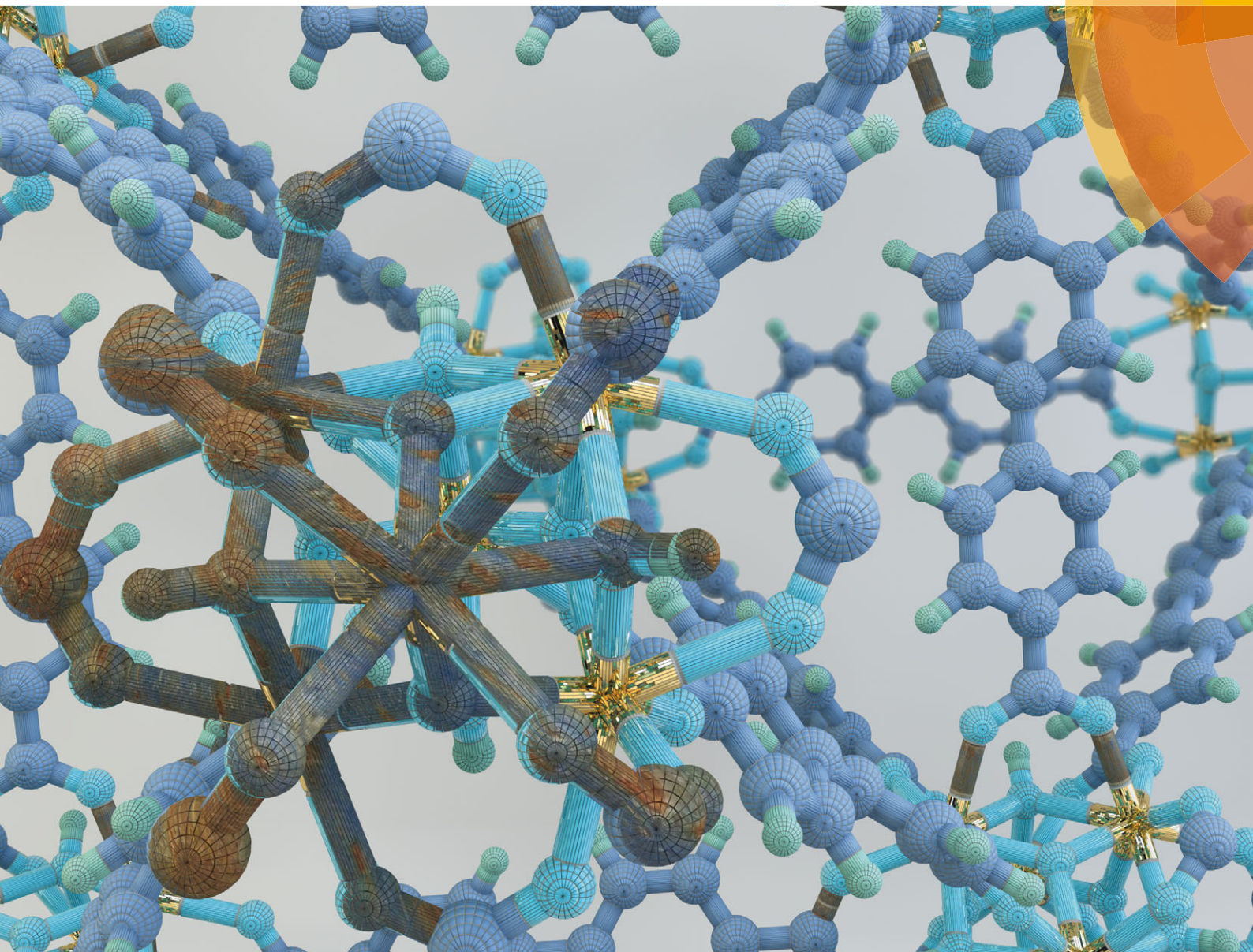


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Determining the structural stability of UiO-67 with respect to time: a solid-state NMR investigation†

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The stability of UiO-67 has been questioned for some time. We have used solid-state NMR to investigate the temporal stability of this MOF. Proper activation is necessary to achieve optimal surface area. However, even with proper activation, the long-term (30+ days) fate of UiO-67 is hydrolysis of the linker-metal bonds and, ultimately, pore collapse.

