

Cite this: *Chem. Sci.*, 2017, 8, 8291

A matrix of heterobimetallic complexes for interrogation of hydrogen evolution reaction electrocatalysts†

Pokhraj Ghosh,^a Shengda Ding,^a Rachel B. Chupik,^a Manuel Quiroz,^a Chung-Hung Hsieh,^b Nattami Bhuvanesh,^a Michael B. Hall^a and Marcetta Y. Darensbourg^{a*}

Experimental and computational studies address key questions in a structure–function analysis of bioinspired electrocatalysts for the HER. Combinations of NiN₂S₂ or [(NO)Fe]N₂S₂ as donors to (η⁵-C₅H₅)Fe(CO)⁺ or [Fe(NO)₂]⁺⁰ generate a series of four bimetallics, gradually “softened” by increasing nitrosylation, from 0 to 3, by the non-innocent NO ligands. The nitrosylated NiFe complexes are isolated and structurally characterized in two redox levels, demonstrating required features of electrocatalysis. Computational modeling of experimental structures and likely transient intermediates that connect the electrochemical events find roles for electron delocalization by NO, as well as Fe–S bond dissociation that produce a terminal thiolate as pendant base well positioned to facilitate proton uptake and transfer. Dihydrogen formation is *via* proton/hydride coupling by internal S–H⁺⋯H–Fe units of the “harder” bimetallic arrangements with more localized electron density, while softer units convert H⁺⋯H⁺ *via* reductive elimination from two Fe–H deriving from the highly delocalized, doubly reduced [Fe₂(NO)₃][–] derivative. Computational studies also account for the inactivity of a Ni₂Fe complex resulting from entanglement of added H⁺ in a pinched –S^{δ–}⋯H⁺⋯S^{δ–} arrangement.

Received 2nd August 2017
Accepted 11th October 2017

DOI: 10.1039/c7sc03378h

rsc.li/chemical-science

Introduction

From protein crystallography the bimetallic active site structures in enzymes such as [NiFe]-, [FeFe]-hydrogenases (H₂ase), CO-dehydrogenases and acetyl coA synthase (ACS) have been convincingly interpreted in terms of characteristics needed for their organometallic-like functions.^{1,2} Whereas most major homogeneous catalytic applications involving redox processes use precious metals that can operate as single site catalysts, the intricate molecular arrangements in nature's biocatalysts harness combinations of at least two first row transition metals, connected by sulfides or thiolates, along with Lewis acid/base sites.^{3–5} Over the past two decades a rich area in synthetic chemistry inspired by such natural products has developed,

yielding biomimetics for insight into enzyme mechanisms. In addition the link between the [NiFe]- and [FeFe]-H₂ase active sites and base metal, sustainable catalysts for the Hydrogen Evolution Reaction (HER) holds promise for the production of H₂ from “solar” (photovoltaic) electrons *via* electrocatalysis.⁶ Specific efforts have been directed towards the use of metal-lodithiolates from MN₂S₂ complexes as bidentate donor ligands (readily deduced from the structure of the ACS enzyme active site), that bind to receiver metal units *via* bridging dithiolates.^{7–25} The electronic requirements of the thiolate sulfurs have a steric consequence in the butterfly M(μ-SR)₂M' cores that are seen in the H₂ase active sites, placing M and M' within close proximity.²⁶

The advancement of chemistry *via* structure/function analysis of sets of compounds with well-known differences in composition and structure is a challenge in the complicated area of HER electrocatalysis. Nonetheless the metallodithiolate-as-synthon approach, inspired from the ACS active site, permits modular design that includes some features of the bimetallic [NiFe]- and [FeFe]-H₂ase active sites beyond the obvious dithiolate core structures. An initial foray explored the properties of the diiron, trinitrosyl complex shown in Fig. 1.^{8,27} With it we intended to exploit the redox-activity of {Fe(NO)}^{7/8} in the (NO)FeN₂S₂ metalloligand bound to a redox-active {Fe(NO)₂}^{9/10}, iron dinitrosyl unit. Electrochemical studies of [(NO)FeN₂S₂·Fe(NO)₂]⁺, [Fe–Fe]⁺, Fig. 1, found two single-

^aDepartment of Chemistry, Texas A & M University, College Station, TX 77843, USA. E-mail: marcetta@chem.tamu.edu

^bDepartment of Chemistry, Tamkang University, New Taipei City, Taiwan 25157

† Electronic supplementary information (ESI) available: Experimental, spectroscopic, additional electrochemical and computational details, X-ray crystallographic data (CIF) from the structure of the complexes [Ni–Fe]⁰, [Ni₂–Fe₂]²⁺, [Ni₂–Fe]⁺, and computational coordinates are available. CCDC crystallographic data for the complexes [Ni–Fe]⁰, [Ni₂–Fe₂]²⁺ and [Ni₂–Fe]⁺ were deposited in the Cambridge Crystallographic Data Centre. CCDC [Ni–Fe]⁰ (CCDC 1045461), [Ni₂–Fe₂]²⁺ (CCDC 1045460) and [Ni₂–Fe]⁺ (CCDC 1565539). For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7sc03378h



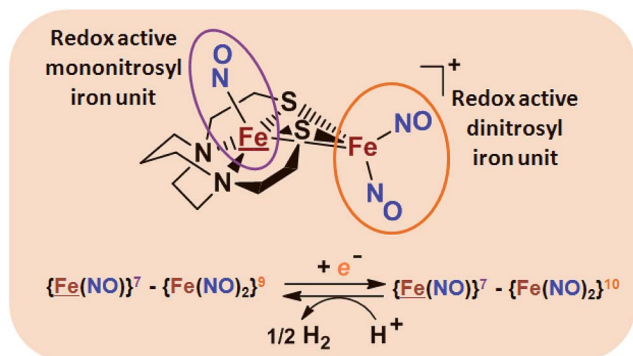


Fig. 1 Structure and redox activity of $[(\text{NO})\underline{\text{Fe}}\text{N}_2\text{S}_2\cdot\text{Fe}(\text{NO})_2]^+$, $[\underline{\text{Fe}}-\underline{\text{Fe}}]^+$; protonation of the one-electron reduced diiron complex yields H_2 .⁸

electron, reversible reduction events, -0.78 V and -1.33 V, assigned to $\{\text{Fe}(\text{NO})_2\}^{9/10}$ and $\{\underline{\text{Fe}}(\text{NO})\}^{7/8}$ couples, respectively.⁸ The Fe of the $(\text{NO})\underline{\text{Fe}}\text{N}_2\text{S}_2$ metalloligand is herein distinguished as $\underline{\text{Fe}}$; the electron count of the iron nitrosyl units uses the Enemark–Feltham approach.²⁸ Consistent with the stoichiometric reaction shown in Fig. 1, the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ couple, at -0.78 V, was the catalytically active center for electrochemical proton reduction in the presence of strong acid, $\text{HBF}_4\cdot\text{Et}_2\text{O}$. Although modest in overpotential and TOF, electrocatalysis for H_2 production was observed at this potential; preliminary computational studies indicated that a hydride-bound $\{\text{Fe}(\text{NO})_2\}^8$ could likely be a transient intermediate, however the overall H_2 releasing mechanism was at that stage incomplete.⁸ Interestingly, the second reduction process, related to the more negative $\{\underline{\text{Fe}}(\text{NO})\}^{7/8}$ couple, showed a current response to added weak acid, however H_2 was not detected. Computational study attributed this to a non-productive reduction event with protonation on the nitrosyl, which terminates the catalytic cycle.⁸

We have made analogues of the diiron trinitrosyl complex making use of NiN_2S_2 and $(\text{NO})\underline{\text{Fe}}\text{N}_2\text{S}_2$ metallodithiolates in combination with $\eta^5\text{-C}_5\text{R}_5$ derivatives ($\text{R} = \text{H}, \text{CH}_3$),^{9,10} of Fe^{II} shown in Fig. 2. The large differences in reduction potential of the MN_2S_2 ligands, with the $d^8\text{-Ni}^{\text{II}}$ being more negative because of a more rigid, less polarizable electronic structure as compared to the delocalized $\{\underline{\text{Fe}}(\text{NO})\}^7$ unit, of greater electronic flexibility, inspired the labels of “hard” for the former MN_2S_2 unit, and “soft” S-donor unit for the latter. The Fe-receivers also differ in electronic flexibility and their ease of electron uptake, the term “soft” describing the highly delocalized $\{\text{Fe}(\text{NO})_2\}^9$ unit, and the indefinite oxidation state of the iron, as compared to the definite Fe^{II} of the $\eta^5\text{-C}_5\text{R}_5$, “hard” receiver derivatives.^{8,9} The hard receiver unit, $(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})^+$, is herein distinguished from the soft $\text{Fe}(\text{NO})_2$ unit by $\underline{\text{Fe}}$ and Fe , respectively.

