Two-coordinate group 14 element(II) hydrides as reagents for the facile, and sometimes reversible, hydrogermylation/hydrostannylation of unactivated alkenes and alkynes†

Terrance J. Hadlington,*a Markus Hermann,b Gernot Frenking,*b and Cameron Jones*a

Reactions of the solution stable, two-coordinate hydrido-tetrylenes, \( \text{E(H)}(L^1) \) (E = Ge or Sn; \( L^1 = \text{Ar} \equiv \text{C}_6\text{H}_{13}(\text{H})\text{Cl} \)) with the protic reagents, \( \text{HCl}, \text{NH}_3 \) and \( \text{EtOH} \), lead to oxidative addition to the germanium(0) centre, and formation of the stable chiral germanium(v) complexes, \( \text{L}^1\text{Ge}(\text{C}_6\text{H}_{13})\text{H}[\text{Cl}] \) and \( \text{L}^1\text{Ge}(\text{Et})\text{H}[\text{IR}] \) (R = NH\(_2\) or OEt). In contrast, related reactions between \( \text{L}^1\text{SnEt} \) and Bu\(_4\)OH or TEMPOH (TEMP = 2,2,6,6-tetramethylpiperidinyl) proceed via ethane elimination, affording the tin(v) products, \( \text{L}^1\text{SnR} \) (R = OBU\(_4\) or OTEMP). In addition, the oxidation of \( \text{L}^1\text{Ge}(\text{C}_6\text{H}_{13}) \) and \( \text{L}^1\text{Sn}(\text{C}_6\text{H}_{13}\text{Bu}) \) with \( \text{O}_2 \) yields the oxo-bridged metal(v) dimers, \( \text{[L}^1\text{C}_6\text{H}_{13}\text{Ge}(	ext{µ-O})\text{]}_2 \) and \( \text{[L}^1\text{Bu}_2\text{C}_6\text{H}_{13}\text{Sn}(	ext{µ-O})\text{]}_2 \), respectively.

π-complex with the unsaturated substrate, prior to its insertion into the polar \( \text{B} = \text{H} \) linkage.† This mechanism has also been used to explain the typically observed cis/anti-Markovnikov addition of boranes to unsaturated hydrocarbons. While much less studied than boranes, a variety of electron deficient, polar hydride complexes of aluminium, the heavier group 13 metals, and the s\(^4\) and early d-block metals, have additionally been shown to be effective for the hydrometallation of alkenes and alkynes.

Considering that neutral group 14 element(IV) hydrides (e.g. \( \text{R}_3\text{EH} \)) do not possess any vacant valence orbitals, it is not surprising that they are poorly effective for the hydroelementation of alkenes and alkynes, at least in their own right. However, reactions of this type (particularly hydrosilylations) are of considerable synthetic importance, and can proceed, for example, in the presence of transition metal catalysts or radical initiators, and/or when subjected to UV irradiation or elevated temperatures.6,7

It would be a significant advantage if the addition of group 14 element-hydrogen bonds to unsaturated hydrocarbons could be effected in the absence of catalysts or initiators, and in a

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Introduction

The 1,2-addition of element-hydrogen bonds across the carbon–carbon unsaturations of alkenes and alkynes is of immense importance to organic synthesis. In this respect, and since Brown’s seminal work on the hydroboration of alkenes in the 1950’s, boranes have become the reagent of choice for the 1,2-addition of element-hydrogen bonds across the carbon unsaturations of alkenes and alkynes. One of the primary reasons for the efficacy of such hydroborations, is that electron deficient, three-coordinate boranes (R\(_3\)BH) possess an empty p-orbital which is thought to allow the formation of a loose

