Concentration dependent effects of urea binding to poly(N-isopropylacrylamide) brushes: a combined experimental and numerical study†

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The binding effects of osmolytes on the conformational behavior of grafted polymers are studied in this work. In particular, we focus on the interactions between urea and poly(N-isopropylacrylamide) (PNIPAM) brushes by monitoring the ellipsometric brush thickness for varying urea concentrations over a broad temperature range. The interpretation of the obtained data is supported by atomistic molecular dynamics simulations, which provide detailed insights into the experimentally observed concentration-dependent effects on PNIPAM–urea interaction. In particular, in the low concentration regime (c_u ≤ 0.5 mol L\(^{-1}\)) a preferential exclusion of urea from PNIPAM chains is observed, while in the high concentration regime (2 ≤ c_u ≤ 7 mol L\(^{-1}\)) a preferential binding of the osmolyte to the polymer surface is found. In both regimes, the volume phase transition temperature (T_v) decreases with increasing urea concentration. This phenomenon derives from two different effects depending on urea concentration: (i) for c_u ≤ 0.5 mol L\(^{-1}\), the decrease of T_v is explained by a decrease of the chemical potential of bulk water in the surrounding aqueous phase; (ii) for c_u ≥ 2 mol L\(^{-1}\), the lower T_v is explained by the favorable replacement of water molecules by urea, which can be regarded as a cross-linker between adjacent PNIPAM chains. Significant effects of the concentration-dependent urea binding on the brush conformation are noticed: at c_u = 0.5 mol L\(^{-1}\), although urea is loosely embedded between the hydrated polymer chains, it enhances the brush swelling by excluded volume effects. Beyond 0.5 mol L\(^{-1}\), the stronger interaction between PNIPAM and urea reduces the chain hydration, which in combination with cross-linking of monomer units induces the shrinkage of the polymer brush.

1 Introduction

Osmolytes are low-weight organic molecules which contribute to the regulation of the osmotic pressure inside the cells. Their presence allows biological cells to counteract a change of the external osmotic pressure in such a way that the isotonic balance is ensured. In addition, osmolytes have a strong influence on the structural conformation of proteins.\(^1\)\(^–\)\(^4\) It was reported that in high concentrated solutions of denaturants like guanidinium chloride or urea, protein unfolding occurs when a critical osmolyte concentration is reached.\(^5\)\(^–\)\(^9\) In contrast, the stabilization of protein structures was validated for protectants like trimethylamine-N-oxide (TMAO) and hydroxyectoine.\(^10\)\(^–\)\(^13\)

Experimental findings further evidenced that low concentrations of urea and guanidinium chloride can even stabilize protein native structures\(^6\) and lead to the fluidization of lipid bilayers and monolayers.\(^12\)\(^–\)\(^14\) The mechanisms and the different accumulation properties\(^15\)\(^–\)\(^18\) for protectants and denaturants around molecular surfaces can be brought into relation with the ‘law of matching water affinities’.\(^15\)\(^,\)\(^19\)\(^,\)\(^20\) As it was also recently demonstrated by computer simulations, the resulting binding behavior is a consequence of a subtle interplay between enthalpic and entropic effects, and the underlying charge of the solvated surface.\(^21\) In addition, further simulations of thermo-sensitive polymers in presence of different salt concentration also reveal a concentration dependent change of binding properties for chaotrope salts.\(^22\) The possibility of investigating these effects is reasoned by the fact that the binding or the exclusion mechanism of the co-solute has a crucial influence on the stability of the macromolecule and therefore on the corresponding transition temperatures. Two distinct osmolyte binding mechanisms have been proposed: (1) direct interactions, related to the binding of the osmolyte to the surface of the macromolecule via hydrogen
bonds, salt bridges or dispersion interactions, and (2) indirect interactions, mostly imposed by the influence of the osmolyte on the local water shell around the surface. However, it is still under debate which mechanism is predominant.

Due to the chemical variety and structural complexity, simplified homopolymers are more preferred than proteins in fundamental studies. Among them, poly(N-isopropylacrylamide) (PNIPAM) received a lot of scientific attention, because the mechanism underlying its thermoresponsive behavior, namely a coil-to-globule transition above the lower critical solution temperature (LCST), is closely related to the loss of protein structure upon addition of denaturing agents and resembles the behavior found for the cold denaturation of proteins. It has been validated that the predominant effects displayed above the LCST rely on the low number of hydrating water molecules bound to the macromolecule. Hence, the analysis above the LCST is presented. The experimental approach consisted in monitoring the optical thickness of the swelling and collapse properties of PNIPAM brushes over a broad range of urea concentrations from 0.1 to 7 mol L$^{-1}$ is presented. The experimental approach consisted in monitoring the optical thickness of PNIPAM brushes by ellipsometry as a function of temperature for urea concentrations between 0.1 and 7 mol L$^{-1}$. From the obtained data, the transition temperature ($T_d$) and the swelling degree ($\phi_{sw}$) of PNIPAM below (288 K) and above (328 K) the critical solution temperature were calculated. The hydration properties with regard to the underlying polymer–urea interactions were deduced from the observed swelling and collapse behavior. The interpretation of the experimental data is validated by atomistic MD simulation of PNIPAM chains in two different concentration regimes (low and high molar concentration) and for temperatures below and above the phase transition. The numerical studies further allow us to elucidate the balance between solvent (water) and co-solute (urea) binding on the conformational equilibrium of the polymer chain for the different combinations of urea concentration and temperature. In particular, at low molar urea concentrations, the observed shift of the transition temperature was governed by the dilution effect of urea on bulk water, which changes its chemical potential from pure solvent to dilute solution, and therefore favors brush dehydration across the phase transition. This effect was due to a preferential exclusion of urea from the macromolecule. In contrast, at high molar urea concentrations, the preferential binding of urea to PNIPAM was responsible for the pronounced decrease of $T_d$, which was reasoned by the weakening of PNIPAM–water interactions, and the cross-linking of the osmolyte between adjacent PNIPAM chains.

