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# Highly durable carbon-supported Pt catalysts prepared by hydrosilane-assisted nanoparticle deposition and surface functionalization†

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**Hydrosilane enabled the formation of Pt nanoparticles and the silane functionalization of a carbon support material in one pot. The metal/Si-modified carbon composites are highly durable during catalytic methane oxidation.**

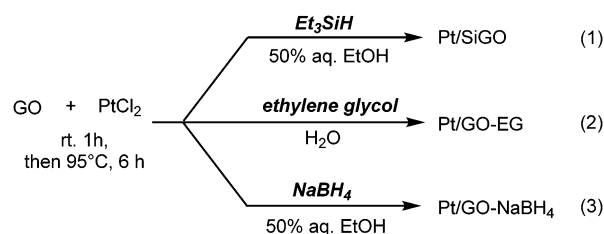
Carbon-supported noble metal composites are used as catalysts in organic synthesis,<sup>1</sup> fuel cells,<sup>2</sup> super capacitors,<sup>3</sup> solar cells,<sup>4</sup> and sensors.<sup>5</sup> Support materials composed of sp<sup>2</sup> hybridized carbon atoms have advantages of high electron and thermal conductivities, which enhance the performance of the metal species.<sup>6</sup> In contrast, carbon reacts readily with oxygen to eliminate gaseous carbon oxides upon heating especially in the presence of metal species, and thus catalytic combustion occurs rapidly. Therefore, the development of thermally durable carbon materials is of great importance.<sup>7</sup> The preparation of metal-supported carbon materials covered with silica layers has been reported and they have been used as catalysts for cyclohexane dehydrogenation.<sup>8</sup> However, multistep functionalization processes such as (1) the deposition of metal nanoparticles onto carbon, (2) surface treatment with a silane coupling agent, and (3) silica-layer formation by the hydrolysis of an alkoxy silane are required.

We focused on the development of a one-pot synthesis of a metal/Si-modified carbon composite with high thermal durability. To achieve this we used hydrosilane (R<sub>3</sub>SiH) as the dual-role reagent for the metal nanoparticle formation and for the surface modification of the carbon support. Hydrosilanes can reduce metal salts

(M<sup>n+</sup>) to metal nanoparticles (M<sup>0</sup>) through transmetallation or by ligand exchange. A successive reductive elimination<sup>9</sup> is accompanied by the formation of R<sub>3</sub>SiX (X = the counter anion of M<sup>n+</sup> or OH) that functions as a silane coupling agent<sup>10</sup> towards OH groups on the carbon support. Among the various carbon-based support materials, we initially focused on graphene oxide (GO) because of its highly oxygenated and two-dimensional nanosheet structure. This allows for uniform nanoparticle distribution and silane functionalization. The thermal resistance of the composites was evaluated by a catalytic methane oxidation reaction. This reaction requires high energy because of the high C–H bond dissociation energy (104 kcal mol<sup>−1</sup>).<sup>11</sup>

A Pt/GO composite (Pt/SiGO) was prepared from GO and PtCl<sub>2</sub> using Et<sub>3</sub>SiH as a reductant.<sup>12</sup> The mixture was stirred at room temperature for 1 h and then heated to 95 °C for 6 h (Scheme 1-1). For comparison, PtCl<sub>2</sub> was reduced by ethylene glycol (Pt/GO–EG, Scheme 1-2) or NaBH<sub>4</sub> (Pt/GO–NaBH<sub>4</sub>, Scheme 1-3).

The amount of Pt on Pt/SiGO was found to be 4.2 wt% by energy dispersive X-ray (EDX) spectroscopy (Table 1, entry 1). 1.5 wt% Si was also detected by EDX. The Pt nanoparticles on the surface of GO were characterized by X-ray diffraction (XRD). The XRD patterns of graphite, GO, and Pt/SiGO are shown in Fig. 1. The peak at 2θ = 10.3° in the XRD pattern corresponds to GO (002), and the broad peak at around 2θ = 23° is attributed to the stacked graphene sheets produced by the reduction of GO.<sup>13</sup> For Pt/SiGO the peaks are at 2θ = 39.7, 46.2, and 67.6° and these peaks correspond to the (111), (200), and (220) planes



Scheme 1 Preparation of the Pt composites.