Notable from the computational approach that guided the interpretation of electrochemical events of the $[\underline{\text{Ni}}-\underline{\text{Fe}}]^+$ and $[\underline{\text{Fe}}-\underline{\text{Fe}}]^+$ complexes in the presence of acid was the indication of a reductive iron–sulfur bond cleavage (shown in red, Fig. 2) that converted the bidentate dithiolate into a monodentate S-donor, thus creating a transient frustrated Lewis pair, *i.e.*, yielding reactive sites for proton and electron uptake on the free thiolate and the open site on iron, respectively, see Fig. 2.⁹ In this way, complexes, that do not have an amine pendant base for proton uptake and storage, or open sites on iron for a hydride, as in the $[\text{FeFe}]-\text{H}_2\text{ase}$ active site,^{29–33} adjust their structures to accommodate coupled electron/proton uptake. While the mild potential for the first EC process for the $[\underline{\text{Fe}}-\underline{\text{Fe}}]^+$ complex required both proton/electron uptake for genesis of the pendant base, the more negative potential that reduces the $[\underline{\text{Ni}}-\underline{\text{Fe}}]^+$ labilizes the sulfur and creates an $\text{Fe}^{\text{III}}\text{-H}$ at the first reduction, Fig. 2.^{8,9}

In this report we provide a more complete matrix of $\text{MN}_2\text{S}_2\text{-Fe}$ complexes of electrocatalytic potential for experimental and computational analysis. Specifically a redox innocent (“hard”)

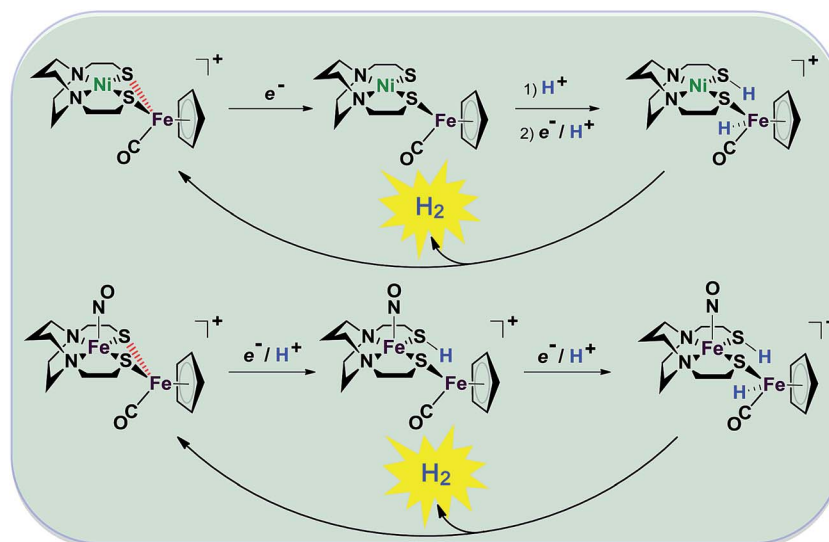


Fig. 2 Abbreviated computational mechanisms for electrocatalysis of H_2 production by the $[\underline{\text{Fe}}-\underline{\text{Fe}}]^+$ and $[\underline{\text{Ni}}-\underline{\text{Fe}}]^+$ electrocatalysts.⁹ Shown in red is the Fe–S bond that undergoes reductive bond cleavage.

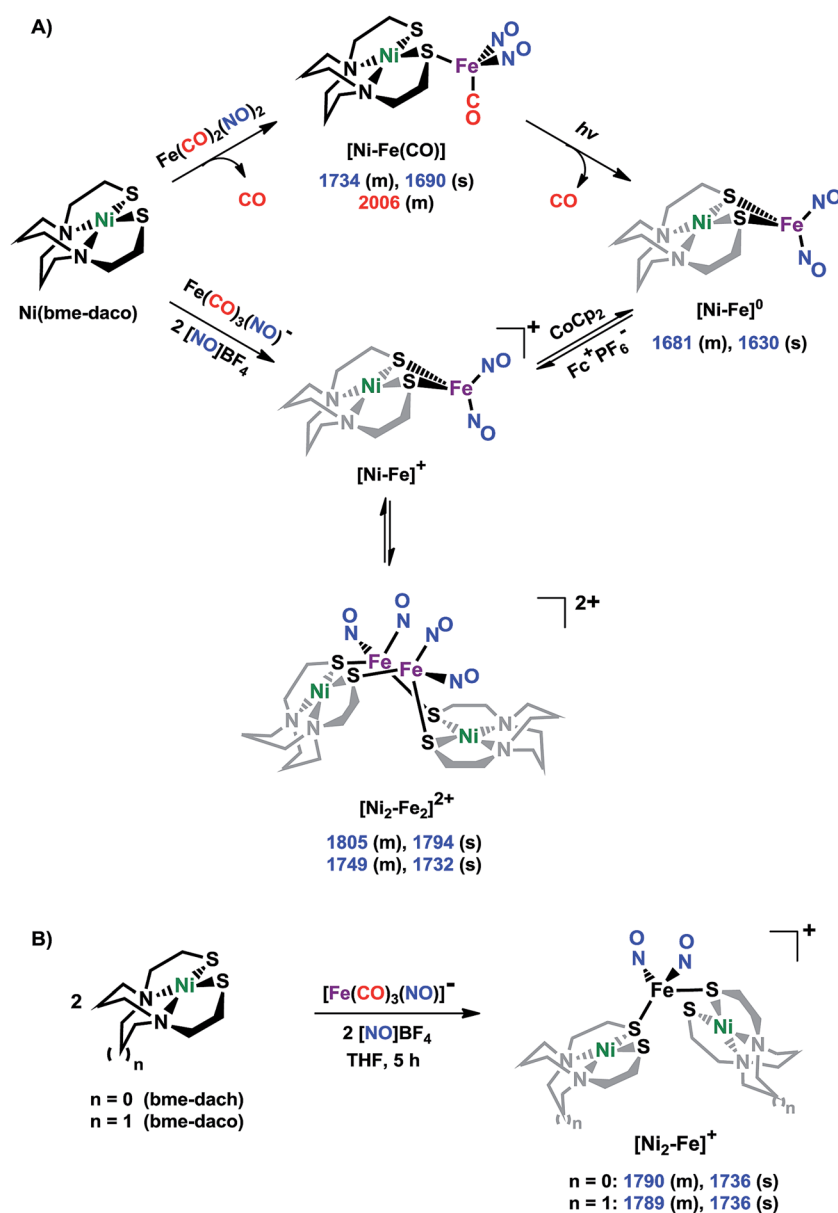


metalloligand, NiN_2S_2 , of more negative reduction potential, is incorporated in place of “soft” $(\text{NO})\text{FeN}_2\text{S}_2$. The thus generated $[\text{Ni}^{\text{II}}\text{N}_2\text{S}_2 \cdot \text{Fe}(\text{NO})_2]^{0/+}$, a “hard”/“soft” complex may be compared to the other members of the matrix. The solid state structures of $[\text{Ni}^{\text{II}}\text{N}_2\text{S}_2 \cdot \text{Fe}(\text{NO})_2]^{0/+}$ in two redox levels and characteristics as an electrocatalyst (robustness and turnover frequency), for proton reduction are also described. Computational study, addresses the diversity of geometries of di- and poly-metallic compounds containing N_2S_2 metalloligands by inspecting the versatile bonding orbitals of the metalloligands. The computational mechanisms contrast the working electrocatalysts against a non-working analogue by exploring possible intermediates in the proposed catalytic cycles. Here important roles for hemi-labile and redox active ligands are revealed.

