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Eu³⁺ and Tb³⁺ doped LaPO₄ nanorods, modified with a luminescent organic compound, exhibiting tunable multicolour emission†

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Co-precipitation reaction followed by hydrothermal treatment were used to synthesise Eu³⁺ or Tb³⁺ doped LaPO₄ nanorods, of 5–10 nm in width and 50–100 nm in length. Surface modification of the as-prepared nanoparticles with a selected luminescent organic compound resulted in formation of hybrid inorganic–organic nanomaterials. The products obtained exhibited tunable multicolour luminescence, dependent on the surface modification and applied excitation wavelength. The colour of their emission can be altered from red-orange to yellow-green. Powder X-ray diffraction (XRD), high resolution transmission electron microscopy (HRTEM) and scanning transmission electron microscopy (STEM) confirmed the structure and morphology of the products synthesized. Successful surface modification of the nanophosphors was evidenced by analytical and spectroscopic techniques such as dynamic light scattering (DLS) – providing size distribution histograms and zeta potentials of the nanoparticles; IR spectroscopy and elemental analysis which proved the presence of an organic phase in the structure; spectrofluorometry (excitation/emission spectra and luminescence decay curves) which confirmed the formation of hybrid, surface modified nanomaterials revealing tunable multicolour emission.

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1. Introduction

Luminescent nanomaterials based on lanthanide ions (Ln³⁺) have fascinated researchers for over the last two decades.^{1–5} The spectroscopic properties of Ln³⁺ ions are unique among the other elements, which result from their electronic configuration [Xe]4f^{*n*} (*n* = 0–14). 4f orbitals of lanthanide ions are shielded by 5p and 6s shells, which makes the 4f–4f electronic transitions nearly insensitive to the coordination environment of Ln³⁺ ion, and provides narrow spectral width of emission or absorption bands as well as long-lived luminescence.⁶ The parity-forbidden character of the 4f–4f transitions results in a very low molar absorption coefficients and also low efficiencies of Ln³⁺ emission, when the direct excitation of 4f–4f absorption bands is applied. However, some Ln³⁺ ions, like Tb³⁺ and Ce³⁺, can be effectively excited *via* the allowed 4f–5d transitions, which strongly enhances the luminescence efficiency.⁷ Also the other,

partially allowed processes like charge transfer (CT) observed in O^{2–} containing host materials and energy transfer (ET), can increase absorption of the excitation light and therefore luminescence intensity.^{8,9} These specific properties of Ln³⁺ ions allow the design of nanomaterials revealing effective luminescence by applying ET and CT phenomena (indirect excitation). Much effort has been made to study ET between, *e.g.*: Ce³⁺ and Tb³⁺, Gd³⁺ and Eu³⁺ or in up-converting systems: Yb³⁺ and Er³⁺, Yb³⁺ and Tb³⁺.^{10–13} Also ET from the host or ligand to Ln³⁺ ion has been extensively investigated.¹⁴ The emission bands of Ln³⁺ ions can be observed in the ultraviolet, visible and near infrared ranges.

The high potential of Ln³⁺-doped nanomaterials is used in many different applications like lighting, phosphors production, organic light emitting diodes (OLEDs and LEDs), lasers, optical amplifiers or waveguides and such areas like medicine and biology.^{4,15–19} Some of these applications result from the strong interest in Ln³⁺-doped hybrid inorganic–organic materials.^{20–22} In general, hybrid materials have altered properties that can be tailored according to the needs, *e.g.* they can show increased mechanical resistance, thermal stability, luminescence efficiency *etc.*^{23–25} Also the multifunctionality and the possibility to modulate their properties are important factors increasing development of such hybrid materials.^{26–31} Such sophisticated bi- or multifunctional hybrid/composite nanomaterials can exhibit simultaneously different desired properties like luminescence and magnetism, which are crucial for

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development of advanced bioimaging, drug delivery, luminescence tracing, document protection *etc.*^{28,29,32,33} One of the areas most recently studied is surface modification of nanocrystals (NCs). Significantly increased stability of colloids, lowered cytotoxicity of NCs and their higher biocompatibility, possibility of NCs use as drug carriers, production of biological sensors and detectors are only a few examples of the numerous applications that result from the inorganic–organic characteristics of hybrid nanomaterials.^{20,34–36}

In this study we report the preparation of hybrid inorganic–organic nanomaterials based on LaPO_4 nanorods doped with Tb^{3+} or Eu^{3+} ions, having organically modified surface. Our aim was to synthesize the nanomaterials exhibiting tunable luminescence dependent on the excitation wavelength used and study their photophysical properties. Such nanocomposites exhibit bright, multicolour luminescence originating from inorganic and organic components. They can be applied as advanced phosphors, luminescence tracers, biomarkers, *etc.* What is more, nanomaterials based on lanthanide phosphates reveal low cytotoxicity,²⁴ and the use of functional organic compound as a surface modifier can alter the surface properties of the nanostructures modified and extend the range of their potential applications.

2. Experimental section

2.1. Materials

Eu_2O_3 , Tb_4O_7 and La_2O_3 (Stanford Materials, 99.99%) were separately dissolved in a concentrated HNO_3 (POCH S.A., ultra-pure) to obtain $\text{Eu}(\text{NO}_3)_3$, $\text{Tb}(\text{NO}_3)_3$ and $\text{La}(\text{NO}_3)_3$ aqueous solutions, respectively. Ammonium phosphate monobasic $\text{NH}_4\text{H}_2\text{PO}_4$ (Sigma-Aldrich, ReagentPlus®, ≥98.5%) was used as a source of phosphate ions. Polyethylene glycol (PEG) 6000 (Alfa Aesar, 98%) was used as a surfactant. The synthetic procedure leading to 2,3-di(3,4-dioctyloxyphenyl)-quinoxaline-6-carboxylic acid starts with the preparation of the appropriate ketone – 3,3,4,4-tetraoctyloxydibenzoyl, which has already been described in literature.³⁷ From this ketone and 3,4-diaminobenzoic acid, final compound was prepared according to the method described by E. J. Foster^{38,39} and H.-J. Chen.⁴⁰ In all experiments, ultra-pure distilled water and absolute ethanol were used.

