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Detection of σ -alkane complexes of manganese by NMR and IR spectroscopy in solution: $(\eta^5\text{-C}_5\text{H}_5)\text{Mn}(\text{CO})_2(\text{ethane})$ and $(\eta^5\text{-C}_5\text{H}_5)\text{-Mn}(\text{CO})_2(\text{isopentane})$

Olga Torres,^a James A. Calladine,^b Simon B. Duckett,^a Michael W. George^{*b} and Robin N. Perutz^{*a}

Irradiation of $\text{CpMn}(\text{CO})_3$ in liquid ethane at 135 K at 355 nm yields a photoproduct that exhibits $\nu(\text{CO})$ bands in the IR spectrum shifted to low wavenumber with respect to $\text{CpMn}(\text{CO})_3$ that are indicative of a Mn(I) dicarbonyl. Parallel experiments employing *in situ* irradiation within an NMR probe (133 K, 355 nm photolysis) reveal the ^1H NMR signals of this product and confirm its formulation as the σ -ethane complex $\text{CpMn}(\text{CO})_2(\eta^2\text{-C1-H-ethane})$. The resonance of its coordinated C–H group is observed at δ –5.84 and decays with lifetime of *ca.* 360 s. Analogous photolysis experiments in isopentane solution with IR detection produce $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-isopentane})$ with similar IR bands to those of $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-ethane})$. ^1H NMR spectra of the same species were obtained by irradiation of $\text{CpMn}(\text{CO})_3$ in a 60 : 40 mixture of propane and isopentane; three isomers of $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-isopentane})$ were detected with coordination of manganese at the two inequivalent methyl positions and at the methylene group, respectively. The lifetimes of these isomers are *ca.* 380 ± 20 s at 135 K and do not vary significantly from each other. These σ -complexes of manganese are far more reactive than those of related $\text{CpRe}(\text{CO})_2(\text{alkane})$ complexes which are stable in solution at 170–180 K. The room temperature lifetimes of $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-ethane})$ and $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-isopentane})$, as determined by TRIR spectroscopy, are 2.0 ± 0.1 and 28 ± 1 μs , respectively.

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Introduction

Alkane complexes¹ have now been demonstrated both experimentally and theoretically to act as reaction intermediates in many fundamental transformations associated with transition metal centres: they include oxidative addition, and its reverse, reductive elimination,^{2,3} protonation,⁴ σ -complex assisted metathesis (σ -CAM) which can lead to alkane functionalisation,⁵ hydrogenation/dehydrogenation⁶ and substitution reactions.⁷ Such alkane complexes were first detected by UV/vis and IR spectroscopy in low temperature matrices and glasses in the 1970's.^{8,9} For many years, time-resolved spectroscopy, especially with IR detection, has provided strong evidence for the existence of σ -alkane complexes, *via* monitoring of the $\nu(\text{CO})$ vibrations of metal carbonyls.^{10–13} This technique also revealed the variation in lifetime of the alkane complexes with the metal and spectator ligands. A breakthrough that enables the NMR detection of these complexes came through the use of UV irradiation inside the NMR probe to study complexes that had first

been identified by IR spectroscopy as relatively long-lived.^{14–21} Low temperature protonation has added a thermal route to the formation of methane and, very recently, ethane complexes detected by NMR spectroscopy.^{4,22} These NMR studies have distinguished the different regioisomers of alkane complexes and revealed their diagnostic ^1H and ^{13}C chemical shifts and $^2J_{\text{CH}}$ coupling constants. The information from NMR spectroscopy has confirmed that these alkane complexes belong to the growing range of σ -bond complexes^{5,23,24} comprising dihydrogen,²⁵ borane,^{26,27} silane^{28–31} and germane³² ligands as well as alkanes.

The first examples of crystallographically characterised alkane complexes were discovered serendipitously in a coordination environment typical of a solvent of crystallisation.^{33,34} More recently, a rhodium(norbornane) complex containing two chelating $\text{Rh}(\eta^2\text{-C,H})$ linkages has been made by hydrogenation of a norbornadiene complex in the solid state and characterised by single crystal X-ray crystallography (Rh-C1 2.494(10) Å; Rh-C2 , 2.480(11) Å).⁶ The structures of some iron-based metal-organic frameworks that adsorb alkanes have been determined by powder neutron diffraction and yield far longer Fe–C distances of *ca.* 3 Å, suggesting weaker interactions.³⁵

^aDepartment of Chemistry, University of York, York YO10 5DD, UK^bSchool of Chemistry, University of Nottingham, Nottingham NG7 2RD, UK. E-mail: robin.perutz@york.ac.uk; Mike.George@nottingham.ac.uk

Some of the special properties of alkane complexes can be deduced from isotope exchange experiments on metal alkyl deuterides that result in metal (deuteroalkyl) hydride complexes.^{1,2,36} Of particular importance, is the ability of the metal to migrate from one CH₂ or CH₃ position to another within a σ -alkane complex, a phenomenon known as chain walking.³⁷ These dynamic processes are usually investigated by product studies with isotope labelling. Recently, chain walking has been followed by dynamic ¹³C NMR spectroscopy for an ethane complex.²² On the other hand, geminal exchange of the coordination position of an η^2 -alkane complex is too fast to resolve by NMR spectroscopy.^{15,16}

