Stress-induced colouration and crosslinking of polymeric materials by mechanochemical formation of triphenylimidazolyl radicals†

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Under mechanical stress, the hexaarylbiimidazole (HABI) motif can cleave to triphenylimidazolyl (TPI) radicals when incorporated into a polymer matrix. The mechanically produced coloured radicals can initiate secondary radical reactions yielding polymer networks. Thus, the HABI mechanophore combines optical reporting of bond scission and reinforcement of polymers in a single molecular moiety.

Mechanical degradation is an inconvenient yet inevitable effect of repeated deformation when using polymer-based materials. Stress-induced bond scission, a microscopic process, eventually causes macroscopic failure of the material. Recent research demonstrates impressively how generally unspecific and undesired bond scission can be altered into constructive processes.1–5 Commonly, this is achieved by incorporating mechanically activatable molecular moieties (mechanophores) into polymer architectures. Those mechanophores then either act as probes for the visualization of stress or initiate the repair of the damaged material. Amongst the motifs for stress-report, functional moieties changing their optical properties upon the application of stress are highly desirable as they allow the visualization of areas under heavy loading.5–7 Consequently, the implementation of both the ability to visualize as well as repair broken bonds in a polymeric material can be seen as a significant advance in the field of materials science, as recently demonstrated by Weng, Boulatov and co-workers.11

Motifs relying on the generation of stabilized radicals are particularly attractive for this endeavour as they combine features favourable for optical detection and for the formation of new chemical bonds. While the delocalization of mechanically generated radicals into a π-system can produce strongly absorbing chromophores, the controlled reactivity of these radicals may be used to initiate constructive secondary reactions, such as polymerizations. The former was recently demonstrated by Otsuka and co-workers relying on the reversibly induced cleavage of the diarylbenzofuranone moiety15–17 while the latter is reported exploiting unspecific, ultrasound-induced polymer scission to unstabilized radicals in solution yielding end-capped, chain-extended or grafted polymers.18–21

So far only little attention has been paid to the use of hexaarylbiimidazole in polymer mechanochemistry. This motif reversibly cleaves into highly coloured triphenylimidazolyl (TPI) radicals upon the application of light, heat or pressure (Scheme 1a).22–25 Moreover, HABI derivatives are already employed as radical photoinitiators in photoimaging systems26 as well as for the realization of de-crosslinkable polymer networks.27,28 Most importantly however, the TPI’s ability to initiate radical polymerizations mediated by thiol chain transfer agents29 (Scheme 1b) renders the implementation of HABI to deliver stress report and bond repair highly promising.

Scheme 1 (a) Reversible dissociation of HABI into TPI radicals. (b) Polymerization initiated by TPI radicals.
Here we show the mechanical activation of this motif by transmitting force employing different polymer architectures, in solution as well as in physically crosslinked networks in the solid state, to achieve optical report over the mechanical activation and exploit the bond-scission to initiate radical reactions.

To gain insight into the mechanochemical activity of the HABI motif, linear poly(methyl acrylate) chains containing the functional moiety in their centre (PMA-1) were synthesized for ultrasonication activation in the dissolved state (Scheme 2).

Alkylation of commercial 4-hydroxybenzaldehyde with 6-chlorohexanol yielded ether 4 which, in turn, reacts with benzil followed up by saponification of the phenyl acetate side-product giving imidazole 3.30 Esterification with z-bromoisobutyryl bromide to initiator 2 and subsequent Cu(0)-mediated living radical polymerization with methyl acrylate (SET-LRP) result in PMA-2 (Mn = 98 kDa, D = 1.06).31 Oxidative coupling to PMA-1 was carried out with PbO2 after the polymerization step instead of before because we found that the dynamic equilibrium of HABI with its TPI radicals is incompatible with the conditions of most living radical polymerizations.32 Relatively high molecular weights for PMA-2 were chosen as we discovered that increasing the molecular weight improved the coupling efficiency. Due to the HABI’s dynamic character and the relatively high dilution of the functional groups, conversion to the coupled product was generally incomplete and roughly 50% of PMA-2 remained in the final product.

For experiments in the solid state physically crosslinked polyurethane PU-1 was prepared by oxidation of imidazole 3 with K3Fe(CN)6 to give diol 1, which then yielded the final product in a polycondensation reaction with hexamethylene diisocyanate and poly(tetrahydrofuran) (Mn = 2 kDa) using a procedure employed by us in an earlier work.33

Mechanochemical activation in solution was carried out by ultrasonication of PMA-1 in benzene. A first indication that the desired cleaving process is occurring at the mechanophore can be gathered by monitoring the gel permeation chromatography (GPC) refractive index (RI) detector traces (Fig. 1a). The high molecular weight peak situated at around 200 kDa decreases rapidly and adds to the original low molecular weight peak of PMA-2 at 100 kDa. We attribute the peak-shoulder arising at around 50 kDa to unspecific secondary chain-scission due to the large contour length of PMA-2.1 H-NMR measurements clearly confirm that scission occurs mostly at the HABI-moiety (Fig. 1b) by revealing the loss of the characteristic protons at 7.8 and 8.0 ppm leaving only the TPI signals of PMA-2. Additionally, thermal activation was ruled out by keeping a benzene solution of PMA-1 at room temperature without applying ultrasound irradiation proving that thermal dissociation to the TPI radicals cannot be the origin of the observed cleaving process as the HABI’s characteristic signals can still be retrieved. To explore the mechanochemical origin of the sonochemical activation of PMA-1 and to further rule out thermal effects small molecule 1 was subjected to similar sonication conditions as PMA-1 (Fig. 1c). Since the limiting molecular weight for mechanochemical activation is typically around 30–40 kDa, we expected this low-molecular weight mechanophore not to show mechanoresponsive behaviour.34,35 Indeed, from the 1H-NMR spectra it becomes clear that neither thermal dissociation nor irradiation with ultrasound of 1 lead to the formation of TPI or other considerable degradation of the HABI-moiety during the monitored time span.