Results and discussion

Synthesis and characterization

Shown in Scheme 1A are the synthetic routes to $\text{NiN}_2\text{S}_2 \cdot \text{Fe}(\text{NO})_2$, $[\text{Ni}-\text{Fe}]^0$, and its one-electron oxidized analogue, $[\text{Ni}-\text{Fe}]^+$, isolated and crystallographically characterized as a dimer, $[\text{NiN}_2\text{S}_2 \cdot \text{Fe}(\text{NO})_2]_2^{2+}$ or $[\text{Ni}_2-\text{Fe}_2]^{2+}$, ($\text{N}_2\text{S}_2 = N,N$ -bis(2-mercaptoethyl)-1,5-diazacyclooctane or bme-daco). Infrared values listed for the diatomic ligands were recorded in CH_2Cl_2 or THF solution. Freshly prepared $\text{Fe}(\text{CO})_2(\text{NO})_2$ in THF readily reacts with NiN_2S_2 at 22°C , with replacement of one CO, releasing the second CO under photolysis, or within 20 min at 40°C , thus converting the NiN_2S_2 from mono- to bidentate ligand.^{34–37}



Scheme 1 (A) The syntheses of $[\text{Ni}-\text{Fe}]^0$ and $[\text{NiFe}(\text{CO})]^0$ as neutral complexes, and $[\text{Ni}_2-\text{Fe}_2]^{2+}$ and (B) $[\text{Ni}_2-\text{Fe}]^+$ as BF_4^- salt. The IR values (in cm^{-1}) of CO and NO are in red and blue, respectively.



From this approach the $[\text{Ni-Fe}]^0$ complex was isolated as a brown solid that is stable at ambient temperature under Ar. Oxidation of $[\text{Ni-Fe}]^0$ by Fc^+PF_6^- at 0 °C resulted in a color change of the THF solution from brown to dark purple with concomitant shifts of the $\nu(\text{NO})$ values by an average of *ca.* 110 cm^{-1} to higher wave numbers. The reversibility of this oxidation was confirmed by reaction with cobaltocene and return to the reduced $[\text{Ni-Fe}]^0$. The $\nu(\text{NO})$ bands listed under $[\text{Ni}_2\text{-Fe}_2]^{2+}$, Scheme 1A, reflect the presence of overlapping components which were resolved into two sets of absorbances, interpreted as a likely mixture of monomeric cation and dicationic dimer, with the set at lower values slightly less in intensity (Fig. S24†). As other experimental data, *vide infra*, as well as computational studies, indicate the predominance of monomeric $[\text{Ni-Fe}]^+$, we postulate that the set of absorbances at slightly lower wavenumbers (as shoulders on the major bands) are due to the dimeric $[\text{Ni}_2\text{-Fe}_2]^{2+}$. We note that the electron-spray ionization mass spectrum of $[\text{Ni-Fe}]^+$ has a parent ion with isotopic bundle distribution at *m/z* that is consistent with a monomeric $[\text{Ni-Fe}]^+$ rather than a dimeric $[\text{Ni}_2\text{-Fe}_2]^{2+}$, Fig. S28.† The difference between two consecutive isotopic mass units is ~ 1 , rather than 0.5, which indicates the predominance of the monomer, $[\text{Ni-Fe}]^+$, in the polar solvents in which they are soluble.

The magnetic moments of $[\text{Ni-Fe}]^+$ and $[\text{Ni}_2\text{-Fe}_2]^{2+}$ are 1.54 B.M. and 1.76 B.M., respectively, measured by Evans' method at room temperature in CD_2Cl_2 . This is consistent with the presence of a single unpaired electron, Fig. S1 and S2.† The EPR spectra for both complexes display the isotropic $g = 2.03$ signal that is prototypical of the $\{\text{Fe}(\text{NO})_2\}^9$ unit, Fig. S21 and S22,† respectively. The 77 K EPR spectrum of the $[\text{Ni-Fe}]^+$ displayed fine structure requiring two signals for simulation: A major isotropic signal of $g = 2.035$ showed coupling with nitrogen of $A(^{14}\text{N}) = 32.74$ MHz and a minor anisotropic signal had parameters of $g_{\text{xyz}} = 2.183, 2.012, 1.908$ and no observable hyperfine coupling, Fig. S21.†

X-ray diffraction quality crystals of the oxidized NiFe compound were obtained from the one-pot reaction of equimolar NiN_2S_2 and (putative) $[\text{Fe}(\text{CO})_2(\text{NO})_2]^+$ (prepared *in situ* by reacting $[\text{Fe}(\text{CO})_3(\text{NO})]^-$ with two equivalents of $[\text{NO}]\text{BF}_4$)³⁸ in CH_2Cl_2 at ambient temperature, Scheme 1. A third Ni-Fe complex, $[\text{Ni}_2\text{-Fe}]^+$, was obtained on combining NiN_2S_2 with $[\text{Fe}(\text{CO})_2(\text{NO})_2]^+$ in 2 : 1 ratio in THF solution, Scheme 1B. X-ray quality crystals of this trimetallic as its BF_4^- salt were obtained from hexane/THF layering at -28 °C. Its $\nu(\text{NO})$ IR spectral

features are typical of monomeric DNICs in the $\{\text{Fe}(\text{NO})_2\}^9$ redox level.

X-ray diffraction studies

The molecular structures of the heterometallic complexes $[\text{Ni-Fe}]^0$, $[\text{Ni}_2\text{-Fe}_2]^{2+}$ and $[\text{Ni}_2\text{-Fe}]^+$ are shown in Fig. 3. The bimetallic complex $[\text{Ni-Fe}]^0$, exhibits an overall butterfly-like $\text{Ni}(\mu\text{-SR})_2\text{Fe}$ core, analogous to the report of Pohl *et al.*, where an open chain N_2S_2 ligand chelated the Ni^{II} .³⁶ The converging lone pairs (see below) on the *cis*-dithiolates engage in bidentate binding and impose a hinge angle (the intersection of the best N_2S_2 plane with the S_2Fe plane) of *ca.* 117°, *vis-à-vis* constricting the $\angle \text{S-Ni-S}$ angle by *ca.* 4° compared to the free metal-ligand.³⁹ The two nitrosyl units bound to the pseudo tetrahedral iron center are slightly bent towards each other, in an "attracto" orientation;⁴⁰ the average $\angle \text{Fe-N-O}$ angle is $\sim 163.8^\circ$. The $\text{Ni}\cdots\text{Fe}$ distance of 3.001 (2) Å, is slightly longer (by 0.022 Å) than that obtained in the Pohl, *et al.* structure.³⁶

The oxidized NiFe complex crystallizes as dimeric $[\text{Ni}_2\text{-Fe}_2]^{2+}$ with two BF_4^- anions; two dinitrosyl iron units are bridged by two NiN_2S_2 metalloligands. The tetrahedral geometry about each $\text{Fe}(\text{NO})_2$ unit is thus completed by two thiolates from different NiN_2S_2 metalloligands, thus generating an abbreviated paddlewheel, or propeller type, complex seen in examples of nickel-gold tetrametallic complexes.⁴¹ Likewise, the molecular structure of $[\text{Ni}_2\text{-Fe}]^+$ demonstrates that one $\text{Fe}(\text{NO})_2$ unit bridges two NiN_2S_2 metalloligands, each acting as a monodentate ligand. As shown in the $[\text{Ni}_2\text{-Fe}]^+$ structure, Fig. 3, the free unbound thiolates of two NiN_2S_2 units are transoid to each other. The addition of a second $\text{Fe}(\text{NO})_2$ unit to generate the dication, $[\text{Ni}_2\text{-Fe}_2]^{2+}$, would require rotation about one Fe-S bond in order to align the two metalloligands.

The average $\text{Ni}\cdots\text{Fe}$ distances in $[\text{Ni}_2\text{-Fe}_2]^{2+}$ and $[\text{Ni}_2\text{-Fe}]^+$, are 3.680 (2) Å and 3.521 (2) Å, respectively, and are longer than in the $[\text{Ni-Fe}]^0$ reduced complex by *ca.* 0.5 Å. The Ni^{II} maintains a square planar geometry in the reduced and oxidized complexes with a displacement of no more than 0.1 Å from the N_2S_2 best plane. Overall these structures demonstrate the impressive adaptability of the NiN_2S_2 metalloligands, and their potential to template clusters through S-based aggregation.⁷

Computational structural study

This computational section uses density functional theory (DFT) analysis to address the structural question in particular

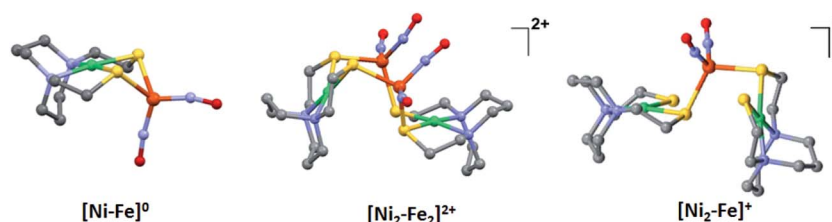


Fig. 3 Molecular structures of (a) $[\text{Ni-Fe}]^0$, (b) $[\text{Ni}_2\text{-Fe}_2]^{2+}$ and (c) $[\text{Ni}_2\text{-Fe}]^+$ from X-ray diffraction. The BF_4^- counter ions of $[\text{Ni}_2\text{-Fe}_2]^{2+}$ and $[\text{Ni}_2\text{-Fe}]^+$ are omitted for clarity.