Alkane complexes represent an important route to alkane functionalisation as demonstrated for borylation,^{38,39} dehydrogenation⁴⁰ and alkane metathesis.^{41,42} They are also likely to participate in the platinum chemistry of alkanes pioneered by Shilov.^{43,44} Simple alkanes are often very reactive towards metal ions in the gas phase, as shown by mass spectrometric methods.⁴⁵ Their reactivity may be understood *via* alkane complexation and, in some cases, the σ -CAM mechanism.^{5,45} Alkane complexes are also postulated in Periana's most recent results on C–H functionalisation catalysed by heavy p-block metals.⁴⁶ However, there are recent examples of methane functionalisation that involve very different chemical systems for which other mechanisms may apply.^{47–50}

The study of alkane complexes and alkane functionalisation reactions by computational methods is an important adjunct to experiment providing geometric information and details of the potential energy surfaces.^{20,51–53} The highest level calculations (composite CCSD(T)/def2-QZVPP) on CpRe(CO)₂(η^2 -CH₄) give Re–C and Re–H distances of 2.60 and 1.92 Å, respectively, and a Re–CH₄ complexation enthalpy of –62.0 kJ mol^{–1} at 298 K.⁵⁴

We have demonstrated previously that we can identify suitable target alkane complexes for NMR spectroscopic investigation by first measuring their lifetimes at low temperature by TRIR or FTIR spectroscopy.^{18,19} We then investigate the same complexes by NMR spectroscopy using the same laser wavelength and the same temperature if the lifetime is estimated to be sufficient for NMR monitoring. Using this technique, we were able to measure the NMR spectra of the isomeric propane complexes, CpMn(CO)₂(η^2 -C1-H-propane) and CpMn(CO)₂(η^2 -C2-H-propane), the first example of the NMR characterisation of a first row transition metal alkane complex.²¹ The spectra of the corresponding butane complexes were also recorded.²¹ Recognising that the lifetime of the manganese propane complexes is only just long enough for NMR characterisation at 134 K, we selected three other alkanes suitable for investigation at such low temperatures, ethane, isopentane and methane and report here their reactivity towards photogenerated CpMn(CO)₂ based on evidence from IR and NMR spectroscopy.

Experimental

Materials

For IR spectroscopy, CpMn(CO)₃ (Aldrich), ethane and CO (BOC, CP grade) were used as received. Isopentane (Aldrich, 99%) was dried by refluxing over CaH₂ for 24 h under an argon

atmosphere. In all cases, the concentration of the CpMn(CO)₃ was in the range 10^{–3} to 10^{–4} M, and the IR and UV absorbances at the probe and pump wavelengths were kept in the range 0.2–0.8. For NMR spectroscopy, chemicals were sourced as follows: ethane (Air Liquide, grade N2.5, 99.5%), isopentane (Aldrich, 99%), propane (BOC, instrument grade N2.5, 99.5%), propane-d₂ (98% atom D, Isotec TM), methane (BOC research grade), CpMn(CO)₃ (Strem). Propane and methane were used as received, but ethane was purified by passing through a U-tube containing Pt (5%)/charcoal catalyst (Johnson Matthey) on a high vacuum line at 195 K to trap ethylene, hydrogen and other reactive impurities.

Room temperature TRIR spectroscopy

A flow system containing the isopentane solution was prepared on a Schlenk line and pressurized to 2 atm with CO. An infrared solution cell (Harrick Scientific Corp.) with CaF₂ windows and 0.5 mm pathlength was used to record FTIR and TRIR spectral data. Fresh sample was flowed into the cell after each UV laser shot from a reservoir vessel into a receiving vessel, held at a slightly lower pressure. The activation energy measurements were performed in a variable temperature IR cell and a recirculating water bath was used to control the temperature. The solutions were allowed to come to thermal equilibrium before any measurements were made, and the internal IR cell temperature was monitored using a 1/16" K-type thermocouple.

FTIR spectroscopy in low temperature solutions

A custom built, high-pressure, low-temperature copper cell (CaF₂ windows, 2.5 mm pathlength), cooled by a CryoTiger cold-finger (Brooks Automation) was used for all low temperature FTIR and time-resolved measurements. Details of this apparatus are given elsewhere.⁵⁵ For the experiments in liquid ethane, a small quantity of CpMn(CO)₃ was placed in the cell and pressurized to 45 atm (no CO was added) using a hand-pump. For the isopentane experiment, a solution of CpMn(CO)₃ pre-made using the method described above under 2 atm Ar was flowed into the cell (held at a slightly lower pressure). In both cases, the apparatus was cooled to 135 K and allowed to thermally equilibrate for 15 min before any measurements were taken.

Time-resolved infrared spectroscopy

Microsecond TRIR spectra and kinetics were recorded on the Nottingham point-by-point quantum cascade laser (QCL) spectrometer.⁵⁶ A Nd : YAG laser (Spectra Physics GCR-12; 355 nm) initiates the photochemical reaction and the laser-induced change in IR transmission through the sample, as a function of probe wavelength, is measured using an MCT detector cooled by liquid nitrogen and coupled to an oscilloscope and PC.