Our study offers a deeper understanding on the impact of direct urea binding and indirect effects on the conformational equilibrium of the polymer. It contributes to widen the previous knowledge in this field and to understand the behavior of proteins in presence of osmolytes.

## 2 Experimental section

### 2.1 Materials

Silicon wafers (Siltronic AG, Germany) were used as solid substrates for the polymer brushes. N-isopropylacrylamide (NIPAM) (98%, stabilized by methylhydroquinone) was purchased from TCI Tokyo Chemical Industry; $N,N,N',N',N''$,pentamethylene-triamine (PMDETA), Cu(II)Cl, methanol and urea were all purchased from Sigma Aldrich (Germany). All reagents were used as received without any further purification.

### 2.2 Synthesis of polymer brushes

Polymer brushes were grown from silicon wafers by Atom Transfer Radical Polymerization (ATRP). First of all, a self-assembled monolayer of the ATRP initiator 2-bromo-2-methyl-$N$-[3-(triethoxysilyl)-propyl]-propanamide (BTPAm)$^{49}$ in dry toluene (4 mL/10 mL) was adsorbed on the substrates for 24 h. A polymerization mixture of 2.83 g NIPAM in 50 mL of methanol and water (1:1 v/v), 183 $\mu$L $N,N,N',N'',N'''$pentamethylene-triamine (PMDETA) and 25 mg copper(i) chloride was prepared and degassed by N$_2$ bubbling for 40 minutes. The initiator-modified substrates were soaked in the reaction mixture and the polymerization was run for 8 minutes, followed by termination in water/methanol mixture and thorough rinsing in Milli-Q water. The grafting density $\sigma$ was calculated considering the molecular structure of the initiator and the measured thickness of the self-assembled monolayer (details are reported in the ESI†). A value of $\sigma \approx 0.4$ nm$^{-2}$ was estimated, which corresponds to a moderate brush regime$^{50}$ and allows urea to access the full solvent accessible surface area of PNIPAM. The brush thickness of the dry state (1.9 ± 0.1% r.h.) was measured by ellipsometry after drying the samples in a sealed cell flushed with nitrogen stream for 20 minutes. The measured value was $d_{dry} = (40.5 \pm 0.9)$ nm and it remained constant for at least 10 minutes.

The molecular weight ($M_W$) of PNIPAM chains was estimated from the grafting density $\sigma$ and the dry thickness $d_{dry}$, according to the relation$^{49,61,52}$

$$M_W = \frac{\rho d_{dry} N_A}{\sigma}$$

(1)
with \( \rho \) the polymer mass density (1.1 g mL\(^{-1}\)) and \( N_A \) the Avogadro number. The obtained value for \( M_w \) was \( \approx 67,000 \text{ g mol}^{-1} \) (ca. 590 monomers per chain). Reported polydispersity indices (PDI) for ATRP polymerization are mostly between 1.2 and 1.5.\(^{53,54}\) A low PDI is also expected for our system, given the low standard deviation of the brush thickness (ca. 0.8% of the brush thickness). The swollen thicknesses of PNIPAM brushes in solutions of urea at different concentrations varied between 100 and 150 nm at 288 K, and between 47 and 80 nm at 328 K. The values are reported in Table S1 of the ESL.\(^{†}\)

2.3 Ellipsometry

Ellipsometry is a non-destructive, optical technique which is based on the detection of the change of polarization of light upon reflection from a substrate. Such a change is described by two parameters, \( \psi \) and \( \Delta \), which are related to the amplitude (\( E \)) and the phase (\( \delta \)) of light by the following relations:

\[
\tan \psi = \frac{|E_p^s|/|E_p^i|}{|E_s^i|/|E_s^r|}
\]

(2)

\[
\Delta = (\delta_p^r - \delta_p^s) - (\delta_s^r - \delta_s^i)
\]

(3)

where the subscripts \( p, s \) indicate the parallel and perpendicular components of the light, while the superscripts \( i, r \) specify the incoming and reflected beam. The reflectivity properties of a sample can be expressed by the reflection coefficients, \( r_p \) and \( r_s \),

\[
\begin{align*}
\rho_p &= \frac{|E_p^s|}{|E_p^i|} e^{i(\delta_p^r - \delta_p^s)} \\
r_s &= \frac{|E_s^i|}{|E_s^r|} e^{i(\delta_s^r - \delta_s^i)}
\end{align*}
\]

(4)

(5)

whose combination with eqn (2) and (3) yield the fundamental equation of ellipsometry:

\[
\tan \psi \cdot e^{i\Delta} = \frac{\rho_p}{r_s}
\]

(6)

The use of a proper layer model, which describes the system under measurement (substrate-film-medium) in terms of its optical properties (thickness \( d \) and refractive index \( n \)) can be obtained from the measured parameters \( \Delta \) and \( \psi \), which in turn describe the interaction between polarized light and sample. Specific literature is available on the topic to which the reader is addressed for a more detailed description of the technique.\(^{55,56}\)