that was raised by the X-ray diffraction report: is there an electronic factor that governs the different μ_2 -SR binding modes found in the three forms of NiFe heterometallic aggregates? The functional/basis set combination, TPSS/6-311++G(d,p), and natural bond orbital (NBO) analysis were applied to the computational structural modeling of the free metalloligand NiN_2S_2 and its derivatives $[\text{Ni-Fe}]^0$, $[\text{Ni}_2\text{-Fe}_2]^{2+}$; more details of the computational methodology is available in ESI.†

The divergent or convergent orientation of S lone pairs of NiN_2S_2 metalloligand and influences on structures of $\text{NiN}_2\text{S}_2 \cdot \text{M}'$ heterobimetallics. Traditional bidentate ligands such as diphosphines, diamines and bipyridyls have a single lone pair on each donor site. These lone pairs are positioned on orbitals originating from sp^x -hybridization and are highly directional.⁴² They provide fixed binding orientations that correspond one-to-one with the coordination sites of the metal. In contrast, the binding between the sulfurs of the metathiolate NiN_2S_2 and an exogenous metal are more geometrically flexible because of the multiple S lone pairs. From NBO bonding analysis, sulfur in the NiN_2S_2 metalloligand is found to use mainly p orbitals for bonding to Ni and C.^{43,44} For example, in a free NiN_2S_2 , p character makes up 83% and 86% of the S contributions in the S–Ni bonds and S–C $_{\alpha}$ bonds (C $_{\alpha}$ and C $_{\beta}$ refer to the C $_2$ H $_4$ linker connecting S and N where C $_{\alpha}$ is directly bound to S, Fig. 4A), which leaves one lone pair in a p orbital and another in an s-dominated orbital on each S. Because a receiver group, a $\text{Fe}(\text{NO})_2$ unit in our case, may bind to either lobe of the p lone pair(s), whose orientation is determined by the Ni–S–C $_{\alpha}$ torsion angle, a diversity of structures results.^{7,26}

The orientation of this remaining p lone pair in the NiN_2S_2 metalloligand is determined by the NiN_2S_2 metalloligand's Ni–S–C $_{\alpha}$ –C $_{\beta}$ –N five-membered rings that adopt a non-planar envelope conformation like cyclopentane. The C $_{\alpha}$ carbon, the “flap” of the envelope conformation, puckers towards one side or the other of the N $_2\text{S}_2$ plane, Fig. 3. Fig. 4 shows how this puckering tilts the remaining 3p-lone pair on each sulfur from the normal to the N $_2\text{S}_2$ plane. This tilt causes two p-orbital lobes (green lobes) to converge on the side to which the flap puckers, while the red lobes diverge on the opposite side. The orientation of the added $\text{Fe}(\text{NO})_2$ receiver unit(s), will be thus determined by such directional property of the donor p lone pairs. The structure of the reduced monomer $[\text{Ni-Fe}]^0$ shows the $\text{Fe}(\text{NO})_2$ fragment is on the same side as the flap; while in the oxidized

dimer $[\text{Ni}_2\text{-Fe}_2]^{2+}$ the flap and the $\text{Fe}(\text{NO})_2$ fragment(s) appear on different sides of each N $_2\text{S}_2$ plane, thus, binding to the other end of the p lone pair. Based on the analysis above, the converging lobes of the p donor lone pairs maximize contact to the $\text{Fe}(\text{NO})_2$ unit in the monomer $[\text{Ni-Fe}]^0$, while the diverging lobes are preferred by two bridging $\text{Fe}(\text{NO})_2$ units between two metalloligands in the dimer $[\text{Ni}_2\text{-Fe}_2]^{2+}$. The utilization of the divergent lobes apparently lessens the steric repulsion between $\text{Fe}(\text{NO})_2$ units. In summary, the binding position of the $\text{Fe}(\text{NO})_2$ unit with respect to the flap in the Ni–S–C $_{\alpha}$ –C $_{\beta}$ –N five-membered rings are correlated by the competition between chemical bond directionality of the binding sulfurs and steric repulsion of the receiver units.

Electrochemistry

The cyclic voltammograms of $[\text{Ni-Fe}]^+$ (in CH_2Cl_2), $[\text{Ni}_2\text{-Fe}]^+$ (in CH_3CN), as BF_4^- salts, and $[\text{Ni-Fe}]^0$ (in CH_2Cl_2), were recorded under Ar at 22 °C, and referenced to $\text{Fc}^{0/+}$ ($E_{1/2} = 0.0$ V) as an internal standard. Both the neutral complex $[\text{Ni-Fe}]^0$ and the cationic analogue, $[\text{Ni-Fe}]^+$, used in the CV studies as its BF_4^- salt, displayed reversible reduction events at ca. -0.73 V (CH_2Cl_2), assigned to the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ couple; see Fig. S13–S15† for CV scans. The NiFe complexes also present two irreversible oxidation events at ca. -0.10 V and ca. $+0.45$ V, with minor differences in intensities according to the neutral or cationic sources. Both of these events are assumed to be S-based.

In CH_3CN , the trimetallic complex $[\text{Ni}_2\text{-Fe}]^+$, showed a reversible event at, $E_{1/2} = -0.75$ V, assigned to the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ couple and an irreversible oxidation event at $E_{1/2} = -0.05$ V, see Fig. S16.† The $E_{1/2}$ value for the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ couple, is anodically shifted by ca. 30 mV relative to the 1 : 1, NiFe complexes, resulting from the greater electron donation of two NiN_2S_2 centers to the $\text{Fe}(\text{NO})_2$ redox marker. The scan rate dependences of the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ couple for all three complexes support the assignments to reversible or quasi-reversible as described above, see Fig. S17–S19.†

Cyclic voltammetry and response to added acid. Electrochemical studies of $[\text{Ni-Fe}]^+$ and $[\text{Ni}_2\text{-Fe}]^+$ were carried out in presence of $\text{HBF}_4 \cdot \text{Et}_2\text{O}$ under a N $_2$ /Ar atmosphere. For comparison the $[\text{Fe-Fe}]^+$ complex was examined under similar experimental conditions. Sequential addition of $\text{HBF}_4 \cdot \text{Et}_2\text{O}$ to a CH_2Cl_2 solution of $[\text{Ni-Fe}]^+$ (2 mM $[\text{Ni}_2\text{-Fe}_2]^{2+}$) showed an increase in the cathodic current at the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ redox event at -0.73 V. The initial cathodic current response at -0.8 V saturates with ~ 20 equivalents of the acid, Fig. 5 (inset). A second rise in cathodic current at -1.10 V, commences upon addition of >12 equivalents of the acid, which continues to rise as the catalytic current response, Fig. 5. The first response is attributed to the reduction of $[\text{Ni-Fe}]^+$ followed by a protonation. The second response is assigned to the up-take of another electron by the reduced and protonated counterpart of $[\text{Ni-Fe}]^+$. The mechanism below connects the successive protonation to the production of H_2 , thus closing the catalytic cycle. Overlays of this response of the NiFe complex in presence of 50 equivalents of $\text{HBF}_4 \cdot \text{Et}_2\text{O}$ (0.1 M), as well as the CV of the free acid, are shown in Fig. 5. The catalytic H_2 produced was confirmed by

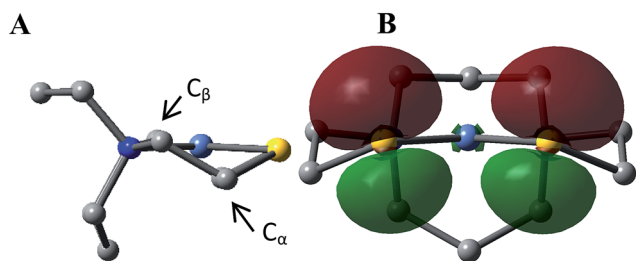


Fig. 4 (A) The geometry of a free metalloligand NiN_2S_2 and (B) its two 3p lone pairs presented one on each sulfur (contour plots at isovalue = 0.05 a.u. by NBO analysis). Note the relative positions of the S–C $_{\alpha}$ /S–Ni bonds and the 3p-lone pair.