Rapid-scan FTIR spectroscopy was performed using an Agilent Cary 680 spectrometer. The cryostat was mounted on an external bench and the UV laser (SpectraPhysics GCR-12; 355 nm) aligned through the sample cell.⁵⁵



Low-temperature NMR spectroscopy with *in situ* laser irradiation

NMR spectra were recorded on a Bruker Avance wide-bore 600 MHz spectrometer with solvent suppression achieved with the excitation sculpting pulse sequence with gradients *zgesgp* 1D sequence (Avance-version, Feb 2012), typically using about 24 scans.⁵⁷ Some additional spectra were obtained with WATERGATE W5 using the pulse program *zgpgpw5*.⁵⁸ ¹H NMR spectra were calibrated by adding a trace of benzene, taken as resonating at δ 7.15. Laser photolysis was carried out with a pulsed Nd : YAG laser (Continuum Surelite II) operating at 355 nm with 10 Hz repetition rate and a laser power of 85 mW. The laser radiation is directed through a hole in the probe housing onto the NMR tube. The sample is contained in a high-pressure sapphire NMR tube (IDEAS!, UVA, B.V.). The samples contained 1–2 mg of CpMn(CO)₃ and approximately 0.4 mL of solvent that was transferred on a high pressure line (ethane and propane) or a high vacuum line (isopentane). A very dilute sample of [Ru(CO)₂(Ph₂PCH₂CH₂PPh₂)(PPh₃)] in C₆D₆ was used for real-time laser alignment in conjunction with *para*-hydrogen amplification.⁵⁹ For kinetics, the integration of the relevant resonance was measured relative to that of benzene added in trace quantities. This apparatus is described in more detail elsewhere.¹⁹

Results and discussion

Infrared spectroscopic studies

The photochemistry of CpMn(CO)₃ in several different alkane solvents including ethane has been investigated previously.^{13,21} A solution of CpMn(CO)₃ in isopentane (2-methylbutane) at room temperature was investigated by TRIR spectroscopy with 355 nm photolysis. The transient difference spectrum measured 1.2 μ s after the laser pulse (Fig. 1a, Table 1) reveals that bleaching of the parent ν (CO) bands occurs together with the formation of two

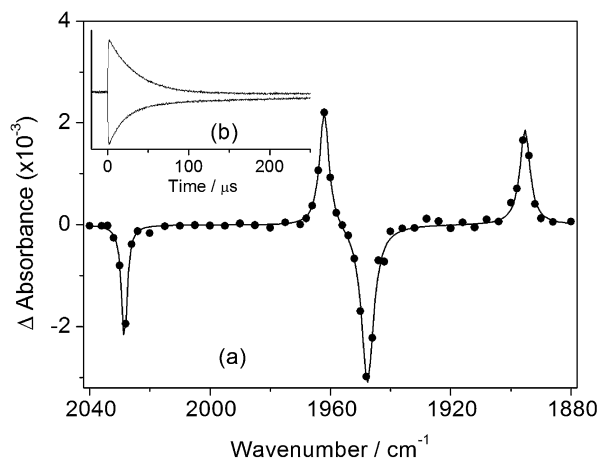


Fig. 1 (a) Point-by-point TRIR difference spectrum of CpMn(CO)₃ in isopentane in the presence of 2 atm CO, 1.2 μ s after 355 nm photolysis. The solid line is a multi-Lorentzian fit of the spectrum. (b) Kinetic traces showing decay of the Mn-isopentane complex and reformation of the parent.

Table 1 IR ν (CO) band positions, lifetimes and activation energies for Mn-alkane complexes

Complex	Solvent, temp/K	ν (CO)/cm ⁻¹	Lifetime	E_a (kJ mol ⁻¹)
CpMn(CO) ₃	Liq. ethane, 298 ^a	2032, 1952		
	Liq. ethane, 135	2029, 1946		
	Isopentane, 298	2030, 1948		
	Isopentane, 135	2028, 1944		
	scCH ₄ , 298 ^a	2035, 1955		
CpMn(CO) ₂ - (alkane)	Liq. ethane, 298 ^a	1967, 1902	2.0 ± 0.1 μ s	38 ± 2
	Liq. ethane, 135	1962, 1895	126 ± 6 s	
	Isopentane, 298	1962, 1896	28 ± 1 μ s	43 ± 2
	Isopentane, 135	1958, 1890	250 ± 3 s	
	scCH ₄ , 298 ^a	1972, 1908	400 ± 32 ns	41 ± 2

^a Ref. 13.

new bands that are shifted 68 cm⁻¹ (ν (CO)_{symm}) and 53 cm⁻¹ (ν (CO)_{antisymm}) to lower wavenumber relative to the a₁ and e modes of CpMn(CO)₃, respectively. The shifts are consistent with a reduction in the number of CO groups with no change in oxidation state. They are assigned to the a' and a'' modes of the σ -alkane complex, CpMn(CO)₂(isopentane) by comparison to the spectra of related alkane complexes formed by [CpMn(CO)₂] and [CpRe(CO)₂].^{13,19,21,60} In the presence of CO (2 atm) the Mn-isopentane complex reacts quickly and cleanly to regenerate the parent with a lifetime of 28 ± 1 μ s (Fig. 1b). The rate of reaction of the σ -alkane complex with CO was measured as a function of temperature between 282 and 295 K and an activation energy (E_a) of 43 ± 2 kJ mol⁻¹ obtained (Fig. 2). This activation energy can be used as a minimum value for the Mn-alkane bond strength, and is similar to those measured for analogous CpMn(CO)₂(alkane) complexes (*viz.* Mn-heptane,⁶⁰ 39 kJ mol⁻¹; Mn-methane,¹³ 41 kJ mol⁻¹).