† Electronic supplementary information (ESI) available: Experimental procedures and characterisation data for all new compounds, full details of the computational studies. Crystal data, details of data collections and refinements. CCDC 1422735–1422742. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5sc03376d
facile manner under ambient conditions. The first hints that this might be possible came with the kinetic stabilisation of group 14 element(n) hydride complexes, a small number of which (e.g. I–V, Scheme 1) have been reported since the turn of the millennium.13 Of these, the three-coordinate silicon(n) hydride, I, has been shown to hydroisylate cyclopentene and a series of terminal olefins at elevated temperatures (70–120 °C) and in the presence of large excesses of the alkene substrate.8 The latter reactions give rise to mixtures of regioisomers, in which the anti-Markovnikov product predominates. In one case, i.e. the reaction with trimethylsilylethylene, the reaction proceeds via an isolated [2 + 1] cycloadduct, viz. the silirane \[ \{\eta^2-H_2C=CH(H)[SiMe_3]\} \], which exists in equilibrium with I and free \( H_2C=CH(H)[SiMe_3] \) at ambient temperature. With respect to hydrogermylation and hydrostannylation reactions, the three-coordinate species, II and III, have been cleanly and hydrolymattel active (ester substituted) terminal and internal alkynes at ambient temperature.11 Furthermore, the dimeric, three-coordinate metal(n) hydride complexes, IV and V, react with tert-butylethylene at ambient temperature over 48 hours to give the alkyl/aryl substituted ditetrelenes \[ [Ar \cdot E(CH_2CH_2Bu)]_2 \] \( (Ar'=C_6H_5(C_6H_5)Pr'2-2,6)_2-2,6; E = Ge or Sn) \). Contrastingly, after 48 hours, the reaction of IV with excess cyclopentene at ambient temperature yielded only a mono-hydrogermylation product, viz. the hydrodi-digermene, \( [Ar(H)Ge=Ge(Cp)Ar'] \) \( (Cp = cyclopentyl) \).14 This suggests that the dissociation of IV to the two-coordinate hydro-germene, \( Ge(H)Ar' \) in solution is minimal.

Recently, we have utilised extremely bulky amide ligands, developed in our group,15 to kinetically stabilise amido/hydrido-digermenes, e.g. \( L^1[H]Ge=Ge(H)L^1 \) 1 \( (L^1 = \text{N}(Ar')) \) \( (SiPri_3) \); \( Ar' = C_6H_5(C_6H_5)Ph_2Pr'-2,6,4) \), and isomeric dimeric hydrod bridged stannylene, e.g. \( L^2Sn(\mu-H)L^2 \) \( 2 \).17 These were shown to significantly dissociate to the uncoordinated two-coordinate hydro-tetrelenes; \( E(H)[L^1] \) \( (E = Ge 3 or Sn 4) \), in hydrocarbon solutions. Subsequently, two-coordinate hydrido-germynes, bearing even bulkier amide ligands, e.g. \( [Ge(H)-L^{10\text{Obu}}] \) [\( L^{10\text{Obu}} = \text{N}(Ar')\text{Si}(OBu')_3] \), were isolated in the solid state.18 Given that the six valence electron compounds 3 and 4 possess empty p-orbitals at their metal centres \( (cf. \) boranes) it was proposed that they would act as effective reagents for the hydrodymatation of unsaturated substrates. This was shown to be the case for aldehydes and ketones, and, indeed, 3 and 4 were also shown to be highly efficient catalysts for the hydroboration of the same substrates.19 Here, we show that these hydrod teenylenes regiospecifically hydrolymate a variety of unactivated alkene and alkyne substrates at ambient temperature, and with unprecedented facility. In some cases, these hydrolymattelation reactions are shown to be cleanly reversible under ambient conditions. Preliminary further reactivity studies of the formed amido/alkyl-tetrelenes are also reported.