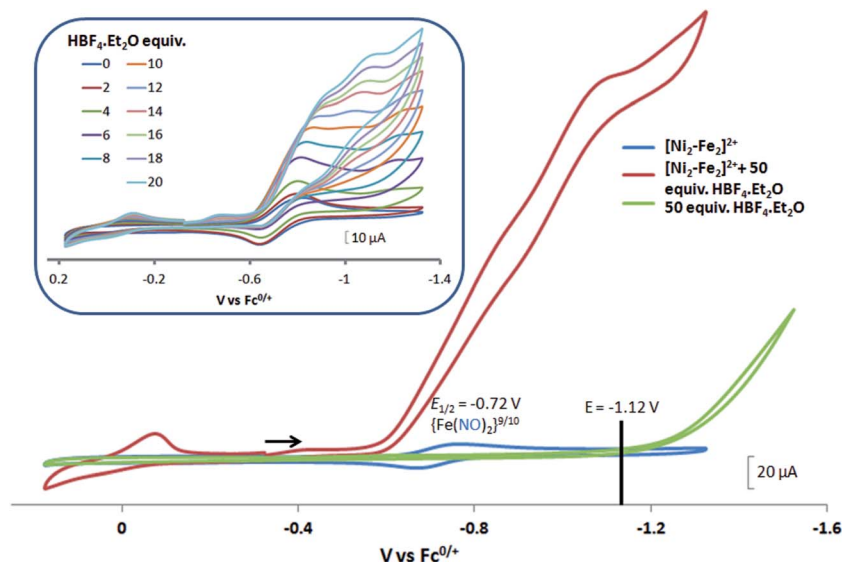


Fig. 5 Cyclic voltammograms of 2.0 mM CH_2Cl_2 solutions of $[\text{Ni}_2\text{-Fe}_2]^{2+}$ (or $[\text{Ni-Fe}]^+$ (blue)); with 50 equiv. (0.1 M) of added $\text{HBF}_4\cdot\text{Et}_2\text{O}$ (red); and, for reference, 50 equiv. (0.1 M) of $\text{HBF}_4\cdot\text{Et}_2\text{O}$ (green). The black line indicates the constant potential applied for bulk electrolysis, -1.12 V. Inset: cyclic voltammograms of $[\text{Ni}_2\text{-Fe}_2]^{2+}$ (or $[\text{Ni-Fe}]^+$) with 2 to 20 equiv. aliquots of $\text{HBF}_4\cdot\text{Et}_2\text{O}$. Crystalline $[\text{Ni}_2\text{-Fe}_2]^{2+}$ was dissolved as its BF_4^- salt, in 0.1 M $^t\text{Bu}_4\text{NPF}_6$ as supporting electrolyte, with a glassy carbon electrode at scan rate of 200 mV s^{-1} . Note: equivalents of $\text{HBF}_4\cdot\text{Et}_2\text{O}$ was calculated with respect to the dimeric $[\text{Ni}_2\text{-Fe}_2]^{2+}$.

applying a constant potential at -1.12 V for 60 min (black bold line in Fig. 5), and analysis of the headspace by gas chromatography. The H_2 was quantified by an average of two consistent constant potential coulometry experiments with subtraction of the H_2 produced from the free acid.^{9,10} The nitrosylated compounds $[\text{Ni-Fe}]^+$ and $[\text{Fe-Fe}]^+$ were found to have low turnover numbers and faradaic efficiencies, $68 \pm 2\%$ and $58 \pm 1\%$, respectively, for H_2 production. In contrast the $[\text{Ni-Fe}]^+$ gave a faradaic efficiency of ca. 96%. We assume that the former involves alternate protonation pathways, particularly at NO, that lead to degradation and hence low F.E. In addition the TON for the robust $\text{Ni-Fe}'$ complex in 50 equiv. of TFA, measured at -1.73 V and over a period of 8 h was found to be 6.7, assuring catalytic proton reduction. The electrocatalytic response of the reduced complex, $[\text{Ni-Fe}]^0$ in the presence of $\text{HBF}_4\cdot\text{Et}_2\text{O}$, is, as expected, the same as $[\text{Ni-Fe}]^+$ and is shown in Fig. S20-B.†

Following the approach of Helm and Appel,⁴⁵ and Wiese,⁴⁶ the turnover frequency (TOF) as calculated from the CV experiment for $[\text{Ni-Fe}]^+$, was 39.7 s^{-1} , which is slightly better than the $[\text{Fe-Fe}]^+$ complex, 26.7 s^{-1} , calculated under similar experimental conditions. The $[\text{Ni-Fe}]^+$ shows a saturation of the more negative catalytic current upon addition of 80 equivalents of $\text{HBF}_4\cdot\text{Et}_2\text{O}$, i.e., ~ 0.16 M CH_2Cl_2 solution. Notably, the behavior of the $[\text{Fe-Fe}]^+$ complex is similar, and further addition of acid leads to decomposition of both catalysts. The precise calculation of overpotential is indeterminable as the thermodynamic potential ($E_{\text{HBF}_4/\text{H}_2, \text{BF}_4^-}$) of 0.1 M $\text{HBF}_4\cdot\text{Et}_2\text{O}$ in CH_2Cl_2 is unavailable.⁴⁷ Using the thermodynamic potential of $\text{HBF}_4\cdot\text{Et}_2\text{O}$ in acetonitrile (-0.26 V),^{48,49} an estimate of the overpotential of $[\text{Ni-Fe}]^+$ and $[\text{Fe-Fe}]^+$ were 711 mV and 660 mV, respectively, which are lower than those of the $[\text{Ni-Fe}]^+$ and $[\text{Fe-Fe}]^+$ electrocatalysts by over 220 mV.

In contrast to the NiFe complexes, addition of $\text{HBF}_4\cdot\text{Et}_2\text{O}$ to a 2.0 mM CH_3CN solution of $[\text{Ni}_2\text{-Fe}]^+$, (the N_2S_2 ligand used in this electrochemical study is bme-dach) did not show an increase in the cathodic current at -0.75 V, the reversible $\{\text{Fe}(\text{NO})_2\}^{9/10}$ redox event. Instead, a new reversible redox event at $E_{1/2} = -0.52$ V, appeared upon addition of two equivalents of $\text{HBF}_4\cdot\text{Et}_2\text{O}$ with a concomitant disappearance of the original redox process, Fig. 6. Further addition of acid resulted in electrode fouling, Fig. S20-A.† A possible explanation, from computational chemistry, *vide infra*, for the positive 230 mV shift is that $[\text{Ni}_2\text{-Fe}]^+$ can be protonated on its exposed thiolate sulfur by $\text{HBF}_4\cdot\text{Et}_2\text{O}$, *vide infra*. Such would account for the greater ease of reduction for the $\{\text{Fe}(\text{NO})_2\}^{9/10}$ couple, compared to the $[\text{Ni}_2\text{-Fe}]^+$ complex. Supporting this conclusion is that addition of 1 equivalent of $\text{HBF}_4\cdot\text{Et}_2\text{O}$ to a CH_3CN solution of $[\text{Ni}_2\text{-Fe}]^+$ produced a small but definite shift of the $\nu(\text{NO})$ in the IR spectrum from 1787 and 1734 cm^{-1} to 1793 and 1737 cm^{-1} , Fig. S27.† In addition, the irreversible oxidation event at 0.07 V, which is assumed to be sulfur-based oxidation, shows a decrease in the anodic current upon addition of acid, indicating disulfide formation is inhibited under acidic conditions.

Computational mechanistic study

The electrochemical study points to additional questions for computational study: (A) how do the calculated electrocatalytic mechanisms compare for the hard-soft *vs.* soft-soft donor/receiver adducts? (B) Can computational analysis clarify those cases of non-catalytic electrochemical responses to added protons? Modeling is extended to $[\text{Ni}_2\text{-Fe}]^+$, along with $[\text{Ni-Fe}]^0$, $[\text{Ni-Fe}]^+$, in various oxidation states and with multiple added protons to answer these questions.