The reactivity of CpMn(CO)₃ in isopentane was also investigated at low temperature by rapid-scan FTIR spectroscopy. A

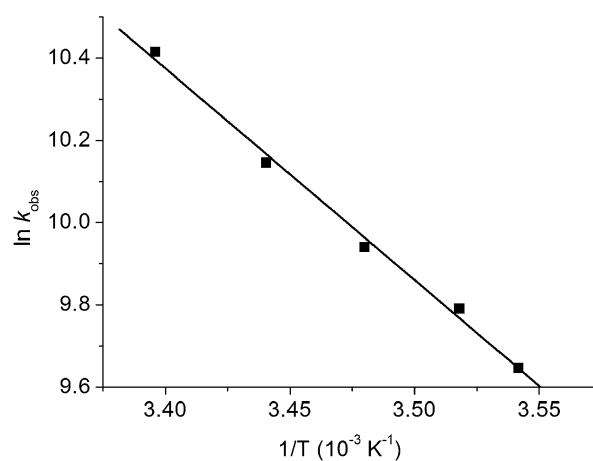


Fig. 2 Arrhenius plot for the reaction of CpMn(CO)₂(isopentane) with CO measured from the decay of the σ -alkane complex ν (CO) band at 1896 cm⁻¹.



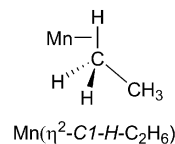
temperature of 135 K was selected in order to be consistent with the NMR experiments. Again, photolysis (355 nm) of an isopentane solution of $\text{CpMn}(\text{CO})_3$ leads to the formation of the σ -alkane complex, $\text{CpMn}(\text{CO})_2(\text{isopentane})$ with bands at 1958 and 1890 cm^{-1} (Fig. 3a). At 135 K, once irradiation is stopped, these bands are observed to decay with a lifetime of 250 ± 3 s (Fig. 3b) in the absence of added CO.

The room temperature spectra and kinetics of $\text{CpMn}(\text{CO})_2$ -ethane have been reported previously.¹³ Analogous results to those for isopentane are obtained upon irradiation of a solution of $\text{CpMn}(\text{CO})_3$ in liquid ethane (650 psi) at 135 K. The rapid-scan FTIR difference spectrum measured 2 s after irradiation (Fig. 3c) illustrates $\nu(\text{CO})$ bands due to the formation of $\text{CpMn}(\text{CO})_2(\text{ethane})$ at very similar wavenumbers of 1962 and 1895 cm^{-1} . The Mn-ethane complex decays with a lifetime of 125 ± 6 s (Fig. 3d), about half the value of the isopentane analogue.

Studies with NMR spectroscopy

The lifetimes of the isopentane and ethane complexes measured by low temperature IR spectroscopy suggest that the NMR detection of these photoproducts is possible. A ^1H NMR spectrum of a sample of $\text{CpMn}(\text{CO})_3$ dissolved in unpurified ethane at 133 K was first measured with solvent suppression at δ 0.65. In addition to the expected Cp resonance at δ 4.58, a considerable amount of ethylene was detected at δ 5.33 together with some hydrogen resonating at δ 4.55. The ethane was therefore purified by passing through a Pt/charcoal trap which removed the H_2 impurity signal and greatly reduced the ethylene contamination. Integration indicated that the ethylene/ $\text{CpMn}(\text{CO})_3$ ratio was *ca.* 1 : 0.7, although the ethane solvent is still in vast excess. Photolysis of $\text{CpMn}(\text{CO})_3$ (10–20 mmol dm^{-3}) within the NMR probe for *ca.* 1–2 min using the pulsed laser operating at 355 nm in neat ethane at 133–135 K results in a decrease in intensity of the corresponding Cp resonance at δ 4.58 (s) and the appearance of a new resonance at

δ 4.19 (s) together with another resonance in the high-field region at δ –5.84 (Fig. 4). The Cp resonance of the product is a clean singlet without any of the broadening or shoulders observed with propane and butane, reflecting the formation of a single isomer.²¹ These two new resonances can be attributed to the Cp and C^1H_3 resonances, respectively, of the product $\text{CpMn}(\text{CO})_2(\eta^2\text{-C1-H-ethane})$ by comparison with previous studies of the photolysis of $\text{CpMn}(\text{CO})_3$ in neat propane.



The high field NMR resonance can be understood as arising from the average of the bound $\eta^2\text{-C-H}$ proton and the two geminal protons since there will be a very low barrier for exchange between them.¹⁶ This resonance is poorly resolved but may be interpreted as a binomial quartet with $J = 7$ Hz (Fig. 4c) arising from resolved coupling to the nuclei of the second CH_3 group. The C^2H_3 resonance is not detected because of overlap with the solvent resonance. We estimated the lifetime of $\text{CpMn}(\text{CO})_2(\eta^2\text{-C1-H-ethane})$ as *ca.* 360 s at 135 K which corresponds to a value of ΔG^\ddagger of 39 kJ mol^{-1} . The ethylene impurities in the ethane probably reduce the lifetime of the photoproduct. This estimate compares with 126 ± 6 s determined by TRIR spectroscopy above with unpurified ethane.

We then turned our attention to isopentane. The first attempt to characterise the products of the photolysis at 355 nm

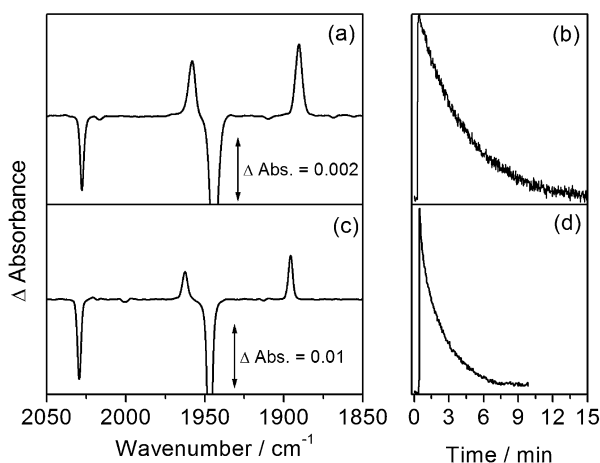


Fig. 3 Rapid-scan FTIR difference spectra obtained 2 s after 355 nm photolysis of $\text{CpMn}(\text{CO})_3$ at 135 K in (a) isopentane and (c) liquid ethane. Kinetic traces showing decay of (b) the Mn-isopentane complex at 1958 cm^{-1} and (d) the Mn-ethane complex at 1962 cm^{-1} .

