Rearrangement of \{\alpha\text{-P}_2\text{W}_{15}\} to \{\text{PW}_6\} moieties during the assembly of transition-metal-linked polyoxometalate clusters†

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Polyoxometalates (POMs) are molecular oxide units generally formed by highly oxidised transition metal atoms such as W\(^{6+}\), Mo\(^{5+/6+}\), etc. that yield in most cases highly charged anionic clusters. These clusters have a wide range of different architectures and may possess properties applicable to diverse fields such as catalysis, material science, nanotechnology, magnetism, and medicine. By incorporating other metals to make transition metal substituted polyoxometalates (TMSPs), the structural diversity and functionality of the anions is increased. Many TMSPs are based upon either Keggin\(^{1}\) or Dawson\(^{6}\) fragments, this is due to the wide library of lacunary building blocks available\(^{10}\) for both of these parent architectures. One of the interesting aspects of using lacunary precursors is the transformations that can occur when reacting with transition metals, but this is more commonly seen for Keggin type starting materials.\(^{11}\)

Herein we report two structurally analogous compounds that show the previously unseen transformation of the Dawson-type \{\alpha\text{-P}_2\text{W}_{15}\} precursor to a Keggin-type \{\text{PW}_6\} fragment. Both the cobalt and manganese containing TMSPs are obtained using a synthesis adapted from previously reported methodology.\(^{12}\) The solution behaviour of the cluster is performed using ESI-MS and structural characterization is performed by XRD crystallography.

Compounds 1 and 2, Na_{14}Li_{14}[Co\text{(P}_2\text{W}_{15}\text{O}_{49})(\text{H}_2\text{O})_2]\text{23}^- and Na_{14}Li_{14}[Mn\text{(P}_2\text{W}_{15}\text{O}_{49})(\text{H}_2\text{O})_2]\text{23}^- \text{are hydrated 5.8H}_2\text{O} and 5.8H}_2\text{O} are analogues, possessing a V-shaped sandwich structure (Fig. 1).

The general formula for the anion is [M\text{(P}_6\text{W}_{20})(\alpha\text{-P}_2\text{W}_{15}\text{O}_{49})(\text{H}_2\text{O})_2]\text{23}^- \text{with } M = \text{Co}^{3+}(1\text{a}), \text{Mn}^{2+}(2\text{a}). These anionic units consist of two \{\alpha\text{-P}_2\text{W}_{15}\} units each connected by three transition metal atoms to a central \{\text{PW}_6\} species, creating a bent architecture. Both clusters belong to the C\text{2v} point group and crystallise in the \text{P1} space group. For these compounds the bending angle as defined between the phosphorus atoms of the \{\alpha\text{-P}_2\text{W}_{15}\} and the phosphorus of the central \{\text{PW}_6\} unit, is 145° for \text{1a} and 141° for \text{2a}. The P–O bond lengths of the central \{\text{PW}_6\} unit are between 1.51–1.55 Å for \text{1a} and 1.54–1.58 Å for \text{2a}. The M–O bond lengths are between 2.00–2.21 Å for \text{1a} (M = \text{Co}) and 2.10–2.29 Å for \text{2a} (M = \text{Mn}). The terminal M–O bond lengths are 2.11–2.13 Å for \text{1a} (M = \text{Co}) and 2.17–2.19 Å for \text{2a} (M = \text{Mn}) which is indicative that these oxygen atoms are water molecules. This is confirmed by bond valence sum
(BVS) calculations, which also confirm that all the Co and Mn atoms are in the $2^+$ oxidation state.

Compounds 1 and 2 were synthesised using $\text{Na}_{12}[\text{x-P}_2\text{W}_{15}\text{O}_{56}]24\text{H}_2\text{O}$ as the lacunary precursor in the presence of the appropriate transition metal ions, lithium chloride and sodium phosphate as a slurry in a minimal amount of solvent in the pH range 6–7. If the pH value is raised to values higher than pH 7.5 then $\text{Na}_3\text{Li}_1\text{H}_2[\text{M}_6(\text{H}_2\text{O})_6][\text{x-P}_2\text{W}_{15}\text{O}_{56}]_3\text{H}_2\text{O}$ crystallises.