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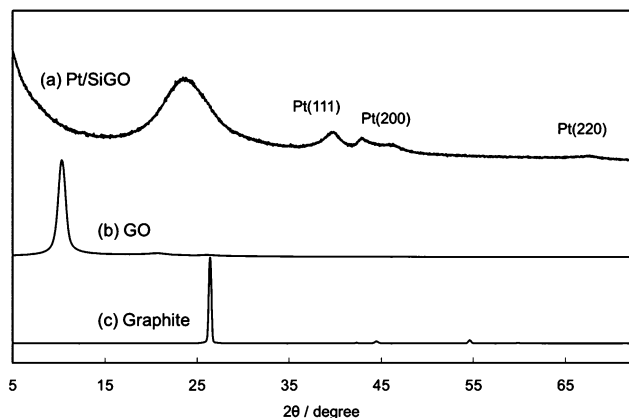
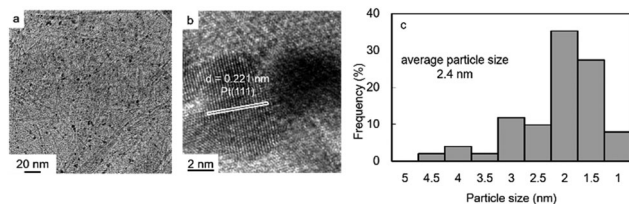
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**Table 1** Pt content of the catalysts as determined by EDX

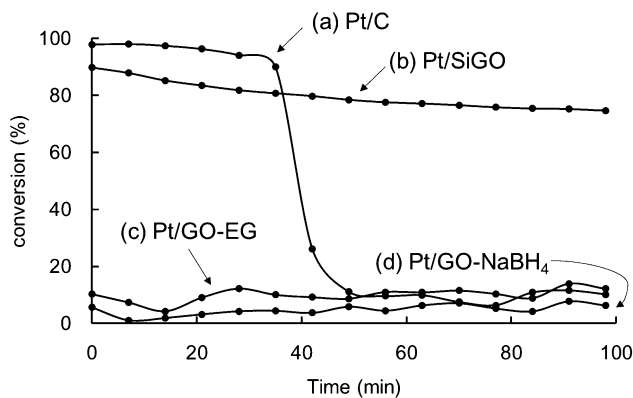
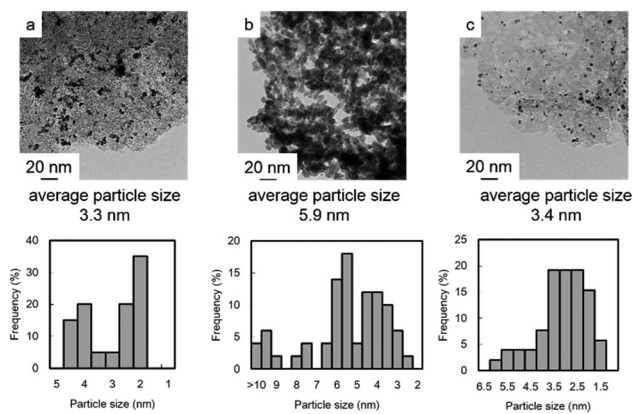
Entry	Catalyst	Reductant	Support	Pt content (wt%)
1	Pt/C	— <sup>a</sup>	C	5.0
2	Pt/SiGO	Et <sub>3</sub> SiH	GO	4.2
3	Pt/GO–NaBH <sub>4</sub>	NaBH <sub>4</sub>	GO	5.9
4	Pt/GO–EG	EG	GO	3.5