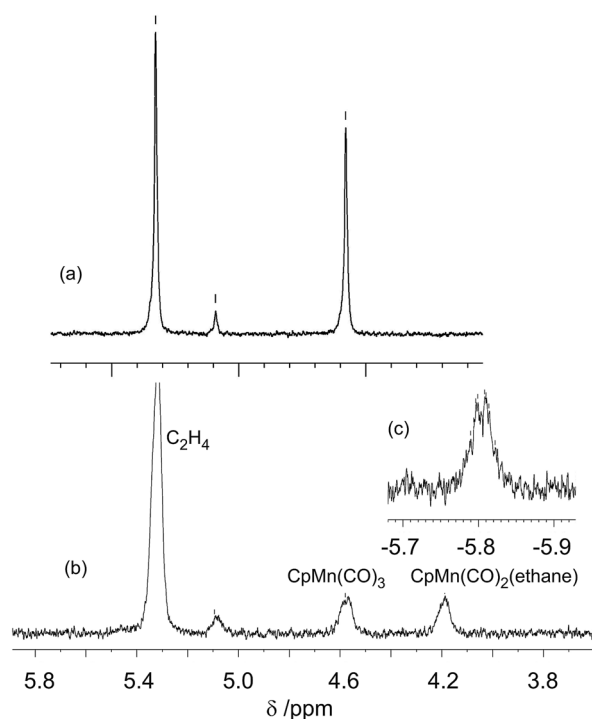
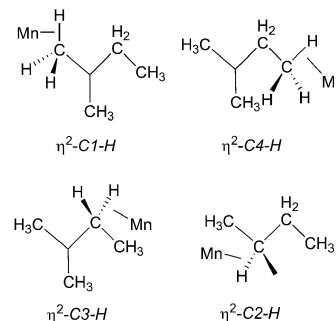


Fig. 4 (a) ^1H NMR spectrum of $\text{CpMn}(\text{CO})_3$ in ethane prior to irradiation in cyclopentadienyl region (b) spectrum after 355 nm laser irradiation. (c) Inset showing high field region after irradiation.



of $\text{CpMn}(\text{CO})_3$ in neat isopentane at 142 K was unsuccessful due to the poor solubility of the starting material in this solvent. In order to increase its solubility, we employed a propane/isopentane (60 : 40) mixture. We chose propane because of its advantageous spectroscopic properties and low freezing point, coupled with the fact that we have already reported the ^1H NMR spectrum of the corresponding propane complexes. The ^1H NMR spectrum measured before photolysis shows the expected Cp resonance at δ 4.58. After photolysis (200 s) of $\text{CpMn}(\text{CO})_3$ in a mixture of liquid propane and isopentane at 133 K a new Cp resonance is observed at δ 4.20 which exhibits a low field shoulder. Three new high field resonances associated with the Mn–CH moiety are observed in addition to the two resonances of $\text{CpMn}(\text{CO})_2(\eta^2\text{-C1-H-propane})$ and $\text{CpMn}(\text{CO})_2(\eta^2\text{-C2-H-propane})$. These three resonances appear at δ –5.77, –5.94 and –9.35 with integration ratios of 3 : 6 : 1 (Fig. 5b). They are consistent with the observation of three isomers of the photoproduct, namely $\text{CpMn}(\text{CO})_2(\eta^2\text{-C4-H-isopentane})$, $\text{CpMn}(\text{CO})_2(\eta^2\text{-C1-H-isopentane})$ and $\text{CpMn}(\text{CO})_2(\eta^2\text{-C3-H-isopentane})$, respectively (Scheme 1). An experiment using 2,2- D_2 -propane in place of H_8 -propane eliminated the resonance at δ –8.2 due to $\text{CpMn}(\text{CO})_2(\eta^2\text{-C2-H-propane})$. The kinetics of decay of the resonances at δ –5.77 and –5.94 were measured as 372 ± 22 s and 395 ± 23 s, respectively (Fig. 6), and compare to the lifetime of the FTIR signal of 250 ± 3 s. Note that these integrations should be considered as indicative because of the use of the solvent suppression sequence. For comparison, Ball has observed all three isomers of $(\eta^5\text{-C}_5\text{H}_4^1\text{Pr})\text{Re}(\text{CO})_2(\eta^2\text{-C-H-pentane})$ and two isomers of $\text{CpRe}(\text{CO})_2(\eta^2\text{-C-H-cyclohexane})$, whereas with $(\eta^6\text{-C}_6\text{Et}_6)\text{W}(\text{CO})_2(\eta^2\text{-C-H-pentane})$ and $(\eta^6\text{-C}_6\text{Et}_6)$



Scheme 1 Isomers of $\text{CpMn}(\text{CO})_2(\text{isopentane})$. The C2 isomer is not observed.

