



Cite this: *Dalton Trans.*, 2015, **44**, 11819

A computational analysis of the apparent *nido* vs. *hypho* conflict: are we dealing with six- or eight-vertex open-face diheteroboranes?†‡

João Pedro F. Nunes,^a Josef Holub,^b David W. H. Rankin,^c Derek A. Wann*^a and Drahomír Hnyk*^b

A series of computational studies have been undertaken to investigate the electronic structures and bonding schemes for six hetero-substituted borane cages, all of which have been presented in the literature as potential *hypho* structures. The six species are *hypho*-7,8-[C₂B₆H₁₃][−] (**1a**), *hypho*-7,8-[CSB₆H₁₁][−] (**1b**), *hypho*-7,8-[S₂B₆H₉][−] (**1c**), *hypho*-7,8-[NSB₆H₁₁] (**1d**), *exo*-7-Me-*hypho*-7,8-[NCB₆H₁₂] (**1e**), and *endo*-7-Me-*hypho*-7,8-[NCB₆H₁₂] (**1f**) and the so-called *mno* rule has been applied to each of them. As no structural data are known for the carbathia-, azathia-, and dithiahexaboranes, we have also applied the *ab initio*/GIAO/NMR structural tool for **1b–1d**, with **1c** having been prepared for this purpose. We conclude that an *mno* count of 10 means that **1a**, **1b**, **1d**, **1e**, and **1f** should be termed *pseudo-nido* or *pseudo-hypho*. Only **1c** can be considered to be correctly termed *hypho*-7,8-[S₂B₆H₉][−].

Received 17th April 2015,
Accepted 2nd June 2015
DOI: 10.1039/c5dt01460c

www.rsc.org/dalton

Introduction

Polyhedral boranes and heteroboranes have long provided interest covering a range of chemical research,¹ where their unexpected structures, innovative bonding schemes, and unusual chemical properties have resulted in them being considered promising candidates for use in medicinal and materials applications. For decades now an understanding of the molecular architectures of these species has led directly and indirectly to practical uses for boranes and future structural studies will undoubtedly lead to potential further uses being identified.

The variety of bonding schemes exhibited by polyhedral boranes and heteroboranes was initially noted by Lipscomb,²

with Williams³ and Wade⁴ also making significant contributions. Williams reported a major breakthrough by recognising that the experimentally isolated *nido*, *arachno*, and *hypho* boranes could be derived from the nearest *closo* structures by the removal of one, two, and three {BH}^{2−} vertices, respectively. This led to the derivation of Wade's *n* + 1 electron-pair rule for a *closo*-type cluster, where *n* is the number of vertices.^{4,5} Subsequently, the number of skeletal electron pairs required for stable *nido*, *arachno*, and *hypho* skeletons was determined to be *n* + 1 + *p*, where *p* is the number of missing vertices; it follows that *closo*-[B_{*n*}H_{*n*}]^{2−} requires such a formal charge to comply with Wade's rule. It should be noted that *closo* systems are known experimentally for *n* = 5–12, where the icosahedral cluster with I_h point-group symmetry (*n* = 12) is the most stable of the *closo* series.

The unusually high stability of *closo*-[B₁₂H₁₂]^{2−}, as well as the existence of condensed B₁₂ units [for example, in the so-called macropolyhedral boron clusters B₂₀H₁₆⁶ and B₂₁H₁₈[−] (ref. 7)], meant that a generalisation of the electron-counting rules was required since such structures are beyond the scope of the Williams–Wade formalism. A generally applicable electron-counting rule – the so-called *mno* rule – was derived by Jemmis,⁸ based on Hückel's rule it allows the structures of macropolyhedral boranes and metallaboranes to be characterised alongside simple boranes. According to the *mno* rule, *m* + *n* + *o* electron pairs are necessary for a macropolyhedral system to be stable, where *m* is the number of individual polyhedral subclusters from which a macropolyhedral cluster is composed, *n* is the number of vertices, and *o* is the number of

^aDepartment of Chemistry, University of York, Heslington, York YO10 5DD, UK. E-mail: derek.wann@york.ac.uk

^bInstitute of Inorganic Chemistry of the Academy of Sciences of the Czech Republic, v.v.i., 250 68 Husinec-Řež, Czech Republic. E-mail: hnyk@iic.cas.cz

^cSchool of Chemistry, University of Edinburgh, Joseph Black Building, David Brewster Road, Edinburgh, EH9 3FJ, UK

† Dedicated to the memory of Professor Ken Wade.

‡ Electronic supplementary information (ESI) available: Cartesian coordinates for the structures of **cB11**, **nB10**, and **aB9** (Tables S1–S3), **1a–1f** (Tables S4–S9), and *nido*-[B₆H₁₁][†] (Table S10), the *s* character, *p* character, and resultant hybridisation of **1a–1f** for all levels of theory (Tables S11–S13), and charges on the heteroatoms of **1a–1f** (Table S14). The molecular structures of **1a–1f** with full atom numbering (Fig. S1), the equivalent for the *closo-nido-arachno-hypho* relationship for *n* = 11 with full atom numbering (Fig. S2), and the molecular structure of *nido*-[B₆H₁₁][†] with full atom numbering (Fig. S3). See DOI: 10.1039/c5dt01460c



single-vertex-sharing condensations. For *nido*, *arachno*, and *hypho* arrangements, one, two, and three additional pairs of electrons are required. Wade's $n + 1$ rule can be considered as a special case of the *mno* rule, where $m = 1$ and $o = 0$.

