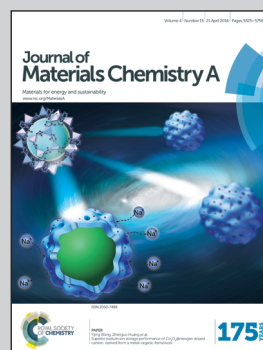


Showcasing a study on the assembly of hierarchical carbon nanotube constructs by Prof. Milo Shaffer at Imperial College London and Prof. Sulaiman Basahel at King Abdulaziz University.

Title: Cross-linked single-walled carbon nanotube aerogel electrodes *via* reductive coupling chemistry

A new approach to cross-linking converts solutions of individualised single-walled carbon nanotube anions into covalent, carbon-bonded aerogels, with low density, high surface area, and high conductivity, showing promise as supercapacitor electrodes.

As featured in:



See Milo S. P. Shaffer et al.,
J. Mater. Chem. A, 2016, 4, 5385.



www.rsc.org/MaterialsA

Registered charity number: 207890

CrossMark
click for updatesCite this: *J. Mater. Chem. A*, 2016, 4, 5385Received 16th December 2015
Accepted 18th February 2016

DOI: 10.1039/c5ta10311h

www.rsc.org/MaterialsA

Cross-linked single-walled carbon nanotube aerogel electrodes *via* reductive coupling chemistry†

Martina De Marco,^a Foivos Markoulidis,^a Robert Menzel,^{bc} Salem M. Bawaked,^d Mohamed Mokhtar,^d Shaeel A. Al-Thabaiti,^d Sulaiman N. Basahel^d and Milo S. P. Shaffer^{*a}

Single-walled carbon nanotube (SWCNT) anions can be cross-linked by a dielectrophile to form covalent, carbon-bonded organogels. Freeze-drying produces cryogels with low density (2.3 mg cm^{-3}), high surface area ($766 \text{ m}^2 \text{ g}^{-1}$), and high conductivity (9.4 S m^{-1}), showing promise as supercapacitor electrodes. Counterion concentration controls debundling, grafting ratio, as well as all the resulting properties.

In the past decade, single-walled carbon nanotubes (SWCNTs) have been assembled into three-dimensional architectures, in order to combine their intriguing properties (mechanical,^{1,2} thermal³ and electrical,⁴ very high aspect ratio and high surface area⁵) with the unique features of porous networks.⁶ To this end, SWCNTs have been trapped within other gel-forming systems such as polymers,^{7–9} DNA,¹⁰ surfactants,¹¹ silica¹² and resorcinol-formaldehyde.¹³ To avoid the contaminating, often insulating, secondary gel-former, SWCNTs have been cross-linked using a range of strategies, including cycloaddition reactions,¹⁴ Sonogashira coupling,¹⁵ linking of pre-introduced carboxyl¹⁶ or hydroxyl¹⁷ functionalities with bi- (or tri-)functional reagents, and even dry *via* electron beam irradiation.¹⁸ However, existing methods tend to degrade the SWCNTs aspect ratio, damage the carbon framework and debundle inefficiently. A network derived from individualised, covalently interconnected, high aspect ratio SWCNTs is the ideal target for maximising surface area and accessible porosity, while retaining high conductivity, important for applications including catalysis, energy storage, controlled release systems and other multifunctional devices.^{19–21}

Overcoming the intrinsic, strong, van der Waals interactions ($\sim 40k_B T \text{ nm}^{-1}$) that drive SWCNT bundle formation is challenging.²² Ultrasonication is damaging, hard to scale and yields only low concentrations of individualised SWCNTs.²³ Covalent modification can limit bundling,^{24–26} though many routes also introduce damage,²⁷ or only functionalise exposed bundle surfaces. In contrast, chemical reduction is a promising functionalisation strategy that can generate true solutions of individualised, negatively-charged SWCNT ions (nanotubides)^{28–30} at high concentrations, avoiding the damage caused by sonication or aggressive oxidation.³¹ Nanotubides can react with an extensive variety of functional groups including alkyl/aryl halides,^{32,33} peroxides,³⁴ disulfides,³⁵ epoxides³⁶ *etc.*; however, cross-linking reactions *via* this route have not yet been explored. The present study uses difunctional reagents to cross-link nanotubide solutions to form fully carbon-bonded covalent organogels.