Metal-organic frameworks (MOFs) are porous materials formed *via* coordination of bridging organic ligands (linkers) with inorganic metal cations/clusters (nodes). With judicious choice of these components, MOFs with applications in gas-storage,¹ chemical separations,² light-harvesting,³ sensing,^{2b,4} and catalysis⁵ have been realized. One family of MOFs which are becoming ubiquitous in these applications is the Zr-cluster-containing family of MOFs, including, but not limited to, UiOs,⁶ NU-1000,⁷ PCN-222,⁸ and MOF-808.⁹ The interest in these MOFs stems from their thermal, chemical, and mechanic stability making them ideal for many applications.^{6d,10}

With respect to both anecdotal evidence as well as literature precedence, UiO-67 (Fig. 1) has had a precarious history. DeCoste *et al.* demonstrated that the internal surface area (SA) of UiO-67 decreased from 2145 m² g⁻¹ to 10 m² g⁻¹ after the MOF was exposed to 90% relative humidity. Similarly, when soaked in water, UiO-67 was found to be unstable; powder X-ray diffraction data (pxrd) indicated the presence of ZrO₂.¹¹ Although the instability is attributed to hydrolysis of the bonds between the linker and node, FTIR data showed no vibrational changes to corroborate this.

In a related manuscript, Mondloch *et al.* have proposed an alternate hypothesis. When UiO-67 was activated (*i.e.*, the process of removing solvent from the porous frameworks) from water, then there was no notable porosity remaining. However,

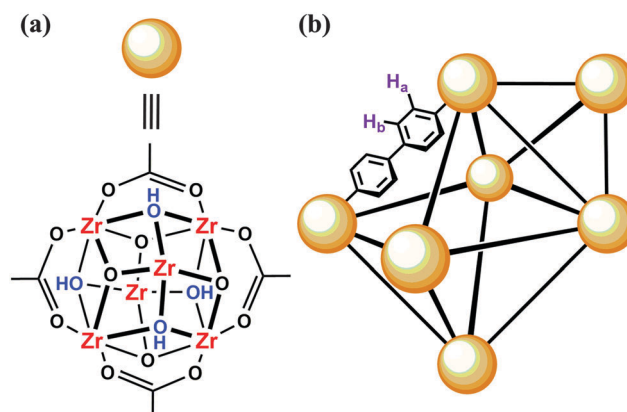


Fig. 1 (a) Zr₆O₄(OH)₄ cluster showing 4 of the 12 μ-BPDC units. (b) Schematic drawing of UiO-67 illustrating how the BPDC units link Zr₆O₄(OH)₄ clusters to one another to form both octahedral- and tetrahedral-shaped pores.

if water in the pore was replaced with acetone prior to activation, then the porosity remained; similar results were obtained when UiO-67 was boiled in water prior to solvent exchange. Thus, rather than an inherent instability in the MOF it was proposed that capillary-force driven collapse, due to improper activation, is responsible for the proposed instability.¹²

Given the utility of UiO-67,^{5b,6h,13} we were interested in further probing its stability. Specifically, we are interested in investigating the long-term stability of UiO-67 with respect to time. We turned our attention to solid-state NMR (SS-NMR) as a probe for the potential structural changes that occur within this MOF. Unlike pxrd, which is sensitive to crystalline materials containing high Z nuclei, SS-NMR is equally sensitive to both amorphous and crystalline materials. Furthermore, SS-NMR has the potential to independently report on each nucleus.

UiO-67 was synthesized by the method of Katz *et al.* (See ESI,† for further details).^{6a} UiO-67 was subsequently filter-dried (*ca.* 1 h). The samples of DeCoste *et al.*¹¹ were similarly filtered.