Some *hypho* and *nido* complexes have very similar structures and electronic structure investigation would seem sensible to ensure clusters are not wrongly classified. In this work we have performed *ab initio* and DFT analyses of the bonding schemes for six potential *hypho* structures {*hypho*-7,8-[C₂B₆H₁₃]⁻ (**1a**), *hypho*-7,8-[CSB₆H₁₁]⁻ (**1b**), *hypho*-7,8-[S₂B₆H₉]⁻ (**1c**), *hypho*-7,8-[NSB₆H₁₁]⁻ (**1d**), *exo*-7-Me-*hypho*-7,8-[NCB₆H₁₂]⁻ (**1e**), and *endo*-7-Me-*hypho*-7,8-[NCB₆H₁₂]⁻ (**1f**)}, also employing the *mno* rule to

see what results it yields. All six structures are shown in Fig. 1, while Fig. S1† also gives the hydrogen-atom numbering. As no structural data are currently available for the carbathia-, azathia-, and dithiahexaboranes, we have also applied the *ab initio*/GIAO/NMR structural method to **1b–1d**.

Experimental section

NMR

Experimental ¹¹B NMR chemical shifts for **1b** and **1d** were taken from ref. 9 and 10, respectively; **1c** was prepared according to ref. 11.

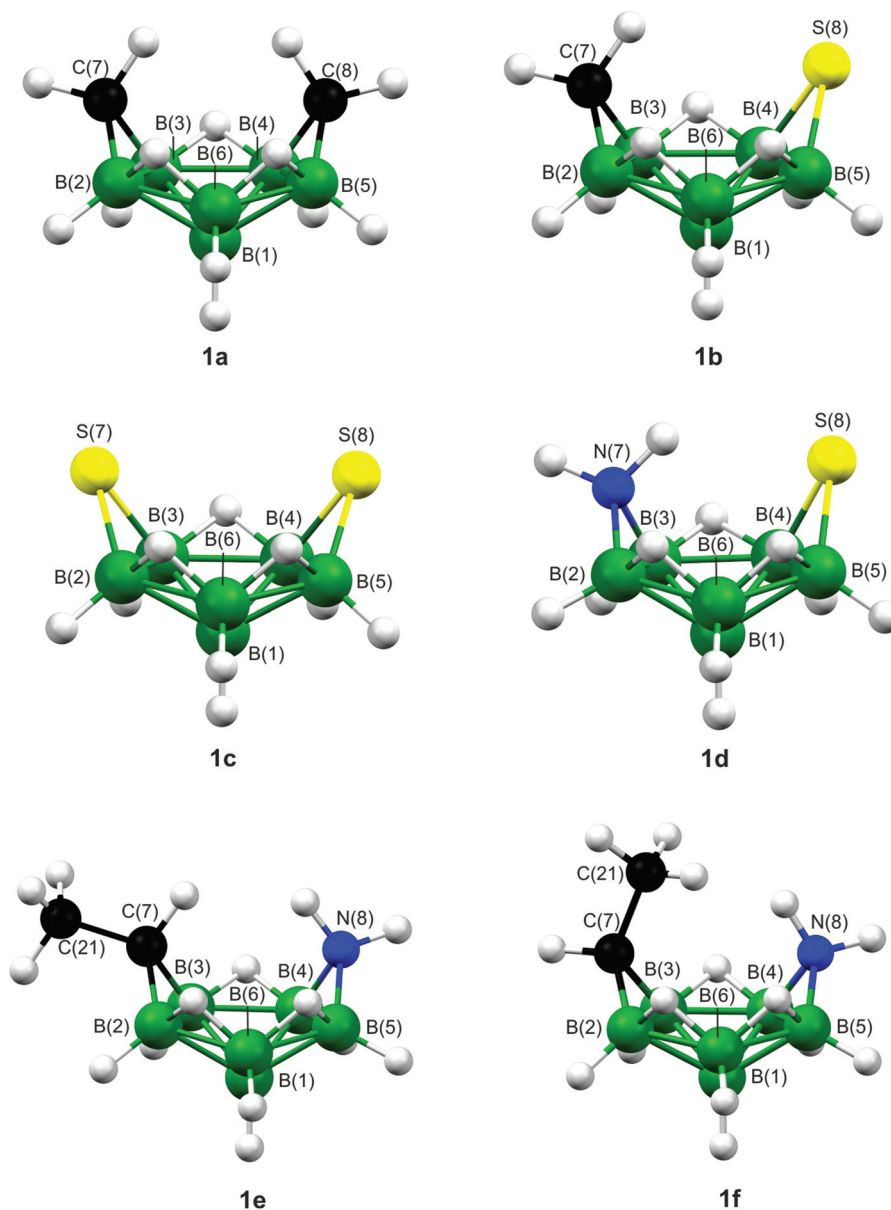


Fig. 1 The molecular structures of eight-vertex diheteroboranes that have been proposed as *hypho*-type clusters. For clarity hydrogen-atom numbering has been omitted. The molecules are *hypho*-7,8-[C₂B₆H₁₃]⁻ (**1a**), *hypho*-7,8-[CSB₆H₁₁]⁻ (**1b**), *hypho*-7,8-[S₂B₆H₉]⁻ (**1c**), *hypho*-7,8-[NSB₆H₁₁]⁻ (**1d**), *exo*-7-Me-*hypho*-7,8-[NCB₆H₁₂]⁻ (**1e**), and *endo*-7-Me-*hypho*-7,8-[NCB₆H₁₂]⁻ (**1f**).



