A new ultrafast superionic Li-conductor: ion dynamics in Li$_{11}$Si$_2$PS$_{12}$ and comparison with other tetragonal LGPS-type electrolytes†

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We report on a new ultrafast solid electrolyte of the composition Li$_{11}$Si$_2$PS$_{12}$, which exhibits a higher room-temperature Li ion diffusivity than the present record holder Li$_{10}$GeP$_2$S$_{12}$. We discuss the high-pressure synthesis and ion dynamics of tetragonal Li$_{10}$SiP$_2$S$_{12}$, and comparison is made with our investigations of related members of the LMePS family, i.e. electrolytes of the general formula Li$_{11-}$_Me$_{2-}$_P$_{1-}$_S$_{12}$ with Me = Ge, Sn : Li$_{10}$GeP$_2$S$_{12}$, Li$_{10}$GePS$_8$, Li$_{10}$SnP$_2$S$_{12}$. The structure and dynamics were studied with multiple complementary techniques and the macroscopic diffusion could be traced back to fast Li ion hopping in the crystalline lattice. A clear correlation between the diffusivity and the unit cell volume of the LGPS-type electrolytes was observed.

In 2011, the new solid lithium electrolyte Li$_{10}$GeP$_2$S$_{12}$ (LGPS) was reported, featuring liquid-like Li ion conduction in a crystalline solid matrix.$^{1,2}$ The ultrafast room temperature transport of tetragonal LGPS with a conductivity of several mS cm$^{-1}$ came as a surprise as it exceeds the values of the best crystalline Li conductors by one order of magnitude. Therefore, there has been a strong upsurge of interest recently in realizing LGPS-type materials based on the homologous elements Si and Sn. A theoretical study published by Ceder and coworkers highlights the potential of such hypothetical tetragonal LGPS-type Li ion conductors.$^3$ Li$_{10}$SnP$_2$S$_{12}$ was recently reported$^{4,5}$ and showed slightly reduced Li mobility as compared to LGPS, in line with the predictions by Ceder et al.$^1$

Here, we (i) report the successful high-pressure synthesis of the Si-analogue, Li$_{11}$Si$_2$PS$_{12}$, and (ii) present a comparative fundamental study of the Li ion dynamics in the four LGPS-type electrolytes reported so far, Li$_{11}$Si$_3$PS$_{12}$ (LSiPS, this study), Li$_{10}$SnP$_2$S$_{12}$ (LSnPS, this study and ref. 4 and 5), Li$_{10}$GeP$_2$S$_{12}$ (ref. 1 and 6) and Li$_{10}$GePS$_8$ (ref. 6). LSnPS and LSiPS were comprehensively characterized both with respect to their structure and Li ion dynamics and compared with the results previously published on LGPS.$^6$ In excellent agreement with the theoretical predictions in ref. 3, LSiPS shows an even higher Li diffusivity than LGPS while LSnPS has a slightly lower Li diffusivity. A clear correlation between the diffusivity and the unit cell volume was observed.

Tetragonal Li$_{10}$SnP$_2$S$_{12}$ was prepared by heating elemental Sn, P, S, and Li$_2$S to 653 K for 10 h with an additional sintering step at 723 K for two days, cf. ref. 4 and 5. The tendency of the larger Sn$^{IV}$ to reside in six-fold rather than four-fold coordination (the latter being desired for LSnPS) is reflected by the presence of side phases with edge-sharing Sn$_6$ building units (such as Li$_4$SnS$_4$) when stoichiometric amounts of the starting materials are used. A 10–20% excess of Li$_2$S, however, completely prevents the formation of these layered side phases. It should be noted that we used a slight excess of S yielding approx. 1 atm S at the reaction conditions in order to ensure complete oxidation of Sn and P.

The preparation of phase-pure hypothetical tetragonal Li$_{10}$SiP$_2$S$_{12}$ was not successful by means of a conventional solid-state approach. Instead, the main phase at all temperatures between 573 K and 1023 K was the orthorhombic modification of the LSiPS solid solution, which was known to possess distinctly less favorable transport properties.$^7$ The fact that the tetragonal modification has a slightly higher density than the orthorhombic one (for LGPS, compare ref. 1 and 10), and that a higher Si content should enhance the stability owing to the lower ionic radius of Si led us to the successful preparation of Li$_{11}$Si$_3$PS$_{12}$ by high-pressure synthesis [see ESI,$^†$ S1 for further details]. We obtained almost phase pure Li$_{11}$Si$_3$PS$_{12}$, which is referred to as LSiPS in the following.