2.2. Synthesis of nanophosphors – $\text{LaPO}_4\text{:Ln}^{3+}$ 10% ($\text{Ln}^{3+} = \text{Eu}^{3+}$ or Tb^{3+})

The typical synthesis was carried out to get 1.5 g of the final product. The co-precipitation reaction was performed in ambient conditions, using a hot-plate magnetic stirrer. The initial pH of the prepared $\text{RE}(\text{NO}_3)_3$ solutions was around 5. Solution A: $\text{La}(\text{NO}_3)_3$ (11.62 mL, 0.494 M) and $\text{Eu}(\text{NO}_3)_3$ (1.52 mL, 0.419 M) were mixed together at 9/1 molar ratio, and filled with water up to 75 mL. 50 mL of ethanol was added to the as-prepared solution. Subsequently, 0.75 g of PEG was dissolved in this solution. Solution B: 25% molar excess of $\text{NH}_4\text{H}_2\text{PO}_4$ was dissolved in the same solvent systems, with addition of the same amount of PEG. The as-prepared solution B was added dropwise to solution A during 20 minutes. Afterwards the

obtained white precipitate was centrifuged and washed with water and ethanol several times. To get well-crystallised nanomaterial, the as-prepared product was dispersed in 70 mL of water, transferred into a Teflon lined vessel and treated under hydrothermal conditions for 120 minutes at 200 °C and 40 bar (microwave autoclave – ERTEC, Magnum II, 600 W). When the reaction was complete, the purification procedure was repeated. The final product was dried under vacuum overnight. Tb^{3+} doped product was prepared in the same way, using $\text{Tb}(\text{NO}_3)_3$ instead of $\text{Eu}(\text{NO}_3)_3$.

2.3. Surface modification of nanorods

The following organic compound was selected as an exemplary surface modifier because of its complex character, presence of functional groups allowing further chemical modification, potential liquid crystal properties and intensive luminescence.⁴¹ 5 mg of 2,3-di(3,4-dioctyloxyphenyl)-quinoxaline-6-carboxylic acid ($\text{C}_{53}\text{H}_{78}\text{N}_2\text{O}_6$ – for the reader's convenience the simple abbreviation the “organic” will be used throughout the article) dissolved in 10 mL of THF was added slowly to a stirred suspension of nanorods (50 mg) in 25 mL of THF. The reaction mixture was stirred at room temperature for further 72 h and the obtained yellow precipitate was then centrifuged (5 min, 13 000 rpm). In the next step, pure THF was added to the nanomaterial. Afterwards, the mixture was sonicated for 60 s and then centrifuged once more (5 min, 13 000 rpm). The procedure was repeated until no trace of excess of the organic compound was found as determined by TLC. The final, hybrid nanomaterials (yellow powders), revealed tunable multicolour luminescence originating from both inorganic and organic components. Here is worth noting, that there was no observed release of the organic compound from the nanoparticles surface, after washing the product in water. Only a long sonication and washing with THF or toluene caused a slow release of the organic compound. Fig. 1 illustrates a scheme of the nanomaterials surface modification.

2.4. Characterization

Electron microscopy measurements were carried out using transmission electron microscope-TEM Zeiss LIBRA 200FE, operating at 200 kV. Powder XRD (X-ray diffractograms) were recorded on a Bruker AXS D8 Advance diffractometer, using $\text{Cu K}\alpha$ radiation ($\lambda = 1.5406 \text{ \AA}$). The elemental analysis of the products was performed using an Elementar Analyser Vario EL III. Setaram Setsys 1200 device, was used for simultaneous thermogravimetric-differential thermal analysis (TG-DTA), with a heating rate of $5 \text{ }^\circ\text{C min}^{-1}$ in air. IR spectra were recorded on FT-IR spectrophotometer, JASCO 4200. The IR spectra were measured in transmission mode, the samples were mixed with KBr, ground and pressed forming transparent discs. The particle size distribution (hydrodynamic diameter) and zeta (ζ) potential of the nanomaterials synthesized were recorded on Malvern Zetasizer Nano ZS, equipped with dynamic light scattering (DLS) module (He–Ne laser 633 nm, max 4 mW). Before measurements each product was dispersed in MiliQ quality water, forming stable aqueous colloid (0.1 mg mL^{-1}). The



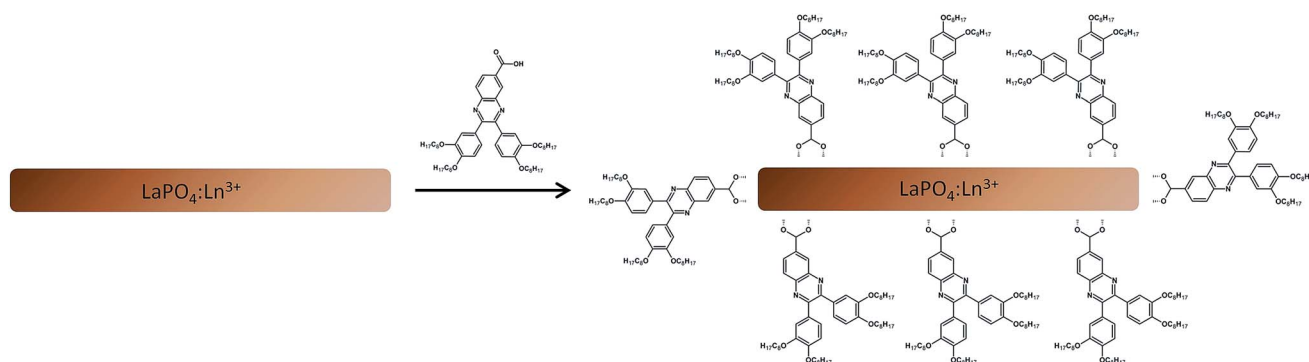


Fig. 1 Scheme of organic surface modification of $\text{LaPO}_4:\text{Ln}^{3+}$ 10% nanoparticles.

excitation/emission spectra and luminescence decay curves of the dried products were measured in ambient conditions, using a Hitachi F-7000 spectrofluorometer. All spectra were appropriately corrected for the apparatus response.

