In this paper we report the preparation and characterization of Sr$_{1-x}$Na$_x$SiO$_3$ \( \cdot \) 0.5$\alpha$ samples, recently proposed as oxide ion conductors. We show that Na-doping unlikely takes place in the silicate phase, and that a secondary glassy phase is at the origin of the transport properties, thereby suggesting that the conductivity is due only to a limited extent to oxide ion migration in the crystalline system.

Research efforts devoted to the improvement of material properties in the field of solid oxide fuel cells (SOFCs) are extremely active: new compositions as well as ways of improving the performances of existing materials are constantly explored.\(^1\) One of the major challenges in this area is lowering the operating temperature to the intermediate temperature regime (\(<700{\,}^\circ\text{C}\)) while maintaining good electrochemical performances. In this picture, the solid electrolyte plays a central role since its ionic conductivity determines the best operating temperature range. Novel compounds include oxygen-deficient materials possessing the perovskite crystal structure,\(^2\) and interstitial-oxygen materials limited extent to oxide ion migration in the crystalline system.

Recently, the Sr$_{1-\alpha}$K$_{\alpha}$MO$_{3-\beta}$ (with M = Si, Ge) family was proposed as a highly promising novel class of electrolyte materials showing good conductivity values \((\sigma \geq 10^{-2}\;\text{S cm}^{-1})\) at intermediate temperatures.\(^4\) K-doped samples in this system were subsequently shown to be highly hygroscopic at room temperature, while Na-doping was proved to be highly effective in introducing purely oxide ion conduction in the system, thanks also to the large range of solubility of Na ions on the Sr site.\(^5,6\) Such a solubility, extending up to the composition Sr$_{0.05}$Na$_{0.45}$SiO$_{2.775}$, gives origin to a remarkable oxide ion conductivity \((\sigma \geq 10^{-2}\;\text{S cm}^{-1})\) at \(T < 500{\,}^\circ\text{C}\) and is associated with a very low activation energy for oxide ion migration of about 0.3 eV; the stability range of the ionic conductivity as a function of the oxygen partial pressure was shown to extend down to 10\(^{-3}\) atm.\(^6\) Based on such values, the Sr$_{0.33}$Na$_{0.48}$SiO$_{2.775}$ composition represents the best performing electrolyte material for SOFC currently known for temperatures lower than 650 \(^{\circ}\text{C}.\(^6\)

The parent SrSiO$_3$ compound is constituted by layers of Sr\(^{2+}\) ions spaced by SiO$_4$ clusters, where each SiO$_4$ unit is linked through two corner oxygen ions to adjacent tetrahedra forming isolated three-fold rings. Na substitution on the Sr site will originate, for charge compensation, oxygen vacancies responsible for the oxide ion conductivity in the system, according to a mechanism still not completely clear at the moment, as both vacancy and interstitial-mediated mechanisms were discussed.\(^4,5\)

A neutron powder diffraction study on Na and K-doped samples excluded the presence of interstitial oxygen ions in both the systems and suggested that oxide ion vacancies, consistent with the nominal compositions, are present at r.t. and 400 \(^{\circ}\text{C}\) for Sr$_{0.4}$Na$_{0.4}$SiO$_{2.8}$, while a larger oxygen under-stoichiometry is expected at 800 \(^{\circ}\text{C}.\(^7\)

In this work we present a detailed study of the structural and transport properties of the Sr$_{1-x}$Na$_x$SiO$_3$ \( \cdot \) 0.5$\alpha$ series \((x = 0.05, 0.25, 0.45)\). We propose that the complexity of the SiO$_2$-Na$_2$O-SrO phase diagram must be taken into consideration, as the actual solubility degree of Na in the SrSiO$_3$ phase may be significantly lower than expected due to the possibility of formation of secondary, spurious and metastable glassy phases. We show that such secondary phases may be at the origin of the promising transport properties for this class of materials, therefore suggesting the possibility that the high conductivity values previously reported were not significantly due to oxide ion migration in the crystalline system.