Liquid ammonia,³⁷ THF³⁸ and DME³⁹ are the most common solvents employed for reductive chemistry; however *N,N*-dimethylacetamide (DMAc) has recently been shown to be particularly suitable for a convenient, one-pot reaction,⁴⁰ using Na/naphthalene as a charge transfer agent. *p*-Diiodobenzene (*p*-DIB) was selected as a soluble, rigid, potentially conjugated, dielectrophilic cross-linking agent, using the established iodide reactivity (Fig. 1). A relatively high nominal SWCNT concentration (0.15 M) was selected to increase chances of network formation, whilst remaining within the range found to be effective for simple nanotubide grafting reactions.^{40,41}

^aDepartment of Chemistry, Imperial College London, UK. E-mail: m.shaffer@imperial.ac.uk

^bSchool of Chemistry, University of Leeds, LS2 9JT, UK

^cBio Nano Consulting, The Gridiron Building, One Pancras Square, London, N1C 4AG, UK

^dDepartment of Chemistry, Faculty of Science, King Abdulaziz University, Jeddah, Saudi Arabia

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c5ta10311h

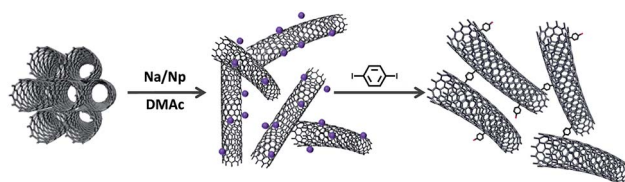


Fig. 1 Representation of the reaction sequence.

Since the sodium concentration ($[Na]$) and/or degree of charge is known to control the level of solubilisation and functionalisation, a range of $[Na]$ was tested, adding one equivalent of *p*-DIB (to Na) for the cross-linking, in order to identify the optimum conditions. At higher $[Na]$, the immediate gelation of the network, following the addition of *p*-DIB, indicated successful cross-linking of the SWCNTs, through the double covalent attachment of phenyl groups. In contrast, using mono-substituted iodobenzene (IB) as the electrophile (SWCNT-IB control) led only to a liquid solution of functionalised CNTs (photograph S1c†).

During thermo-gravimetric analysis, under nitrogen, the grafted species decompose and desorb from the surface, as identified by mass spectroscopy (TGA-MS, Fig. 2a and S1†). A SWCNT control, treated in the same reaction conditions but without adding *p*-DIB (treated SWCNTs), has a very similar TGA trace to the pristine SWCNTs, indicating that the charging process has little effect on the CNT structure. A control adding *p*-DIB to the quenched nanotubes showed a very similar TGA to the treated SWCNTs, excluding a contribution from physisorption. The grafting ratio (GR), the weight ratio of the grafted species to the nanotube framework, was determined from the TGA data. GR and tendency to gel is correlated with $[Na]$, and the corresponding C : Na ratio (Fig. 2b). Although TGA shows that all the samples are functionalised, the samples at low sodium concentration ($[Na] < 9$ mM; (C : Na > 20)) did not gel. Selecting the correct $[Na]$ (and hence C : Na ratio) is crucial to maximise the availability of negative charges on the nanotube

surface and to prevent 'salting out' due to cation condensation on the nanotube surface.⁴² Recent work has shown that $[Na]$ is the critical parameter determining the level of functionalisation across all carbon concentrations.^{40,41} In the present study, the optimum is observed for an absolute metal concentration of 15 mM, providing a GR of 17 wt%.

Raman spectroscopy confirmed that covalent modification has occurred, without introducing excessive damage due to side reactions (Fig. S2†). The intensity ratio (I_D/I_G) of the D band (1293 cm^{-1}) over the tangential mode G band (1590 cm^{-1}) is commonly taken as an indication of the introduction of sp^3 (or other) defects. The I_D/I_G ratio is consistently slightly raised (0.16–0.20) after the cross-linking reaction, compared to the treated (0.06) and pristine (0.05) SWCNTs, and is maximised at $[Na] \sim 15$ mM, following the trend in GR.