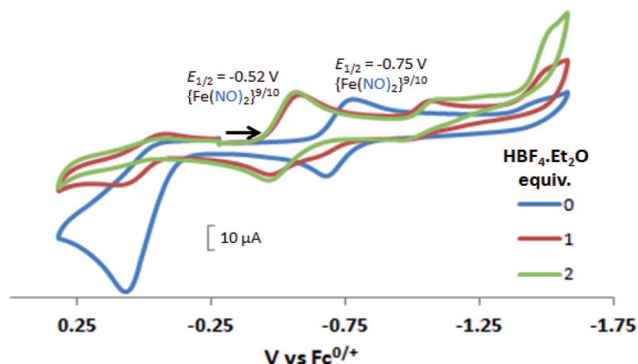


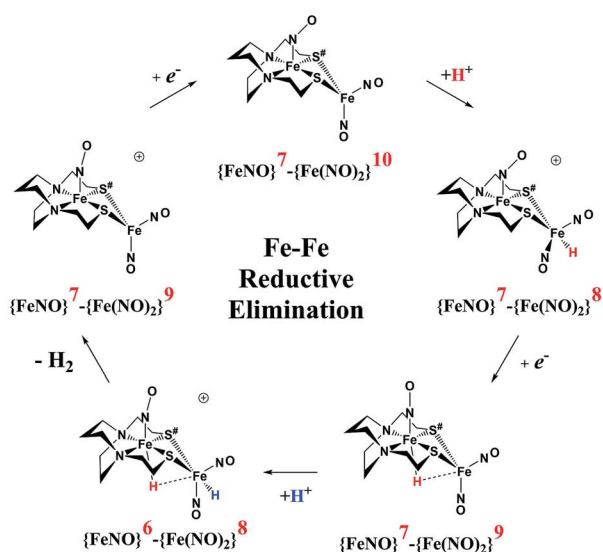
Fig. 6 Cyclic voltammograms of 2.0 mM CH_3CN solution of $[\text{Ni}_2\text{-Fe}]^+$ (blue), with 1 and 2 equiv. of added $\text{HBF}_4 \cdot \text{Et}_2\text{O}$ (red and green, respectively). Note: the N_2S_2 ligand used in this compound is bme-dach.

Mechanistic perspectives of the proton reduction electrocatalysis by $[\text{Ni-Fe}]^+ / [\text{Ni-Fe}]^0$ and comparison to $[\text{Fe-Fe}]^+$. An earlier computational analysis of a HER electrocatalysis mechanism proposed for the $[\text{Fe-Fe}]^+$ complex, Fig. 1, focused on the first reduction event with a strong acid proton source.⁸ The successive reduction event ultimately allowed for double proton addition to the $\text{Fe}(\text{NO})_2$ unit and formation of a dihydride, panel A of Fig. 7.^{8,50} Making use of electron transfer from the reduced $\{\text{Fe}(\text{NO})\}^8$, reductive elimination from the $\{\text{Fe}(\text{NO})\}^6\text{-}\{\text{Fe}(\text{NO})_2\}^8$ morphed into an $\eta^2\text{-H}_2\text{-Fe}(\text{NO})_2$, restoring $\{\text{Fe}(\text{NO})\}^7\text{-}\{\text{Fe}(\text{NO})\}^9$, with H_2 formation and loss. Note that no hemi-lability of the metallodithiolate ligand⁹ is necessary here as the mechanism does not entail hydride/proton coupling to H_2 , but rather reductive elimination from two hydrides.⁷

The $[\text{Ni-Fe}]^+$ and its reduced counterpart $[\text{Ni-Fe}]^0$ are determined to be electrocatalysts at -0.73 V for H_2 production with $\text{HBF}_4 \cdot \text{Et}_2\text{O}$, Fig. 5. [Note: The computational study finds that the $[\text{Ni}_2\text{-Fe}_2]^{2+}$, whose dimeric structure was established in the solid state by crystallography, finds greater stability in solution as the monomeric form, $[\text{Ni-Fe}]^+$. Experimental evidences including ESI-MS and determination of μ_{eff} support this thesis, *vide supra*.] The catalytic cycle is thus initiated with the monomer $[\text{Ni-Fe}]^+$ (Fig. 7B). As indicated in panel B of Fig. 7, the $\{\text{Fe}(\text{NO})_2\}^9$ in the $[\text{Ni-Fe}]^+$ unit accepts the first incoming electron, followed by the first proton, to create a hydride on the now $\{\text{Fe}(\text{NO})_2\}^8$ unit. Addition of a second electron activates the hemi-lability of the bridging thiolate to break one S-Fe bond, while the terminal hydride becomes bridging between Fe and Ni. The cleavage of the S-Fe dative bond essentially releases one p lone pair of the thiolate so that S can act as a pendant base to accept the second proton and guide it to a coupling position with the hydride and produce H_2 . Details of the full catalytic cycle with energetics and analysis of electronic structure evolution for both $[\text{Fe-Fe}]^+$ and $[\text{Ni-Fe}]^+$ are presented in a separate report.⁵⁰

Explanation for the absence of catalytic activity of $[\text{Ni}_2\text{-Fe}]^+$. While one might have expected the dangling thiolates in the 2 : 1 complex $[\text{Ni}_2\text{-Fe}]^+$ to act as a pendant base, in fact this complex does not show any catalytic activity in the presence of strong acid, $\text{HBF}_4 \cdot \text{Et}_2\text{O}$, within the solvent potential window. A computational study, summarized in Fig. 8, reveals that while reduction still occurs on the $\text{Fe}(\text{NO})_2$ unit, the protonation process is diverted from the $\text{Fe}(\text{NO})_2$ unit. In this 2 : 1 complex, the computations show that only one thiolate from each NiN_2S_2 binds to $\text{Fe}(\text{NO})_2$, while the other thiolate, is “free” to interact with other electron acceptors; thus it may also be protonated, even before the reduction of the $\{\text{Fe}(\text{NO})_2\}^9$ unit occurs, see

A Soft donor unit – Soft receiver unit



B Hard donor unit – Soft receiver unit

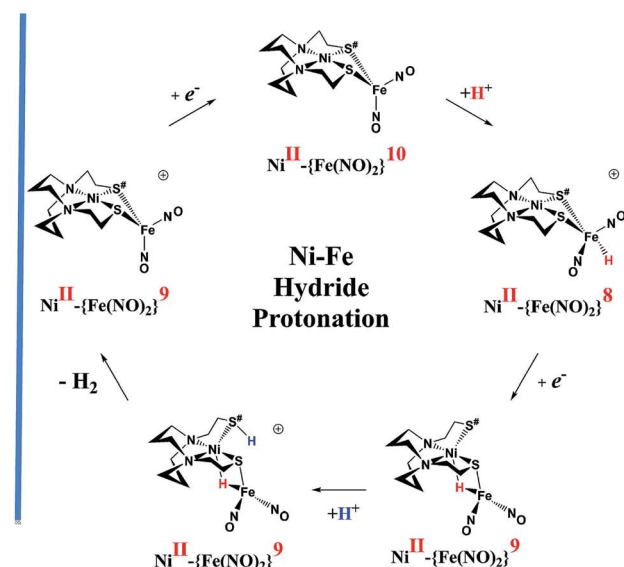
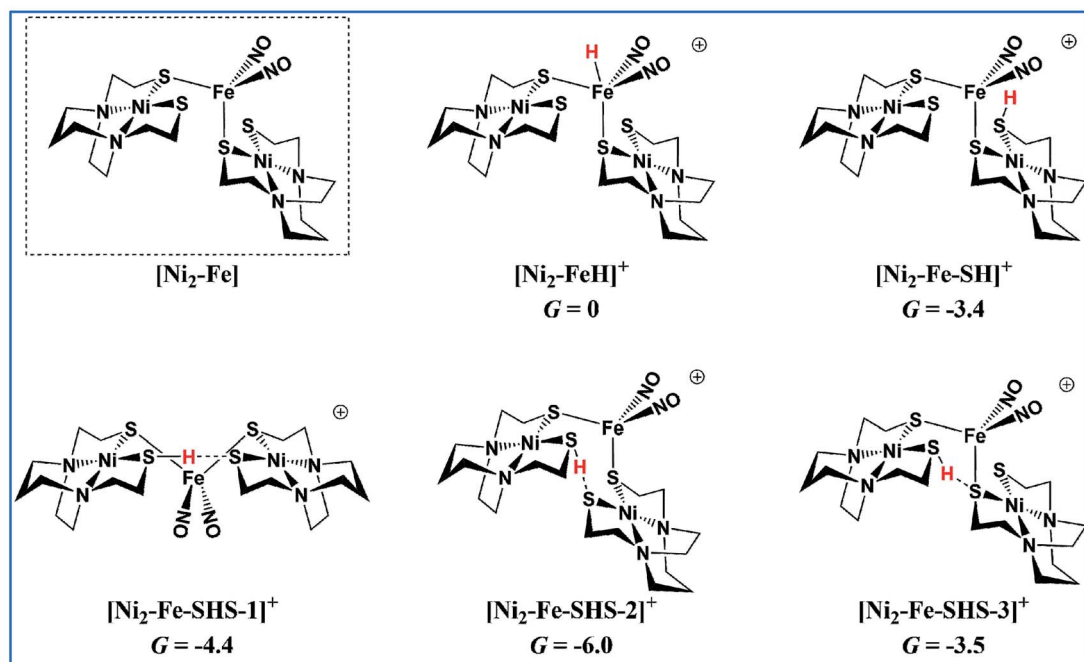