As expected, despite the apparently dry UiO-67, the SS-NMR indicates that freshly-prepared UiO-67 (Fig. 2 red trace) shows

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† Electronic supplementary information (ESI) available: Experimental details, ¹³C-NMR spectra, and ¹H-NMR spectra of CH₂Cl₂-activated UiO-67. See DOI: 10.1039/c5cc09919f



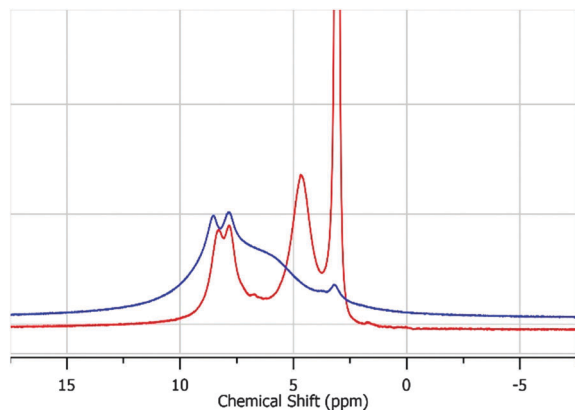


Fig. 2 (red) ^1H -NMR of UiO-67 immediately after vacuum filtration. (blue) ^1H -NMR of UiO-67 6 days after vacuum filtration. The corresponding ^{13}C -NMRs can be seen in Fig. S2 (ESI †).

remanence of methanol (3.05 ppm, 4.65 ppm) within the pores of the MOF; the remaining resonances at 7.82 and 8.29 ppm belong to the biphenyl protons (Fig. 1 H_a and H_b).

When freshly-prepared UiO-67 was left out for 6 days, the SS-NMR indicated that the majority of the methanol signals were greatly diminished.¹⁴ However, a new broad-featureless resonance upfield of the linker protons (6.29 ppm) with a concomitant broad resonance buried at 8.51 ppm was observed. This feature is indicative of the formation of an amorphous material. Concomitantly, the Brunauer–Emmett–Teller (BET) SA of the 6-day old sample was a mere $500\text{ m}^2\text{ g}^{-1}$ (Fig. S1 in the ESI †); this is in contrast with the SA of freshly-prepared and thermally activated UiO-67 which exhibited a BET SA of $2000\text{ m}^2\text{ g}^{-1}$ (Fig. S1 in the ESI †). These results are consistent with the work by DeCoste *et al.*¹¹

In order to further probe whether hydrolysis of the Zr-carboxylate bonds or capillary-force driven collapse is the culprit, we repeated the experiment with UiO-67 which was solvent exchanged (4 days) and subsequently filtered and thermally activated. As illustrated in Fig. 3, there are three

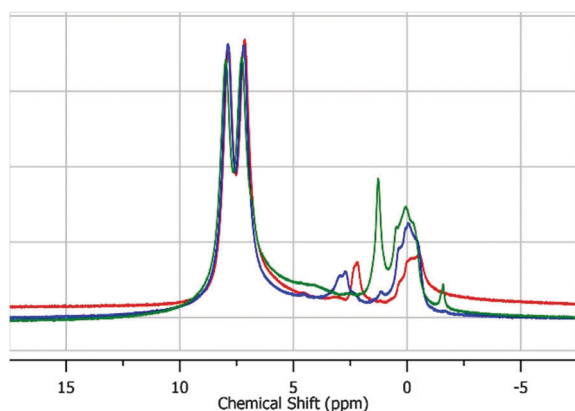


Fig. 3 (red) ^1H -NMR of UiO-67 immediately after thermal activation from methanol. (blue) ^1H -NMR of UiO-67 4 days after thermal activation. (green) ^1H -NMR of UiO-67 1 month after thermal activation. The corresponding ^{13}C -NMRs can be seen in Fig. S3 (ESI †).

distinct regions at *ca.* 0 ppm, 2.5 ppm, and 7.5 ppm; the latter two resonances have been attributed to the linker (7.5 ppm) and the bridging hydroxides (Fig. 1) on the node (2.5 ppm).¹⁵ The remaining resonance at 0 ppm, which is only slightly visible in the spectra by Dolbecq *et al.*,¹⁵ we attribute to linker deficiencies (*i.e.*, defect sites comprised of Zr-bound OH and H_2O moieties on the $\text{Zr}_6(\text{OH})_4\text{O}_4^{12+}$ node) within the porous framework.^{6a,h,16}

Unlike the filter-dry sample which contains pore-bound solvent (Fig. 2), over the course of a month, the BET SA of activated UiO-67 merely decreased to $1500\text{ m}^2\text{ g}^{-1}$ (Fig. S1 in the ESI †). The SS-NMR (Fig. 3) shows nearly no evidence for the broad featureless hump in Fig. 2 suggesting that the origin of the decrease in SA for filter-dried UiO-67 (Fig. 2) is due to capillary-force driven collapse.¹²