<sup>a</sup> Unknown (commercial product).**Fig. 1** XRD spectra of (a) Pt/SiGO, (b) GO, and (c) graphite.**Fig. 2** (a, b) TEM images of Pt/SiGO, and (c) the particle size distribution.

of a face-centered cubic (fcc) Pt crystal, respectively.<sup>14</sup> The morphology of the Pt nanoparticles was observed using a transmission electron microscope (TEM) (Fig. 2). The average particle size of the Pt nanoparticles was 2.4 nm. The lattice space of the Pt nanoparticles was 0.22 nm, which corresponds to the (111) plane of a Pt crystal.<sup>14</sup>

The thermal durability of Pt/SiGO was evaluated by the catalytic methane oxidation reaction (Fig. 3), because this reaction requires high temperature. For comparison, commercially available Pt/C and Pt/GO composites prepared using ethylene glycol (Pt/GO–EG) or NaBH<sub>4</sub> (Pt/GO–NaBH<sub>4</sub>)<sup>15</sup> were also evaluated.<sup>16</sup> Catalysts containing 10 mg of Pt, as determined by EDX (Table 1), were subjected to treatment with a gas containing 7% methane, 33% O<sub>2</sub>, and 60% He at 400 °C. The flow rate was 30 mL min<sup>−1</sup> and the conversion of methane was measured by gas chromatography (Fig. S2, ESI†). For Pt/C (Fig. 3a) the catalytic activity decreased significantly after 40 min while Pt/SiGO maintained its activity (Fig. 3b). The Pt/GO prepared using ethylene glycol (Pt/GO–EG) (Fig. 3c) or NaBH<sub>4</sub> (Pt/GO–NaBH<sub>4</sub>) (Fig. 3d) did not show high activity at any point.

The recovered catalysts were analyzed by TEM (Fig. 4). The original Pt/C contained 3.3 nm Pt nanoparticles (Fig. 4a); however, after methane oxidation almost all the carbon support

**Fig. 3** Time course of the catalytic activity of methane conversion under oxygen at 400 °C. (a) Pt/C, (b) Pt/SiGO, (c) Pt/GO–EG, and (d) Pt/GO–NaBH<sub>4</sub>.**Fig. 4** TEM images and particle size distributions of (a) fresh Pt/C, (b) recovered Pt/C, and (c) recovered Pt/SiGO.

disappeared and nanoparticle grain growth occurred (Fig. 4b). In contrast, the recovered Pt/GO that was prepared using Et<sub>3</sub>SiH did not undergo such a drastic structural change. Its particle size increased slightly from 2.4 nm (Fig. 2) to 3.4 nm (Fig. 4c).

XPS was measured before and after the methane oxidation reaction. The intensities of each spectra were normalized to that of the Pt 4f<sub>7/2</sub> and 4f<sub>5/2</sub> regions (65–76 eV), because we hypothesized that the amount of Pt does not change before and after the reaction. For Pt/SiGO, the intensity of the Si 2p and 2s regions (99 and 151 eV) increased after the reaction at 400 °C (Fig. 5A). This suggests that silyl functional groups were covalently attached to GO because the boiling points of Et<sub>3</sub>SiH and Et<sub>3</sub>SiOH are only 107 and 158 °C, respectively. The increase in Si indicates that other elements such as C and O were removed and, therefore, XPS spectra of the C 1s regions were analyzed. As expected, the intensities of the C 1s regions of Pt/SiGO after the methane oxidation reaction decreased; however, the decrease is much smaller than that of Pt/C (Fig. 5B).