3. Results and discussion

3.1. Structure and morphology

The recorded powder XRD patterns of $\text{LaPO}_4:\text{Tb}^{3+}$ 10% and $\text{LaPO}_4:\text{Eu}^{3+}$ 10% were compared with the pattern from ICDD (International Centre for Diffraction Data) standards database (Fig. 2). Both diffractograms are similar and fit well to that of the hexagonal, hydrated lanthanum phosphate, $\text{LaPO}_4 \cdot 0.5\text{H}_2\text{O}$ (ICDD 000-046-1439). However, the diffractograms reveal also some very small reflexes (overlapping with the background noise), which are probably related to the monoclinic phase (monazite). The obtained XRD patterns exhibit broadened reflexes, which indicate the nanocrystallinity of the products synthesized, and at some extent, it can be due to the superposition of reflexes of hexagonal and monoclinic phase.

Fig. 3 presents electron microscope images of the nano-materials synthesized. TEM images of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (a) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10% (b), HR-TEM image of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% including FFT (Fast Fourier Transform) inset (c), STEM image of $\text{LaPO}_4:\text{Eu}^{3+}$ 10%. TEM images (a and b) reveal numerous

phosphate nanorods similar in shape. The nanorods synthesized are of 5–10 nm in width and 50–100 nm in length, revealing high aspect ratio. The presented HRTEM image of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (c) shows its interplanar distances, namely 0.35 nm (110) and 0.31 nm (200). The observed orientation of the planes and their calculated interplanar distances are consistent with FFT (Fig. 3c inset) of HR-TEM image and XRD analysis. The characteristic hexagonal pattern of the performed FFT undoubtedly confirms the crystal structure of the product obtained. The STEM image provides additional information about the $\text{LaPO}_4:\text{Eu}^{3+}$ 10% morphology (d), confirming the elongated shape of the synthesized nanoparticles.

Surface modification. Surface modification of the prepared inorganic nanophosphors by coating with luminescent organic modifier was carried out to get hybrid nanomaterials, exhibiting tunable multicolour luminescence, dependent on the excitation

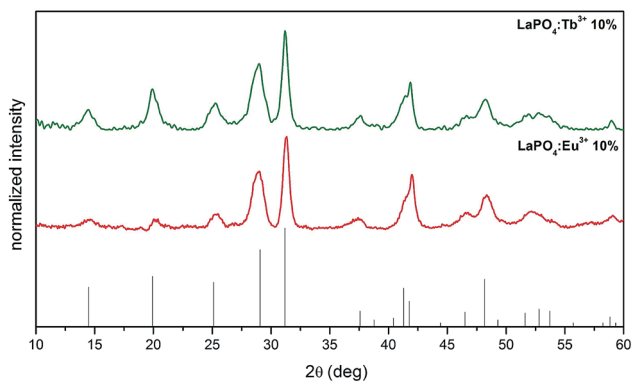


Fig. 2 XRD patterns of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% ($0.5\text{H}_2\text{O}$) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10% ($0.5\text{H}_2\text{O}$).

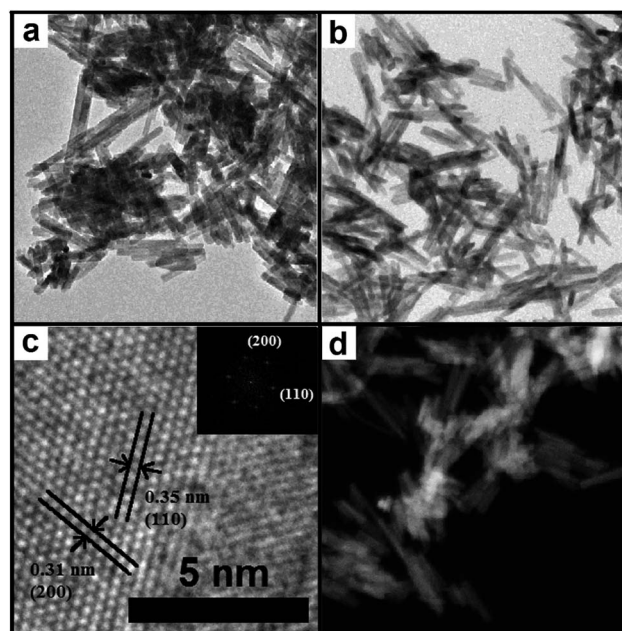


Fig. 3 TEM images of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (a), $\text{LaPO}_4:\text{Tb}^{3+}$ 10% (b); HR-TEM image and its FFT inset for $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (c); STEM image of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (d).



wavelength. The morphology and structure of the nanoparticles obtained, were not affected by the organic surface modification. The presence of the organic modifier molecules on the nanoparticles surface was confirmed and examined by elemental analysis, FT-IR spectroscopy, DLS studies (ζ -potential and particle size distribution) and luminescence spectroscopy.

In order to determine the amount of organic modifiers in the nanomaterials obtained, the elemental analysis of $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%, $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%, $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic was performed. For $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10% product, the content of N, C, H was 0.004, 0.157 and 0.732 (wt%), respectively, whereas for $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10% product, the corresponding contents were 0.005, 0.155 and 0.746 (wt%), respectively. In the products modified with the organic compound ($\text{C}_{53}\text{H}_{78}\text{N}_2\text{O}_6$), the contents of N, C, H for $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic was 0.085, 2.267 and 0.942 (wt%), respectively, whereas for $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic product, the values obtained were 0.087, 2.249 and 0.955 (wt%), respectively. On the basis of the results, the molar concentration of the organic compound bound to the surface is 0.0332 mmol (2.79 wt%) and 0.0329 mmol (2.76 wt%) per one gram of $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic products, respectively. The results presented confirm a successful functionalization of the nanomaterials surface. The excessive hydrogen content in their structure was related to the adsorbed water molecules.

Additionally, TG-DTA measurements were performed to investigate the stability of the organic surface layer and confirm the composition of the nanomaterials synthesized. The results obtained agree well with the elemental analysis data, and they are presented in full in ESI (Fig. S1†).