For the data analysis, the measured values of \( \Delta \) and \( \psi \) were fitted with a five box model consisting, from the bottom to the top, of (i) silicon substrate \( (n = 3.8850, d = \infty) \), (ii) silicon oxide \( (n = 1.46, d = 1.5 \text{ nm}) \), (iii) initiator BTPAm monolayer \( (n = 1.50, d = 0.7 \text{ nm}) \), (iv) polymer brushes \( (n = n_{\text{brush}}, d = d_{\text{brush}}) \), (v) liquid medium \( (n_{\text{calc}}, d = \infty) \). From the measured values of \( \Delta \) and \( \psi \), the corresponding combination of \( n = n_{\text{brush}}, d = d_{\text{brush}} \) for the polymer brushes is obtained. The refractive index of solutions containing different urea concentrations was estimated by the empirical relations elaborated by Warren and Gordon\(^{57}\) from experimental data on urea solutions. The corresponding values are reported in Table S2 of the ESL.\(^{†}\) The degree of swelling \( \phi_{sw} \) for PNIPAM brushes was calculated from the thickness \( d_{dry} \) of the dry brush and the thickness \( d_{sw}(T) \) of the swollen brush at a given temperature \( T \) by

\[
\phi_{sw} = \frac{d_{sw}(T) - d_{dry}}{d_{dry}} \times 100\%.
\]

(7)

The brush collapse from the swollen conformation triggered by the increase of temperature was quantified from the difference between the measured thickness at \( T > 288 \text{ K} \) and the initial value measured at 288 K:

\[
\frac{d_{sw}(T) - d_{sw}(288 \text{ K})}{d_{sw}(288 \text{ K})} \times 100 = \frac{\Delta d_{sw}}{d_{sw}(288 \text{ K})} \%
\]

(8)

where \( d_{sw}(288 \text{ K}) \) is the swollen brush thickness measured at 288 K and \( d_{sw}(T) \) the PNIPAM thickness at temperatures \( T > 288 \text{ K} \). The transition temperature \( T_{tr} \) was taken as the temperature at half of the brush collapse.

Ellipsometry measurements were carried out using a Multiscope Null-Ellipsometer from Optrel GbR (Germany). The instrument is equipped with a red laser \( (\lambda = 632.8 \text{ nm}) \) and a PCSA (polarizer-compensator-sample-analyzer) setup. For the measurements in dry condition, a home-built humidity cell was used, whose inside is flushed by nitrogen stream, reaching a value of \( (1.9 \pm 0.1)\% \) relative humidity. The angle of incidence was set to \( 70^\circ \) and the sample was left drying for 20 minutes before measurements. For the measurements in liquid, the sample was soaked in a stainless steel cell filled with the urea solution. The instrument was equipped with light guides to drive the incident beam directly at the substrate/water interface and avoid the reflection at the liquid/air interface. Prior to the measurements, the sample was swollen in urea solution at 288 K for at least 2 h. A thermal cycle was applied by heating the liquid by means of a copper plate underneath the sample holder. The temperature of the liquid environment was measured continuously with a precision of \( \pm 0.01 \text{ K} \) by means of a temperature sensor immersed in the sample cell.

2.4 Simulation details

A PNIPAM molecule with 24 monomers was studied by atomistic molecular dynamics (MD) simulations with the GROMACS 4.6.2 software package.\(^{58-60}\) The force field parameters for PNIPAM presented in ref. 35 have been used to guarantee the occurrence of the coil-to-globule transition at temperatures below 328 K\(^{34,35}\) while generalized AMBER force fields were used for urea.\(^{61,62}\) Typical swollen and collapsed PNIPAM configurations have been obtained by using the original Metadynamics approach\(^{53-65}\) as implemented within the PLUMED package.\(^{66}\) We extracted the most swollen (radius of gyration \( R_g = 1.4 \text{ nm} \)) and the collapsed state of PNIPAM (radius of gyration \( R_g = 0.8 \text{ nm} \)) as found during the Metadynamics simulation and inserted them into cubic simulation boxes to study the influence of pure water, low and high aqueous urea.
concentrations, respectively. The position of the backbone $C_2$ carbon atoms have been kept fixed by using position restraints. We randomly inserted the corresponding number of urea and TIP3P water molecules$^{67}$ to model the required urea concentrations. In addition, we also simulated a PNIPAM molecule in pure water as a reference system. Electrostatic interactions for all systems have been calculated by the Particle Mesh Ewald method$^{68}$ and all bonds have been constrained by the LINCS algorithm.$^{69}$ First, we performed an energy minimization followed by a 2 ns warm-up run in a NPT ensemble at 288 K and 328 K with the Berendsen barostat and thermostat,$^{70}$ followed by NPT simulations of 20 ns for the production run with the same parameter sets. The minimum and maximum values for the length of the cubic simulation box are given by $(l)_{\text{min}} = 4.7 \text{ nm}$ and $(l)_{\text{max}} = 6.49 \text{ nm}$ in the NPT simulation depending on the system and the temperature. The corresponding effective urea concentrations for the different systems can be found in Table 1.

Hydrogen bonds were defined by a maximum distance criterion of 0.35 nm between acceptor and donor pairs with a maximum angle of 35°.