Si atoms were well dispersed over GO as determined by scanning transmission electron microscopy (STEM)-EDX mapping (Fig. S3, ESI†). The hydroxyl groups on GO are known to react with silane coupling reagents (silyl chlorides) to form silylated GO.<sup>17</sup> Here, we used Et<sub>3</sub>SiH to reduce Pt(II) to Pt nanoparticles and, as a



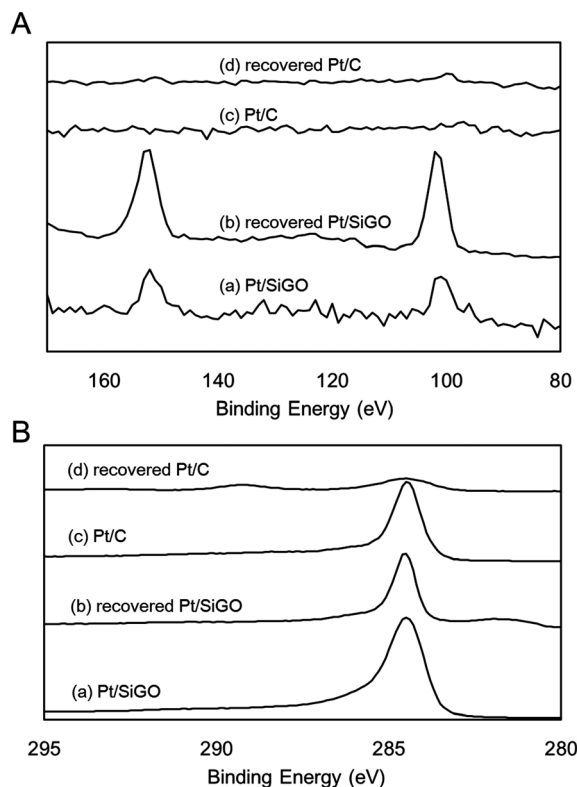


Fig. 5 XPS of (A) Si 2p and 2s, and (B) the C 1s regions of Pt/SiGO and Pt/C before and after the methane oxidation reaction.

result,  $\text{Et}_3\text{SiOH}$  is produced.<sup>18</sup>  $\text{Et}_3\text{SiOH}$  can also function as a silane coupling reagent<sup>10</sup> and, therefore, the OH groups on GO are functionalized with  $\text{Et}_3\text{Si}$  groups *via* dehydrative condensation. To determine the chemical state of Si on GO, an X-ray absorption near edge structure (XANES) spectrum of Pt/SiGO was obtained. Si(0) and  $\text{SiO}_2$  have peaks at 1842 and 1848 eV, respectively (Fig. 6a and b).<sup>19</sup> The XANES spectrum of the  $\text{Et}_3\text{Si}$ -modified GO ( $\text{Et}_3\text{Si-GO}$ ), produced using  $\text{Et}_3\text{SiCl}$ , had peaks at 1845 eV and 1848 eV (Fig. 6c). For Pt/SiGO a similar spectrum to that of  $\text{Et}_3\text{Si-GO}$  was obtained (Fig. 6d). This suggests that the  $\text{Et}_3\text{Si}$  groups are attached to Pt/SiGO by silyl ether bonds.<sup>20</sup> Fourier transform infrared

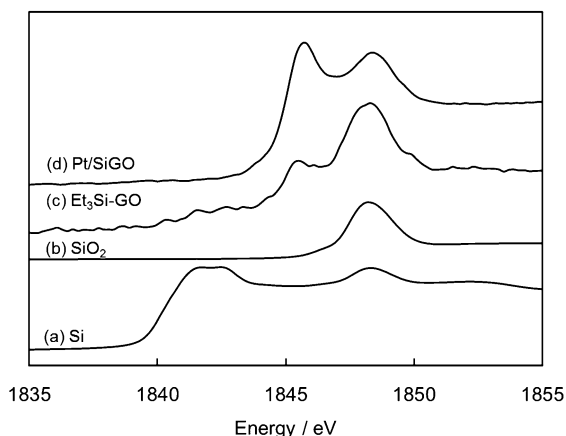


Fig. 6 Si K-edge XANES spectra of (a) Si(0), (b)  $\text{SiO}_2$ , (c)  $\text{Et}_3\text{Si}$ -modified GO, and (d) Pt/SiGO.