Fig. 4 illustrates the IR spectra of $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%, $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%, $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic, $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic, and pure organic compound. All spectra recorded reveal broad absorption peaks around 3400 cm^{-1} and 1640 cm^{-1} , corresponding to the O–H stretching (ν) and deformation (σ) vibrations, respectively. The observed O–H bonds

correspond to water molecules adsorbed on the nanorods surface and structural water molecules (hydration of phosphates). The absorption peaks around 2924 and 2865 cm^{-1} are related to $\nu\text{C-H}$ vibrations of $-\text{CH}_2$ groups of organic compound molecules (adsorbed on the surface of phosphate nanorods). The spectra of the pure and modified phosphates exhibit very intensive and broad peaks assigned to the vibrations of phosphate groups. The peaks around 1050 , 950 cm^{-1} were assigned to the stretching vibrations within PO_4 groups, and around 615 , 542 cm^{-1} to the bending vibrations within these groups.^{42,43} What is more, the quite intensive band observed at $\sim 960\text{ cm}^{-1}$ confirms the coexistence of monoclinic phosphate (minor phase) together with the hexagonal lanthanum phosphate (major phase), which was mentioned during discussion of the XRD patterns. The peaks below 1700 cm^{-1} in the spectrum of the pure organic compound, correspond to the vibrations of numerous bonds in the very complex structure of this compound. The presence of peaks around 2900 and 1200 cm^{-1} in the spectra of the organically modified phosphates ($\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic), confirms their successful modification with the organic compound used.

The ζ -potential and average particle size distribution of the nanomaterials synthesized were measured by the DLS method. All of the measurement were carried out at $\text{pH} = 7$. The ζ -potential recorded for $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10% was $+29.3\text{ mV}$, and for $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10% it was $+34.6\text{ mV}$. The ζ -potential recorded for the surface modified $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic nanorods decreased to $+17.8\text{ mV}$ and $+27.3\text{ mV}$, respectively. The decreased ζ -potential values confirmed surface modification of the nanomaterials synthesized. What is more, all of the nanomaterials exhibited relatively high surface charge, additionally confirming their stability at neutral pH, which is important in potential bioapplications. The average particle size distribution of the products synthesized is presented in Fig. 5. The approximate sizes of $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10% (a) and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10% (b) particles are about 300 nm , and increase after surface modification for $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic (c) and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic (d) to about 500 nm . This fact clearly confirms surface alterations of the nanomaterials modified, manifested by increased hydrodynamic diameter of the nanoparticles. However, the recorded DLS curves revealed the polydispersity of the nanomaterials obtained, which was caused by particles agglomeration and their “sticking” to bigger clusters after surface modification. The hydrodynamic radius/diameter of the particles analysed is usually larger in comparison to the real particle sizes from TEM data, since the DLS method takes into account surface solvation of the particles, and their agglomeration/aggregation in the colloidal solution. One must remember, that the data presented are only a rough approximation of the nanoparticles sizes because of their highly anisotropic shape (DLS size measurements assume the spherical shape of the analysed objects).

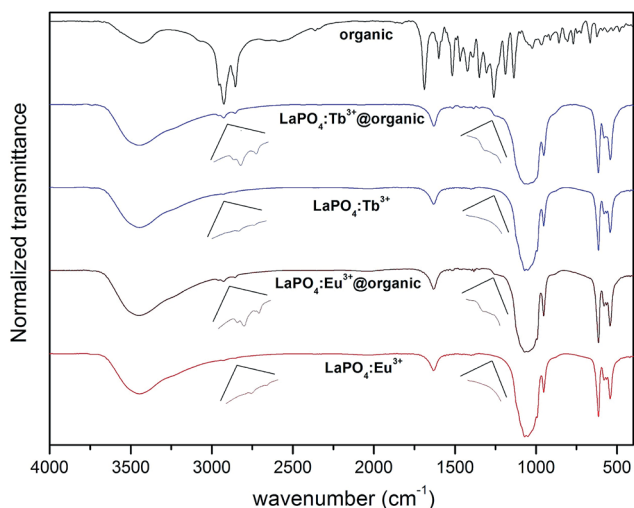


Fig. 4 IR spectra of pure organic compound (surface modifier), unmodified $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10% and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%, surface modified $\text{LaPO}_4\cdot\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4\cdot\text{Tb}^{3+}$ 10%@organic.

3.2. Luminescent properties

Fig. 6–8 show the spectroscopic properties of Eu^{3+} and Tb^{3+} doped phosphate nanorods and the corresponding surface



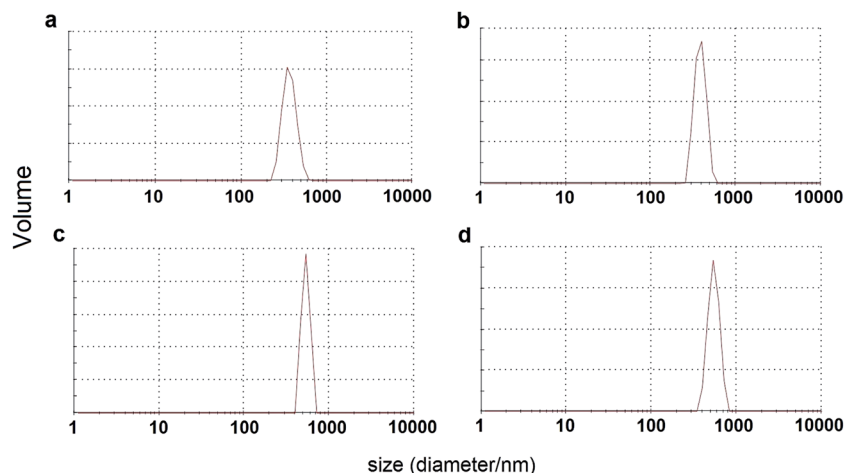


Fig. 5 Particle size distribution for $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (a) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10% (b) $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic (c) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic (d), measured by DLS method.

modified products, namely $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic. All spectra were recorded for solid products (dried powders). Fig. 6a presents three excitation spectra of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% ($\lambda_{\text{em}} = 620$ nm) and $\text{LaPO}_4:\text{Eu}^{3+}$ 10%