Quantum chemical calculations

The geometries of *closo*-[B₁₁H₁₁]²⁻ (**cB11**), *nido*-[B₁₀H₁₀]⁴⁻ (**nB10**) and *arachno*-[B₉H₉]⁶⁻ (**aB9**) were fully optimised using Gaussian09.¹² Each of these structures, as well as that of the related *hypho* species (**hB8**), is depicted in Scheme 1. (The scheme is reproduced in Fig. S2,† where the atom numbering is given.) The geometry optimisations for each of **cB11**, **nB10**, and **aB9** were performed at the B2PLYP,¹³ MP2,^{14–18} B98,¹⁹ B97d,²⁰ PBE,^{21,22} PW91,^{23–27} and HFS^{28–30} levels of theory using the 6-311+G(d,p)^{31,32} basis set on all atoms; the nature of any stationary points on the potential-energy surfaces were investigated using frequency calculations. The initial geometries for **nB10** and **aB9** were optimised from geometries obtained by removing one and two vertices from **cB11**, respectively, using GaussView 5.0.³³ The bonding orbitals for **cB11**, **nB10**, and **aB9** were investigated using natural bond orbital (NBO) analyses to look for three-centre bonds and resonance structures.

The geometries of the heteroboranes **1a–1f** were also fully optimised using Gaussian09 and the method and basis set combinations described above, with the character of each stationary point verified by frequency calculations. Magnetic shieldings were calculated for **1b–1d** by running GIAO³⁴ jobs with TZP basis set II by Huzinaga,³⁵ which is well suited for this purpose. In order to investigate the relationships between **1a–1f** and their true borane analogues, the heteroatoms for each borane cage were replaced by hydrogen atoms using GaussView 5.0. The geometries of the species formed by the inclusion of hydrogen atoms in place of heteroatoms were optimised using the same levels of theory and basis sets previously employed, and frequency calculations were performed to verify the nature of any stationary points; the distances between the substituted hydrogen atoms and the boron-cage atoms were fixed at values optimised for the B–C/N/S distances. NBO analyses were also performed for all of the H-substituted structures.

Calculating *mno* values

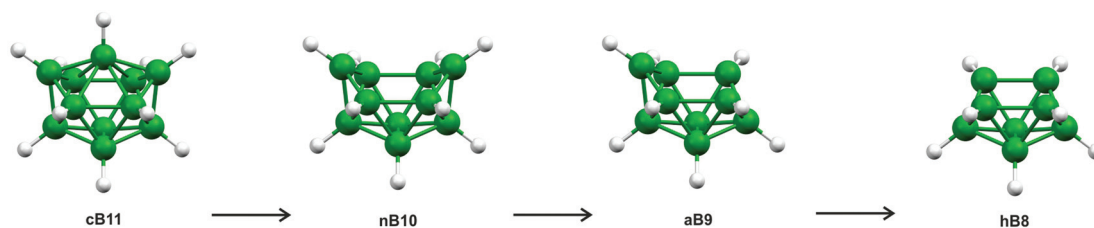
The *mno* rule is used to determine the number of pairs of electrons required for a species to be considered stable by calculating $m + n + o$ (for the meanings of m , n , and o , see above). As an illustration the *mno* rule is applied to an icosahedral borane by taking the values $m = 1$, $n = 12$, and $o = 0$, yielding

$m + n + o = 13$ electron pairs, the same result as the $n + 1$ electron-pair rule. Each vertex with one terminal atom contributes all-but-one of its electrons to cluster bonding, with the remaining electron involved in the *exo* covalent bond. However, in the three isomeric icosahedral carbaboranes *closo*-C₂B₁₀H₁₂, the BH groups each donate one pair of electrons and each CH moiety contributes three electrons (1.5 electron pairs) to the polyhedral bonding, thus satisfying the *mno* rule [*i.e.* $10 + (2 \times 1.5) = 13$ for these neutral carbaboranes].³⁶ Similarly, in the icosahedral azaborane *closo*-NB₁₁H₁₂³⁷ there are eleven BH groups that contribute eleven electron pairs as well as an NH group that donates two electron pairs to the skeletal bonding. Considering molecules containing third-row elements, the neutral icosahedral *closo*-SB₁₁H₁₁ thiaborane is known,³⁸ in which BH groups provide eleven electron pairs; sulfur has a tendency to retain one of its lone pairs, contributing four electrons (two electron pairs) to the cluster.

The following *closo* dianions, presented in order of stability, B₁₂H₁₂²⁻, B₁₁H₁₁²⁻, and B₅H₅²⁻, can be used to derive *nido*, *arachno*, and *hypho* heteroboranes *via* the removal and substitution of vertices.³⁹ However, according to the Williams–Wade concept, heteroboranes cannot exist with two, three, or four BH moieties. *Closo*-[B₅H₅]²⁻ is, therefore, not a suitable “parent” species for any of the *nido*, *arachno*, or *hypho* heteroborane structures.

There are heteroboranes with molecular shapes based on *nido*-[B₁₀H₁₀]⁴⁻ (**nB10**)⁴⁰ and *arachno*-[B₉H₉]⁶⁻ (**aB9**),⁴¹ which are derived from the C_{2v}-symmetric *closo*-[B₁₁H₁₁]²⁻ (**cB11**), as shown in Scheme 1. As well as *closo*, *nido*, and *arachno* clusters, Scheme 1 depicts the hypothetical *hypho*-[B₈H₈]⁸⁻ (**hB8**). In contrast, our so-called *hypho* eight-vertex-type heteroboranes resemble the six-vertex arrangement adopted by *nido*-[B₆H₁₁]⁺ (**nB6**),⁴² shown in Fig. 2. (Fig. S3† depicts the same structure but also includes hydrogen-atom numbering.)