Results and discussion

(i) Hydrometallation of alkenes and alkynes

Treatment of toluene solutions of compound 3 (as an equilibrium mixture with 1) with 1–1.5 equivalents of a range of unactivated terminal or cyclic alkenes, or 1 atm. of ethylene, led to almost instantaneous changes in the colour of the reaction solutions from orange to yellow at ambient temperature. This indicated that the hydrodymetallation reactions were complete in well under 1 minute. \( \text{H} \) NMR spectroscopic analyses of the reaction mixtures after ca. 10 min confirmed that the hydrodymetallation reactions were essentially quantitative after that time, affording the amido/alkyl germynes, 5–10 (Scheme 2). These could be isolated as yellow crystalline solids in moderate to excellent yields upon work-up. Several corresponding reactions involving the \( \text{in}(n) \) hydride 4 (as an equilibrium mixture with 2) were carried out at low temperature \( (–80 \degree \text{C}) \) due to the mild thermal instability of 4 at room temperature (solutions decompose over 2 days\( ^{19} \)). Upon warming the reaction mixtures to \( 20 \degree \text{C} \), \( \text{H} \) NMR spectroscopic analyses revealed that hydrostannylation of the substrates had cleanly occurred to give 11–13, which were isolated as crystalline solids in good yields (Scheme 2).

The facility of these uncatalysed olefin hydrogermylation and hydrostannylation reactions is unprecedented, and is likely a result of them preferentially involving the monomeric hydrod teenylenes, 3 and 4, over the dimeric species, 1 and 2 \( (cf. \) related carbonyl hydrodymetallations\( ^{19} \)). These coordinately unsaturated species possess an empty p-orbital, which likely lowers the energy barrier to olefin hydrodymetallation, by allowing an interaction between the substrate and the group 14 metal center prior to the hydrodymetallation reaction \( (cf. \) alkene hydroelementations by boranes\( ^{4} \) and the hydridosilylene 1\( ^{4} \)).

Scheme 1 Examples of previously reported group 14 element(n) hydride complexes \( (\text{Dip} = C_6H_5Pr'2-2,6; \text{Ar}' = C_6H_5(C_6H_5)Pr'2-2,6; 2-2,6) \).

Scheme 2 Synthesis of compounds 5–13.
The fact that all of the alkene hydrometallation products reported here are monomeric, provides further evidence for the more active species in these reactions being the hydrido-tetrylenes, 3 and 4 (cf. dimeric products from alkene hydrometallations involving dimeric IV and V). It is also of note that the hydrometallations of all of the terminal alkenes regiospecifically yielded the anti-Markovnikov product, as is typically the case with alkene hydroborations.

All of the alkene hydrometallation products 5–13 are thermally stable in the solid state. The solution state NMR spectroscopic data for the compounds are fully consistent with their proposed monomeric structures, and do not warrant further comment here. X-ray crystallographic studies were used to confirm the monomeric nature of all compounds, which represent the first structurally characterised examples of two-coordinate amido/alkyl germynes and stannylenes. Illustrative examples of the molecular structures of the compounds are depicted in Fig. 1 (see ESI† for the molecular structures of 8, 10, 11 and 13), while selected geometrical parameters are collected in Table 1. It should be noted that, although the crystal structures of 7 and 12 confirmed the molecular connectivity of the compounds, they are not of a quality suitable for publication, and their geometrical parameters will not be discussed here.26 The geometries of all the structurally characterised compounds are similar, in that all of their N–E–C (E = Ge or Sn) angles are suggestive of stereochemically active lone pairs at the metal centre. Moreover, in each, the trigonal planar Si–N–C unit is close to co-planar with the C–E–N fragment. This potentially allows for overlap of the N p-orbital lone pair with an empty p-orbital at the E-centre, which would help prevent the already sterically bulky compounds from dimerising to diteetlenes. There are no contacts between the Ge centres and any of the phenyl carbon atoms that would suggest significant Ge–aryl interactions in 5–10 (sum of van der Waals radii for Ge and C = 3.81 Å (ref. 21)). The closest Sn–Cphenyl contacts in 11–13 are considerably shorter, and are within the sum of van der Waals radii for Sn and C (3.87 Å (ref. 21)), which may indicate weak Sn–aryl interactions in those cases.