@organic ($\lambda_{\text{em}} = 500, 620$ nm). For bare $\text{LaPO}_4:\text{Eu}^{3+}$ 10% ($\lambda_{\text{em}} = 620$ nm – the maximum of $^5\text{D}_0 \rightarrow ^7\text{F}_2$ transition in Eu^{3+} ion), a dominant broad band centred at 256 nm related to $\text{O}^{2-} \rightarrow \text{Eu}^{3+}$ charge transfer (CT) transition can be observed. This band is the

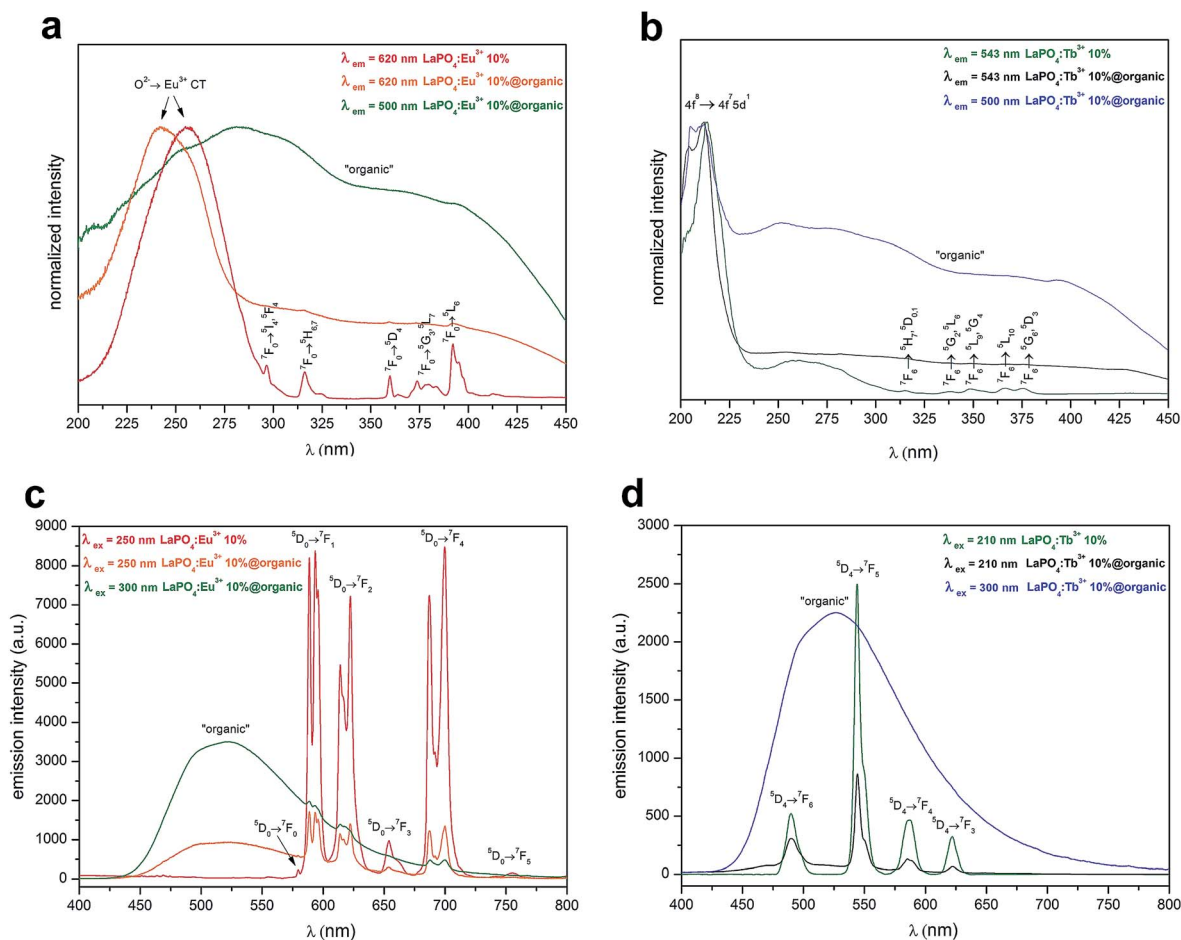


Fig. 6 Excitation spectra of $\text{LaPO}_4:\text{Eu}^{3+}$ 10%, $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic (a) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%, $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic (b); emission spectra of $\text{LaPO}_4:\text{Eu}^{3+}$ 10%, $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic (c) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%, $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic (d).

most intensive one in the whole spectrum because the mentioned CT transition is allowed by selection rules, in contrast to the forbidden 4f–4f transitions within Eu^{3+} ion (the series of narrow bands observed at 300–400 nm). For the surface modified $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic nanomaterial, the same dominant band can be observed ($\lambda_{\text{em}} = 620$ nm). The series of intrinsic 4f–4f transitions is hardly visible in this spectrum, because they overlap with the appearing absorption band of the organic surface modifier. The third plot obtained for the same modified nanomaterial, whose excitation spectrum was recorded at $\lambda_{\text{em}} = 500$ nm (the maximum of emission band for the organic compound) does not reveal the characteristic bands related to the transitions in Eu^{3+} ions. In this spectrum only a very broad absorption band centred at 282 nm corresponding to the surface organic compound can be observed in the whole presented wavelength range.

Fig. 6b shows three excitation spectra of $\text{LaPO}_4:\text{Tb}^{3+}$ 10% ($\lambda_{\text{em}} = 543$ nm) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic ($\lambda_{\text{em}} = 500, 543$ nm). The spectrum of $\text{LaPO}_4:\text{Tb}^{3+}$ 10% ($\lambda_{\text{em}} = 543$ nm is the maximum of the most intensive in the emission spectrum $^5\text{D}_4 \rightarrow ^7\text{F}_5$ transition, within Tb^{3+} ions), shows a dominant broad band centred at 213 nm corresponding to $4\text{f}^6 \rightarrow 4\text{f}^75\text{d}^1$ (f–d) allowed transition. The bands at higher wavelengths are assigned to the forbidden 4f–4f transitions in Tb^{3+} ion. In the spectrum of $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic recorded at the same $\lambda_{\text{em}} = 543$ nm, the slightly shifted f–d transition is also dominant. The other 4f–4f transitions are not visible. The last excitation spectrum of this modified nanomaterial ($\lambda_{\text{em}} = 500$ nm) also reveals the dominant band related to f–d transition in Tb^{3+} ion. However, the very broad absorption band (with maximum ranging from 250 to 280 nm) assigned to the surface organic modifier, can be observed in the whole spectrum, as well.