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obtained the tetragonal modification only for values of $x$(Sn) very close to 0.5. For other Sn/P ratios, two phases were obtained: tetragonal LSnPS of the composition Li$_{10}$SnP$_2$S$_{12}$ and the orthorhombic modification as side phase. This can be explained by the larger ionic radius of Sn$^{IV}$ as compared to Ge$^{IV}$, rendering a higher occupancy of this position relative to P$^V$ energetically unfavourable. The opposite is true for LSiPS: Si$^{IV}$ is only slightly larger than the isoelectronic P$^V$ ion. Therefore, in order to stabilize the tetragonal modification with two sets of differently-sized tetrahedra, the 4$d$ site has to be occupied by Si to a much higher extent as compared to Ge or Sn. This explains why the tetragonal modification is obtained for the stoichiometry Li$_{11}$Si$_2$PS$_{12}$ rather than Li$_{10}$SiP$_2$S$_{12}$.

$^{31}$P MAS NMR was used in order to probe the relative amount of P residing on the 4$d$ and 2$b$ sites (cf. Fig. 1a). This is of special importance for the structural elucidation of the LSiPS sample because here, this information is not accessible from X-ray diffraction since Si$^{IV}$ and P$^V$ are isoelectronic. Fig. 2 shows the $^{31}$P MAS NMR spectra of the tetragonal LGPS-type electrolytes. The spectrum of Li$_{10}$SnP$_2$S$_{12}$ is very similar to that of Li$_{10}$GeP$_2$S$_{12}$: as expected from Rietveld refinement, the relative intensity of the signals assigned to the 4$d$ and 2$b$ sites is 1:1 within 5% error. The third signal shows the chemical shift typical of the orthorhombic modification, which is, however, not observed in the X-ray patterns. Therefore, we assign it to the presence of a side phase of low crystallinity, which resembles the local structure of the orthorhombic modification (see also ref. 6). For Li$_{11}$Si$_2$PS$_{12}$, the structural model assumed above – judging only from the stoichiometry of the sample – is largely verified. Only a small amount of P resides on the mixed-occupied 4$d$ site (~10%). Thus, the occupancy of the 4$d$ site for all LGPS-type samples follows a clear trend. For the largest ion, Sn, the occupancy is close to $x$(Sn) ~ 0.50. For Ge, the occupancy ranges from 0.5 < $x$(Ge) < 0.75, while for the small Si, $x$(Si) ~ 0.95.

In order to characterize the Li diffusivity in the LGPS-type materials Li$_{10}$SnP$_2$S$_{12}$ and Li$_{11}$Si$_2$PS$_{12}$, we performed $^7$Li PFG NMR. The accessible range of the LSnPS solid solution is much narrower than in case of LGPS – we

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**Fig. 1** (a) Crystal structure of tetragonal LGPS as obtained from single-crystal X-ray diffraction. (b) X-ray powder diffraction and Rietveld refinement of Li$_{11}$Si$_2$PS$_{12}$ and Li$_{10}$SnP$_2$S$_{12}$ in comparison to previously reported Li$_{10}$GeP$_2$S$_{12}$ and Li$_7$GePS$_8$. The side phase is marked by a green asterisk.

