond 24 h. The increase in reduction current at short times during chronoamperometry (0–2 hours) and improvement in CV measurements (Fig. S7, ESI†) is consistently observed for all SrTiO₃/np-GaAs samples, and suggests an activation of the SrTiO₃/electrolyte interface. After 24 hours of operation under simulated sunlight, XPS analysis shows negligible changes in the SrTiO₃ surface chemistry (Fig. S8, ESI†). In contrast, bare np-GaAs photoelectrodes quickly corrode under these conditions (Fig. S9, ESI†). We thus demonstrate that 16 nm-thick epitaxial SrTiO₃...
provides sustainable chemical protection and passivation to the underlying semiconducting solar cell.

Comparing the probability that incident photons lead to electrons collected at the cathode (IPCE) of the STOPC in solution where H₂ is generated (Fig. 3a) and in a photovoltaic configuration (Fig. 3b) reveals several important insights. First, in the photovoltaic configuration (Fig. 3b), the oxide has no impact on the GaAs IPCE which exceeds 50% between 700–850 nm. We further note that the similarity of the dry and PEC IPCE behaviour means that the reduction in the AM1.5g photocurrent in Fig. 2 relative to the photocurrent measured for the dry cell is not due to carrier recombination at the SrTiO₃/n-GaAs interface. Rather, the current reduction is the result of an effective series resistance due to the barrier at the SrTiO₃/n-GaAs interface (Fig. S14, ESI†) and kinetic limitations at the SrTiO₃/electrolyte interface. Front surface reflectivity is responsible for much of the loss of incident photons, thus nearly all the photoexcited electrons make it to the SrTiO₃ surface. At 0 V RHE, Fig. 3a shows that the IPCE for H₂ generation is virtually identical to the photovoltaic IPCE, confirming robust electron transport from the SrTiO₃ surface to the electrolyte for H⁺ reduction. The voltage dependence of the H₂ generation IPCE suggests barriers to charge injection into the electrolyte.

Unlike previous work on protected Si or GaAs photocathodes that requires the use of Pt, the STOPC shows efficient carrier collection for hydrogen evolution without an additional metal catalyst. In order to rule out the impact of trace metal impurities such as platinum, on the catalytic performance of the photocathode device during PEC, an extended PEC and XPS study was conducted (Fig. S10–S13, ESI†). The conclusion from these studies is that the activity of the SrTiO₃ surface cannot be attributed to surface Pt contamination. The IPCE values demonstrated here (Fig. 3) are comparable to those of oxide-protected Si photocathodes that required nano-structured Pt catalyst at a much higher overpotential of −0.4 V RHE. Moreover, the addition of ~3 nm of Pt on top of SrTiO₃/np-GaAs photocathode did not lead to significant changes in the hydrogen evolution current and IPCE. The large electronic driving force (~0.37 eV) for electron transfer from the oxide conduction band to the hydrogen–water redox couple is shown to be sufficient to promote hydrogen evolution kinetics.

Photoelectrochemical IPCE of the STOPC is comparable to photovoltaic (PV) IPCE of bare np-GaAs solar cells and is not reduced by the addition of the epitaxial oxide and the collection of current via HER instead of a metal contact (Fig. 3b). The IPCE at 0 V RHE is in excellent agreement with the short-circuit dry two-contact PV IPCE. We expect that there is substantial opportunity for further gains in solar-to-hydrogen efficiency by improvements in the GaAs cell design.

Energy losses at the SrTiO₃/n-GaAs and water–SrTiO₃ interfaces (Fig. 1c) lead to a reduction in the photometric power saved of the photoelectrode, which can be defined as the ratio of HER current times the HER potential relative to RHE to solar energy input. As the most conservative estimate, the photometric power saved compared to an ideally nonpolarizable dark electrode (E = 0 V RHE) was used in part because our scheme is blocking for electron flow in the dark and exhibits a lack of superposition (Fig. S14, ESI†). This efficiency, calculated from Fig. 2a at the maximum power point to be ~0.55%, is reduced from the maximum available from the solar cell, ~10%. If there were no conduction band offset at the SrTiO₃/n-GaAs interface, the 0.94 V open-circuit voltage of the solar cell could shift the onset potential from the −0.37 V RHE flatband of the SrTiO₃ in dark to ~−0.57 V RHE under illumination. Due to the loss of electron energy at the conduction band offset between SrTiO₃ and n-GaAs (see Fig. S14 and S15, ESI†), the current device achieves an onset at ~−0.30 V RHE (Fig. 2a). Thus, ~0.27 V could be gained by tuning the interface chemistry to reduce the conduction band offset using functional oxide engineering techniques applicable to epitaxial oxides like SrTiO₃ or by creating a thin heavily-doped tunnelling layer of n⁺-GaAs at the interface. Several hundred mV of further improvement could be gained by reducing the 0.37 V offset at the water–oxide interface by adjusting the oxide’s band structure to increase the electron affinity just enough so that some offset at the water–oxide interface remains to drive hydrogen evolution without a catalyst. The quantitative band alignments developed here outline a clear path toward greater photoelectrode efficiency.

Conclusions

This work demonstrates the robustness of integrating III–V technology with a high-quality single-crystal, epitaxial oxide as
a platform for further development of photocathodes for solar fuel production. Using a catalyst-free 16 nm-thick SrTiO$_3$ on np-GaAs, a stable hydrogen evolution current is produced under 1 Sun with IPCE reaching 50% at the thermodynamic potential of 0 V_RHE. Because of the high-quality of the SrTiO$_3$/GaAs interface, the IPCE for hydrogen production matches the photovoltaic performance of the GaAs solar cell. Extending this approach to high-efficiency tandem solar cells offers the potential to generate sufficient photovoltage for stable unassisted water splitting. For example, a III–V on a Si dual-junction cell operating as a photocathode could achieve photovoltages well in excess of the 1.23 V thermodynamic minimum required for water-splitting while leveraging low-cost Si substrates with high efficiency III–V materials. Combining the tunability of complex oxides and the sophisticated engineering of III–V solar cells, large gains in solar hydrogen production should be readily accessible using the catalyst-free oxide-stabilized III–V platform demonstrated here.

Acknowledgements

The authors (CHA, LK, MDAM, and FJW) acknowledge support from NSF DMR1309868 and EIA acknowledges support from MRSEC DMR-1119826 (CRISP). Support for MIT research is acknowledged from the MIT Energy Initiative seed fund and the Cooperative Agreement between the Masdar Institute of Science and Technology, Abu Dhabi, UAE and MIT, Reference Number 02/MI/MIT/CP/11/07633/GEN/G/00. DPF acknowledges the support of the MIT/Battelle postdoctoral program. JF and MLL acknowledge support from ARPA-E Award DE-AR0000508. The work at Brookhaven National Laboratory was supported by the Materials Science and Engineering Division, Office of Basic Energy Sciences, of the US Department of Energy, under Contract No. DE-AC02-98CH10886. The authors are grateful to Nir Pour for his expertise in preparing Fig. 1. The authors also thank Dr Ifan E. L. Stephens for discussions and assistance in measuring Pt wire, and Dr B. R. Lukano for his contributions to the early stages of this project.

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