The molecule *hypho*-7,8-[C₂B₆H₁₃]⁻ (**1a**) has previously been structurally characterised by applying the *ab initio*/GIAO/NMR structural tool,⁴³ which confirmed its C_s-symmetric “helmet-like” structure (Fig. 1). Three other proposed *hypho* species (also shown in Fig. 1) have been calculated to have the same structural motif as **1a**; these are *hypho*-7,8-[CSB₆H₁₁]⁻ (**1b**),⁹ *hypho*-7,8-[S₂B₆H₉]⁻ (**1c**),¹¹ and *hypho*-7,8-[NSB₆H₁₁]⁻ (**1d**).¹⁰ The last of these examples (**1d**) illustrates how nitrogen can be accommodated in an eight-vertex *hypho* arrangement, with the same arrangement found in *exo*- and *endo*-7-Me-*hypho*-7,8-



Scheme 1 Schematic representation of the *closo*–*nido*–*arachno*–*hypho* relationship for $n = 11$.



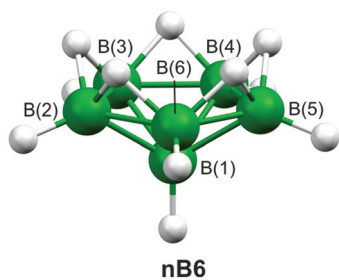


Fig. 2 The molecular structure of *nido*-[B₆H₁₁]⁺, whose structure resembles a *hypho* eight-vertex heteroborane. For clarity hydrogen-atom numbering has been omitted.

[NCB₆H₁₂] (**1e** and **1f**, respectively). Again, GIAO calculations of the shielding tensors have been used for structural characterisation.⁴⁴

Results and discussion

Calculations were initially performed to investigate the bonding patterns of the parent **cb11**, **nb10**, **ab9**, and **hb8** systems, as shown in Scheme 1. Apart from the standard HF calculations perturbed with the MP2 model chemistry, we employed six density functional theory (DFT) approaches to span all the possibilities afforded by DFT, *i.e.* B98 (stand-alone hybrid functional), B97D (stand-alone pure functional), HFS (exchange-only functional), PW91 (correlation functional), PBE (exchange-combined functional), and B2PLYP (double-hybrid functional), the latter utilising HF exchange and an MP2-like correlation. All seven model chemistries used the 6-31G(d) basis set for the initial geometry optimisations. However, all attempts to identify a minimum on the potential-energy hypersurface of **hb8** failed because of the formation of either a series of flattened triangular borane architectures or a random assembly of eight boron atoms. The final basis set used for optimising **cb11**, **nb10**, and **ab9** was 6-311G(d,p). Cartesian coordinates for these structures are given in Tables S1–S3.‡ Attempts to use extra diffuse functions on these basis sets caused the optimisations to diverge. Second-derivative analyses of these systems showed that they represented minima on the respective potential-energy hypersurfaces. The seven model chemistries were also used to calculate Hartree–Fock and Kohn–Sham natural orbitals and no two-centre two-electron (2c–2e) bonds were predicted for the cage motifs, only for the terminal B–H bonds.

In order to get a deeper insight into the bonding patterns of **1a–1f**, we performed entirely the same computational procedures as for **cb11**, **nb10**, **ab9**, and **hb8**. All of these diheterohexaboranes structures were identified as potential minima at *ab initio* and all DFT levels. Cartesian coordinates relating to the MP2/6-311+G(d,p) calculations are given in Tables S4–S9.‡ The strong resemblance of each of the **1a–1f** molecular geometries to that of *nido*-[B₆H₁₁]⁺ prompted us to replace each

heteroatom with a hydrogen atom (H_r), with the B–H_r distances fixed at the values optimised for the B–C/N/S bond lengths. Any substituents originally bonded to C or N were omitted. The new structures generated were termed **1a/H-1e/H** (no calculations were required for **1f/H** which is identical to **1e/H**). The structure of *nido*-[B₆H₁₁]⁺ (Fig. 2) was calculated as a comparison as it has the same arrangement of H-bridges; Cartesian coordinates relating to its geometry at the MP2/6-311+G(d,p) level are given in Table S10.‡ As might be expected, frequency calculations performed for each of **1a/H-1e/H** showed imaginary frequencies relating to the H_r atoms.

Performing NBO analyses for each of **1a/H-1e/H** revealed that three-centre two-electron (3c–2e) bonding is present throughout each of the clusters, *i.e.* also in B–H_r–B bridges. When these structures were allowed to relax further by optimising the B–H_r distances, 3c–2e bonding persisted for all five species **1a/H-1e/H**. The geometry optimisations performed using the 6-311+G** basis sets demonstrate that 3c–2e bonding exists regardless of the lengths of the B–H bridging distances; this was true for all seven methods employed. Similar analyses for **1a–1f** (where the heteroatoms are present) showed a different picture, where all B–C/N/S distances were classified as 2c–2e bonds. If the NBO analyses had revealed 2c–2e bonds for B–H_r–B bridges in **1a/H-1e/H**, we could have stated that the heteroatoms present in **1a–1f** were *not* part of multicentre bonding. However, as this was not the case, we therefore had to follow another way of finding the nature of bonding in **1a–1f**.