Powder samples of nominal composition Sr$_{1-x}$Na$_x$SiO$_3$ \( \cdot \) 0.5$\alpha$ \((x = 0.05, 0.25, 0.45)\) were prepared by the solid state reaction according to the procedure described by P. Singh \textit{et al.}, at a maximum synthesis temperature of 1050 \(^{\circ}\text{C}.\(^7\) Samples were slowly cooled down to room temperature (1 \(^{\circ}\text{C}\) min\(^{-1}\)). These samples will be named Na$_{0.05}$, Na$_{0.25}$ and Na$_{0.45}$, respectively. One sample of nominal composition Sr$_{0.53}$Na$_{0.45}$SiO$_{2.73}$, was subsequently retreated at 650 \(^{\circ}\text{C}\) for 12 hours and slowly cooled down to room temperature (named Na$_{0.45}$-RC in the following).
Room temperature X-ray powder diffraction (XRPD) patterns of all the prepared samples were acquired on a Bruker D8-Advance diffractometer with Cu Kα radiation in the 2θ range 10–100 degrees, with fixed steps of 0.02 degrees and counting time per steps of 10 s. Fig. 1 shows the XRPD patterns of the Sr1−xNa2Si2O5−x series prepared at 1050 °C. All the patterns can be indexed in the monoclinic C12/c1 space group (no. 15). No spurious reflections can be identified for Na0.05, Na0.25 and Na0.45, and the patterns are consistent with the structure of the parent SrSiO3 compound.1 Sample Na0.45-RC clearly presents additional peaks, mainly compatible with the α polymorph of Na2Si2O5 (see below).

Rietveld refinement9 was performed using the FullProf software.10 Table 1 shows the Rietveld refined cell parameters and unit cell volume obtained from the XRPD measurements for the Sr1−xNa2Si2O5−x series prepared at 1050 °C. No clear trend is observed for the evolution of the unit cell volume along with doping. This result suggests that the substitution of Na ions for Sr might not be complete and the actual stoichiometry of the crystalline phase could therefore be different from the nominal one.

In order to assess this aspect, room temperature neutron powder diffraction (NPD) measurements were acquired on the HRPT instrument at the Paul Scherrer Institute (PSI, Villigen) in the 2θ range 10–100 degrees. Rietveld refinement allowed in this case to extract the occupancies of Na ions on the Sr site, as well as to estimate the degree of possible oxygen under-stoichiometry. Fig. 2 shows the Rietveld refined pattern for the sample of nominal composition Sr0.75Na0.25SiO3. Rietveld refinements and the structural parameters for all the samples are reported as Fig. S1 and Table S1, respectively (ESI†). Analysis of the refined parameters (Table S1, ESI†) confirms that this series does not obey the Vegard’s law for solid solutions. If allowed to change in the refinement, the Na occupancy on the Sr sites tends to zero (and the Sr site keeps the full occupancy), while the oxygen sites tend to full occupancy, so that the composition derived from the analysis of the neutron diffraction data is in agreement, within the experimental errors, with nominal SrSiO3 composition for all the three samples.

Differential Scanning Calorimetry (DSC) measurements were performed from room temperature up to 600 °C with a scan rate of 20 °C min−1 in N2 flux on the Na0.45 sample, and showed a sigmoidal feature suggestive of a glass transition just below 500 °C (Fig. S2, ESI†). Following such an indication, the sample Na0.45 was subsequently retreated at 650 °C for 12 hours and slowly cooled down to room temperature (sample Na0.45-RC). The XRD pattern of this sample in Fig. 1 clearly shows the appearance of additional peaks compared to those ascribed to the crystalline SrSiO3 phase.

The most intense peaks among such reflections can be attributed to α-Na2Si2O5. Fig. S3 (ESI†) shows the two-phase Rietveld refined NPD pattern of this sample considering SrSiO3 and α-Na2Si2O5. A good agreement between the calculated and the experimental pattern is found by considering these two phases, with minor contributions possibly related to different Na2Si2O5 polymorphs or other secondary phases still unexplained.