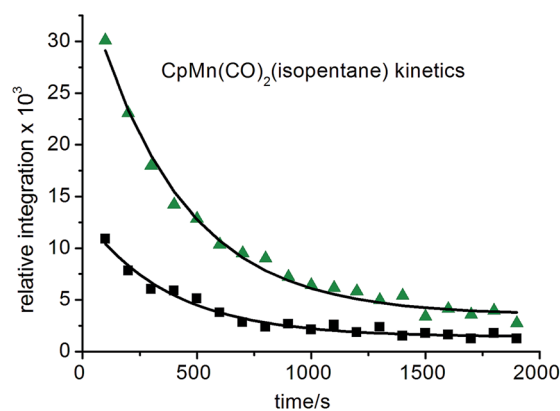


Fig. 6 Time profile of the high field resonance integration of $\text{CpMn}(\text{CO})_2(\text{isopentane})$ relative to the internal benzene standard. \blacktriangle δ –5.94 signal with the fit showing the exponential decay of lifetime 395 ± 23 s; \blacksquare corresponding data for δ –5.77 signal and fit to a lifetime of 372 ± 22 s.

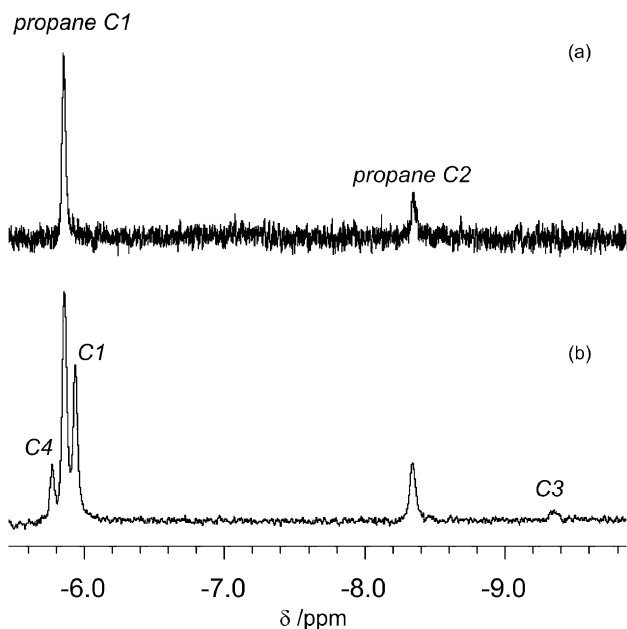


Fig. 5 ^1H NMR spectrum in the high-field region obtained after 355 nm laser irradiation of (a) $\text{CpMn}(\text{CO})_3$ in liquid propane; (b) a 60 : 40 mixture of liquid propane and isopentane at 133 K. The labels indicate the positions of Mn coordination for propane (above) and isopentane (below).

$\text{W}(\text{CO})_2(\eta^2\text{-C-H-2,2-dimethylbutane})$ only the isomers associated with interactions with primary C–H bonds bound were evident.^{15,20}

We carried out similar photochemical experiments with $\text{CpMn}(\text{CO})_3$ dissolved in a 90 : 10 mixture of propane and methane at 133 K. Under these conditions, we were only able to detect the corresponding resonances of the propane complexes mentioned above. Given the signal-to-noise ratio obtained in these data we can deduce that the corresponding methane complex is less stable than the propane adduct.

Conclusions

A combination of IR spectroscopy and NMR spectroscopy at 135 K has been used to detect highly unstable ethane and isopentane complexes of the type $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-alkane})$. As expected, $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-ethane})$ exists as a single isomer whereas $\text{CpMn}(\text{CO})_2(\eta^2\text{-C-H-isopentane})$ is detected as a mixture of three isomers by NMR spectroscopy. These species cannot be differentiated through their IR data, indicating that the electronic effects of the different coordination modes on the $\text{CpMn}(\text{CO})_2$ unit are very similar. The most sterically hindered



isomer, with the tertiary C–H bond coordinated, *viz.* CpMn(CO)₂(η^2 -C2-H-isopentane), is not observed. The only other ethane complex that has currently been observed by NMR spectroscopy is [(PONOP)Rh(η^2 -C-H-ethane)]⁺ where PONOP is a PNP pincer ligand.²² The chemical shifts of the coordinated η^2 -C-H-CH₃ group of the manganese alkane complexes occur close to δ –6, whilst those of a coordinated η^2 -C-H-CH₂ bond lie appreciably to higher field, in line with the reported observations on the analogous propane complex. These values represent averages of shifts for coordinated and uncoordinated C–H bonds because of rapid exchange of geminal protons. The CpMn(CO)₂(η^2 -C-H-ethane) lifetime is *ca.* 360 s whilst that of two of the CpMn(CO)₂(η^2 -C-H-isopentane) isomers are 395 ± 23 s and 372 ± 22 s. The similarity of the lifetimes could be interpreted in terms of interconversion of the isomers by a chain walking process that is slow on the NMR timescale and slow compared to that of the rhodium ethane complexes studied by Brookhart *et al.*²² Very satisfyingly, the lifetime of CpMn(CO)₂(η^2 -C-H-ethane) equates to a value of ΔG^\ddagger of 39 kJ mol^{–1} compared to 39 ± 2 kJ mol^{–1} for the activation energy (*E*_a) as measured by TRIR spectroscopy around room temperature.⁶¹ As demonstrated previously, the manganese complexes are vastly more reactive than their rhenium analogues which are stable in solution of the corresponding alkanes at 170–180 K.