Having confirmed that the HABI-motif can be employed as potent mechanophore we then advanced our system to the solid state. Initial solid state activation experiments on PU-1 were carried out by compression with a pellet press and a blue coloration of the compressed area indicating formation of the TPI radicals can be observed (Fig. S4, ESI†). For reference purposes a similar polyurethane with solution-blended instead

![Scheme 2](#) Synthesis of HABI-containing target polymers PMA-1 and PU-1.
of covalently attached HABI was prepared and also gave a slight coloration during compression. Even though the magnitude of activation appears to be far less in the reference experiment, thermo- or piezochromic effects during compression could hence not completely be ruled out as a source of TPI radical generation.

Inspired by Otsuka and co-workers, we eliminated these effects by rapid freezing of solvent-swollen networks to mechanochemically activate PU-1. Freezing constrains the polymer chains in their translational motion rendering rate of macromolecule migration considerably lower than the rate of solvent migration. Consequently the solvent crystallizes thus straining the polymer chains. For this purpose PU-1 as well as polyurethane containing solution-swollen HABI were first swollen in DMF and then immersed into liquid N₂. From Fig. 2a and c it becomes clear that the reference polyurethane with HABI not covalently attached to the polymer backbone does not change its colour upon freezing and thus does not cleave the HABI-moiety through the application of mechanical stress. However, Fig. 2b and d clearly indicate the formation of coloured TPI radicals after freezing of PU-1 unequivocally confirming the mechanochemical origin of the observed activation. Despite the fast radical recombinination of TPI radicals to form HABI in systems with limited translational diffusion, we could achieve stress-sensing with a weight-fraction of only 4%.

As freezing of the swollen gel yields the most unambiguous results, proof of concept that a radical reaction can be induced mechanochemically was provided by freezing PU-1 swollen with a mixture of ethylene glycol dimethacrylate (EGDMA) and 50 mol% methyl 3-mercaptopropionate (MMP). Once the TPI radicals are generated, hydrogen abstraction of the thiol occurs which then initiates the polymerization to a poly(EGDMA) network (compare Scheme 1b). Note that comparatively high amounts of thiol were needed to achieve a sufficient amount of growing chains as the initiation is mostly limited by diffusion of the thawing MMP to the TPI moieties. A first indication that this cascade is initiated by mechanical stress was gathered by comparing the colours of the frozen samples of PU-1 swollen in EGDMA in the presence or absence of MMP upon thawing (Fig. S5a, ESI†). While a reference sample of PU-1 swollen in EGDMA turns brightly green, PU-1 swollen in EGDMA with 50 mol% MMP decolourizes rapidly suggesting that TPI radicals are formed in the latter system but quickly abstract a hydrogen atom from MMP. The consecutive initiation of crosslinking by MMP radicals is strongly indicated by the formation of a DMF-insoluble white residue after 3 successive freezing and thawing cycles in PU-1 swollen with both MMP and EGDMA (Fig. S5b, ESI†), but not in PU-1 swollen with EGDMA alone.

FTIR-spectroscopy of this residue confirms the absence of the signals indicative for PU-1 suggesting complete dissolution of the polyurethane (Fig. 3a and b). However, the typical C=O stretch (1719 cm⁻¹) and C=C stretch (1638 cm⁻¹) absorptions of EGDMA can be retrieved in the solid residue pointing towards the formation of a poly(EGDMA) network. Moreover, the successful polymerization of EGDMA is indicated by the decreasing intensity of the C=C stretch absorption as compared to the C=O stretch signal confirming the loss of double bonds relative to carbonyl moieties (Fig. 3b).

To further verify that this residue is actually produced by HABI-initiated radical polymerization, we subjected PU-1 to swelling in a mixture of methyl acrylate (MA) and 50 mol% MMP. After mechanical activation by freezing this allowed us to analyse the obtained sample by ¹H-NMR spectroscopy. Fig. 3c clearly shows the vanishing of MA’s characteristic acrylic protons (5.80–6.50 ppm) as well as a shift of MA’s methyl-group (3.77 ppm). Concurrently, broad signals indicative of
PMA’s alkylic backbone (1.8–2.5 ppm) appear. Together with the remaining triplet at 1.65 ppm, suggesting incomplete consumption of the thiol, and GPC measurements (Fig. S6, ESI†) this unequivocally demonstrates a successful mechanically induced radical oligomerization.

In conclusion, employing the mechanoresponsive HABI moiety we showed here for the first time that mechanochemically generated radicals in polymeric architectures can be used for stress-sensing giving rise to coloured stable radicals as well as for the initiation of a secondary radical reaction. We believe that this proof of concept contributes considerably to the fundamental understanding of how we can alter the destructive nature of mechanical stress to perform constructive chemistry and anticipate that stress-induced crosslinking with the HABI mechanophore may be applied even more broadly when combined with non-volatile acrylate and thiol species.

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References