Fig. 7 Comparative catalytic cycles for H_2 production catalyzed by $[\text{Fe-Fe}]^+$ and $[\text{Ni-Fe}]^+$. All pK_a , thermodynamic, and metric data for the two mechanisms are available in a separate report.⁵⁰



A



B

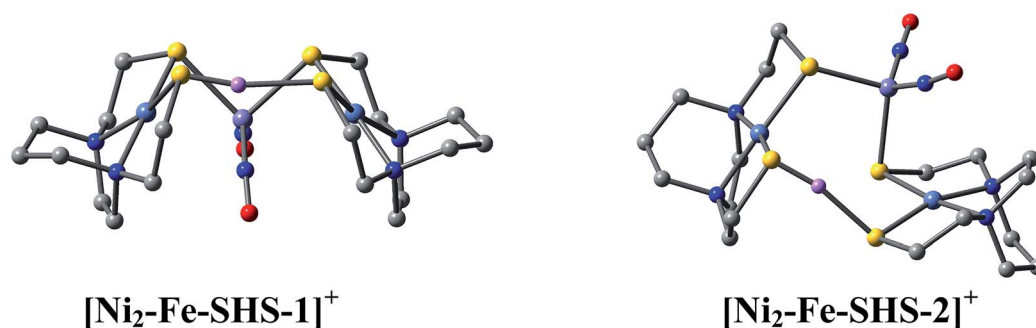


Fig. 8 The protonation of $[\text{Ni}_2\text{-Fe}]$, the reduced form of $[\text{Ni}_2\text{-Fe}]^+$ and (a) possible protonated products with (b) 3D geometric presentations of selected species featuring the pinched proton. The computationally derived structures are rendered so as to show the NiN_2S_2 metalloligand without altering the rigidity of the N_2S_2 planar structure. All hydrogens except the one pinched between two sulfurs are omitted.

Table S9,[†] and shifts the its potential, which is supported by experiment, *vide supra*.

According to the computations, in the reduced $[\text{Ni}_2\text{-Fe}]^0$ the “free” thiolate competes with the reduced $\{\text{Fe}(\text{NO})_2\}^{10}$ unit for the incoming proton (Fig. 8A); in addition, by rotation around an Fe–S bond, the two NiN_2S_2 ligands may orient their “free” thiolate sulfurs to pinch the proton, *i.e.*, consequently forming a strong hydrogen bond (Fig. 8A and B). Spectroscopic evidence supports protonation on S even before reduction, *i.e.*, in $[\text{Ni}_2\text{-Fe}]^+$, Fig. S27.[†] Two geometries of the pinched proton by two “free” thiolates, $[\text{Ni}_2\text{-Fe-SHS-1}]^+$ and $[\text{Ni}_2\text{-Fe-SHS-2}]^+$ can be achieved by either translating or rotating one NiN_2S_2 unit of $[\text{Ni}_2\text{-Fe}]$, respectively. Precedent in Dubois’ $\text{Ni}(\text{P}_2\text{N}_2)_2$ catalysts,⁵¹ a proton pinched between two amine N bases is relatively stable; in our case, the pinched proton is even more stable than a hydride on $\text{Fe}(\text{NO})_2$ (Fig. 8A). However, the mechanistic clue from the $[\text{Ni-Fe}]$ complex⁵⁰ indicates the requirement for

a proton to be reduced into a hydride, by $\{\text{Fe}(\text{NO})_2\}^{10}$, before the H_2 can be produced by the proton–hydride coupling mechanism. Therefore, the formation of a stable pinched proton likely prevents the generation of the hydride and cuts off the catalytic cycle. The thiolate already bound to $\text{Fe}(\text{NO})_2$ also helps stabilize the proton on a “free” thiolate, to a smaller extent, with the example of $[\text{Ni}_2\text{-Fe-SHS-3}]^+$ (Fig. 8A).

Conclusions

Our collection of hydrogen evolution reaction catalysts is summarized in Fig. 9. While the small differences in donor units and acceptor units do not influence the overall structures of the S-bridged bimetallics; all have butterfly-like $[\text{M}(\mu\text{-SR})_2\text{Fe}]$ core and the potential for opening up sites for proton addition *via* the hemi-lability of the metallothiolate donors. Nevertheless, demonstrable and explicable differences are seen in their



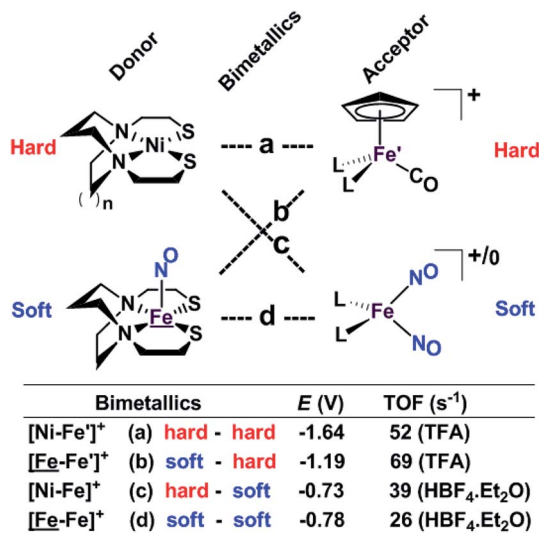


Fig. 9 A comparative schematic for a matrix of bimetallic electrocatalysts containing hard/soft donor/acceptor units.

catalytic performances as indicated by catalytic potential, required acid strength, and TOF.

Analogous to the HSAB (Hard and Soft (Lewis) Acids and Bases) concept, we offer an electronic parallel, “soft vs. hard donor/receiver units”, in this case directed towards the number of NO ligands in the bimetallics ranging from 0 to 3, with increasing flexibility (*i.e.*, soft) of electronic structure within each unit. The non-innocence of the NO ligand confers electron uptake at milder potentials, which we have seen in both the donor units and acceptor units. Thus the incorporation of NO ligands on the acceptor units, the ‘hard–soft’ and ‘soft–soft’ electrocatalysts lead to energetically more accessible catalytic current, however, at the cost of a stronger acid and a lower TOF in comparison to the bimetallics with hard acceptor units.

While these electrocatalysts are only moderately efficient for H₂ production, they are well-behaved and have demonstrated reproducibility. Two of the catalysts, c and d, with soft receivers, are isolated and crystallized in both oxidized and reduced forms at ambient conditions lending confidence to the presumed catalytic cycle.

Features in the electrochemical scans may be reasonably ascribed to protonation products whose identities are further described by computational chemistry. The resulting computational mechanisms identify key features that may guide future synthetic targets. For example, the hemi-lability of the S-donors may be optimized by steric constraints; the usefulness of the Fe(NO)₂ unit as electron depot and protonation site with low redox potential, should encourage explorations with other redox-active, soft acceptors. The computations also suggest a mechanistic paradigm of heterolytic H⁻/H⁺, hydride–proton, coupling for bimetallics a, b and c from the chart, and reductive elimination from d arising in the soft–soft construct. Such a supposition derives from extreme electron delocalization in the trinitrosylated [Fe–Fe]⁺ complex and argues that suitably constructed first row, bimetallic complexes may take on two-electron processes that emulate noble metals.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

We are thankful for financial support from the National Science Foundation (CHE-1266097, CHE-1665258 to M. Y. D., and CHE-1300787, CHE-1664866 to M. B. H.), the Robert A. Welch Foundation (A-0924 to M. Y. D., and A-0648 to M. B. H.). A portion of the salary of Pokhraj Ghosh was made possible by an NPRP award (NPRP 6-1184-1-224) from the Qatar National Research Fund (a member of Qatar Foundation). The authors acknowledge the Laboratory for Molecular Simulation at Texas A&M University for providing computing resources. Appreciation is expressed to Dr David Russell and Dr Doyong Kim for interpretation of the high resolution ESI-MS. The authors would also like to acknowledge Dr Allen M. Lunsford for helpful discussions.