Fig. 6c presents three emission spectra of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% ($\lambda_{\text{ex}} = 250$ nm) and $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic ($\lambda_{\text{ex}} = 250, 300$ nm). In the first spectrum of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% recorded at $\lambda_{\text{ex}} =$

250 nm (the position of the most intensive CT transition in the excitation spectrum), six narrow, split bands corresponding to the $^5\text{D}_0 \rightarrow ^7\text{F}_J$ ($J = 0-5$) transitions can be observed. The hypersensitive electric dipole $^5\text{D}_0 \rightarrow ^7\text{F}_2$ transition is sensitive to the site symmetry alterations.^{44,45} The ratio between integrated areas of the $^5\text{D}_0 \rightarrow ^7\text{F}_2$ and $^5\text{D}_0 \rightarrow ^7\text{F}_1$ transition bands is informative about the presence of symmetry centre in the site occupied by the Eu^{3+} ions. The values calculated both for $\text{LaPO}_4:\text{Eu}^{3+}$ 10% and $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic are close to 1 (1.04 and 1.01 respectively). The ratio higher than 1 indicates that the Eu^{3+} ions are situated at sites without inversion symmetry. However, in the $\text{LaPO}_4:\text{Eu}^{3+}$ material the Eu^{3+} ions occupy sites with D_2 symmetry in the LaPO_4 structure which is non-centrosymmetric and the presence of an inversion centre cannot be assumed.^{46–48}

The product exhibits an intense, bright red luminescence. When the surface modified nanomaterial was excited at the same wavelength, the intensity of the mentioned transitions decreased, and a new broad band appeared in the range of 500–550 nm. This band corresponds to the emission of the organic compound. As a consequence of these alternations in the spectrum shape, the observed luminescence of the product was tuned to yellowish emission. Upon exciting the organic modified nanomaterial at 300 nm (absorption range of the organic compound), the characteristic bands of Eu^{3+} ions can hardly be observed, in contrast to the very high intensity and broad emission band of the organic compound. The resulting emission of the product is green.

Fig. 6d shows three emission spectra of $\text{LaPO}_4:\text{Tb}^{3+}$ 10% ($\lambda_{\text{ex}} = 210$ nm) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic ($\lambda_{\text{ex}} = 210, 300$ nm). The spectrum of $\text{LaPO}_4:\text{Tb}^{3+}$ 10%, recorded at $\lambda_{\text{ex}} = 210$ nm (the position of the most intensive transition in the excitation spectrum), presents four narrow bands assigned to the $^5\text{D}_4 \rightarrow ^7\text{F}_J$ ($J = 6-3$) transitions, characteristic of Tb^{3+} ions.⁷ The product exhibits bright green luminescence. The spectrum of the $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic nanophosphor excited at the same wavelength reveals four bands typical of Tb^{3+} ions, as well.

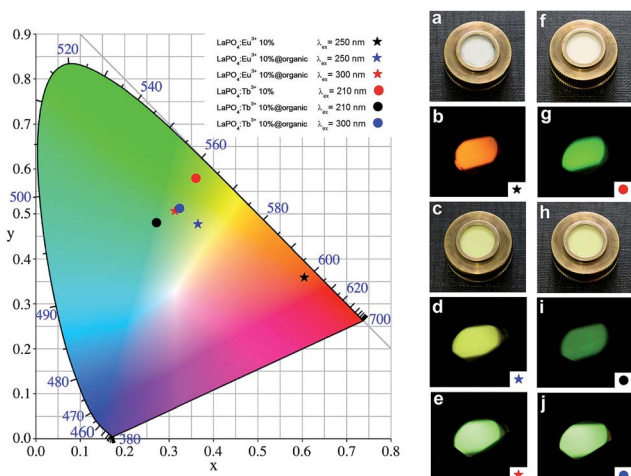


Fig. 7 Chromaticity diagram and photographs of $\text{LaPO}_4:\text{Eu}^{3+}$ 10% (a and b), $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic (c–e), $\text{LaPO}_4:\text{Tb}^{3+}$ 10% (f and g) and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic (h–j), taken in daylight (a, c, f and h) and under UV light irradiation ($\lambda_{\text{ex}} = 210$ nm – g and i; $\lambda_{\text{ex}} = 250$ nm – b and d; $\lambda_{\text{ex}} = 300$ nm – e and j).

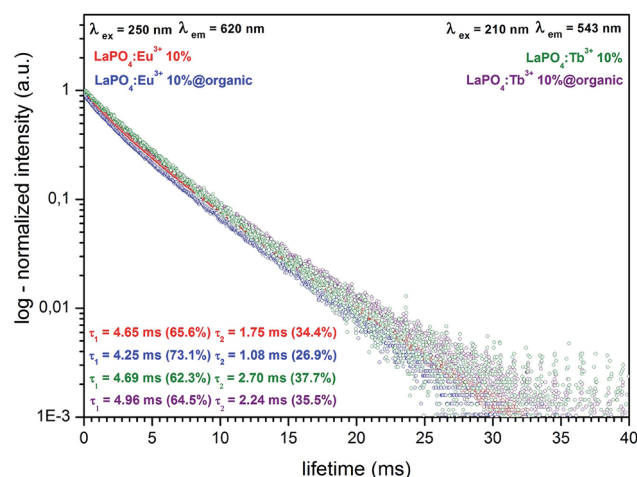


Fig. 8 Luminescence decay curves of $\text{LaPO}_4:\text{Eu}^{3+}$ 10%, $\text{LaPO}_4:\text{Eu}^{3+}$ 10%@organic and $\text{LaPO}_4:\text{Tb}^{3+}$ 10%, $\text{LaPO}_4:\text{Tb}^{3+}$ 10%@organic.