Table 1 shows the hybridisation of the heteroatoms in **1a–1f** when transforming canonical orbitals to natural ones using NBO analysis. (Such calculations were performed for all levels of theory specified in the Experimental section and then averaged. The range of values is also shown. The characters and hybridisations of the B–C/N/S bonds for each of **1a–1f** at all levels of theory are given in Tables S11–S13.‡) Table 1 shows that the heteroatoms in each species are more or less sp³-hybridised, as is the case for textbook examples such as CH₄, NH₃, and H₂S. This means that, as well as 2c–2e bonds to boron atoms, all of the heteroatoms are covalently bonded to hydrogen atoms and any carbon atoms that are external to the cage motif. In addition, NBO analyses showed that heteroatoms in **1a–1f** are negatively charged. The average calculated charges are presented in Table 2 (with the corresponding values for each level of theory in Table S14.‡) and are in line with those values for classical covalent species such as CH₄, NH₃, and H₂S. However, such an observation is in contrast to the electron distribution experimentally determined for the icosahedral species, in which the midpoint of CC vector and sulfur atom were found to be positively charged as revealed by vector algebra of experimental dipole moments measured for the *exo*-substituted icosahedra.^{36d,38b}

Armed with these results, we applied the *mno* rule to *nido*-[B₆H₁₁]⁺ as well as to **1a–1f**. For *nido*-[B₆H₁₁]⁺ *m* = 1, *n* = 6, *o* = 0, and *p* = 1, meaning that eight electron pairs are required to stabilise this system. There are six B–H fragments, each of which contributes one electron pair. The remaining two elec-



Table 1 Hybridisation of the heteroatoms (C/N/S) bonded to the boron atoms. The bond characters were calculated using NBO analyses and are averaged across all levels of theory specified in the experimental section^a

| | Bond | s character/% | p character/% | Hybridisation ^b |
|-----------|-----------|---------------|---------------|----------------------------|
| 1a | B(2)–C(7) | 28.45 | 71.53 | 2.51 ± 0.01 |
| | B(3)–C(7) | 26.84 | 73.14 | 2.73 ± 0.06 |
| | B(4)–C(8) | 26.84 | 73.14 | 2.73 ± 0.06 |
| 1b | B(5)–C(8) | 28.45 | 71.53 | 2.51 ± 0.01 |
| | B(2)–C(7) | 28.62 | 71.34 | 2.49 ± 0.01 |
| | B(3)–C(7) | 26.08 | 73.89 | 2.83 ± 0.07 |
| 1c | B(4)–S(8) | 16.96 | 82.87 | 4.91 ± 0.37 |
| | B(5)–S(8) | 18.49 | 81.22 | 4.40 ± 0.24 |
| | B(2)–S(7) | 18.50 | 81.31 | 4.40 ± 0.23 |
| 1d | B(3)–S(7) | 16.72 | 83.10 | 4.99 ± 0.36 |
| | B(4)–S(8) | 16.72 | 83.10 | 4.99 ± 0.36 |
| | B(5)–S(8) | 18.50 | 81.31 | 4.40 ± 0.23 |
| 1e | B(2)–N(7) | 29.38 | 70.61 | 2.40 ± 0.02 |
| | B(3)–N(7) | 27.46 | 72.53 | 2.64 ± 0.06 |
| | B(4)–S(8) | 16.61 | 83.20 | 5.02 ± 0.32 |
| 1f | B(5)–S(8) | 17.08 | 82.71 | 4.86 ± 0.32 |
| | B(2)–C(8) | 26.28 | 73.70 | 2.80 ± 0.03 |
| | B(3)–C(8) | 25.80 | 74.19 | 2.88 ± 0.05 |
| 1f | B(4)–N(7) | 28.69 | 71.31 | 2.49 ± 0.05 |
| | B(5)–N(7) | 29.48 | 70.52 | 2.39 ± 0.02 |
| | B(2)–C(8) | 26.12 | 73.85 | 2.83 ± 0.03 |
| 1f | B(3)–C(8) | 26.00 | 73.98 | 2.85 ± 0.05 |
| | B(4)–N(7) | 28.49 | 71.50 | 2.51 ± 0.05 |
| | B(5)–N(7) | 29.41 | 70.59 | 2.40 ± 0.02 |

^aThe characters and hybridisations from each of the individual calculations are given in Tables S11–S13. ^bThe errors quoted are the standard deviation of the values calculated for the different model chemistries.

Table 2 Atomic charges of the heteroatoms (C/N/S) within the borane cage motif. The charges are averaged across all levels of theory specified in the experimental section^a

| | Atom | Charge ^b |
|-----------|------|---------------------|
| 1a | C(7) | −0.92 ± 0.02 |
| | C(8) | −0.92 ± 0.02 |
| 1b | C(7) | −0.91 ± 0.02 |
| | S(8) | −0.27 ± 0.03 |
| 1c | S(7) | −0.22 ± 0.03 |
| | S(9) | −0.22 ± 0.03 |
| 1d | N(7) | −0.95 ± 0.02 |
| | S(8) | −0.18 ± 0.04 |
| 1e | N(7) | −0.93 ± 0.02 |
| | C(8) | −0.73 ± 0.02 |
| 1f | N(7) | −0.94 ± 0.02 |
| | C(8) | −0.74 ± 0.02 |

^aThe charges from each of the individual calculations are given in Table S14. ^bThe errors quoted are the standard deviation of the values calculated for the different model chemistries.

tron pairs are available from the five bridging hydrogen atoms with one spare electron present; the total number of electron pairs is $6 + 2.5 = 8.5$. As there is one excess electron the hypothetical *nido*-[B₆H₁₁]⁺ is obviously a cation. (For *nido*-B₆H₁₀ we have eight electron pairs and, consequently, the molecule is neutral.)