Considering the effectiveness of 3 and 4 as reagents for the hydrometallation of alkenes, it seemed likely that they would
pure samples of the compounds reproducibly exhibited signals due to the presence of small amounts of the germanium hydride equilibrium mixture, 1 and 3, and the free cycloalkene. In addition, recrystallisation of the compounds, and 8, repeatedly led to co-crystallisation with small amounts of 1. Moreover, C6D6 solutions of the tin cyclopentyl compound, 13, decomposed over several days at ambient temperature (to L1H, H2 and elemental tin), yet were stable for extended periods, even at 80 °C, in the presence of excess cyclopentene. All of these observations point to the hydrometallation products from the reactions of 3 and 4 with cycloalkenes being in equilibria with significant amounts of those reactants at room temperature (see Scheme 4 for the equilibrium between 3 and 9). The net decomposition of 13 can be explained by the mild instability of the tin hydride 4 at ambient temperature, which upon decomposition, inextricably leads to loss of 13 from the equilibrium mixture in that case.

The reversibility of the reaction that gave 9 was explored by a VT 1H NMR spectroscopic study of a sample of the compound which was prepared by reaction of CyMgBr (Cy = cyclohexyl) with L1GeCl. This ensured the absence of any 1 in the purified

Table 1  Selected bond lengths (Å) and angles (°) for 5, 6, 8–11, and 13 (E = Ge or Sn)

<table>
<thead>
<tr>
<th></th>
<th>5</th>
<th>6</th>
<th>8</th>
<th>9</th>
<th>10</th>
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<tr>
<td>E-N</td>
<td>1.874(2)</td>
<td>1.865(2)</td>
<td>1.885(3)</td>
<td>1.904(2)</td>
<td>1.898(2)</td>
<td>2.123(3)</td>
<td>2.127(4)</td>
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<tr>
<td>E-C</td>
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<td>1.991(2)</td>
<td>1.999(3)</td>
<td>2.031(4)</td>
<td>2.054(5)</td>
<td>2.182(4)</td>
<td>2.168(5)</td>
</tr>
<tr>
<td>N-E-C</td>
<td>105.4(1)</td>
<td>106.1(1)</td>
<td>110.6(1)</td>
<td>108.0(1)</td>
<td>109.8(3)</td>
<td>101.7(1)</td>
<td>102.8(2)</td>
</tr>
<tr>
<td>SiNEC (torsion)</td>
<td>5.4(1)</td>
<td>2.3(1)</td>
<td>10.7(1)</td>
<td>3.2(1)</td>
<td>1.3(1)</td>
<td>18.2(1)</td>
<td>13.5(3)</td>
</tr>
<tr>
<td>E·...Cpheno at</td>
<td>3.461(3)</td>
<td>3.664(2)</td>
<td>3.257(3)</td>
<td>3.307(3)</td>
<td>3.330(3)</td>
<td>3.330(3)</td>
<td>3.239(1)</td>
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The closest contact is given.

also be very reactive toward unactivated alkenes. To assess this, and to investigate if the metal hydrides could doubly hydrometallate alkenes, 1-phenyl-1-propyne was treated with two equivalents of either 3 or 4. Analysis of the reaction mixtures indicated that only one hydrometallation event occurred in both cases, and within several minutes at ambient temperature. These reactions afforded compounds 14 and 15 respectively, in close to quantitative NMR spectroscopic yields, and moderate isolated yields. The reactions proceeded with complete regio-specificity, giving the cis-isomer with the L1E fragment bonded to the phenyl substituted alkeneic carbon (Scheme 3). It is likely that double hydrometallations do not occur in these reactions due to the imposing steric bulk of the L1E fragments. For sake of comparison, the hydroboration of 1-phenyl-1-propyne with boranes typically gives mixtures of regioisomers, the composition of which is dependent upon the borane employed.22

Both 14 and 15 are thermally stable in the solid state and in solution. Their NMR spectra are in line with the proposed formulations of the compounds. The structures of the compounds were confirmed by X-ray crystallographic studies (see Fig. 2 for the molecular structure of 14) which reveal both to be monomeric in the solid state with the alkeneic phenyl and methyl substituents cis- to one another. Their C(45)-C(46) distances reflect localised double bonds, while the geometries about the metal centres are similar to those in 5–13. That is, the CNSi and NEC fragments are close to co-planar with one another, which allows for the possibility of N → E π-bonding in the compounds. In contrast, their C=C units are close to orthogonal to the NEC fragments, which discounts the possibility of any π-delocalisation over those fragments.