However, the intensity of these bands decreased in comparison to that in the spectrum of the unmodified product. Besides these bands, less intensive bands around 500–550 nm, corresponding to the organic modifier can also be observed. The colour of the product emission is still green, however shifted towards blue. The spectrum of the modified nanomaterial excited at 300 nm reveals only the intensive, broad band assigned to the organic compound. Its luminescence is greenish, namely the colour of emission is slightly altered when compared to that of the unmodified nanophosphor.

For more detailed specification of the luminescence colours one can refer to the included chromaticity diagram (CIE 1964 10 deg observer), presented in Fig. 7. The photographs of the products synthesized, taken in daylight and under UV light, showing their multicolour emission, are included in Fig. 7, as well.

Fig. 8 presents the luminescence decay curves and calculated radiative lifetimes for the $^5D_0 \rightarrow ^7F_2$ and $^5D_4 \rightarrow ^7F_5$ transitions of Eu^{3+} and Tb^{3+} ions, respectively. All data were recorded at 293 K; $\lambda_{\text{em}} = 620 \text{ nm}$, $\lambda_{\text{ex}} = 250 \text{ nm}$ for the Eu^{3+} doped compounds and $\lambda_{\text{em}} = 543 \text{ nm}$ and $\lambda_{\text{ex}} = 210 \text{ nm}$, for the Tb^{3+} doped compounds. In hexagonal lanthanum phosphate, all lanthanide ions should be at the sites of the same type (coordination environment).²⁴ However, the experimental profiles were successfully fitted to the biexponential function of decay, namely $y = A_1 \exp(-x/\tau_1) + A_2 \exp(-x/\tau_2) + y_0$. The nanomaterials synthesized exhibit high surface-to-volume ratio. Therefore a large part of ions forming the material is placed on or near the surface of nanocrystals. Hence, after surface modification with the organic compound used, a significant number of the surface/near surface ions were localised in a new coordination environment (altered local site symmetry). The reason for this phenomenon was a strong coordination/binding of the organic molecules to the nanoparticles surface. This is why, the nanophosphors obtained exhibit a shorter second component of luminescence decay. Here is worth noting, that the presence of a small amount of monoclinic LaPO_4 in the sample can affect the lifetime components and disturb their decay profiles. However, because of the large contribution ($\approx 30\%$) of the second lifetime components and their significant shortening after the surface modification, we assume that the discussed biexponential character of the luminescence decay is predominantly caused by the differently emitting surface ions. The calculated luminescence lifetimes for the products synthesized are in the range of 4.25–4.96 ms (τ_1) and 1.08–2.70 ms (τ_2). The detailed values are presented in Fig. 8. Such relatively long radiative lifetimes are in line with literature data for lanthanide doped inorganic phosphors.^{49,50} The observed lifetimes for Eu^{3+} and Tb^{3+} doped products are generally similar, however the lifetimes of Eu^{3+} ions are slightly shorter when compared to those of Tb^{3+} ions. Analysis of the decay profiles leads to a conclusion that the modified nanomaterials exhibit a shorter average lifetime, in comparison to their unmodified analogues. The lifetime shortening is particularly pronounced in the second lifetime component (τ_2), assigned to the surface ions (shortening from 1.75 to 1.08 ms and from 2.70 to 2.24 ms for Eu^{3+} and Tb^{3+} ions, respectively). This phenomenon can be

explained by a strong interaction between surface ions and organic molecules attached to the nanoparticles surface, resulting in enhanced luminescence quenching. The results obtained are in agreement with the data on the emission decrease of the modified nanorods (see the emission spectra in Fig. 6a and b).

4. Conclusions

The highly luminescent, crystalline nanomaterials doped with Eu^{3+} and Tb^{3+} ions were synthesized *via* the co-precipitation approach followed by hydrothermal treatment. The nanomaterials formed were in the form of elongated nanorods (5–10 nm in width and 50–100 nm in length) composed of hexagonal $\text{LaPO}_4 \cdot 0.5\text{H}_2\text{O}$. Subsequently, the products obtained were modified with a luminescent organic compound. The surface modification resulted in a formation of hybrid inorganic–organic nanomaterials, which exhibited tunable and multimodal luminescence. The products emission could be tuned from red-orange to yellow-green luminescence. Successful modification of the surface of nanocrystals was checked by DLS, IR spectroscopy, elemental analysis, TG-DTA and spectrofluorimetry. These novel, functional nanomaterials can be applied in luminescence tracing, detection techniques, multicolour imaging, as novel light sources and in many other special applications requiring sophisticated, hybrid nanomaterials exhibiting tunable emission. The products synthesized can be also used in biomedical applications requiring multifunctionality of nanomaterials.

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Notes and references

- 1 R. Bhargava, *J. Lumin.*, 1996, **70**, 85–94.
- 2 S. Heer, O. Lehmann, M. Haase and H.-U. Güdel, *Angew. Chem., Int. Ed.*, 2003, **42**, 3179–3182.
- 3 Z.-G. Yan and C. Yan, *J. Mater. Chem.*, 2008, **18**, 5046–5059.
- 4 W. Feng, C. Han and F. Li, *Adv. Mater.*, 2013, 1–17.
- 5 Q. Lü, A. Li, F. Guo, L. Sun and L. Zhao, *Nanotechnology*, 2008, **19**, 145701.
- 6 P. A. Tanner, *Chem. Soc. Rev.*, 2013, **42**, 5090–5101.