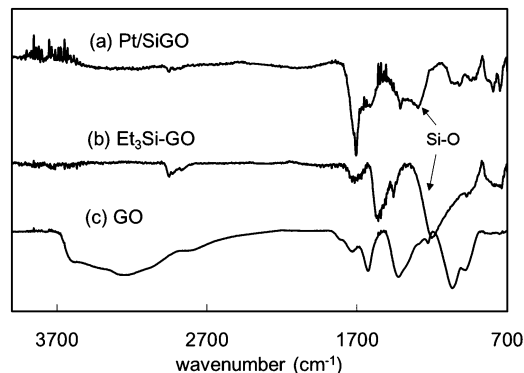


Fig. 7 FT-IR spectra of (a) Pt/SiGO, (b)  $\text{Et}_3\text{Si-GO}$ , and (c) GO.

(FT-IR) spectroscopy of Pt/SiGO also supports the presence of the  $\text{Et}_3\text{Si}$  groups; peaks of alkyl chains were observed at 2740–2980  $\text{cm}^{-1}$  (Fig. 7). Si–O bond has a strong IR absorption at 1100–1200  $\text{cm}^{-1}$ , which is observed for both Pt/SiGO (Fig. 7a) and  $\text{Et}_3\text{Si-GO}$  (Fig. 7b).

To clarify the durability of Pt/SiGO under heating, thermogravimetric analysis (TGA) was conducted (DTA is shown in Fig. S4, ESI†). A drastic weight loss was observed at 350 °C on Pt/C under air. Pt/SiGO and Pt/GO– $\text{NaBH}_4$  showed higher thermal durability; however, the obvious effect of silyl functionalization was not observed (Fig. 8). These results suggest that the silyl group does not affect carbon supports, but stabilizes Pt particles. The weight loss of Pt/SiGO at 50–100 °C would derive from the desorption of water, and the loss at 100–350 °C would derive from the removal of residual oxygen functional groups and alkyl chains, which were also supported by IR analysis (Fig. S5, ESI†).

The hydrosilane-assisted deposition of Pt also enhanced the thermal durability of activated carbon (Pt/SiC) (Fig. 9a). Interestingly, treatment of the commercially available Pt/C with  $\text{Et}_3\text{SiH}$  (Pt/C– $\text{Et}_3\text{SiH}$ ) also showed improved durability (Fig. 9b) compared with the as-obtained compound (Fig. 9c).

We developed a one-pot synthesis of Pt/Si-modified carbon composites. Si atoms were uniformly present on the carbon

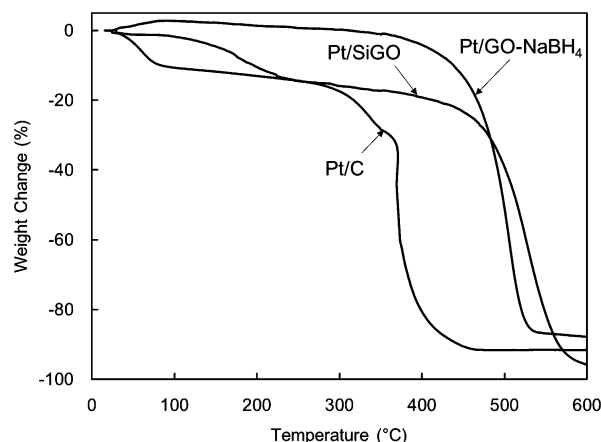


Fig. 8 TGA of Pt/SiGO, Pt/C, and Pt/GO– $\text{NaBH}_4$  under air. Heating rate: 2 °C  $\text{min}^{-1}$ .