Acknowledgements

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Notes and references

- C. Hall and R. N. Perutz, *Chem. Rev.*, 1996, **96**, 3125.
- M. Lersch and M. Tilset, *Chem. Rev.*, 2005, **105**, 2471.
- W. D. Jones, *Acc. Chem. Res.*, 2003, **36**, 140.
- W. H. Bernskoetter, C. K. Schauer, K. I. Goldberg and M. Brookhart, *Science*, 2009, **326**, 553.
- R. N. Perutz and S. Sabo-Etienne, *Angew. Chem., Int. Ed.*, 2007, **46**, 2578.
- S. D. Pike, A. L. Thompson, A. G. Algarra, D. C. Apperley, S. A. Macgregor and A. S. Weller, *Science*, 2012, **337**, 1648.
- L. Biber, L. D. Reuvenov, T. Revzin, T. Sinai, A. Zahavi and R. H. Schultz, *Dalton Trans.*, 2007, **41**; R. Krishnan and R. H. Schultz, *Organometallics*, 2001, **20**, 3314; R. H. Schultz, *Organometallics*, 2004, **23**, 4349.
- R. N. Perutz and J. J. Turner, *J. Am. Chem. Soc.*, 1975, **97**, 4791.
- M. Poliakoff and J. J. Turner, *Dalton Trans.*, 1974, 2276.
- J. A. Calladine, K. Q. Vuong, X. Z. Sun and M. W. George, *Pure Appl. Chem.*, 2009, **81**, 1667.
- M. K. Kuimova, W. Z. Alsindi, J. Dyer, D. C. Grills, O. S. Jina, P. Matousek, A. W. Parker, P. Portius, X. Z. Sun, M. Towrie, C. Wilson, J. X. Yang and M. W. George, *Dalton Trans.*, 2003, 3996.
- M. Besora, J.-L. Carreon-Macedo, A. J. Cowan, M. W. George, J. N. Harvey, P. Portius, K. L. Ronayne, X.-Z. Sun and M. Towrie, *J. Am. Chem. Soc.*, 2009, **131**, 3583.
- A. J. Cowan, P. Portius, H. K. Kawanami, O. S. Jina, D. C. Grills, X. Z. Sun, J. McMaster and M. W. George, *Proc. Natl. Acad. Sci. U. S. A.*, 2007, **104**, 6933.
- S. Geftakis and G. E. Ball, *J. Am. Chem. Soc.*, 1998, **120**, 9953.
- D. J. Lawes, S. Geftakis and G. E. Ball, *J. Am. Chem. Soc.*, 2005, **127**, 4134.
- G. E. Ball, C. M. Brookes, A. J. Cowan, T. A. Darwish, M. W. George, H. K. Kawanami, P. Portius and J. P. Rourke, *Proc. Natl. Acad. Sci. U.S.A.*, 2007, **104**, 6927.
- D. J. Lawes, T. A. Darwish, T. Clark, J. B. Harper and G. E. Ball, *Angew. Chem., Int. Ed.*, 2006, **45**, 4486.
- S. B. Duckett, M. W. George, O. S. Jina, S. L. Matthews, R. N. Perutz, X. Z. Sun and K. Q. Vuong, *Chem. Commun.*, 2009, 1401.
- J. A. Calladine, O. Torres, M. Anstey, G. E. Ball, R. G. Bergman, J. Curley, S. B. Duckett, M. W. George, A. I. Gilson, D. J. Lawes, R. N. Perutz, X.-Z. Sun and P. C. Vollhardt, *Chem. Sci.*, 2010, **1**, 622.
- R. D. Young, A. F. Hill, W. Hillier and G. E. Ball, *J. Am. Chem. Soc.*, 2011, **133**, 13806; R. D. Young, D. J. Lawes, A. F. Hill and G. E. Ball, *J. Am. Chem. Soc.*, 2012, **134**, 8294.
- J. A. Calladine, S. B. Duckett, M. W. George, S. L. Matthews, R. N. Perutz, O. Torres and K. Q. Vuong, *J. Am. Chem. Soc.*, 2011, **133**, 2303.
- M. D. Walter, P. S. White, C. K. Schauer and M. Brookhart, *J. Am. Chem. Soc.*, 2013, **135**, 15933.
- G. J. Kubas, *Metal Dihydrogen and σ -Bond Complexes*, Kluwer Academic Publishers, New York, 2001.
- R. H. Crabtree, *Angew. Chem., Int. Ed.*, 1993, **32**, 789.
- G. J. Kubas, *J. Organomet. Chem.*, 2014, **751**, 33.
- I. A. I. Mkhaldid, J. H. Barnard, T. B. Marder, J. M. Murphy and J. F. Hartwig, *Chem. Rev.*, 2010, **110**, 890.
- G. Alcaraz, M. Grellier and S. Sabo-Etienne, *Acc. Chem. Res.*, 2009, **42**, 1640.
- J. Y. Corey and J. Braddock-Wilking, *Chem. Rev.*, 1999, **99**, 175.
- J. Y. Corey, *Chem. Rev.*, 2011, **111**, 863.
- Z. Y. Lin, *Struct. Bonding*, 2008, **130**, 123.
- G. I. Nikonov, *Adv. Organomet. Chem.*, 2005, **53**, 217.
- J. L. Vincent, S. Luo, B. L. Scott, R. Butcher, C. J. Unkefer, C. J. Burns, G. J. Kubas, A. Lledós, F. Maseras and J. Tomàs, *Organometallics*, 2003, **22**, 5307.
- D. R. Evans, T. Drovetskava, R. Bau, C. A. Reed and P. D. W. Boyd, *J. Am. Chem. Soc.*, 1997, **119**, 3633.
- I. Castro-Rodriguez, H. Nakai, P. Gantzel, L. N. Zakharov, A. L. Rheingold and K. Meyer, *J. Am. Chem. Soc.*, 2003, **125**, 15734.
- E. D. Bloch, W. L. Queen, R. Krishna, J. M. Zadrozny, C. M. Brown and J. R. Long, *Science*, 2012, **335**, 1606; S. J. Geier, J. A. Mason, E. D. Bloch, W. L. Queen, M. R. Hudson, C. M. Brown and J. R. Long, *Chem. Sci.*, 2013, **4**, 2054.
- D. D. Wick, K. A. Reynolds and W. D. Jones, *J. Am. Chem. Soc.*, 1999, **121**, 3974.