In fact, the study of polymer brushes is a challenging task for atomistic simulations. For the minimization of finite-size effects, a large number of grafted PNIPAM chains have to be considered which limits the applicability of atomistic approaches. In contrast, insights into polymer brush behavior can be provided by computationally efficient coarse-grained (CG) simulations$^{71}$ in which specific chemical details are usually neglected. Due to these reasons, CG approaches are not suitable for the detailed study of which specific chemical details are usually neglected. Due to these reasons, CG approaches are not suitable for the detailed study of.

<table>
<thead>
<tr>
<th>$T$ [$K$]</th>
<th>$R_g$ [nm]</th>
<th>$c_u$ [mol L$^{-1}$]</th>
<th>Conformation</th>
<th>Label</th>
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<td>0.5</td>
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<td>288</td>
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<td>0.5</td>
<td>Swollen</td>
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<td>328</td>
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<td>328</td>
<td>1.4</td>
<td>6.1</td>
<td>Swollen</td>
<td>hs328</td>
</tr>
</tbody>
</table>

2.5 Implications of the Kirkwood–Buff theory and local/bulk partition coefficients

The Kirkwood–Buff (KB) integral is given by

$$G_{ab} = 4\pi \int_0^r r^2 \left( \frac{g_{ab}^{KT}(r)}{g_{ab}^0} - 1 \right) \, dr$$

with the cut-off radius $r_c$. The value of $r_c = 1.8 \text{ nm}$ was used for the data analysis, since the results of the radial distribution functions showed converged values from this point onwards. The physical meaning of Kirkwood–Buff integrals can be interpreted as the excess volume of a co-solute $\beta$ around the central particle of a solute $\alpha$, and they can be also used to study the local solvent accumulation behavior. In terms of notation, in the following description the solvent (water) will be denoted by the subscript $\alpha$, and $\beta$ will denote the systems.

The preferential binding coefficient $\nu_{23}$ between PNIPAM and urea can be obtained from the difference of the KB integrals describing the distribution of water and urea around a PNIPAM chain, according to

$$\nu_{23} = \rho_3 (\Delta G_{23} - \Delta G_{21})$$

where $\rho_3$ denotes the urea bulk number density. Note that the argument $r_c$ for the KB integrals has been omitted. The preferential binding coefficient is connected with the transfer free energy

$$\Delta F_{23} = -RT\nu_{23}$$

with $R$ the molar gas constant and $T$ the temperature. The value for the transfer free energy estimates the amount of free energy which is needed to transfer a co-solute from the bulk to the surface of the solute.

The effects of osmolytes on macromolecular conformations can be evaluated in terms of the chemical equilibrium constant $K = n_\alpha/n_\beta$, where $n_\alpha$ and $n_\beta$ denote the fraction of PNIPAM molecules in the swollen (s) and the collapsed state (c). From the relation

$$\left( \frac{\partial \ln K}{\partial \ln a_1} \right) = \Delta \nu_{23} = \nu_{23}^{C} - \nu_{23}^{S}$$

the chemical activity of urea $a_3$ and the chemical equilibrium constant can be associated with the preferential binding.
coefficients of urea to the swollen and the collapsed PNIPAM conformation ($\nu_{23}^{sw}$ for the swollen and $\nu_{23}^{c}$ for the collapsed conformation). The combination of eqn (13) with eqn (12) allows us to connect the chemical equilibrium in presence of osmolytes with the difference in the transfer free energy $\Delta F_{23}$, 

$$\Delta F_{23} = \Delta F_{23}^{sw} - \Delta F_{23}^{c} \tag{14}$$

to yield the final relation 

$$\Delta F_{23} = -RT\left(\frac{\partial \ln K}{\partial \ln c}\right) \tag{15}$$

In accordance with the definition of the reaction constant $K$, it can be concluded that a change of the free energy can be associated with a change of the temperature if it is assumed that the free energy is mostly dominated by enthalpic contributions. Moreover, with regard to eqn (13), it can be also concluded, that the chemical equilibrium is shifted to the conformation with a stronger osmolyte accumulation.

Another robust parameter to study the osmolyte accumulation behavior is given by the local/bulk partition coefficient according to 

$$K_p(r) = \frac{\langle n_p(r) \rangle / \langle n_p \rangle}{n_p^{tot} / n_p^{tot}} \tag{16}$$

where $\langle n_p(r) \rangle$ denotes the average number of water ($X = w$) or urea molecules ($X = u$) within a distance $r$ and $n_p^{tot}$ the total number of urea or water molecules in the simulation box. Thus, a preferential exclusion behavior leads to $K_p(r) < 1$ and a preferential binding of the osmolyte to $K_p(r) > 1$ at short distances $r$. Considering the local/bulk partition coefficient around the swollen PNIPAM configuration $K_p^{sw}(r) = \exp(-\Delta F_p^{sw}(r)/RT)$ and around the collapsed configuration $K_p^{c}(r) = \exp(-\Delta F_p^{c}(r)/RT)$, respectively, the local/bulk partition coefficient free energy difference can be expressed by 

$$\Delta \Delta F_{K_p}(r) = -RT\log\left(\frac{K_p^{sw}(r)}{K_p^{c}(r)}\right) \tag{17}$$

Eqn (17) implies that a negative value of $\Delta \Delta F_{K_p}(r)$ indicates a stronger accumulation of urea around the swollen state, and therefore the shift of the chemical equilibrium towards chain swelling, whereas a positive value reveals the stabilization of the collapsed conformation due to a stronger urea accumulation around the collapsed chain. Herewith, we exactly follow the implications given by eqn (13). Noteworthy, it can be also seen that even for negative preferential binding coefficients (eqn (11)) or local/bulk partition coefficients $K_p(r) > 1$, a shift of the chemical equilibrium for $\Delta \Delta F_{K_p}(r) \neq 0$ at short distances $r$ can be observed, which is purely induced by the accumulation behavior of urea.