Following oxygen quenching, the SWCNT DMAc-organogels were converted to hydrogels by a careful sequence of solvent exchanges (DMAc–THF–water, Fig. S3†) then freeze-dried, in order to preserve the 3D network structure from capillary collapse during solvent evaporation. The resulting highly porous (99.8%, Table S6†), low density ($2\text{--}3\text{ mg cm}^{-3}$) cryogels exhibited minimal shrinkage ($\sim 1\%$) and retained defined, cylindrical shapes (Fig. S3b†), suggesting a network robustness derived from successful chemical cross-linking of the SWCNTs. Scanning Electron Microscopy (SEM) showed a hierarchical interconnected, ordered porous honeycomb structure, with straight and micrometer-sized parallel channels, formed during freezing, and walls of networked SWNTs (Fig. 3a and b). Transmission Electron Microscopy (TEM) images (Fig. 3c) confirmed successful network formation between small bundles and individual nanotubes, with much lower degree of bundling than the starting material (Fig. S5†).

XPS analysis identified residual iodine in the samples (Fig. S4†); however, direct iodination of the nanotube sidewalls is unlikely considering the absence of iodine in the XPS of the SWCNT-IB control. In the single electron transfer mechanism proposed for nanotube grafting,³⁵ the higher reactivity of

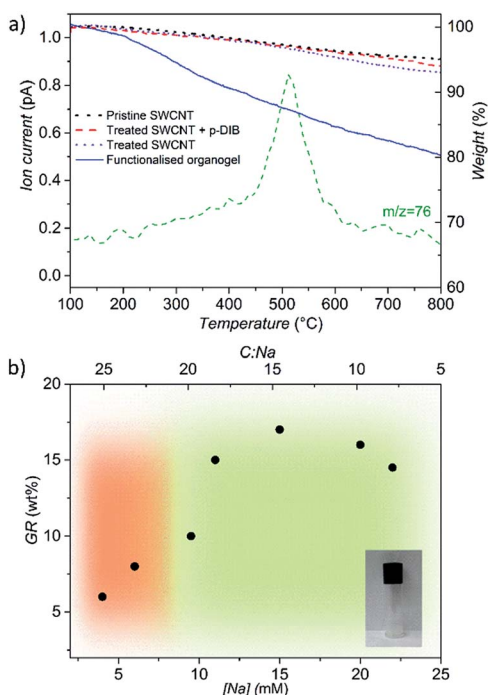


Fig. 2 (a) TGA of pristine SWCNTs (dotted line) and control samples (dashed lines). TGA-MS of representative functionalised organogel (10 wt%, solid line) shows a weight loss related to the presence of the functional groups (m/z 76, $C_6H_4^+$). (b) grafting ratios (wt%) versus $[Na]$. Green area indicates the zone of gel formation.

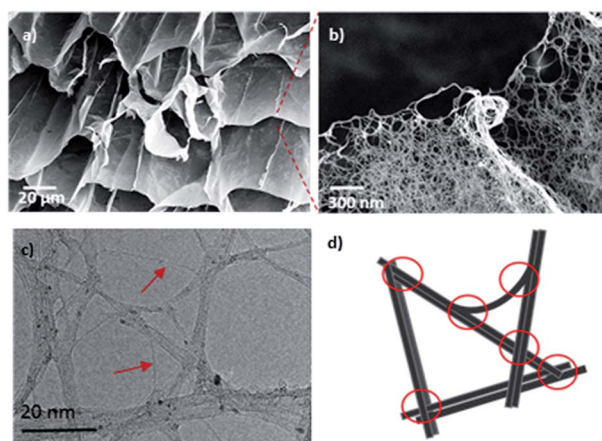


Fig. 3 Dried cryogel (17 wt%) (a and b) SEM images at different magnifications, (c) TEM and (d) corresponding schematic, red arrows indicating individual nanotubes, red circles indicating the junction points. Black particles are residual Fe catalyst.