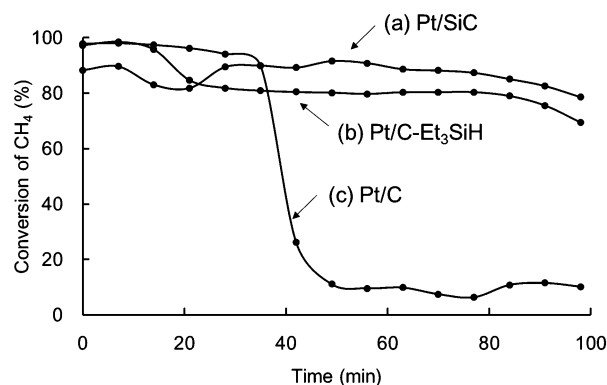


Fig. 9 Catalytic activities of the hydrosilane-treated catalysts.

surface upon dehydrative silyl ether formation between Et<sub>3</sub>SiOH and the OH group of the carbon. The thermal durability of the composite during methane oxidation improved upon silyl group modification. This work was partly supported by the Nanotechnology Platform Program, JST PRESTO, KAKENHI of the Ministry of Education, Culture, Sports, Science and Technology, and the Cooperative Research Program of Catalysis Research Centre, Hokkaido University. We acknowledge Dr Hideki Hashimoto for STEM-EDS measurements.