### 3 Results

#### 3.1 Experimental results

The relative change of ellipsometric thickness $\Delta d_{sw}/d_{sw}(288 K)$ calculated according to eqn (8) is reported in Fig. 1 as a function of temperature for different urea concentrations ($c_u$). Two important features in the obtained trends can be noticed: a shift of $T_{tr}$ to lower values for increasing urea concentration $c_u$ as denoted by the left arrow, and a weaker brush collapse in urea solutions compared to pure water as denoted by the right arrow. The weaker brush collapse above the transition temperature at high urea concentrations can be discussed with the arguments given by Cremer and co-workers in ref. 38, where increasing PNIPAM radii observed at increasing urea concentration were explained by bridging effects of the osmolyte in order to cross-link PNIPAM chains. An analogous effect in this case might be responsible for the reduced brush collapse at high $c_u$. In order to analyze the effects of urea on the thermal behavior of PNIPAM in more detail, $T_{tr}$ and $\phi_{sw}$ were extracted from the ellipsometry curves. The values of $T_{tr}$ for PNIPAM brushes in presence of different urea concentrations are presented in Fig. 2. A systematic decrease of $T_{tr}$ for increasing urea concentrations was found, with the evidence of two distinct regimes: a low concentration regime between 0 and 0.5 mol L$^{-1}$, with a linear decrease of $T_{tr}$ as a function of $c_u$ (inset in Fig. 2), and a high urea concentration regime between 2 and 7 mol L$^{-1}$ with a weaker, but constant decrease of $T_{tr}$. Two distinct regimes, although not so noticeable, were found also for single PNIPAM chains. Thus, it can be concluded that this effect is not exclusively characteristic for PNIPAM brushes and it demonstrates that molecular PNIPAM–urea interactions dominate the chain behavior, independent of the system geometry. This allows us to apply the knowledge of single chain–urea interactions to more complex and crowded molecular systems, like polymer brushes, which might be of paramount importance for the application of grafted biomolecular (e.g. proteins) systems. Remarkably, the influence of urea on the transition temperature of PNIPAM may have different origins: on the one hand, the presence of urea changes the properties of the aqueous medium; on the other hand, the osmolyte could
directly interact with the polymer, as recently demonstrated for high urea concentrations.\textsuperscript{38,82}

While at high $c_u$, the direct urea binding to the polymer is very likely, it is important to verify which mechanism dominates the change of $T_tr$ for a PNIPAM brush in presence of a low molar urea concentration. As a first hypothesis, it is assumed that urea does not directly interact with PNIPAM, \textit{i.e.} it is preferentially excluded from a PNIPAM surface. Under this condition, the shift of $T_tr$ might be reasoned by a change of the PNIPAM hydration properties.

From the decrease of the chemical potential of bulk water upon addition of urea, and from the equilibrium condition at $T_tr$ with identical chemical potentials for H$_2$O (bulk) and H$_2$O (brush) phase, the following relation between the temperature shift $\Delta T_tr$ and the dehydration enthalpy $\Delta H_{dehy}$ can be derived:

$$\Delta T_tr = -\frac{R(T_tr^0)^2}{\Delta H_{dehy}} \gamma_u$$

where $T_tr^0$ is the transition temperature in pure water ($c_u = 0$ mol L$^{-1}$), $\Delta H_{dehy}$ corresponds to the dehydration enthalpy per mole of water and $\gamma_u$ denotes the urea–water mole fraction. The full derivation of this relation is given in the ESI.\textsuperscript{†}

It has to be noted that a similar expression was also derived in ref. 22.

It is worth to mention that this approach represents a simplified picture of the system, where the transition enthalpy of the macromolecule from swollen to collapsed state is omitted. The validity of eqn (18) to describe the experimental data of Fig. 2 at low $c_u$ would confirm the dilution effect of bulk water by urea, \textit{i.e.} indirect effects. Previous studies on the coil-to-globule transition of PNIPAM reported calorimetric enthalpy values $\Delta H_{cal}$ between 5.5 and 7 kJ mol$^{-1}$ for a single (NIPAM) repeating unit.\textsuperscript{83,84}

This can be related to the release of one bound water molecule from each monomer unit, in agreement with previous experimental results.\textsuperscript{83,85} Other literature works report the release of a much higher number of water molecules per NIPAM unit in bulk or in gels.\textsuperscript{86,87} Nevertheless, in our evaluation only the water molecules directly bound to the monomer unit \textit{via} H-bonding are considered.

Assuming that the same process occurs for the dehydration of PNIPAM brushes in presence of low molar urea concentrations, where direct urea–PNIPAM interactions are negligible, it follows that $\Delta H_{cal} \sim \Delta H_{dehy}$ and therefore it is possible to calculate the theoretical trend of $T_tr$ as a function of $c_u$ for the reported dehydration enthalpies of $\Delta H_{cal}(1) = 5.5$ kJ mol$^{-1}$ and $\Delta H_{cal}(2) = 7$ kJ mol$^{-1}$. The corresponding functions are shown as solid lines in the inset of Fig. 2. The agreement between experimental data and theoretical trends corroborates the assumption that the dehydration mechanism of PNIPAM in low concentrated urea solutions is comparable to the dehydration in pure water in accordance to ref. 83 and 84. Thus, the shift of the transition temperature can be explained by the dilution of water by urea as induced by the mole fraction $\gamma_u$.