Solid-state MAS NMR measurements were acquired on a 9.4 T magnet using a 7.0 mm probe.29Si spectra were recorded at a spinning rate of 5 kHz, at ambient temperature with a single pulse experiment (3.00 μs pulse), a delay of 300 seconds and 4k scans. The spectra were referred to TMS.

In Fig. 3 the 29Si spectra of samples Na0.05 (d), Na0.25 (c) and Na0.45 (b) are shown, together with the 29Si spectra of the samples Na0.45-RC (a) and of a properly prepared Na2Si2O5 glassy sample (e). The spectrum of the Na0.05 sample presents a single sharp peak centered at −86 ppm. With the introduction of higher Na content in the structure, a new broad peak centered at −89 ppm appears; the relative intensity of the broad peak with respect to that of the signal at −86 ppm grows along with the increase of the Na content. Fig. 3a reports the spectrum of the sample Sr0.55Na0.45SiO3.
recrystallized at 650 °C (Na0.45-RC). For this sample, the broad signal disappears while a new sharp resonance at −95.7 ppm is observed. Other weak signals appear at −87.7, −89.6 and −91.9 ppm. We recall that for this sample, recrystallized at 650 °C, also the XRD data show additional peaks (Fig. 1) compared to the starting sample Na0.45.

The peak at −86 ppm observed in each sample is related to the crystalline SrSiO3-based phase.11,12 The position and the broadening of this signal do not seem to be affected by the nominal composition of the samples, suggesting the formation of pure SrSiO3, independently of the increasing amount of Na present in the starting reagent mixture. The solubility of Na in the SrSiO3 phase thus seems to be extremely low, in agreement with the NPD analysis. This hypothesis is also supported by the appearance of the broad peak at −89 ppm with the introduction of Na in the system, indicative of the formation of a new disordered phase (see below and Fig. 3e) whose content increases as the nominal Na content in the sample does. This second phase has been identified with glassy Na2Si2O5.14 This graph shows that a considerable drop in conductivity for the Na0.45 samples, together with data of glassy-Na2Si2O5.14 This graph shows that a considerable drop in conductivity for the Na0.25 and Na0.45 samples, together with data of glassy-Na2Si2O5.14 This graph shows that a considerable drop in conductivity for the Na0.45 samples, together with data of glassy-Na2Si2O5.

In order to support the identification of the glassy phase, a sample of nominal composition Na2SiO3 was prepared from the recrystallized species, as also observed in the XRD patterns. This hypothesis is also supported by the appearance of the broad peak at −89 ppm with the introduction of Na in the system, indicative of the formation of a new disordered phase (see below and Fig. 3e) whose content increases as the nominal Na content in the sample does. This second phase has been identified with glassy Na2O2SiO2. In fact, it can be formed with respect to the nominal stoichiometry of the initial mixtures for the different compositions, if we assume limited solubility of Na on the Sr site of SrSiO3. This is confirmed by the spectrum of Na0.45-RC (Fig. 3a). The peak at −95.7 ppm has been attributed to silicon in the crystalline Na4-xNa2Si2O5 phase.13 The small peaks between −87 and −92 ppm can be attributed to minor contributions from other unidentified recrystallized species, as also observed in the XRD patterns.

The results presented here clearly evidenced that the actual solubility degree of Na in the SrSiO3 phase is significantly lower than expected due to the formation of a secondary and metastable glassy phase, thus leading to a very low amount of oxygen vacancies with respect to the nominal stoichiometry. We showed that such a secondary phase may be at the origin of the promising transport properties of these materials. Most probably, the high conductivity values previously reported are due only to a limited...
extent to oxide ion migration in the crystalline system, in addition to transport of Na ions in the glassy phase, in agreement with a very recent work. This suggestion could also be consistent with the very low activation energy reported for this family of oxides compared to conventional oxide ion conductors. We recognize that the degree of crystallization–amorphization of the system may be strongly dependent on the synthesis conditions and therefore different results, especially for what concern the conductivity data, may be obtained. However, this study suggests that attention should be paid in the interpretation of data related to compositions that may easily give rise to glassy phases, such as alkali-doped silicates, as suggested also in our previous work.

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Notes and references