References

- W. Lubitz, H. Ogata, O. Rüdiger and E. Reijerse, *Chem. Rev.*, 2014, **114**, 4081.
- M. Can, F. A. Armstrong and S. W. Ragsdale, *Chem. Rev.*, 2014, **114**, 4149.
- Special issue on “Hydrogenases”, *Eur. J. Inorg. Chem.*, 2011, 915.
- Special issue on “Renewable Energy”, *Chem. Soc. Rev.*, 2009, **38**, 1–300.
- Special issue on “Hydrogen”, *Chem. Rev.*, 2007, **107**, 3900–4435.
- M. L. Helm, M. P. Stewart, R. M. Bullock, M. R. DuBois and D. L. DuBois, *Science*, 2011, **333**, 863.
- J. A. Denny and M. Y. Darensbourg, *Chem. Rev.*, 2015, **115**, 5248.
- C. H. Hsieh, S. Ding, O. F. Erdem, D. J. Crouthers, T. Liu, C. C. L. McCrory, W. Lubitz, C. V. Popescu, J. H. Reibenspies, M. B. Hall and M. Y. Darensbourg, *Nat. Commun.*, 2014, **5**, 3684.
- S. Ding, P. Ghosh, A. M. Lunsford, N. Wang, N. Bhuvanesh, M. B. Hall and M. Y. Darensbourg, *J. Am. Chem. Soc.*, 2016, **138**, 12920.
- P. Ghosh, M. Quiroz, N. Wang, N. Bhuvanesh and M. Y. Darensbourg, *Dalton Trans.*, 2017, **46**, 5617.
- W. Zhu, A. C. Marr, Q. Wang, F. Neese, D. J. E. Spencer, A. J. Blake, P. A. Cooke, C. Wilson and M. Schröder, *Proc. Natl. Acad. Sci. U. S. A.*, 2005, **102**, 18280.
- S. Ogo, R. Kabe, K. Uehara, B. Kure, T. Nishimura, S. C. Menon, R. Harada, S. Fukuzumi, Y. Higuchi, T. Ohhara, T. Tamada and R. Kuroki, *Science*, 2007, **316**, 585.
- S. Ogo, K. Ichikawa, T. Kishima, T. Matsumoto, H. Nakai, K. Kusaka and T. Ohhara, *Science*, 2013, **339**, 682.
- T. Kishima, T. Matsumoto, H. Nakai, S. Hayami, T. Ohta and S. Ogo, *Angew. Chem., Int. Ed.*, 2016, **55**, 724.



- 15 D. Brazzolotto, M. Gennari, N. Queyriaux, T. R. Simmons, J. Pécaut, S. Demeshko, F. Meyer, M. Orio, V. Artero and C. Duboc, *Nat. Chem.*, 2016, **8**, 1054.
- 16 S. Canaguier, M. Field, Y. Oudart, J. Pecaut, M. Fontecave and V. Artero, *Chem. Commun.*, 2010, **46**, 5876.
- 17 M. A. Reynolds, T. B. Rauchfuss and S. R. Wilson, *Organometallics*, 2003, **22**, 1619.
- 18 D. Schilter, J. M. Camara, M. T. Huynh, S. Hammes-Schiffer and T. B. Rauchfuss, *Chem. Rev.*, 2016, **116**, 8693.
- 19 K. Weber, Ö. F. Erdem, E. Bill, T. Weyhermüller and W. Lubitz, *Inorg. Chem.*, 2014, **53**, 6329.
- 20 D. Yang, Y. Li, L. Su, B. Wang and J. Qu, *Eur. J. Inorg. Chem.*, 2015, **2015**, 2965.
- 21 N. T. Nguyen, Y. Mori, T. Matsumoto, T. Yatabe, R. Kabe, H. Nakai, K. S. Yoon and S. Ogo, *Chem. Commun.*, 2014, **50**, 13385.
- 22 T. Matsumoto, R. Kabe, K. Nonaka, T. Ando, K.-S. Yoon, H. Nakai and S. Ogo, *Inorg. Chem.*, 2011, **50**, 8902.
- 23 K. Ichikawa, T. Matsumoto and S. Ogo, *Dalton Trans.*, 2009, 4304.
- 24 K. Kim, T. Kishima, T. Matsumoto, H. Nakai and S. Ogo, *Organometallics*, 2013, **32**, 79.
- 25 K. Kim, T. Matsumoto, A. Robertson, H. Nakai and S. Ogo, *Chem.-Asian J.*, 2012, **7**, 1394.
- 26 J. A. Denny and M. Y. Darensbourg, *Coord. Chem. Rev.*, 2016, **324**, 82.
- 27 C. H. Hsieh and M. Y. Darensbourg, *J. Am. Chem. Soc.*, 2010, **132**, 14118.
- 28 J. H. Enemark and R. D. Feltham, *Coord. Chem. Rev.*, 1974, **13**, 339.
- 29 A. Silakov, B. Wenk, E. Reijerse and W. Lubitz, *Phys. Chem. Chem. Phys.*, 2009, **11**, 6592.
- 30 G. Berggren, A. Adamska, C. Lambertz, T. R. Simmons, J. Esselborn, M. Atta, S. Gambarelli, J. M. Mouesca, E. Reijerse, W. Lubitz, T. Happe, V. Artero and M. Fontecave, *Nature*, 2013, **499**, 66.
- 31 J. Esselborn, C. Lambertz, A. Adamska-Venkatesh, T. Simmons, G. Berggren, J. Noth, J. Siebel, A. Hemschemeier, V. Artero, E. Reijerse, M. Fontecave, W. Lubitz and T. Happe, *Nat. Chem. Biol.*, 2013, **9**, 607.
- 32 J. F. Siebel, A. Adamska-Venkatesh, K. Weber, S. Rumpel, E. Reijerse and W. Lubitz, *Biochemistry*, 2015, **54**, 1474.
- 33 J. Esselborn, N. Muraki, K. Klein, V. Engelbrecht, N. Metzler-Nolte, U. P. Apfel, E. Hofmann, G. Kurisu and T. Happe, *Chem. Sci.*, 2016, **7**, 959.
- 34 W. F. Liaw, C. Y. Chiang, G. H. Lee, S. M. Peng, C. H. Lai and M. Y. Darensbourg, *Inorg. Chem.*, 2000, **39**, 480.
- 35 C. H. Hsieh, R. B. Chupik, S. M. Brothers, M. B. Hall and M. Y. Darensbourg, *Dalton Trans.*, 2011, **40**, 6047.
- 36 F. Osterloh, W. Saak, D. Haase and S. Pohl, *Chem. Commun.*, 1997, 979.
- 37 J. A. W. Verhagen, M. Lutz, A. L. Spek and E. Bouwman, *Eur. J. Chem.*, 2003, 3968.
- 38 C. H. Hsieh, S. M. Brothers, J. H. Reibenspies, M. B. Hall, C. V. Popescu and M. Y. Darensbourg, *Inorg. Chem.*, 2013, **52**, 2119.
- 39 M. V. Rampersad, S. P. Jeffery, M. L. Golden, J. Lee, J. H. Reibenspies, D. J. Darensbourg and M. Y. Darensbourg, *J. Am. Chem. Soc.*, 2005, **127**, 17323.
- 40 G. B. Richter-Addo and P. Legzdins, *Metal Nitrosyls*, 1992, ch. 5.
- 41 T. A. Pinder, S. K. Montalvo, A. M. Lunsford, C. H. Hsieh, J. H. Reibenspies and M. Y. Darensbourg, *Dalton Trans.*, 2014, **43**, 138.
- 42 B. P. Hay and R. D. Hancock, *Coord. Chem. Rev.*, 2001, **212**, 61.
- 43 M. B. Hall, *Inorg. Chem.*, 1978, **17**, 2261.
- 44 D. J. Crouthers, S. Ding, J. A. Denny, R. D. Bethel, C. H. Hsieh, M. B. Hall and M. Y. Darensbourg, *Angew. Chem.*, 2015, **127**, 11254.
- 45 A. M. Appel and M. L. Helm, *ACS Catal.*, 2014, **4**, 630.
- 46 S. Wiese, U. J. Kilgore, M. H. Ho, S. Raugei, D. L. DuBois, R. M. Bullock and M. L. Helm, *ACS Catal.*, 2013, **3**, 2527.
- 47 V. Fourmond, P. A. Jacques, M. Fontecave and V. Artero, *Inorg. Chem.*, 2010, **49**, 10338.
- 48 S. Daniele, P. Ugo, G. A. Mazzocchin and G. Bontempelli, *Anal. Chim. Acta*, 1985, **173**, 141.
- 49 G. A. N. Felton, R. S. Glass, D. L. Lichtenberger and D. H. Evans, *Inorg. Chem.*, 2006, **45**, 9181.
- 50 S. Ding, P. Ghosh, M. Y. Darensbourg and M. B. Hall, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, DOI: 10.1073/pnas.1710475114, Accepted manuscript.
- 51 D. L. DuBois, *Inorg. Chem.*, 2014, **53**, 3935.