A different mechanism is likely to occur in the high concentration regime, where urea is known to have a significant influence on the stability of proteins and a direct osmolyte–polymer binding has been reported in previous studies.\textsuperscript{38,82}

Indeed, the profile of $T_tr$ shown in Fig. 2 is very similar to the results presented by Cremer and coworkers\textsuperscript{38} obtained by Fourier transformed infrared (FTIR) spectroscopy measurements. The reported FTIR data showed a strong amide band for PNIPAM in 6 M urea solution, with a major contribution from C\equiv O(PNIPAM)–H$_2$O hydrogen bonds and a second minor one from C\equiv O(PNIPAM)–NH(urea) bonds. An analogous effect is likely to dominate the collapse of PNIPAM brushes at $c_u \geq 2$ M, as indicated by our results, where the direct binding of urea with (NIPAM) monomers might introduce the aggregation of adjacent chains, such that urea acts as a cross-linking agent.\textsuperscript{38}

Evidence of direct polymer–osmolyte binding was also found by Wang et al.\textsuperscript{82} for poly(N,N-dimethylacrylamide) (PDMA) and poly(N,N-diethylacrylamide) (PDEA) brushes by atomic force microscopy-based single molecule force spectroscopy (SMFS). The enhanced stiffness of the polymer chains in presence of 2 M and 8 M urea solution was explained by the formation of hydrogen bonds between the polymer side groups and urea molecules. Cross-linking of urea \textit{via} hydrogen bonding was also reported by Lu et al.\textsuperscript{88} to explain urea-induced aggregation of PNIPAM chains observed by static and dynamic light scattering. According to Cremer and co-workers,\textsuperscript{32,38} a schematic representation of the bridging mechanism between PNIPAM chains in presence of urea is shown in Fig. 3. This aspect will be discussed in more detail in the next section.

To study the effect of urea on the conformation of grafted chains, the swelling degree $\phi_{sw}$ of PNIPAM brushes below (288 K) and above (328 K) the phase transition temperature was calculated according to eqn (7) for the different urea concentrations. As it is shown in Fig. 4, a stronger swelling percentage compared to pure water ($c_u = 0$ mol L$^{-1}$) is achieved...
for low urea concentrations both at 288 K and 328 K. In contrast, in the high concentration regime ($c_u \geq 2 \text{ mol L}^{-1}$) $\phi_{sw}$ is lower (288 K) or similar (328 K) to the swelling degree in pure water. Analogous to the previous assumption, we rule out a strong accumulation of urea inside the brush. Therefore indirect effects might be responsible for the observed enhancement of brush swelling at $c_u = 0.5 \text{ mol L}^{-1}$. More specifically, steric effects on the chain conformation due to the presence of urea in the swelling medium might lead to higher chain stretching. This aspect is analyzed and clarified by the numerical studies presented in the next section. It is worth to mention that this effects might not be visible at concentrations below 0.5 M due to the very low amount of the osmolyte in the swelling medium. In the case of high $c_u$, there are two important implications: at low temperature, increasing urea concentrations decrease the brush swelling, likely due to a decrease of water bonding (lower hydration) to the polymer chain; at high temperature, the presence of urea molecules around the macro-molecule might be responsible for the stabilization of the collapsed state. It has to be noticed that the considered temperatures of 288 K and 328 K are significantly below and above the transition temperature, such that temperature dependent dehydration effects, which could in principle influence $T_{tr}$, can be omitted.

An analogous representation of the structural transition of PNIPAM brushes across $T_{tr}$ is given by the percentage of the total brush collapse (i.e. reached at 328 K), which is reported in the ESI.† The obtained trend as a function of $c_u$ demonstrates a reduced brush collapse for any urea concentration compared to pure water. While in the high concentration regime this can be reasonably explained by the reduced water content in the brush, which would confirm the direct polymer–urea interaction reported by other authors,38 in the low concentration regime it is likely the result of a more subtle interplay between hydration properties and bulk effects of urea.

In summary, the experimental results indicate concentration-dependent properties of urea on transition temperatures and resulting conformations of PNIPAM brushes above and below $T_{tr}$. The understanding of the molecular origins of the observed effects on phase transition and swelling behavior motivated the numerical studies presented in the next subsection. Herewith, we mostly focus on the interactions between PNIPAM, water and urea in the low/high concentration regimes and below/above the phase transition temperature.

### 3.2 Numerical results

The interpretation of the experimental results presented in the previous section suffer of some speculative hypotheses to deduce a suitable molecular mechanism to explain the observed effects. This is an intrinsic limitation of most experimental methods, but as we demonstrated in a previous publication,89 it can be overcome when experiments and simulations are interpreted in a synergistic way. Following this approach, simulation results are presented to complement the experimental findings and to validate their interpretation.

Typical snapshots of a collapsed and a swollen PNIPAM configuration in a low molar urea solution are presented in Fig. 5. By introducing specific position restraints to fix the PNIPAM configurations, it was possible to avoid transitions into metastable states and to obtain reliable and well-sampled statistically averaged values. A detailed discussion of this approach is given in the section Simulation details.