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**Fig. 2** $^{31}$P MAS NMR spectra ($v_{rot}$ = 12 kHz, $B_0$ = 9.4 T) of the different isostructural materials crystallizing in the tetragonal LGPS-type. The spectra for Li$_7$GePS$_8$ and Li$_{10}$GeP$_2$S$_{12}$ as well as the assignment of the lines were taken from ref. 6. The asterisks denote impurities with a chemical shift typical of the orthorhombic modification.
NMR measurements. The obtained Li tracer-diffusion coefficients \( D^T \) (3D diffusion, cf. ref. 6) are shown in Fig. 3a in comparison with those previously reported for \( \text{Li}_{10}\text{GeP}_2\text{S}_{12} \).\(^6\) Note that in all cases, the measured diffusion coefficients can clearly be assigned to diffusion in the bulk of the tetragonal LGPS-type electrolytes, since the quadrupolar structure of the decaying NMR signal showed the fingerprint (see Fig. 4a and b) of the tetragonal modification, which is distinct from that of orthorhombic or amorphous side phases. The diffusivity of \( \text{Li}_{10}\text{SnP}_2\text{S}_{12} \) is slightly lower than that of \( \text{Li}_{10}\text{GeP}_2\text{S}_{12} \) and the activation energy is slightly higher \((0.23(1) \text{ eV} \text{ vs. } 0.21(1) \text{ eV})\). In contrast, the diffusivity of \( \text{Li}_{11}\text{Si}_2\text{PS}_{12} \) is even higher than that of \( \text{Li}_{10}\text{GeP}_2\text{S}_{12} \) with a slightly lower activation energy \((0.19(1) \text{ eV} \text{ vs. } 0.21(1) \text{ eV})\). We would like to point out that this trend is – both qualitatively and quantitatively – in very good agreement with theoretical calculations by Ong et al.\(^3\) as presented in Table 1. Note that the theoretical diffusion coefficients taken from the MD simulation in ref. 3 have been extrapolated from 600 K down to room temperature in order to compare them with our experimental data. As shown in Fig. 3b, a clear correlation between the unit cell volume and the diffusion parameters is observed.
Table 1  Experimental Li+ diffusion coefficients at room temperature and activation energies for tetragonal LGPS-type materials obtained from 7Li PFG NMR measurements in comparison with the theoretical ones obtained from MD simulations in ref. 3.

<table>
<thead>
<tr>
<th>Compound</th>
<th>Dexp/m2s⁻¹</th>
<th>Dsim/m2s⁻¹</th>
<th>EArr/eV</th>
<th>EArropt/eV</th>
</tr>
</thead>
<tbody>
<tr>
<td>Li10SiP2S12</td>
<td>—</td>
<td>3.5 x 10⁻¹²</td>
<td>0.19(1)</td>
<td>a</td>
</tr>
<tr>
<td>Li10SiPS2S12</td>
<td>1.9 x 10⁻¹¹</td>
<td>2.2 x 10⁻¹²</td>
<td>0.21(1)</td>
<td>b</td>
</tr>
<tr>
<td>Li10GePS2S12</td>
<td>1.4 x 10⁻¹²</td>
<td>1.8 x 10⁻¹²</td>
<td>0.23(1)</td>
<td>b</td>
</tr>
<tr>
<td>Li3GePS3S12</td>
<td>4.4 x 10⁻¹²</td>
<td>1.4 x 10⁻¹²</td>
<td>0.24(1)</td>
<td>a</td>
</tr>
</tbody>
</table>

a This work. b From ref. 6. c Extrapolated from Fig. 3 in ref. 3. d From ref. 3. e From ref. 4.

Thus, the appearance of this distinct quadrupolar powder pattern indicates that the fast Li dynamics measured in fact occurs within crystallites of tetragonal LSnPS or LSiPS.

Impedance spectroscopy was applied in order to study the ionic conductivity of Li10SnP2S12 for which dense ceramic samples could be obtained. Fig. 5a shows the temperature-dependent total conductivity and the bulk conductivity extracted from the impedance spectroscopy measurements for Li10SnP2S12 (see ESI† for details). The bulk conductivity of LSiPS as calculated from NMR diffusivity data is included as a dotted line. For comparison, the dashed dotted line represents the conductivity of the best oxidic solid electrolyte, Li10LaZr2O12. (b) Galvanostatic dc polarization measurement on a symmetric Aul|Li10SnPS12|Au cell at 473 K. (c) The polarization curve from (b), as a function of linear fit. See text and ESI† for further details. (d) Comparison of the diffusion coefficients obtained from long-range sensitive methods (PFG NMR, impedance spectroscopy) and short-range sensitive methods (NMR relaxometry) for LSnPS and LSiPS. The data derived from NMR relaxometry stem from [a] motional narrowing of the 7Li central transition (see ESI† S4), [b] motional averaging of the quadrupolar interaction, (see ESI† S4), and [c] longitudinal relaxation (see ESI† S4). The Arrhenius line obtained for LGPS is included as well for comparison.
the electronic contribution to the total conductivity, a dc polarization measurement was carried out at 473 K by applying a small current of 1 nA to a symmetric cell with ion blocking Au electrodes Au|LSnPS|Au. Fig. 5b shows the polarization curve. For a good solid electrolyte with vanishingly small electronic contribution one not only expects a considerable polarization but also a small chemical diffusion coefficient. Accordingly, a steady state could not be attained within reasonable waiting time, but an upper limit of the electronic conductivity corresponding to an electronic transference number of $t_{\text{ION}} < 1$ PPM could be safely determined from the absolute voltage values as well as from the time behaviour (linear fit in Fig. 5c, cf. ESI† S5). Thus, LSnPS can be considered as a purely ionic conductor. The impedance of the pellet prior to and after the dc polarization measurement was equal within 1%. For LSiPS, owing to the instability at sintering temperatures, the preparation of phase pure dense ceramic samples suitable for precise impedance spectroscopic experiments failed and we hence concentrate on the NMR results.