benzene radicals should dominate the reaction. The iodine content (at%) in the cross-linked samples follows the general trend in grafting ratio (Fig. 4a), and can be related to the fraction of grafted phenyl groups that are cross-linked rather than monofunctionalised ('CL:mono', Fig. 4b & S4†). Di-reacted *p*-DIB at junctions (Fig. 3d) creates porosity by cross-linking the SWCNTs in a defined network, and all the associated iodine is subsequently washed out as NaI. However, in other locations, as the network forms, the SWCNTs are far from contact, and the rigid *p*-DIB can only react once, leaving phenyl iodide grafted to the sidewalls. The quantity of and distance between the junctions is presumably determined by extent of debundling, concentration and geometric constraints (related to the length of the CNTs and small size of *p*-DIB molecule). At higher [Na], where more salting out of bundles is expected, a higher proportion of cross-links are formed (greater CL:mono), most likely between parallel SWCNTs; randomly arranged, individualised SWCNTs offer a smaller fraction of sites for crosslinking. Interestingly, the degree of cross-linking correlates closely with the average pore size (nm) of the cryogels, calculated using Barrett, Joyner and Halenda analysis of the nitrogen adsorption data. Increased pore size can be attributed to a coarser network of more bundled/salted out SWCNTs, which is also reflected in the pore size distribution (Fig. S7†).

Nitrogen adsorption/desorption isotherms of the cryogels (Fig. 4d) are type IV, indicative of a mesoporous material. Consistent with a hierarchical structure, micropores are also present, evidenced by the steep isotherm profile at low relative pressure. Specific surface areas (SSAs) calculated according to the Brunauer–Emmett–Teller (BET) equation, range between 695 and 766 m² g⁻¹ (Table S6†), more than three times the SSA of the treated SWCNT control (217 m² g⁻¹) and double the SWCNT-IB control (366 m² g⁻¹), reflecting the effectiveness of cross-linking in preventing rebundling. The highest surface

area is achieved for the sample charged with [Na] 15 mM (17 wt%), further confirming the importance of the sodium concentration in effective dissolution.

Interestingly, the volumetric electrical conductivity through the monolith is high (9.4 S m⁻¹ at 10% of compression, Table S6†), given the low density, reflecting the intimate SWCNT cross-links. Smaller diameter bundles in networks have been reported to have improved electrical contacts, and therefore higher conductivity,^{43,44} consistent with the trend observed in Fig. 4c. It is important to note that all the trends in Fig. 4a–c, grafting ratio, degree of cross-linking, surface area, and electrical conductivity, as well as I_D/I_G , are closely correlated, and consistent with a maximised degree of individualisation in the network at an optimum [Na] of 15 mM.

As a preliminary illustration of the opportunities for electrochemical applications, a full working supercapacitor cell was assembled employing cryogel derived from [Na] = 15 mM as the active electrodes. The monolithic aerogels were combined with whisker-based current collectors to form a complete, symmetric, binderless electric double layer capacitor (EDLC), with 1 cm² active electrode area (SI), and low charge transfer resistance (ESI, Fig. S9†). The specific electrode capacitance, calculated from cyclic voltammetry at 1 V s⁻¹, was 55 F g⁻¹ (Fig. S8a†), which is similar to recently reported values for denser aerogels incorporated in EDLC devices.⁴⁵ One of the attractive features of such devices is their excellent cycling stability; indeed, after 10⁵ cycles of galvanostatic charge-discharge (GCD), the nanotubide-derived electrode retained 88% of the initial capacitance and a Coulombic efficiency of 95% (Fig. 5), in agreement with similar energy storage devices,⁴⁶ and confirming a robust and inherently porous structure.

Moreover, due to the open porous network, the GCD curves display a symmetrical shape at high applied current, highlighting the high rate stability of the cryogel for EDLC applications (Fig. S8b†). Promising results were also obtained for GCD-derived energy density (E_d) and power density (P_d) (Fig. S8c†), showing good stability over long term cycling even in the absence of binders; the present device has comparable performance to current other CNT-based supercapacitors (Fig. S8d†). The results are a proof of concept that the directly networked SWCNTs can form fully functioning, binderless supercapacitors. Overall performance could be improved, for example by using different electrolytes (e.g. low-viscosity ionic liquids or aqueous systems), or tuning the electrode geometry.