(ii) Reversible alkenyl hydrometallations, and alkenyl isomerisations

During characterisation of the cycloalkene hydrogermylation products 9 and 10, it was noticed that 1H NMR spectra (C6D6) of pure samples of the compounds reproducibly exhibited signals

![Scheme 3 Synthesis of compounds 14 and 15.](image)

![Fig. 2 Molecular structure of 14 (25% thermal ellipsoids are shown; hydrogen atoms, except the alkene proton, omitted). Selected bond lengths (Å) and angles (°) for 14: Ge(1)–N(1) 1.881(6), Ge(1)–C(45) 1.991(1), C(45)–C(46) 1.35(1), closest Ge(1)···Cphenyl 3.27(1), N(1)–Ge(1)–C(45) 110.3(3). Selected bond lengths (Å) and angles (°) for 15: Sn(1)–N(1) 2.122(5), Sn(1)–C(45) 2.229(8), C(45)–C(46) 1.32(1), closest Sn(1)···Cphenyl 3.26(1), N(1)–Sn(1)–C(45) 107.1(2).](image)
As far as we are aware, the van der Waal’s radii for E and H (E atom (range: 2.62 Å–3.29 Å)) are significantly less than the sum of the van der Waal’s radii for E and H (E = Ge 3.21 Å, Sn 3.28 Å [ref. 21]). As far as we are aware, β-hydride elimination process involving germanium or tin alkyls, which are facile and reversible at room temperature, are unknown.24 With that said, there is one report of the hydrogermylation of a phosphaalkyne, P=CBu’, by a three-coordinate germanium(II) hydride, [(MesNacnac)GeH] [(MesNacnac = [(MesNCMe)2]CH)–, Mes = mesityl], which reversibly affords the phosphaalkenyl complex, [(MesNacnac) Ge(C(Bu’)=PH)].24 The reversibility of this reaction at ambient temperature was proposed to involve migration of the phosphorus-bound β-hydride to the germanium centre in [(MesNacnac)GeC(Bu’)=PH].

To explore the possibility of facile β-hydrogen elimination processes being the origin of the reversibility of the reactions that gave 8–10, DFT calculations were carried out at several levels of theory on the hydrogermylation reaction that gave 9 (see ESI† for full details). These indicated that the reaction is exergonic by an amount [ΔG = −42.3 kJ mol⁻¹ at M06-2X+D3/def2-TZVPP/TPSS+D3/def2-TZVPP] that is small, and not dissimilar to that found from the experimental van’t Hoff analysis of the reaction. Importantly, the reverse reaction was, indeed, found to proceed via a β-hydride elimination process, involving a transition state with a four-membered GeC₂H₆ ring (see Scheme 4 and ESI†). In combination with the small free energy of the reaction, the relatively low free energy of activation for this transition state (Δ‡G = 76.6 kJ mol⁻¹), is fully consistent with the experimentally observed equilibrium for the reaction that gave 9.

Further evidence that the reversibility of the cyclic alkene hydrogermylation reactions proceed via β-hydride elimination processes, comes from the reactions of 3 with 1,5-cyclooctadiene (1,5-COD) and 2-methyl-2-butene (Scheme 5). In both cases the expected hydrometallation products were not observed, and instead products, 16 and 17, that apparently result from the hydrogermylation of isomerised alkenes, were isolated in moderate yields. It is worth mentioning that treatment of 1,5-cyclooctadiene with two equivalents of 3 did not lead to a double hydrogermylation product, and no reaction occurred between 3 and the tetra-substituted alkenes, R₄C=C₆R₄ (R = Me or Ph). These observations presumably result from the considerable steric bulk of the monomeric hydrido-germylene, 3.
It is possible that the formation of 16 proceeds via the expected hydrogermylation product, 18, as an intermediate. This then undergoes a β-hydrogen elimination to give 3 and 1,4-cyclooctadiene (1,4-COD). Hydrogermylation of this by 3, and another β-hydrogen elimination event, yields 3 and 1,3-cyclooctadiene (1,3-COD), the latter of which is then hydrogermylated to give the observed product, 16. Similarly, hydrogermylation of 2-methyl-2-buten affords the initially expected product, 19, which β-hydride eliminates to give 3 and 3-methyl-1-butene. Hydrogenation of this olefin then leads to the observed product, 17. The facility of these reactions highlights the potential that 3, and related reagents, have for the selective stoichiometric isomerisation of alkenes. While such isomerisations are common for transition metal systems,3 they are rare for main group compounds.