Table 3 Computed (GIAO) and experimental NMR chemical shifts for **1b–1d**

| | Vertex | | | | | |
|-----------------------------|--------|-------|-------|-------|-------|------|
| | B(1) | B(3) | B(6) | B(4) | B(2) | B(5) |
| 1b GIAO ^a | −56.5 | −36.3 | −33.5 | −25.1 | −4.6 | 3.1 |
| Exp. ^{b,c} | −54.8 | −33.2 | −29.4 | −23.9 | −3.1 | 2.0 |
| 1c GIAO ^a | −55.8 | 5.3 | −26.1 | −31.4 | −26.1 | 5.3 |
| Exp. ^{b,d} | −52.2 | 6.8 | −22.4 | −25.0 | −22.4 | 6.8 |
| 1d GIAO ^a | −56.3 | −26.4 | −24.9 | −21.7 | −2.0 | 2.2 |
| Exp. ^{b,e} | −55.1 | −25.6 | −22.4 | −20.8 | −1.6 | 1.8 |

^aGIAO-MP2/II//6-311+G**. ^bMeasured in CDCl₃. ^cRef. 9. ^dThis work. ^eRef. 10.

To apply the *mno* rule to **1a–1f** we need to determine whether they adopt a *hypho* or a *nido* electron count. Let's start by supposing that they are considered to be *hypho*, giving $m = 1$, $n = 8$, $o = 0$, $p = 3$, meaning that 12 electron pairs are required. In order to comply with a *hypho* electron count the heteroatoms must contribute the following number of electron pairs: sulfur 2, carbon 2, and nitrogen 2.5 for the charges to be “*hypho*” correct.

Therefore, for **1a** we have $7.5 + 4 = 11.5$ electron pairs, thus requiring one extra electron to stabilise the system. This C_s skeleton is therefore characterised with a single negative charge. The same applies to the other non-nitrogen-containing compounds (**1b** and **1c**), where sulfur and/or carbon contributes 4 electron pairs, resulting in an overall single negative charge for these species. However, such electron pair contributions from the heteroatoms are in conflict with the NBO results shown in Table 1, where C, N, and S are more or less sp³-hybridised and are connected to the boron atoms by conventional 2c–2e bonds. Moreover, the carbon and nitrogen atoms are not naked, but rather are bonded also to two *exo* atoms or groups, *i.e.* the corresponding one electron pair from each of these heteroatoms cannot contribute to the skeletal moieties.

Secondly, we may formally consider **1a–1f** to be *nido* clusters, where values of $n = 6$, $m = 1$, $o = 0$, and $p = 1$ yields an *mno* value of eight and, consequently, eight electron pairs are required to stabilise these molecules, as was the case for *nido*-[B₆H₁₁]⁺. In the latter, six B–H bonds contribute six electron pairs, and five hydrogen bridges contribute 2.5 electron pairs, *i.e.* this system must have a single positive charge ($6 + 2.5 = 8.5$ and so one electron must be removed to get 8). There is also another hypothetical system, *nido*-[B₆H₉][−], which has three H-bridges and complies with the *mno* rule ($mno = 8$; $6 + 1.5 = 7.5$, so the formal charge is −1). On that basis, there is no scope for the two additional bridging bonds that would be required for this to be a *nido* system with a single negative charge. Such a charge is, however, unambiguously observed for **1a**, **1b**, and **1c**, in which there are five bridges (three B–H–B bridges and two B–C/S–B bridges). The same argument can be applied to relate the neutral molecules **1d**, **1e**, and **1f**



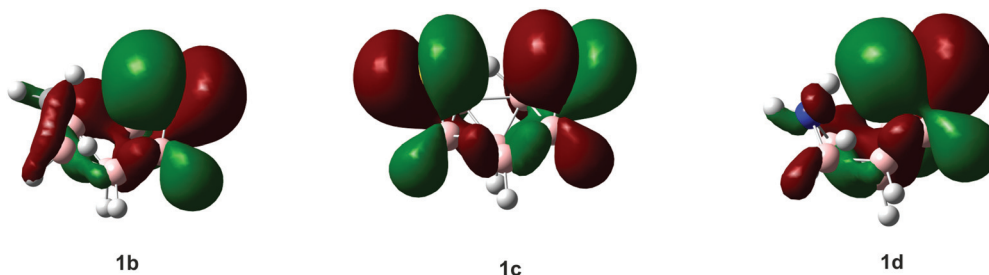


Fig. 3 HOMO for **1b–1d** at HF/6-311+G(d,p).

to *nido*-B₆H₁₀ ($6 + 4/2 = 8$; there is no scope for further bridge-type bonding to comply simultaneously with the *nido* electron count requirement and to keep the system neutral).

Only **1c** can truly be considered to be correctly classified as *hypho*-7,8-[S₂B₆H₉][−], since **1c** can afford to accept four electron pairs from the sulfur atoms into cluster bonding leaving one lone pair of electrons on each sulfur atom.

We must therefore conclude that **1a**, **1b**, **1d**, **1e**, and **1f** are, in reality, neither formally *hypho* skeletons nor *nido* structures. To be formally *hypho*, an *mno* value of 12 should be satisfied for **1a**, **1b**, **1d**, **1e**, and **1f**. Moreover, we would actually need to have an *mno* value of 10 to “accommodate” two extra B–C/N/S–B bridges and to comply with *nido* requirement. Since each of **1a**, **1b**, **1d**, **1e** and **1f** is analysed in terms of having a *mno* value of 12 or 8 (and not 10), we might call these systems *pseudo-nido* or *pseudo-hypho*. As stated above, **1c** can be correctly classified as *hypho*-7,8-[S₂B₆H₉][−].

Small clusters are known to be very sensitive to the inclusion of electron dynamic correlation when shielding tensors are being calculated.⁴⁵ The GIAO-MP2 calculations predict ¹¹B chemical shifts that compare well with experimental values. Table 3 compares computed and experimental ¹¹B chemical shifts for **1b–1d**.