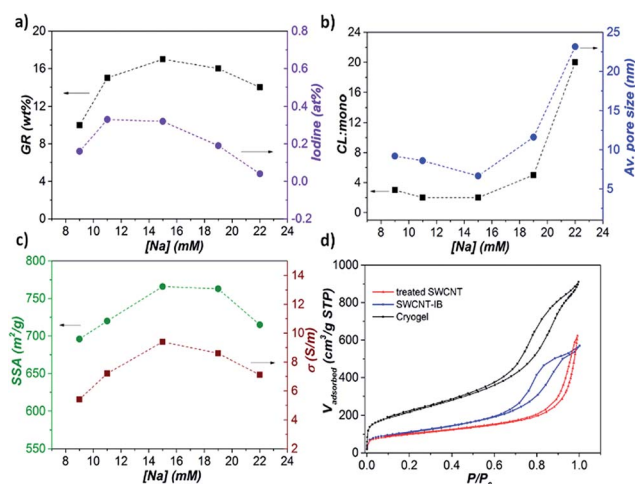


Fig. 4 (a) Plot of GR (wt%) and iodine (at%) versus [Na], (b) estimated ratio of cross-linked/mono reacted phenyl groups (CL:mono, refer to Fig. S4† for the calculation) and average pore size (nm, Fig. S7†) versus [Na], (c) specific surface area (SSA, m² g⁻¹) and electrical conductivity (S m⁻¹) versus [Na], (d) N₂ adsorption/desorption isotherm of cryogel (17 wt%), SWCNT-IB and treated SWCNT.

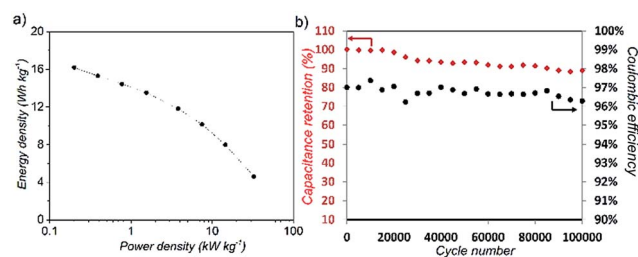


Fig. 5 (a) Ragone plot of GCD-derived power and energy densities (per kg of active material), (b) plot of capacitance retention (%) and Coulombic efficiency (%) over 10⁵ cycles of charge discharge at 50 mA.



The use of supercritical rather than freeze-drying may enhance accessible area further, and offers routes to improve volumetric performance as energy storage electrodes.

Conclusions

The present work reports a successful new method of producing SWCNTs cryogels by cross-linking reduced salts with *p*-DIB and subsequent freeze-drying. This approach is particularly attractive compared to previous cross-linking strategies, as the SWCNTs can be individualised without shortening or other framework damage. The cross-linking reaction is intrinsically enabled by the charge-based dissolution mechanism, and leads to a direct carbon-carbon bonded network throughout the gel, maximising robustness and electrochemical stability. The unreacted iodide component may provide a useful site for further modification. The behaviour of the system expands and is consistent with developing understanding of nanotubide polyelectrolytes. A sodium concentration of 15 mM is found to maximise the debundling and the grafting ratio, leading to material with enhanced surface area ($766 \text{ m}^2 \text{ g}^{-1}$) and an abundant pore distribution at lower pore sizes (10–22 nm) estimated from BET (N_2) measurements. These characteristics, in combination with good electrical conductivity (9.4 S m^{-1}), low mass density (2.3 mg cm^{-3}) and chemical stability, make the cryogels promising candidates for a variety of electrochemical energy-conversion technologies, as well as environmental applications such as sorption, separation and filtration. Preliminary electrochemical tests of EDLCs assembled with the current cryogel electrodes show good performance and stability over long cycles (10^5) of charge-discharge. The use of the same strategy with more concentrated, cleaner, more exfoliated and higher aspect ratio nanotubide solutions, different types and degrees of cross-linkers, and different drying protocols will allow the morphology of the aerogels to be optimised for different applications.