As a function of time however, the spectra in Fig. 3 show that the $\mu^3\text{-OH}$ resonance (2.5 ppm) shifts with a concomitant increase in the intensity of the defect-based protons at 0 ppm. The latter implies that hydrolysis occurs over time leading to an increased defect density. Thus, in addition to capillary-force-driven collapse, the process of node-hydrolysis occurs slowly over time even in properly-activated UiO-67. However, given the nominal decrease in SA over the course of a month, the MOF is clearly able to tolerate some hydrolysis of the linker-Zr bonds.

In order to probe the generality of our observations, we repeated the experiment with acetone and dichloromethane as the exchanged solvent (Fig. 4, Fig. S4 and S5, ESI †).¹⁷ Samples activated from dichloromethane (Fig. S5, ESI †) showed evidence of amorphous material and was not further examined; we hypothesize that the low miscibility of water with dichloromethane, and thus a less-efficient solvent exchange, is responsible for the degradation of the MOF.

However, when acetone was utilized (Fig. 4) the SS-NMR of the MOF initially indicated a more stable MOF with respect to hydrolysis (*i.e.*, the peaks at 0 ppm do not shift or increase in intensity). However, after a month, the MOF was found to be completely amorphous with a BET SA of $500\text{ m}^2\text{ g}^{-1}$ indicating

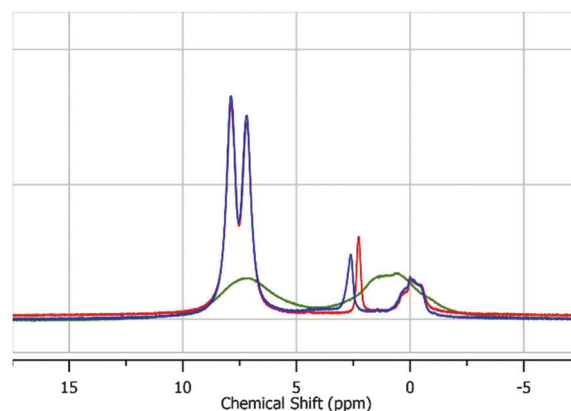


Fig. 4 (red) ^1H -NMR of UiO-67 immediately after thermal activation from acetone. (blue) ^1H -NMR of UiO-67 4 days after thermal activation. (green) ^1H -NMR of UiO-67 1 month after thermal activation. The corresponding ^{13}C -NMRs can be seen in Fig. S4 (ESI †).



Table 1 Summary of SS-NMR observations correlated to BET SA and activation method and solvent

| Activation method | Time (days) | BET SA (m ² g ⁻¹) | SS-NMR observations |
|------------------------|-------------|--|--|
| MeOH Wash | 0 | 2000 | MeOH present |
| | 6 | 500 | Crystalline and amorphous |
| Thermal from MeOH | 0 | 2000 | Defects present (0 ppm) |
| | 4 | — | μ ³ -OH shifts downfield, increased defect density |
| | 30 | 1500 | μ ³ -OH shifts upfield, increased defect density, onset of amorphous material |
| Thermal from acetone | 0 | 2000 | Defects present (0 ppm) |
| | 4 | — | μ ³ -OH shifts downfield, no change in defect density |
| | 30 | 500 | Completely amorphous |
| Thermal from DCM (ESI) | 0 | — | Amorphous material observed |

that, eventually, the MOF succumbs to hydrolysis;¹⁸ ultimately, we expect a similar fate to methanol-exchanged UiO-67.

As summarized in Table 1, SS-NMR in combination with SA measurements were used to examine the stability of UiO-67 with respect to time.¹⁹ As evident by the changing chemical shift of the μ³-OH and defect-based resonance, SS-NMR is a key tool for the understanding of the dynamic behaviour within MOFs. With respect to the stability of UiO-67, we observed that when solvent molecules remain inside the pore for a few days, then UiO-67 collapses rapidly. However, when solvent is removed at elevated temperatures, then UiO-67 remains stable for at least a month. Inevitably, hydrolysis, caused by the relative humidity, degrades the MOF beyond its structural integrity. In our hands, if kept dry or in solution, UiO-67 remains intact.

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