In order to study the hydration behavior, the local Kirkwood–Buff integrals $G_{21}(r)$ for PNIPAM and water molecules at 288 K for $c_u = 0.5 \text{ mol L}^{-1}$ were calculated. The values for $G_{21}(r)$ in pure water were also estimated for comparison. The results are presented at the top of Fig. 6. Only small differences between the $G_{21}(r)$ in urea solution and the value in pure water can be observed. Hence, it becomes clear that low molar urea concentrations lead to a weak change in the PNIPAM hydration behavior at low temperatures, as it was also concluded from the experimental results.

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*(Schematic representation of the urea bridging effect between adjacent PNIPAM chains occurring in the high concentration regime ($c_u \geq 2 \text{ M}$). Urea as a cross-linking agent significantly enhances the stability of the collapsed state. The image was inspired by ref. 32 and 38.)*
Changes in the PNIPAM–water interactions in presence of urea can be easily evaluated considering the differences of the Kirkwood–Buff integrals according to

$$\Delta G_{21} = G_{21}^u - G_{21}^w$$  \hspace{1cm} (19)$$

with $G_{21}^w$ for pure water and $G_{21}^u$ in presence of urea. The corresponding results at $r_c = 1.8$ nm for all investigated combinations of chain conformation, urea concentration and temperature are shown at the bottom of Fig. 6. Larger values of $\Delta G_{21}$ were observed for all systems at 328 K and in particular for the swollen configurations (ls328 and hs328). In contrast, a slightly modified water accumulation behavior is evident for all systems at 288 K. It can be therefore concluded that at low temperature the presence of urea only weakly perturbs the water structure around PNIPAM, in good agreement with previous assumptions in the literature\(^8^,2^0\) where it was discussed that urea can be well integrated into the water structure network.

To study the direct and local interactions between PNIPAM and urea or water molecules, the number of hydrogen bonds was calculated. The corresponding results and further findings are shown in Fig. 7 and in Table S3 of the ESI.\(^\dagger\) At low molar urea concentrations, the number of water–PNIPAM hydrogen bonds is comparable to pure water ($\Delta$H-bonds (water)) for both temperatures (ls288, lc288, ls328 and lc328). This means that the first solvent shell around PNIPAM at low $c_u$ has a similar composition as in pure water, and only a small number of urea–PNIPAM hydrogen bonds (H-bonds (urea)) are formed. These results validate the previous assumptions of a minor PNIPAM–urea interaction in presence of low $c_u$, and support the hypothesis of indirect effects from the water–urea network on the chain conformation at low $c_u$.

In contrast, a significant decrease of water hydrogen bonds compared to pure water ($\Delta$H-bonds (water)) is found at high $c_u$ (hs288, hc288, hs328 and hc328), and also the total number of hydrogen bonds is smaller compared with pure water (i.e. large negative values of $\Delta$H-bonds (total)). Interestingly, at high temperature (hs328 and hc328) the number of urea hydrogen bonds even exceeds the number of water hydrogen bonds ($\text{H-bonds (urea)} > \text{H-bonds (water)}$) meaning that in this condition urea replaces water at molecular interfaces.\(^8^8\) Moreover, from the data reported in Table S3 of the ESI it can be seen that urea hydrogen bonds are energetically stronger than water hydrogen bonds. From these results it can be concluded that at high $c_u$ urea binding dominates the behavior of PNIPAM chains across the phase transition, while the effects arising from the change of PNIPAM–water interactions are negligible.

In order to classify the hydrogen bonds with respect to their related energetic contributions according to the transition state theory with the Luzar–Chandler approach,\(^9^0,9^1\) it is possible to calculate the forward rate constants $k$ of hydrogen bonds via

$$k \sim \exp(-\Delta F^*/k_B T),$$

where $\Delta F^*$ denotes the activation free binding energy. The corresponding values for water–PNIPAM and urea–PNIPAM hydrogen bonds are shown in Table S3 of the ESI.\(^\dagger\) The most relevant result is the decrease of the activation free energies of water for increasing urea concentrations. The replacement of water–PNIPAM hydrogen bonds with urea–PNIPAM hydrogen bonds is therefore energetically favorable. Furthermore, for increasing urea concentration, the strength of
water hydrogen bonds is slightly lowered, which facilitates the release of water molecules from the first solvent shell, in agreement with the experimental results (Fig. 2). Moreover, the increasing amount of urea molecules favors PNIPAM dehydration also due to the replacement of water by urea in the first solvent shell. These conclusions are additionally supported by the stagnating values of \( \phi_{\text{sw}} \) at high urea concentrations (Fig. 4).