Acknowledgements

The authors would like to thank Edward White for TEM images and Adam Clancy for discussion during the experiments. Financial support for this project was provided by the Deanship of Scientific Research at King Abdulaziz University (Grant D-1-434), and EPSRC grant EP/L001896/1.

Notes and references

- J. P. Salvetat, J. M. Bonard, N. H. Thomson, A. J. Kulik, L. Forró, W. Benoit and L. Zuppiroli, *Appl. Phys. A*, 1999, **69**, 255–260.
- J.-P. Salvetat, G. A. D. Briggs, J.-M. Bonard, R. R. Bacsa, A. J. Kulik, T. Stöckli, N. A. Burnham and L. Forró, *Phys. Rev. Lett.*, 1999, **82**, 944–947.
- J. Hone, M. Whitney, C. Piskoti and A. Zettl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, R2514–R2516.
- A. Javey, J. Guo, Q. Wang, M. Lundstrom and H. Dai, *Nature*, 2003, **424**, 654–657.
- A. Peigney, C. Laurent, E. Flahaut, R. R. Bacsa and A. Rousset, *Carbon*, 2001, **39**, 507–514.
- S. Nardecchia, D. Carriazo, M. L. Ferrer, M. C. Gutierrez and F. del Monte, *Chem. Soc. Rev.*, 2013, **42**, 794–830.
- Y. Li, D. Yang, A. Adronov, Y. Gao, X. Luo and H. Li, *Macromolecules*, 2012, **45**, 4698–4706.
- A. Koganemaru, Y. Bin, Y. Agari and M. Matsuo, *Adv. Funct. Mater.*, 2004, **14**, 842–850.
- J. Zou, J. Liu, A. S. Karakoti, A. Kumar, D. Joung, Q. Li, S. I. Khondaker, S. Seal and L. Zhai, *ACS Nano*, 2010, **4**, 7293–7302.
- G. N. Ostojic, *ChemPhysChem*, 2012, **13**, 2102–2107.
- M. B. Bryning, D. E. Milkie, M. F. Islam, L. A. Hough, J. M. Kikkawa and A. G. Yodh, *Adv. Mater.*, 2007, **19**, 661–664.
- J. G. Duque, C. E. Hamilton, G. Gupta, S. A. Crooker, J. J. Crochet, A. Mohite, H. Htoon, K. A. D. Obrey, A. M. Dattelbaum and S. K. Doorn, *ACS Nano*, 2011, **5**, 6686–6694.
- M. A. Worsley, J. H. Satcher and T. F. Baumann, *Langmuir*, 2008, **24**, 9763–9766.
- M. Holzinger, J. Steinmetz, D. Samaille, M. Glerup, M. Paillet, P. Bernier, L. Ley and R. Graupner, *Carbon*, 2004, **42**, 941–947.
- R. Kumar and C. N. R. Rao, *J. Mater. Chem. A*, 2015, **3**, 6747–6750.
- A. Satti, A. Perret, J. E. McCarthy and Y. K. Gun'ko, *J. Mater. Chem.*, 2010, **20**, 7941–7943.
- M. B. Jakubinek, B. Ashrafi, J. Guan, M. B. Johnson, M. A. White and B. Simard, *RSC Adv.*, 2014, **4**, 57564–57573.
- S. G. Miller, T. S. Williams, J. S. Baker, F. Solá, M. Lebron-Colon, L. S. McCorkle, N. G. Wilmoth, J. Gaier, M. Chen and M. A. Meador, *ACS Appl. Mater. Interfaces*, 2014, **6**, 6120–6126.
- K. H. An, W. S. Kim, Y. S. Park, J. M. Moon, D. J. Bae, S. C. Lim, Y. S. Lee and Y. H. Lee, *Adv. Funct. Mater.*, 2001, **11**, 387–392.
- E. Frackowiak and F. Béguin, *Carbon*, 2001, **39**, 937–950.
- L. Li, Z. Wu, S. Yuan and X.-B. Zhang, *Energy Environ. Sci.*, 2014, **7**, 2101–2122.
- L. Girifalco, M. Hodak and R. Lee, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2000, **62**, 13104–13110.
- H. C. Yau, M. K. Bayazit, J. H. G. Steinke and M. S. P. Shaffer, *Chem. Commun.*, 2015, **51**, 16621–16624.
- J. E. Riggs, Z. Guo, D. L. Carroll and Y.-P. Sun, *J. Am. Chem. Soc.*, 2000, **122**, 5879–5880.
- R. Bandyopadhyaya, E. Nativ-Roth, O. Regev and R. Yerushalmi-Rozen, *Nano Lett.*, 2002, **2**, 25–28.
- C. A. Dyke and J. M. Tour, *J. Phys. Chem. A*, 2004, **108**, 11151–11159.
- J. Zhang, H. Zou, Q. Qing, Y. Yang, Q. Li, Z. Liu, X. Guo and Z. Du, *J. Phys. Chem. B*, 2003, **107**, 3712–3718.
- F. Liang, A. K. Sadana, A. Peera, J. Chattopadhyay, Z. Gu, R. H. Hauge and W. E. Billups, *Nano Lett.*, 2004, **4**, 1257–1260.
- S. Fogden, C. A. Howard, R. K. Heenan, N. T. Skipper and M. S. P. Shaffer, *ACS Nano*, 2011, **6**, 54–62.
- D. Wunderlich, F. Hauke and A. Hirsch, *J. Mater. Chem.*, 2008, **18**, 1493–1497.