Both 16 and 17 were crystallographically characterised, and their molecular structures are depicted in Fig. 4. These show them to be monomeric with geometries that are comparable to those of the other amido/alkyl germylenes, 5–10, reported here. In the case of 16, the presence and location of the residual double bond of its cyclooctenyl moiety is confirmed by the shortness of the C(51)–C(52) linkage (1.372(6) Å).

(iii) Reactivity studies of amido/alkyl germylenes and stannylene

Preliminary further reactivity studies were carried out on examples of the amido/alkyl germylene and stannylene complexes prepared here, with a view to utilising these compounds in catalytic synthetic protocols. Initially, the oxidative addition (OA) of H₂ and hydridic reagents (e.g. HBpin, HBcat, PhSiH₃, Et₃SiH, (EtO)₃SiH and DIBAL) to 5–13 was explored, but in no case was a reaction observed under ambient conditions. Attention then turned to the reaction of stoichiometric amounts of protonic reagents with 5, 8 and 11. In the case of the germylenes, the oxidative addition of HCl, NH₃ or EtOH to the Ge centres of 5 or 8 occurred, to give a few crystals of 20 (amongst several other unidentified products), and good isolated yields of 21 and 22, respectively (Scheme 6). Solutions of 21 and 22 are resistant to reductive elimination (RE) of ethane or cyclopentane, even when heated to 50 °C for one hour. It is noteworthy that the reactions that gave 20–22 are comparable to related oxidative additions of HF,26 NH₃27 and EtOH28 to germylenes, that have appeared in the literature.

Different outcomes resulted from the reactions of the stannylene 11 with protic reagents. These were not clean, and typically generated product mixtures that contained significant amounts of the secondary amine, L₁H. The two exceptions here were the reactions with stoichiometric amounts of the bulkier reagents TEMPOH (TEMP = 2,2,6,6-tetramethylpiperidinyl) and Bu’OH. These afforded moderate to good isolated yields of the piperidinyl N-oxide product, 23, and the known tin tert-butoxide, 24,29 respectively (Scheme 6). When these reactions were followed by ³H NMR spectroscopy, the generation of significant amounts of ethane (and smaller quantities of L₁H) was
observed. It cannot be sure if these reactions proceed via initial oxidative additions of the O–H bond of the reagents to the SnII center of 11, prior to reductive elimination of ethane, but given the formation of the stable germanium(IV) ethoxide, 22, this is certainly a possibility \(\text{cf.} \) related “OA/RE” reactions of H₂ and NH₁ with Ar’₂Sn: (ref. 27).

In attempts to form a rare example of a three-coordinate germanone \((\text{R}_2\text{Ge}=\text{O})\)\(^{39}\) and the first example of a corresponding stannanone \((\text{R}_2\text{Sn}=\text{O})\), toluene solutions of the germylene, 9, and stannylene, 12, were reacted with excess O₂. Instead of yielding monomeric products, the dimeric oxo-bridged species, 25 and 26, were obtained in moderate isolated yields (Scheme 6). It is apparent that the steric shielding of the metal centres in the compounds is not sufficient to prevent dimerisation of the target heavier ketone products. In this respect, there is ample literature precedent for the oxidation of germynes and stannylenes to give dimeric systems comparable to 25 and 26.\(^{30,31}\)