While the computed and experimental ¹¹B NMR chemical shifts generally compare well, the presence of sulfur in **1b–1d** makes these fits for some atoms slightly worse than for clusters reported in the literature that do not contain a third-row element. Such discrepancies can be attributed to the inadequacy of using a triple-zeta Huzinaga type-II basis set on sulfur.⁴⁶ The most striking features of the individual spectra are shifts to low frequencies for the “bottom” boron atom, B(1). When comparing these values we clearly see the difference between the nature of the ¹¹B chemical shifts for B(2) and B(3) in **1b** and **1d** and those of **1c**. Inspecting the highest occupied molecular orbitals (HOMOs; see Fig. 3) for **1b–1d** offers an explanation in terms of the entirely different shieldings of B(2) and B(3) when comparing **1b** and **1d** with C_s-symmetric **1c**.

The decent agreement between theory and experiment suggests that MP2/6-311+G(d,p) geometries serve as valid representations of the molecular geometries in solution, which are characterised by very long B–B distances (values are between 1.94 and 1.98 Å) bridged by heteroatoms.

Conclusions

According to these computational efforts, we have concluded that *hypho*-7,8-[S₂B₆H₉][−] (**1c**) is correctly classified as a formal *hypho* structure. However, the other compounds studied here should be considered as *pseudo-hypho* or *pseudo-nido*, i.e. there is a *nido vs. hypho* conflict, and also a disagreement with the literature. Conceivably, according to the *mno* rule all valence electrons of carbon and nitrogen are accounted for in skeletal bonding when assuming *hypho* electron count. On the contrary, we would need to have 10 electron pairs instead of eight to comply with the formal *nido* electron count while also having sufficient electrons to allow for the bonding of carbon and nitrogen cage atoms to *exo* carbon and hydrogens atoms. The fact that *hypho*-[B₈H₈]^{8−} has been shown not to exist complements such a conclusion.

Acknowledgements

D.A.W. thanks the EPSRC for funding a Career Acceleration Fellowship (EP/I004122), and for partially funding, along with the Department of Chemistry, University of York, a studentship for J.P.F.N. We acknowledge the use of the EPSRC National Service for Computational Chemistry Software hosted at Imperial College in carrying out this work, which also made use of the York Advanced Research Computing Cluster (YARCC) provided by the University of York. We acknowledge the Czech Science Foundation (project number P207/11/07505) for financial support. All data supporting this study are provided either in the results section of this paper or as supplementary information accompanying this paper.

References

- See, for example, (a) K. Wade, *Polymer*, 1997, **38**, 4539; (b) J. M. Oliva, D. J. Klein, P. v. R. Schleyer and L. Serano-Andres, *Pure Appl. Chem.*, 2009, **81**, 719; and many other references emerging these days.
- W. N. Lipscomb, *Adv. Inorg. Chem. Radiochem.*, 1950, **1**, 117.