- 31 H. Yoon, M. Yamashita, S. Ata, D. N. Futaba, T. Yamada and K. Hata, *Sci. Rep.*, 2014, **4**, 3907.
- 32 B. Gebhardt, Z. Syrgiannis, C. Backes, R. Graupner, F. Hauke and A. Hirsch, *J. Am. Chem. Soc.*, 2011, **133**, 7985–7995.
- 33 J. Chattopadhyay, A. K. Sadana, F. Liang, J. M. Beach, Y. Xiao, R. H. Hauge and W. E. Billups, *Org. Lett.*, 2005, **7**, 4067–4069.
- 34 P. S. Engel, W. E. Billups, D. W. Abmayr, K. Tsvaygboym and R. Wang, *J. Phys. Chem. C*, 2008, **112**, 695–700.
- 35 J. Chattopadhyay, S. Chakraborty, A. Mukherjee, R. Wang, P. S. Engel and W. E. Billups, *J. Phys. Chem. C*, 2007, **111**, 17928–17932.
- 36 Y. Martinez-Rubi, B. Ashrafi, J. Guan, C. Kingston, A. Johnston, B. Simard, V. Mirjalili, P. Hubert, L. Deng and R. J. Young, *ACS Appl. Mater. Interfaces*, 2011, **3**, 2309–2317.
- 37 K. F. Kelly and W. E. Billups, *Acc. Chem. Res.*, 2013, **46**, 4–13.
- 38 D. Voiry, C. Vallés, O. Roubeau and A. Pénicaud, *Carbon*, 2011, **49**, 170–175.
- 39 J. M. Englert, C. Dotzer, G. Yang, M. Schmid, C. Papp, J. M. Gottfried, H.-P. Steinrück, E. Spiecker, F. Hauke and A. Hirsch, *Nat. Chem.*, 2011, **3**, 279–286.
- 40 A. J. Clancy, J. Melbourne and M. S. P. Shaffer, *J. Mater. Chem. A*, 2015, **3**, 16708–16715.
- 41 T. Morishita, A. J. Clancy and M. S. P. Shaffer, *J. Mater. Chem. A*, 2014, **2**, 15022–15028.
- 42 G. S. Manning, *J. Chem. Phys.*, 1969, **51**, 924–933.
- 43 P. E. Lyons, S. De, F. Blighe, V. Nicolosi, L. F. C. Pereira, M. S. Ferreira and J. N. Coleman, *J. Appl. Phys.*, 2008, **104**, 044302.
- 44 D. Hecht, L. Hu and G. Grüner, *Appl. Phys. Lett.*, 2006, **89**, 133112.
- 45 E. Wilson and M. F. Islam, *ACS Appl. Mater. Interfaces*, 2015, **7**, 5612–5618.
- 46 P.-C. Chen, G. Shen, Y. Shi, H. Chen and C. Zhou, *ACS Nano*, 2010, **4**, 4403–4411.