In order to evaluate the influence of urea binding on the shift of the chemical equilibrium towards the swollen or the collapsed state, the local/bulk partition coefficients were estimated according to eqn (16). The results reported in Fig. 8 prove that the local/bulk partition coefficients \( K_p(r) \) for low \( c_u \) at both temperatures are smaller than 1 at all distances. This confirms that urea is preferentially excluded from the PNIPAM backbone at low concentrations, in perfect agreement with the small number of urea–PNIPAM hydrogen bonds reported in Fig. 7. The highly stretched PNIPAM conformations observed at low urea concentrations \( (c_u = 0.5 \text{ mol L}^{-1}) \) in Fig. 4 can be therefore attributed to the weaker urea exclusion around the swollen conformation in accordance to eqn (13). Although urea is very poorly bound to PNIPAM, the interaction with the water molecules around the macromolecule leads to volume exclusion effects, which contribute to enhance the chain stretching. The swollen conformation is therefore more favored compared to the collapsed state. A slightly stronger urea accumulation at short distances \( (r \leq 1 \text{ nm}) \) is found for high \( c_u \), which is more pronounced at high temperature, and demonstrates the preferential binding of urea to PNIPAM chains. However, the preferential exclusion of urea for low \( c_u \) results in a slight dehydration, as demonstrated by the water accumulation behavior according to the KB integrals presented in Fig. 6 and the negative values for ΔH-bonds (water). This finding is in good agreement with the general explanation for the occurrence of a coil to globule transition for PNIPAM and the shift to lower transition temperatures. It can be assumed that urea molecules are preferentially excluded from PNIPAM at low \( c_u \) where it is energetically unfavorable to replace water with urea molecules. The situation changes for high \( c_u \) at high temperature, where a significant dehydration upon phase transition occurs and the strong binding of urea molecules is energetically favorable to compensate the loss of the water molecules. These two mechanisms are able to explain the two different slopes for the decrease of the transition temperature observed for different urea concentrations in Fig. 3.

The shift of the chemical equilibrium towards swollen or collapsed PNIPAM conformations as shown in Fig. 4 can be rationalized with regard to the local/bulk partition coefficient free energy difference (eqn (17)). It becomes obvious that only the high molar urea concentration at 328 K indicates a positive value of \( \Delta F_{K_p} \approx 0.2 \text{ kJ mol}^{-1} \) at a PNIPAM distance of \( r = 0.5 \text{ nm} \), which validates the energetic stability of the collapsed PNIPAM conformation. For all the other temperatures and concentrations, swollen conformations are preferred. Based on these findings, it can be concluded that the chemical equilibrium is slightly shifted towards the more swollen conformations for low \( c_u \) at both temperatures, in agreement with the experimental results (Fig. 4), whereas only high temperatures and high \( c_u \) induce a shift of the chemical equilibrium towards the collapsed state. The strong attraction of urea to PNIPAM in high concentrated solutions and at high temperatures may rationalize the previously discussed effects for urea in terms of a cross-linking agent between adjacent PNIPAM chains above a critical urea concentration. The presence of a direct urea binding can be also assumed with regard to the stagnating values for \( \phi_{\text{sw}} \) at \( c_u \geq 2 \text{ mol L}^{-1} \) which can be rationalized by a fully saturated urea shell around PNIPAM.
4 Summary and conclusions

In this work, the effect of urea on the conformational behavior of PNIPAM brushes was studied by combining optical experiments and atomistic molecular dynamics simulations. The reported results indicated a concentration-dependent binding behavior between urea and PNIPAM. In the low concentration regime (\(c_u \leq 0.5 \text{ mol L}^{-1}\)), urea is preferentially excluded from the PNIPAM backbone. Hence, the linear decrease of \(T_c\) was rationalized by indirect effects of urea, which influence the dehydration process at high temperature by dilution of bulk water. However, despite the low number of urea molecules around PNIPAM in the first solvent shell, it was found that a more stretched conformation is stabilized due to a more favorable urea accumulation around the swollen conformation and excluded volume effects, as a consequence of the water network around the macromolecule. 92 In contrast, in the high concentration regime a preferential binding of urea to PNIPAM was found at both temperatures. Due to the small number of water molecules bound to PNIPAM, it can be assumed that the shift of the transition temperatures at high \(c_u\) is mostly dominated by direct urea interactions with PNIPAM. Moreover, the high positive local/bulk partition coefficient of urea at 328 K supports the hypothesis of urea acting as a crosslinking agent between adjacent PNIPAM chains. 32,38,88

Thus, two basic mechanisms for different urea concentrations are proposed: (1) a preferential exclusion of urea from PNIPAM surface at low molar concentrations. This was rationalized by the experimental results, which supported a similar dehydration mechanism at low \(c_u\) as in pure water, proving the absence of a direct binding; (2) a preferential binding of urea to PNIPAM above a critical concentration (\(c_u \geq 2 \text{ mol L}^{-1}\)). This produces a further decrease of the transition temperature in agreement to previous findings, 38,82 and a cross-linking effect due to a favorable binding between adjacent chains with a similar mechanism as proposed earlier. 38,88 Although several differences between numerical and experimental studies exist, for instance the consideration of a single chain in contrast to polymer brushes, a reasonable agreement between the results was found. This qualitative coincidence is of crucial importance, as it demonstrates that the experimental outcomes are strongly influenced by single chain properties and not by collective effects of the brush. Specific effects like cross-linking between the polymer chains which have been reported for high urea concentrations 38,88 are not detectable by the simulations. However, indirect hints were found, like the strong accumulation of urea molecules around PNIPAM chains by a preferential binding mechanism, providing a reasonable explanation of the experimental observations. With regard to the two distinct accumulation regimes of urea, it has to be mentioned that explicit reasons for a similar behavior have been also recently discussed, 22 and with regard to charged co-solutes, the presence of specific binding mechanisms for chaotropic ions was demonstrated. 33,94

In conclusion, the conformational behavior and the transition temperatures of PNIPAM brushes are strongly affected by the presence of urea, either by indirect effects or by direct polymer–osmolyte interactions. This article demonstrates that the combined results of experiments and simulations point towards the most reliable methodology to study molecular mechanisms and their consequences for macromolecular conformations.

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