- 3 R. E. Williams, *Adv. Inorg. Chem. Radiochem.*, 1976, **18**, 67.
- 4 (a) K. Wade, *J. Chem. Soc. D*, 1971, 792; (b) K. Wade, *Adv. Inorg. Chem. Radiochem.*, 1976, **18**, 1.
- 5 K. Wade, *Inorg. Nucl. Chem. Lett.*, 1972, **8**, 563.
- 6 D. Hnyk, J. Holub, T. Jelínek, J. Macháček and M. G. S. Londesborough, *Collect. Czech. Chem. Commun.*, 2010, **75**, 1115, and references therein.
- 7 E. Bernhardt, D. J. Brauer, M. Finze and H. Willner, *Angew. Chem., Int. Ed.*, 2007, **46**, 2927.
- 8 See, for example, (a) M. M. Balakrishnarajan and E. D. Jemmis, *J. Am. Chem. Soc.*, 2000, **122**, 4516; (b) E. D. Jemmis, M. M. Balakrishnarajan and P. D. Pancharatna, *J. Am. Chem. Soc.*, 2001, **123**, 4313.
- 9 J. Holub, J. D. Kennedy, T. Jelínek and B. Štíbr, *J. Chem. Soc., Dalton Trans.*, 1994, 1317.
- 10 T. Jelínek, J. D. Kennedy and B. Štíbr, *J. Chem. Soc., Chem. Commun.*, 1993, 1628.
- 11 S. O. Kang and L. G. Sneddon, *J. Am. Chem. Soc.*, 1989, **111**, 3281.
- 12 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, T. Keith, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc., Wallingford, CT, 2013.
- 13 S. Grimme, *J. Chem. Phys.*, 2006, **124**, 034108.
- 14 M. Head-Gordon, J. A. Pople and M. J. Frisch, *Chem. Phys. Lett.*, 1988, **153**, 503.
- 15 S. Saebø and J. Almlöf, *Chem. Phys. Lett.*, 1989, **154**, 83.
- 16 M. J. Frisch, M. Head-Gordon and J. A. Pople, *Chem. Phys. Lett.*, 1990, **166**, 275.
- 17 M. J. Frisch, M. Head-Gordon and J. A. Pople, *Chem. Phys. Lett.*, 1990, **166**, 281.
- 18 M. Head-Gordon and T. Head-Gordon, *Chem. Phys. Lett.*, 1994, **220**, 122.
- 19 H. L. Schmider and A. D. Becke, *J. Chem. Phys.*, 1998, **108**, 9624.
- 20 S. Grimme, *J. Comput. Chem.*, 2006, **27**, 1787.
- 21 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865.
- 22 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1997, **78**, 1396.
- 23 J. P. Perdew, in *Electronic Structure of Solids '91*, ed. P. Ziesche and H. Eschrig, Akademie Verlag, Berlin, 1991, p. 11.
- 24 J. P. Perdew, J. A. Chevary, S. H. Vosko, K. A. Jackson, M. R. Pederson, D. J. Singh and C. Fiolhais, *Phys. Rev. B: Condens. Matter*, 1992, **46**, 6671.
- 25 J. P. Perdew, J. A. Chevary, S. H. Vosko, K. A. Jackson, M. R. Pederson, D. J. Singh and C. Fiolhais, *Phys. Rev. B: Condens. Matter*, 1993, **48**, 4978.
- 26 J. P. Perdew, K. Burke and Y. Wang, *Phys. Rev. B: Condens. Matter*, 1996, **54**, 16533.
- 27 K. Burke, J. P. Perdew and Y. Wang, in *Electronic Density Functional Theory: Recent Progress and New Directions*, ed. J. F. Dobson, G. Vignale and M. P. Das, Plenum, 1998.
- 28 P. Hohenberg and W. Kohn, *Phys. Rev. B*, 1964, **136**, 864.
- 29 W. Kohn and L. J. Sham, *Phys. Rev. A*, 1965, **140**, 1133.
- 30 J. C. Slater, *The Self-Consistent Field for Molecular and Solids, Quantum Theory of Molecular and Solids*, McGraw-Hill, New York, 1974, vol. 4.
- 31 R. Krishnan, J. S. Binkley, R. Seeger and J. A. Pople, *J. Chem. Phys.*, 1980, **72**, 650.
- 32 T. Clark, J. Chandrasekhar and P. v. R. Schleyer, *J. Comput. Chem.*, 1983, **4**, 294.
- 33 http://www.gaussian.com/g_tech/gv5ref/gv5ref_toc.htm.
- 34 (a) R. Ditchfield, *Mol. Phys.*, 1974, **27**, 789; (b) K. Wolinski, J. F. Hinton and P. Pulay, *J. Am. Chem. Soc.*, 1990, **112**, 8251; (c) J. Gauss, *J. Chem. Phys.*, 1993, **99**, 3629.
- 35 W. Kutzelnigg, U. Fleischer and M. Schindler, *NMR Principles and Progress*, Springer, Berlin, 1990.
- 36 Structural data accumulated on the three dicarbadodecaboranes 1,2-, 1,7-, and 1,12-C₂B₁₀H₁₂ are exemplified in the following publications: (a) A. R. Turner, H. E. Robertson, K. B. Borisenko, D. W. H. Rankin and M. A. Fox, *Dalton Trans.*, 2005, 1310; (b) S. Samdal, H. Møllendal, D. Hnyk and J. Holub, *J. Phys. Chem. A*, 2011, **115**, 3380; (c) A. R. Campanelli, A. Domenicano and D. Hnyk, *J. Phys. Chem. A*, 2015, **119**, 205; (d) For the charge distribution in the 1,2-isomer see: D. Hnyk, V. Všetěčka, L. Drož and O. Exner, *Collect. Czech. Chem. Commun.*, 2001, **66**, 1375.
- 37 D. Hnyk, M. Bühl, P. v. R. Schleyer, H. V. Volden, S. Gundersen, J. Müller and P. Paetzold, *Inorg. Chem.*, 1993, **32**, 2442.
- 38 (a) D. Hnyk, E. Vajda, M. Bühl and P. v. R. Schleyer, *Inorg. Chem.*, 1992, **31**, 2464; (b) J. Macháček, J. Plešek, J. Holub, D. Hnyk, V. Všetěčka, I. Císařová, M. Kaupp and B. Štíbr, *Dalton Trans.*, 2006, 1024.
- 39 P. v. R. Schleyer, K. Najafian and A. M. Mebel, *Inorg. Chem.*, 1998, **37**, 6765.
- 40 Decaborane(14), B₁₀H₁₄, is one of the building blocks of boron cluster chemistry.
- 41 D. A. Wann, P. D. Lane, H. E. Robertson, J. Holub and D. Hnyk, *Inorg. Chem.*, 2013, **52**, 4502, and references therein.
- 42 In essence, there exists neutral *nido*-B₆H₁₀ structurally characterised in the solid state, see: F. L. Hirshfeld,



- K. Eriks, R. E. Dickerson, E. L. Lippert Jr. and W. L. Lipscomb, *J. Chem. Phys.*, 1958, **28**, 56.
- 43 M. G. S. Londesborough, Z. Janoušek, B. Štíbr, D. Hnyk, J. Plešek and I. Císařová, *Dalton Trans.*, 2007, 1221.
- 44 T. Jelínek, B. Štíbr, J. D. Kennedy, D. Hnyk, M. Bühl and M. Hofmann, *Dalton Trans.*, 2003, 1326.
- 45 D. Hnyk and E. G. Jayasree, *J. Comput. Chem.*, 2013, **34**, 656, and references therein.
- 46 In *closo*-1-SB₁₁H₁₁ (see ref. 38*b*) and *closo*-1-SB₉H₉, see D. Hnyk, D. A. Wann, J. Holub, S. Samdal and D. W. H. Rankin, *Dalton Trans.*, 2011, **40**, 5734, the difference between the computed and experimental $\delta(^{11}\text{B})$ values of B(12) and B(10) atoms, respectively, atoms that are antipodally coupled with S, using almost the same model chemistries and solvent as for **1d** is also *ca.* 5 ppm; when using C₆D₆ for the second system such a difference even increases.