## Notes and references

- (a) H. Sakurai, T. Tsukuda and T. Hirao, *J. Org. Chem.*, 2002, **67**, 2721; (b) T. Maegawa, Y. Kitamura, S. Sako, T. Udzu, A. Sakurai, A. Tanaka, Y. Kobayashi, K. Endo, U. Bora, T. Kurita, A. Kozaki, Y. Monguchi and H. Sajik, *Chem. – Eur. J.*, 2007, **13**, 5937; (c) L. Yin and J. Liebscher, *Chem. Rev.*, 2007, **107**, 133; (d) F. X. Felpin, T. Ayad and S. Mitra, *Eur. J. Org. Chem.*, 2006, 2679.
- (a) P. J. Ferreira, G. J. La Ó, Y. Shao-Horn, D. Morgan, R. Makharia, S. Kocha and H. A. Gasteiger, *J. Electrochem. Soc.*, 2005, **152**, A2256; (b) Y. Shao, G. Yin, Y. Gao and P. Shi, *J. Electrochem. Soc.*, 2006, **153**, A1093; (c) Z. Zhou, S. Wang, W. Zhou, G. Wang, L. Jiang, W. Li, S. Song, J. Liu, G. Sun and Q. Xin, *Chem. Commun.*, 2003, 394; (d) J. R. Varcoe, R. C. T. Slade, G. L. Wright and Y. Chen, *J. Phys. Chem. B*, 2006, **110**, 21041.
- G. Y. Yu, W. X. Chen, Y. F. Zheng, J. Zhao, X. Li and Z. D. Xu, *Mater. Lett.*, 2006, **60**, 2453.
- Y. H. Ng, I. Lightcap, K. Goodwin, M. Matsumura and P. V. Kamat, *J. Phys. Chem. Lett.*, 2010, **1**, 2222.
- I. V. Lightcap, S. Murphy, T. Schumer and P. V. Kamat, *J. Phys. Chem. Lett.*, 2012, **3**, 1453.
- (a) A. Kongkanand, K. Vinodgopal, S. Kuwabata and P. V. Kamat, *J. Phys. Chem. B*, 2006, **110**, 16185; (b) R. Muszynski, B. Seger and P. V. Kamat, *J. Phys. Chem. C*, 2008, **112**, 5263; (c) K.-S. Ha, G. Kwak, K.-W. Jun, J. Hwang and J. Lee, *Chem. Commun.*, 2013, **49**, 5141.
- L. D. Rogatis, M. Cargnello, V. Gombac, B. Lorenzuti, T. Montini and P. Fornasiero, *ChemSusChem*, 2010, **3**, 24.
- (a) K. Nakagawa, T. Okayama, Y. Tanimoto, K. Sotowa, S. Sugiyama, T. Moriga, S. Takenaka and M. Kishida, *Appl. Catal., A*, 2012, **419**, 13; (b) K. Nakagawa, T. Hoshino, K. Sotowa, S. Sugiyama, L. Rkiouak, V. Dubois and S. Hermans, *Chem. Lett.*, 2012, **41**, 1308; (c) K. Nakagawa, Y. Tanimoto, T. Okayama, K. Sotowa, S. Sugiyama, S. Takenaka and M. Kishida, *Catal. Lett.*, 2010, **136**, 71; (d) S. Takenaka, T. Iguchi, E. Tanabe, H. Matsune and M. Kishida, *Carbon*, 2009, **47**, 1251; (e) S. Takenaka, H. Matsumori, K. Nakagawa, H. Matsune, E. Tanabe and M. Kishida, *J. Phys. Chem. Lett.*, 2007, **111**, 15133; (f) S. Takenaka, T. Arike, H. Matsune, E. Tanabe and M. Kishida, *J. Catal.*, 2008, **257**, 345.
- (a) L. N. Lewis and N. Lewis, *J. Am. Chem. Soc.*, 1986, **108**, 7228; (b) B. Fu, M. N. Missaghi, C. M. Downing, M. C. Kung, H. H. Kung and G. Xiao, *Chem. Mater.*, 2010, **22**, 2181.
- Y. Xie, C. A. S. Hill, Z. Xiao, H. Militz and C. Mai, *Composites, Part A*, 2010, **41**, 806.
- S. J. Blanksby and G. B. Ellison, *Acc. Chem. Res.*, 2003, **36**, 255.
- Pt nanoparticle formation with other hydrosilanes was also investigated. See ESI†.
- P. Cui, J. Lee, E. Hwang and H. Lee, *Chem. Commun.*, 2011, **47**, 12370.
- Powder Diffraction File No. 04-0802.
- (a) B. Seger and P. V. Kamat, *J. Phys. Chem. Lett.*, 2009, **113**, 7990; (b) Y. Li, L. Tang and J. Li, *Electrochem. Commun.*, 2009, **11**, 846; (c) J. D. Qiu, G. C. Wang, R. P. Liang, X. H. Xia and H. W. Yu, *J. Phys. Chem. C*, 2011, **115**, 15639; (d) Y. Li, W. Gao, L. Ci, C. Wang and P. M. Ajayan, *Carbon*, 2010, **48**, 1124.
- The surface area of graphene oxide derivatives determined by the BET method is generally small (<100 m<sup>2</sup> g<sup>-1</sup>). We measured the surface area by methylene blue adsorption. Pt/SiGO: 534 m<sup>2</sup> g<sup>-1</sup>, Pt/GO-NaBH<sub>4</sub>: 744 m<sup>2</sup> g<sup>-1</sup>, Pt/GO-EG: 921 m<sup>2</sup> g<sup>-1</sup>, Pt/C: 1162 m<sup>2</sup> g<sup>-1</sup>, and GO: 1292 m<sup>2</sup> g<sup>-1</sup>. See ESI† for more details.
- (a) Y. Matsuo, T. Tabata, T. Fukunaga, T. Fukutsuka and Y. Sugie, *Carbon*, 2005, **43**, 2875; (b) Y. Matsuo, T. Fukunaga, T. Fukutsuka and Y. Sugie, *Carbon*, 2004, **42**, 2113.
- (a) K. Shimizu, T. Kubo and A. Satsuma, *Chem. – Eur. J.*, 2012, **18**, 2226; (b) B. P. S. Chauhan, A. Sarkar, M. Chauhan and A. Roka, *Appl. Organomet. Chem.*, 2009, **23**, 385; (c) M. Jeon, J. Han and J. Park, *ACS Catal.*, 2012, **2**, 1539.
- Y. Baba, T. Sekiguchi, I. Shimoyama and N. Hirao, *Surf. Sci.*, 2013, **612**, 77.
- Si K-edge XANES was also studied by the DV-X $\alpha$  molecular-orbital calculation method. The presence of two peaks in XANES spectra suggested the existence of Si–O and Si–C bonds. See ESI†.