- 7 T. Grzyb, M. Runowski, A. Szczeszak and S. Lis, *J. Phys. Chem. C*, 2012, **116**, 17188–17196.
- 8 T. Grzyb, R. J. Wiglusz, V. Nagirnyj, A. Kotlov and S. Lis, *Dalton Trans.*, 2014, **43**, 6925–6934.
- 9 F. Wang, X. Fan, M. Wang and Y. Zhang, *Nanotechnology*, 2007, **18**, 25701–25706.
- 10 T. Grzyb, *RSC Adv.*, 2014, **4**, 2590–2595.
- 11 T. Grzyb, M. Runowski, K. Dąbrowska, M. Giersig and S. Lis, *J. Nanopart. Res.*, 2013, **15**, 1958–1972.
- 12 C. Cao, S. Guo, B. K. Moon, B. C. Choi and J. H. Jeong, *Solid State Sci.*, 2013, **19**, 99–103.
- 13 F. Auzel, *Chem. Rev.*, 2004, **104**, 139–173.
- 14 S. Lis, M. Elbanowski, B. Mąkowska and Z. Hnatejko, *J. Photochem. Photobiol. A*, 2002, **150**, 233–247.
- 15 S. Wang, J. Feng, S. Song and H. Zhang, *CrystEngComm*, 2013, **15**, 7142–7151.
- 16 C. C. Lin and R.-S. Liu, *J. Phys. Chem. Lett.*, 2011, **2**, 1268–1277.
- 17 Q. Y. Zhang and X. Y. Huang, *Prog. Mater. Sci.*, 2010, **55**, 353–427.
- 18 S. V. Eliseeva and J.-C. G. Bünzli, *Chem. Soc. Rev.*, 2010, **39**, 189–227.
- 19 G. Boulon, *Opt. Mater.*, 2012, **34**, 499–512.
- 20 K. Binnemans, *Chem. Rev.*, 2009, **109**, 4283–4374.
- 21 P. Escrivano, B. Julián-López, J. Planelles-Aragó, E. Cordoncillo, B. Viana and C. Sanchez, *J. Mater. Chem.*, 2008, **18**, 23–40.
- 22 L. D. Carlos, R. A. S. Ferreira, V. de Zea Bermudez, B. Julián-López and P. Escrivano, *Chem. Soc. Rev.*, 2011, **40**, 536–549.
- 23 J.-N. Park, P. Zhang, Y.-S. Hu and E. W. McFarland, *Nanotechnology*, 2010, **21**, 225708.
- 24 M. Runowski, K. Dąbrowska, T. Grzyb, P. Miernikiewicz and S. Lis, *J. Nanopart. Res.*, 2013, **15**, 2068.
- 25 E. He, H. Zheng, J. Dong, W. Gao, Q. Han, J. Li, L. Hui, Y. Lu and H. Tian, *Nanotechnology*, 2014, **25**, 045603–045611.
- 26 S. A. Corr, Y. P. Rakovich and Y. K. Gun'ko, *Nanoscale Res. Lett.*, 2008, **3**, 87–104.
- 27 S. T. Selvan, T. T. Y. Tan, D. K. Yi and N. R. Jana, *Langmuir*, 2009, **26**, 11631–11641.
- 28 S. X. C. Mi, J. Zhang, H. Gao, X. Wu, M. Wang, Y. Wu, Y. Di, Z. Xu and C. Mao, *Nanoscale*, 2010, **2**, 1141–1148.
- 29 M. Runowski, T. Grzyb and S. Lis, *J. Nanopart. Res.*, 2012, **14**, 1188.
- 30 B. Lu, Z. Zhang, J. Hao and J. Tang, *RSC Adv.*, 2014, **4**, 21909.
- 31 Y. He, S. Vasiraju and L. Que, *RSC Adv.*, 2014, **4**, 2433.
- 32 M. Liong, J. Lu, M. Kovichich, T. Xia, S. G. Ruehm, A. E. Nel, F. Tamanoi and J. I. Zink, *ACS Nano*, 2008, **2**, 889–896.
- 33 P. Kulpinski, M. Namyslak, T. Grzyb and S. Lis, *Cellulose*, 2012, **19**, 1271–1278.
- 34 J. Wang, J. Hu, D. Tang, X. Liu and Z. Zhen, *J. Mater. Chem.*, 2007, **17**, 1597–1601.
- 35 L. Lou, K. Yu, Z. Zhang, B. Li, J. Zhu, Y. Wang, R. Huang and Z. Zhu, *Nanoscale*, 2011, **3**, 2315–2323.
- 36 M. Runowski, A. Ekner-Grzyb, L. Mrówczyńska, S. Balabhadra, T. Grzyb, J. Paczesny, A. Zep and S. Lis, *Langmuir*, 2014, **30**, 9533–9543.
- 37 B. Mohr, V. Enkelmann and G. Wegner, *J. Org. Chem.*, 1994, **59**, 635–638.
- 38 E. J. Foster, C. Lavigueur, Y.-C. Ke and V. E. Williams, *J. Mater. Chem.*, 2005, **15**, 4062–4068.
- 39 E. J. Foster, R. B. Jones, C. Lavigueur and V. E. Williams, *J. Am. Chem. Soc.*, 2006, **128**, 8569–8574.
- 40 C.-J. Chen, Y.-C. Wu, H.-S. Sheu, G.-H. Lee and C. K. Lai, *Tetrahedron*, 2011, **67**, 114–124.
- 41 J. Szydłowska, P. Krzyczkowska, M. Salamończyk, E. Górecka, D. Pociecha, B. Maranowski and A. Krówczyński, *J. Mater. Chem. C*, 2013, **1**, 6883–6889.
- 42 S. Lucas, E. Champion, D. Bregiroux, D. Bernache-Assollant and F. Audubert, *J. Solid State Chem.*, 2004, **177**, 1302–1311.
- 43 Z. Xu, Y. Cao, C. Li, P. Ma, X. Zhai, S. Huang, X. Kang, M. Shang, D. Yang, Y. Dai and J. Lin, *J. Mater. Chem.*, 2011, **21**, 3686–3694.
- 44 B. R. Judd, *J. Chem. Phys.*, 1966, **44**, 839.
- 45 M. Runowski, S. Balabhadra and S. Lis, *J. Rare Earths*, 2014, **32**, 242–247.
- 46 T. Anfimova, Q. Li, J. O. Jensen, and N. J. Bjerrum, *Int. J. Electrochem. Sci.* 2014, **9**, 2285–2300.
- 47 J. a. Diaz-Guillén, A. F. Fuentes, S. Gallini and M. T. Colomer, *J. Alloys Compd.*, 2007, **427**, 87–93.
- 48 K. Binnemans and C. Görller-Walrand, *J. Rare Earths*, 1996, **14**, 173–180.
- 49 I. Omkaram, G. Seeta Rama Raju and S. Buddhudu, *J. Phys. Chem. Solids*, 2008, **69**, 2066–2069.
- 50 M. Runowski and S. Lis, *J. Alloys Compd.*, 2014, **597**, 63–71.