No spectroscopic data could be obtained for the HCl oxidative addition product, 20, due to the very low yield of the compound. The NMR spectroscopic data for the other products, 21–23, are as would be expected, though it is worthy of mention that the chiral Ge centres in 21 and 22 give rise to two multiplet resonances for the diastereotropic CH₂ units of both the ethyl and ethoxide ligands. The X-ray crystal structures of 20, 22 and 23 were determined and their molecular structures are shown in

![Molecular structures](image-url)
It is worthy of mention that there is only one other structurally characterised tin piperidinyl N-oxide complex reported in the literature, \([\text{CH}_2\text{C(SiMe}_3)_2\text{]}\text{Sn(OTEMP)}_2\)\textsuperscript{12} though the metal centre in this is in the +4 oxidation state. The dimeric nature of \(25\) and \(26\) was confirmed by X-ray crystallographic studies (see Fig. 6 for the molecular structure of \(26\)), which also showed their oxide ligands to essentially symmetrically bridge two distorted tetrahedral metal centres.

**Conclusion**

In summary, reactions of solution stable two-coordinate hydrido-tetrylenes with a variety of unactivated cyclic and acyclic alkenes, and one internal alkyne, lead to the unprecedentedly rapid and regioselective hydrometallation of the unsaturated substrate at ambient temperature. The products of the alkene hydrometallations represent the first structurally characterised examples of two-coordinate amido/alkyl germynes and stannylenes. In the cases of the cycloalkene hydrometallations, the reactions were shown to be cleanly reversible under ambient conditions. The results of computational and experimental van’t Hoff analyses of one such reaction, strongly suggest that its reversal proceeds via \(\beta\)-hydride elimination from the cycloalkyl ligand, regenerating the cycloalkene and hydrido-tetrylene starting materials. Further evidence for this proposal comes from the reactions of a hydrido-germylene with 1,5-cyclooctadiene and 2-methyl-2-buten, both of which seemingly proceed via intermediate \(\beta\)-hydride elimination processes, and the clean isomerisation of the alkene involved, prior to its ultimate hydrogermylation. In addition, the element-hydrogen bonds of several protic compounds have been shown to oxidatively add to the germanium(\(\text{II}\)) centre of two of the amido/alkyl germynes prepared in this study, while similar reactions with an equivalent stannylene proceed via alkane elimination, and generation of tin(\(\text{II}\)) products. Oxidations of two amido/alkyl tetrylenes with \(\text{O}_2\) have been shown to give four-coordinate, oxo-bridged metal(\(\text{IV}\)) dimers. We continue to explore the stabilisation and synthetic utility of low oxidation state group 14 element compounds.

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**References**


**Edge Article**


20 N.B. A few crystals of the compound, [L’Ge(μ-C5H5)C4]=, were isolated from the reaction of 3 with ethylene. It is believed these arose from the double hydrogermylation of acetylene, which was an impurity in the ethylene gas used for the experiment. Details of the X-ray crystal structure of [L’Ge(μ-C5H5)C4]= can be found in the ESI†


23 N.B. We have recently shown that β-hydride elimination from the dinuclear compound, [L’Ge(μ-C5H5)C4]=, occurs at 70 °C to give 3 and benzene. This reaction is, however, not reversible. See T. J. Hadlington, J. Li, M. Hermann, A. Davey, G. Frenking and C. Jones, Organometallics, 2015, 34, 3175.


27 Y. Ping, J.-D. Guo, B. D. Ellis, Z. Zhu, J. C. Fettinger, S. Nagase and P. P. Power, J. Am. Chem. Soc., 2009, 131, 16272. N.B. It is well known that tin(II) species are more reluctant than germanium(II) species to undergo oxidative addition processes, due to the inert pair effect.


31 N.B. A few crystals of [[L’Ge(μ-C5H5)C4]=O}] were isolated from the mixture that gave compound 6. This compound presumably arose from the presence of a small amount of adventitious O2 in the reaction mixture. Details of the X-ray crystal structure of [[L’Ge(μ-C5H5)C4]=O}] can be found in the ESI†