- 37 W. D. Jones, *Inorg. Chem.*, 2005, **44**, 4475; T. O. Northcutt, D. D. Wick, A. J. Vetter and W. D. Jones, *J. Am. Chem. Soc.*, 2001, **123**, 7257.
- 38 J. F. Hartwig, K. S. Cook, M. Hapke, C. D. Incarvito, Y. Fun, C. E. Webster and M. B. Hall, *J. Am. Chem. Soc.*, 2005, **127**, 2538.
- 39 C. S. Wei, C. A. Jimenez-Hoyos, M. F. Videa, J. F. Hartwig and M. B. Hall, *J. Am. Chem. Soc.*, 2010, **132**, 3078.
- 40 J. Choi, A. H. Roy MacArthur, M. Brookhart and A. S. Goldman, *Chem. Rev.*, 2011, **111**, 1761.
- 41 A. S. Goldman, A. H. Roy, Z. Huang, R. Ahuja, W. Schinski and M. Brookhart, *Science*, 2006, **312**, 257.
- 42 D. C. Leitch, Y. Choi Lam, J. A. Labinger and J. E. Bercaw, *J. Am. Chem. Soc.*, 2013, **135**, 10302.
- 43 J. E. Bercaw, G. S. Chen, J. A. Labinger and B. L. Lin, *J. Am. Chem. Soc.*, 2008, **130**, 17654; J. E. Bercaw, G. S. Chen, J. A. Labinger and B. L. Lin, *Organometallics*, 2010, **29**, 4354.
- 44 G. S. Chen, J. A. Labinger and J. E. Bercaw, *Proc. Natl. Acad. Sci. U. S. A.*, 2007, **104**, 6915.
- 45 M. Schlangen and H. Schwarz, *Dalton Trans.*, 2009, 10155.
- 46 B. G. Hashiguchi, M. M. Konnick, S. M. Bischof, S. J. Gustafson, D. Devarajan, N. Gunsalus, D. H. Ess and R. A. Periana, *Science*, 2014, **343**, 1232.
- 47 R. A. Periana, R. A. Mironov, O. D. J. Taube, G. Bhalla and C. J. Jones, *Science*, 2003, **301**, 814.
- 48 E. V. Kudrik, P. Afanasiev, L. X. Alvarez, P. Dubourdeaux, M. Clémancey, J.-M. Latour, G. Blondin, D. Bouchu, F. Albrieux, S. E. Nefedov and A. B. Sorokin, *Nat. Chem.*, 2012, **4**, 1024.
- 49 A. Caballero, E. Despagnet-Ayoub, M. M. Díaz-Requejo, A. Díaz-Rodríguez, M. E. González-Núñez, R. Mello, B. K. Muñoz, W.-S. Ojo, G. Asensio, M. Etienne and P. J. Pérez, *Science*, 2011, **332**, 835.
- 50 M. V. Kirillova, M. L. Kuznetsov, P. M. Reis, J. A. L. da Silva, J. J. R. Fraústo da Silva and A. J. L. Pombeiro, *J. Am. Chem. Soc.*, 2007, **129**, 10531.
- 51 D. Balcells, E. Clot and O. Eisenstein, *Chem. Rev.*, 2010, **112**, 749.
- 52 B. A. Vastine and M. B. Hall, *Coord. Chem. Rev.*, 2009, **253**, 1202.
- 53 M. D. Walter, P. S. White, C. K. Schauer and M. Brookhart, *New J. Chem.*, 2011, **35**, 2884.
- 54 B. Chan and G. E. Ball, *J. Chem. Theory Comput.*, 2013, **9**, 2199.
- 55 J. A. Calladine, A. Love, P. A. Fields, R. G. M. Wilson and M. W. George, *Appl. Spectrosc.*, 2014, **68**, 324.
- 56 J. A. Calladine and M. W. George, *Spectrosc. Eur.*, 2009, **21**, 6.
- 57 T.-L. Hwang and A. J. Shaka, *J. Magn. Reson.*, 1995, **112**, 275.
- 58 M. L. Lui, X. A. Mao, C. H. Ye, H. Huang, J. K. Nicholson and J. C. Lindon, *J. Magn. Reson.*, 1998, **132**, 125.
- 59 R. A. Green, R. W. Adams, S. B. Duckett, R. E. Mewis, D. C. Williamson and G. G. R. Green, *Prog. Nucl. Magn. Reson. Spectrosc.*, 2012, **67**, 1.
- 60 G. I. Childs, C. S. Colley, J. Dyer, D. C. Grills, X. Z. Sun, J. X. Yang and M. W. George, *J. Chem. Soc., Dalton Trans.*, 2000, 1901.
- 61 A. J. Cowan, *PhD Thesis*, University of Nottingham, 2007.