For manganese, this trimer (compound 3, see Fig. 3) was observed only a few times in low yield and full characterisation has not been possible due to lack of material. We can thus see that a small pH change can lead to a different set of structures and we deduce that pH control is the most important factor for the synthesis of these compounds. It is worth noting that in all these reactions, the $\text{M}_i(\text{P}_2\text{W}_{15})_3$ sandwich was also obtained as a by-product, which is expected given that $\text{P}_2\text{W}_{15}$ is not known to reorganise easily.

The nature of the synthetic strategy used makes it difficult to follow the reaction using analytical methods that would give an insight into the mechanism. This is because the reaction requires such a large excess of $\text{Na}_{12}[\text{x-P}_2\text{W}_{15}\text{O}_{56}]24\text{H}_2\text{O}$ that any $\text{in situ}$ analysis was unfeasible due to the super-saturation and turbidity of the reaction mixture.

There are a few examples in the literature that report banana or V-shaped sandwiches, but all of these architectures are based on Keggin building blocks. All of them have the common or V-shaped sandwiches, but all of these architectures are based on Keggin building blocks.15 All of them have the common template generating a $\{\text{PW}_6\}$ unit.18

In order to investigate the stability of compounds 1 and 2 in solution, electrospray ionisation mass spectrometry (ESI-MS) analysis was conducted. Ionisation for compound 1 was successful and a full spectrum of assignable envelopes was obtained. The ESI-MS spectrum (see Fig. 2, top) shows three envelopes that correspond to the intact cluster. These can be found at the charge values of $-6$, $-7$, and $-8$ at $m/z$ of 1599.9 $[\text{Na}_5\text{Li}_1\text{H}_2\text{Co}(\text{PW}_8\text{O}_{26})(\text{x-P}_2\text{W}_{15}\text{O}_{56})(\text{H}_2\text{O})_3][\text{H}_2\text{O}]_3\text{H}_2\text{O}$, 1363.5 $[\text{Na}_3\text{Li}_2\text{H}_2\text{Co}(\text{PW}_8\text{O}_{26})(\text{x-P}_2\text{W}_{15}\text{O}_{56})(\text{H}_2\text{O})_3][\text{H}_2\text{O}]_3\text{H}_2\text{O}$ and 1187.2 $[\text{Li}_1\text{H}[\text{PW}_8\text{O}_{26}(\text{x-P}_2\text{W}_{15}\text{O}_{56})(\text{H}_2\text{O})_2][\text{H}_2\text{O}]_3\text{H}_2\text{O}]$. Other peaks at lower $m/z$ values can also be found and they correspond to fragment $[\text{Mn}_3\text{P}_2\text{W}_{15}]$ species. These can be found at the charge values of $-5$, $-5$ and $-4$ at $m/z$ of 817.0 $[\text{Li}_2\text{H}_2\text{Co}(\text{PW}_8\text{O}_{26})(\text{x-P}_2\text{W}_{15}\text{O}_{56})(\text{H}_2\text{O})_2][\text{H}_2\text{O}]_3\text{H}_2\text{O}$ and 1025.5 $[\text{Na}_5\text{Li}_1\text{H}_2\text{Co}(\text{PW}_8\text{O}_{26})(\text{x-P}_2\text{W}_{15}\text{O}_{56})(\text{H}_2\text{O})_3][\text{H}_2\text{O}]_3\text{H}_2\text{O}$ (see ESI† Section 5).

Compound 3 is a trimeric structure that consists of a core of nine manganese centres templated by two phosphate units (see Fig. 3). This core is encased by three $[\text{x-P}_2\text{W}_{15}]$ units. The full formula of this compound cannot be given due to a lack of reproducibility and very low yields resulting in a full characterisation not being possible. The formula derived solely from crystallographic analysis is
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References