Fig. 5d summarizes the results obtained from PFG NMR, NMR relaxometry, and conductivity measurements for LiSiPS and LSnPS. For comparison, the jump rates $\tau^{-1}$ and the conductivities $\sigma_{dc}$ were appropriately transformed into diffusion coefficients $D^{ac}$ (uncorrelated diffusion coefficient) and $D^c$ (conductivity diffusion coefficient) using the Einstein–Smoluchowski relation $D^{ac} = \alpha^2/6 \times \tau^{-3} \text{[jump distance $a$]}$ and the Nernst–Einstein relation $D^c = k_B T (Nq^2) / \sigma_{dc} \text{[number density of Li$^+$, charge of Li$^+$, $q$, Boltzmann's constant $k_B$]}$. For the calculation, $N$ and an average jump distance of $a \sim 2\text{Å}$ were deduced from the structure. Obviously, for both Li$_{10}$SnP$_2$S$_{12}$ and Li$_{11}$Si$_2$PS$_{12}$ the macroscopically observed tracer diffusion can be traced back to 3D Li hopping in the bulk lattice with activation energies of 0.25(1) eV and 0.20(1) eV, respectively. The correlation factor $f$ and the Haven ratio $H_R$, connecting $D^{ac}$ and $D^c$ with $D^c$ via $D^c = f \times D^{ac} = H_R \times D^c$ are on the order of unity as expected for simple diffusion mechanisms. In view of the very good agreement between the results obtained from NMR techniques and impedance spectroscopy in the case of LGPS and LSnPS, we have calculated the expected bulk conductivity of LiSiPS from the Arrhenius line obtained from NMR data (Fig. 5d) using the Nernst–Einstein relation and included it in Fig. 5a. For comparison with other non-LGPS-type electrolytes, the conductivity of the best oxidic solid electrolyte reported to date, Li$_3$La$_2$Zr$_2$O$_{12}$, is included as well. For further comparison, Li$_3$La$_2$Zr$_2$O$_{12}$ shows a conductivity which is almost identical with that of the well-known superionic conductor Li$_3$N, but significantly lower than that of the LGPS-type electrolytes reported in this study.

Conclusions

The transport properties of Li$_{11}$Si$_2$PS$_{12}$, the new tetragonal Si-analogue of the ultrafast electrolyte Li$_{10}$GeP$_2$S$_{12}$ are reported, which was obtained under high-pressure conditions for the first time and shown to exhibit the highest room temperature conductivity of any Li solid electrolyte known to date. The structure and Li ion dynamics were characterized in comparison with the Ge-containing parent compound and the Sn analogue, which was comprehensively characterized as well. The Li ion dynamics were elucidated by both long-range sensitive and short-range sensitive techniques and the long-range transport was traced back to Li hopping in the crystalline lattice. In agreement with theoretical predictions, the diffusivity of the Si-compound exceeds that of the Ge-compound while the Sn-analogue shows a lower diffusivity. A clear correlation between the unit cell volume and the diffusivity was observed. Studies of how far the LiMePS family is applicable for specific problems of energy research is primarily a question of mechanical stability (e.g. sinterability) and chemical stability of specific contacts. Such studies are currently underway, but out of scope of this report. In summary, the observed high ionic conductivity and earth-abundance of its constituents renders the Ge-free LGPS-type electrolyte Li$_{11}$Si$_2$PS$_{12}$ a promising candidate for the development of a new generation of all-solid-state batteries.

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Notes and references

5 Preliminary results of us on LSnPS and LSiPS had been reported on the occasion of the 112th Bunsentagung, Karlsruhe, Germany (May 10th 2013), as well as at the 19th International Conference on Solid State Ionics, Kyoto, Japan (June 3rd 2013).


